

Electronic Supplementary Information

**“On water” palladium catalyzed diastereoselective boronic acid addition to structurally
diverse cyclopropane nitriles**

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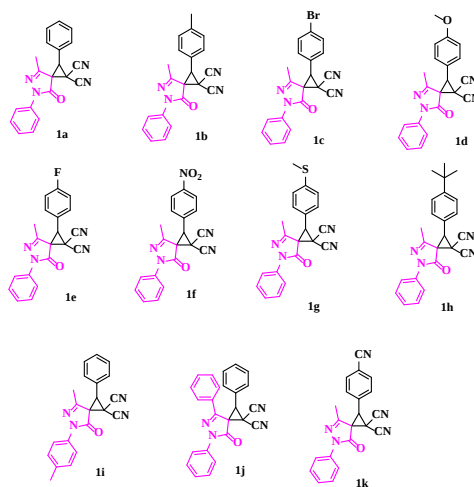
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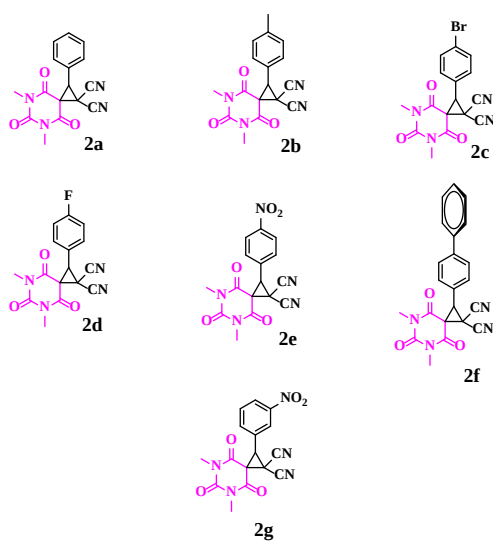
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I. List of starting materials:

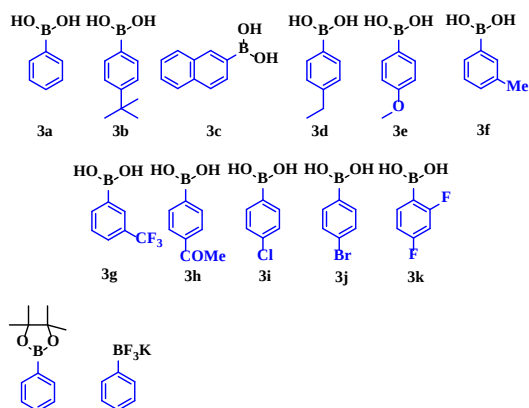
a. Pyrazolo cyclopropanes:



b. Barbiturate cyclopropanes:

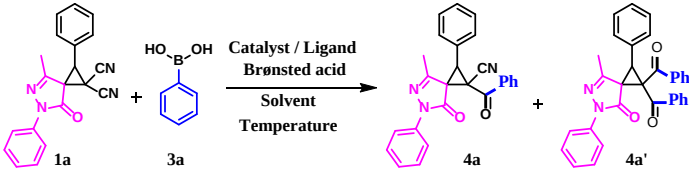


c. Organo-boron Compounds:

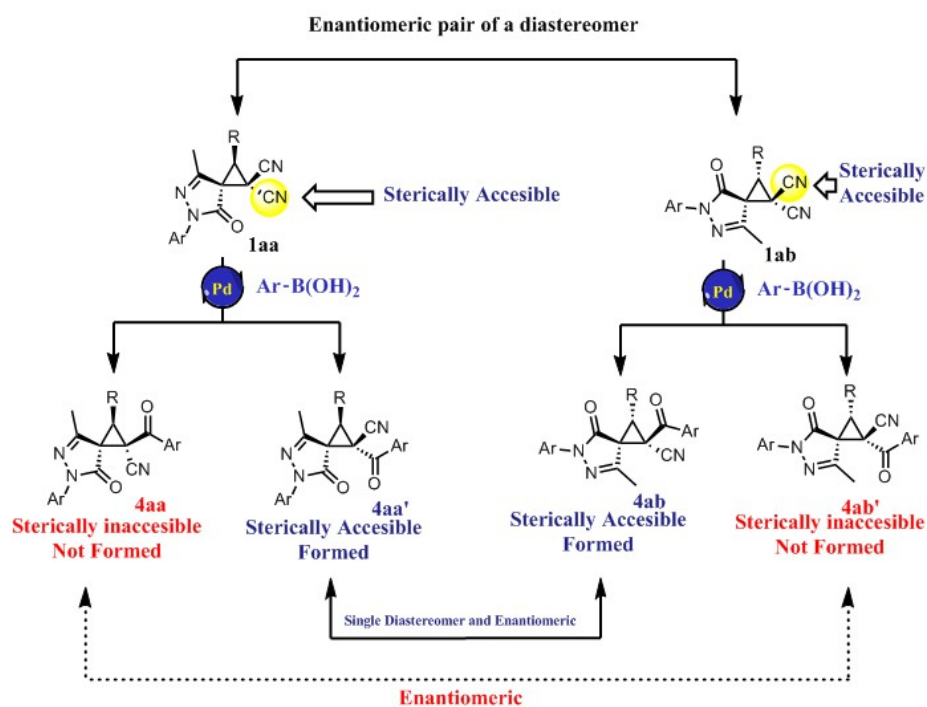


II. Optimization of the reaction conditions and study of stereoselectivity of compound 4:

Table S1. Screening of catalysts, ligands, acids, solvents and optimization of reaction conditions^a

							
Entry	Catalyst (mol%)	Ligand (mol%)	Brønsted acid (equiv.)	Solvent	Temp. (°C)	Time (h)	Yield of 4a ^b (%)
1 ^c	Pd(OAc) ₂ (5)	L1(5)	AcOH(2.0)	1,4-dioxane+H ₂ O (1:1)	100	3.0	10
2 ^c	Pd(OAc) ₂ (5)	L3(5)	AcOH(2.0)	1,4-dioxane+H ₂ O (1:1)	100	3.0	18
3	Pd(TFA) ₂ (3)	L2(3)	DBSA(0.2)	H ₂ O	90	2.0	84
4	Pd(CH ₃ CN) ₂ Cl ₂ (3)	L2(3)	DBSA(0.2)	H ₂ O	90	2.0	61
5	Pd(acac) ₂ (3)	L2(3)	DBSA(0.2)	H ₂ O	90	2.0	88
6	Pd(OAc) ₂ (3)	L1(3)	DBSA(0.2)	H ₂ O	90	2.0	42
7	Pd(OAc) ₂ (3)	L3(3)	DBSA(0.2)	H ₂ O	90	2.0	87
8	Pd(OAc) ₂ (3)	L4(3)	DBSA(0.2)	H ₂ O	90	2.0	-
9	Pd(OAc) ₂ (3)	L5(3)	DBSA(0.2)	H ₂ O	90	2.0	-
10	Pd(OAc) ₂ (3)	L2(3)	(±)CSA(1.1)	THF + H ₂ O (1:1)	90	2.0	50
11	Pd(OAc) ₂ (3)	L2(3)	TFA(1.1)	THF + H ₂ O (1:1)	90	2.0	31
12	Pd(OAc) ₂ (3)	L2(3)	Phenylacetic Acid(1.1)	THF + H ₂ O (1:1)	90	2.0	12
13	Pd(OAc) ₂ (3)	L2(3)	Benzoic Acid(1.1)	THF + H ₂ O (1:1)	90	2.0	-
14	Pd(OAc) ₂ (3)	L2(3)	p-nitro Benzoic Acid(1.1)	THF + H ₂ O (1:1)	90	2.0	-
15	Pd(OAc) ₂ (3)	-	DBSA(0.2)	H ₂ O	90	2.0	42
16	Pd(OAc) ₂ (3)	L2(3)	-	H ₂ O	90	2.0	12
17	Pd(OAc) ₂ (3)	L2(3)	-	H ₂ O	90	24.0	38
18	Pd(OAc) ₂ (3)	L2(3)	-	THF	90	24.0	14
19	Pd(OAc) ₂ (3)	L2(3)	-	1, 4-dioxane	90	24.0	12
20	Pd(OAc) ₂ (3)	L2(3)	-	Toluene	90	24.0	-

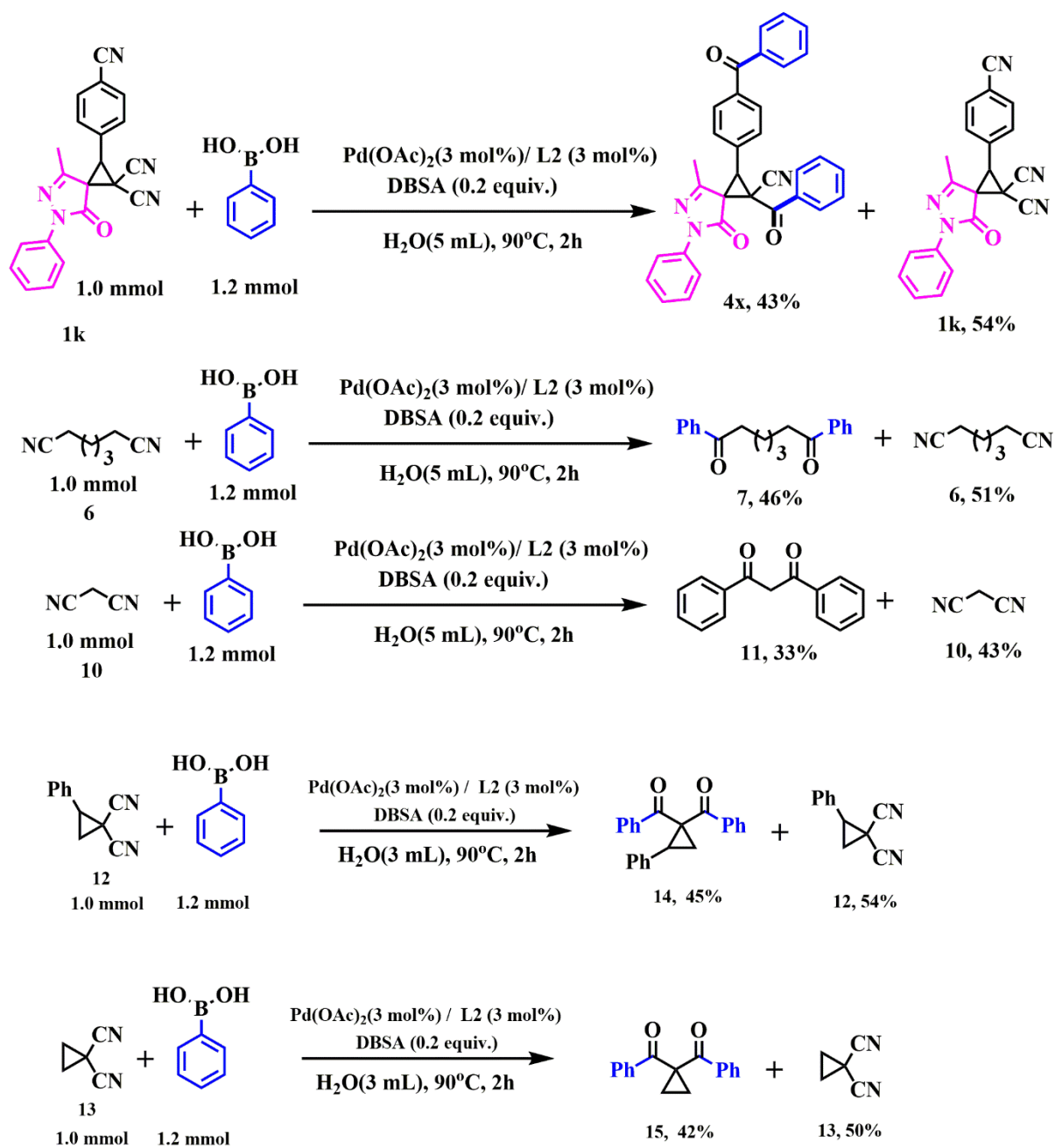
^a all the reactions were carried out with **1a** (0.5 mmol) and **2a** (0.6 mmol) in 3 mL of solvent or solvent mixture; ^b isolated yields of **4a** and **4a'** was not obtained in every case. ^c 2.0 equiv. of phenylboronic acid(**3a**) was used instead of 1.2 equiv.



Scheme S1. Diastereoselectivity of compound **4**

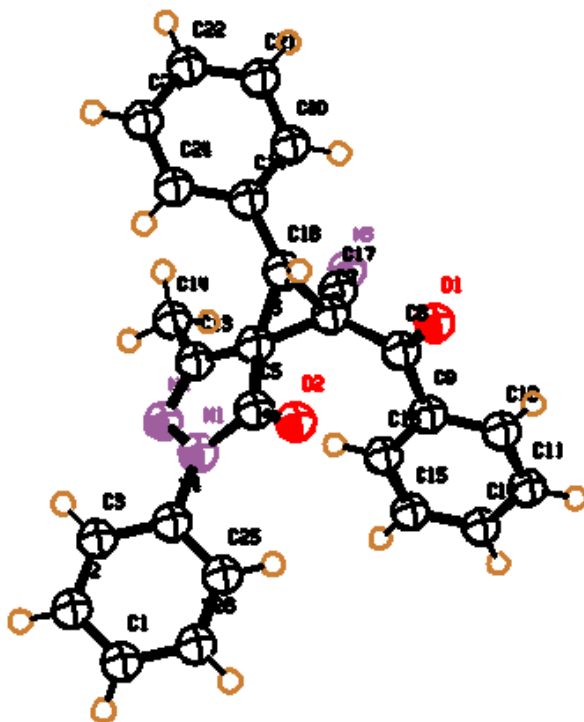
We have started from compound **1** which has been prepared according to our previous work¹ and has two diastereomers of which we have only considered the major one and have only explained the diastereoselectivity of it. Since **1aa** and **1ab** are in enantiomeric relationship, during the reaction four stereoisomers can be obtained, at least theoretically. However, since formation of isomer **4aa** and **4ab'** are sterically forbidden they were not obtained during the reaction which is confirmed by NMR and X-ray. Only **4aa'** and **4ab** were found to be as the products which are essentially enantiomers with respect to each other. Interestingly, **4aa** and **4ab'**, bearing enantiomeric relationship to each other, possess diastereomeric relationship with **4aa'** and **4ab** respectively. Thus, it is quite clear that this reaction is indeed a diastereoselective reaction and affording only one diastereomer as an enantiomeric mixture.

III. Control Experiments:

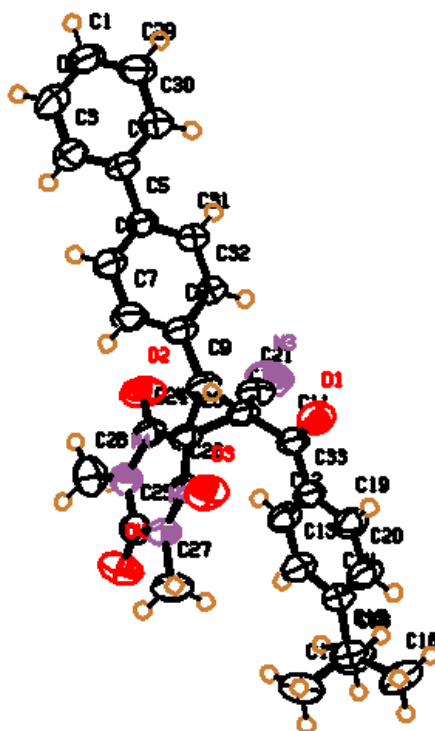


Scheme S2: Some control experiments performed varying the amount of boronic acids for different dinitrile compounds

IV. X-ray Crystallography Data of Compound 4a (CCDC 1952896) and 5g (CCDC 1952797):



The X-ray structure of **4a**. The ellipsoid contour percent probability level is 70%.



The X-ray structure of **5g**. The ellipsoid contour percent probability level is 70%.

Single crystal X-ray data for compound Compound4a (CCDC 1952896) and 5g (CCDC 1952797):

Single crystals suitable for X-ray diffraction of **4a** and **5g** were grown from ethyl acetate. The crystals were carefully chosen using a stereo zoom microscope supported by a rotatable polarizing stage. In all the cases the data were collected at 296(2) K on a CCD diffractometer with graphite monochromated Mo-K α radiation (0.71073 Å). The data were processed using the package SAINT.¹ Structures were solved by direct and Fourier methods and refined by full-matrix least squares based on F² using SHELXTL² and SHELXL-97³ packages.

Table 1. Crystallographic data for the compound **4a** and **5g**

Compounds	4a	5g
Empirical formula	C ₂₆ H ₁₉ N ₃ O ₂	C ₃₂ H ₃₀ N ₃ O ₄
Formula weight	405.44	520.59
crystal system	Monoclinic	Monoclinic
space group	<i>P</i> 21/ <i>c</i>	<i>P</i> 2(1)/ <i>c</i>
<i>a</i> (Å)	10.086(2)	18.0183(7)
<i>b</i> (Å)	16.995(3)	7.4454(3)
<i>c</i> (Å)	12.365(3)	20.6732(7)
α (°)	90.00	90.00
β (°)	98.074(3)	104.0510(10)
γ (°)	90.00	90.00
<i>V</i> (Å³)	2098.5(7)	2690.40(18)
<i>Z</i>	4	4
<i>T</i>, K	296(2)	296(2)
Wavelength (Å)	0.71073	0.71073
2θ (°)	2.37-20.47	2.30-35.46
μ (mm⁻¹)	0.083	0.086
ρ_{calcd} (g cm⁻³)	1.283	1.285
<i>F</i> (000)	848	1100
absorption correction	multi-Scan	multi-Scan
index ranges	-11 ≤ <i>h</i> ≤ 12	-29 ≤ <i>h</i> ≤ 29
	-21 ≤ <i>k</i> ≤ 21	-10 ≤ <i>k</i> ≤ 12
	-15 ≤ <i>l</i> ≤ 15	-31 ≤ <i>l</i> ≤ 34
reflections collected	15560	45612

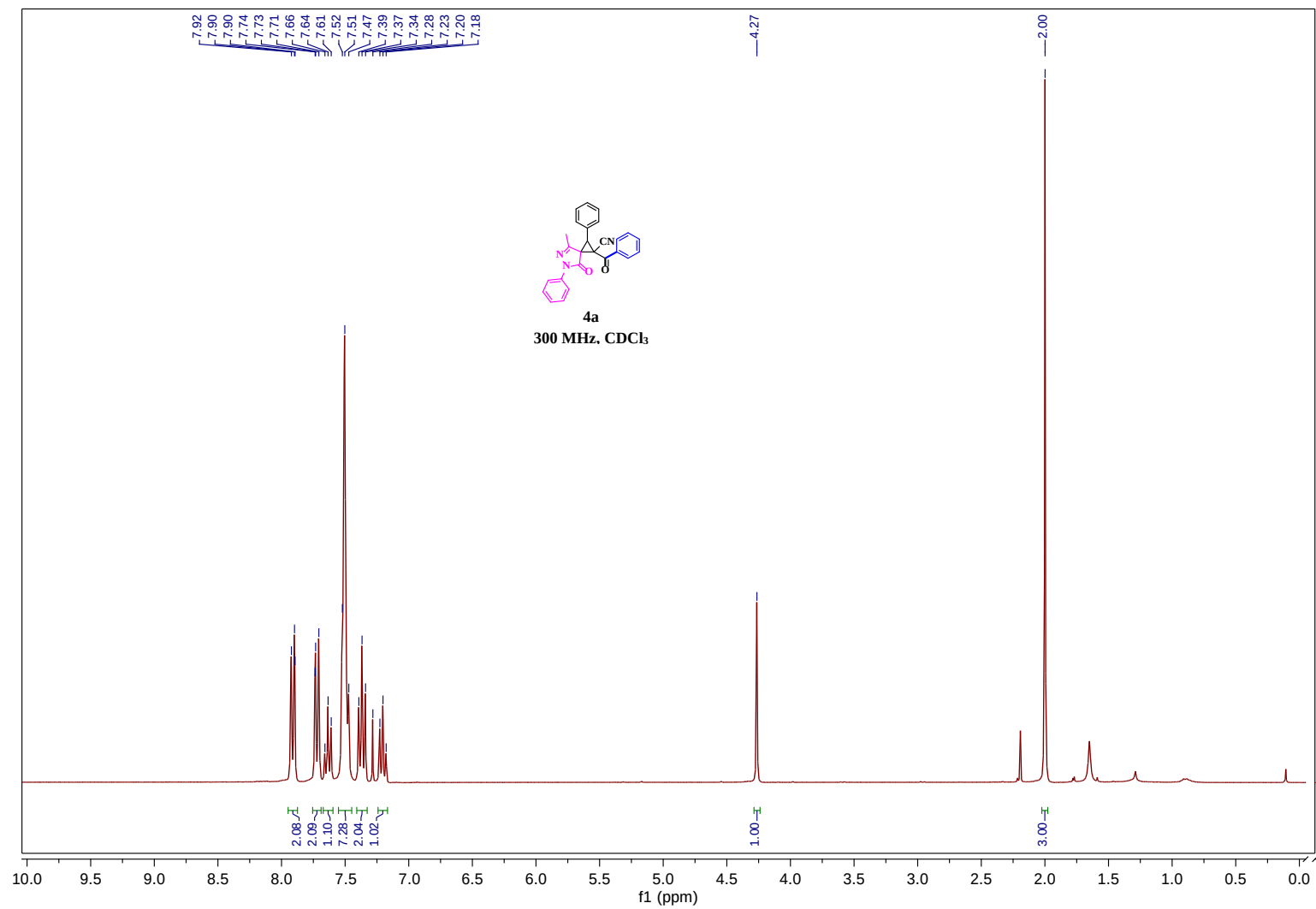
independent reflections (R_{int})	2902(0.0405)	9986(0.0281)
Goodness-of-fit on F^2	0.900	0.927
R_1^a/wR_2^b ($I > 2\sigma(I)$)	0.0439/0.1249	0.0616/ 0.1812
R_1^a/wR_2^b (for all data)	0.0854/0.1555	0.1040/0.2218
Largest diff. peak/hole / e $\text{\AA}^{-3}$	0.141/-0.163	0.393/-0.213

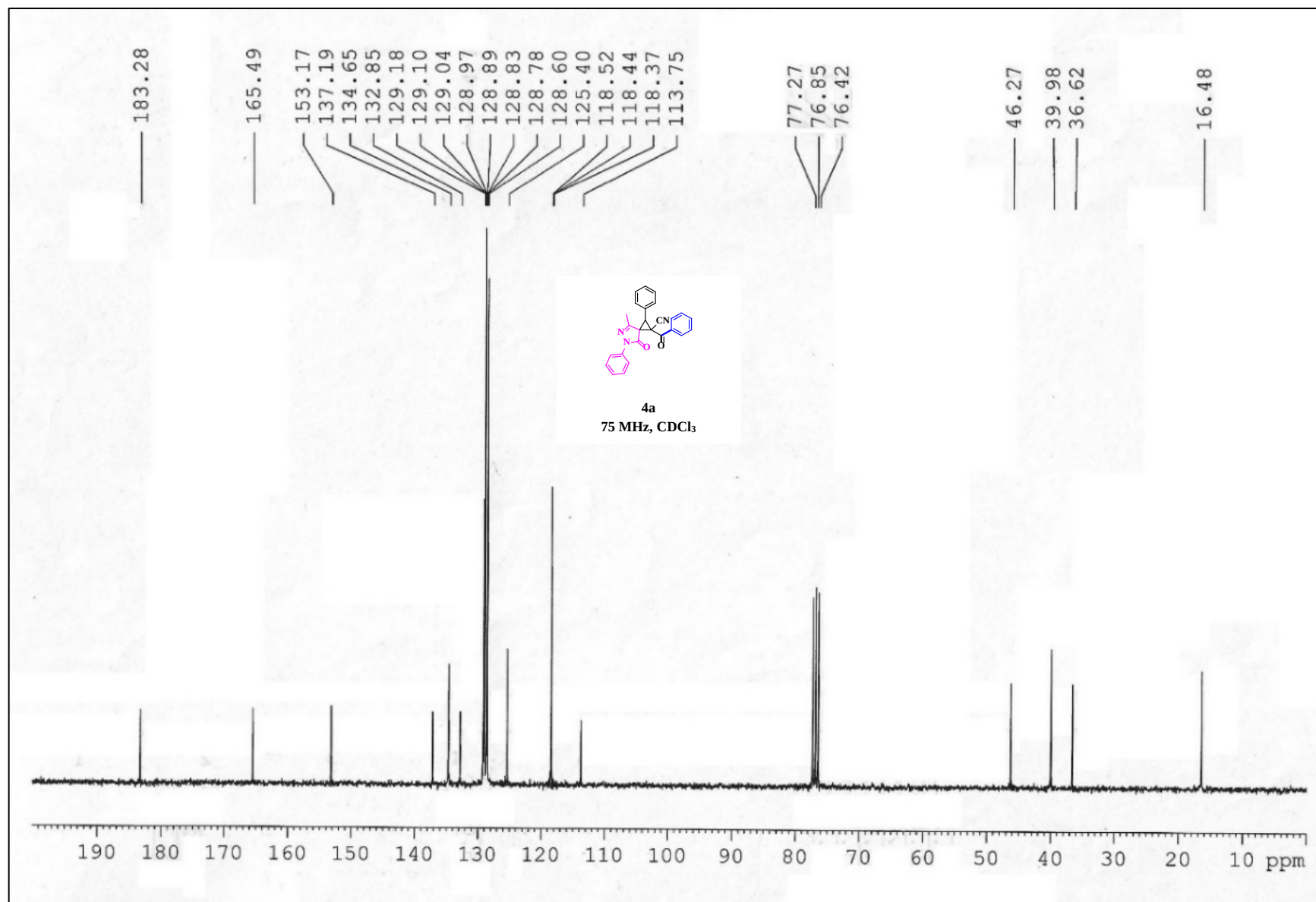
$$^aR_1 = [\sum||F_o| - |F_c||/\sum|F_o|]. \quad ^b wR_2 = [\sum w(F_o^2 - F_c^2)^2/\sum wF_o^4]^{1/2}$$

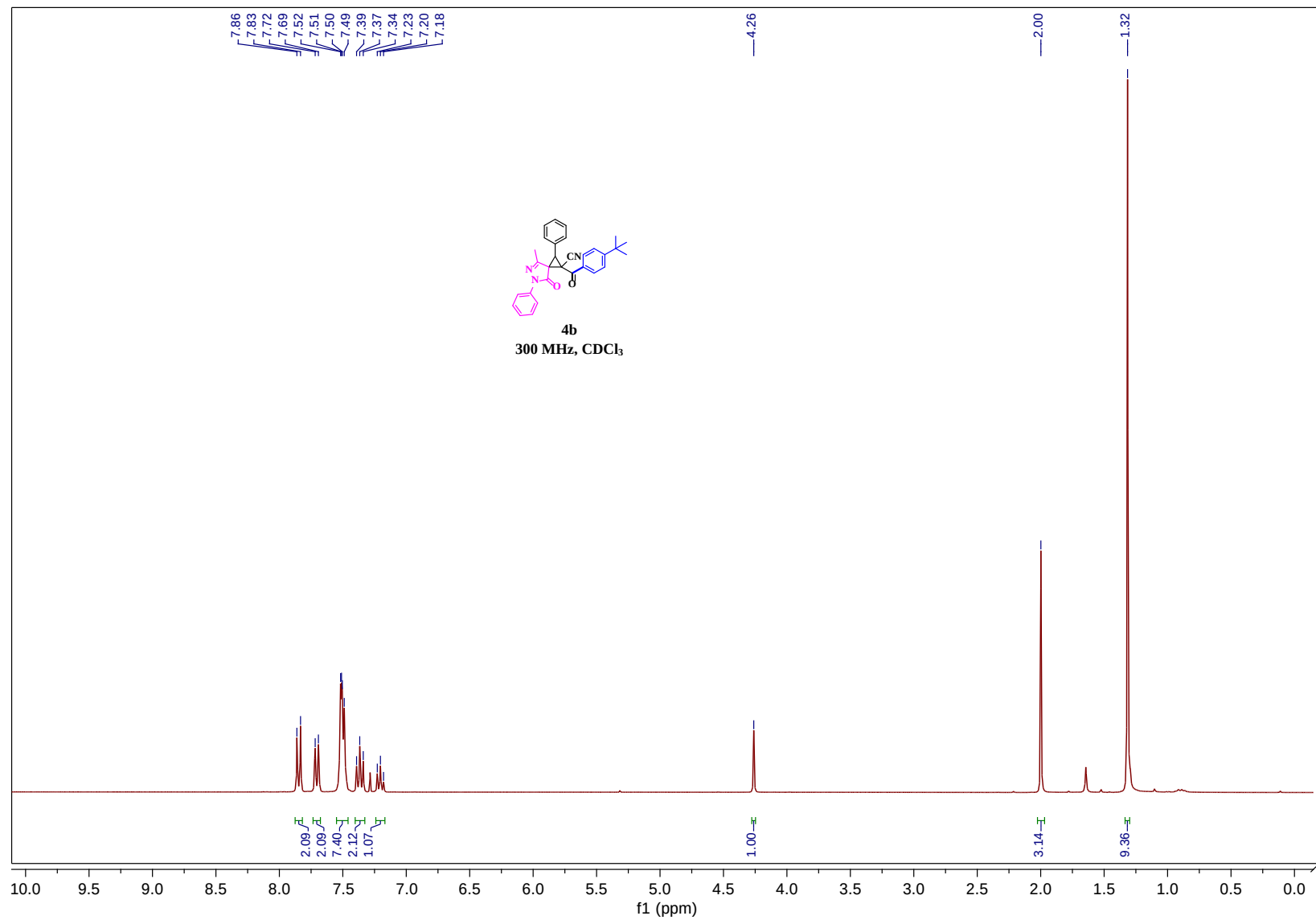
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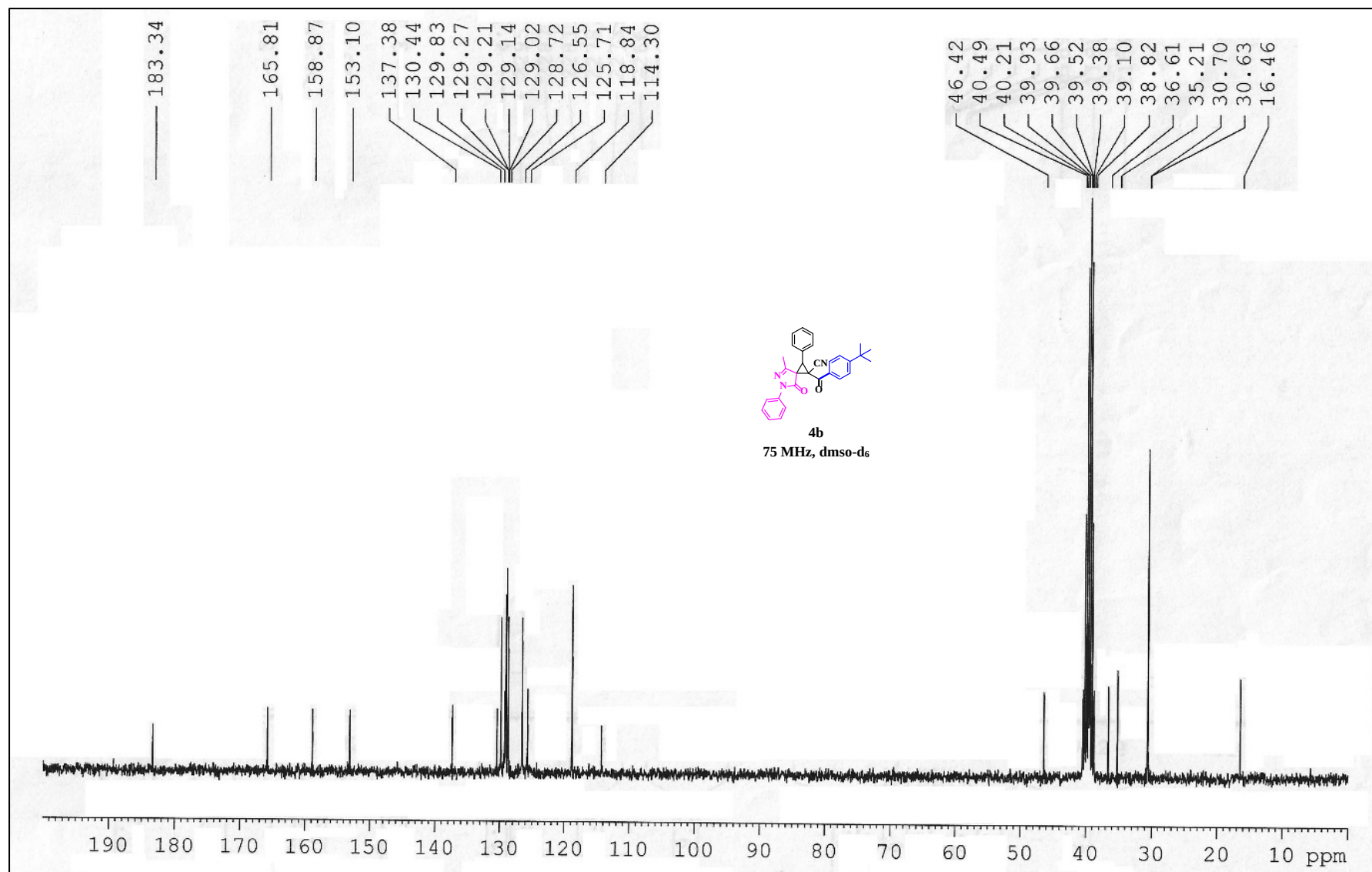
1. P. Mukherjee, A. R. Das. *J. Org. Chem.*, 2017, **82**, 2794
2. *APEX-II*, *SAINT-Plus*, and *TWINABS*; Bruker-Nonius AXS Inc.: Madison, WI, 2004.
3. *SHELXTL*, version 6.10; Bruker AXS Inc.: Madison, WI, 2002.
4. Sheldrick, G. M. *SHELXL-97, Crystal Structure Refinement Program*; University of Göttingen: Göttingen, Germany, 1997.

