Supporting Information

Chemo- and Regioselective Hydroformylation of Alkenes with CO_2/H_2 over a Bifunctional Catalyst

Kaimin Hua, Xiaofang Liu*, Baiyin Wei, Zilong Shao, Yuchao Deng, Liangshu Zhong, Hui Wang*, Yuhan Sun*

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1 Material and Methods

General. All reactions and operations involving air- or moisture-sensitive compounds were performed in Ar-filled glove box. Unless otherwise noted, all commercially available reagents were used without further purification. Among these, extra-dry solvents with water \leq 50 ppm (by K.F.) were bought from Energy Chemical. And all reactions were carried out in a Teflon-lined stainless steel reactor of 50 mL capacity equipped with a magnetic stirrer.

Instrumentation. GC analysis was performed on Shimadzu-2014 with a FID detector and a Stabilwax®-DA column (60 m, 0.53 mm, 0.5 μ m). The gaseous mixture was collected and analyzed by gas chromatography equipped with thermal conductivity detector (Shimadzu GC-2014C) with Argon as the carrier gas. ESI-MS analysis was performed on WATERS 2695+ZQ2000. Detected masses are given in m/z and correlated to calculated masses of the respective species. High resolution MS (omni star gas analysis system) analyses were performed to detect ¹³CO after ¹³CO₂ Isotope experiment. Detected masses are given in m/z and correlated to calculated masses of the respective species. ¹H- and ¹³C- NMR spectra were recorded on Bruker Avance III 400 MHz spectrometers. Chemical shifts (δ values) were reported in ppm relative to internal TMS.

Calculation of the CO amount in the gaseous product. We quantified the gas from the reactor by the internal standard method (nitrogen as the internal standard). The relative correction factor (f) and correlation coefficients of each component are shown in the following table:

component	H_2	CO	CH_4	CO_2
Correction factor(f)	0.111738704	1.099931423	0.37969925	1.223846589
Correl	0.999999858	0.99933951	0.999991074	0.999999152

GC analysis of the gas sample clearly indicates the presence of CO, residual H_2 and CO_2 in the gas phase of the reaction system (Figure S1).



Figure S1 GC spectra of the gas phase after the reaction

2 Typical procedure for hydroformylation of alkenes with CO₂/H₂

General procedure: All reactions were carried out in a Teflon-lined stainless steel reactor of 50 mL capacity equipped with a magnetic stirrer. Typically, in a glovebox, 1-decene (1.0 mmol), Rh(acac)(CO)₂ (0.02 mmol), ligand (0.12 mmol), Ac₂O (2.0 mmol) and NMP (4 mL) were loaded into the reactor. The autoclave was sealed and purged three times with CO₂ gas, subsequently charged with CO₂ (20 bar), then H₂ (5 bar) to total pressure of 5 bar. The autoclave was then heated at 80 °C for 12 h. After the reaction, the autoclave was cooled in ice water and then the gas was carefully vented into the airbag. The resulting clear solution was diluted with N-methyl-2-pyrrolidinone (10 mL) and added benzyl alcohol (0.5 mmol) or isooctane (0.3 mmol) as internal standard. The sample was filtered through a short cotton plug, and the filtrate was firstly analyzed by GC-MS to determine the structures of the aldehyde products, then immediately analyzed by GC to determine the conversion of alkenes, yields of the aldehydes, and the regioselectivity (n/iso ratio) as well as percentage of isomerization and hydrogenation products. Yields were found to be reproducible within Y = ±5% in three independent runs for selected experiments.

Gas chromatography method: Stabilwax®-DA column (60 m, 0.53 mm, 0.5 μ m). Temperature: 35 °C for 4 min, then heated to 100 °C at 10 °C min⁻¹ for 2 min, rated 10 °C/min to 240 °C and held 10 min, t_{r major}=21.2 min.



Table S1. The effect of monodentate phosphine ligands^[a]

[a] Reaction conditions: 1-decene (1 mmol), $Rh(acac)(CO)_2$ (0.02 mmol), ligand (0.12 mmol), Ac_2O (2.0 mmol), $H_2/CO_2 = 5/20$ bar, 80 °C, 12 h, NMP (4 mL). [b] Determined by GC with benzyl alcohol or isooctane as the internal standard. GC at page 18-25.

