

Computational Insights into Selective CO₂ Hydrogenation to CH₃OH Catalysed by ZnO Based Nanocages

Shyama Charan Mandal,[†] Biswarup Pathak,^{†,*}

[†]Discipline of Chemistry, Indian Institute of Technology Indore, Simrol, Indore 453552, India

*Email: biswarup@iiti.ac.in

Contents

Figure S1: Adsorption patterns of the considered CO₂ hydrogenation reaction intermediates on ZnO NC: (a) *CO₂, (b) *COOH, (c) *CO, (d) *CHO, (e) *COH, (f) *CHOH, (g) *CH₂O, (h) *CH₂OH, (i) *CH₃O, (j) *CH₃OH, (k) *CH₂, (l) *H₂O, (m) *H, (n) *O and (o) *OH.

Figure S2: Adsorption patterns of the considered CO₂ hydrogenation reaction intermediates on Cu@ZnO NC: (a) *CO₂, (b) *COOH, (c) *CO, (d) *CHO, (e) *COH, (f) *CHOH, (g) *CH₂O, (h) *CH₂OH, (i) *CH₃O, (j) *CH₃OH, (k) *CH₂, (l) *H₂O, (m) *H, (n) *O and (o) *OH.

Table S1: Calculated electronic energy, ZPE correction and ZPE corrected reaction energy of considered intermediates for CO₂ hydrogenation reaction to CH₃OH by ZnO NC. All energies are in eV.

Table S2: Calculated electronic energy, ZPE correction and ZPE corrected reaction energy of considered intermediates for CO₂ hydrogenation reaction to CH₃OH by Cu@ZnO NC. All energies are in eV.