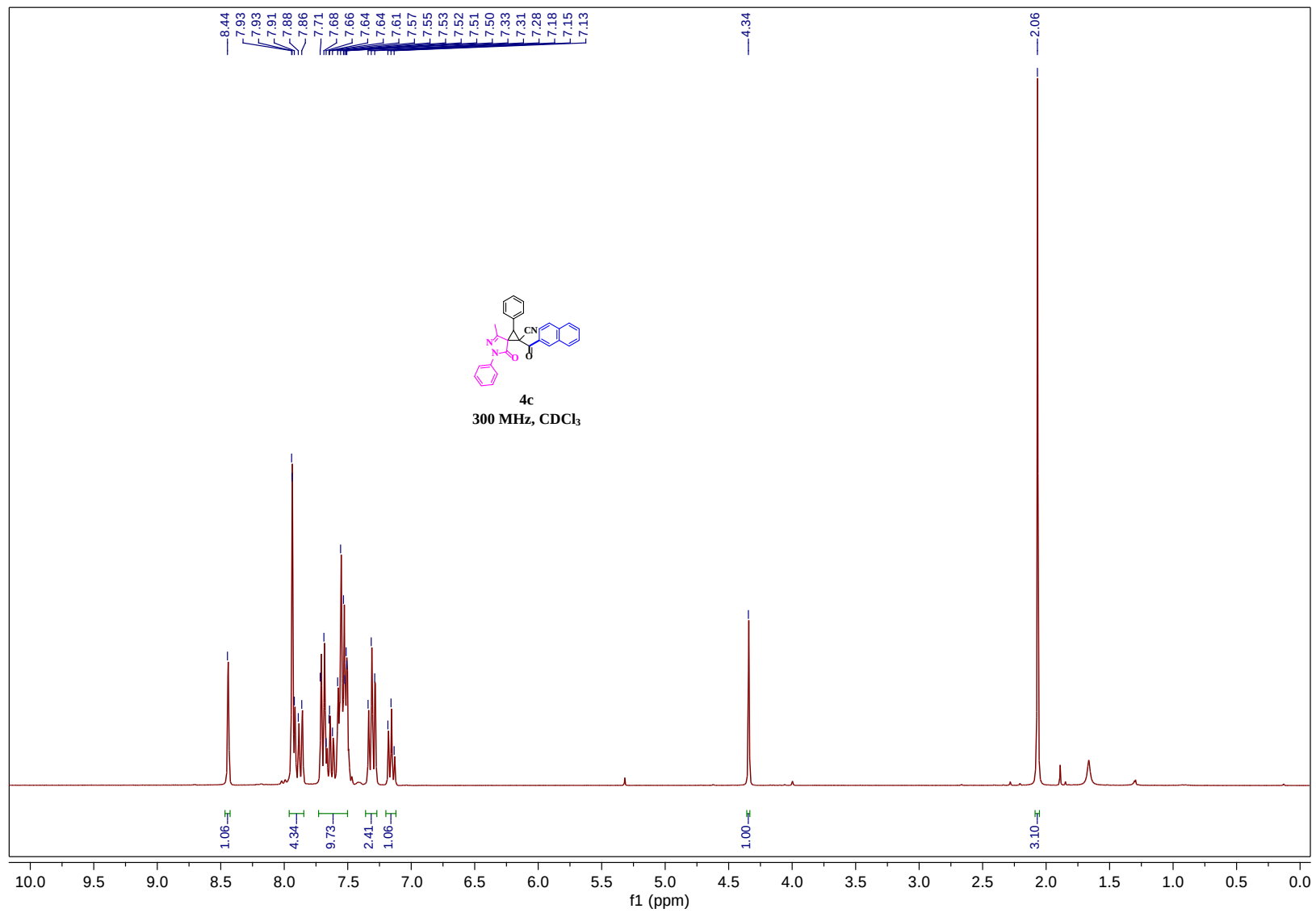
V. ^1H and ^{13}C NMR spectra for the products 4a-4w, 5a-i