nC ₈ H ₁₇	\approx + CO ₂ + H ₂	$\frac{Rh(acac)(CO)}{Rh/L}$	(2 mol%) (1/6) $v), \text{NMP}$ nC_8H_{17}	∼~CHO + nC	CHO C ₈ H ₁₇
1a	20 bar 5 b	ar 80°C, 1	12n	2a	2a'
Ligand	Ph ₂ P ^{Pph} 2	Ph ₂ P PPh ₂	Ph ₂ P PPh ₂	Ph ₂ PPPh ₂	Ph ₂ P
	L1	L2	L3	L4	L5
$Y/l/b^{[b]}$	6%/2.0	0/-	57%/1.1	87%/4.0	11%/3.0
Ligand	Fe PPh ₂	PPh ₂ PPh ₂	PPh ₂ PPh ₂		PPh ₂ PPh ₂
	L6	L7	L8	L9	L10
$Y/l/b^{[b]}$	6%/13.0	20%/12.0	15%/6.0	15%/4.7	5%/1.2
Ligand	H N O PPh ₂ PPh ₂				
	L11	L12			
$Y/l/b^{[b]}$	6%/1.4	0/-			

Table S2 The effect of bidentate phosphine ligands^[a]

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[a] Reaction conditions: 1-decene (1 mmol), $Rh(acac)(CO)_2$ (0.02 mmol), ligand (0.12 mmol), Ac_2O (2.0 mmol), $H_2/CO_2 = 5/20$ bar, 80 °C, 12 h, NMP (4 mL). [b] Determined by GC with benzyl alcohol or isooctane as the internal standard. GC at page 26-37.

Table S3 The effect of Rh precursor^[a]

Entry	Rh source	PTA/Rh	2a+2a'[b]/%	l/b ^[b]	1a+Iso ^[b] /%	H ^[b] /%	GC at page
1	(PPh ₃) ₃ (CO)Rh-H	-	17	65/35	47	36	38
2	Rh(OAc) ₃	6	80	90/10	11	9	39
3	RhCl(CO) ₂	6	17	93/7	77	6	40
4	Rh(NO ₃) ₃	6	0	-	98	2	41
5	Rh(acac)(CO) ₂	6	90	89/11	7	3	42

[a] Reaction conditions: 1-decene (1 mmol), Rh precursor (0.02 mmol), PTA (0 or 0.12 mmol), Ac₂O (2.0 mmol), H₂/CO₂ = 5/20 bar, 80 °C, 12 h, NMP (4 mL). [b] Determined by GC with benzyl alcohol or isooctane as the internal standard.

Entry	PTA/Rh	2a+2a'[b]/%	l/b ^[b]	1a+Iso ^[b] /%	H ^[b] /%	GC at page
1	0	3	85/15	83	14	43
2	3	25	68/32	22	42	44
3	6	90	89/11	7	3	45
4	9	68	90/10	28	5	46

Table S4 The effect of PTA/Rh^[a]

[a] Reaction conditions: 1-decene (1 mmol), $Rh(acac)(CO)_2(0.02 \text{ mmol})$, PTA (0 - 0.18 mmol), Ac₂O (2.0 mmol), H₂/CO₂= 5/20 bar, 80 °C, 12 h, NMP (4 mL). [b] Determined by GC with benzyl alcohol or isooctane as the internal standard.

Entry	Solvent	2a+2a'[b]/%	l/b ^[b]	1a+Iso ^[b] /%	H ^[b] /%	GC at page
1	NMP	90	89/11	7	3	47
2	DMI	83	91/9	8	9	48
3	DMSO	73	89/11	26	1	49
4	DMF	68	86/14	26	6	50
5	THF	34	97/3	14	52	51
6	toluene	27	95/5	59	14	52
7	CH ₃ CN	26	90/10	49	25	53
8	1,4-Dioxane	16	92/8	17	68	54

Table S5 The effect of the solvent[a]

[a] Reaction conditions: 1-decene (1 mmol), $Rh(acac)(CO)_2$ (0.02 mmol), PTA (0.12 mmol), Ac₂O (2.0 mmol), H₂/CO₂ = 5/20 bar, 80 °C, 12 h, solvent (4 mL). [b] Determined by GC with benzyl alcohol or isooctane as the internal standard.

Table S6 The effect of the temperature ^[a]	
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Entry	T/ºC	2a+2a'[b]/%	$l/b^{[b]}$	1a+Iso ^[b] /%	H ^[b] /%	GC at page
1	60	65	93/7	19	16	55
2	70	84	87/13	9	7	56
3	80	90	89/11	7	3	57
4	90	85	89/11	10	5	58
5	100	75	90/10	11	14	59
6	120	78	84/16	12	10	60

[a] Reaction conditions: 1-decene (1 mmol), $Rh(acac)(CO)_2$ (0.02 mmol), PTA (0.12 mmol), Ac_2O (2.0 mmol), $H_2/CO_2 = 5/20$ bar, 12 h, NMP (4 mL). [b] Determined by GC with benzyl alcohol or isooctane as the internal standard.

