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Supporting information

Reusable ZIF-8@Chitosan Sponge for Efficiently Selective Removal of Congo Red

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Anionic dyes	<p>Alizarin Yellow (AY)</p> <p>Methyl Orange (MO)</p> <p>Congo Red (CR)</p>
Neutral dyes	<p>Dimethyl Yellow (DY)</p> <p>Sudan-1 (SD-1)</p>
Cationic dyes	<p>Rhodamine B (RhB)</p> <p>Methylene Blue (MB)</p>

Figure S1 Structures of organic dyes with different charges.

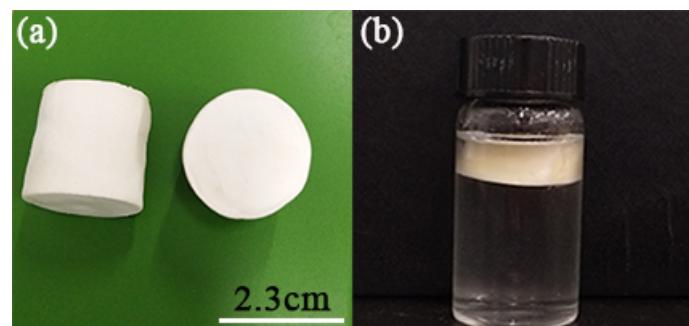


Figure S2 (a) Digital photo of the ZIF-8@CS composite sponge; (b) Digital photo of ZIF-8@CS composite sponge placed in water.