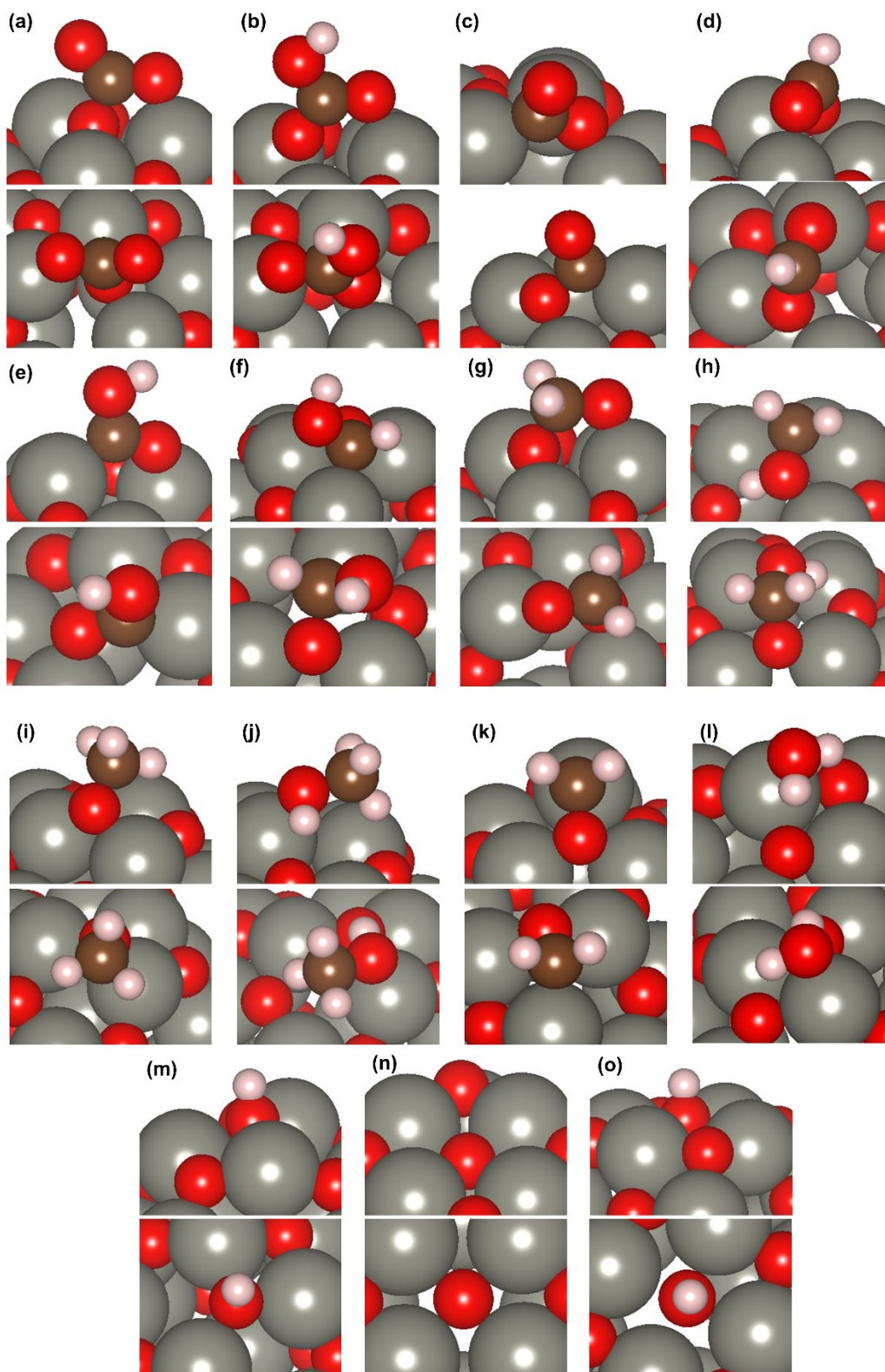


Figure S1: Adsorption patterns of the considered CO₂ hydrogenation reaction intermediates on ZnO NC: (a) $\ast\text{CO}_2$, (b) $\ast\text{COOH}$, (c) $\ast\text{CO}$, (d) $\ast\text{CHO}$, (e) $\ast\text{COH}$, (f) $\ast\text{CHOH}$, (g) $\ast\text{CH}_2\text{O}$, (h) $\ast\text{CH}_2\text{OH}$, (i) $\ast\text{CH}_3\text{O}$, (j) $\ast\text{CH}_3\text{OH}$, (k) $\ast\text{CH}_2$, (l) $\ast\text{H}_2\text{O}$, (m) $\ast\text{H}$, (n) $\ast\text{O}$ and (o) $\ast\text{OH}$.

