

Supporting Information

Influence of Ligand Field on Magnetic Anisotropy in a Family of Pentacoordinate Co^{II} Complexes

*Joydev Acharya,^a Arup Sarkar,^b Pawan Kumar,^a Vierandra Kumar,^a Jessica Flores Gonzalez,^c Olivier Cador,^c Fabrice Pointillart,^{*c} Gopalan Rajaraman,^{*b} and Vadapalli Chandrasekhar^{*a,d}*

^aDepartment of Chemistry, Indian Institute of Technology Kanpur, Kanpur-208016, India. E-mail: vc@iitk.ac.in, <http://www.iitk.ac.in>

^bDepartment of Chemistry, Indian Institute of Technology Bombay, Powai,Mumbai-400 076, India. E-mail: rajaraman@chem.iitb.ac.in

^cInstitut des Sciences Chimiques de Rennes, UMR 6226 CNRS-Université de Rennes 1,263 Avenue du Général Leclerc, 35042 Rennes Cedex, France. E-mail: fabrice.pointillart@univ-rennes1.fr

^dTata Institute of Fundamental Research, 36/P, Gopanpally Village, Serilingampally Mandal, Ranga Reddy District, Hyderabad 500107, India.

Table S1. Selected bond lengths (\AA) and angels ($^\circ$) in complex **1-3**.

Complex	Bond Length (\AA)							
	Co-N1	Co-N2	Co-N3	Co-X1 (X= Cl/Br)	Co-X2 (X= Cl/Br)	Co-N4	Co-N5	
1	2.069(2)	2.243(2)	2.069(2)	2.289(5)	2.323(5)	-	-	
2	2.063(3)	2.313(3)	2.079(3)	2.4384(6)	2.4948(6)	-	-	
3	2.038(1)	2.301(1)	2.0614(1)	-	-	1.984(2)	2.026(2)	
Bond angle($^\circ$)								
Complex 1			Complex 2			Complex 3		
Cl1—Co1—Cl2	100.478(2)		Br1—Co1—Br2	106.51(2)		N3—Co1—N2	77.34(5)	
N1—Co1—Cl1	136.06(4)		N2—Co1—Br2	95.08(7)		N1—Co1—N3	115.76(5)	
N1—Co1—Cl2	96.35(4)		N2—Co1—Br1	158.35(7)		N1—Co1—N2	76.90(5)	
N1—Co1—N2	76.25(6)		N1—Co1—Br2	99.96(8)		N4—Co1—N3	105.86(6)	
N2—Co1—Cl1	90.14(4)		N1—Co1—Br1	100.78(8)		N4—Co1—N1	129.89(6)	
N2—Co1—Cl2	169.34(4)		N1—Co1—N2	76.56(1)		N4—Co1—N2	86.69(5)	
N3—Co1—Cl1	103.45(5)		N1—Co1—N3	146.25(1)		N4—Co1—N5	100.77(6)	
N3—Co1—Cl2	99.86(5)		N3—Co1—Br2	100.01(8)		N5—Co1—N3	99.20(6)	
N3—Co1—N1	113.23(6)		N3—Co1—Br1	99.25(8)		N5—Co1—N1	98.93(6)	
N3—Co1—N2	76.59(6)		N3—Co1—N2	74.81(1)		N5—Co1—N2	172.44(6)	

Table S2. Crystal data and structure refinement parameters of complexes.

Complex	1	1_{Zn}	2	2_{Zn}	3
Formula	C ₂₃ H ₂₉ Cl ₂ CoN ₃ O	C ₂₂ H ₂₅ Cl ₂ ZnN ₃	C ₂₂ H ₂₅ Br ₂ CoN ₃	C ₂₂ H ₂₅ Br ₂ N ₃ Zn	C ₂₄ H ₂₅ CoN ₅ S ₂
g/mol	493.32	467.72	550.20	556.64	506.54
Crystal system	monoclinic	triclinic	triclinic	triclinic	triclinic
Space group	<i>P</i> 21/n	<i>P</i> -1	<i>P</i> -1	<i>P</i> -1	<i>P</i> -1
<i>a</i> /Å	16.4698(8)	8.3183(14)	8.4526(6)	8.4626(5)	8.47120(10)
<i>b</i> /Å	8.1305(4)	10.3241(18)	10.2684(8)	10.2744(6)	9.4205(2)
<i>c</i> /Å	17.6932(9)	13.067(2)	13.3275(11)	13.2348(7)	15.8080(2)
α (°)	90	98.325(5)	99.010(2)	98.374(2)	101.4480(10)
β (°)	100.979(2)	92.150(5)	90.814(2)	90.925(2)	91.1720(10)
γ (°)	90	110.353(5)	110.221(2)	109.403(2)	102.3220(10)
<i>V</i> /Å ³	2325.9(2)	1036.3(3)	1069.15(14)	1071.24(11)	1205.34(3)
<i>Z</i>	4	2	2	2	2
ρ_c /g cm ⁻³	1.409	1.499	1.709	1.726	1.396
μ /mm ⁻¹	0.987	1.456	4.553	4.889	0.907
<i>F</i> (000)	1028.0	484.0	550.0	556.0	526.0
Crystal size (mm ³)	0.02 × 0.02 × 0.02	0.02 × 0.02 × 0.02	0.02 × 0.02 × 0.02	0.02 × 0.02 × 0.02	0.36 × 0.308 × 0.179
θ range (deg)	4.69 to 56.634	4.268 to 56.768	4.292 to 56.636	4.858 to 56.602	5.438 to 57.738
Limiting indices	-21 ≤ <i>h</i> ≤ 21, -10 ≤ <i>k</i> ≤ 10, -23 ≤ <i>l</i> ≤ 23	-11 ≤ <i>h</i> ≤ 11, -13 ≤ <i>k</i> ≤ 13, -17 ≤ <i>l</i> ≤ 17	-11 ≤ <i>h</i> ≤ 11, -13 ≤ <i>k</i> ≤ 13, -17 ≤ <i>l</i> ≤ 17	-11 ≤ <i>h</i> ≤ 11, -13 ≤ <i>k</i> ≤ 13, -17 ≤ <i>l</i> ≤ 17	-11 ≤ <i>h</i> ≤ 11, -12 ≤ <i>k</i> ≤ 11, -14 ≤ <i>l</i> ≤ 21
Reflns collected	36319	15060	17222	16819	20127
Ind reflns	5773 [R _{int} = 0.0427]	5151 [R _{int} = 0.0766]	5251 [R _{int} = 0.0384]	5317 [R _{int} = 0.0653]	5549 [R _{int} = 0.0398]
Completeness to θ (%)	100	100	98	100	100
Refinement method	Full-matrix least-squares on <i>F</i> ²				
Data/restraints/parameters	5773/0/287	5151/12/250	5251/0/256	5317/0/256	5549/0/303
Goodness-of-fit on <i>F</i> ²	1.060	1.041	0.920	1.047	1.071
Final R indices [<i>I</i> > 2θ(<i>I</i>)]	<i>R</i> ₁ = 0.0335, <i>wR</i> ₂ = 0.0783	<i>R</i> ₁ = 0.0664, <i>wR</i> ₂ = 0.1808	<i>R</i> ₁ = 0.0297, <i>wR</i> ₂ = 0.0980	<i>R</i> ₁ = 0.0389, <i>wR</i> ₂ = 0.0835	<i>R</i> ₁ = 0.0329, <i>wR</i> ₂ = 0.0825
R indices (all data)	<i>R</i> ₁ = 0.0420, <i>wR</i> ₂ = 0.0831	<i>R</i> ₁ = 0.0789, <i>wR</i> ₂ = 0.1948	<i>R</i> ₁ = 0.0445, <i>wR</i> ₂ = 0.1249	<i>R</i> ₁ = 0.0552, <i>wR</i> ₂ = 0.0993	<i>R</i> ₁ = 0.0381, <i>wR</i> ₂ = 0.0857