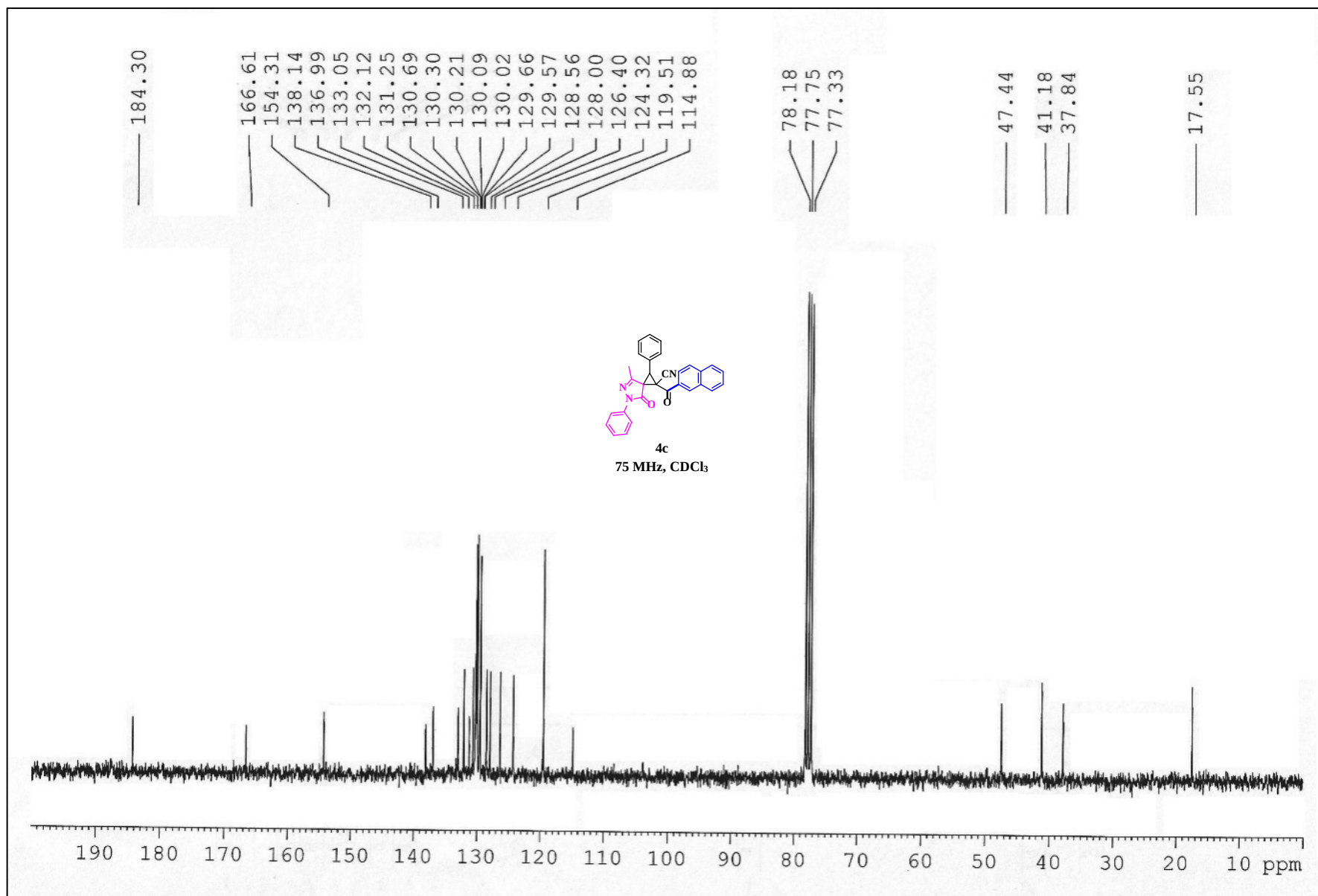


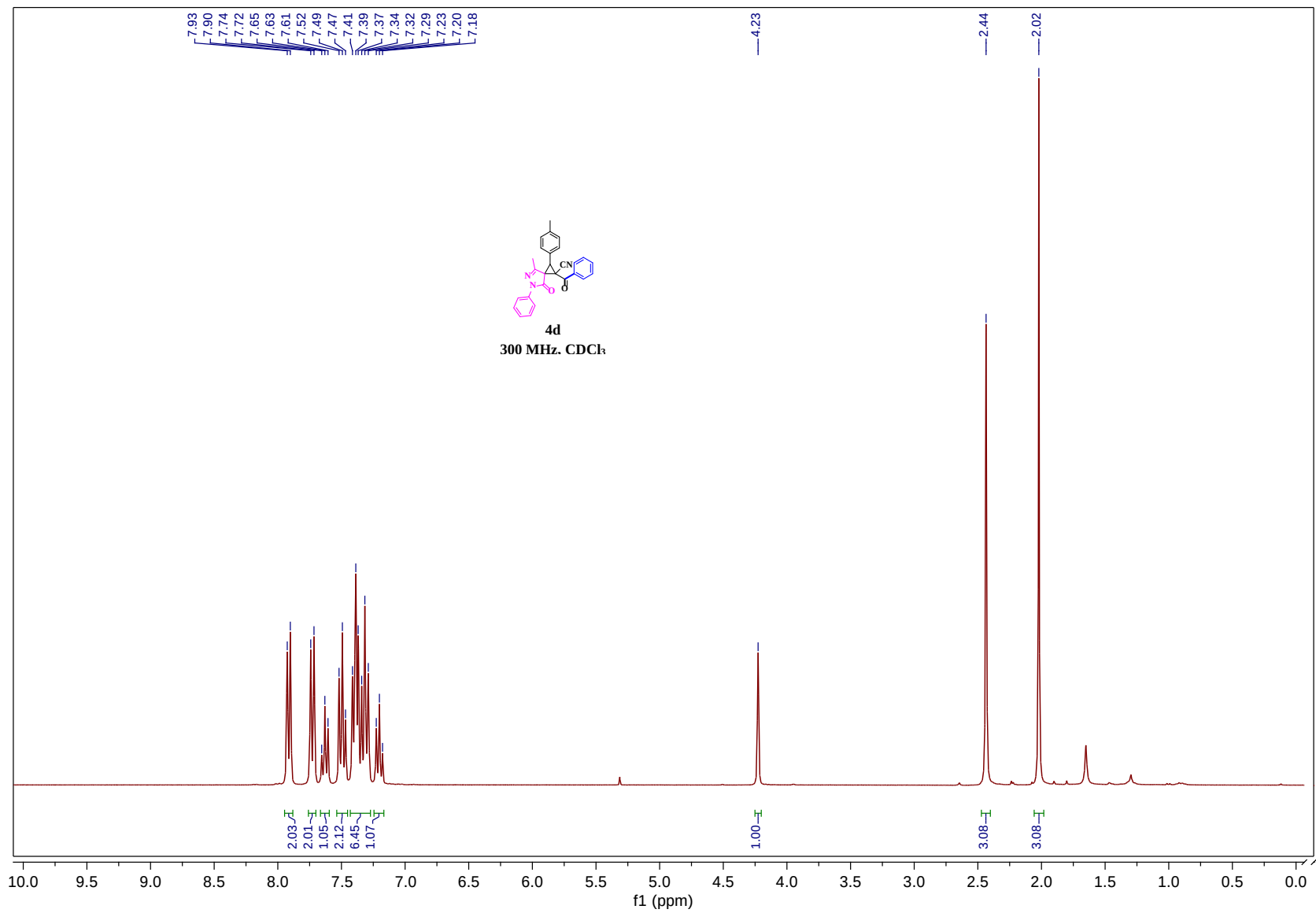


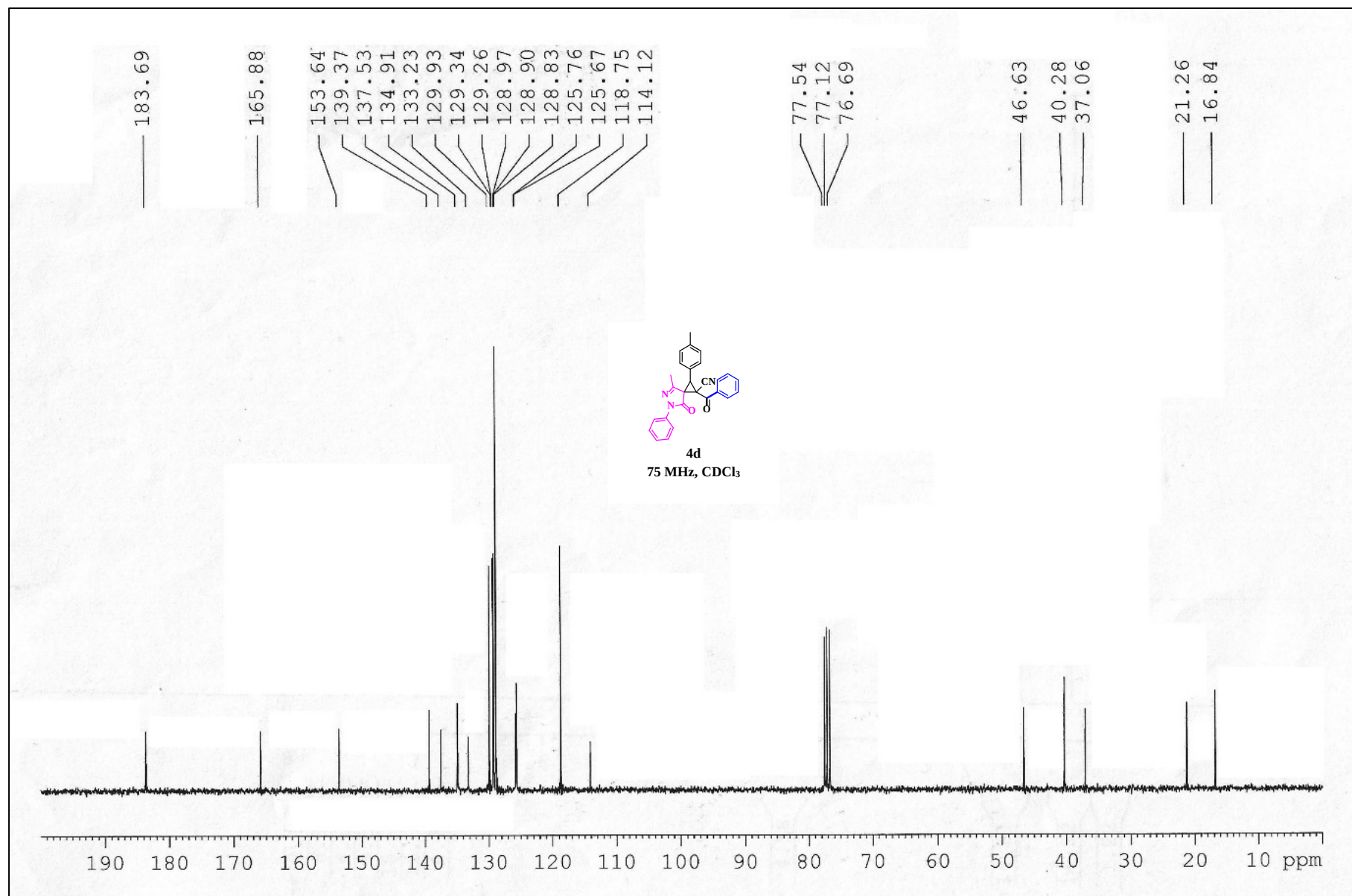


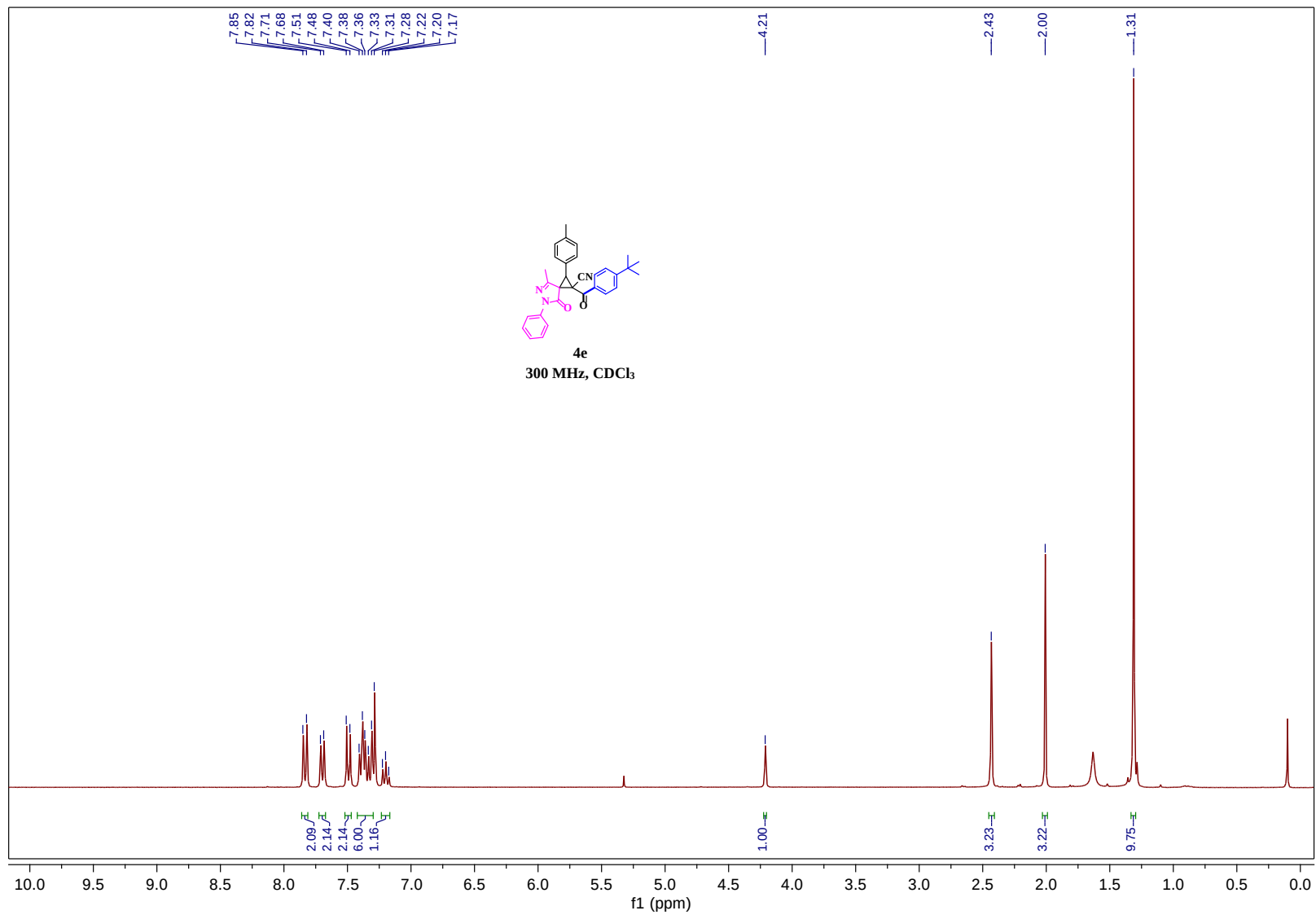


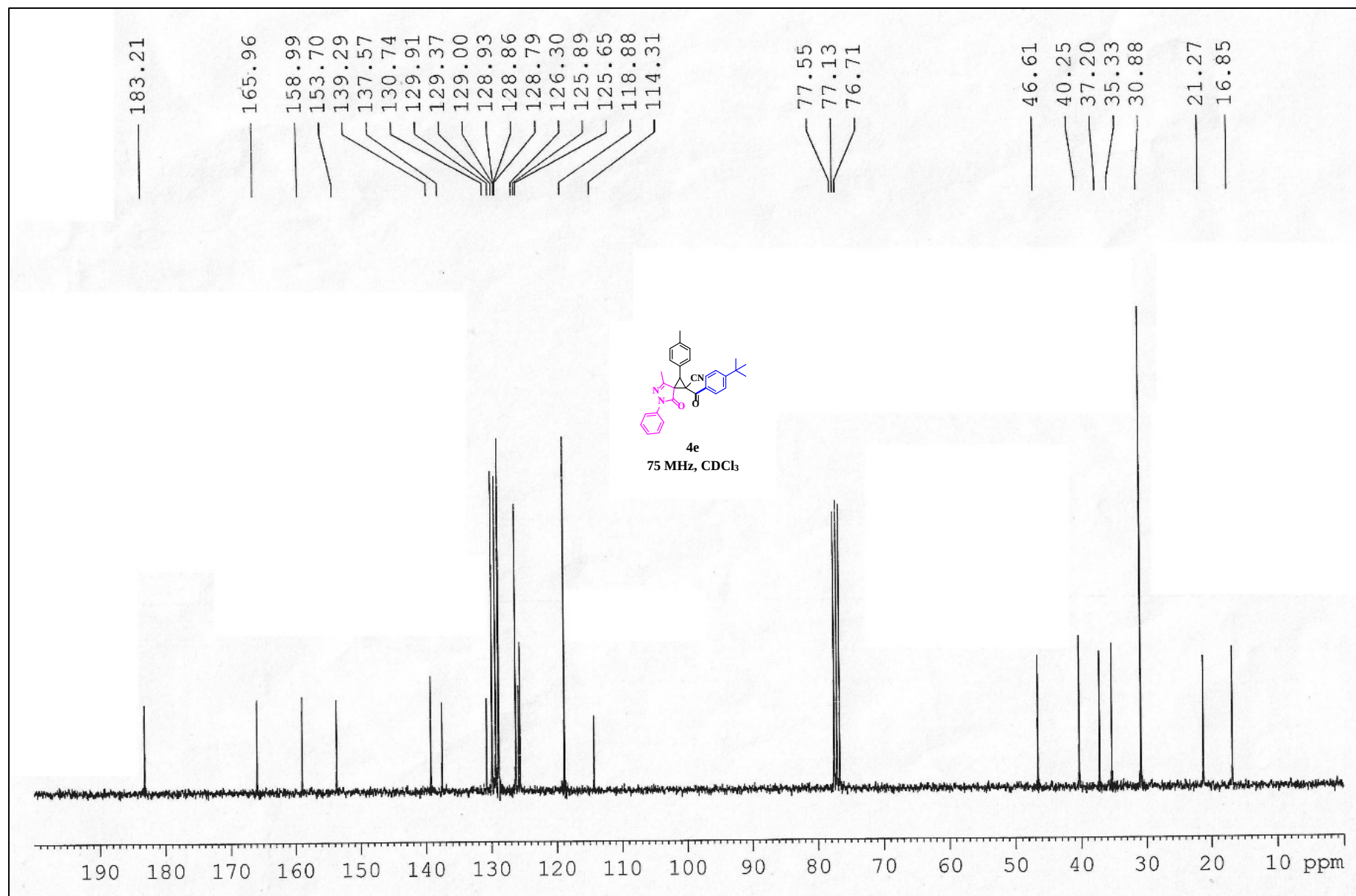


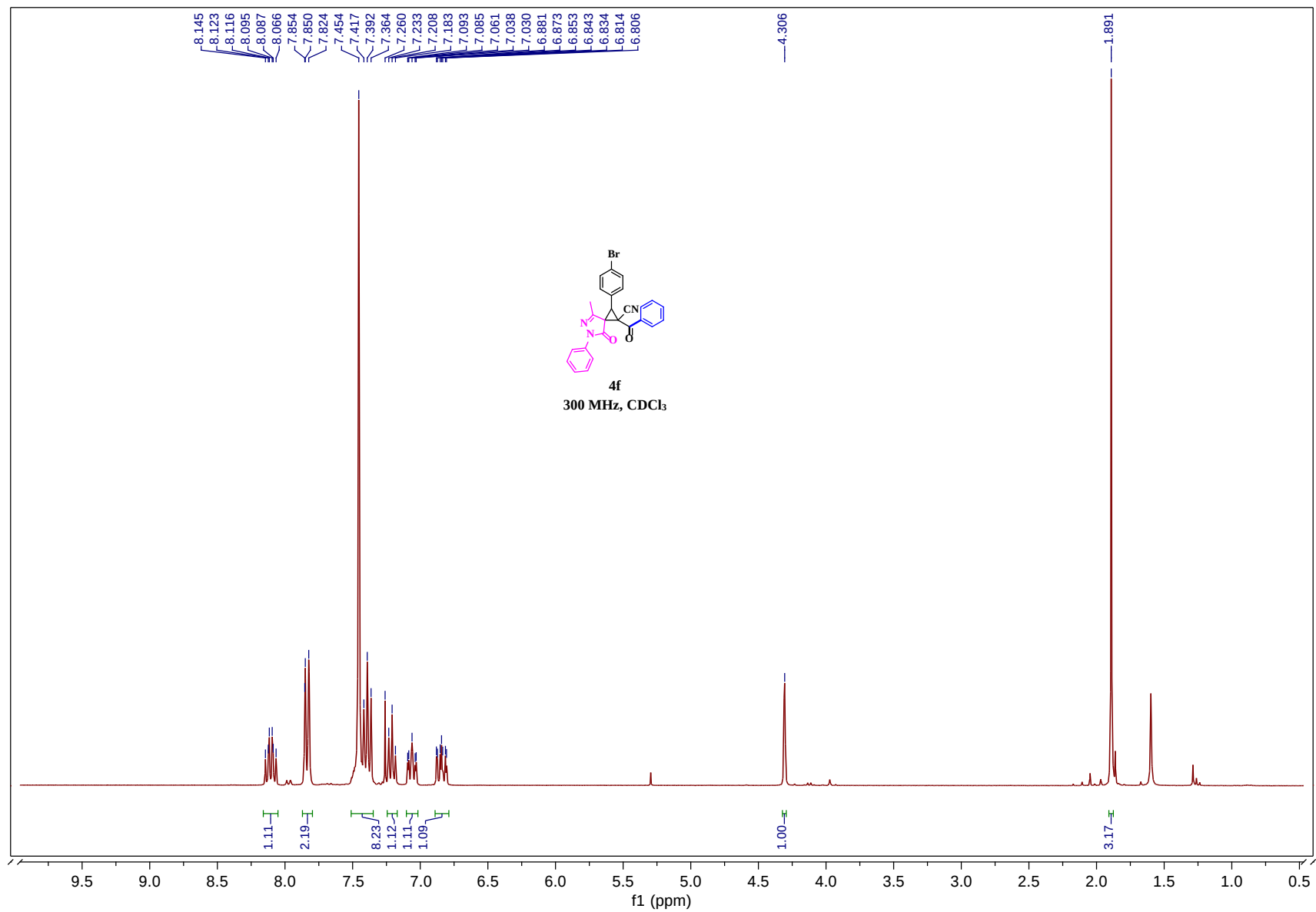


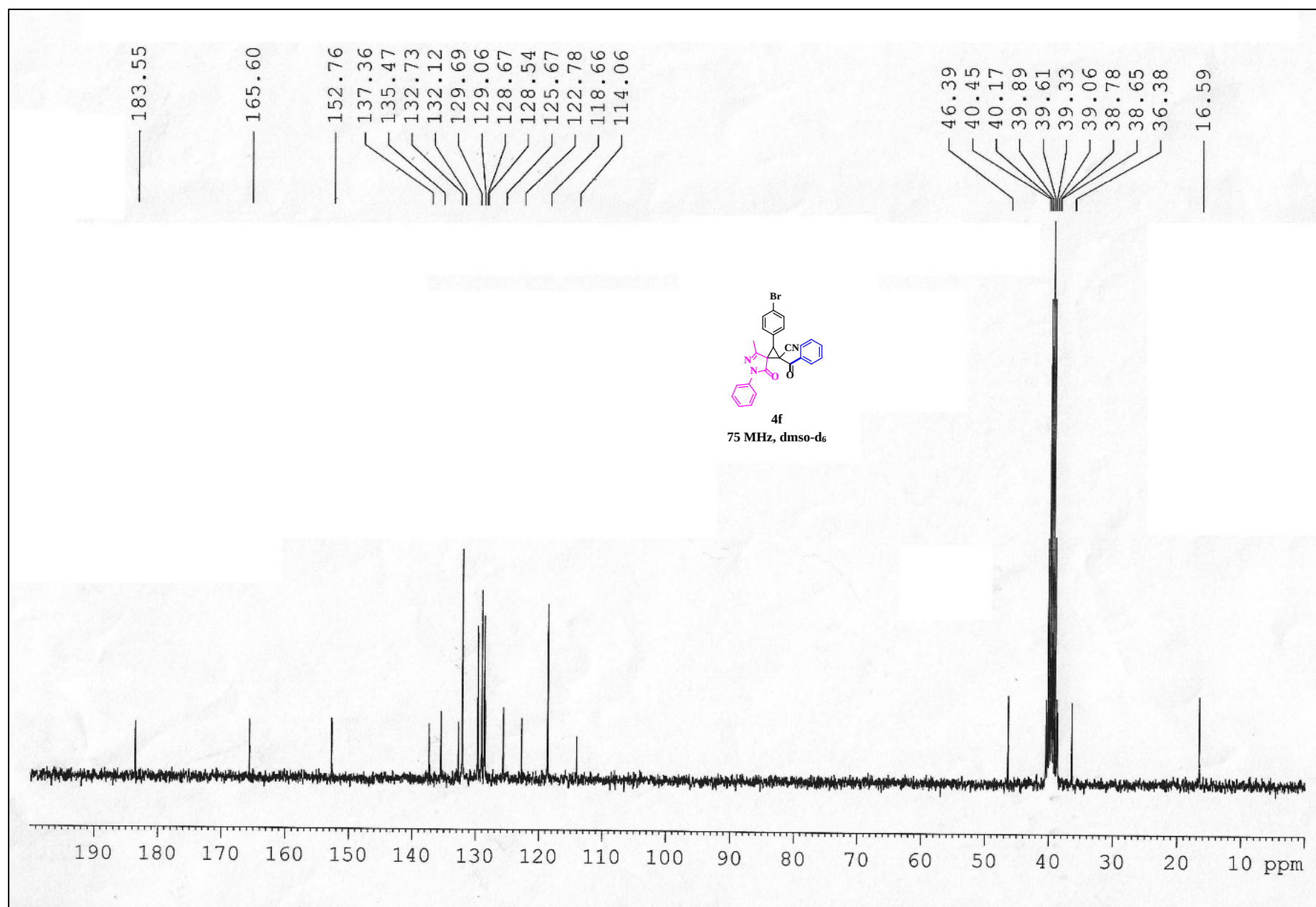


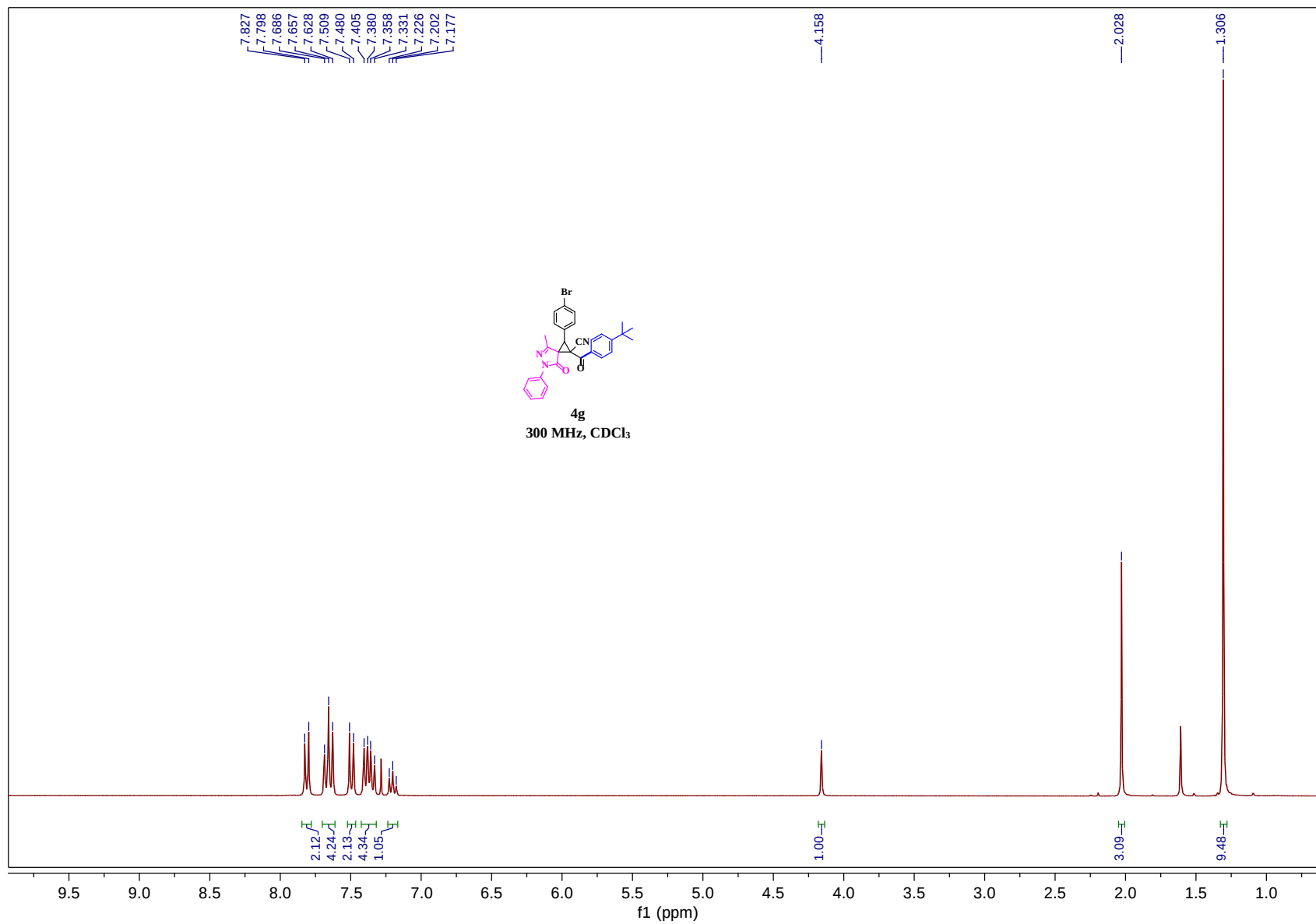


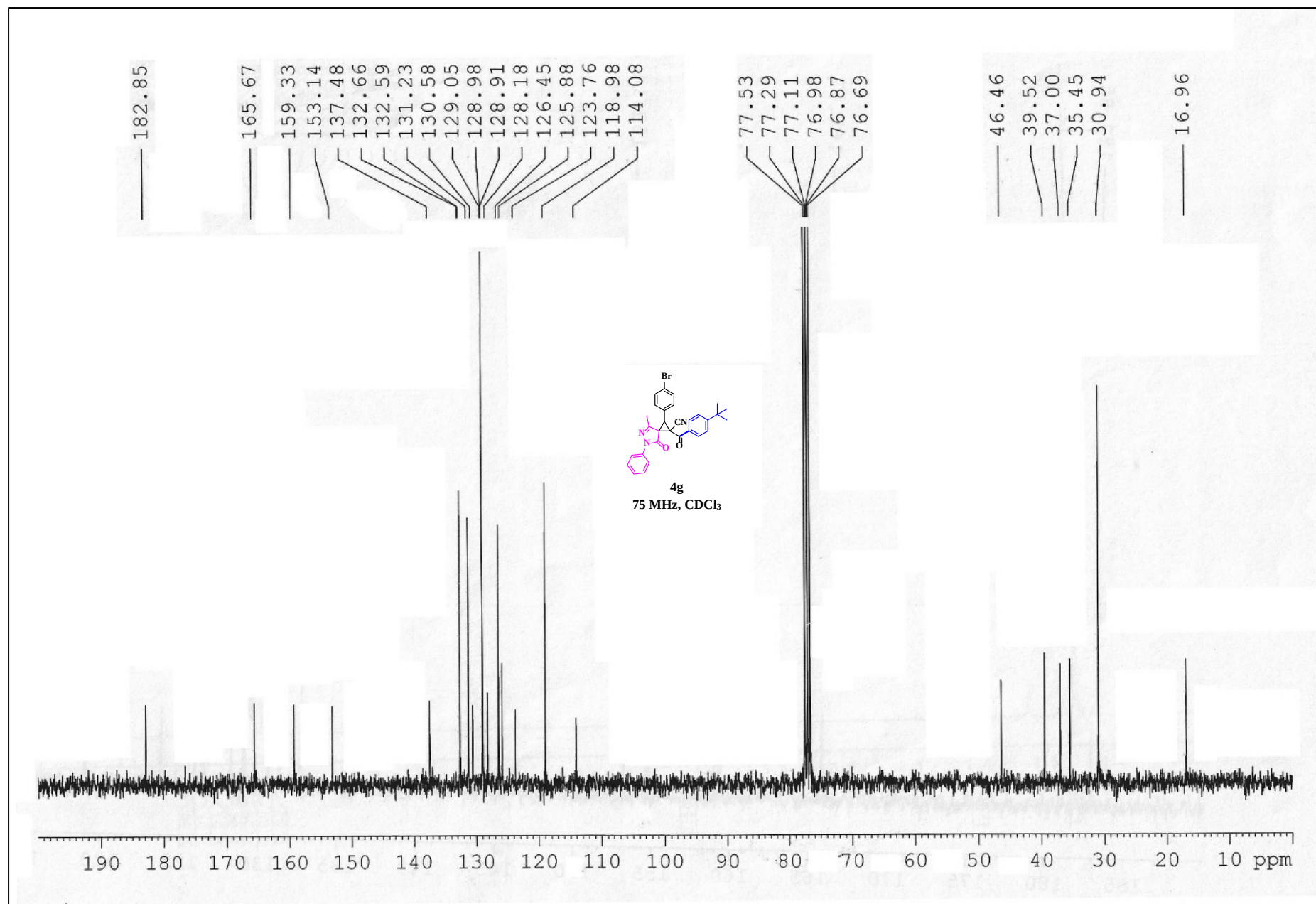


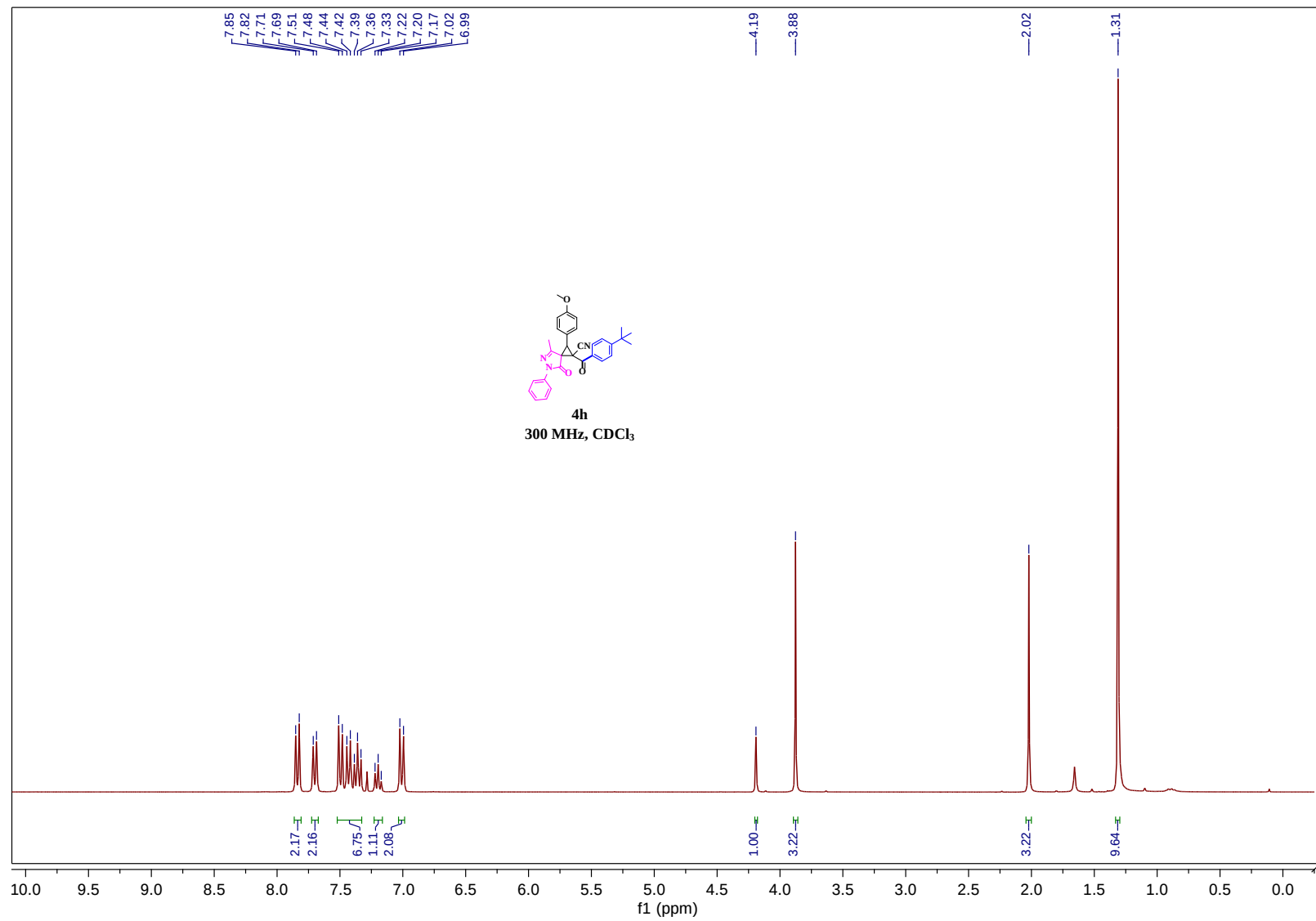


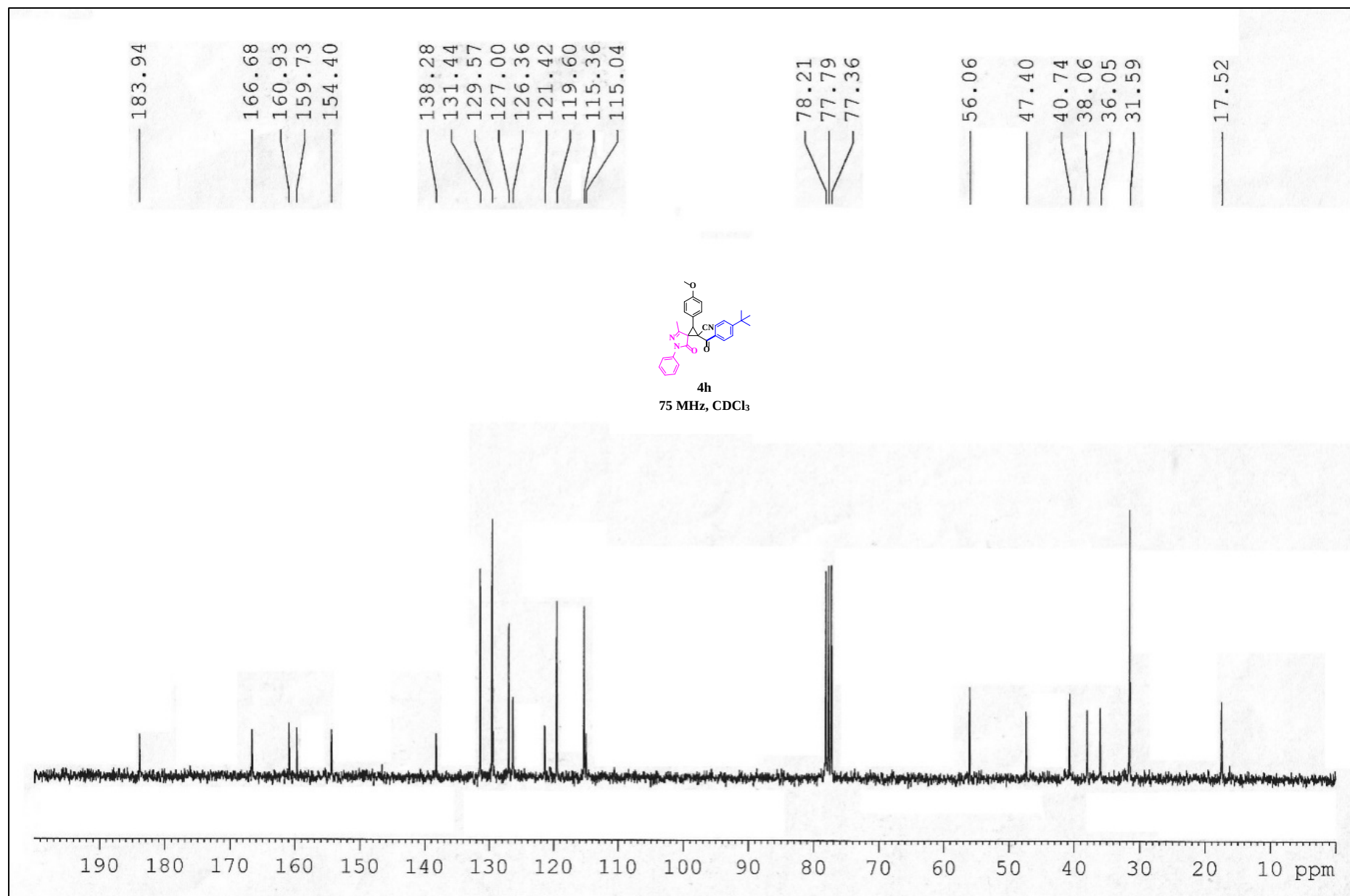


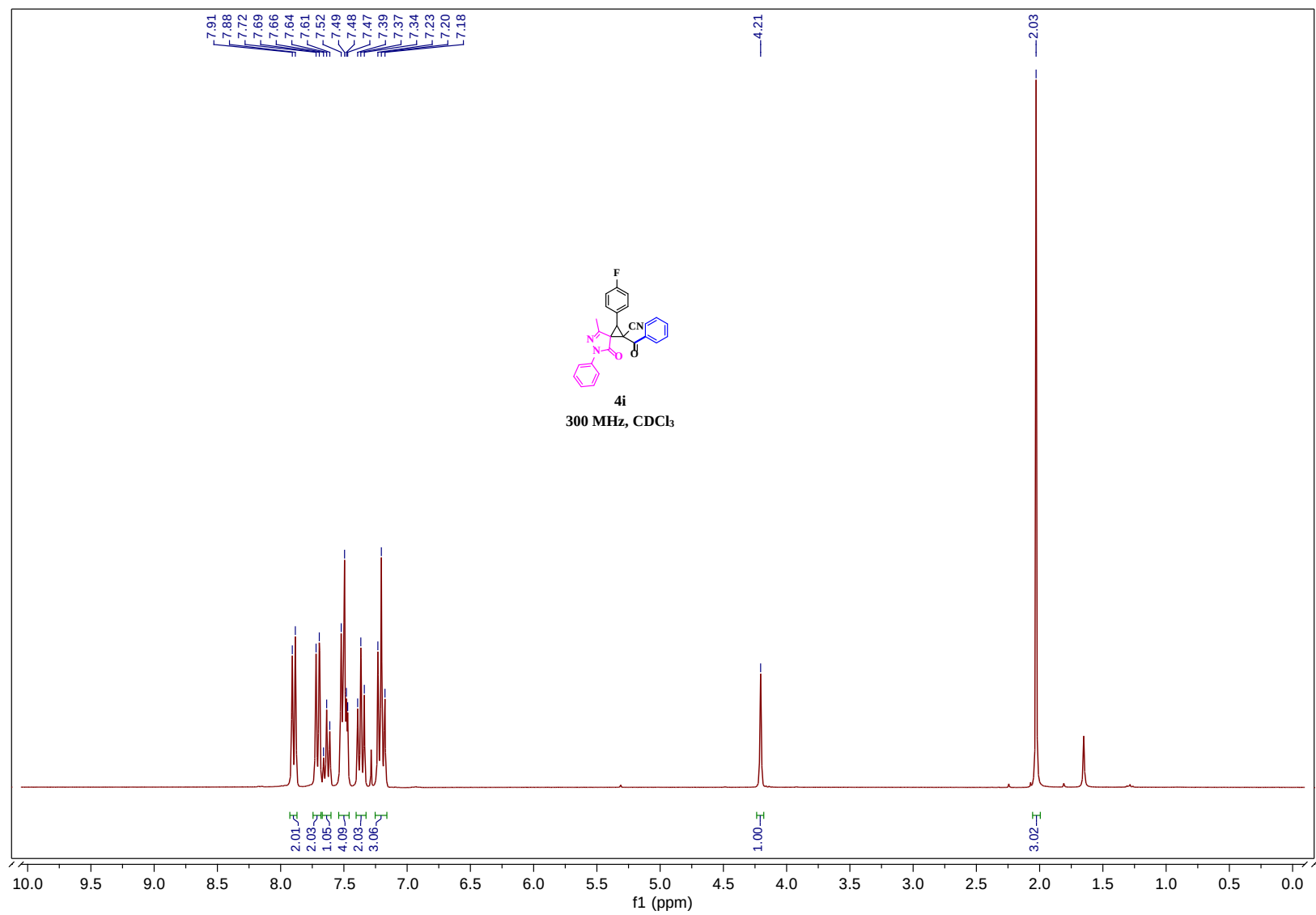


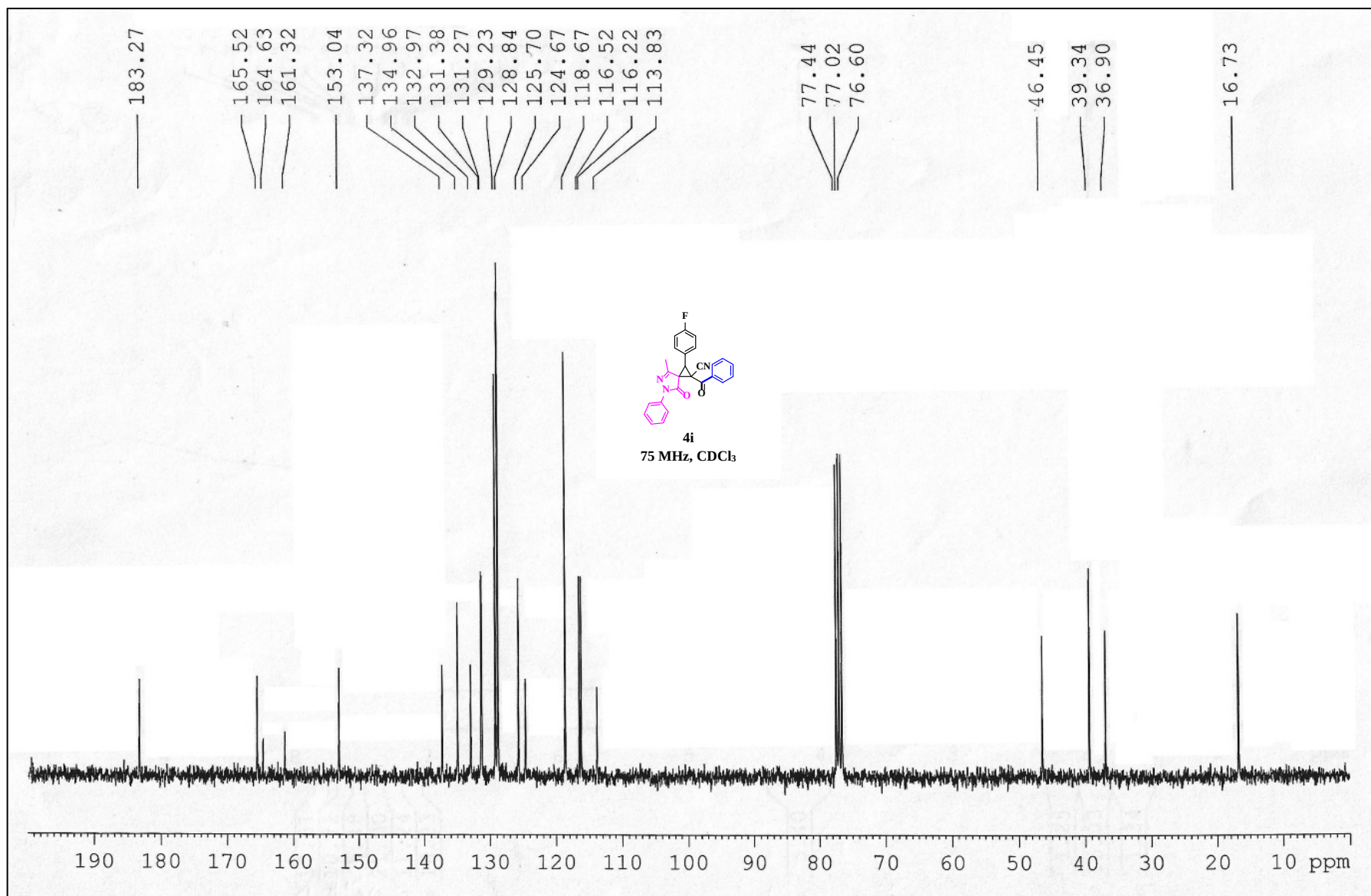


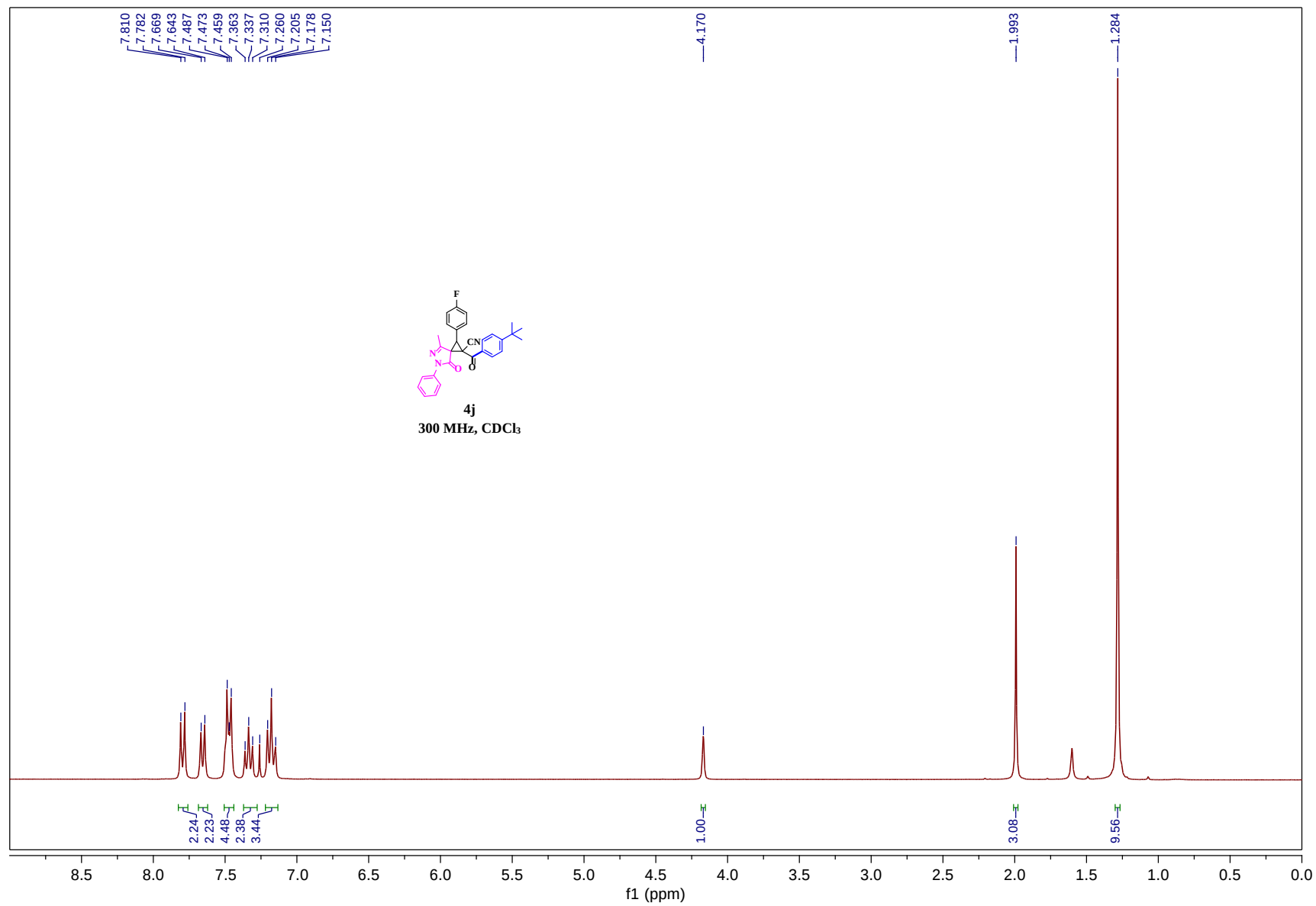


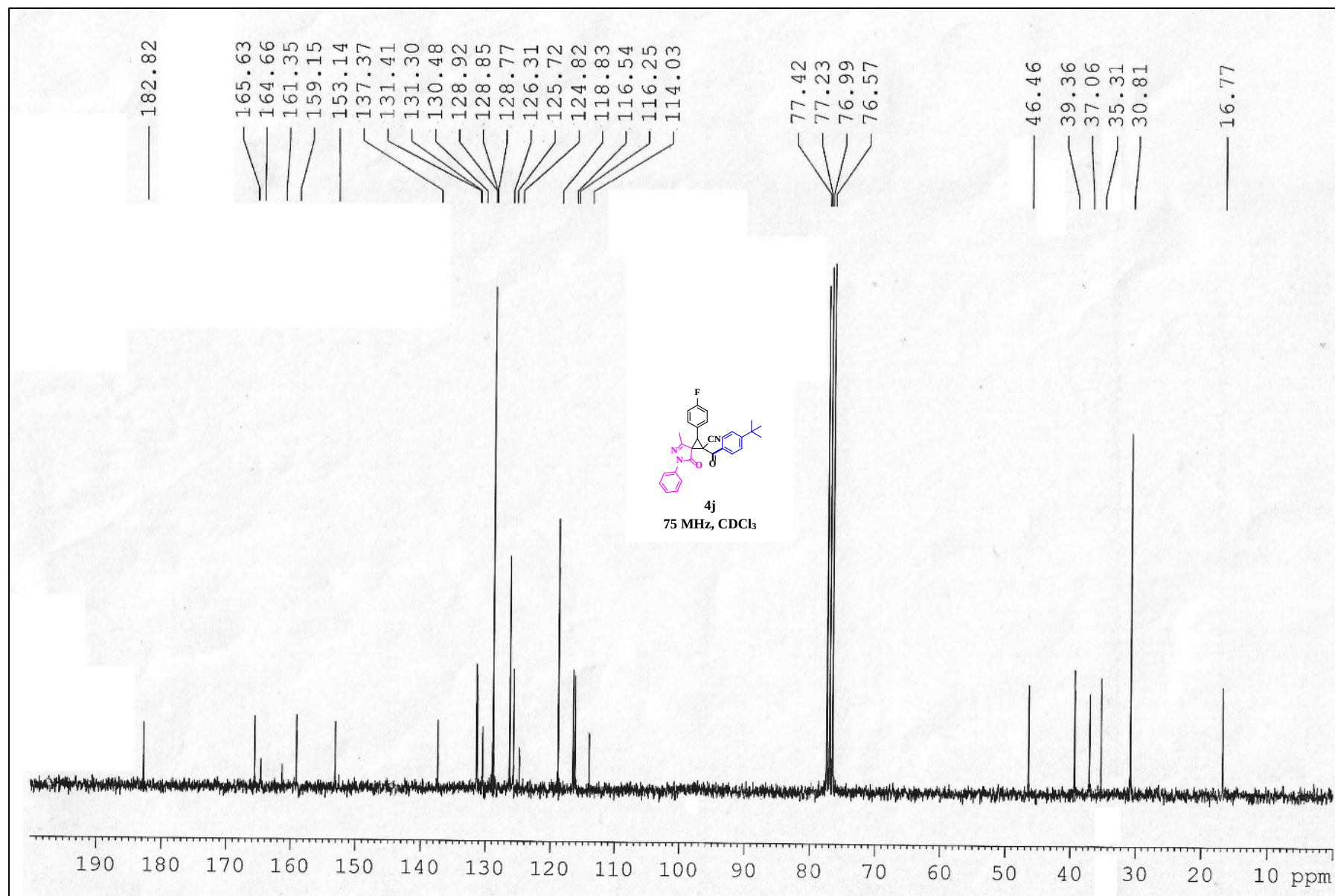


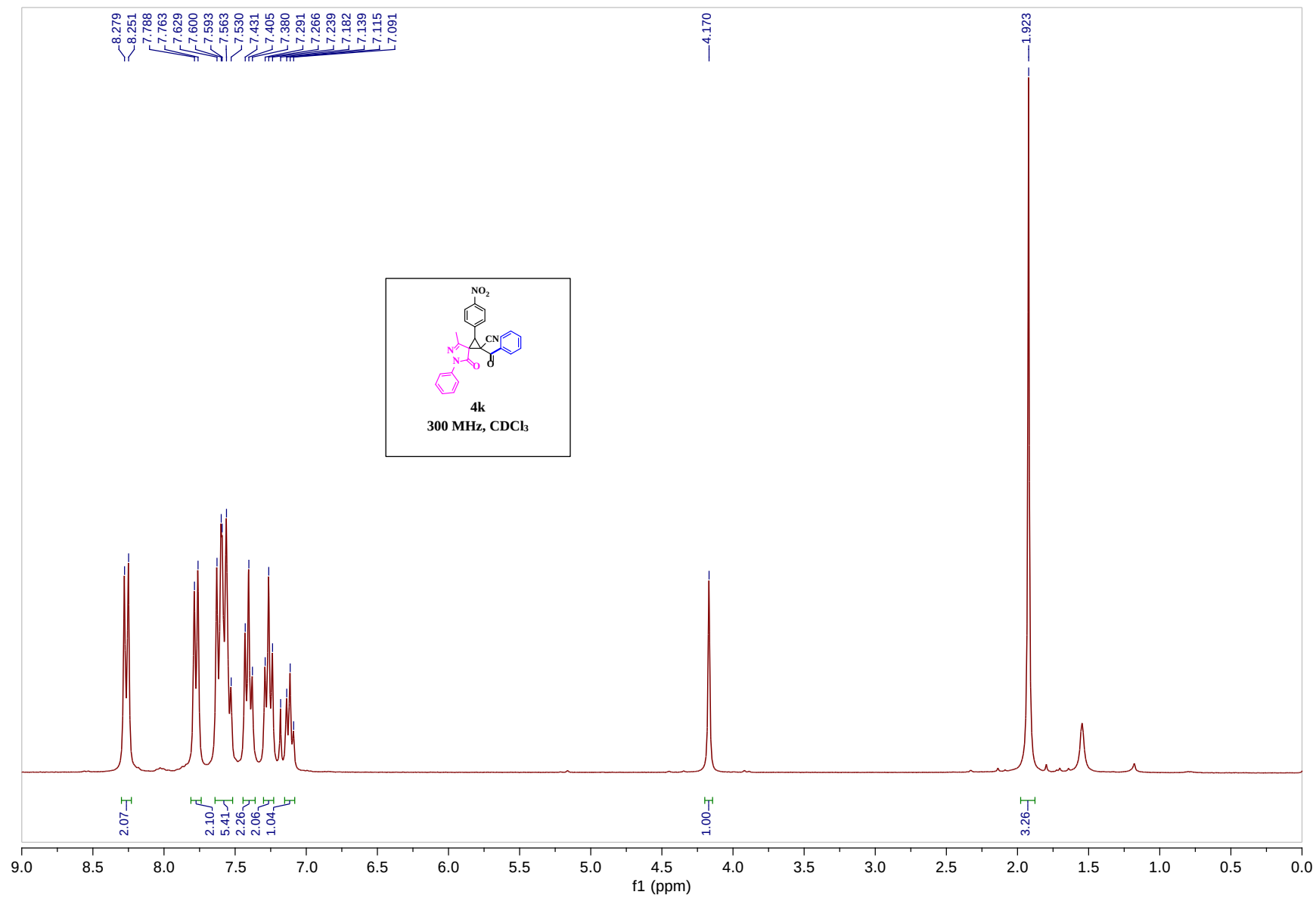