Entry	P _{CO2} /P _{H2} bar/bar	2a+2a' ^[b] /%	l/b ^[b]	1a+Iso ^[b] /%	H ^[b] /%	GC at page
1	40/10	88	87/13	7	5	61
2	20/5	90	89/11	7	3	62
3	22/3	83	92/8	11	6	63
4	10/3	87	91/9	9	4	64
5	7/3	74	94/6	13	13	65
6	12/12	42	87/13	19	39	66

Table S7 The effect of the $P_{\rm CO2}/P_{\rm H2}{}^{[a]}.$

[a] Reaction conditions: 1-decene (1 mmol), $Rh(acac)(CO)_2$ (0.02 mmol), PTA (0.12 mmol), Ac₂O (2.0 mmol), H₂/CO₂, 80 °C, 12 h, NMP (4 mL). [b] Determined by GC with benzyl alcohol or isooctane as the internal standard.

Entry	Alkene	Conv./%[b]	Yield ^[b] /%	l/b	Iso ^[b] /%	H ^[b] /%	GC at page
1	~~~~//	99	92	90/10	3	3	67
2	$\sim \sim \sim \sim$	99	96	90/10	1	1	68
3		99	96	88/12	2	1	69
4	$\sim\sim\sim\sim$	99	95	86/14	2	1	70
5		100	97	85/15	2	<1	71
6	\sim	99	95	86/14	2	2	72
7[c]		-	77	93/7	-	-	73
8[c]	\sim	-	73	85/15	-	-	74
9[c]	//	-	72	-	-	-	75
10	Cl	98	87	87/13	2	8	76
11	HO	99	92	73/27	1	6	77
12		98	90	85/15	3	4	78
13 ^[d]		76	75	-	0	1	79
14 ^[e]	3 1 2	90	64	16/84	25	1	80
15		99	90	10/90	0	1	81
16	CI	98	47	17/83	0	4	82
17	H ₂ CO	94	91	16/84	0	<1	83

Table S8: Substrate scope.^[a]

[a] Alkene (1 mmol), Rh(acac)(CO)₂ (0.022 mmol), PTA (0.132 mmol), Ac₂O (2.0 mmol), NMP (4 mL), H₂/CO₂=5/20 bar, 80 °C, 12 h,. [b] Determined by GC analysis with benzyl alcohol or isooctane as the internal standard. [c] alkene (3 mmol), Ac₂O (4.5 mmol). [d] T = 90 °C. [e] T=100 °C, 24 h. 1-CHO: 2-CHO: 3-CHO=16:60:24.

Table S9	Control	experiments	with	different	ligands.	[a]

a)	PPh_3

Entry	t/h	2a+2a'[b]/%	l/b ^[b]	CO/mmol	GC at page
1	0.5	8	82/18	0.10	84
2	1	10	90/10	0.12	86
3	2	13	76/24	0.08	88
4	6	13	80/20	0.08	90
5	12	15	74/26	trace	92

b) Xantphos

Entry	t/h	2a+2a'[b]/%	l/b ^[b]	CO/mmol	GC at page
1	0.5	0	0	0	94
2	1	3	60/40	trace	96
3	2	3	62/38	0.01	98
4	6	4	59/41	0.03	100
5	12	5	50/50	0.05	102

c) dppb

Entry	t/h	2a+2a'[b]/%	l/b ^[b]	CO/mmol	GC at page
1	0.5	58	79/21	0.25	104
2	1	61	82/18	0.39	106
3	2	73	80/20	0.66	108
4	6	87	80/20	0.61	110
5	12	92	78/22	0.51	112

[a] Reaction conditions: 1-decene (1 mmol), $Rh(acac)(CO)_2$ (0.02 mmol), ligand (0.12 mmol), Ac_2O (2.0 mmol), $H_2/CO_2 = 5/20$ bar, 80 °C, 12 h, NMP (4 mL). [b] Determined by GC with benzyl alcohol or isooctane as the internal standard. The data of PTA please see Table S12.