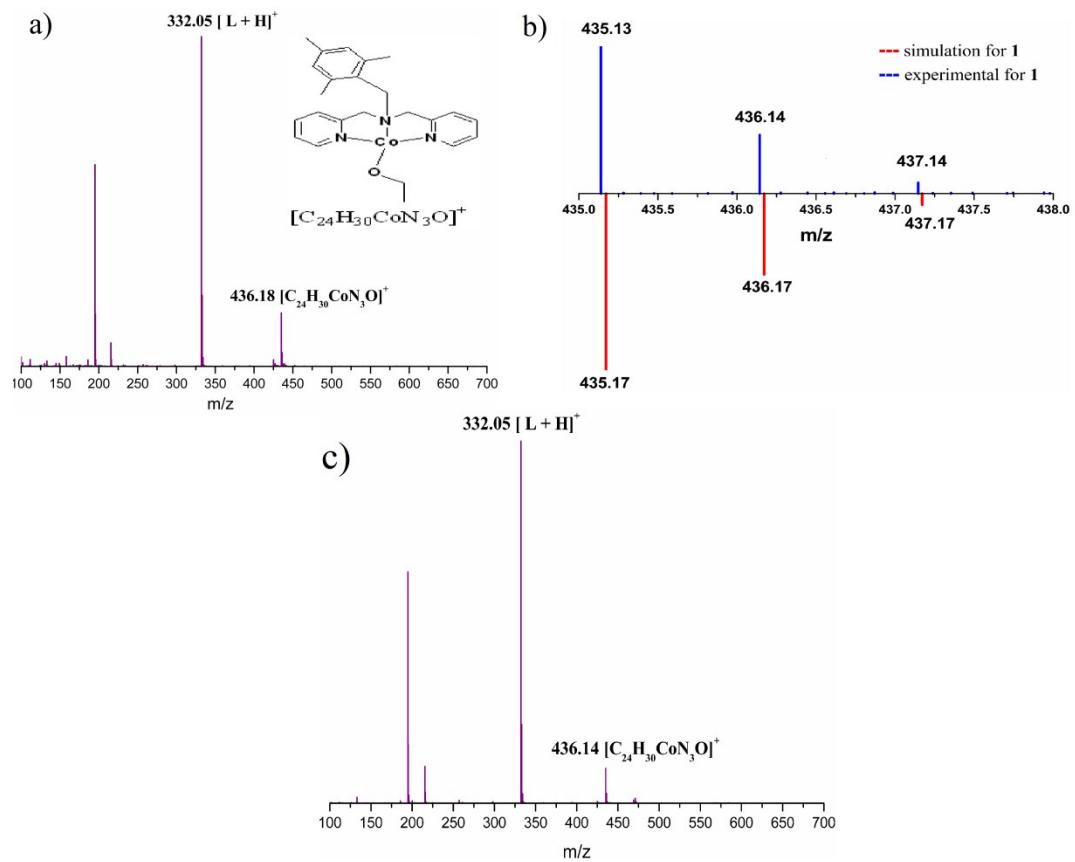


Figure S1. a) Full range ESI-MS spectra of **1** and structure of $[C_{24}H_{30}CoN_3O]^+$ (inset), b) isotopic distribution pattern of the species combining experimental and simulation pattern, c) similar full range spectra of **2**.

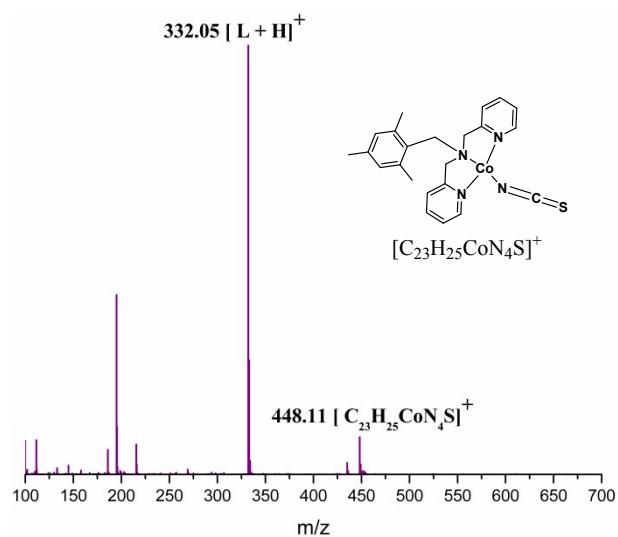


Figure S2. a) Full range ESI-MS spectra of **3** and structure of $[C_{22}H_{25}CoN_4S]^+$ (inset).

S H A P E v2.1 Continuous Shape Measures calculation

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Contact: llunell@ub.edu

Sl. No.	Geometry	CShM value for 1	CShM value for 2	CShM value for 3
1	Pentagon(D_{5h})	30.704	33.010	31.689
2	Vacant octahedron(C_{4v})	4.967	3.367	5.604
3	Trigonal bipyramidal(D_{3h})	2.371	6.452	1.238
4	Spherical square pyramid(C_{4v})	4.992	1.741	4.956
5	Johnson trigonal bipyramidal J12(D_{3h})	3.897	7.601	2.778

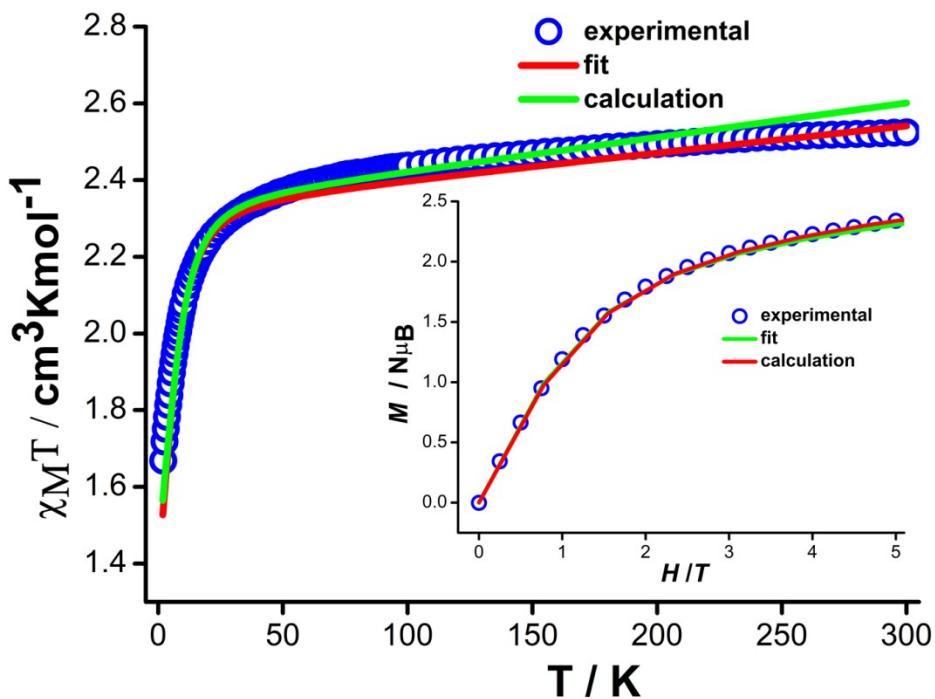


Figure S3. Temperature-dependent dc magnetic susceptibility plots of **3**, respectively; (inset) Field-dependent magnetization plots for **3** respectively at 2K. The solid red lines in all the panels represent the best fits obtained after simultaneous fitting of the magnetic data taking negative D value of using PHI. The solid green lines in all the panels represent simulations of

the experimental magnetic data considering negative D value of the computed SH parameters obtained from CASSCF/NEVPT2 calculation.

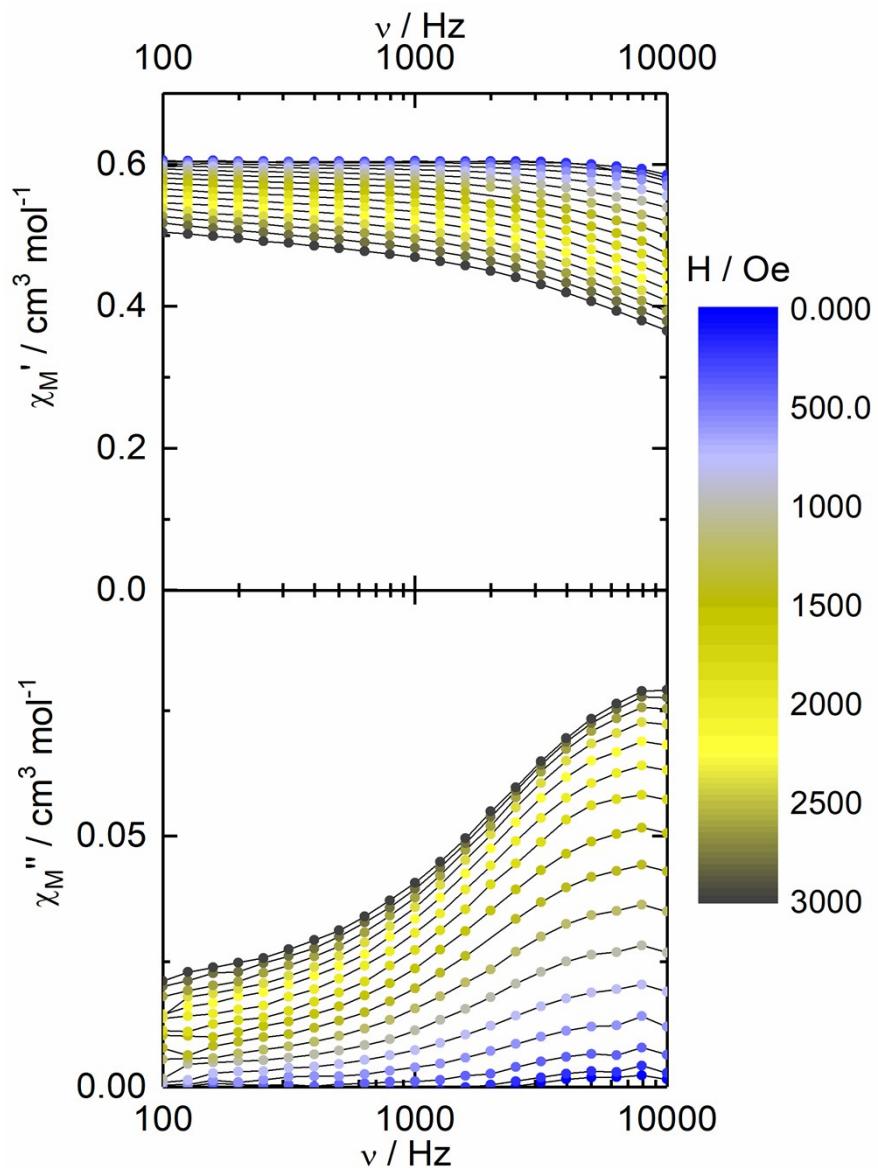


Figure S4. Frequency dependences of the in-phase, χ_M' , and out-of-phase, χ_M'' , components of the AC susceptibility measured at 2 K at various external DC fields for complex **1**.