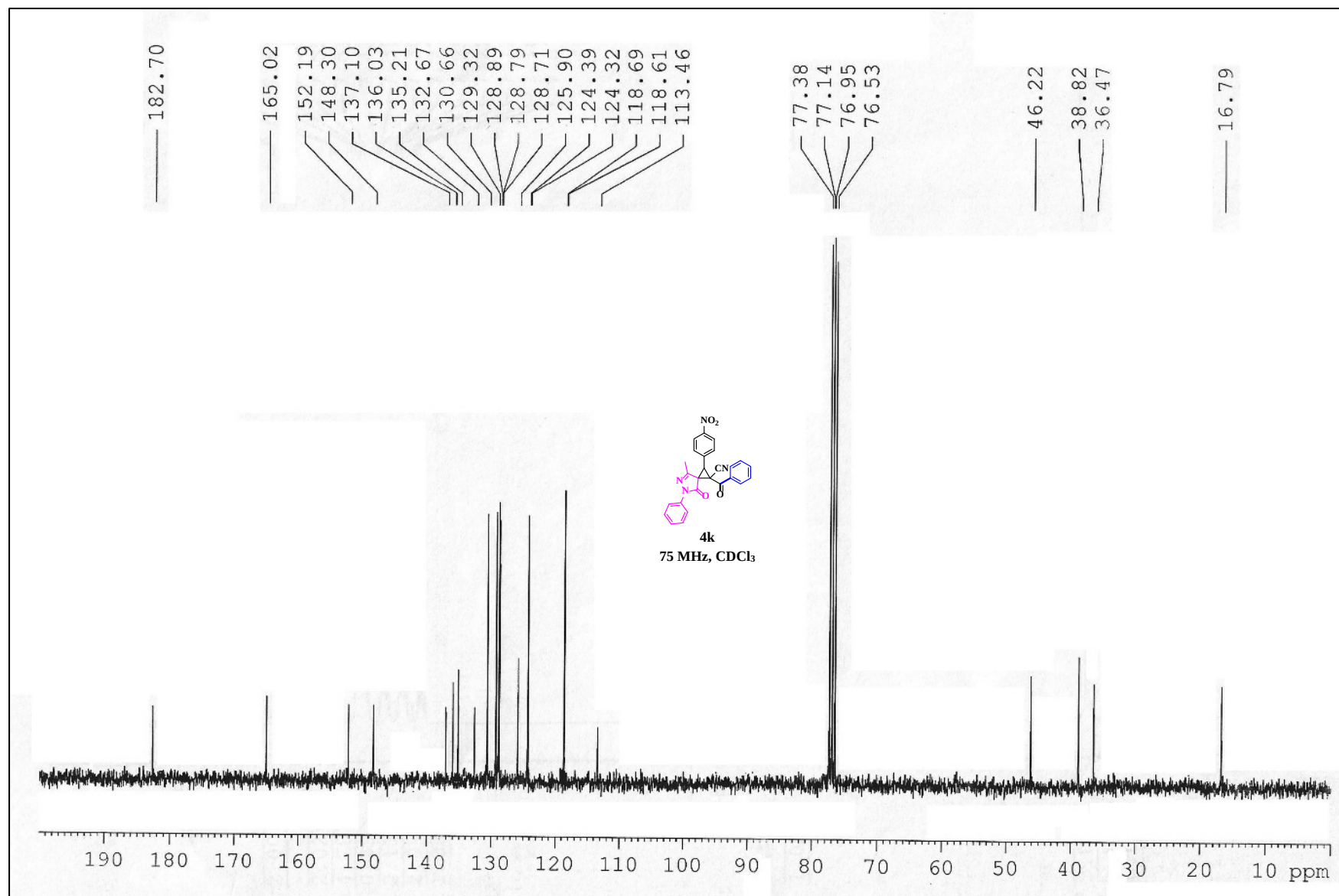


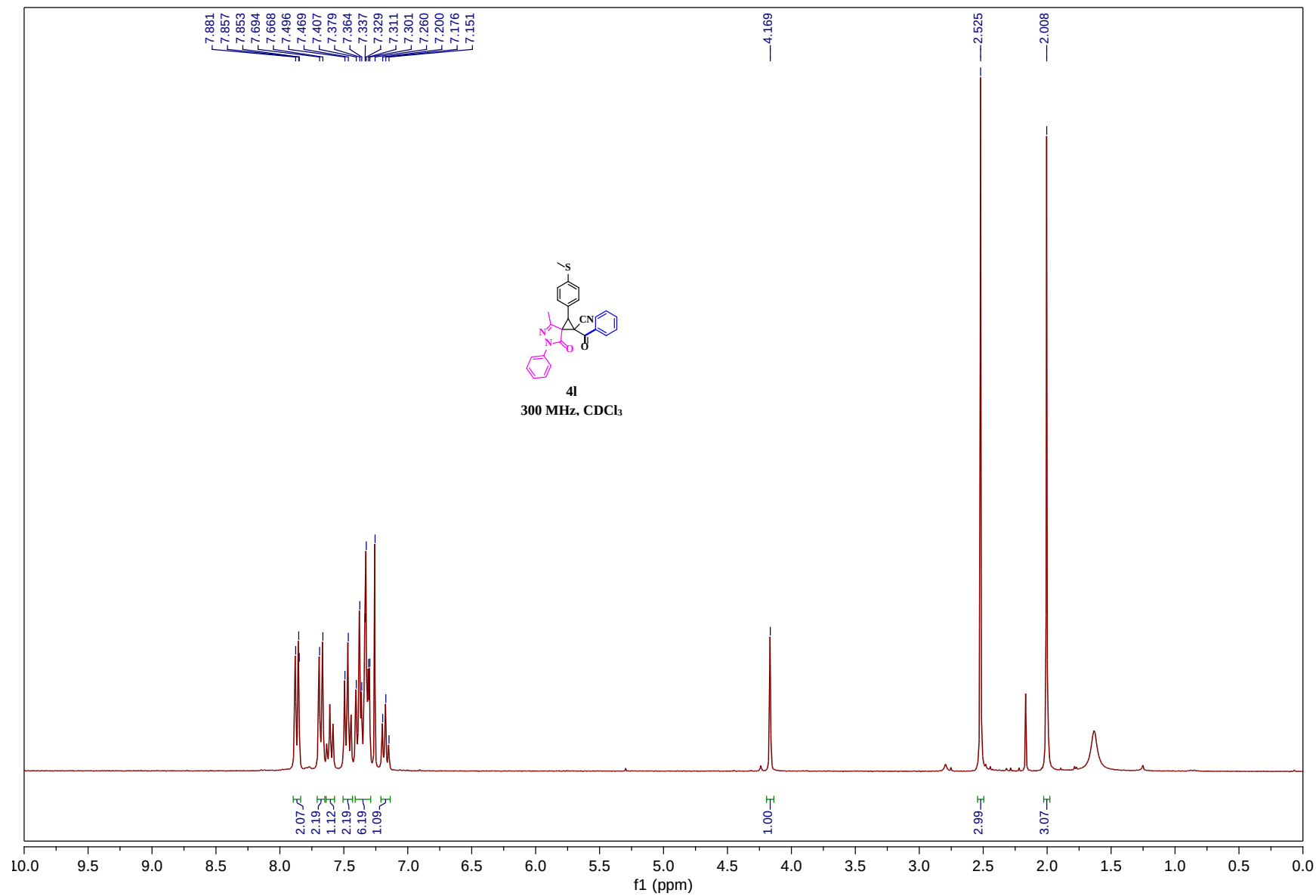


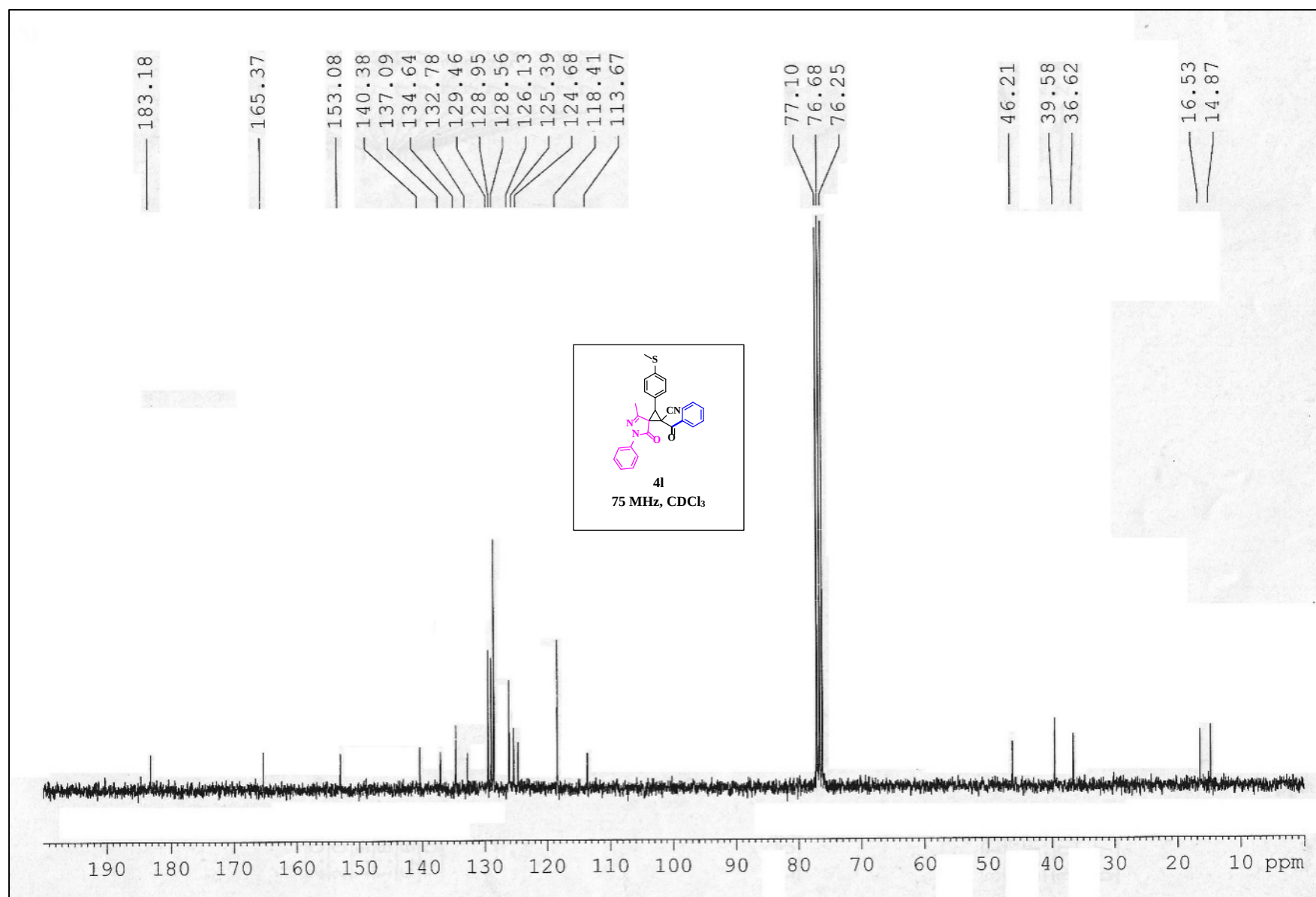


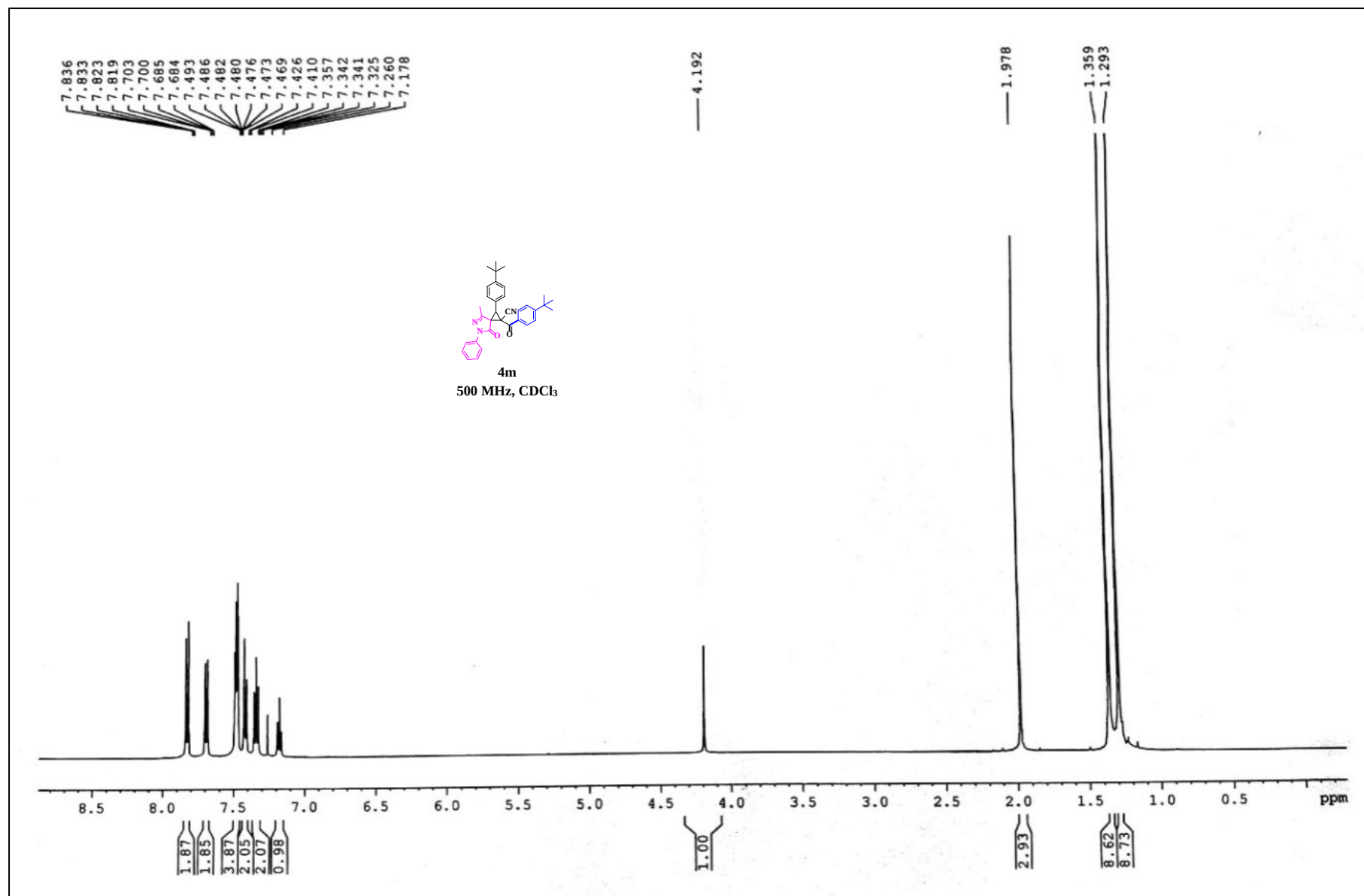


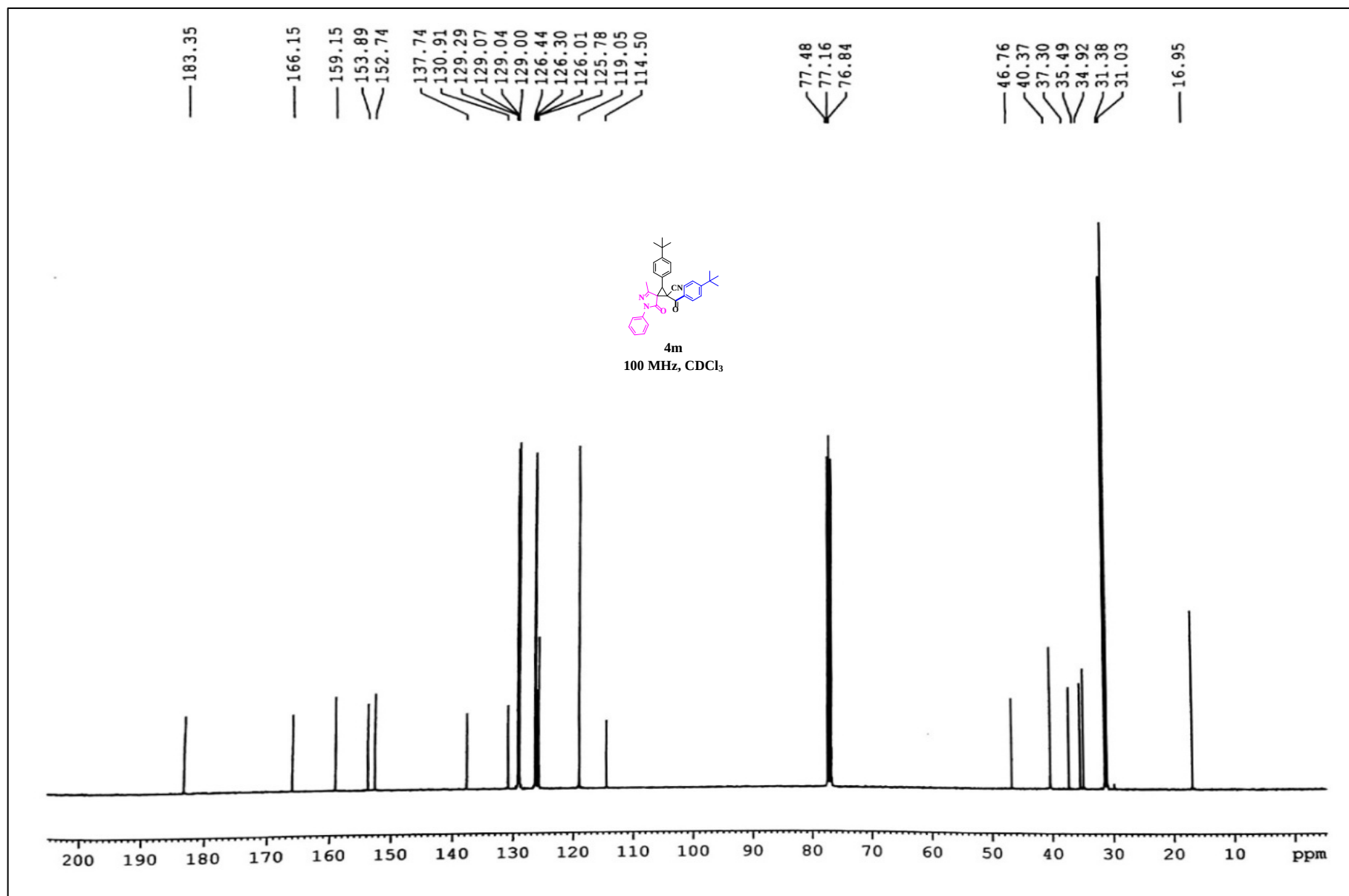


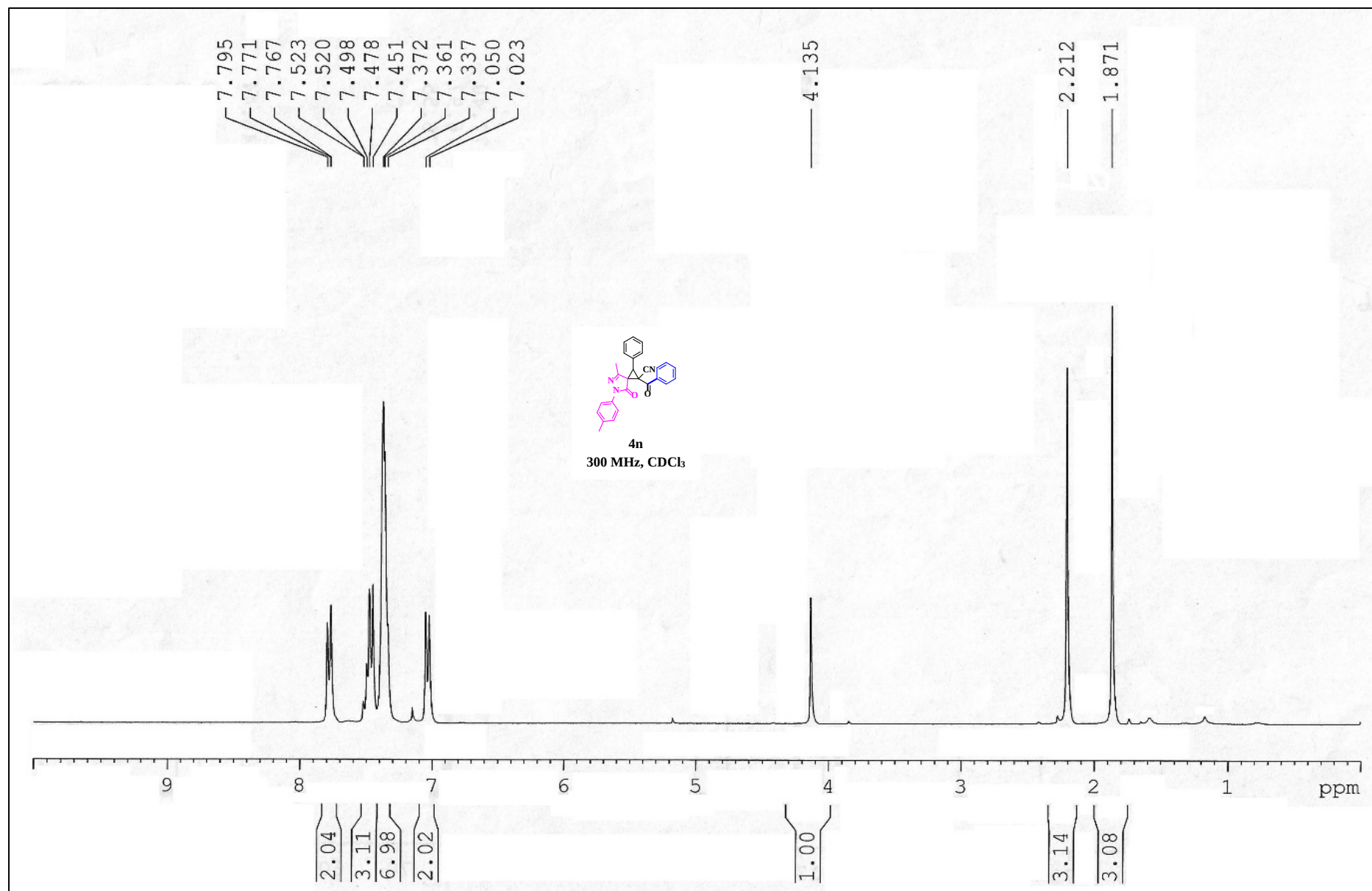


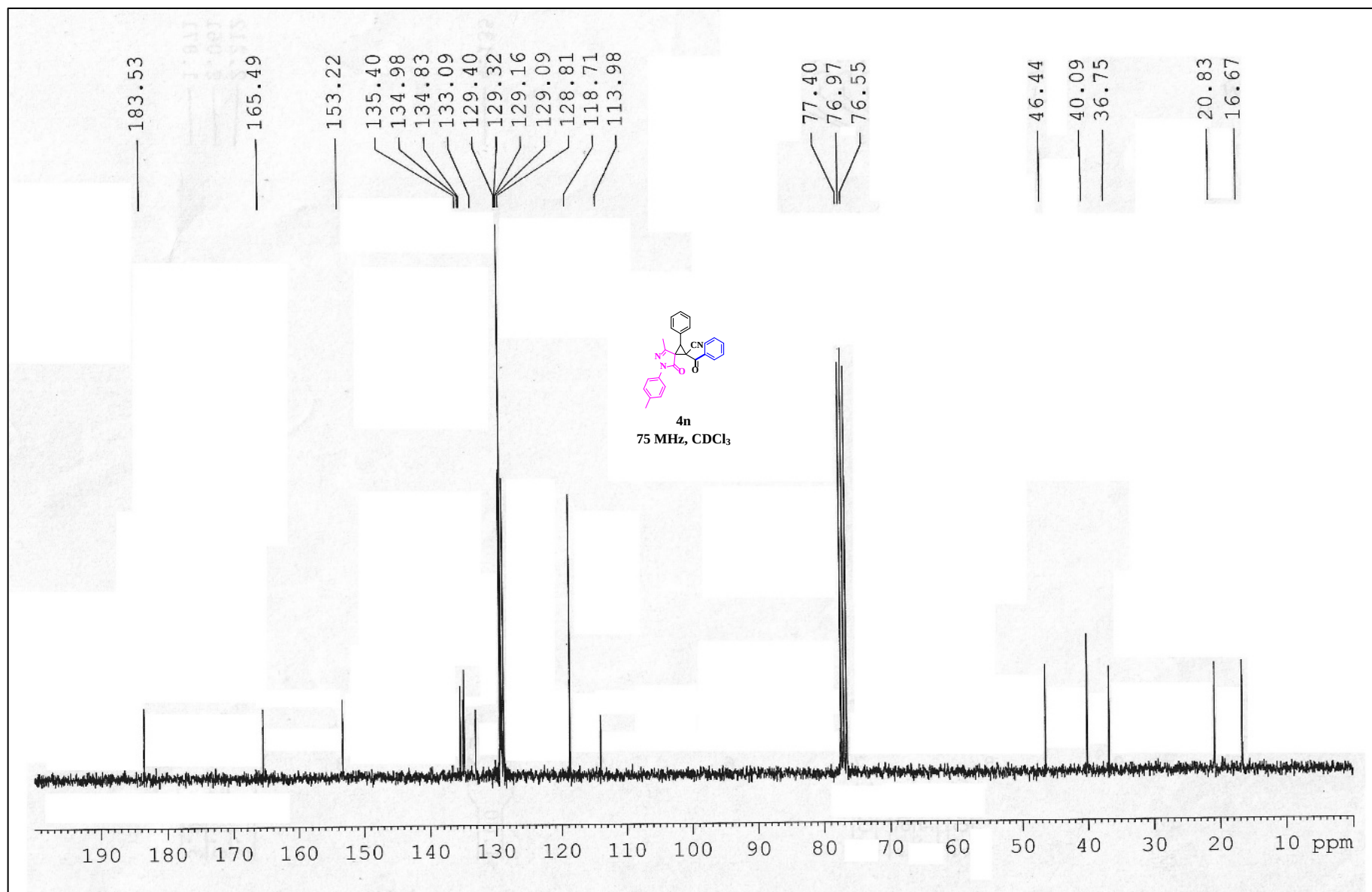


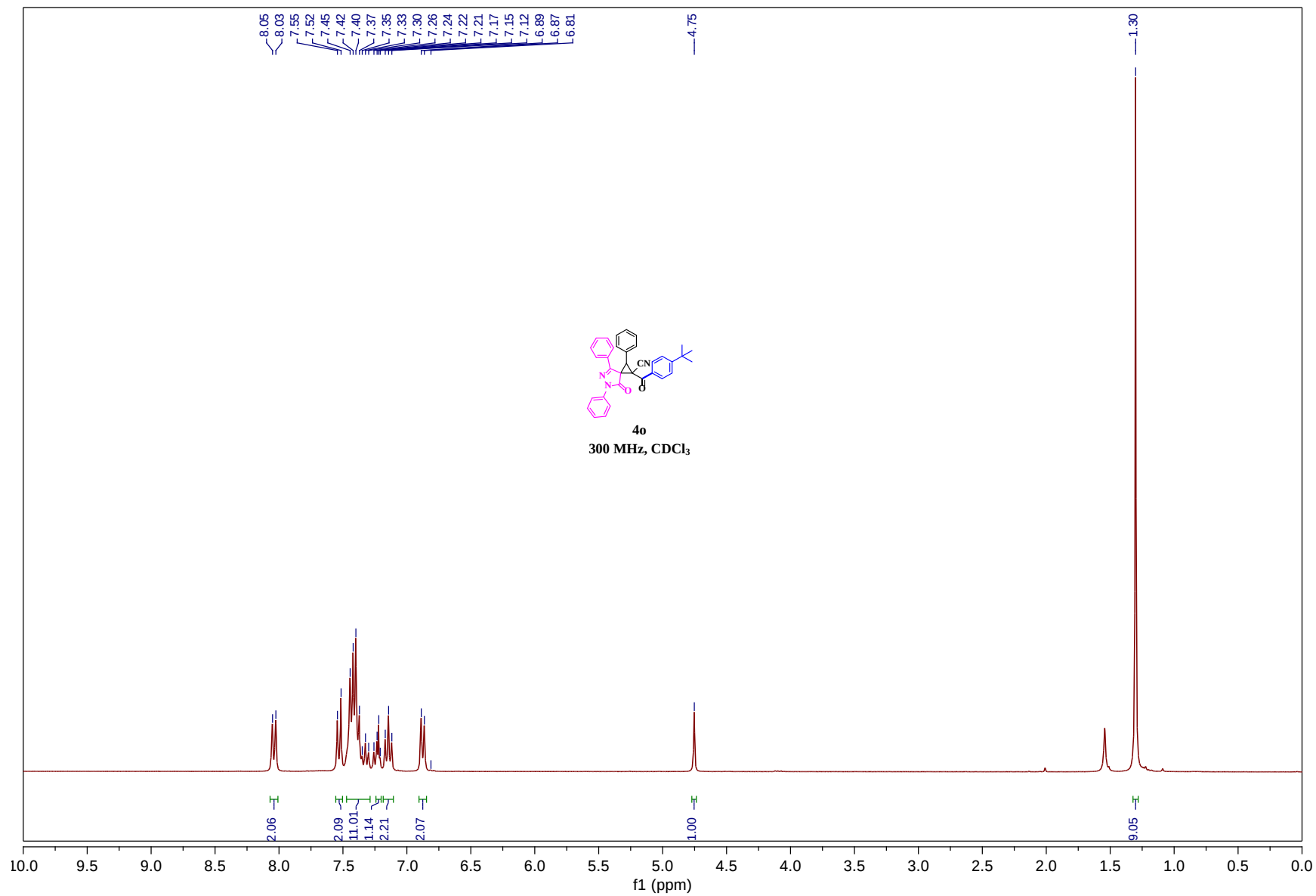




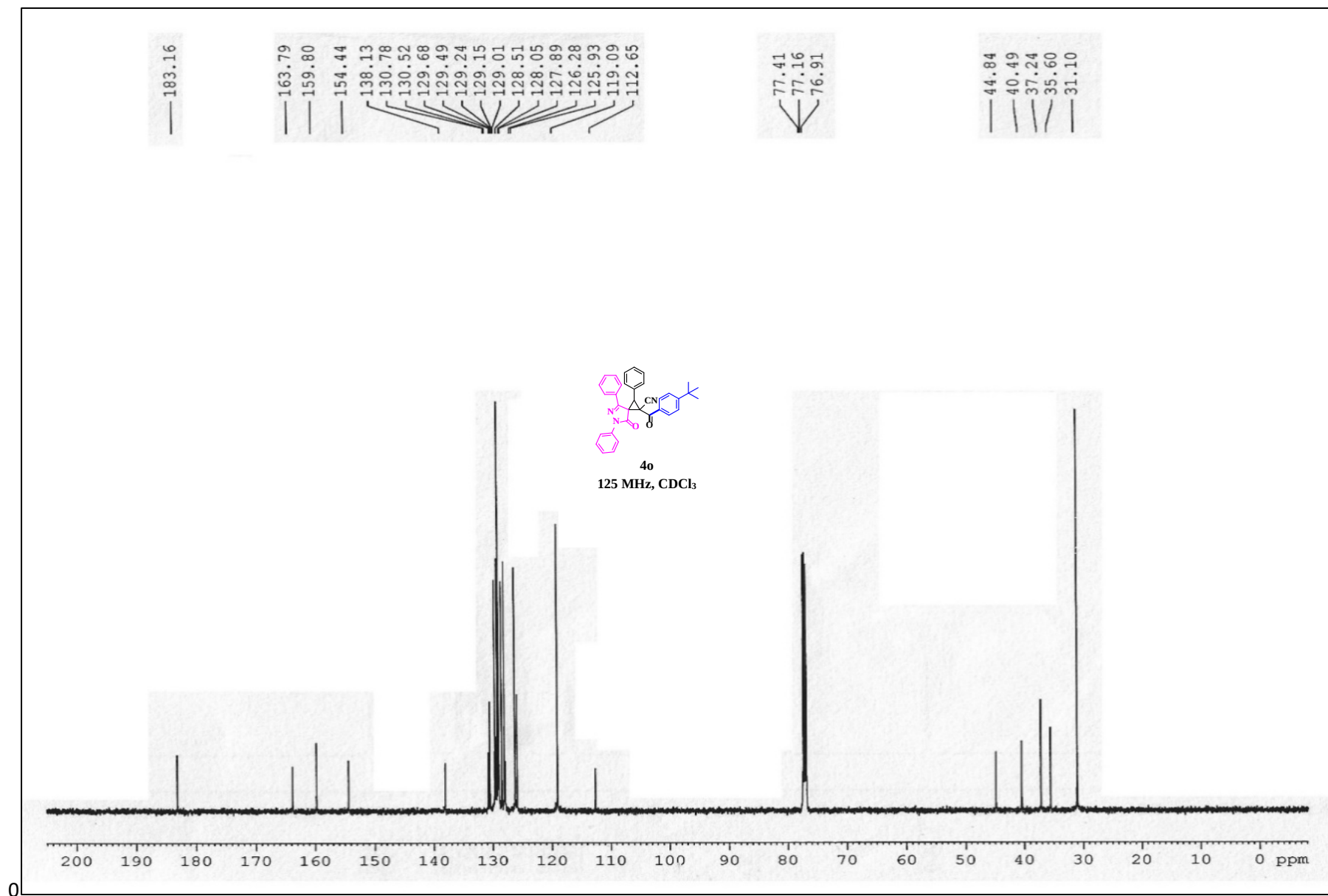


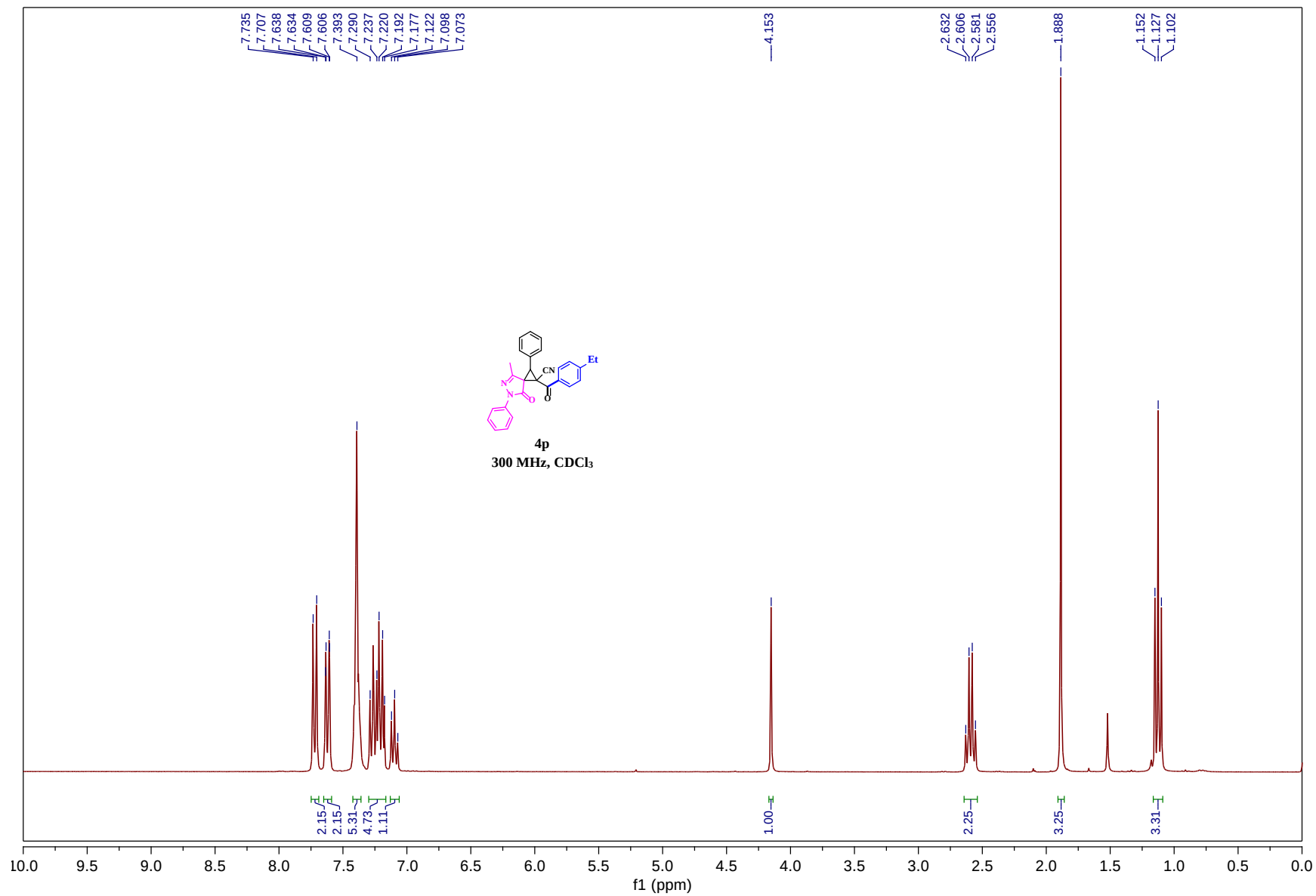


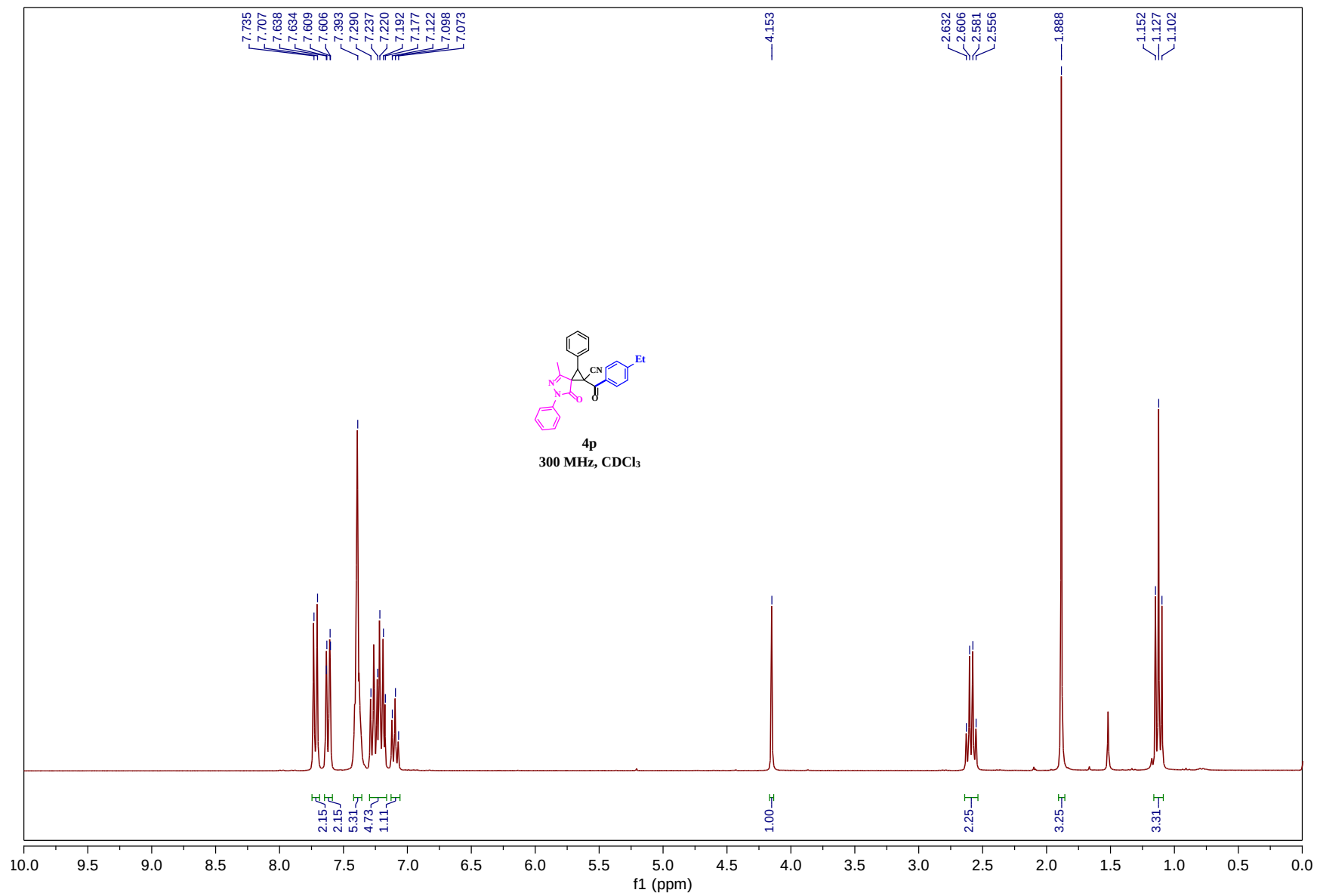


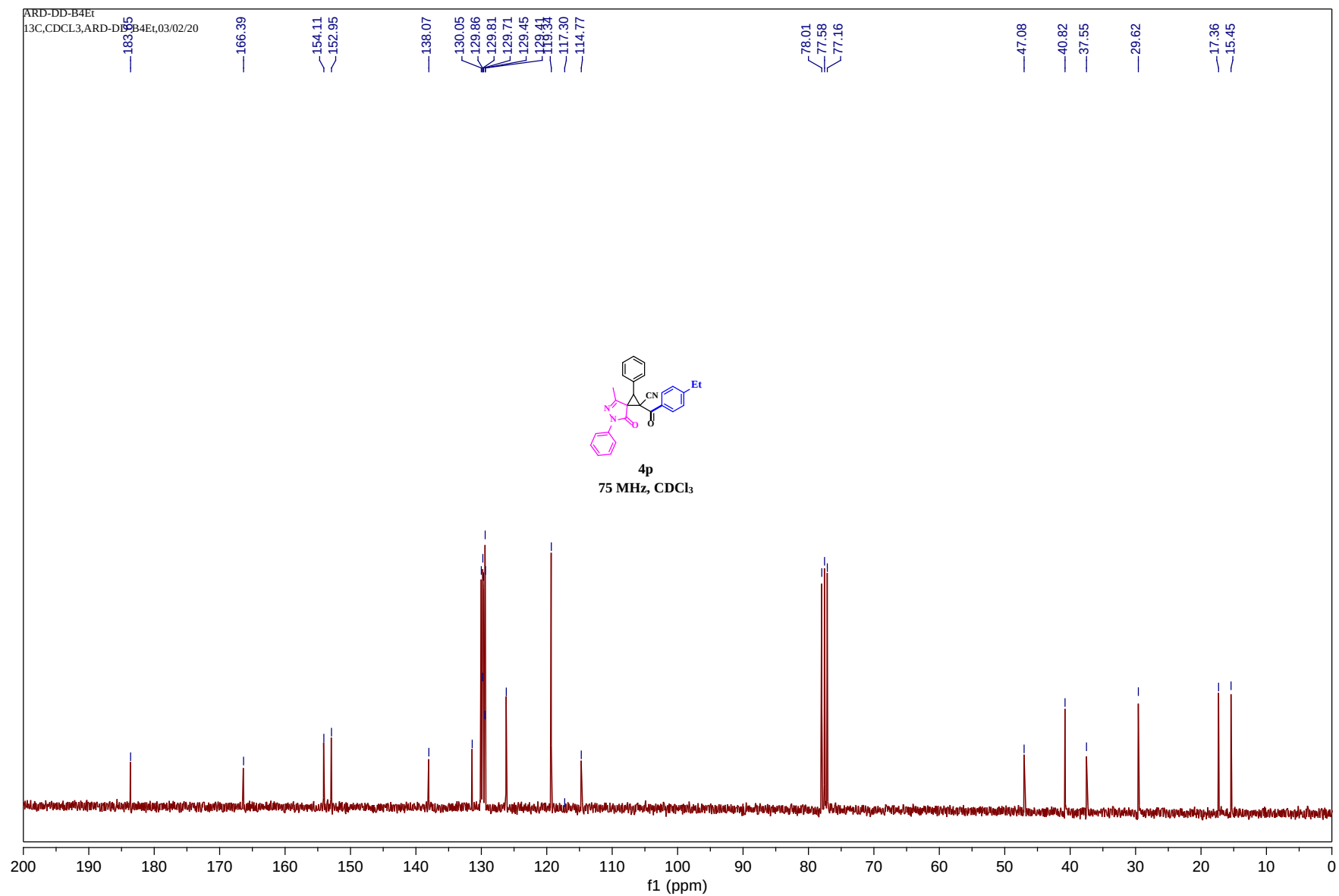


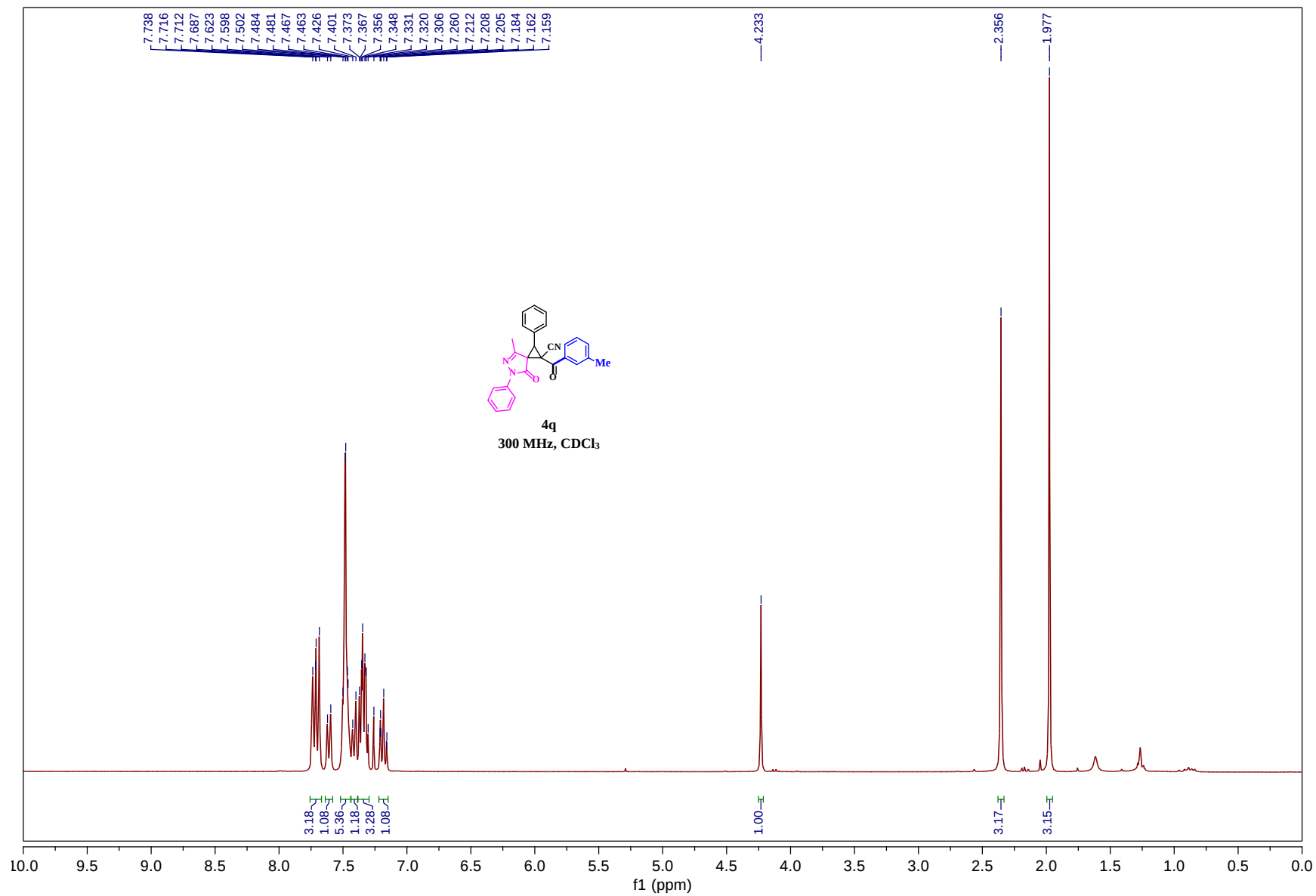
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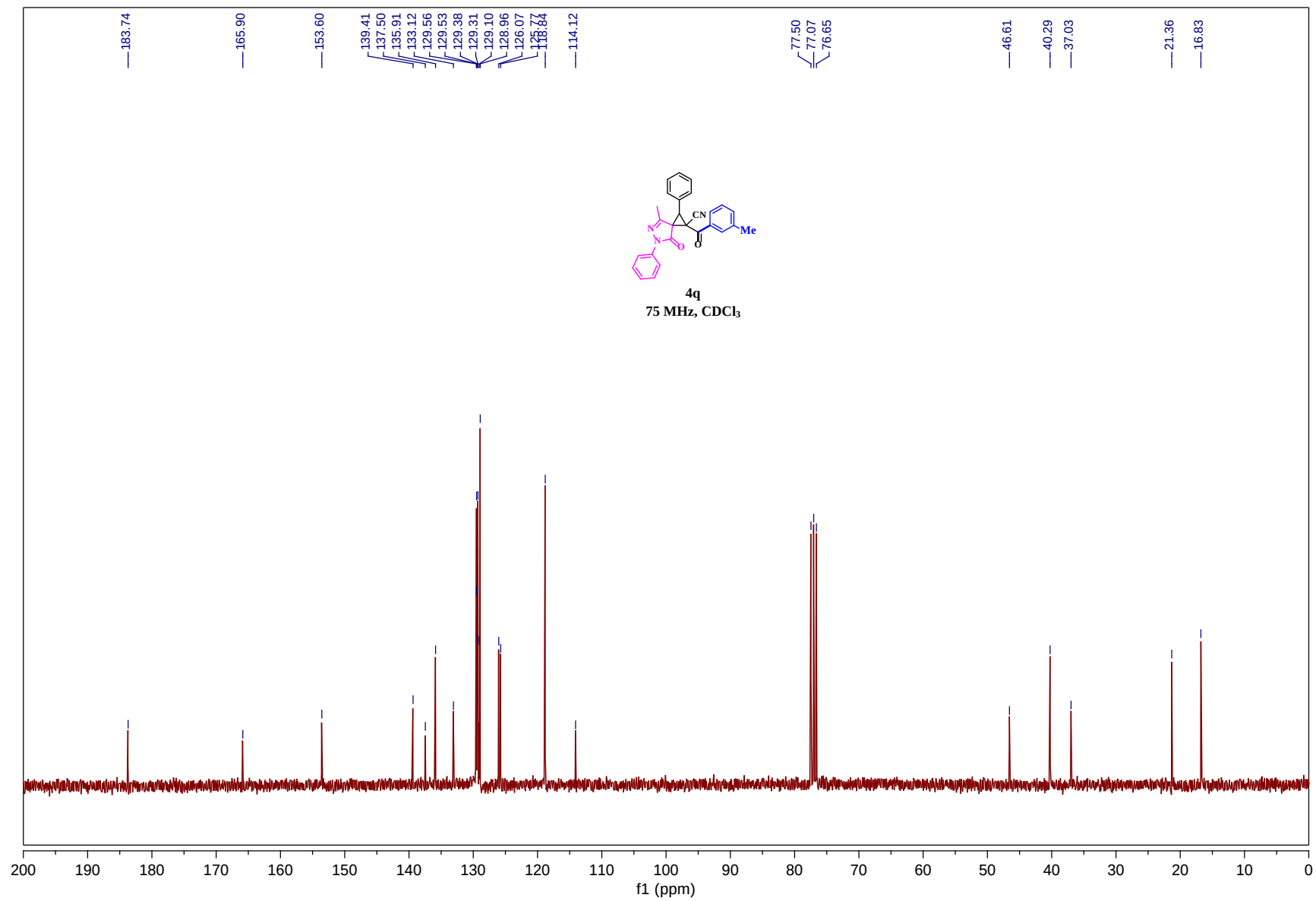




