

1 **Electronic supplementary information**

2 **Adsorption and Oxidation of SO₂ on the Surface of TiO₂ Nanoparticles: The**
3 **Role of Terminal hydroxyl, Oxygen Vacancy-Ti³⁺ States**

4

5 Bing Wang,^a Xue Li,^{a,b} Shanshan Liang,^{a,c} Runxuan Chu,^{a,c} Dan Zhang,^{a,b} Hanqing
6 Chen,^a Meng Wang,^a Shuang Zhou,^{a,b} Wei Chen,^{a,b} Xingzhong Cao^d and Weiyue
7 Feng^{*,a}

8 ^aCAS Key Laboratory for Biomedical Effects of Nanomaterials and Nanosafety,
9 Institute of High Energy Physics, Chinese Academy of Sciences, Beijing 100049,
10 China

11 ^bUniversity of Chinese Academy of Sciences, Beijing 100049, China

12 ^cInstitute of Health Sciences, Anhui University, Hefei, Anhui 230601, China

13 ^dMuti-disciplinary Research Division, Institute of High Energy Physics, Chinese
14 Academy of Sciences, Beijing 100049, China

15 *Corresponding author: *E-mail address*: fengwy@ihep.ac.cn

16

17 **Experimental Section**

18 **Chamber experiment**

19 The TiO₂NPs was flattened in a quartz tube in the box. The gas flow rate of SO₂
20 and O₃ was controlled by mass flow meters. O₃ was generated by O₃ generator and the
21 concentration was monitored by O₃ sensor. The concentrations of SO₂ were controlled
22 at 0.1 ppm (280 μg/m³), 0.5 ppm (1400 μg/m³) and 2 ppm (5600 μg/m³). The
23 chamber RH was controlled *via* introducing humidified N₂ gas that flowed through a
24 tank filled with deionized water. The kinetics of SO₂ oxidation on TiO₂NPs at
25 different time points (10, 20, 30, 120 and 240 min) were studied. All the experiments
26 were operated under ambient pressure at room temperature.