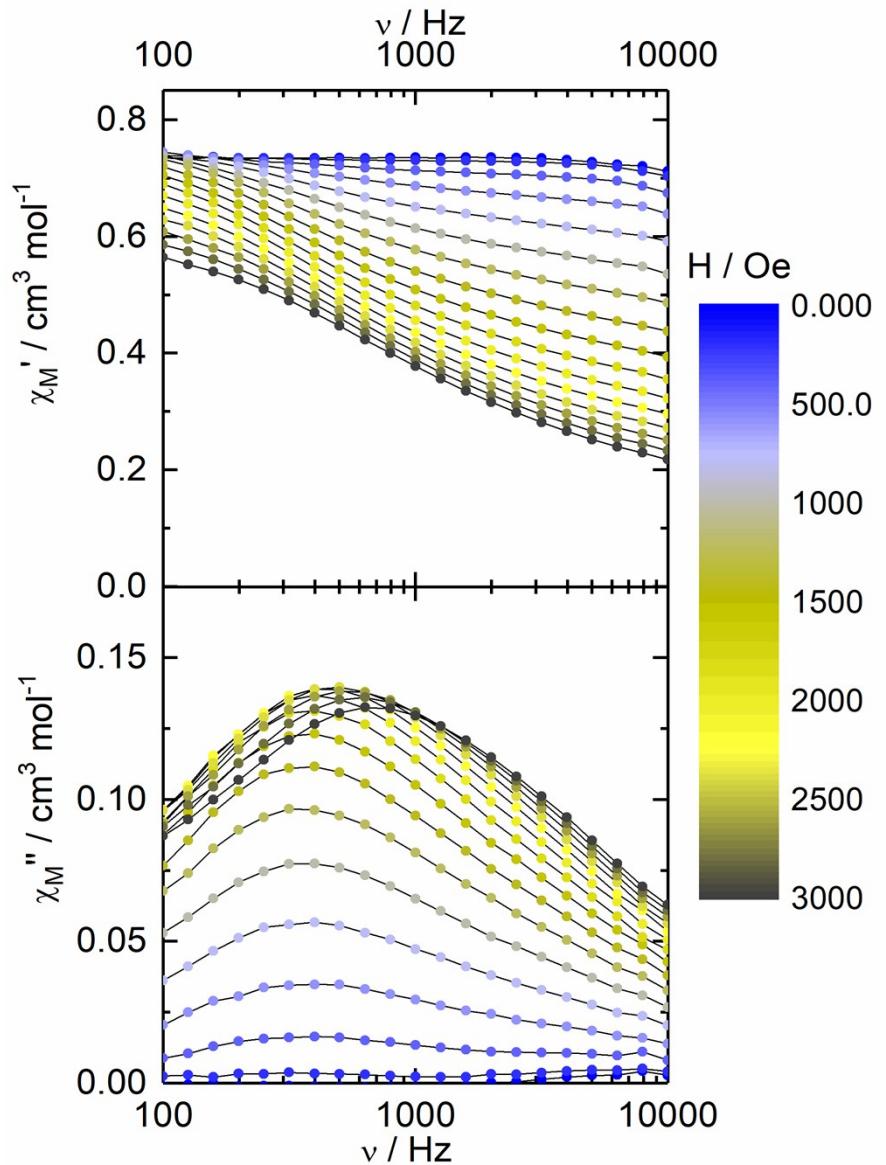


Figure S5. Frequency dependences of the in-phase, χ_M' , and out-of-phase, χ_M'' , components of the AC susceptibility measured at 2 K at various external DC fields for complex **2**.

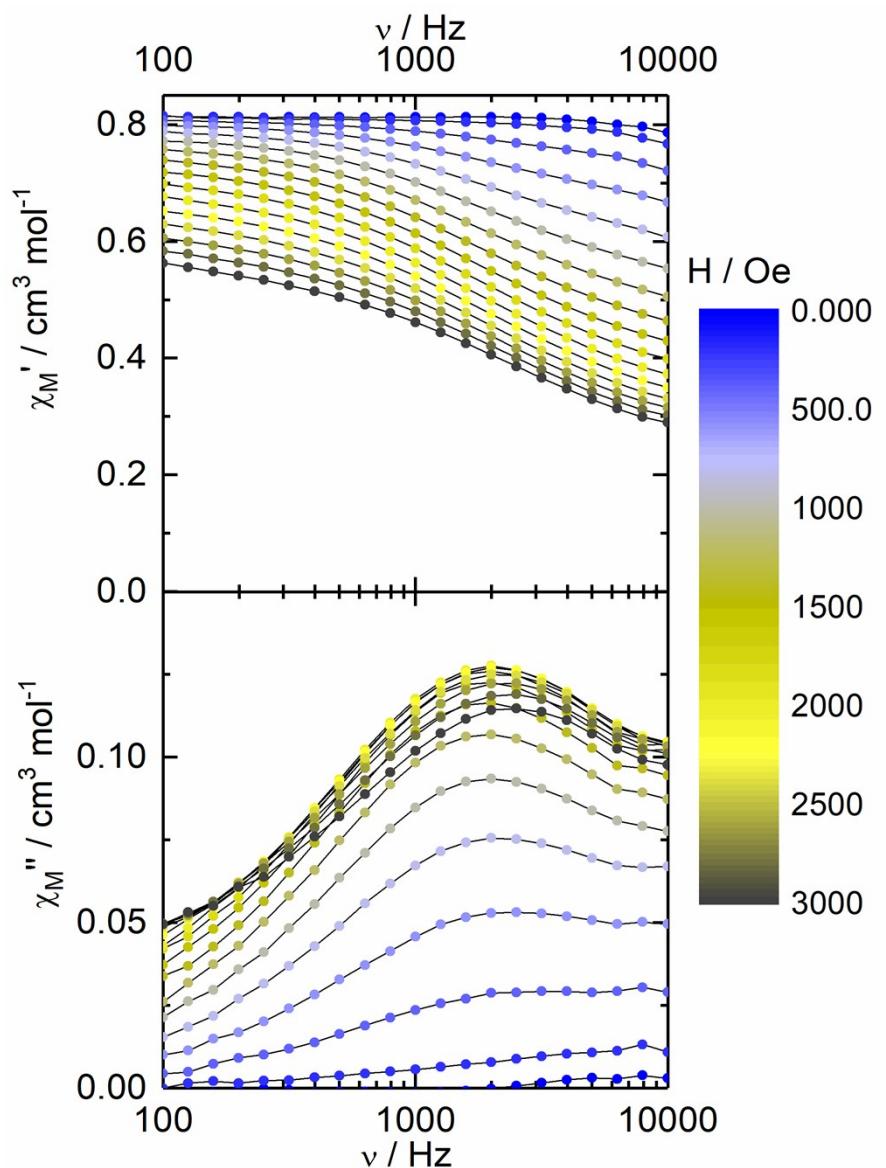


Figure S6. Frequency dependences of the in-phase, χ_M' , and out-of-phase, χ_M'' , components of the AC susceptibility measured at 2 K at various external DC fields for complex 3.