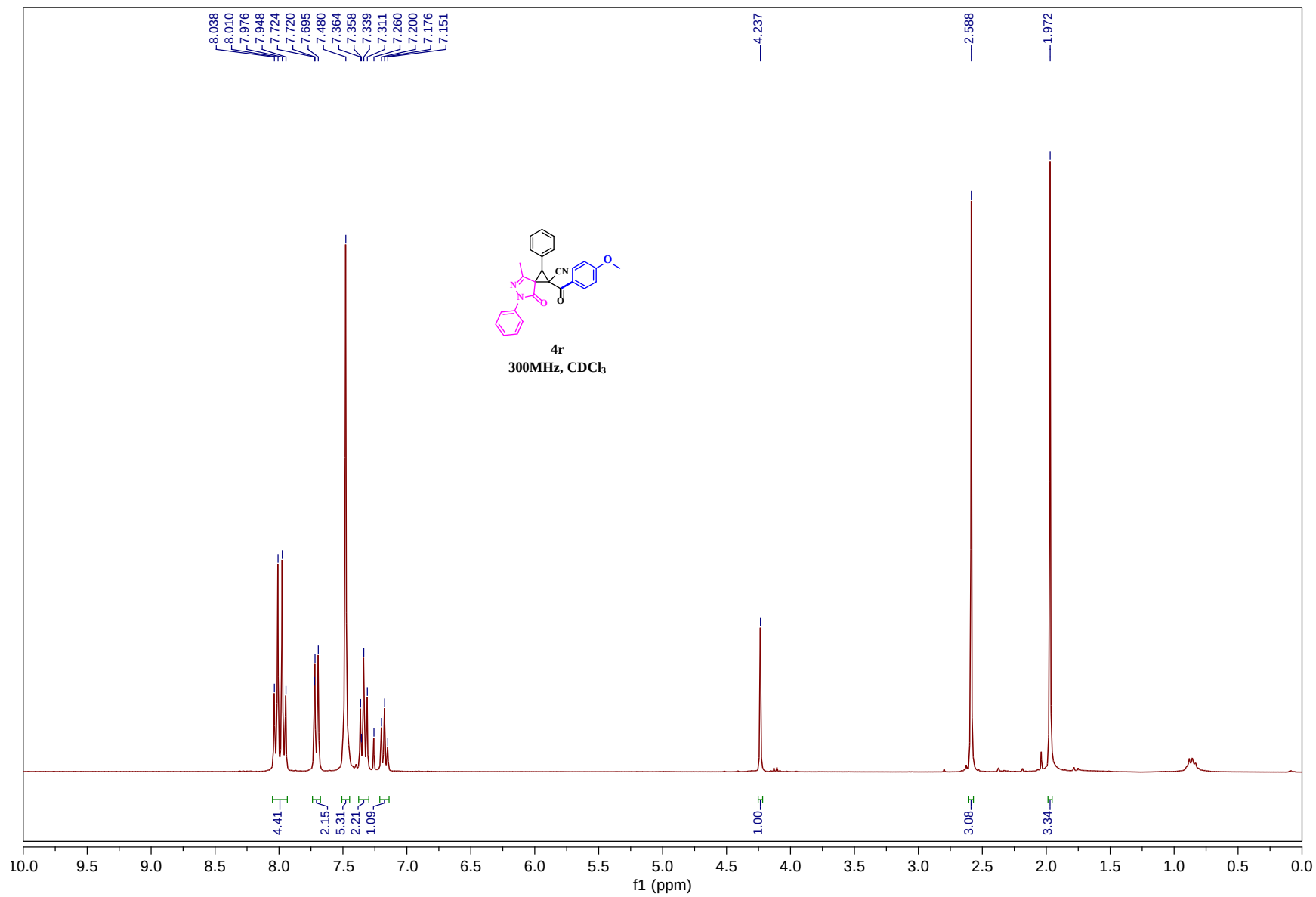


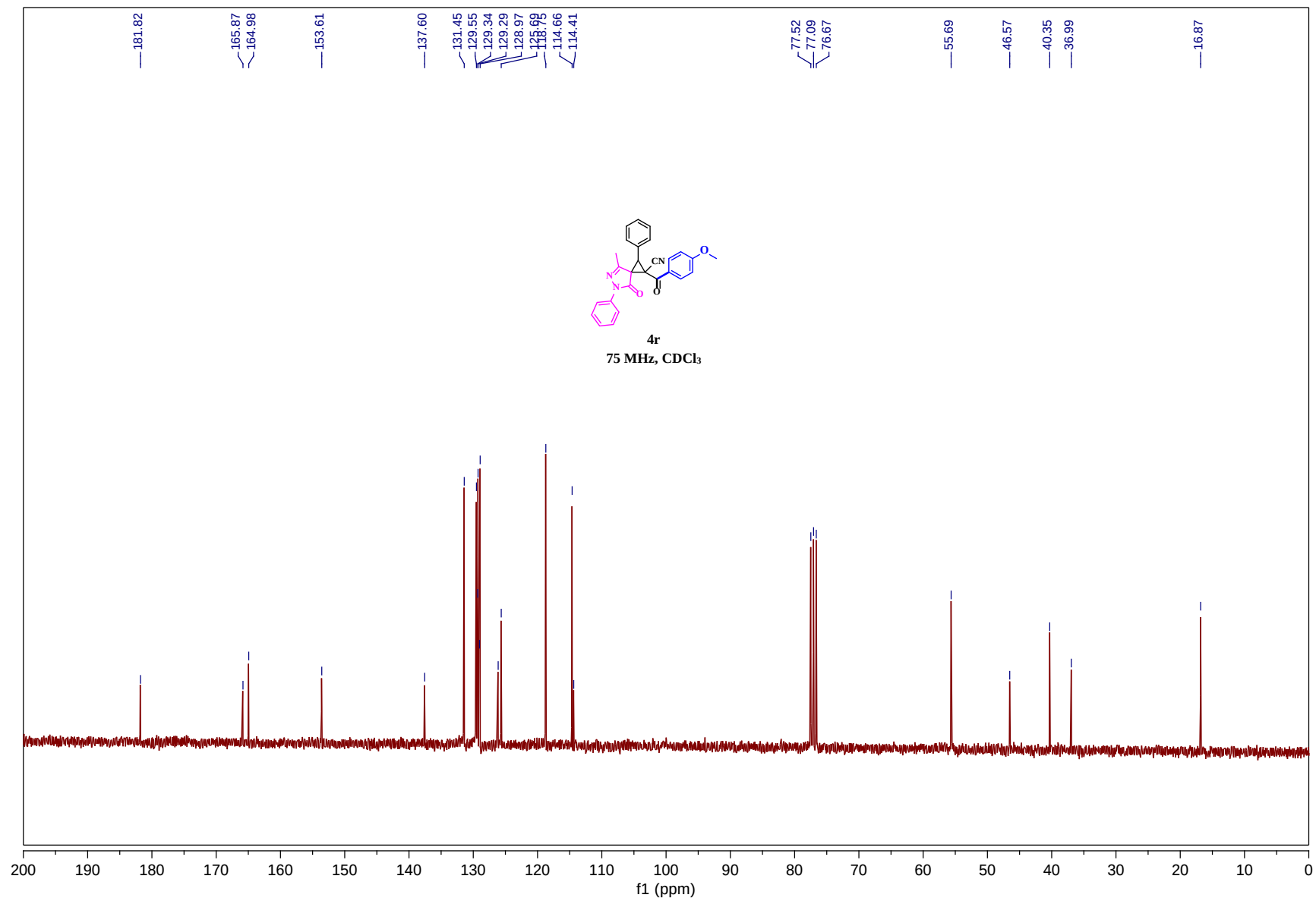


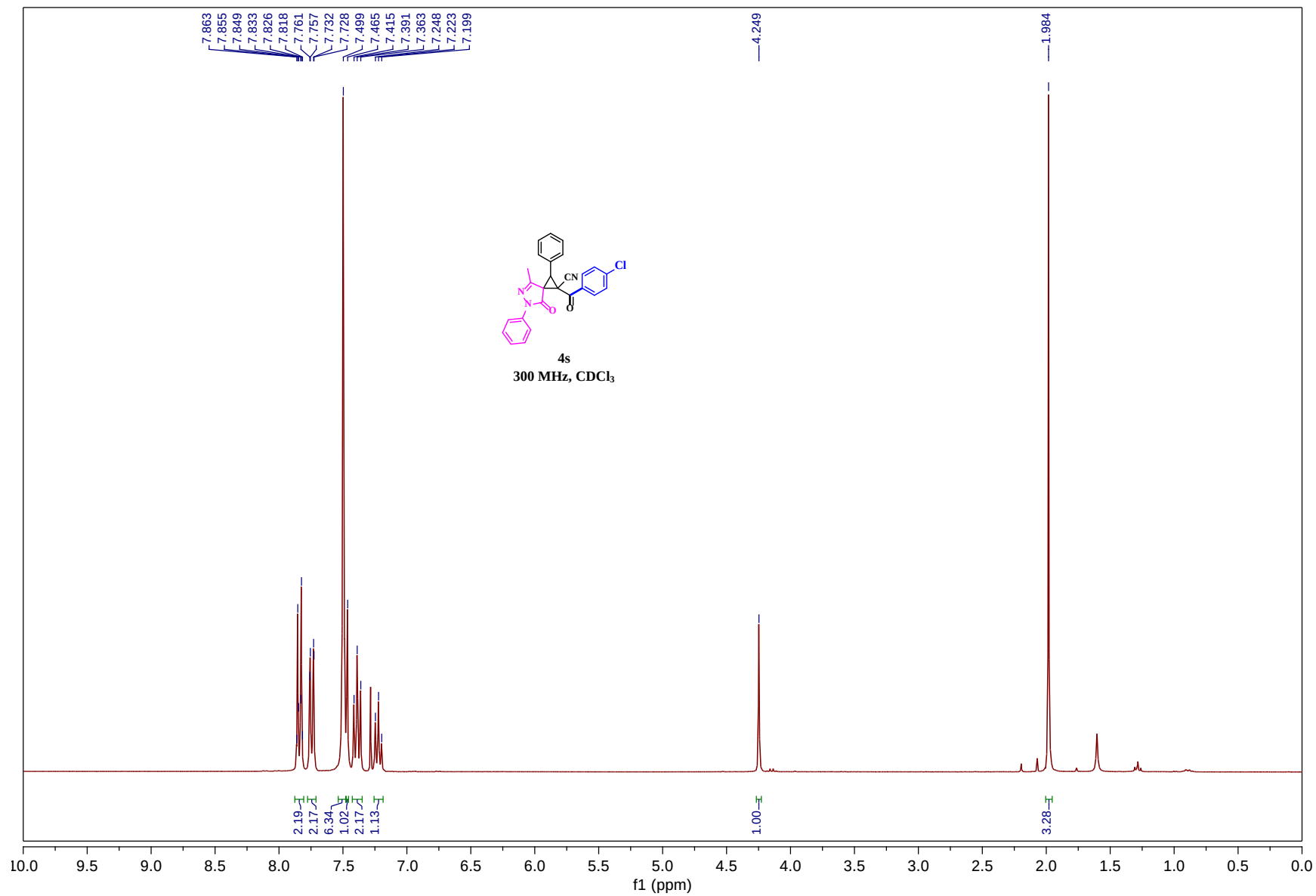
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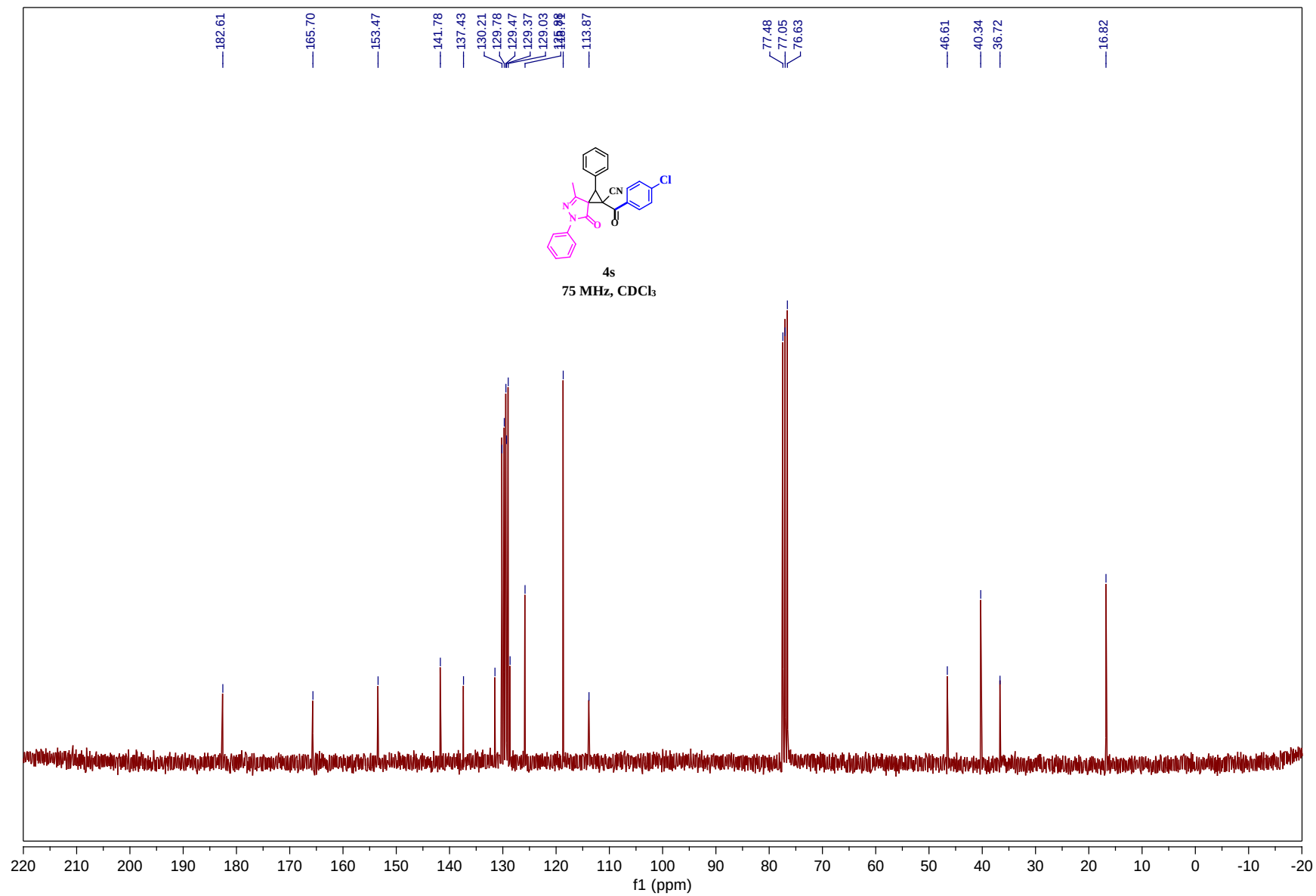


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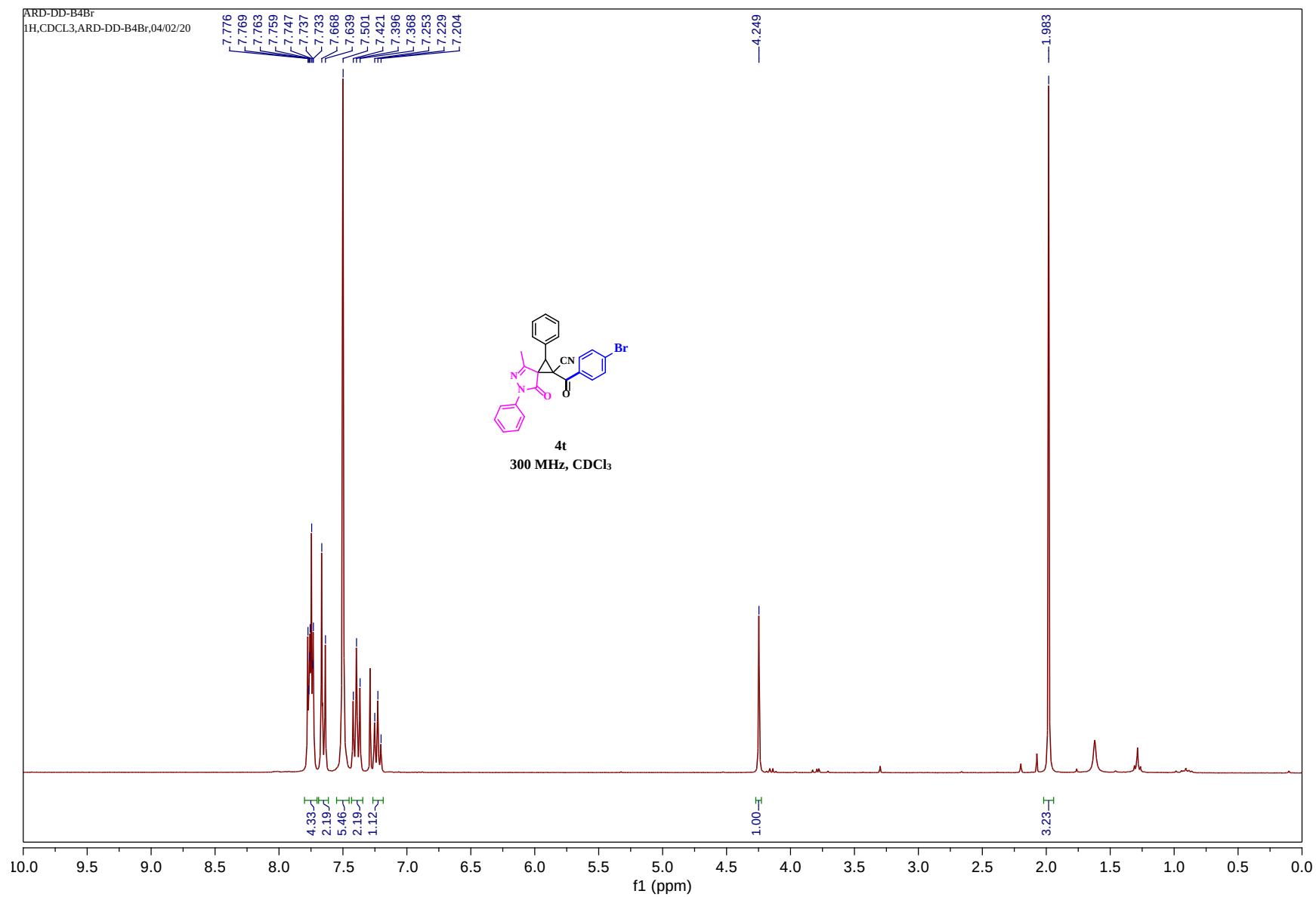


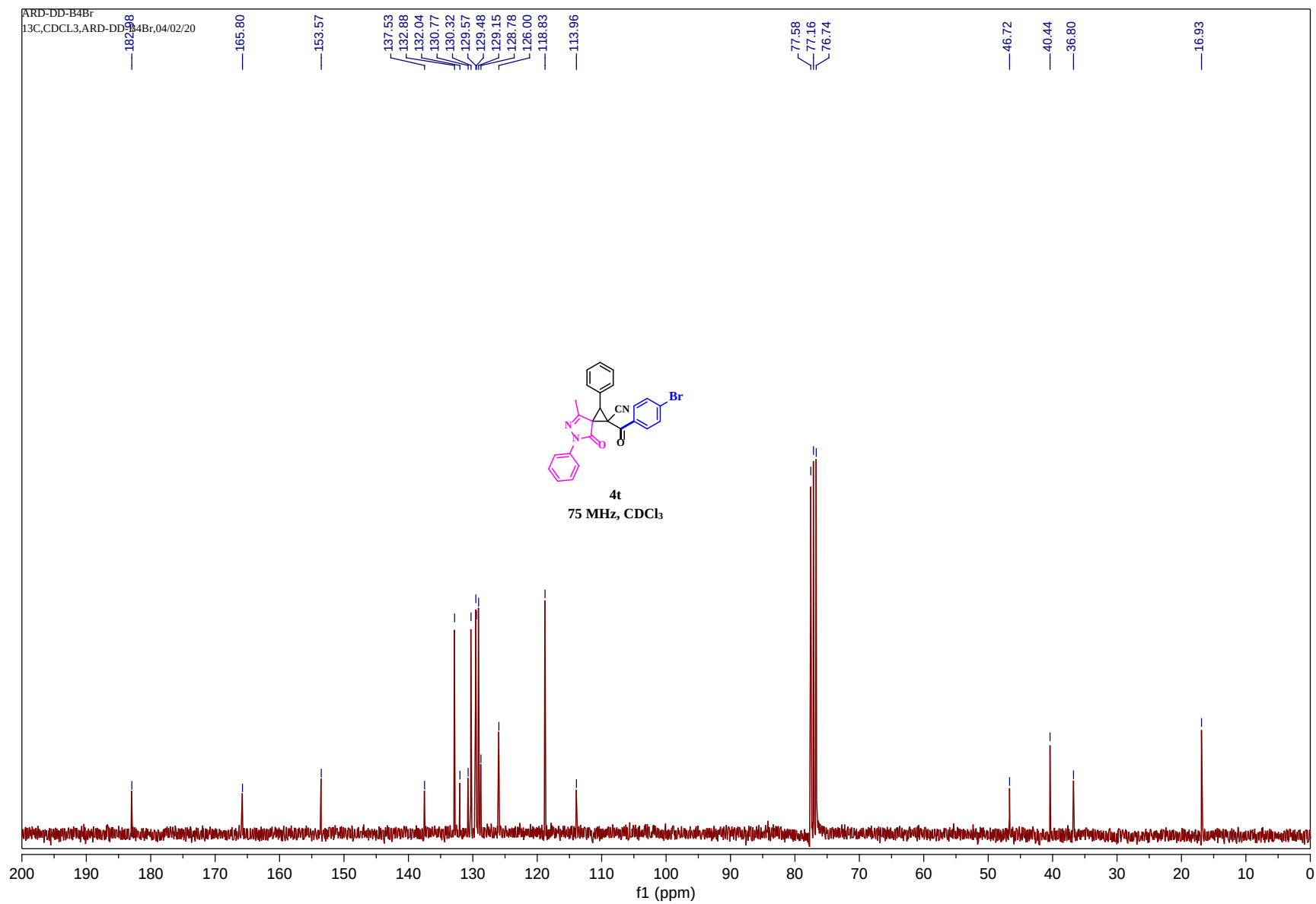




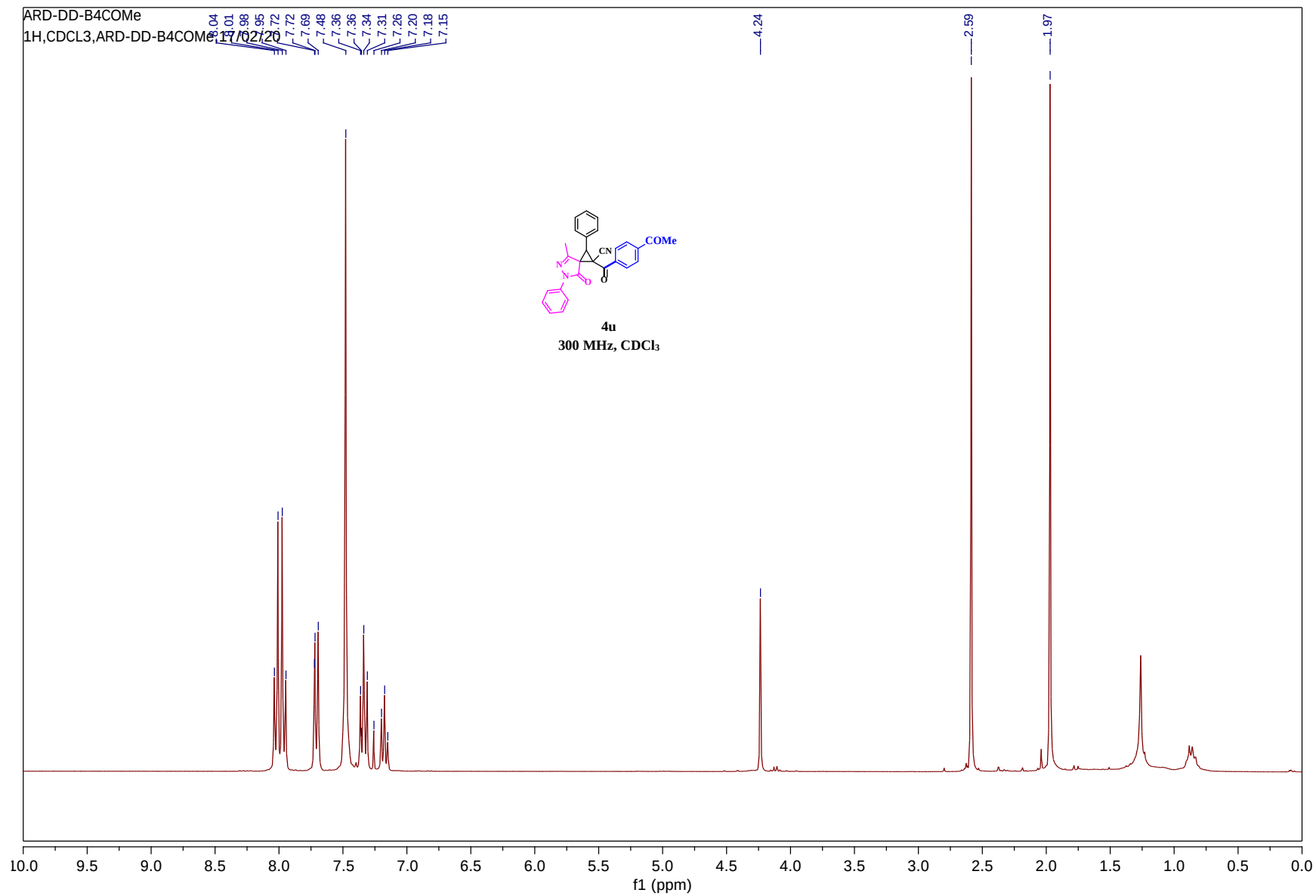


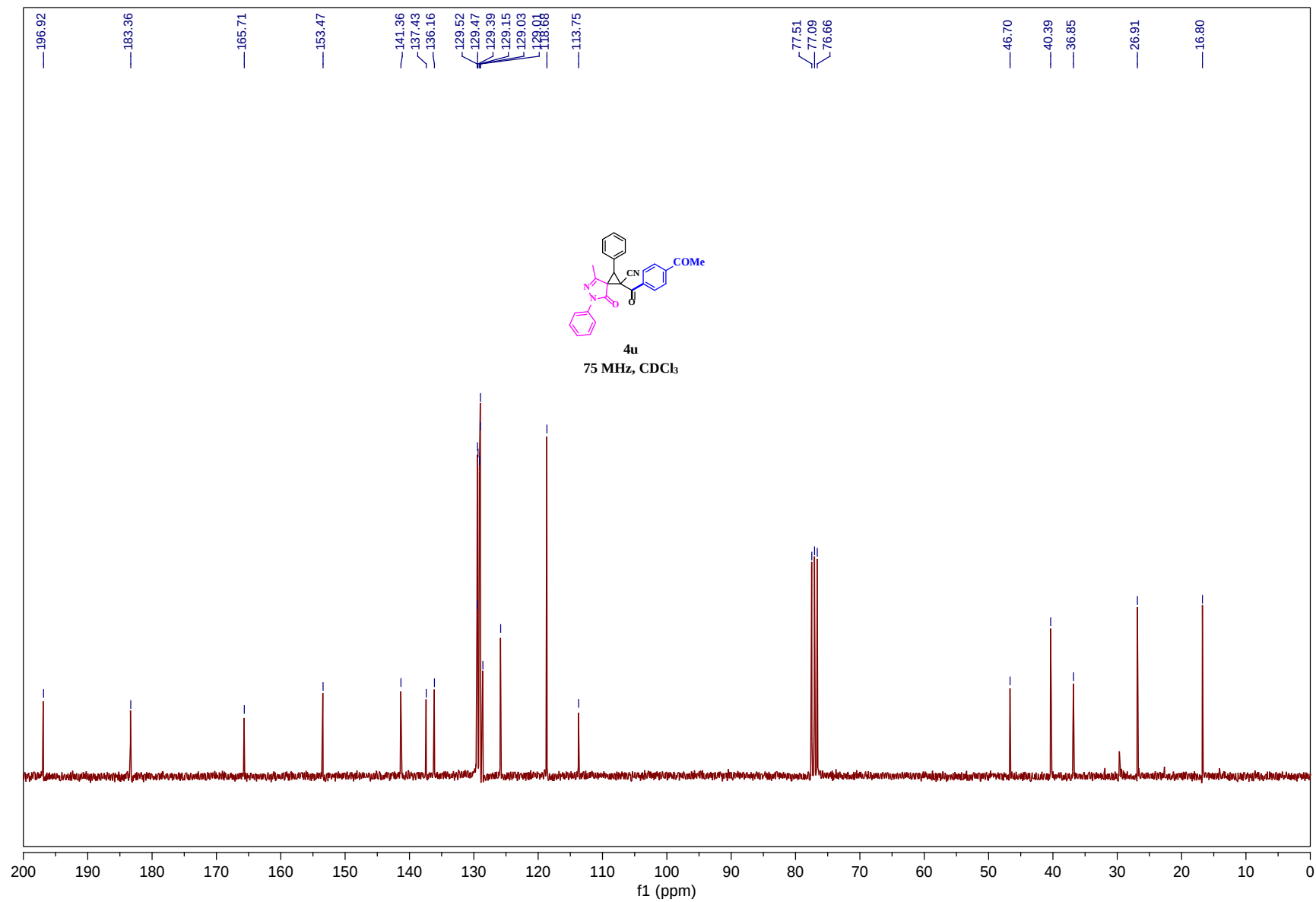
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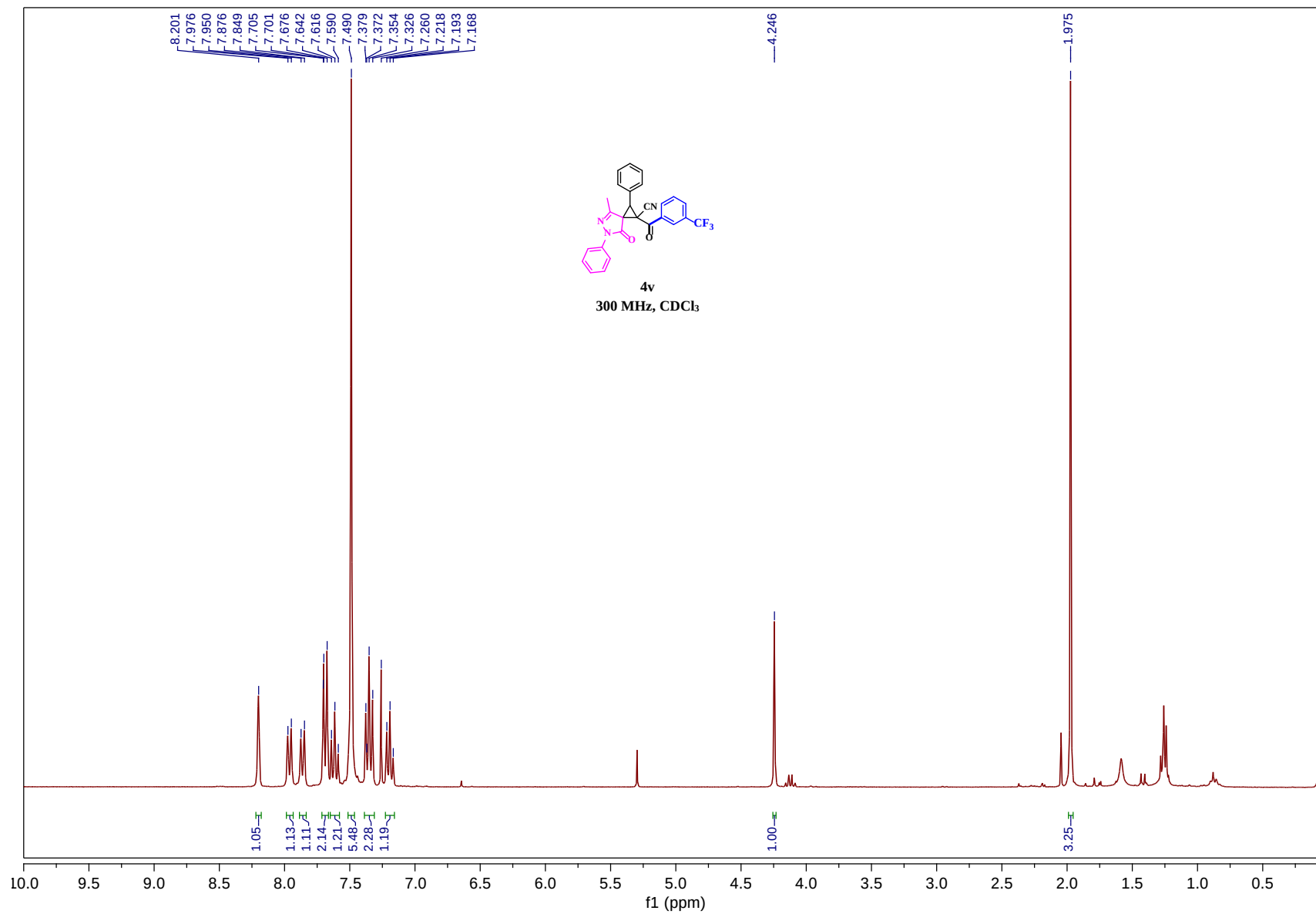




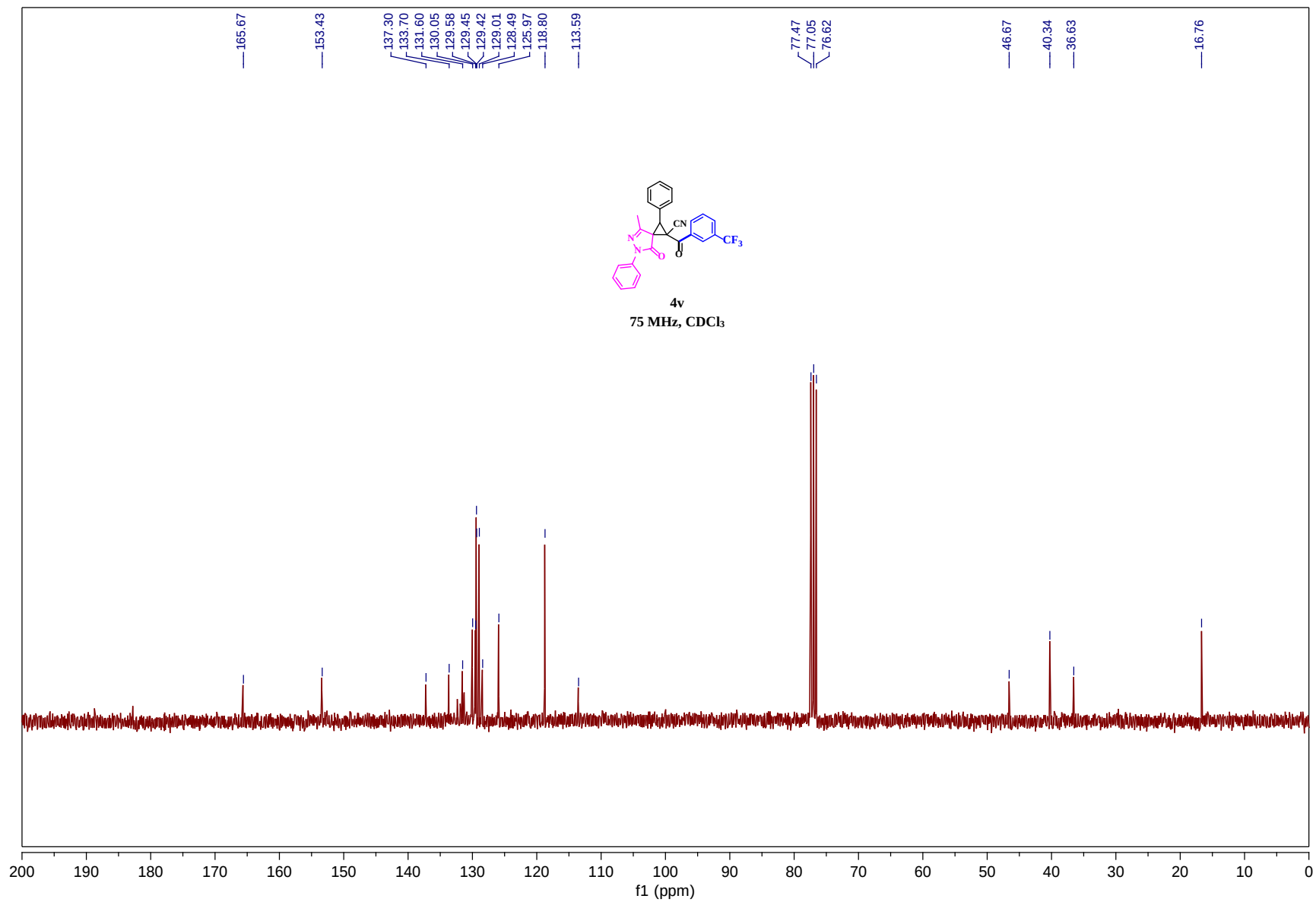
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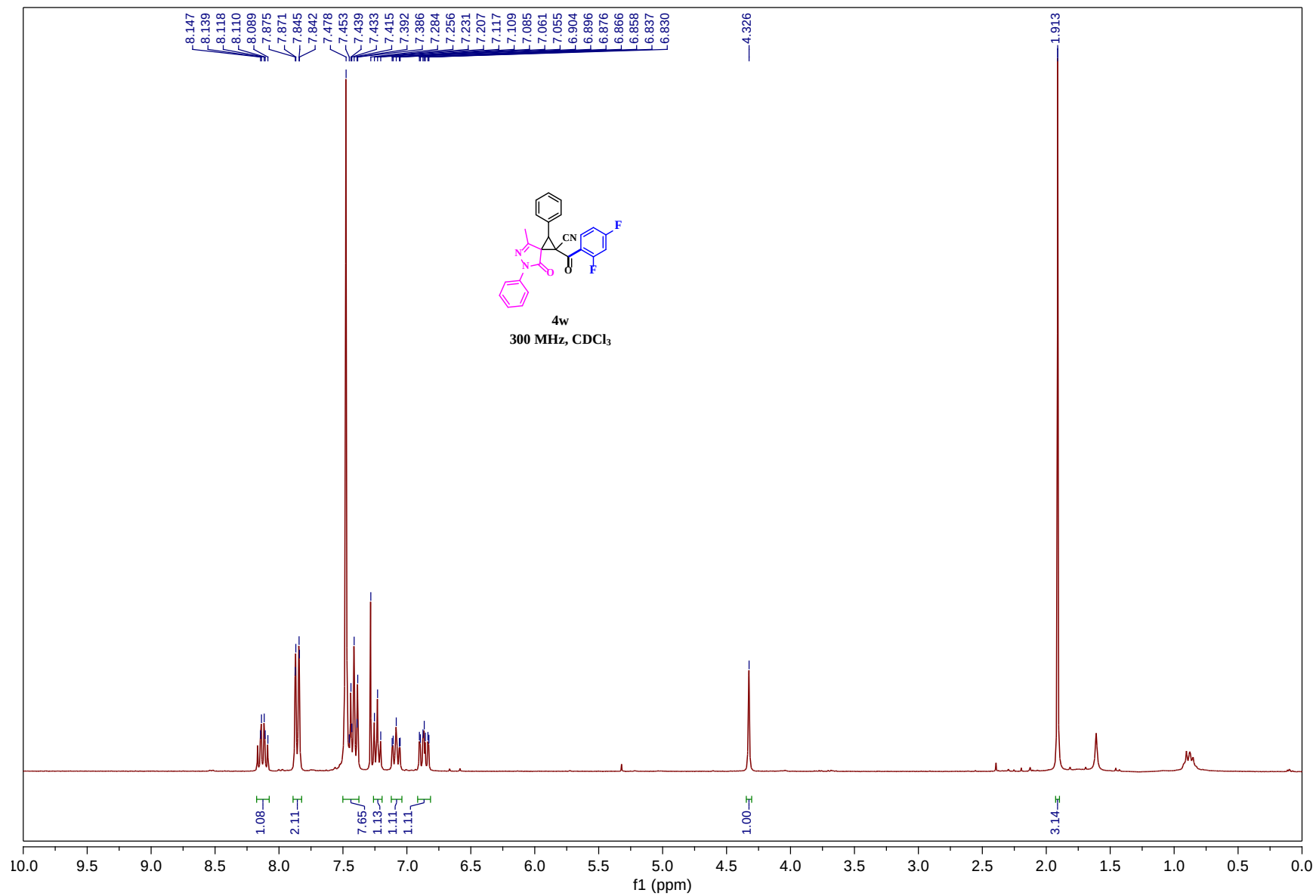