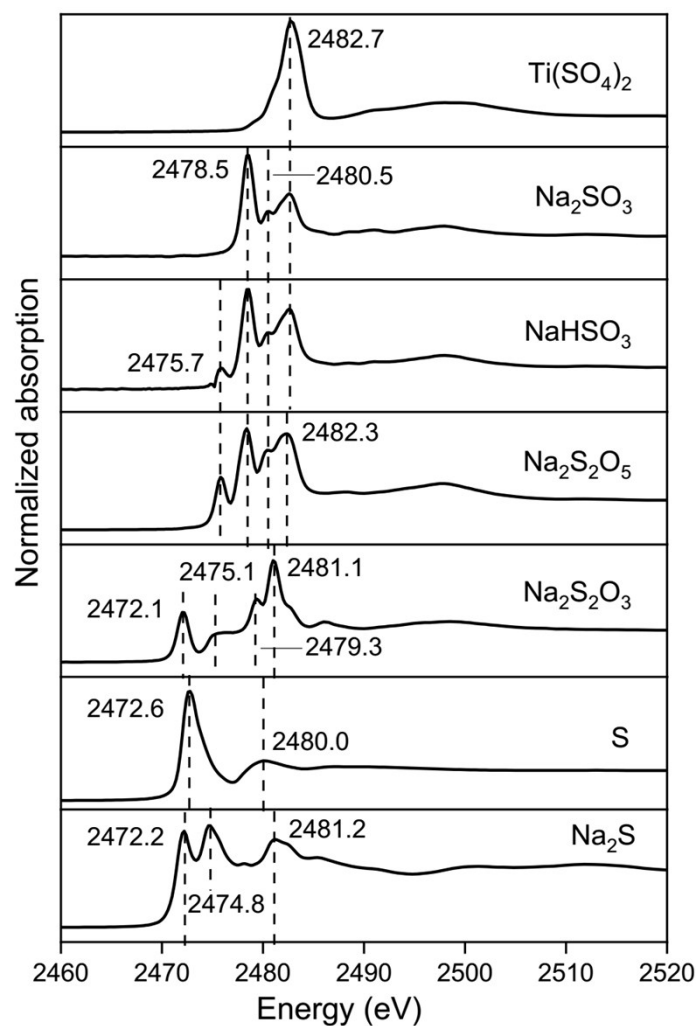
27

28 Table S1. Positron annihilation lifetimes and intensities of TiO₂NPs analyzed by LT-9 program

Samples	τ_1 (ps)	I_1 (%)	τ_2 (ps)	I_2 (%)	τ_3 (ns)	I_3 (%)
TiO ₂ NPs	214 ± 2	59.5 ± 1.3	393 ± 7	38.7 ± 1.3	2.34 ± 0.06	1.79 ± 0.07

29

30

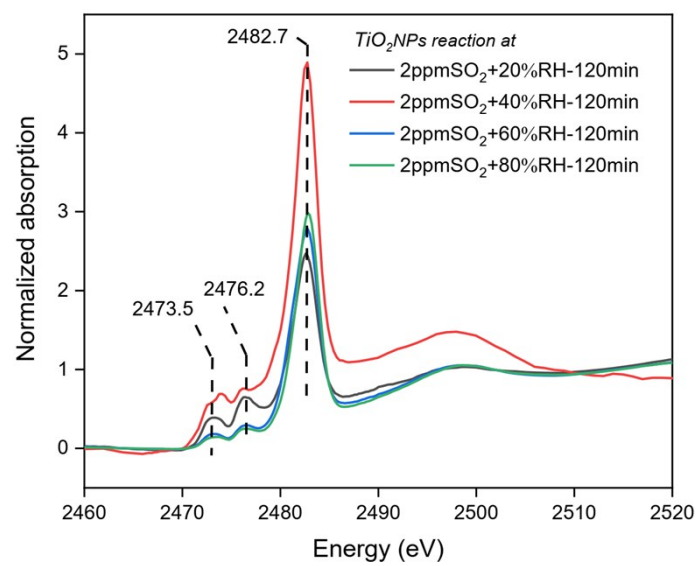


31

32

33 Fig. S1. Sulfur K-edge XANES spectra of S-containing standard references, including
 34 titanium sulfate [Ti(SO₄)₂], sodium sulfite (Na₂SO₃), sodium bisulfite (NaHSO₃),
 35 sodium metabisulfite (Na₂S₂O₅), sodium thiosulfate (Na₂S₂O₃), elemental sulfur (S)
 36 and sodium sulfide (Na₂S).

37

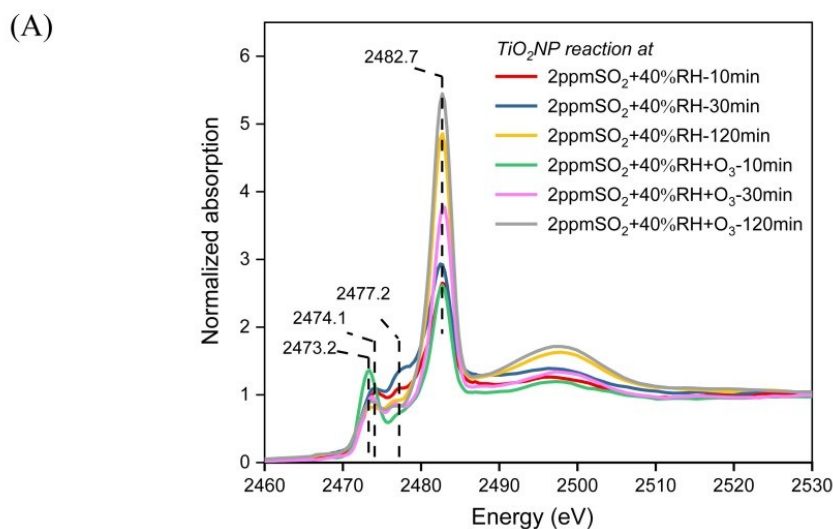


38

39 Fig. S2. Sulfur K-edge XANES measurement for the effects of relative humidity (RH)
40 on SO₂ oxidation in TiO₂NPs after 2-ppm SO₂ reaction with TiO₂NPs for 120 min.