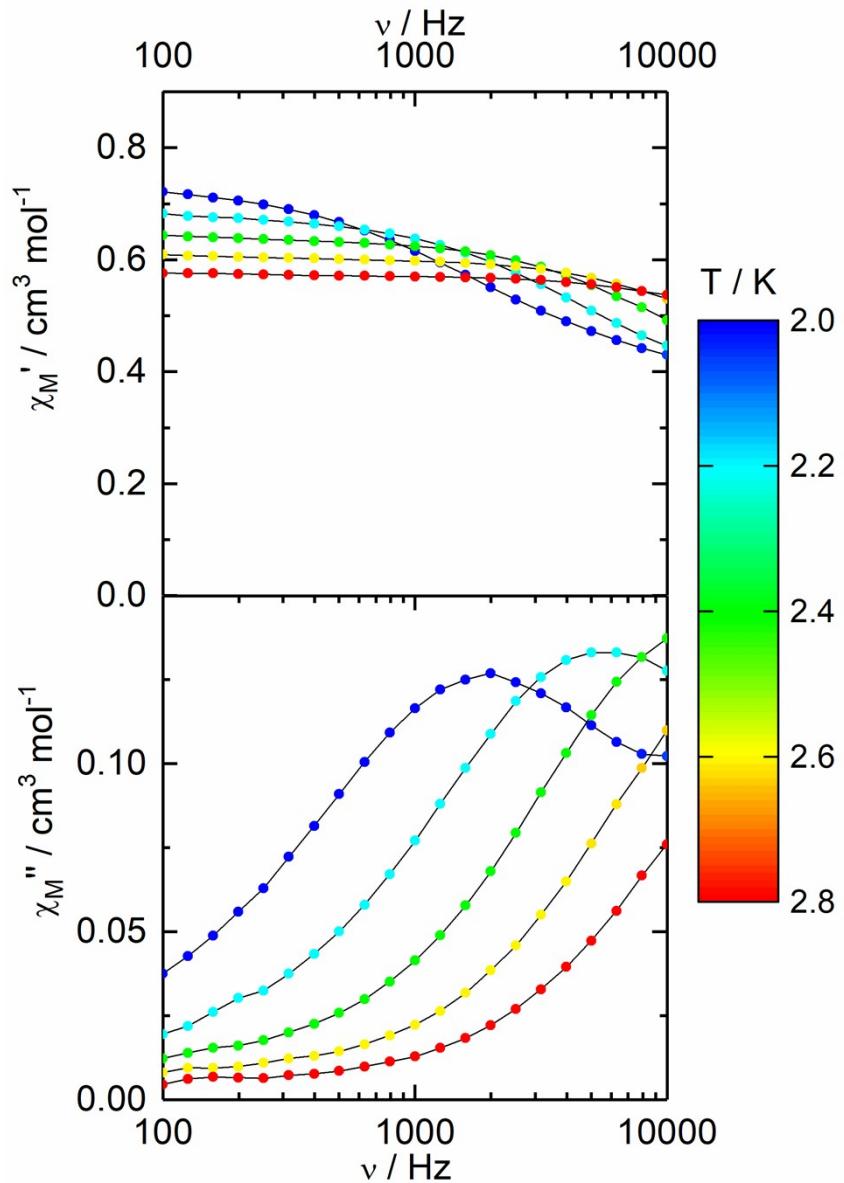


Figure S7. Temperature variations of χ_M' and χ_M'' for complex **3** as a function of the frequency, ν , of the oscillating AC field measured at an external DC field of 1600 Oe.

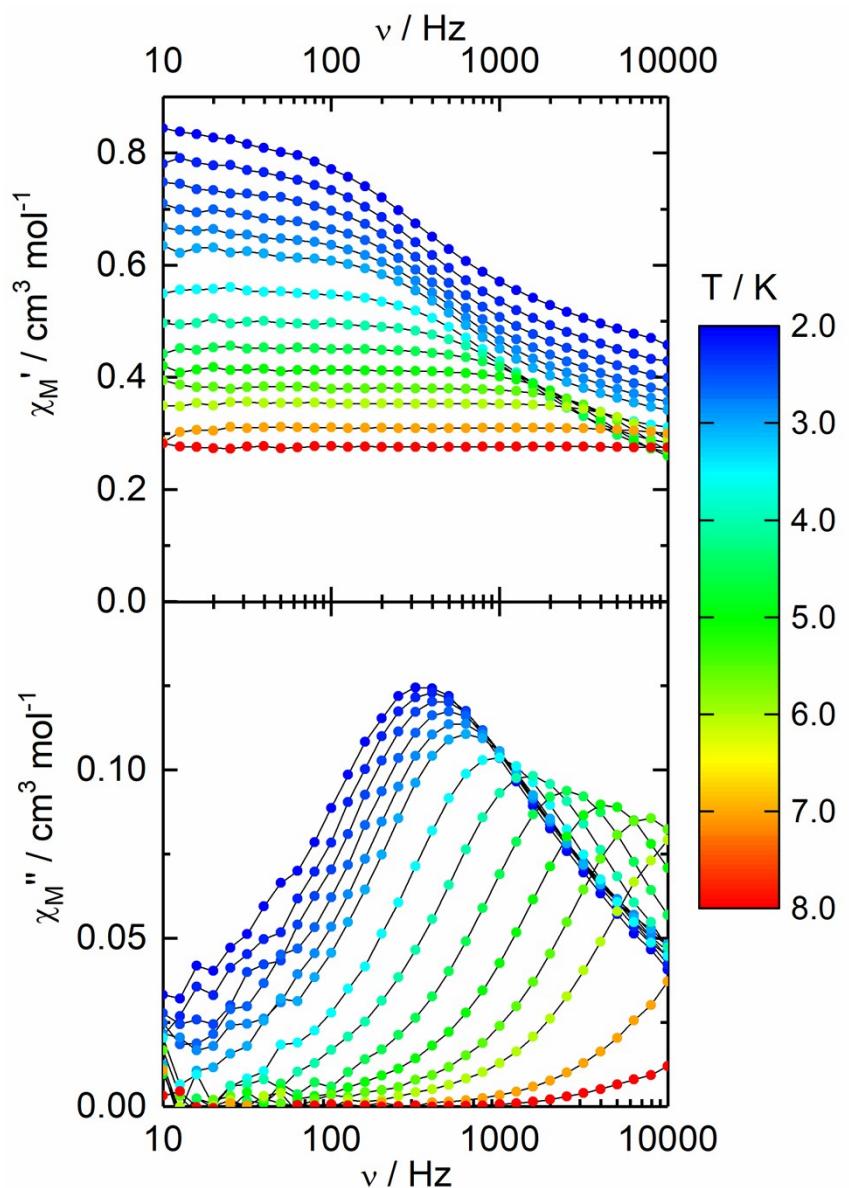


Figure S8. Temperature variations of χ_M' and χ_M'' for compound **2** as a function of the frequency, ν of the oscillating AC field measured at an external DC field of 1400 Oe.

Extended Debye model

$$\chi_M' = \chi_s + (\chi_t - \chi_s) \frac{1 + (\omega\tau)^{1-\alpha} \sin\left(\alpha \frac{\pi}{2}\right)}{1 + 2(\omega\tau)^{1-\alpha} \sin\left(\alpha \frac{\pi}{2}\right) + (\omega\tau)^{2-2\alpha}}$$

$$\chi_M'' = (\chi_t - \chi_s) \frac{(\omega\tau)^{1-\alpha} \cos\left(\alpha \frac{\pi}{2}\right)}{1 + 2(\omega\tau)^{1-\alpha} \sin\left(\alpha \frac{\pi}{2}\right) + (\omega\tau)^{2-2\alpha}}$$

With χ_t the isothermal susceptibility, χ_s the adiabatic susceptibility, τ the relaxation time and α an empiric parameter which describe the distribution of the relaxation time. For SMM with only one relaxing object α is close to zero. The extended Debye model was applied to fit simultaneously the experimental variations of χ_M' and χ_M'' with the frequency ν of the oscillating field ($\omega = 2\pi\nu$). Typically, only the temperatures for which a maximum on the χ_M'' vs. ν curves, have been considered (see figure here below for an example). The best fitted parameters τ , α , χ_t , χ_s are listed in table S3 with the coefficient of determination R^2 .

Table S3. Best fitted parameters (χ_t , χ_s , τ and α) with the extended Debye model for compound 2 at 1400 Oe in the temperature range 2-5.5 K.