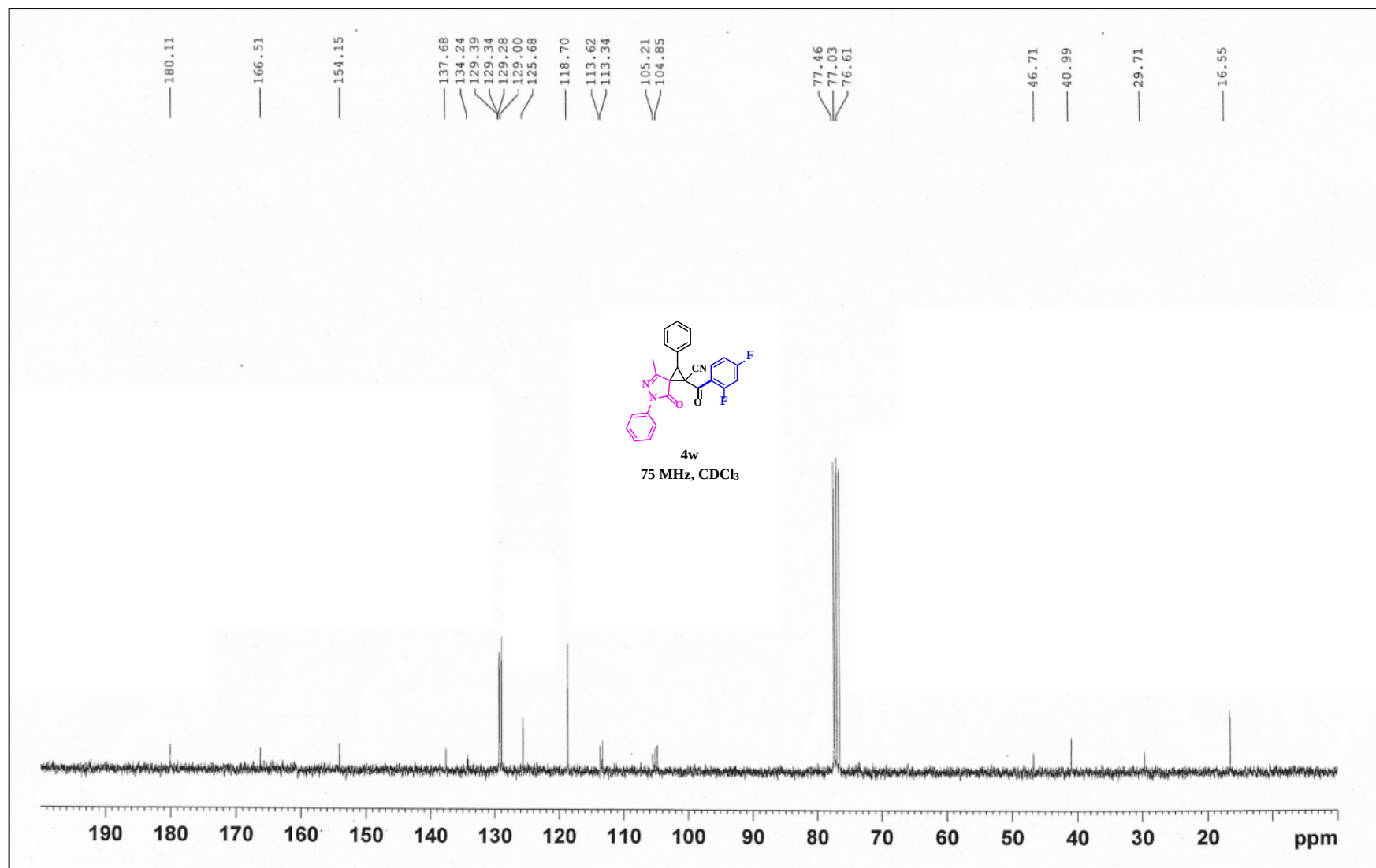


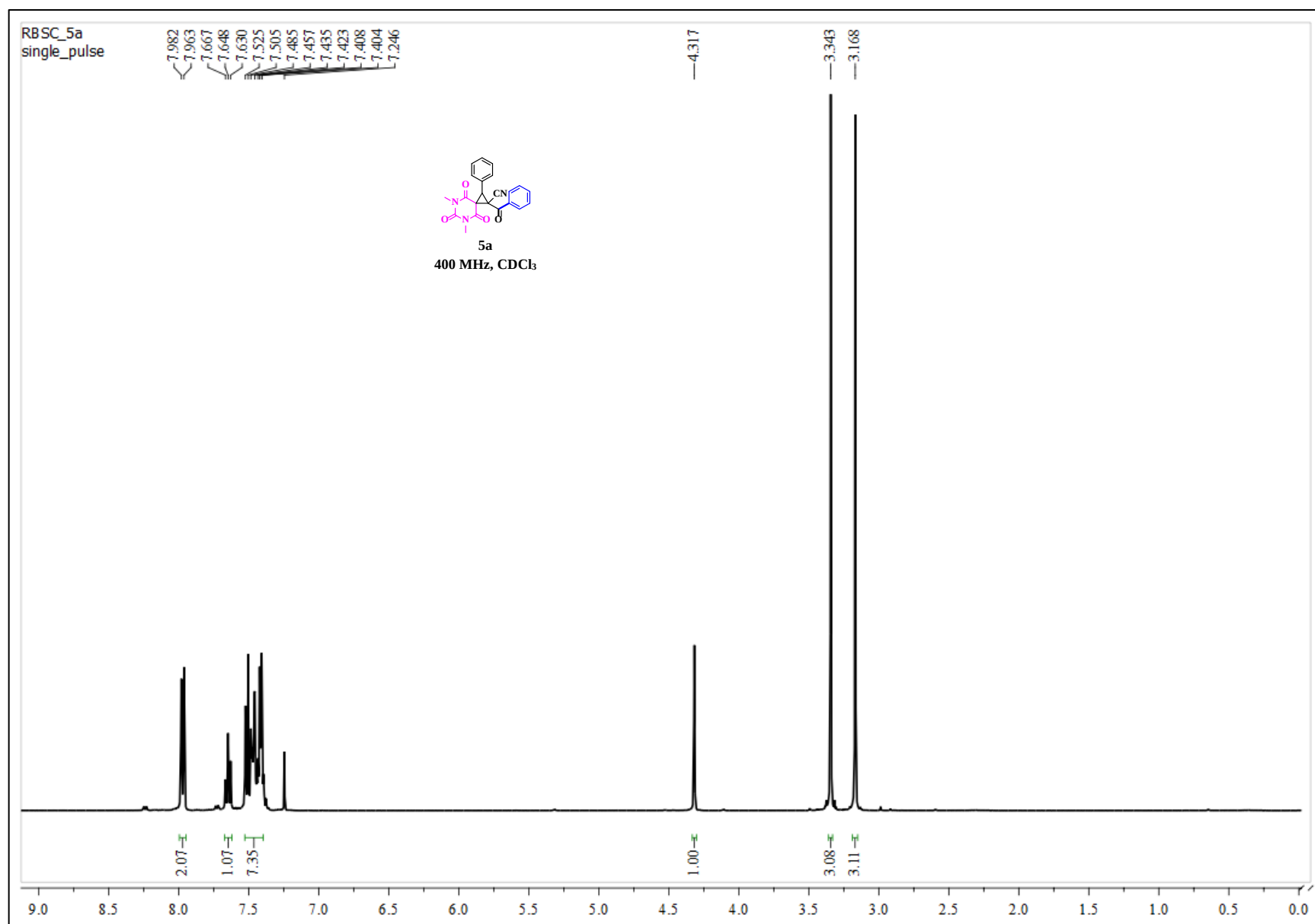


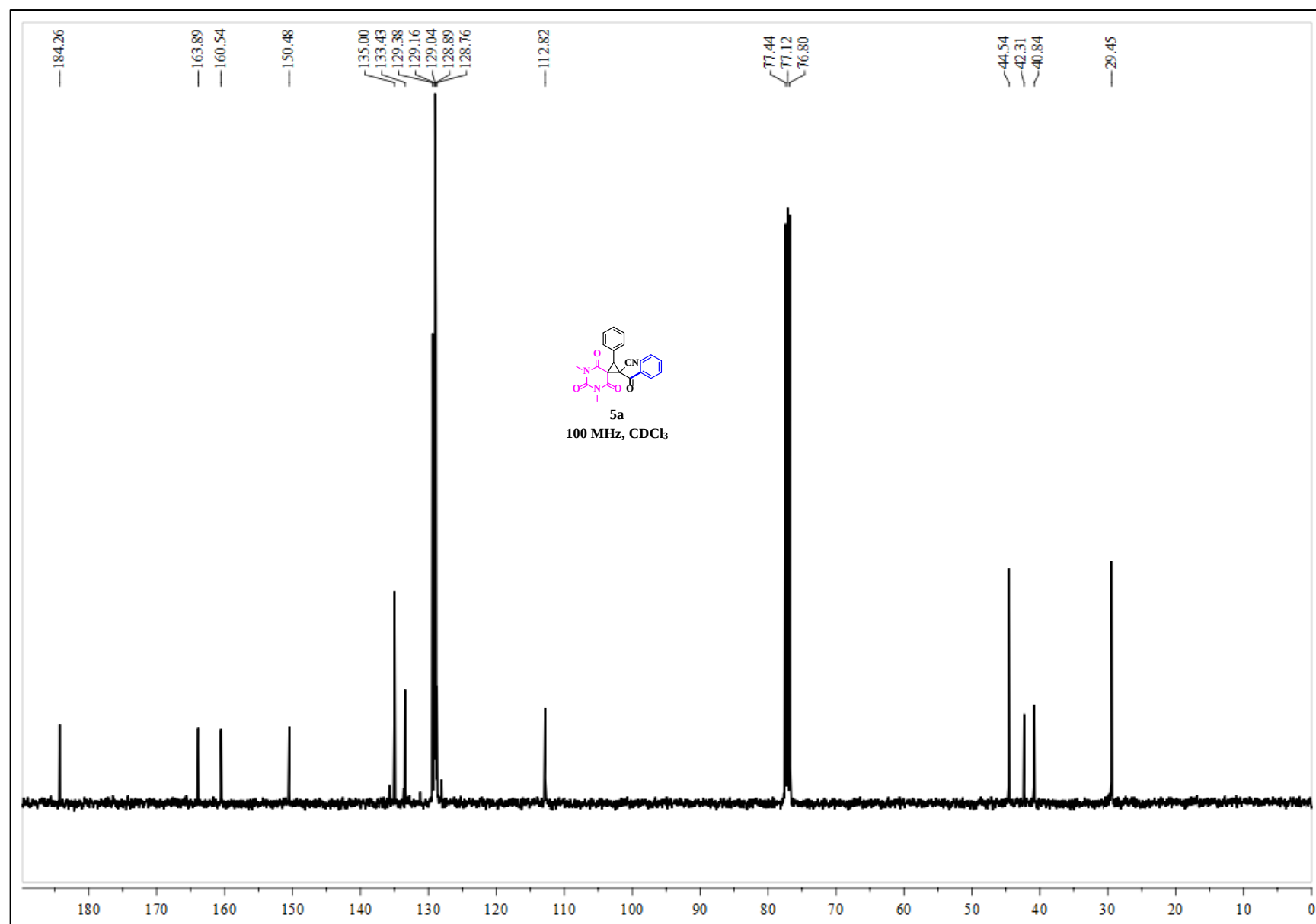
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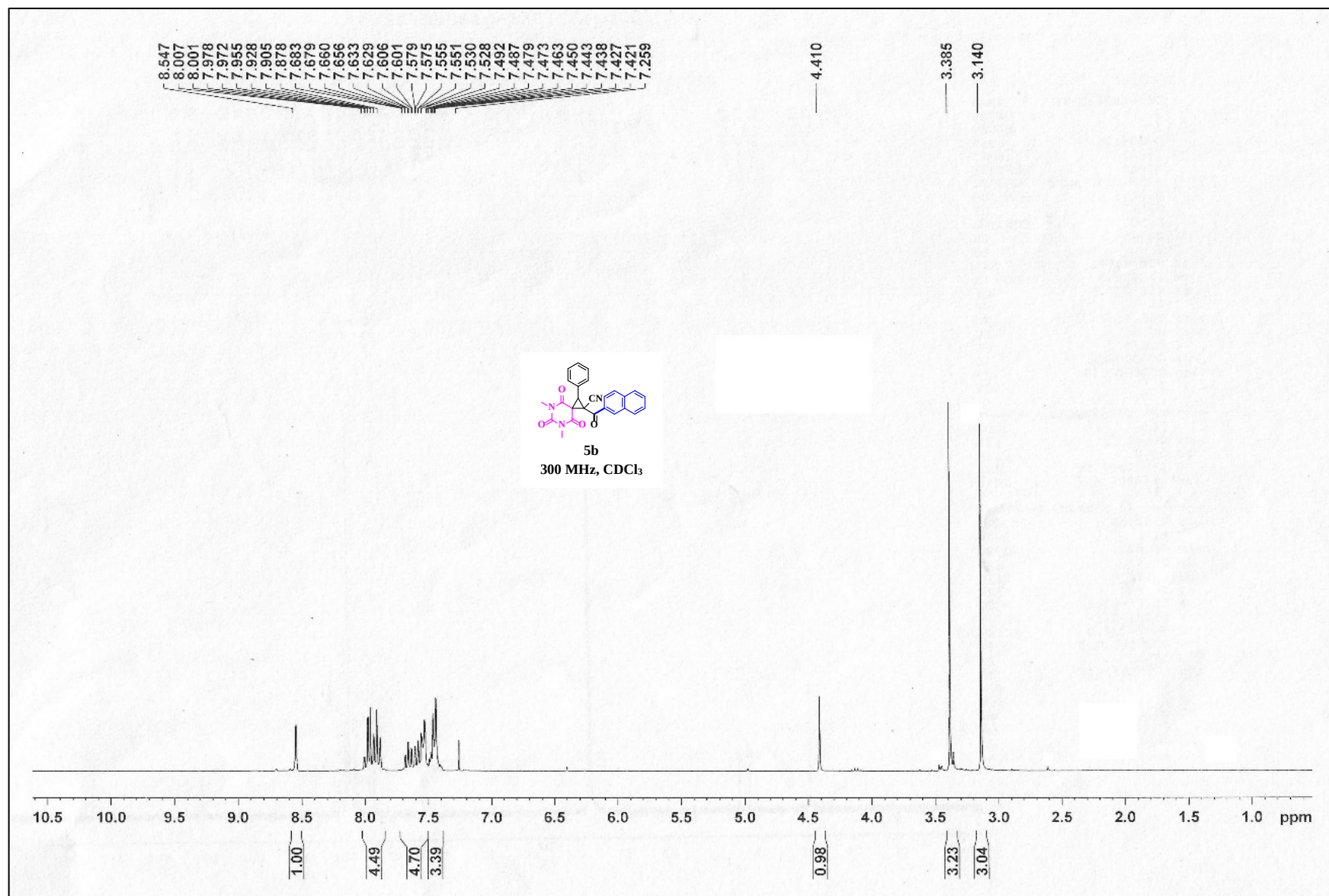