Entry	additive	2a+2a'[b]/%	l/b[b]	1a+Iso[b]/%	H ^[b] /%	GC at page
1	-	90	89/11	7	3	114
2	PTSA	0	-	97	1	115
3	CH ₃ SO ₃ H	trace	-	84	15	116
4 ^[c]	-	92	78/22	7	trace	117
5[c,d]	НСООН	94	77/23	5	1	118
6 ^[c,e]	НСООН	94	74/26	5	trace	119
7[c,e]	Ac ₂ O	98	80/20	1	trace	120
8 ^[c,f]	HCOOH+Ac ₂ O	91	84/16	9	trace	121

Table S10 Control experiments with different acids additives.^[a]

[a] Reaction conditions: 1-decene (1 mmol), Rh(acac)(CO)₂ (0.02 mmol), ligand (0.12 mmol), additive (0.1 mmol), Ac₂O (2.0 mmol), H₂/CO₂ = 5/20 bar, 80 °C, 12 h, NMP (4 mL). [b] Determined by GC with benzyl alcohol or isooctane as the internal standard. [c] Reaction conditions: 1-decene (1 mmol), Rh(acac)(CO)₂ (0.02 mmol), ligand (0.12 mmol), H₂/CO = 5/5 bar, 80 °C, 12 h, NMP (4 mL). [d] additive (0.3 mmol). [e] additive (2.0 mmol). [f] HCOOH (0.3 mmol) and Ac₂O (2.0 mmol).

Table S11: Isotope-Labeling Experiments and Control Experiment

(a)
$$nC_8H_{17} + {}^{13}CO_2 + H_2 \xrightarrow{Rh(aca)(CO)_2/PTA}{Ac_2O(2 \text{ equiv.}), NMP} nC_8H_{17} \xrightarrow{H_2}CHO + {}^{13}CO + {}^{13}$$

 $\sqrt{}$ indicated the typical substance was detected.

GC at page 122-130. NMR and Mass Spectra at page 147-153.

Entry	t/h	2a+2a'[b]/%	$l/b^{[b]}$	1a+Iso ^[b] /%	H ^[b] /%	CO/mmol	GC at page
1	1/6	14	96/4	83	3	0.46	131
2	0.5	20	>96/4	75	5	0.91	133
3	1	57	<96/4	35	7	0.71	135
4	2	73	93/7	20	7	0.40	137
5	4	82	91/9	12	6	0.39	139
6	6	85	89/11	8	7	0.37	141
7	8	85	89/11	9	6	0.44	143
8	12	89	89/11	8	4	0.41	145

Table S12 Reaction profiles for the hydroformylation of 1a with CO₂/H₂ catalyzed by Rh/PTA^[a].

[a] Reaction conditions: 1-decene (1 mmol), $Rh(acac)(CO)_2$ (0.02 mmol), PTA (0.12 mmol), Ac_2O (2.0 mmol), $H_2/CO_2 = 5/20$ bar, 80 °C, NMP (4 mL). [b] Determined by GC with benzyl alcohol or isooctane as the internal standard.

3 Gas chromatographic Data

In the following GC chromatograms, **H** represented the hydrogenation product of the corresponding alkene, and **Iso** represented the isomerization product of the alkene.



3.1 Gas chromatograms to Table S1





L3



L4



L5







without Ligand



3.2 Gas chromatograms to Table S2




























3.3 Gas chromatograms to Table S3











3.4 Gas chromatograms to Table S4



Entry 3





3.5 Gas chromatograms to Table S5



















3.6 Gas chromatograms to Table S6

Entry 1













3.7 Gas chromatograms to Table S7



Entry 2











3.8 Gas chromatograms to Table S8






















Entry 11















Entry 17



3.9 Gas chromatograms to Table S9

























PPh₃ Entry 5



Xantphos Entry 1



















Xantphos Entry 5













dppb Entry 3








dppb Entry 5





3.10 Gas chromatograms to Table S10

Entry 1





















3.11 Gas chromatograms to Table S11

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equation (c) without Ac<sub>2</sub>O
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equation (c) with Ac₂O











equation (e) Entry 2





equation (f) Entry 1

3.12 Gas chromatograms to Table S12






























Entry 8



4 NMR, ESI-MS and Mass Spectra

Figure S2 1 H NMR (CDCl₃, 400 MHz) spectra of the reaction mixture for CO₂ hydrogenation in NMP





Figure S3 ESI-MS spectra of the Rh/PTA



Figure S4 13 C NMR (CDCl₃, 100 MHz) spectra of the reaction system with 13 CO₂ (Table S10 (a))



Figure S5 ¹H and ¹³C NMR (CDCl₃, 400 MHz) spectra of the reaction system (Table S10 (b))



Figure S6 MS spectra of the reaction gas for the isotope-labeling experiments



Figure S7 ¹H NMR (CDCl₃, 400 MHz) spectra of the reaction system (c, Table S10)



Figure S8 ¹H NMR (CDCl₃, 400 MHz) spectra of HCOOAc