T / K	$\chi_s / \text{cm}^3 \text{ mol}^{-1}$	$\chi_t / \text{cm}^3 \text{ mol}^{-1}$	τ / s	α	R^2
2	0.44267	0.85998	3.8616E-4	0.32528	0.99987
2.2	0.41012	0.80184	3.36478E-4	0.2961	0.99983
2.4	0.38103	0.75372	2.95939E-4	0.28018	0.99976
2.6	0.35735	0.71094	2.62713E-4	0.26686	0.99974
2.8	0.3397	0.67078	2.32467E-4	0.24406	0.99968
3	0.32621	0.6341	2.06276E-4	0.21467	0.99964
3.5	0.29652	0.55891	1.48711E-4	0.15057	0.99968
4	0.26764	0.50005	9.71502E-5	0.10368	0.99978
4.5	0.24361	0.45239	6.0143E-5	0.06404	0.99981

5	0.22193	0.41417	3.62249E-5	0.04173	0.99985
5.5	0.20208	0.38289	2.14462E-5	0.03285	0.9997

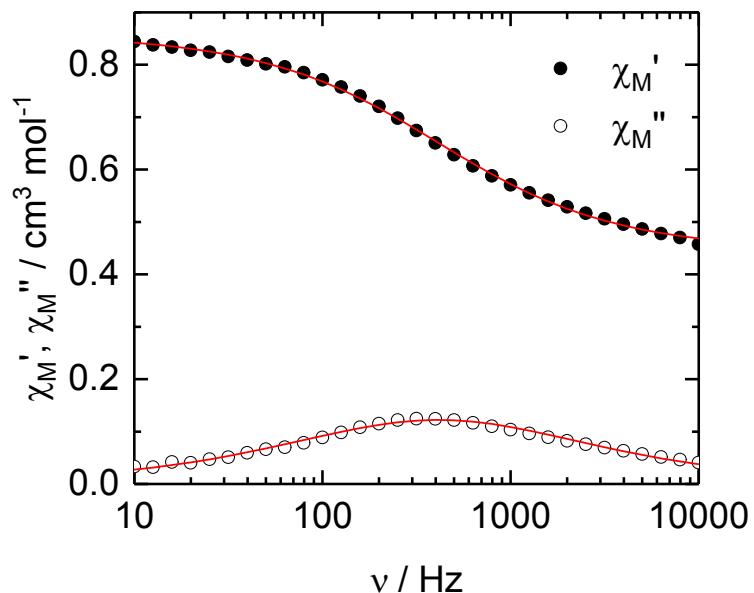


Figure S9. Frequency dependence of the in-phase (χ_M') and out-of-phase (χ_M'') components of the AC susceptibility measured on powder at 2 K at 1400 Oe with the best fitted curves (red lines) for **2**.

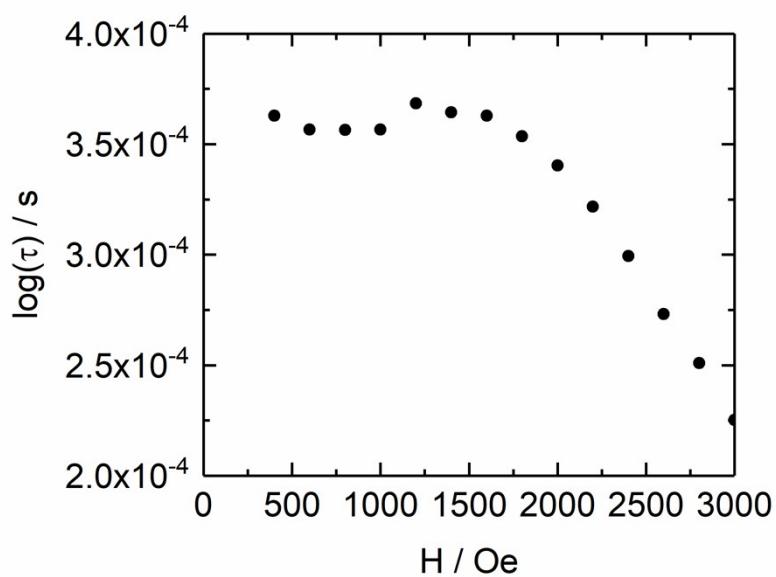


Figure S10. Field dependence of the magnetic relaxation time for **2** at 2 K.

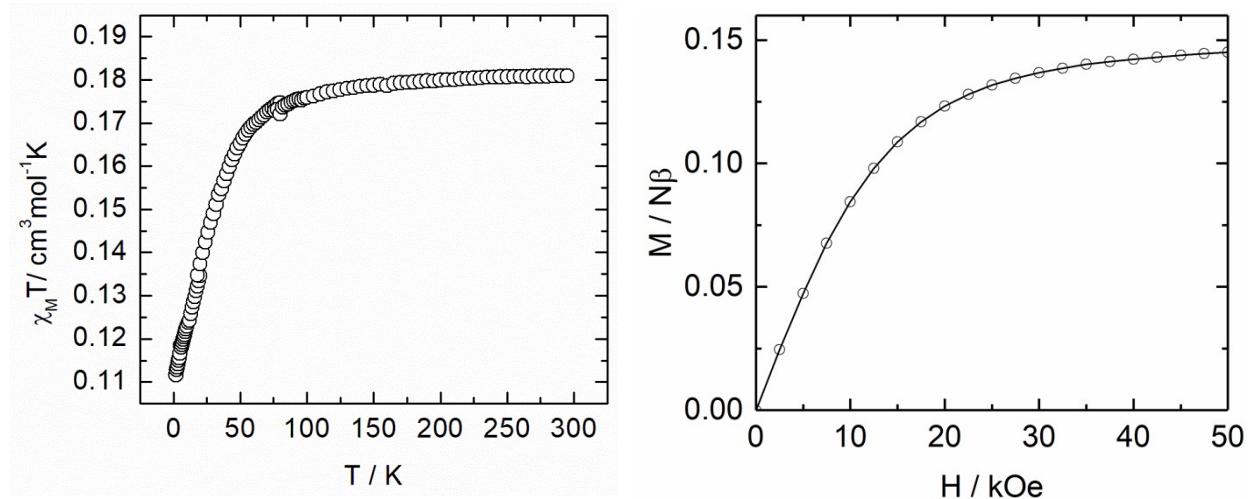


Figure S11. Temperature-dependent magnetic susceptibility and magnetization curve at 2 K of **2₁₀**.

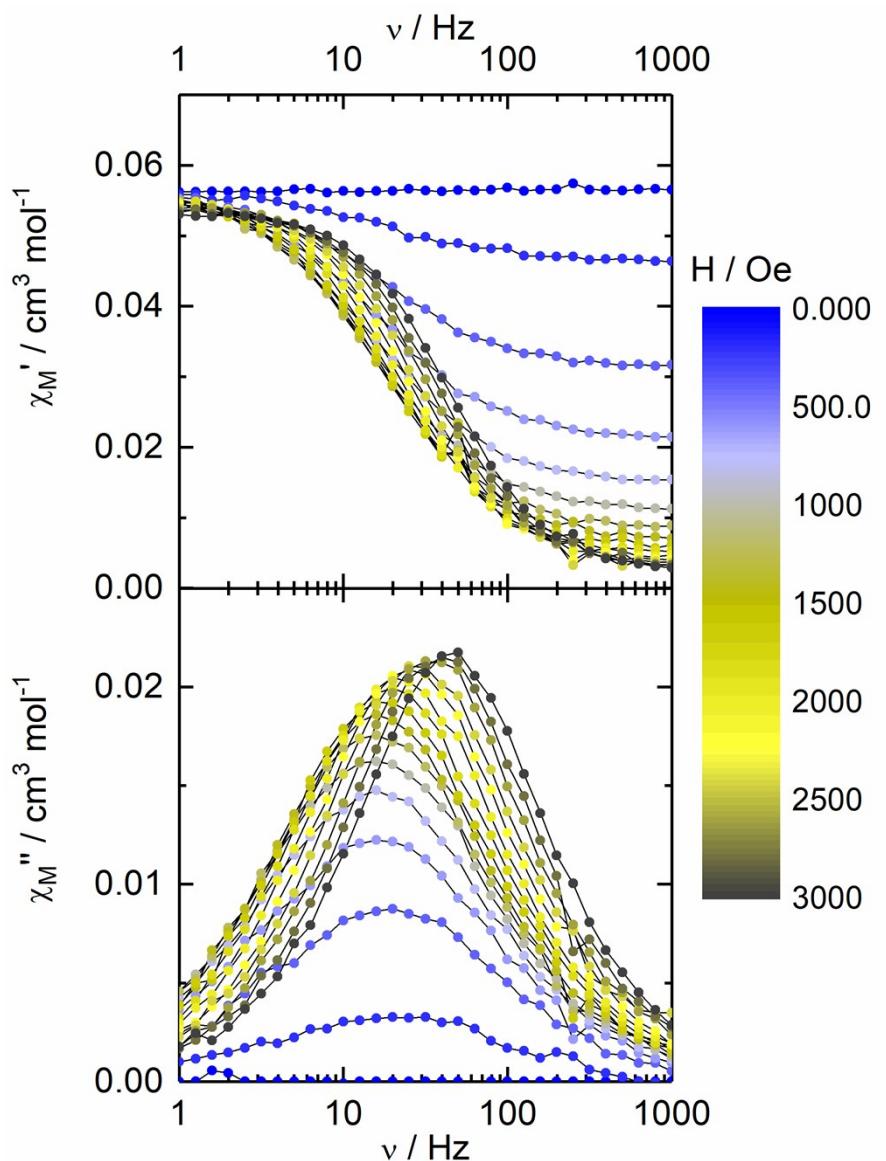


Figure S12. Frequency dependences of the in-phase, χ_M' , and out-of-phase, χ_M'' , components of the AC susceptibility measured at 2 K at various external DC fields for compound **2₁₀**.