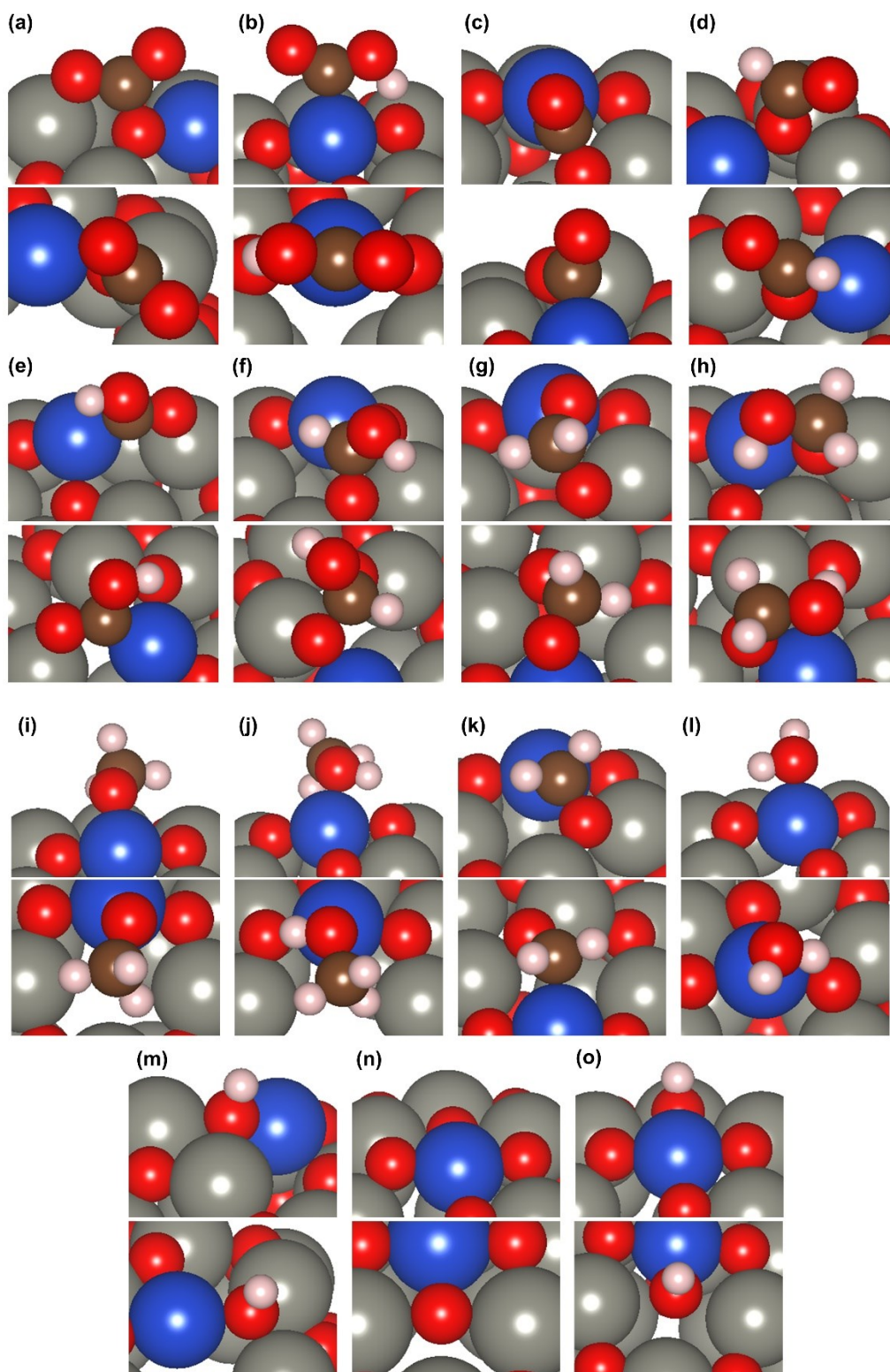


Figure S2: Adsorption patterns of the considered CO₂ hydrogenation reaction intermediates on Cu@ZnO NC: (a) *CO₂, (b) *COOH, (c) *CO, (d) *CHO, (e) *COH, (f) *CHOH, (g) *CH₂O, (h) *CH₂OH, (i) *CH₃O, (j) *CH₃OH, (k) *CH₂, (l) *H₂O, (m) *H, (n) *O and (o) *OH.

Table S1: Calculated electronic energy, ZPE correction and ZPE corrected reaction energy of considered intermediates for CO₂ hydrogenation reaction to CH₃OH by ZnO NC. All energies are in eV.

Intermediates	Energy	ZPE corrections	ZPE corrected reaction energy
ZnO NC	-198.8320	-	-198.8320
CO ₂	-22.9622	0.3062	-22.6560
H ₂	-6.7686	0.2679	-6.5007
H	-3.3843	0.1339	-3.2504
*CO ₂	-222.0834	0.2296	-221.8538
*COOH	-225.4550	0.5755	-224.8795
*CO	-214.0428	0.1087	-213.9341
H ₂ O	-14.2188	0.5684	-13.6504
*CHO	-217.9717	0.3625	-217.6092
*COH	-217.1852	0.3756	-216.8096
*CH ₂ O	-221.8826	0.6717	-221.2109
*CHOH	-221.0812	0.6997	-220.3815
*CH ₃ O	-224.2807	1.0103	-223.2704
*CH ₂ OH	-225.2234	1.0312	-224.1922
*CH ₂	-214.2784	0.4617	-213.8197
*CH ₃ OH	-229.7218	1.3697	-228.3521
O ₂	-9.8454	-	-9.8454
O	-4.9227	-	-4.9227
CH ₃ OH	-30.2253	1.3567	-28.8686

Table S2: Calculated electronic energy, ZPE correction and ZPE corrected reaction energy of considered intermediates for CO₂ hydrogenation reaction to CH₃OH by Cu@ZnO NC. All energies are in eV.

Intermediates	Energy	ZPE corrections	ZPE corrected reaction energy
Cu@ZnO NC	-199.6432	-	-199.6432
CO ₂	-22.9622	0.3062	-22.6560
H ₂	-6.7686	0.2679	-6.5007
H	-3.3843	0.1339	-3.2504
*CO ₂	-222.9462	0.2269	-222.7193
*COOH	-226.0777	0.4953	-225.5824
*CO	-215.3012	0.1090	-215.1922
H ₂ O	-14.2188	0.5684	-13.6504
*CHO	-219.2358	0.3680	-218.8678
*COH	-219.0440	0.3810	-218.6630
*CH ₂ O	-222.6953	0.6636	-222.0317
*CHOH	-222.4853	0.7126	-221.7727
*CH ₃ O	-225.7869	0.9964	-224.7905
*CH ₂ OH	-226.5876	1.0233	-225.5643
*CH ₂	-215.5456	0.4641	-215.0815
*CH ₃ OH	-230.4229	1.3615	-229.0614
O ₂	-9.8454	-	-9.8454
O	-4.9227	-	-4.9227
CH ₃ OH	-30.2253	1.3567	-28.8686