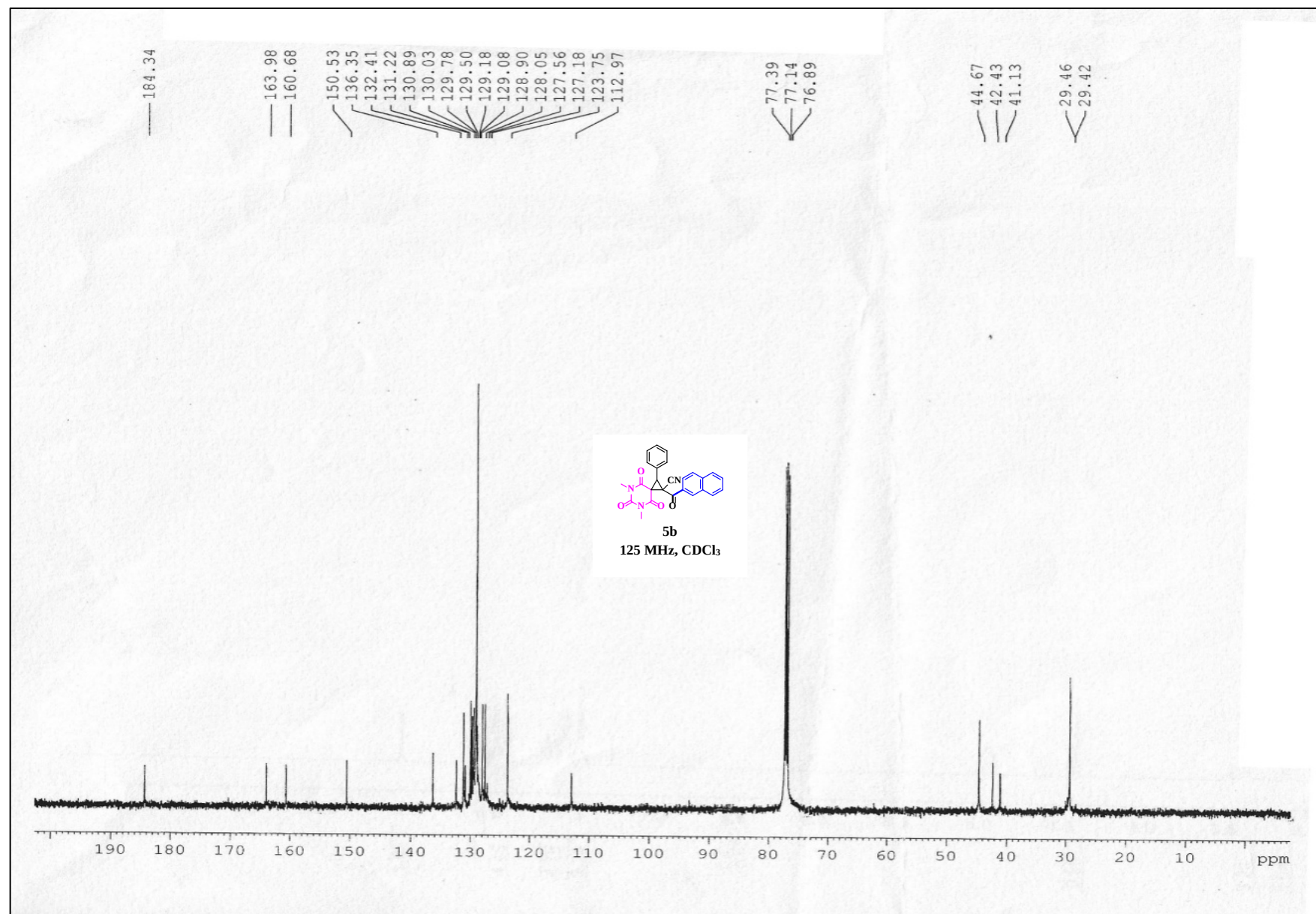


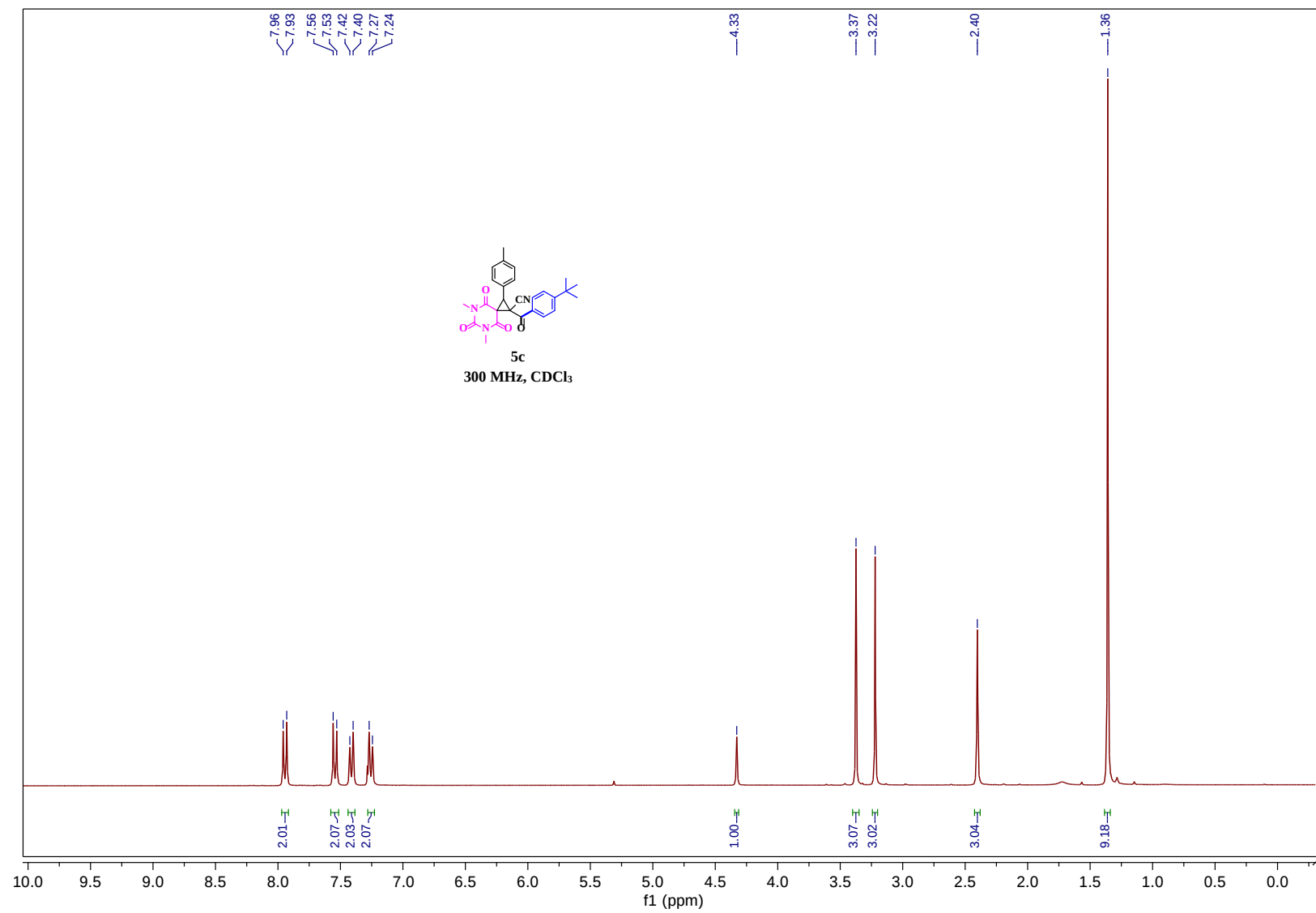


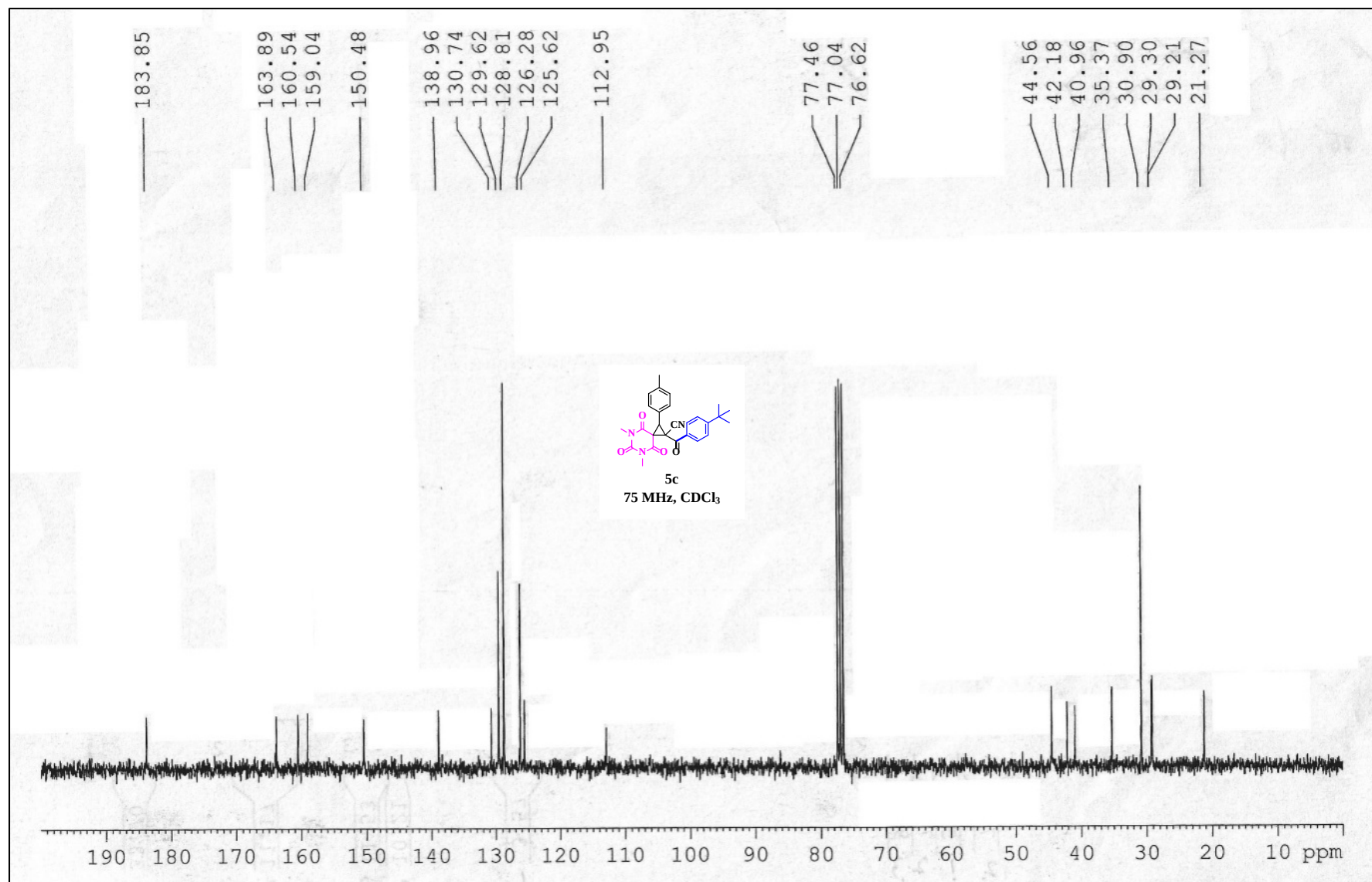


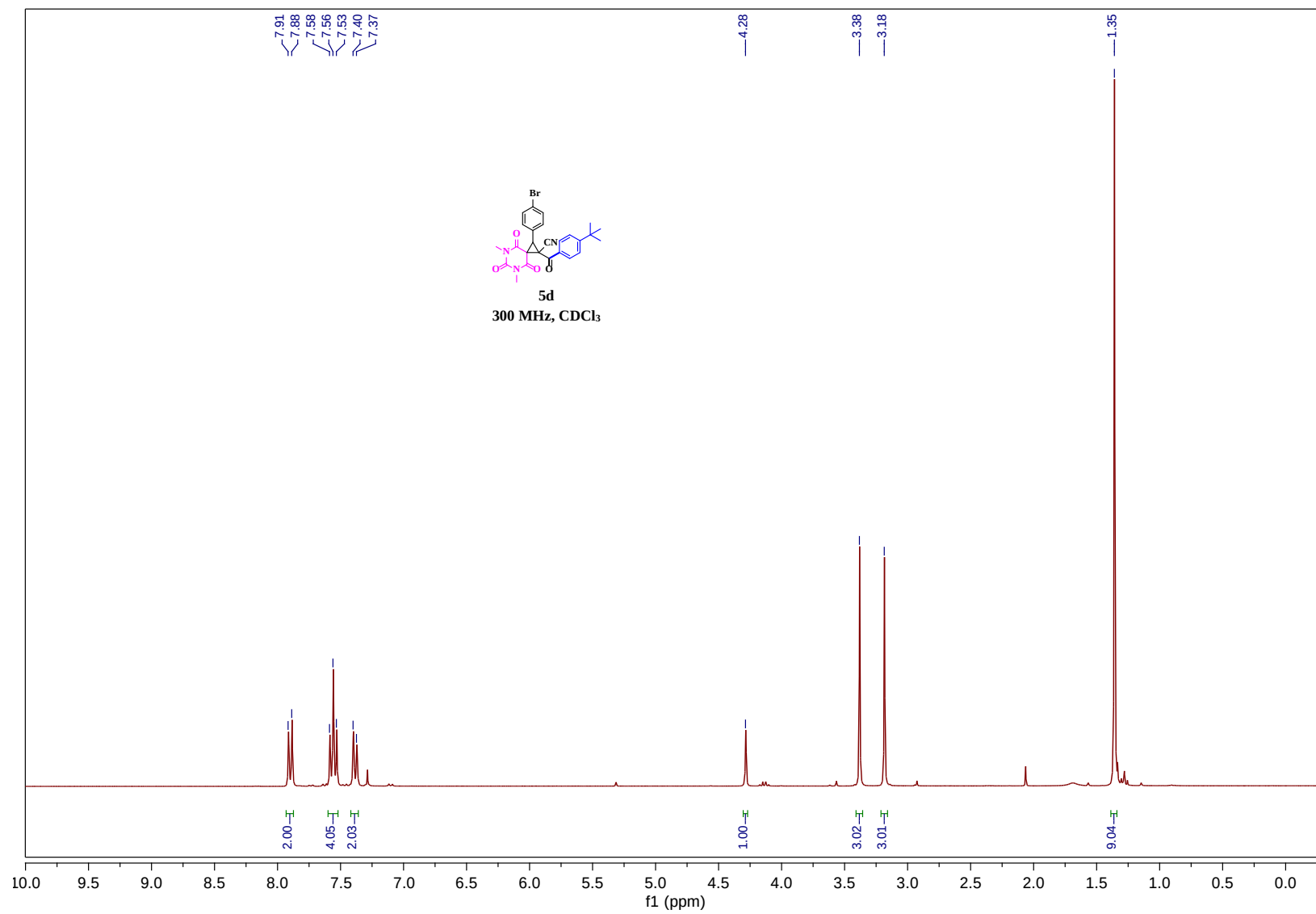


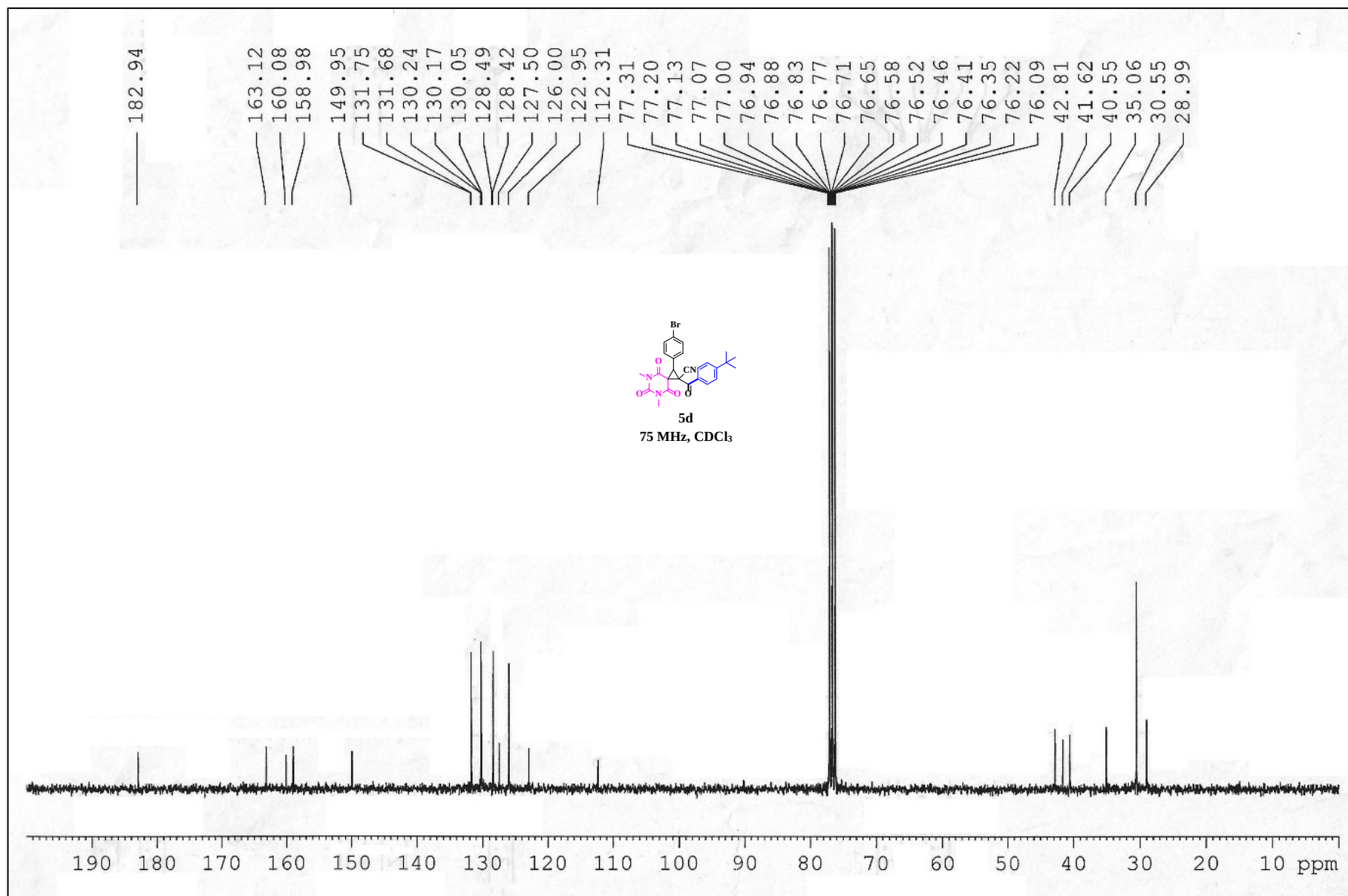


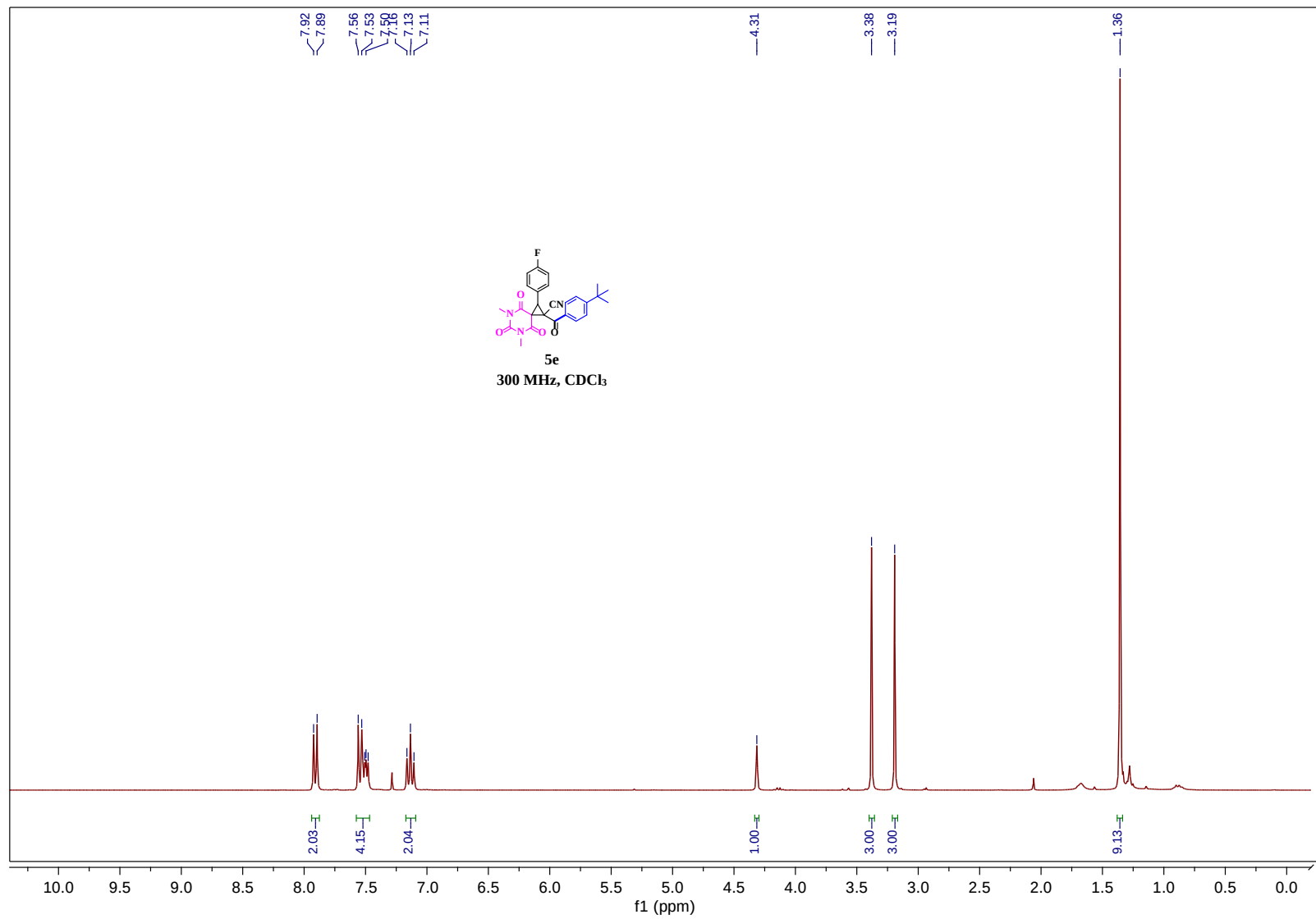


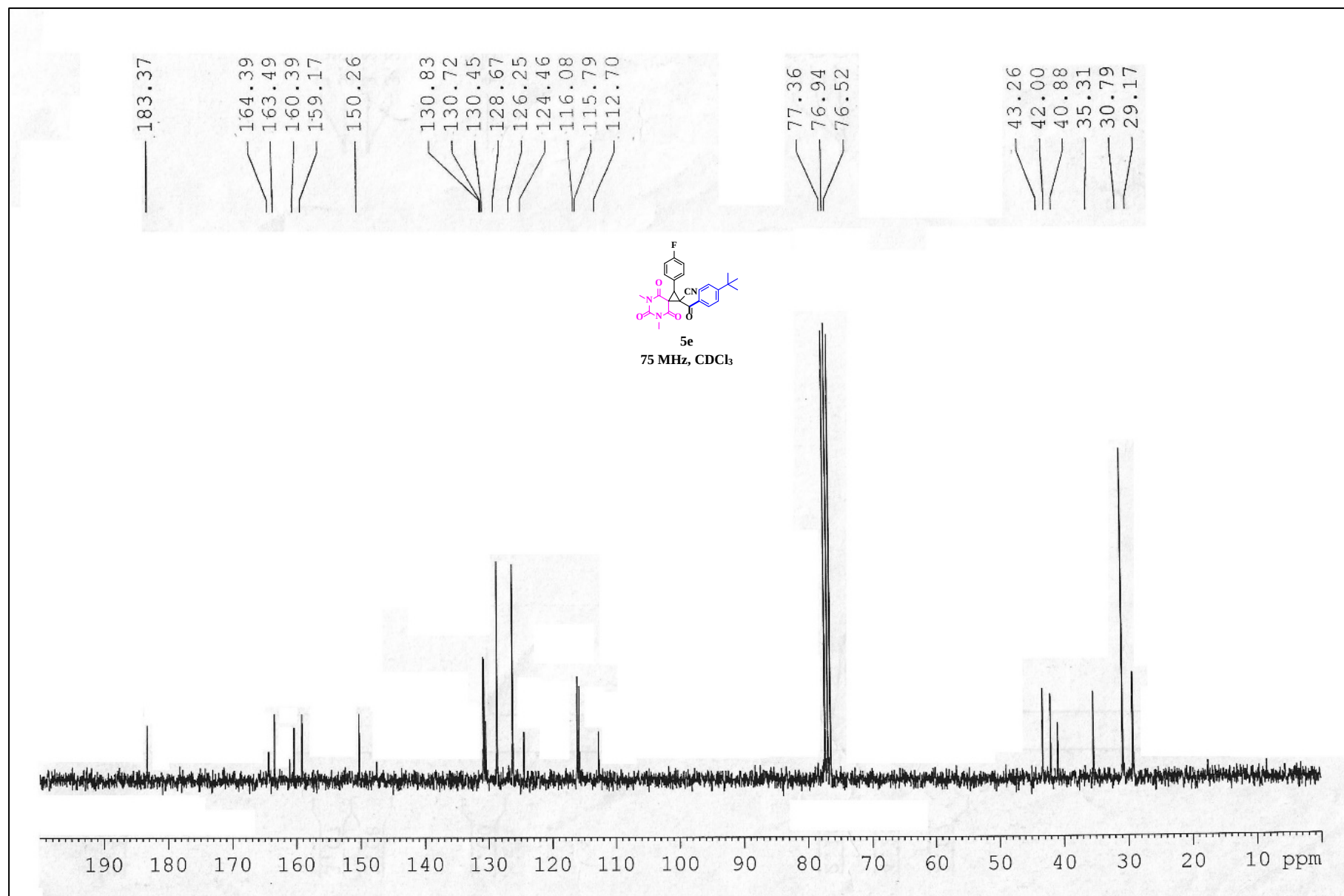


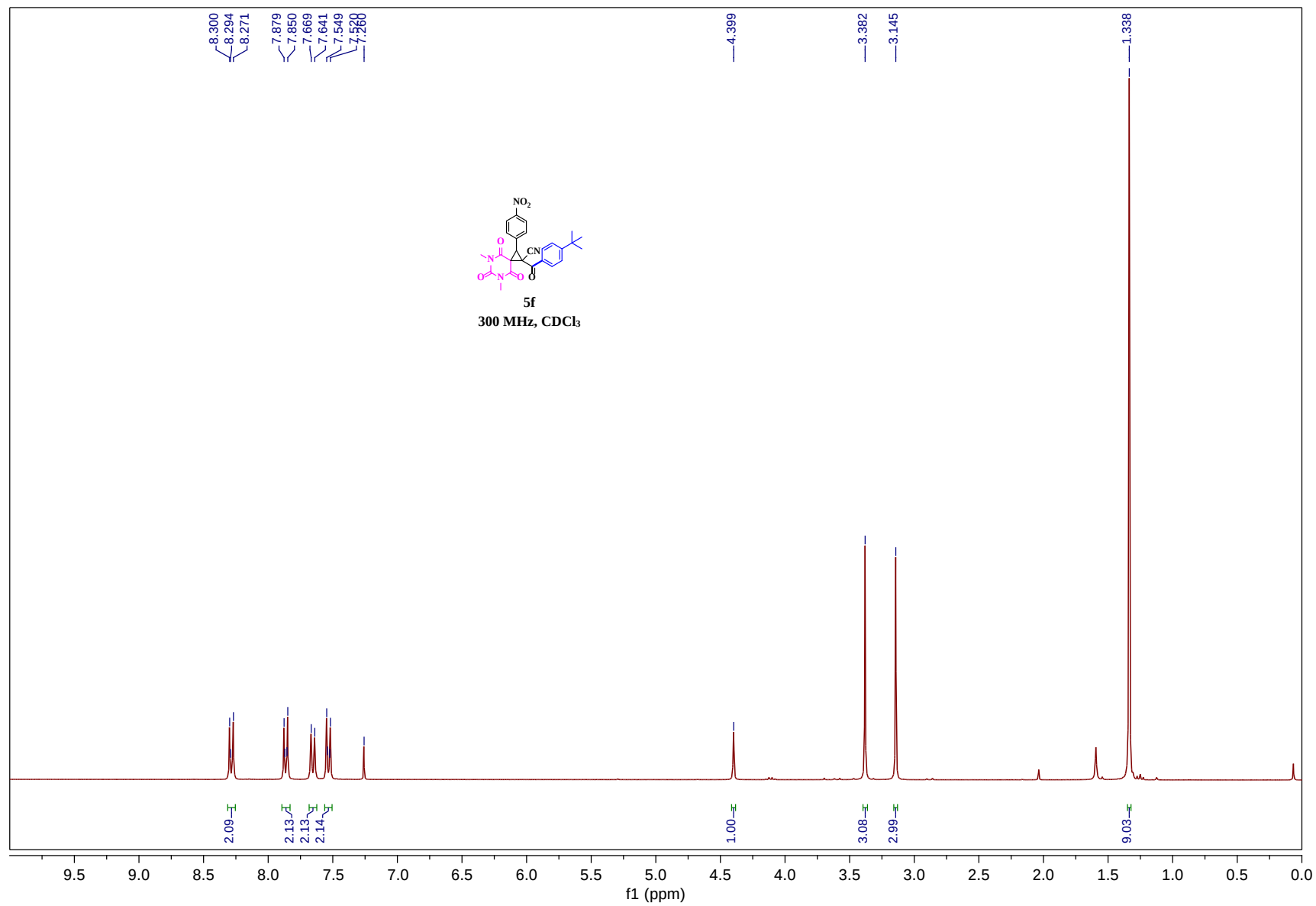


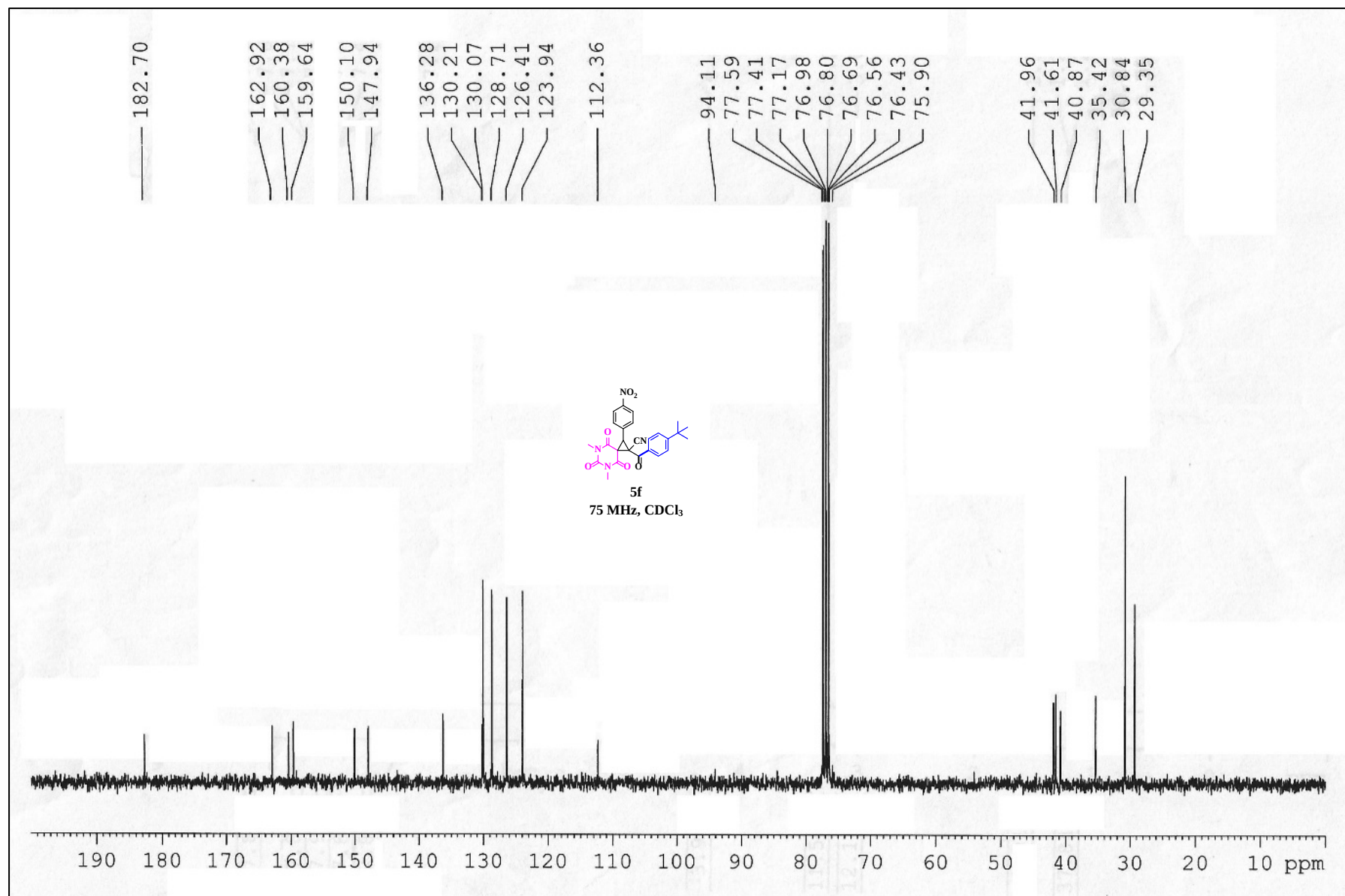


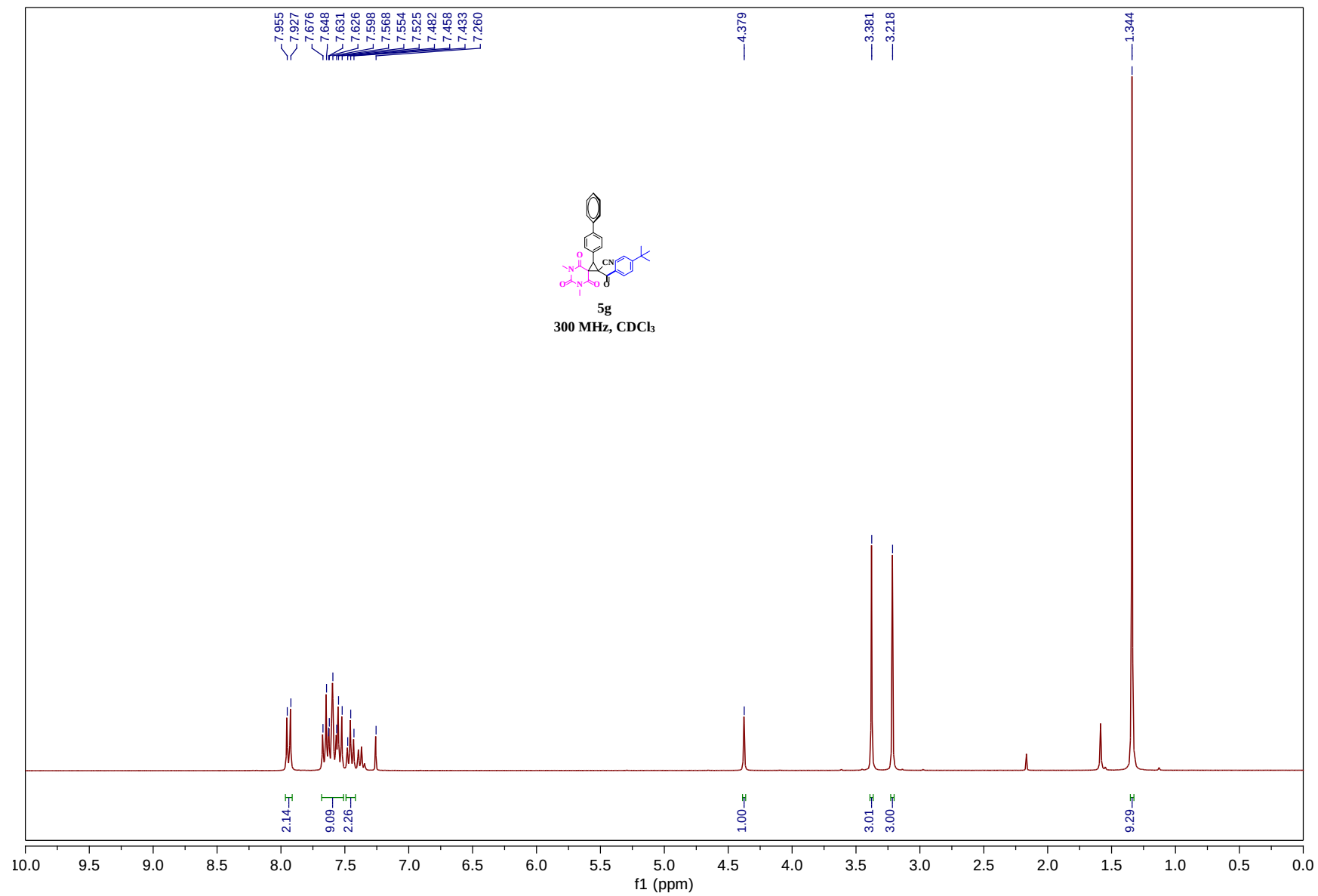


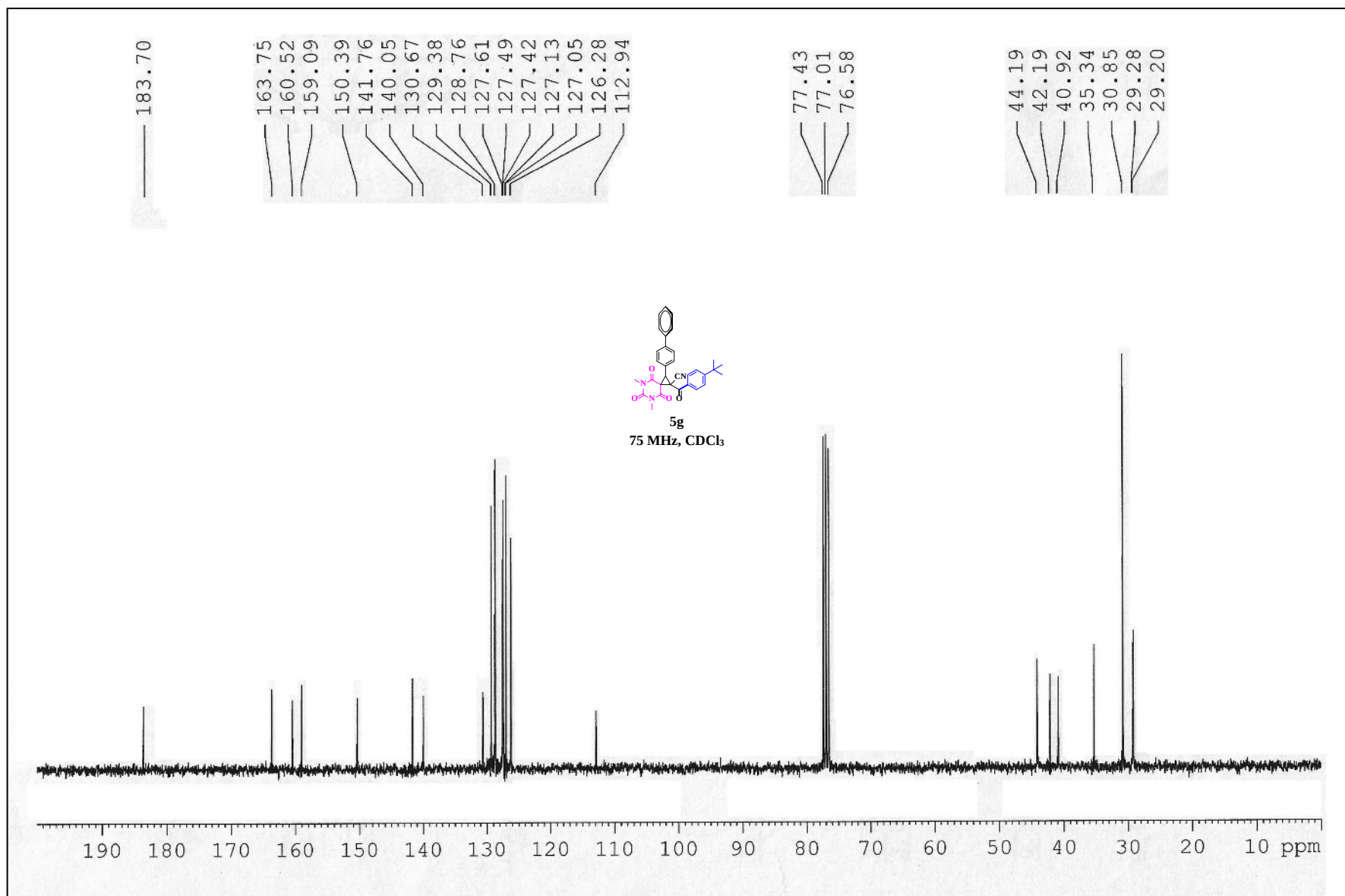


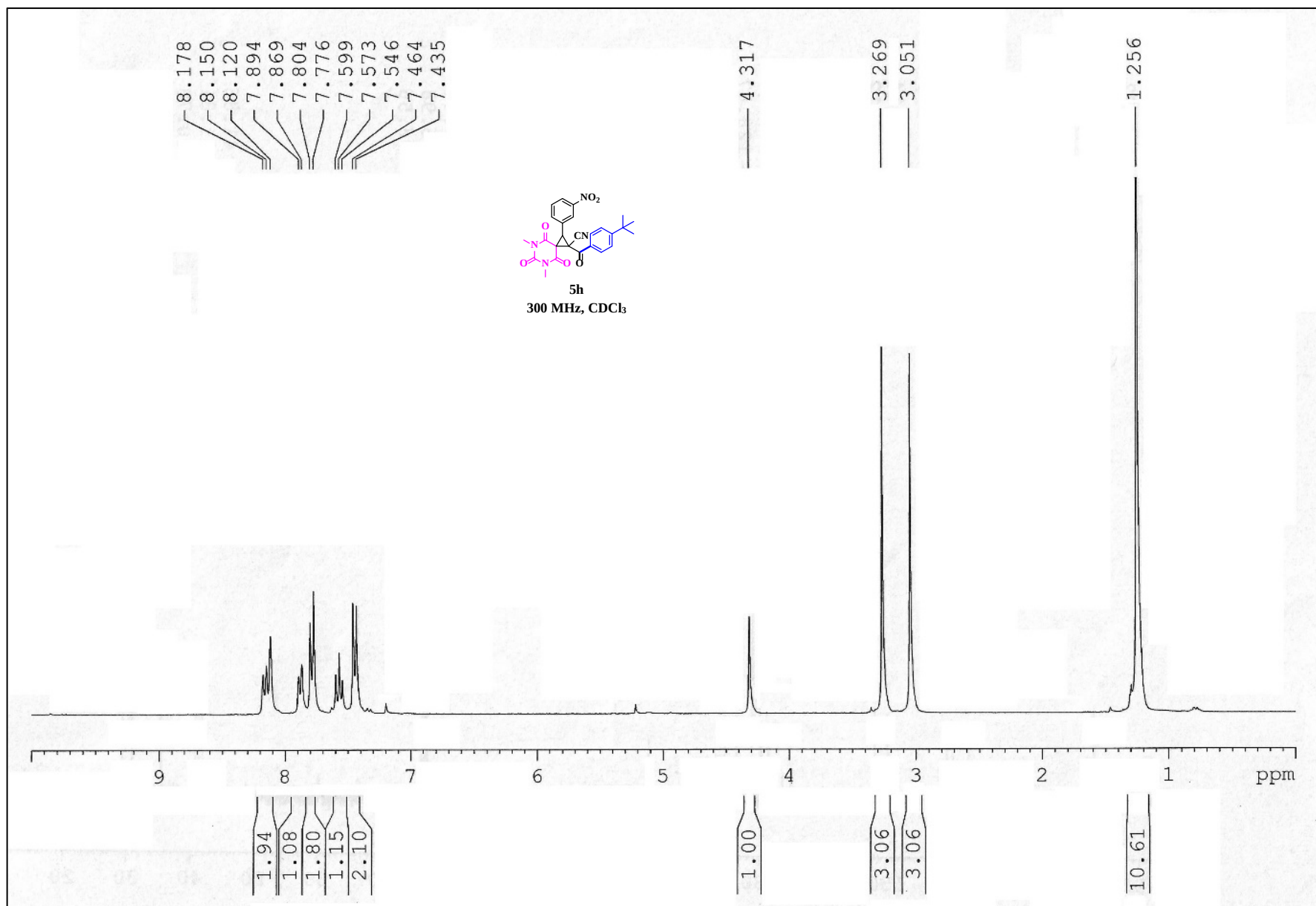


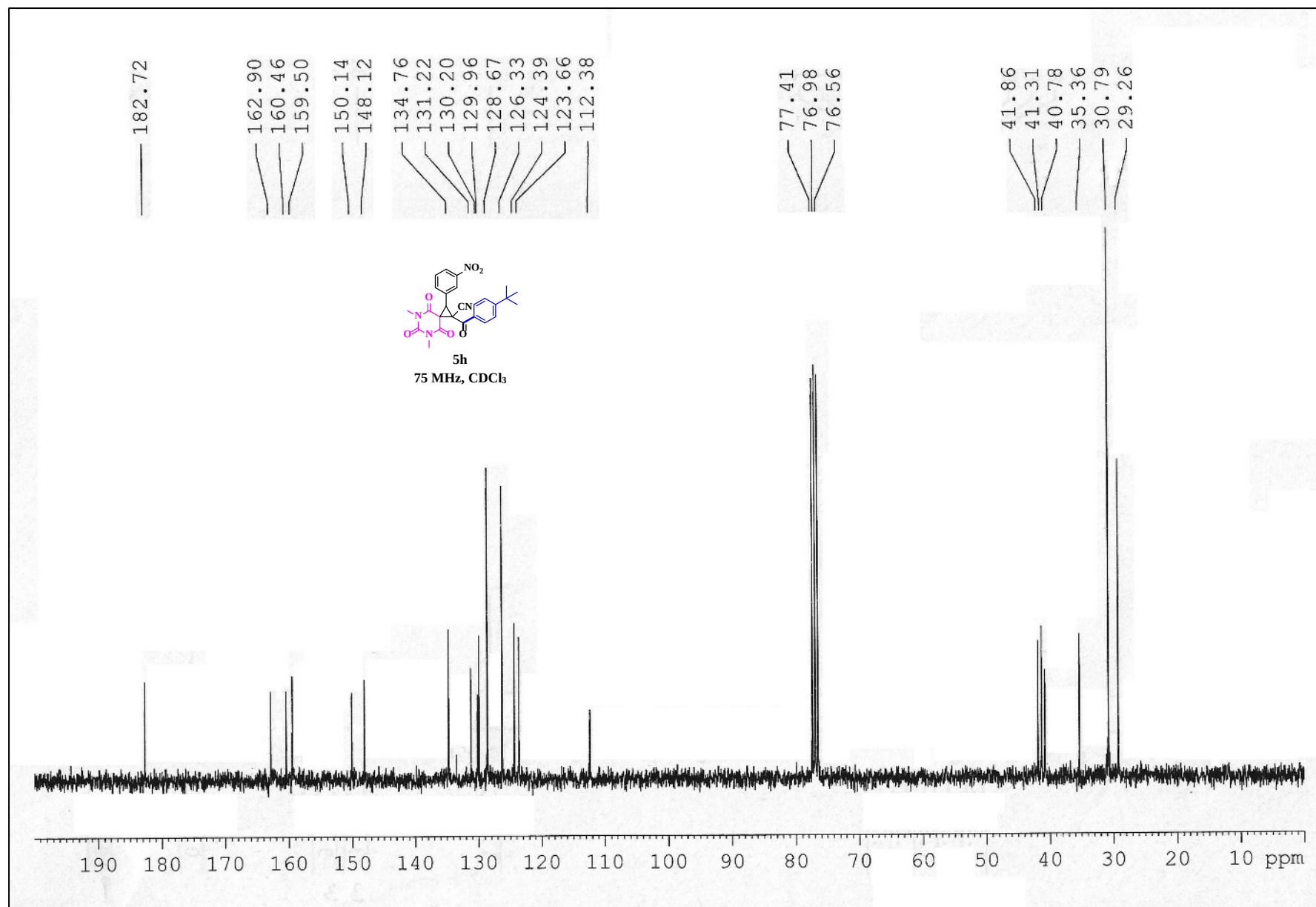


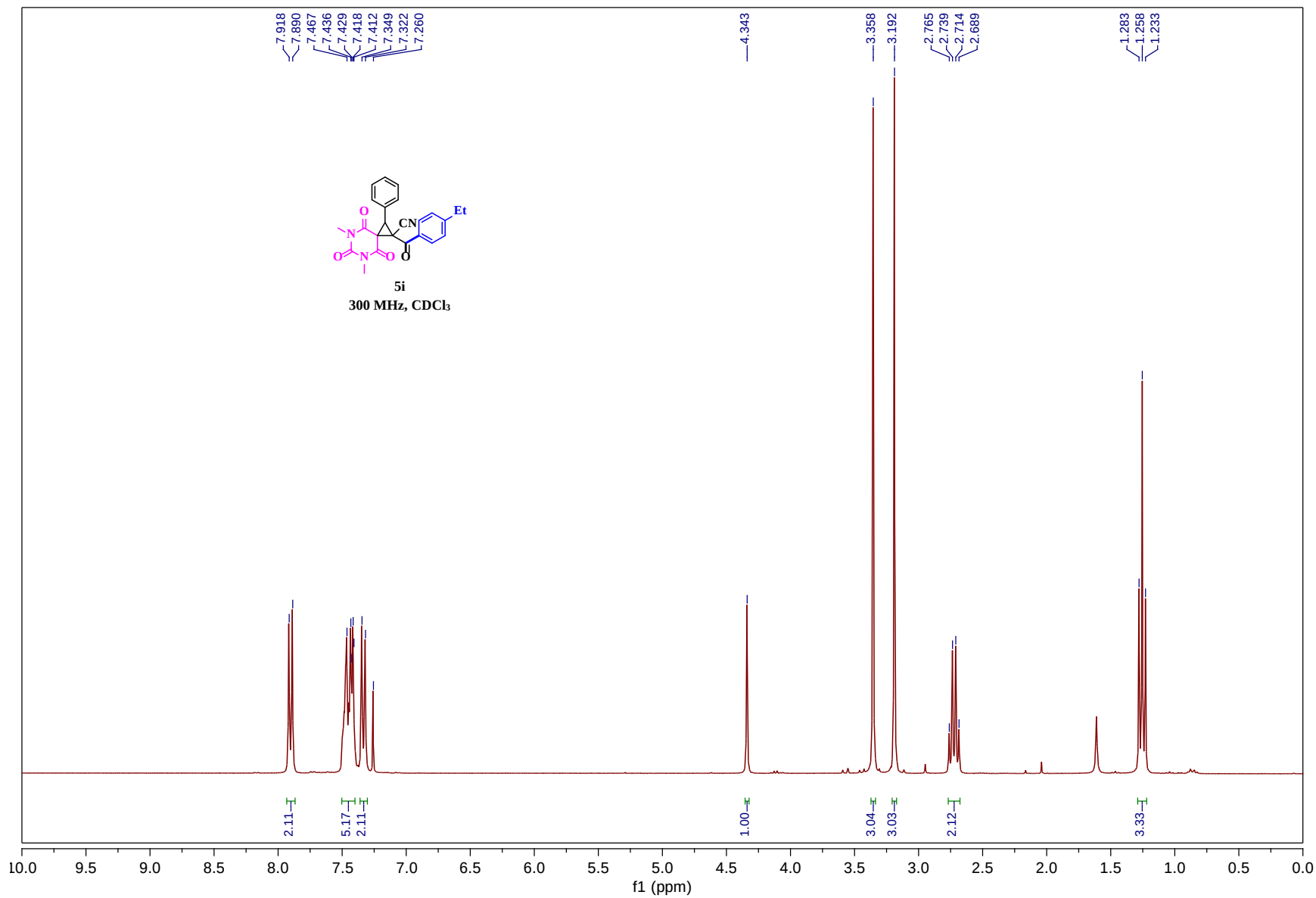