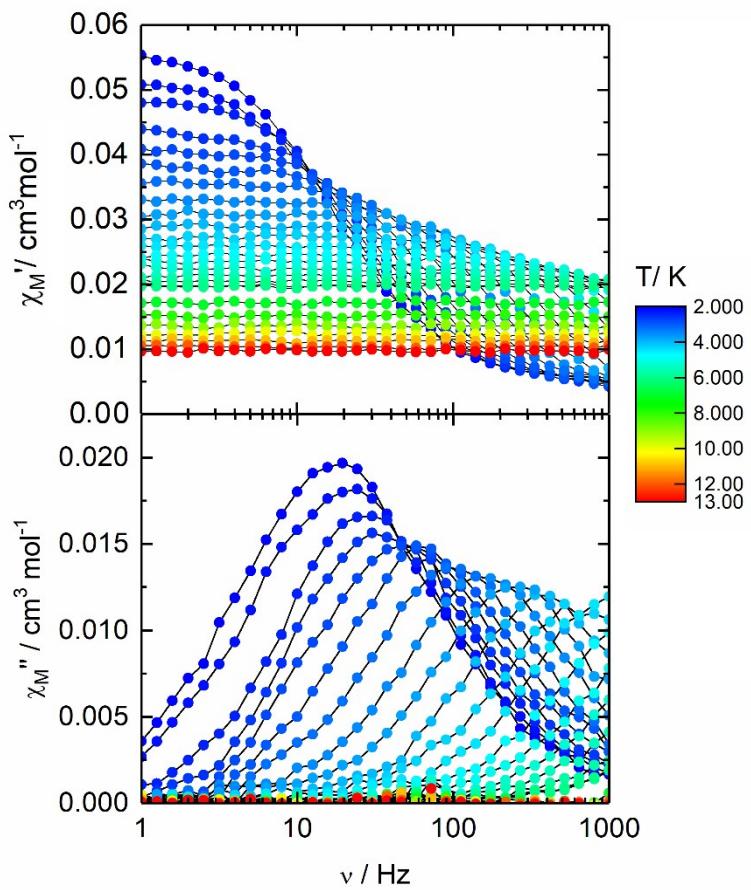


Figure S13. Temperature variations of χ_M' and χ_M'' for compound **2₁₀** as a function of the frequency, ν of the oscillating AC field measured at an external DC field of 1600 Oe.

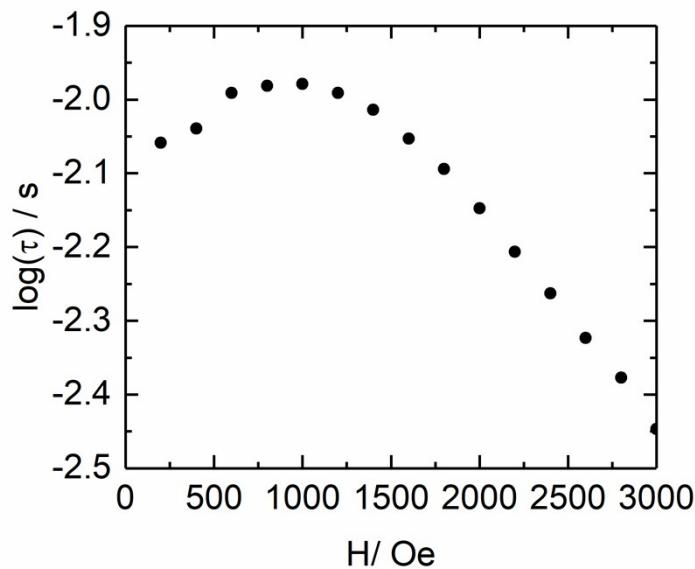


Figure S14. Field dependence of the magnetic relaxation time for **2₁₀** at 2 K.

Table S4. Best fitted parameters (χ_T , χ_S , τ and α) with the extended Debye model for compound **2₁₀** at 2 K between 0 and 3000 Oe.

H / Oe	χ_S / cm ³ mol ⁻¹	χ_T / cm ³ mol ⁻¹	τ / s	α	R ²
0	NA	0.05646 (average)	NA	NA	NA
200	0.04645	0.05614	0.00874	0.21887	0.9999
400	0.03109	0.05608	0.00914	0.21435	0.99981
600	0.02096	0.05648	0.0102	0.22033	0.99967
800	0.01454	0.05645	0.01044	0.21779	0.99982
1000	0.01054	0.05668	0.01049	0.20911	0.99976
1200	0.00808	0.05672	0.01021	0.19497	0.99945
1400	0.0066	0.05677	0.00968	0.18862	0.99916
1600	0.00553	0.05647	0.00885	0.18074	0.99611
1800	0.00467	0.05629	0.00805	0.16011	0.99923
2000	0.00396	0.05554	0.00712	0.14028	0.99977
2200	0.00325	0.05524	0.00621	0.12959	0.99947
2400	0.00266	0.05495	0.00546	0.13244	0.99935
2600	0.0028	0.05454	0.00475	0.11715	0.9998
2800	0.00286	0.05389	0.00419	0.10753	0.99958

3000	0.00233	0.05359	0.00358	0.11021	0.99975
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Table S5. Best fitted parameters (χ_T , χ_S , τ and α) with the extended Debye model for compound **2₁₀** at 1600 Oe in the temperature range 2-4.25 K.

T / K	χ_S / cm ³ mol ⁻¹	χ_T / cm ³ mol ⁻¹	τ / s	α	R ²
2	0.0052	0.05676	0.00903	0.16629	0.99956
2.2	0.00495	0.05221	0.00749	0.15998	0.99963
2.4	0.00496	0.04786	0.00546	0.13026	0.99711
2.6	0.0047	0.04354	0.00416	0.11772	0.99808
2.8	0.00418	0.04087	0.00311	0.11425	0.99816
3	0.00423	0.03822	0.00224	0.09409	0.99881
3.25	0.0044	0.03568	0.00153	0.07565	0.99962
3.5	0.00407	0.033	9.64E-04	0.05839	0.99973
3.75	0.00404	0.03088	6.37E-04	0.02949	0.99944
4	0.00428	0.02894	4.24E-04	0.00969	0.99947
4.25	0.00347	0.02729	2.59E-04	0.02786	0.99949

Table S6. Literature reports of five coordinated Co^{II} complexes and their Spin Hamiltonian parameters. We can include this Table after we get the SH parameters

Compound	Geometry	D (cm ⁻¹)	E (cm ⁻¹)	g	U_{eff} (cm ⁻¹)	Reference
[Co(NS ₃ ^{iPr})Cl]BPh ₄	TBPY	-19.1	1.5	2.43	32	S1
[Co(NS ₃ ^{tBu})Cl] ClO ₄	TBPY	-21.4	-	2.29		S2
[Co(NS ₃ ^{tBu})Br] ClO ₄	TBPY	-20.2	-	2.29	21	
[Co(NS ₃ ^{tBu})NCS] ClO ₄	TBPY	-11.0	-	2.29	20	
[Co(TPMA)(CH ₃ CN)] (BF ₄) ₂ ·CH ₃ CN	TBPY	9.6	0.2	2.29		S3
[Co(TPMA)Cl]Cl.2·4(H ₂ O)	TBPY	-6.9	-1.78	2.35	-	
[Co(TPMA)Cl]Cl	TBPY	-8.4	0	2.30		
[Co(TPMA)Br]Br.2·0(H ₂ O)	TBPY	-6.3	1.6	2.34		
[Co(TPMA)Br]Br	TBPY	-7.1	0	-		

[Co(TPMA)I]I	TBPY	-7.5	1			
[Co(PP ₃)Cl]ClO ₄	SPY	46.4	10.1	2.31	37.8	S4
[Co(PP ₃)Br]ClO ₄	SPY	40.7	9.3	2.28	34.5	
[Co ^{II} (NNN ^{-pyr})Cl ₂]	SPY	150	11.6	$g_z=2.00$; $g_{xy}=3.28$	-	S5
[Co ^{II} (NNN ^{-ip-Me})SCN ₂]	SPY	-28.8	-	-	11.1	S6
[Co ^{II} (NNN ^{-ip-ph})SCN ₂]	SPY	-28.8	-	-	16.7	
[Co(Me ₄ cyclam)N ₃] ⁺	SPY	30	9.8	$g_z=2.03$; $g_{xy}=2.35$	2.3	S7
[Co ^{II} (NNN ^{-bim})Cl]	SPY	14.5	0	$g_x=2.41$; $g_y=2.25$; $g_z=2.01$	19.6	S8
[Co ^{II} (NNN ^{-bim})Br]	SPY	8.4	0	$g_x=2.3$; $g_y=2.2$; $g_z=1.99$	8.2	
[Co(tpa)Cl]ClO ₄	SPY	-10.1	1.8	-	12	S9
[Co(tpa)Br]ClO ₄	SPY	-7.8	2.1	-	8.7	
[Co(tbta)Cl](ClO ₄) (MeCN) ₂ (H ₂ O)	SPY	-7.5	0.4	-	8.1	
[Co(tbta)Br]ClO ₄	SPY	-4.3	0.03	-	5	
[Co(N ₃)LH ₃] ⁴⁺	SPY	-7.1	-	$g_z=2.24$; $g_{xy}=2.36$	14.2	S10
[Co(terpy)Cl ₂]	SPY	-	-	$g_x=1.35$; $g_y=1.93$; $g_z=7.75$	19.5 2.8	S11
[Co(terpy)(NCS) ₂]	SPY	-	-	-	11.8 2.1	
[Co(bzimpy)Cl ₂]	SPY	71.7	1.4	$g_x=2.5$; $g_y=2.62$; $g_z=1.47$	-	S12
[CoL ¹ Cl ₂]	SPY	71.7	0	$g_z=2.0$; $g_{xy}=2.5$	-	S13
[CoL ² Cl ₂]	SPY	46.8	0	$g_z=2.0$; $g_{xy}=2.35$	-	
[Co(L)Cl ₂]·CH ₃ OH]	TBPY					This work
[Co(L)Br ₂]	SPY					
[Co(L)(NCS)2]	TBPY					