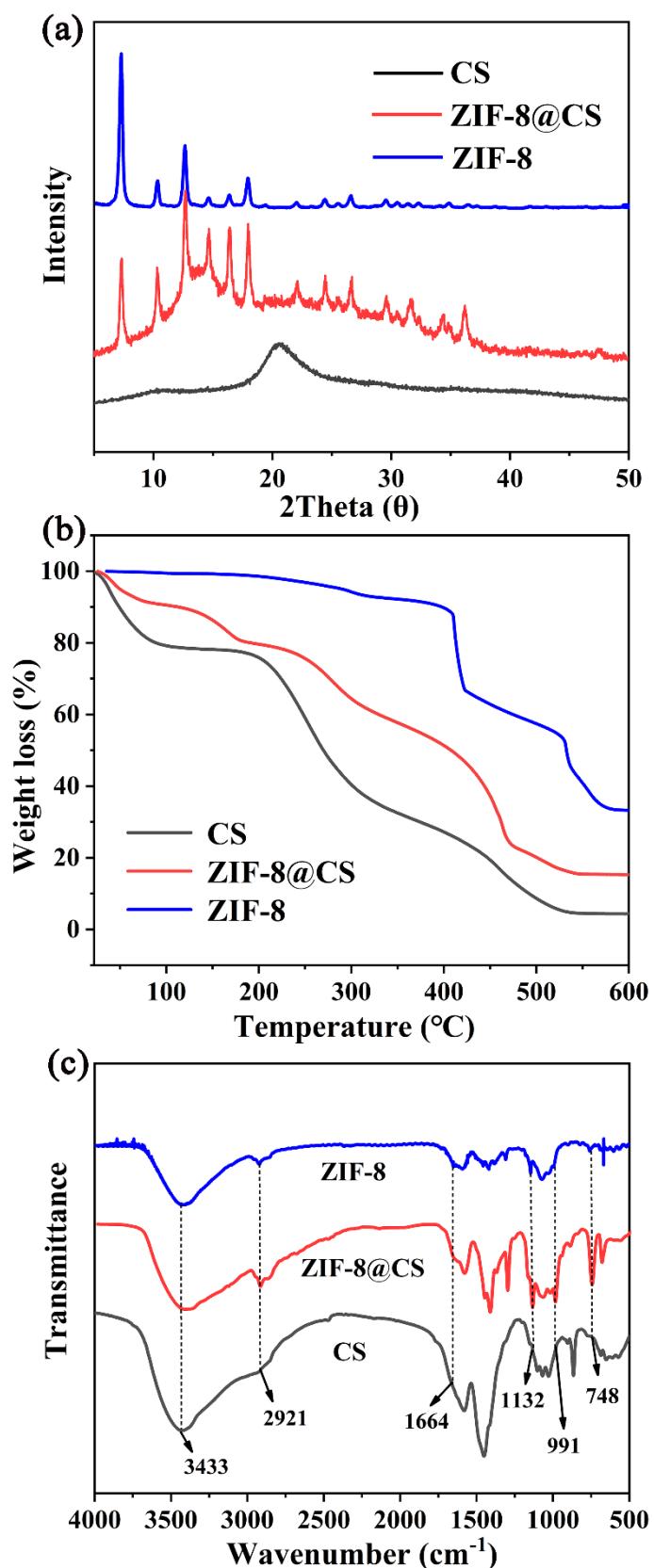


Figure S3 (a) PXRD patterns, (b) TG curves and (c) FTIR spectra of CS, ZIF-8 and ZIF-8@CS sponges.

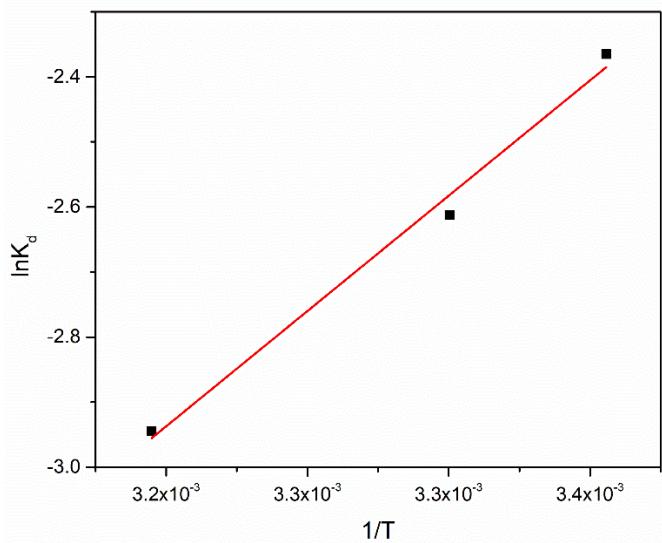


Figure S4 The influence of temperature on the CR sorption relationship curve between $\ln K_d$ and $1/T$.

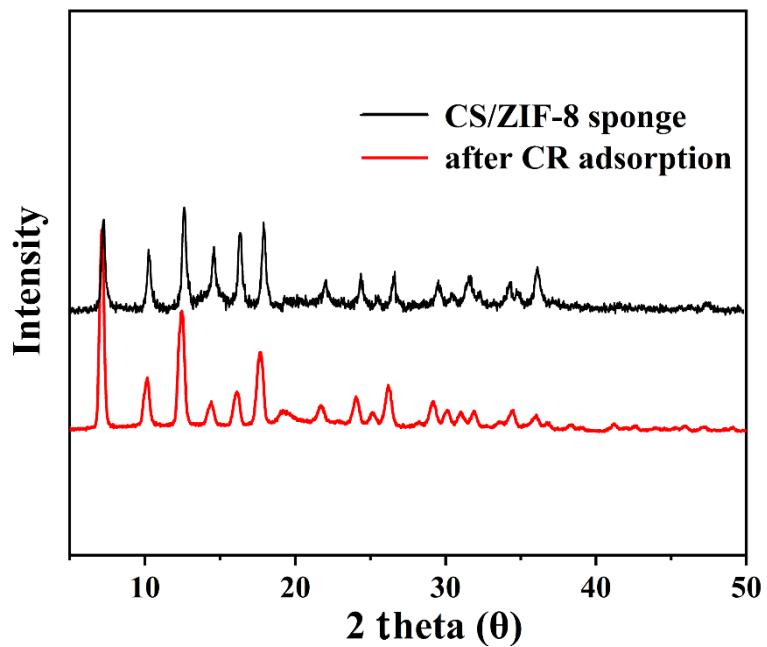


Figure S5 PXRD patterns of ZIF-8@CS sponges before and after CR adsorption.