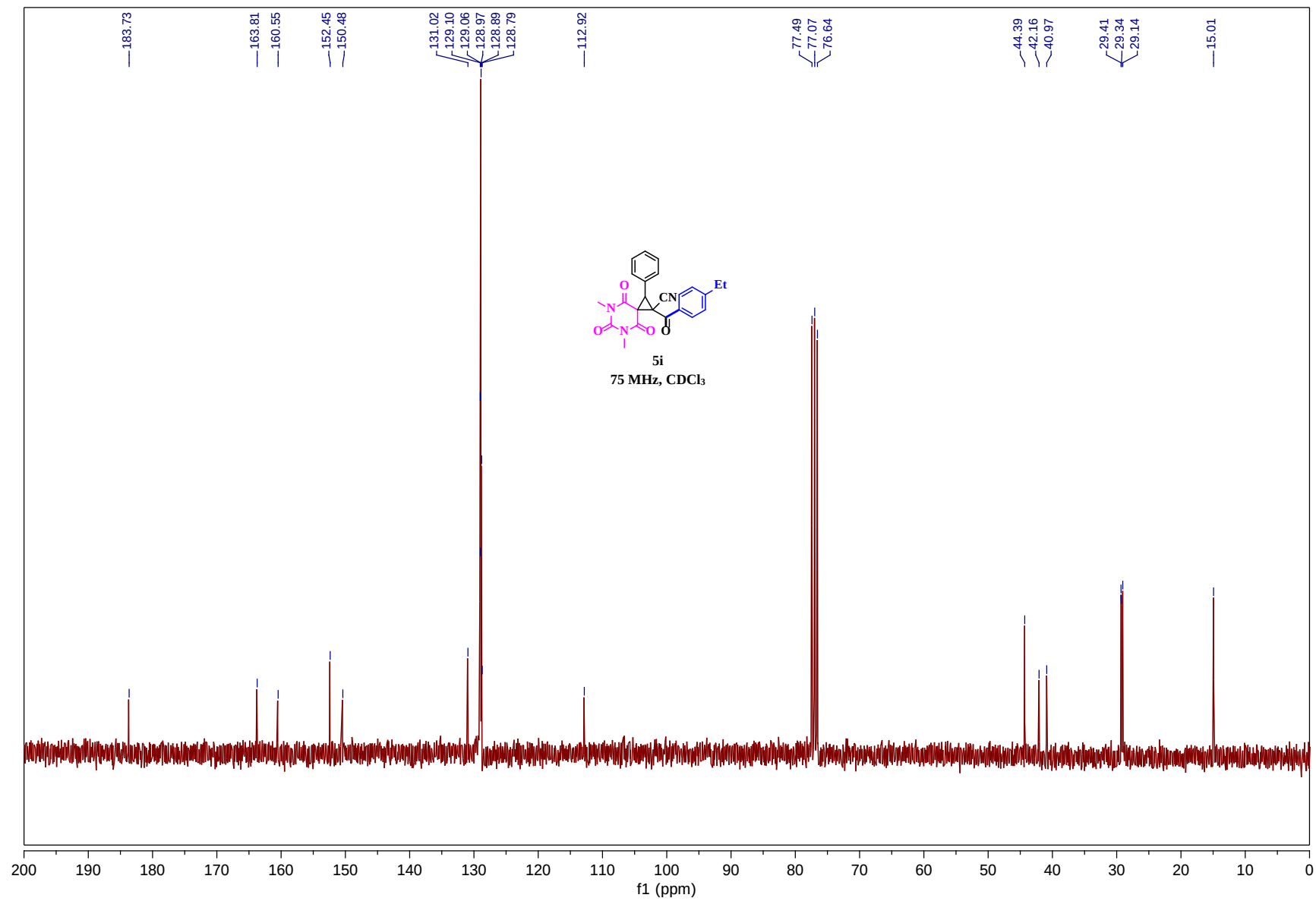




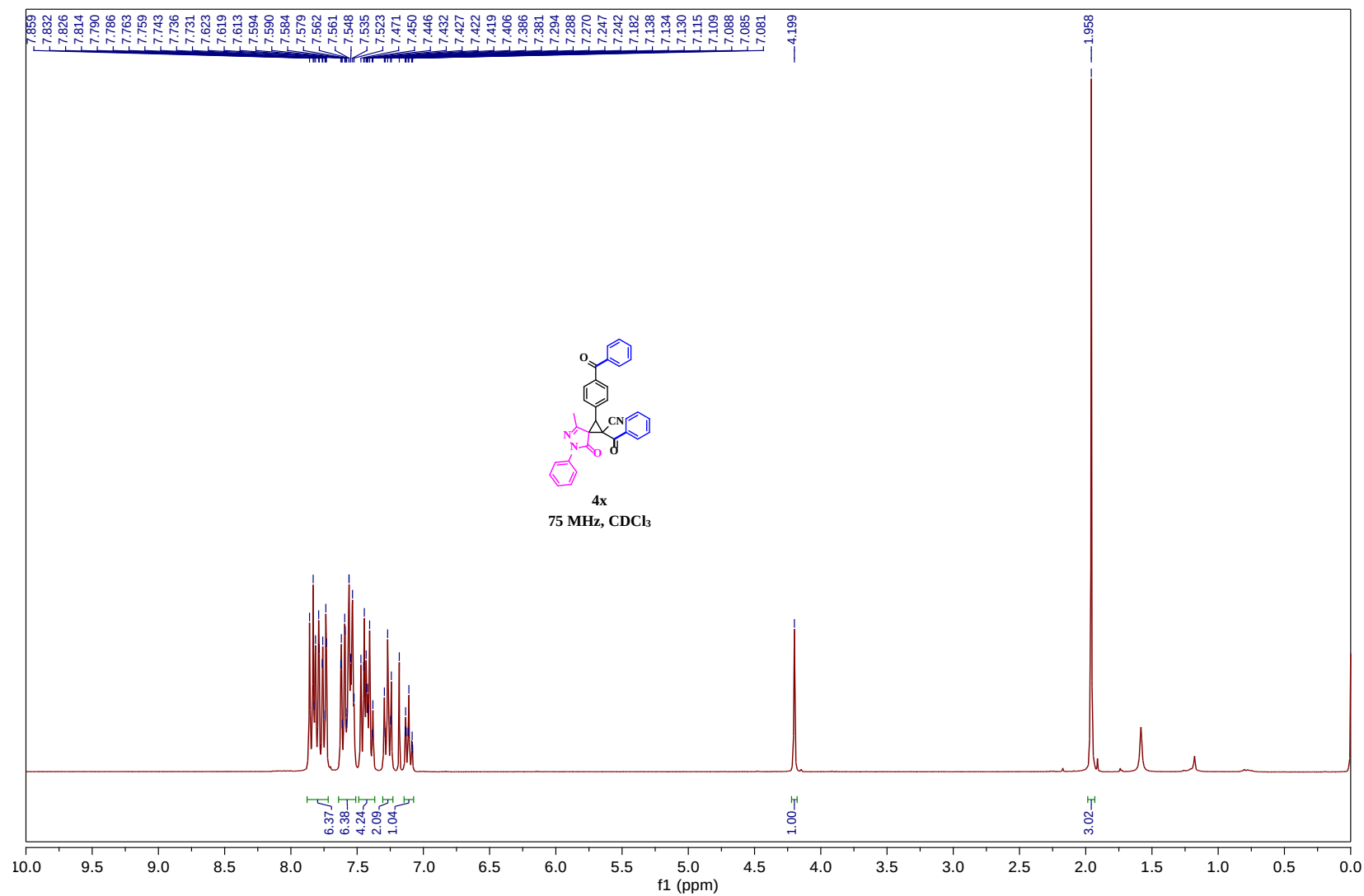


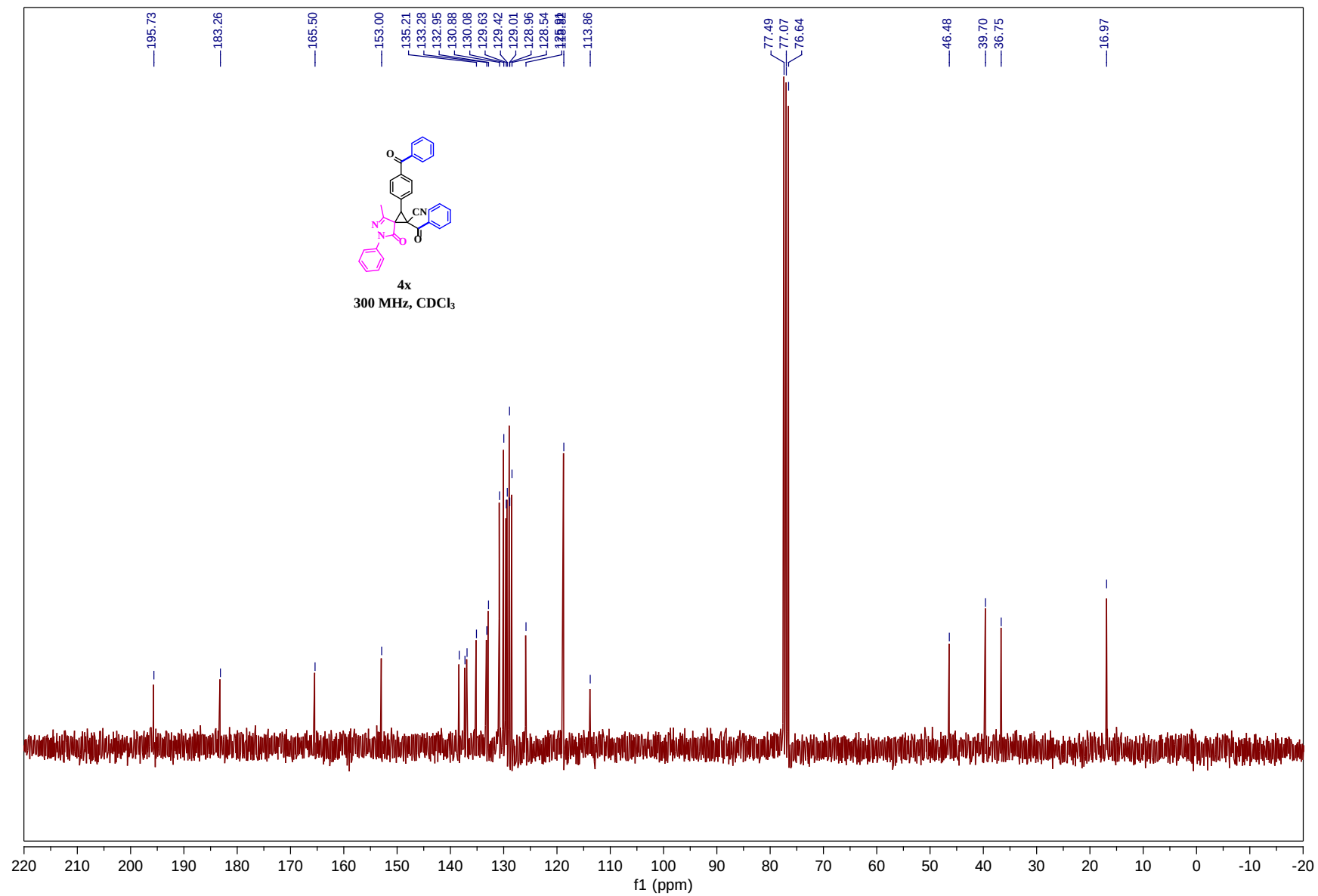


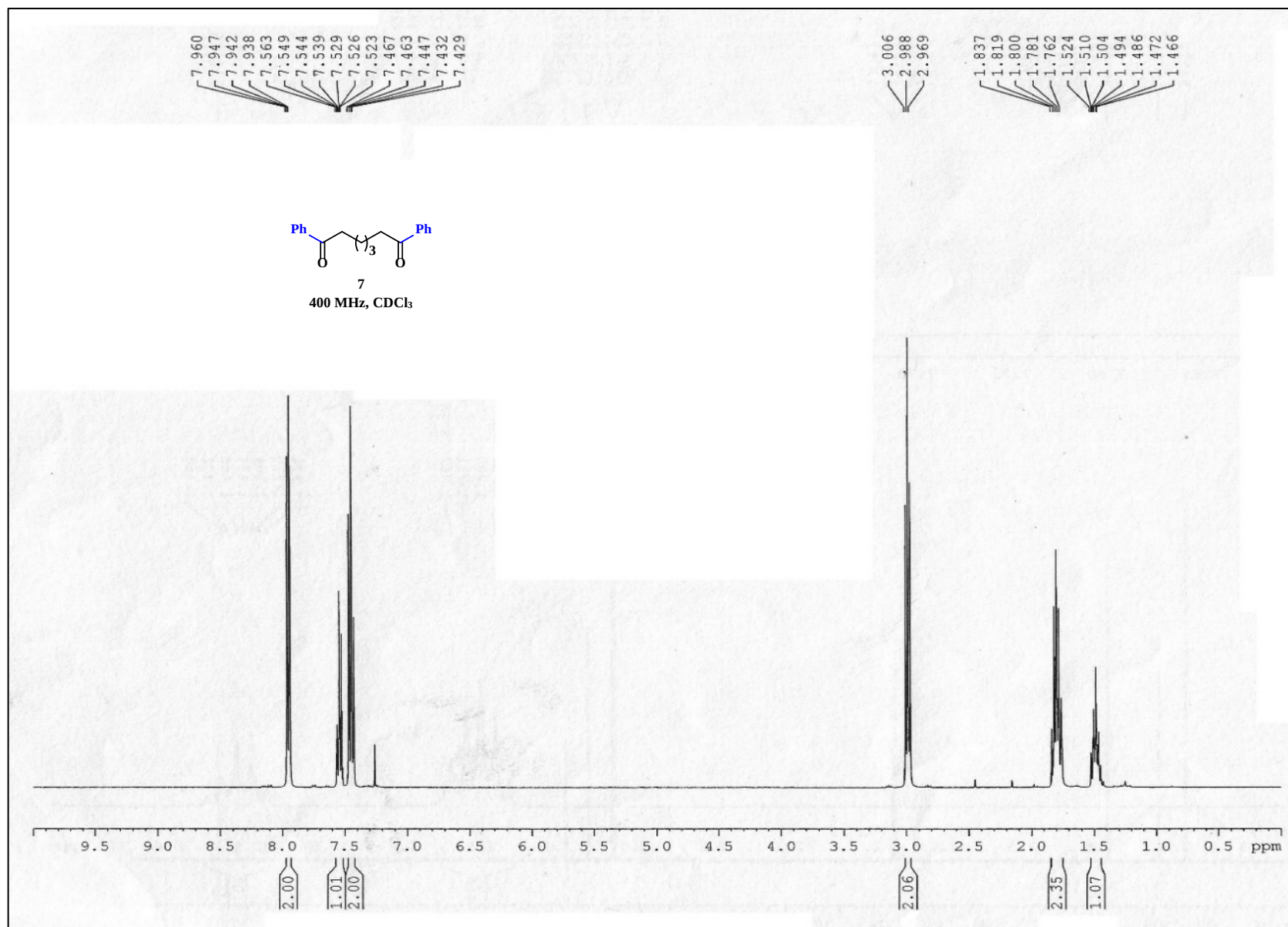


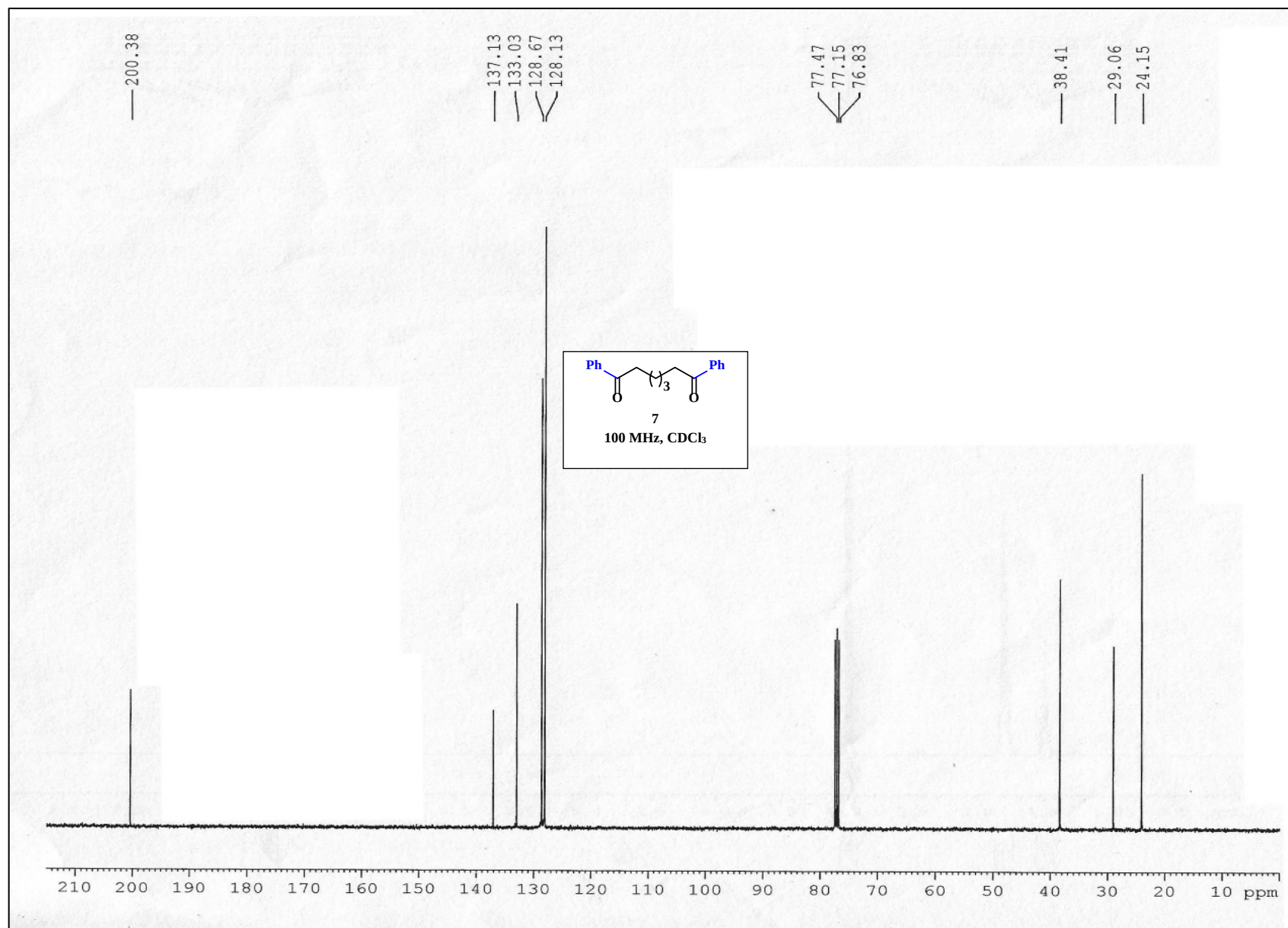


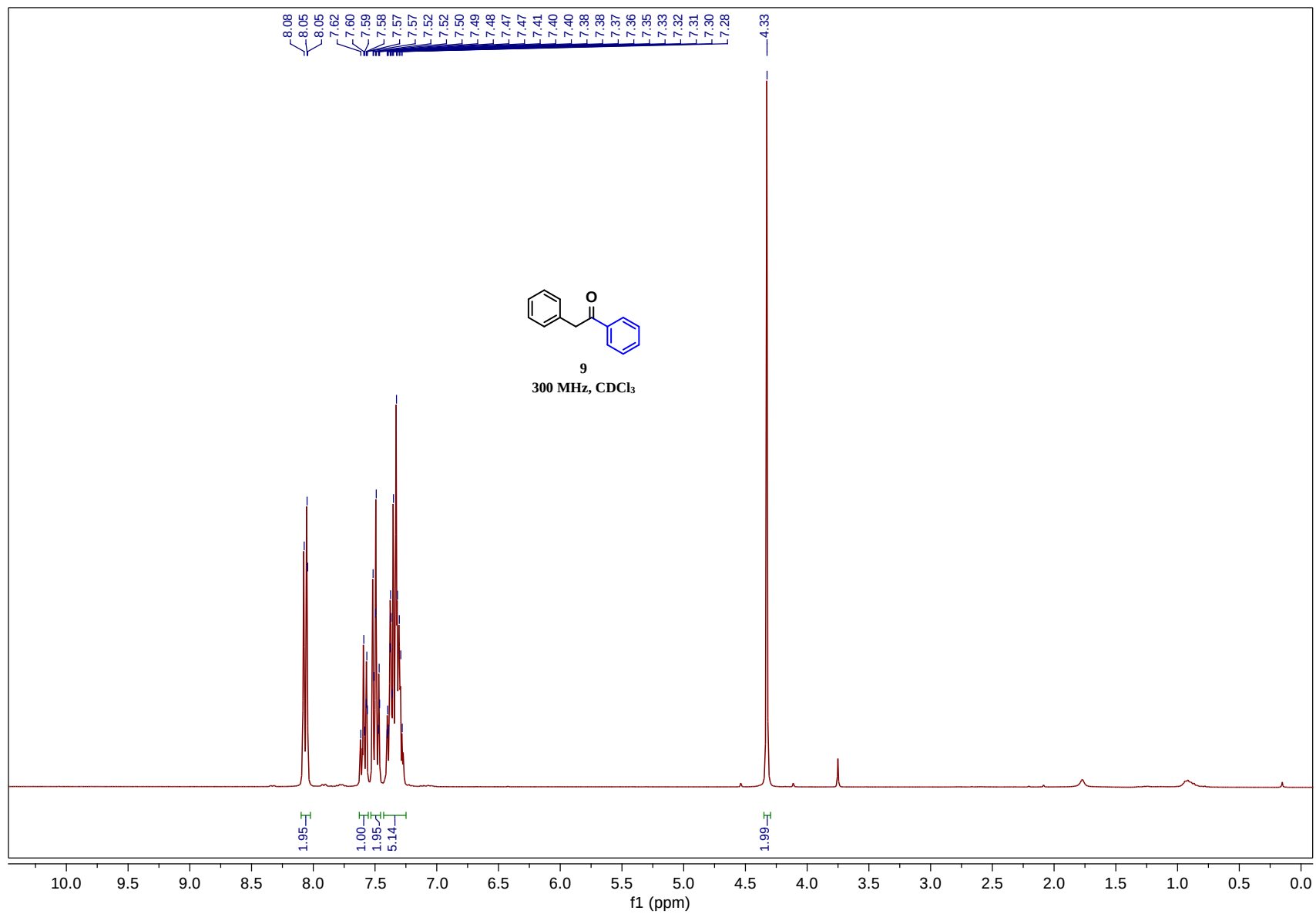
VI. ^1H and ^{13}C NMR spectra for the products 4x, 7, 9, 11, 14 and 15

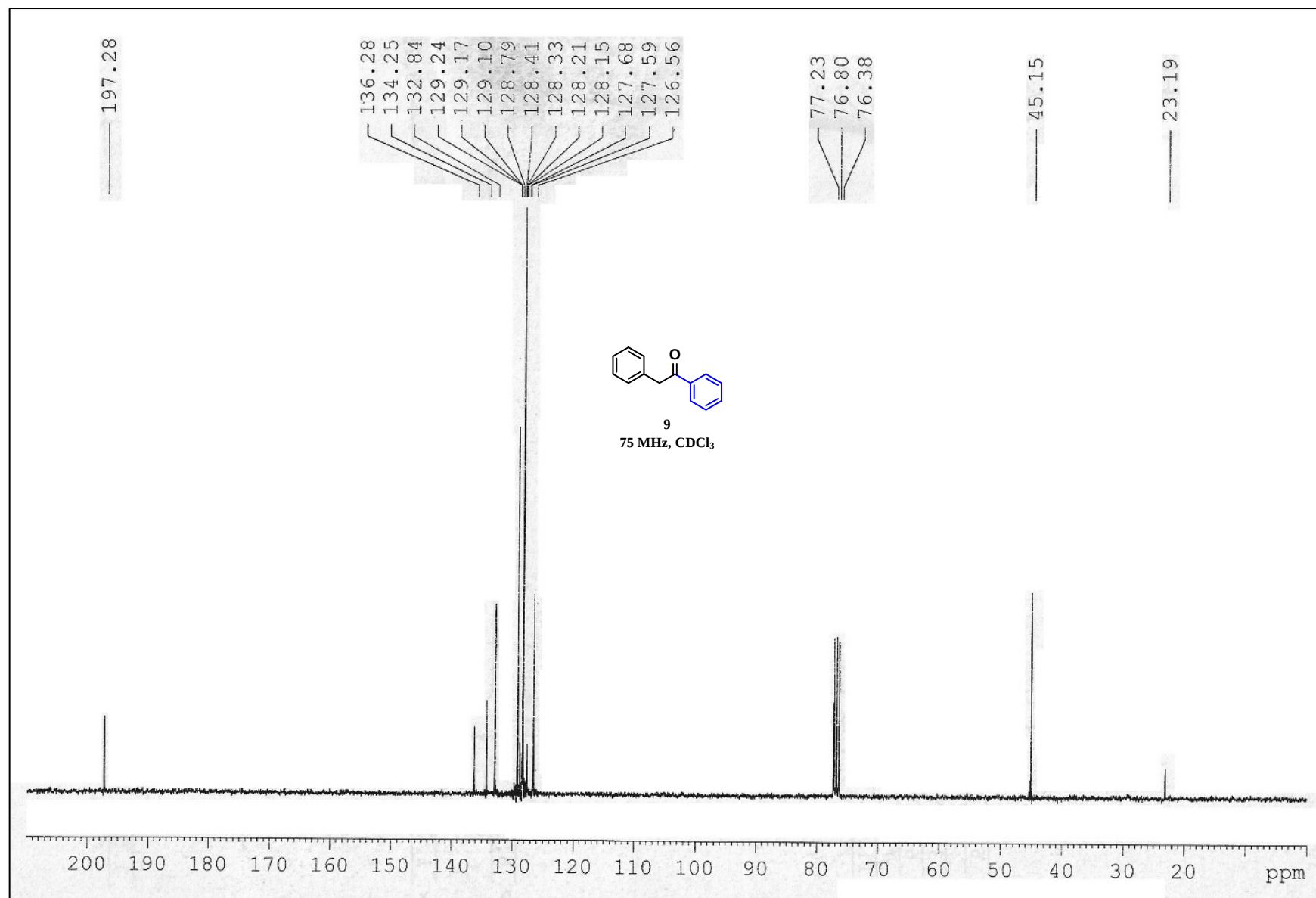


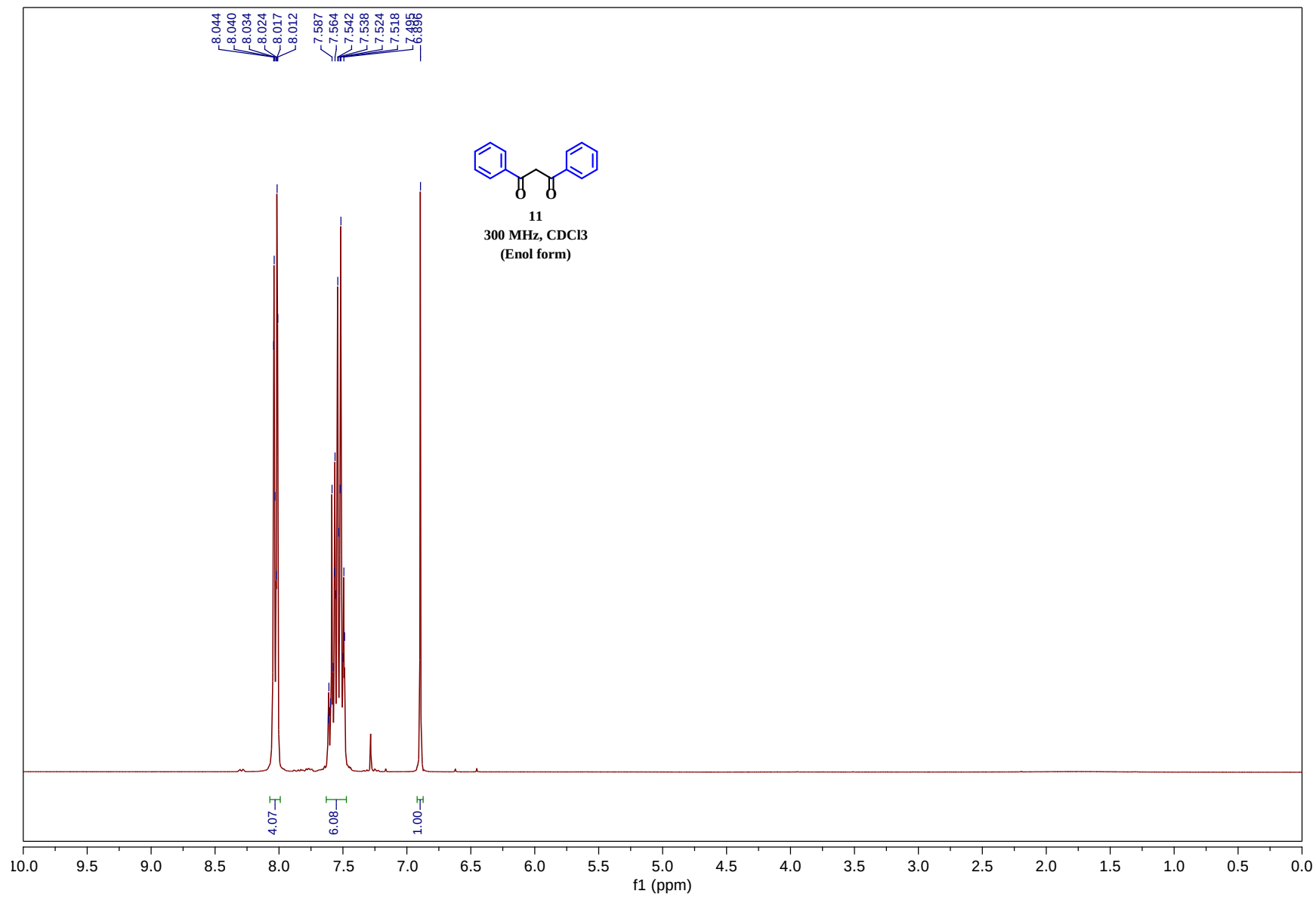




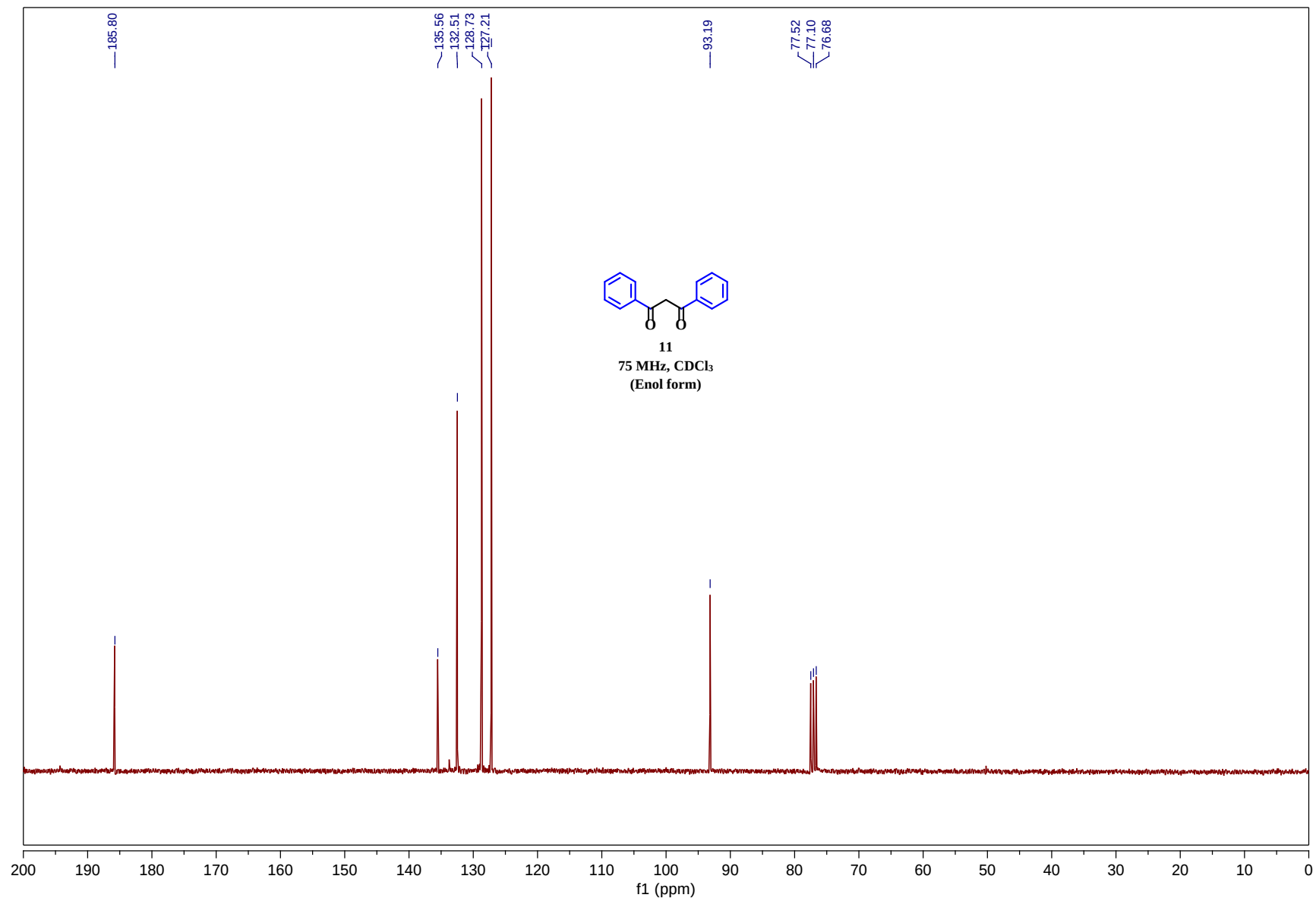




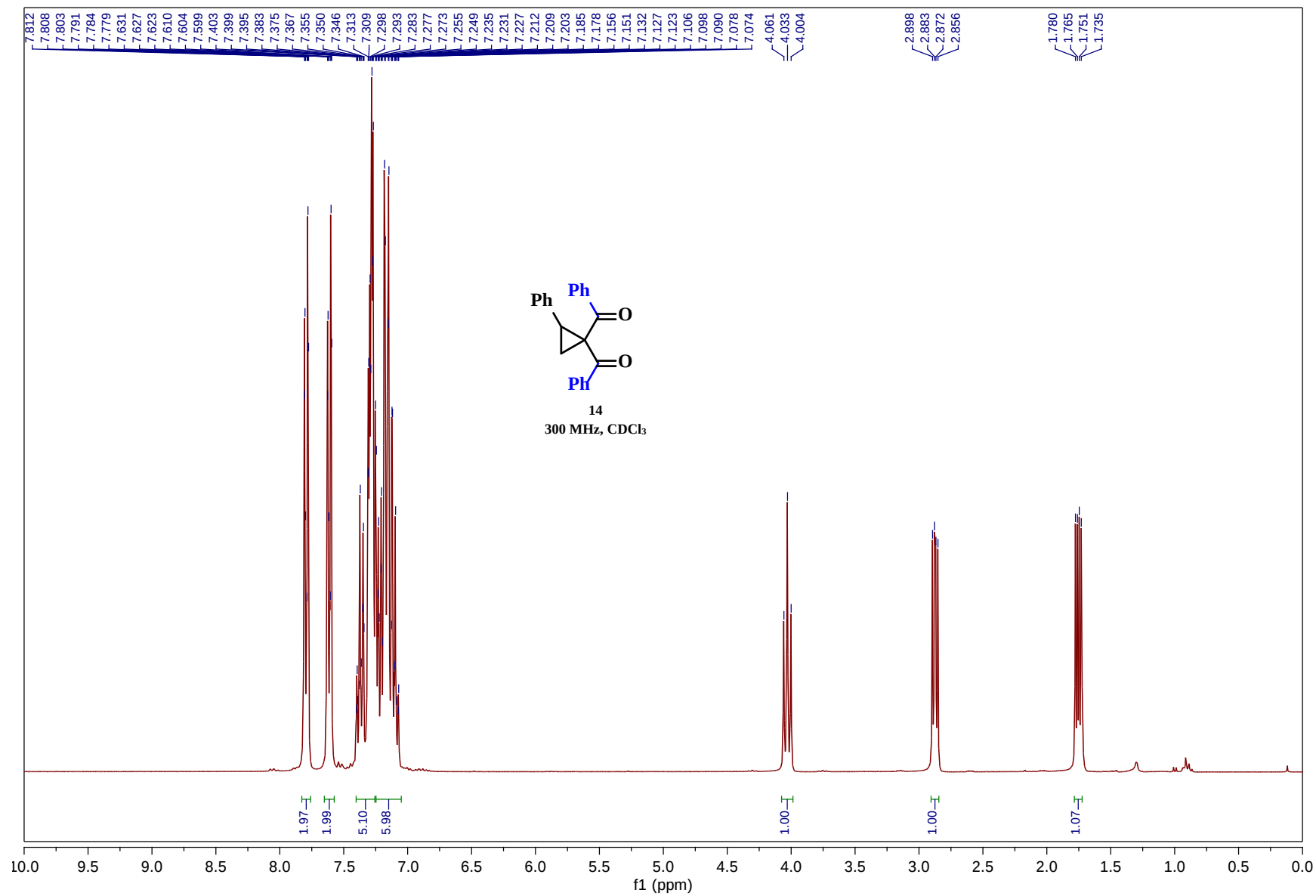


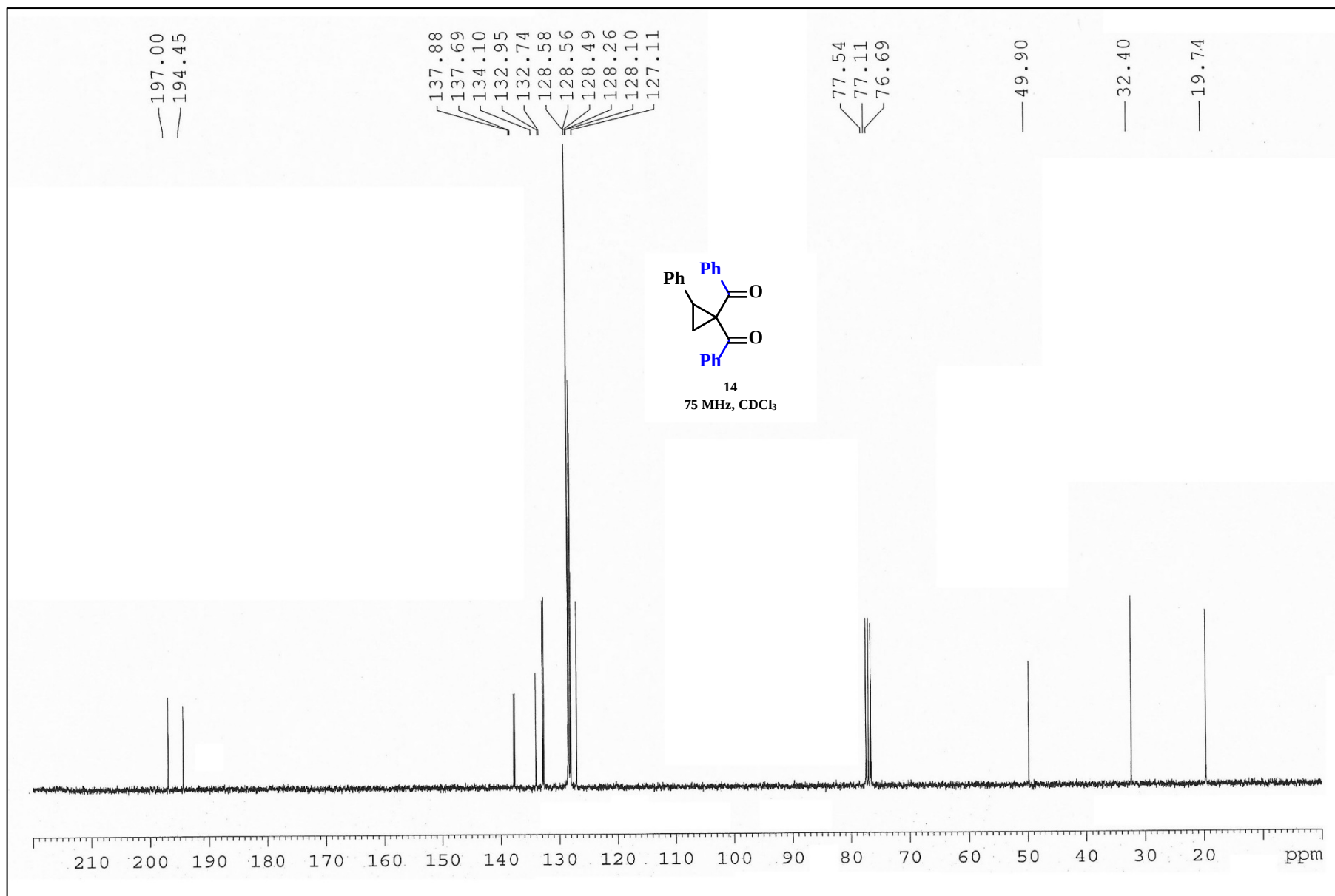


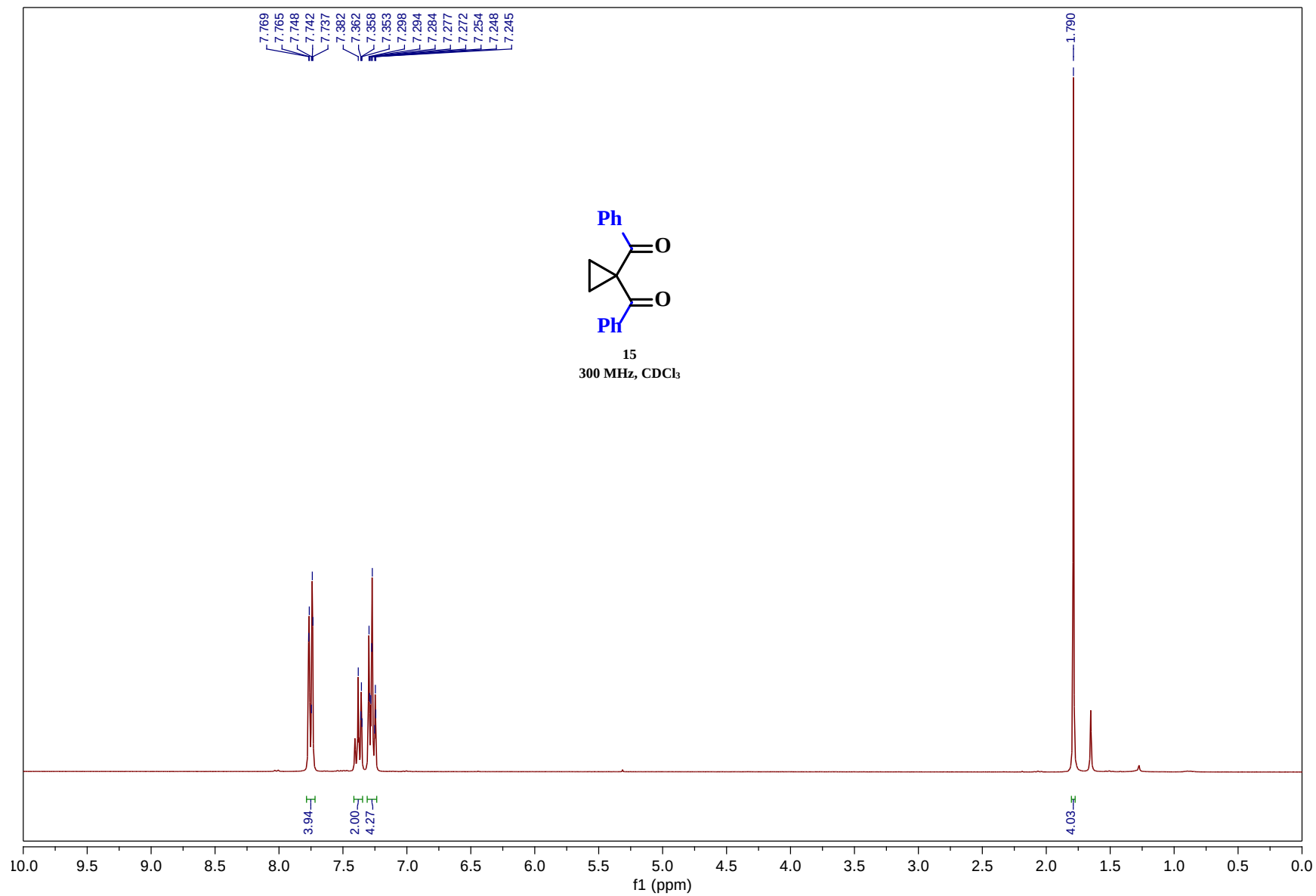
S81



S82







S85