Me₆tren: hexamethyltris(aminoethyl)amine; NS₃^{iPr}: tris-(2-(isopropylthio)ethyl)amine;
 NS₃^{tBu}: tris-(2-(tertbutyl)thio)ethyl)amine; Me₄cyclam: tetramethylcyclam, N₃: azido; PP₃:
 tris[2-(diphenylphosphino)ethyl]-phosphine; NNN^{-pyr}: 4-hept-1-ynyl-2,6-dipyrazol-1-

ylpyridine; (NNN^{-ip}) : bis(imino)pyridine; terpy: terpyridine; bzimpy: 2,6-bis(benzimidazol-2-yl)pyridine; L¹: 4'-ido-2',6'-dipyrazolyl-pyridine; L²: 4'-dodecynyl-2',6'-dipyrazolyl pyridine; LH₃:6,16,2,5-tribenzena(1,4)-1,4,8,11,14,18,23,27-octaaazabicyclo[9.9.9]nonacosaphane; TBP: Trigonal Bipyramidal; SqPy: Square Pyramid

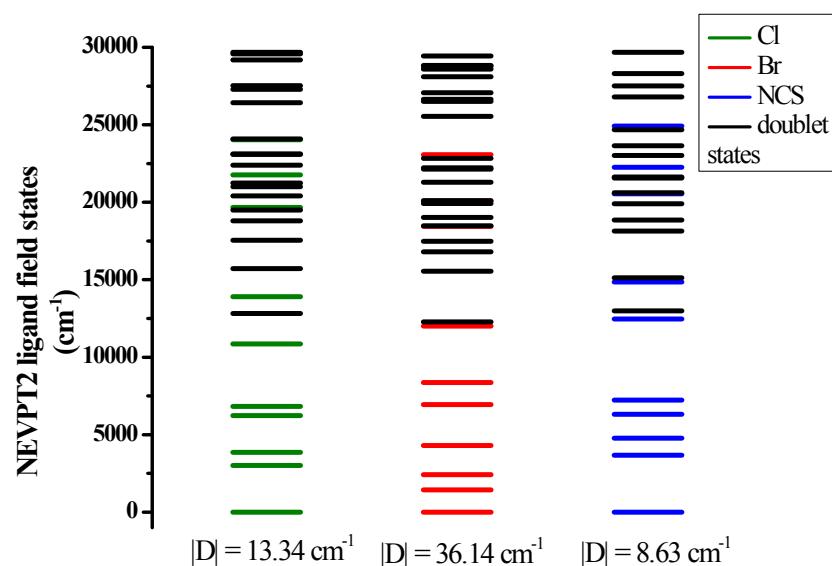


Figure S15. NEVPT2 transition energies of quartet and doublet states (up to 30000 cm⁻¹ shown) of the three complexes.

Table S7 . NEVPT2 transition energies, their corresponding wavefunction and their individual contributions towards D and E value for the ground state and 1st four excited states for complex 1.

Spin free energy states of Complex 1	Major CASSCF electronic configuration	NEVPT2 transition energy (cm ⁻¹)	Contribution to D (cm ⁻¹)	Contribution to E (cm ⁻¹)
GS	d _{xz} ² d _{xy} ² d _{yz} ¹ d _z ²¹ d _{x^{2-y}} ²¹ (51%)	0.0	0.0	0.0

	$d_{xz}^1 d_{xy}^2 d_{yz}^1 d_{z^2}^2 d_{x^2-y^2}^1$ (24%) $d_{xz}^2 d_{xy}^1 d_{yz}^1 d_{z^2}^2 d_{x^2-y^2}^1$ (16%)			
1 st ES	$d_{xz}^1 d_{xy}^2 d_{yz}^2 d_{z^2}^1 d_{x^2-y^2}^1$ (35%) $d_{xz}^1 d_{xy}^2 d_{yz}^1 d_{z^2}^2 d_{x^2-y^2}^1$ (27%) $d_{xz}^2 d_{xy}^2 d_{yz}^1 d_{z^2}^1 d_{x^2-y^2}^1$ (14%)	3017.3	14.2	12.2
2 nd ES	$d_{xz}^2 d_{xy}^1 d_{yz}^2 d_{z^2}^1 d_{x^2-y^2}^1$ (71%)	3858.1	9.7	-9.0
3 rd ES	$d_{xz}^1 d_{xy}^2 d_{yz}^2 d_{z^2}^1 d_{x^2-y^2}^1$ (42%) $d_{xz}^1 d_{xy}^2 d_{yz}^1 d_{z^2}^2 d_{x^2-y^2}^1$ (35%)	6236.3	1.0	0.0
4 th ES	$d_{xz}^2 d_{xy}^1 d_{yz}^1 d_{z^2}^1 d_{x^2-y^2}^2$ (40%) $d_{xz}^1 d_{xy}^1 d_{yz}^1 d_{z^2}^2 d_{x^2-y^2}^2$ (19%)	6823.1	-12.1	-0.4

Table S8. NEVPT2 transition energies, their corresponding wavefunction and their individual contributions towards D and E value for the ground state and 1st five excited states for complex 2.

Spin free energy states of Complex 2	Major CASSCF electronic configuration	NEVPT2 transition energy (cm ⁻¹)	Contribution to D (cm ⁻¹)	Contribution to E (cm ⁻¹)
GS	$d_{xy}^2 d_{z^2}^2 d_{yz}^1 d_{xz}^1 d_{x^2-y^2}^1$ (57%) $d_{xy}^1 d_{z^2}^2 d_{yz}^1 d_{xz}^1 d_{x^2-y^2}^2$ (19%)	0.0	0.0	0.0
1 st ES	$d_{xy}^2 d_{z^2}^1 d_{yz}^2 d_{xz}^1 d_{x^2-y^2}^1$ (45%) $d_{xy}^1 d_{z^2}^2 d_{yz}^1 d_{xz}^2 d_{x^2-y^2}^1$ (15%)	1439.4	22.7	-20.2
2 nd ES	$d_{xy}^1 d_{z^2}^2 d_{yz}^2 d_{xz}^1 d_{x^2-y^2}^1$ (22%) $d_{xy}^2 d_{z^2}^1 d_{yz}^2 d_{xz}^1 d_{x^2-y^2}^1$ (18%) $d_{xy}^1 d_{z^2}^1 d_{yz}^2 d_{xz}^2 d_{x^2-y^2}^1$ (16%)	2423.4	29.2	27.4
3 rd ES	$d_{xy}^1 d_{z^2}^1 d_{yz}^2 d_{xz}^2 d_{x^2-y^2}^1$ (48%) $d_{xy}^2 d_{z^2}^1 d_{yz}^1 d_{xz}^2 d_{x^2-y^2}^1$ (20%)	4297.3	0.2	0.2
4 th ES	$d_{xy}^1 d_{z^2}^1 d_{yz}^2 d_{xz}^1 d_{x^2-y^2}^2$ (35%) $d_{xy}^1 d_{z^2}^1 d_{yz}^2 d_{xz}^2 d_{x^2-y^2}^1$ (21%)	6947.5	0.5	0.5
5 th ES	$d_{xy}^1 d_{z^2}^2 d_{yz}^1 d_{xz}^1 d_{x^2-y^2}^2$ (51%) $d_{xy}^2 d_{z^2}^2 d_{yz}^1 d_{xz}^1 d_{x^2-y^2}^1$ (17%)	8363.7	-11.9	-0.1

Table S9. NEVPT2 transition energies, their corresponding wave function and their individual contributions towards D and E value for the ground state and 1st four excited states for complex 3.

