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Supporting Information

Zinc Triflate Catalyzed 1,3-Indolylation of Cyclohexanones: Tandem Condensation, Dehydrogenation and Aromatization Sequence

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2. Hydrogen gas Evolution Experiment:

In control experiment, hydrogen gas evolution experiment was performed using GCMS technique. To a stirred mixture of *N*-methyl indole (**1a**) (200 mg, 1.52 mmol) in 1,4-Dioxane, cyclohexanone (**2a**) (90 mg, 0.92 mmol) and $Zn(OTf)_2$ (111 mg, 20 mol%) were added in ace pressure tube and covered by rubber septum attached with a tedlar bag. The mixture was stirred at 120 °C temperature. After 24 h, the tedlar bag was removed and reattached into the inlet of GCMS instrument for detecting the gases formed in the reaction mixture. The peak for hydrogen gas was observed from the reaction mixture, which exactly matched with the standard for identifying hydrogen gas (at 0.088 minutes' retention time) (**Figure S1-S2**).





3. Time dependent ¹H NMR of reaction mixture:

The mechanism was probed by time dependent NMR studies of the reaction mixture (**Figure S3**). The reaction of *N*-methyl indole (**1a**) (200 mg, 1.52 mmol) was performed with cyclohexanone (**2a**) (90 mg, 0.92 mmol) under the standard reaction conditions. At different time intervals, 20 μ L aliquot was drawn out from reaction mixture and ¹H NMR of the sample was carried out. The time dependent ¹H NMR of reaction mixture showed appearance of new olefinic proton peaks at 4.1 and 6.2 ppm at 6 h and 12 h of reaction time respectively.



Figure S3. Time dependent ¹H NMR spectra of reaction mixture.

4. HRMS (ESI-MS) of intermediate 1D:

The formation of intermediate **1D** is detected by HRMS of reaction mixture. For this, a reaction of *N*-methyl indole (**1a**) (200 mg, 1.52 mmol) was performed with cyclohexanone (**2a**) (90 mg, 0.92 mmol) under the standard reaction conditions. After 12 h, an aliquot was drawn out, extracted extracted with ethyl acetate and concentrated. The resulting suspension was diluted with acetonitrile, filtered and injected for HRMS study. HRMS of the reaction mixture showed mass/charge $[M+H]^+$ at 339.1869 for the intermediate **1D** (Figure S4).



Figure S4. HRMS of intermediate 1D.