41

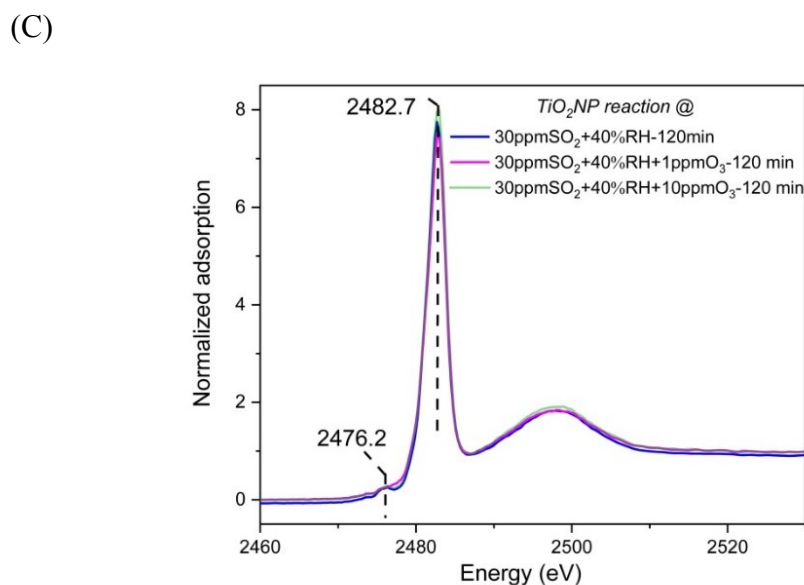
42



(B) Quantitative analysis of sulfur chemical species on TiO_2NPs when O_3 co-presence

TiO_2NP reaction at	Sulfite/Bisulfite	Sulfate
2ppm SO_2 +40%RH-10min	11.5	88.5
2ppm SO_2 +40%RH-30min	8.3	91.7
2ppm SO_2 +40%RH-120min	3.4	96.6
2ppm SO_2 +40%RH+ O_3 -10min	9.9	90.1
2ppm SO_2 +40%RH+ O_3 -30min	7.3	98.8
2ppm SO_2 +40%RH+ O_3 -120min	2.4	99.4

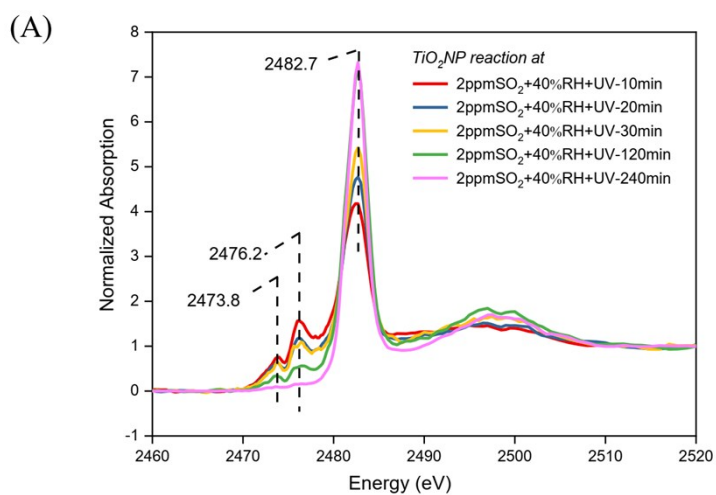
43



44

45

46 Fig. S3. Sulfur K-edge XANES measurement for SO_2 oxidation on TiO_2NPs (at 40%
47 RH) when 1 ppm and 10 ppm O_3 copresence.



(B) Quantitative analysis of S chemical species on TiO₂NPs

Reaction time	Sulfite/Bisulfite	Sulfate
10min	6.9	93.1
20min	4.4	95.6
30 min	2.8	97.2
120 min	1.1	98.9
240 min	0.1	99.9

49

50 Fig. S4. Sulfur K-edge XANES measurement for the effects of UV-irradiation time on
 51 SO₂ oxidation in TiO₂NPs after reactions (at 40% RH).

52

53

54