Spin free energy states of Complex 3	Major CASSCF electronic configuration	NEVPT2 transition energy (cm ⁻¹)	Contribution to D (cm ⁻¹)	Contribution to E (cm ⁻¹)

GS	$d_{xy}^2 d_{xz}^2 d_{yz}^1 d_{x^2-y^2}^1 d_z^2$ (31%) $d_{xy}^2 d_{xz}^1 d_{yz}^1 d_{x^2-y^2}^1 d_z^2$ (22%)	0.0	0.0	0.0
1 st ES	$d_{xy}^1 d_{xz}^2 d_{yz}^2 d_{x^2-y^2}^1 d_z^2$ (42%) $d_{xy}^1 d_{xz}^1 d_{yz}^1 d_{x^2-y^2}^2 d_z^2$ (18%)	3674.9	-0.3	-7.8
2 nd ES	$d_{xy}^1 d_{xz}^1 d_{yz}^2 d_{x^2-y^2}^2 d_z^2$ (47%) $d_{xy}^2 d_{xz}^1 d_{yz}^2 d_{x^2-y^2}^1 d_z^2$ (31%)	4781.4	-9.3	-2.7
3 rd ES	$d_{xy}^1 d_{xz}^1 d_{yz}^1 d_{x^2-y^2}^2 d_z^2$ (22%) $d_{xy}^1 d_{xz}^2 d_{yz}^1 d_{x^2-y^2}^2 d_z^2$ (19%) $d_{xy}^1 d_{xz}^2 d_{yz}^1 d_{x^2-y^2}^1 d_z^2$ (17%)	6316.9	-5.1	-0.3
4 th ES	$d_{xy}^2 d_{xz}^1 d_{yz}^1 d_{x^2-y^2}^1 d_z^2$ (28%) $d_{xy}^1 d_{xz}^2 d_{yz}^1 d_{x^2-y^2}^2 d_z^2$ (21%) $d_{xy}^1 d_{xz}^1 d_{yz}^1 d_{x^2-y^2}^2 d_z^2$ (19%)	7229.2	7.6	7.5

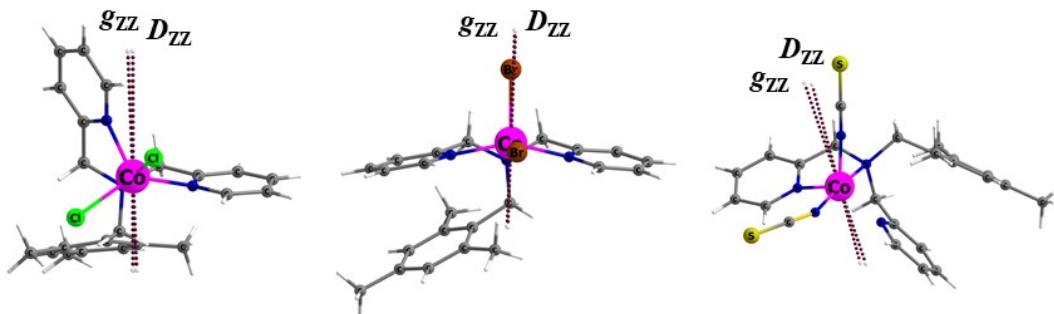


Figure S16. Z-component of the D and g -anisotropy axes (g_{zz} and D_{zz}) of the three complexes **1-3** respectively.

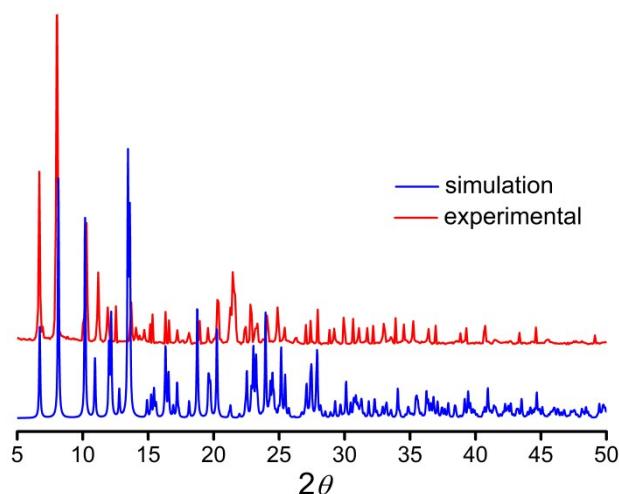


Figure S17. Experimental and simulated PXRD pattern for complex **1**.

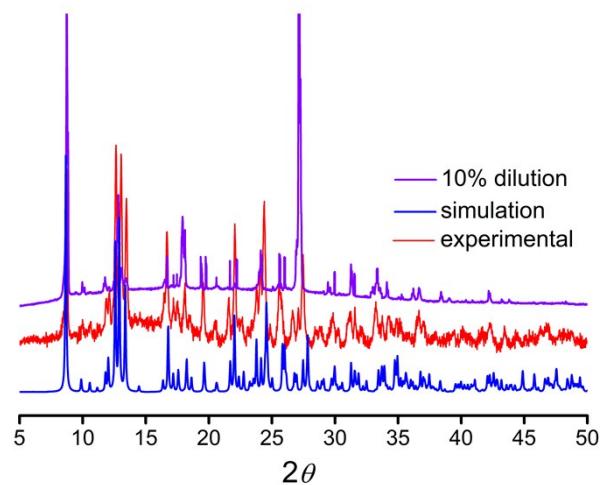


Figure S18. Experimental and simulated PXRD pattern for complex **2** and 10% diluted sample, **2₁₀**.

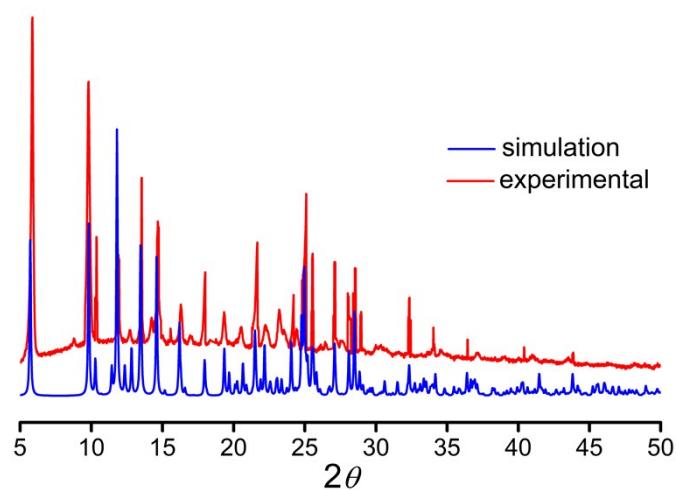


Figure S19. Experimental and simulated PXRD pattern for complex **3**.

References

- S1. F. Shao, B. Cahier, N. Guihéry, E. Rivière, R. Guillot, A.-L. Barra, Y. Lan, W. Wernsdorfer, V. E. Campbell and T. Mallah, *Chem. Commun.*, 2015, **51**, 16475-16478.
- S2. F. Shao, B. Cahier, E. Rivière, R. Guillot, N. Guihéry, V. E. Campbell and T. Mallah, *Inorg. Chem.*, 2017, **56**, 1104-1111.
- S3. T. J. Woods, M. F. Ballesteros-Rivas, S. Gómez-Coca, E. Ruiz and K. R. Dunbar, *J. Am. Chem. Soc.*, 2016, **138**, 16407-16416.
- S4. A. K. Mondal, J. Jover, E. Ruiz and S. Konar, *Chem. Commun.*, 2017, **53**, 5338-5341.
- S5. C. Rajnák, J. Titiš, O. Fuhr, M. Ruben and R. Boča, *Inorg. Chem.*, 2014, **53**, 8200-8202.
- S6. T. Jurca, A. Farghal, P.-H. Lin, I. Korobkov, M. Murugesu and D. S. Richeson, *J. Am. Chem. Soc.*, 2011, **133**, 15814-15817.
- S7. B. Cahier, M. Perfetti, G. Zakhia, D. Naoufal, F. El-Khatib, R. Guillot, E. Rivière, R. Sessoli, A.-L. Barra, N. Guihéry and T. Mallah, *Chem.-Eur. J.*, 2017, **23**, 3648-3657.
- S8. A. K. Mondal, T. Goswami, A. Misra and S. Konar, *Inorg. Chem.*, 2017, **56**, 6870-6878.
- S9. A. K. Mondal, J. Jover, E. Ruiz and S. Konar, *Chem.-Eur. J.*, 2017, **23**, 12550-12558.
- S10. F. El-Khatib, B. Cahier, F. Shao, M. López-Jordà, R. Guillot, E. Rivière, H. Hafez, Z. Saad, J.-J. Girerd, N. Guihéry and T. Mallah, *Inorg. Chem.*, 2017, **56**, 4601-4608.
- S11. F. Habib, O. R. Luca, V. Vieru, M. Shiddiq, I. Korobkov, S. I. Gorelsky, M. K. Takase, L. F. Chibotaru, S. Hill, R. H. Crabtree and M. Murugesu, *Angew. Chem. Int. Ed.*, 2013, **52**, 11290-11293.
- S12. R. Boča, L. u. Dlháň, W. Linert, H. Ehrenberg, H. Fuess and W. Haase, *Chem. Phys. Lett.*, 1999, **307**, 359-366.

S13. C. Rajnák, J. Titiš, I. Šalitroš, R. Boča, O. Fuhr and M. Ruben, *Polyhedron*, 2013, **65**, 122-128.