Spectrum 2. 75 MHz ¹³C NMR of compound 3a







Spectrum 5. 75 MHz ¹³C NMR of compound 3b



Spectrum 6. 75 MHz DEPT-135 NMR of compound 3b



Spectrum 8. 100 MHz ¹³C NMR of compound 3c







Spectrum 12. 75 MHz DEPT-135 NMR of compound 3d



Spectrum 13. 400 MHz ¹H NMR of compound 3e



Spectrum 14. 75 MHz ¹³C NMR of compound 3e



Spectrum 16. 400 MHz ¹H NMR of compound 3f



Spectrum 17. 75 MHz ¹³C NMR of compound 3f



Spectrum 18. 75 MHz 135-DEPT NMR of compound 3f



Spectrum 20. 75 MHz ¹³C NMR of compound 3g









Spectrum 23. 75 MHz ¹³C NMR of compound 3h



Spectrum 24. 75 MHz DEPT-135 NMR of compound 3h



Spectrum 26. 75 MHz ¹³C NMR of compound 3i



Spectrum 28. 400 MHz ¹H NMR of compound 3j



Spectrum 29. 75 MHz ¹³C NMR of compound 3j



Spectrum 30. 75 MHz DEPT-135 NMR of compound 3j



Spectrum 32. 100 MHz ¹³C NMR of compound 3k





Spectrum 35. 75 MHz ¹³C NMR of compound 31



Spectrum 36. 75 MHz DEPT-135 NMR of compound 31





Spectrum 38. 75 MHz ¹³C NMR of compound 3m



Spectrum 40. 400 MHz ¹H NMR of compound 3n



Spectrum 41. 75 MHz ¹³C NMR of compound 3n



Spectrum 42. 75 MHz DEPT-135 NMR of compound 3n



Spectrum 44. 75 MHz ¹³C NMR of compound 30



Spectrum 45. 75 MHz DEPT-135 NMR of compound 30



Spectrum 46. 400 MHz ¹H NMR of compound 4a



Spectrum 48. 100 MHz DEPT-135 NMR of compound 4a



Spectrum 50. 100 MHz ¹³C NMR of compound 4b



Spectrum 52. 400 MHz ¹H NMR of compound 4c



Spectrum 54. 75 MHz DEPT-135 NMR of compound 4c



Spectrum 56. 75 MHz ¹³C NMR of compound 4d



Spectrum 58. 400 MHz ¹H NMR of compound 4e



Spectrum 59. 75 MHz ¹³C NMR of compound 4e



Spectrum 60. 75 MHz DEPT-135 NMR of compound 4e



Spectrum 62. 75 MHz ¹³C NMR of compound 4f



Spectrum 64. 400 MHz ¹H NMR of compound 4g



Spectrum 65. 100 MHz ¹³C NMR of compound 4g



Spectrum 66. 100 MHz DEPT-135 NMR of compound 4g



Spectrum 68. 75 MHz ¹³C NMR of compound 4h



Spectrum 70. 400 MHz ¹H NMR of compound 4i



Spectrum 72. 75 MHz DEPT-135 NMR of compound 4i





Spectrum 74. 100 MHz ¹³C NMR of compound 4j







Spectrum 77. 75 MHz ¹³C NMR of compound 5a



Spectrum 78. 75 MHz DEPT-135 NMR of compound 5a



Spectrum 80. 75 MHz ¹³C NMR of compound 5b



Spectrum 82. 400 MHz ¹H NMR of compound 5c







Spectrum 84. 75 MHz DEPT-135 NMR of compound 5c



Spectrum 85. 400 MHz ¹H NMR of compound 6



Spectrum 86. 75 MHz ¹³C NMR of compound 6

9. Crystallographic Description:

Data collection and refinement single-crystal X-ray data of compounds was collected on Bruker SMART CCD Diffractometer using graphite monochromated MoK α radiation ($\lambda = 0.71073$ Å). Frames were collected at T = 297 K by ω , φ , and 2 θ -rotations with full quadrant data collection strategy (four domains each with 600 frames) at 10s per frame with SMART. The measured intensities were reduced to F² and corrected for absorption with SADABS.¹ Structure solution, refinement, and data output were carried out with the SHELXTL package by direct methods.² Non-hydrogen atoms were refined anisotropicallyusing the WinGX (version 1.80.05) program package.³ All non-hydrogen atoms were refined anisotropically and hydrogen atoms were treated as riding atoms using SHELX default parameters. Molecular structures have drawn using ORTEP software shown in **figure S5**. Further information on the crystal structure determination (excluding structure factors) has been given as **table S1** and also deposited in the Cambridge Crystallographic Data Centre as supplementary publications number 1918563. Copies of the data can be obtained free of charge upon application to CCDC, 12 Union Road, Cambridge CB2 1EZ, UK (fax: (+44) 1223-336-033. e-mail: deposit@ccdc.cam.ac.uk) or via internet.



Figure S5. ORTEP diagram of 3a.

Identification code	DD_NA_0m_a	
Empirical formula	C ₂₄ H ₂₀ N ₂	
Formula weight	336.42	
Temperature	297 K	
Wavelength	0.71073 Å	
Crystal system	Monoclinic	
Space group	C2/c	
Unit cell dimensions	a = 15.931(2) Å	$\alpha = 90^{\circ}$.
	b = 14.7572(15) Å	$\beta = 106.219(4)^{\circ}.$
	c = 8.0025(9) Å	$\gamma = 90^{\circ}$.
Volume	$1806.5(4) \text{ Å}^3$	
Ζ	4	
Density (calculated)	1.237 g/cm^3	
Absorption coefficient	0.073	
F(000)	712.0	
Crystal size	0.22 x 0.12 x 0.10 mm ³	
Theta range for data collection	2.66 to 24.68°.	
Index ranges	-16<=h<=18, -15<=k<=17, -	
	9<=1<=7	
Reflections collected	1572	
Independent reflections	4820 [R(int) = 0.0397]	
Completeness to theta = 24.68°	97.7%	
Absorption correction	None	
Refinement method	Full-matrix least-squares on F ²	
Data / restraints / parameters	4820 / 0 / 120	
Goodness-of-fit on F ²	1.017	
Final R indices [I>2sigma(I)]	R1 = 0.0551, WR2 = 0.1448	
R indices (all data)	R1 = 0.0900, wR2 = 0.1312	
CCDC	1918563	

Crystallographic description of 1,3-bis(1-methyl-1H-indol-3-yl)benzene (3a) (Table S1):

10. References:

(1) SADABS V2.10 (Sheldrick, G. M. 2003).

- (2) (a) Sheldrick, G. M. Acta Crystallogr., Sect. A: Found. Crystallogr., 1990, 46, 467; (b) Sheldrick, G. M.
- SHELXL-NT Version 6.12, University of Gottingen, Germany, 2000.

(3) Farrugia, L. J. J. Appl. Cryst. 1999, 32, 837.