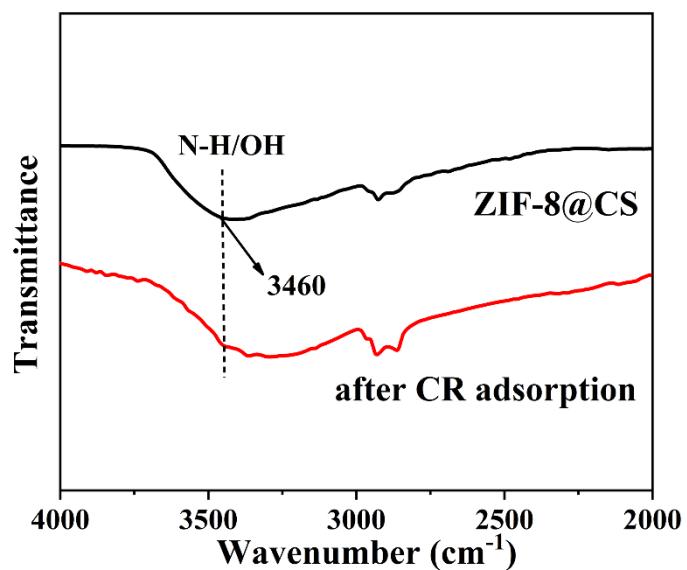


Figure S6 FT-IR spectra of ZIF-8@CS sponges before and after CR adsorption.

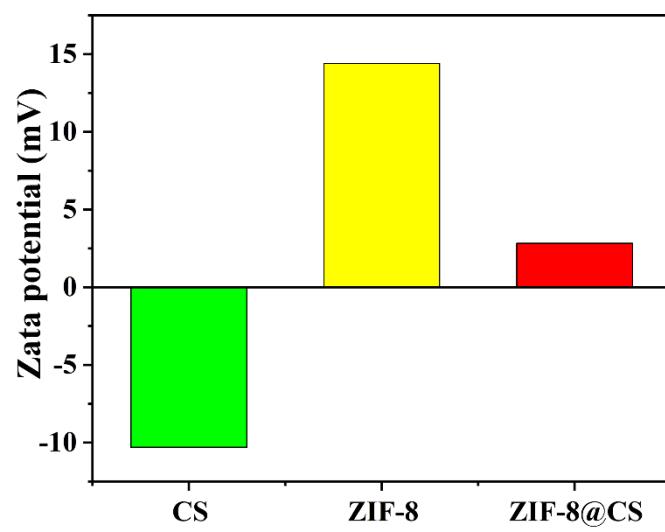


Figure S7 Zeta potentials of different materials.

Table S1 The adsorption kinetic fitting parameters.

	Pseudo-second-order model			Pseudo-first-order model		
	q _e	R ²	K ₂	q _e	R ²	K ₁
CS	105.12	0.98	3.31×10 ⁻⁴	100.90	0.76	1.11×10 ⁻²
ZIF-8	468.00	0.99	4.70×10 ⁻⁵	459.96	0.97	1.32×10 ⁻²
ZIF-8@CS	935.52	0.97	8.96×10 ⁻⁶	932.57	0.91	1.48×10 ⁻²

Table S2 Langmuir, Freundlich and Langmuir-Freundlich Constants and Correlation Coefficients.

	Langmuir			Freundlich			Langmuir-Freundlich			
	q _{max}	K _L	R ²	K _F	n	R ²	q _{max}	K _a	n	R ²
CS	273.24	0.013	0.996	5.145	1.254	0.934	344.72	0.009	0.948	0.999
ZIF-8	726.77	0.050	0.991	64.253	1.792	0.968	546.83	0.089	1.459	0.999
ZIF-8@CS	1221.06	0.093	0.982	195.216	2.253	0.945	987.01	0.141	1.612	0.995

Table S3 Langmuir Constants and Correlation Coefficients for CR adsorption upon ZIF-8@CS sponge at different temperature.

	q _{max} (mg/g)	K _L	R ²
298 K	1221.06	0.093	0.982
303 K	1064.08	0.073	0.984
313 K	791.86	0.052	0.968

Table S4 Thermodynamic parameters of CR sorption by ZIF-8@CS sponge.

ΔH (kJ mol ⁻¹)	ΔG (kJ mol ⁻¹)			ΔS (J mol ⁻¹ k ⁻¹)
	298k	308k	318k	
-29.47	-8.54	-8.19	-7.49	-70.22