Insight into the structure-property relation of UO₂ nanoparticles

Received 00th January 20xx, Accepted 00th January 20xx

DOI: 10.1039/x0xx00000x

Evgeny Gerber^{a,b,c}, Anna Yu. Romanchuk^c, Stephan Weiss^b, Stephen Bauters^{a,b}, Bianca Schacherl^d, Tonya Vitova^d René Hübner^b, Salim Shams Aldin Azzam^b, Dirk Detollenaere^{e,f}, Dipanjan Banerjee^{f,g}, Sergei M. Butorin^h, Stepan N. Kalmykov^c and Kristina O. Kvashnina^{a,b,c*}

- ^{a.} The Rossendorf Beamline at ESRF The European Synchrotron, CS40220, 38043 Grenoble Cedex 9, France. ^{b.} Helmholtz-Zentrum Dresden-Rossendorf (HZDR), Institute of Resource Ecology, PO Box 510119, 01314, Dresden.
- ^{c.} Lomonosov Moscow State University, Department of Chemistry, 119991 Moscow, Russia.
- d. Institute for Nuclear Waste Disposal (INE), Karlsruhe Institute of Technology, P.O. 3640, D-76021 Karlsruhe, Germany
- e. Department of Chemistry, X-ray Imaging and Microspectroscopy Research Group, Ghent University, Ghent, Belgium
- f Dutch-Belgian Beamline (DUBBLE), European Synchrotron Radiation Facility, 71 Avenue des Martyrs, CS 40220, 38043 Grenoble Cedex 9, France
- ^{g.} Department of Chemistry, KU Leuven, Celestijnenlaan 200F, Box 2404, B-3001 Leuven, Belgium

h Molecular and Condensed Matter Physics, Department of Physics and Astronomy, Uppsala University, P.O. Box 516, Uppsala, Sweden

Supplementary information



Experimental

Fig. S1. UV-vis spectra of 0.1M uranium solution in 0.5M before and after reduction.

Journal Name



Fig. S2. Photo of the sample cell, including triple holders.

Table S1. pH and Eh values after	r the end of the reactions.
----------------------------------	-----------------------------

Sample	рH	Eh,	
Sample	рп	mV	
0.01 M U(IV) pH 8	8.0	-272	
0.01 M U(IV) pH >11	11.6	-287	
0.1 M U(IV) pH 8	8.0	-84	
0.1 M U(IV) pH >11	10.9	-482	



Fig. S3. HRTEM images of NPs and corresponding SAED patterns (inset): a) 0.01M U(IV) pH >11, b) 0.1M U(IV) pH 8, c) 0.1M U(IV) pH >11.



Fig. S4. XRD patterns of the precipitates from U(IV) with different pH and concentrations, UO_2 and U_4O_9 references.

Sample	Size, nm (XRD)	Size, nm (HRTEM)
0.01 M U(IV) pH 8	2.5±0.6	2.7±0.8
0.01 M U(IV) pH >11	1.7 (one peak)	3.1±0.7
0.1 M U(IV) pH 8	2.5±0.5	3.8±0.9
0.1 M U(IV) pH >11	1.8±0.3	2.7±0.8
0.1 M U(IV) pH 8 aged	6.0±0.2	6±1
0.1 M U(IV) pH >11 aged	4.8±0.1	6±1

Table S2. Particle size estimation of the samples by XRD and HRTEM.



Fig. S5. U M₄ HERFD experimental data. The increase of the peak asymmetry is notable.

Calculations of HERFD spectra:

The calculations of the U(IV) M₄ HERFD spectra were performed using crystal field theory. To obtain the HERFD spectra, the coreto core (3d-4f) resonant inelastic x-ray scattering (RIXS) intensity maps were calculated on the emission versus incident photon energy scales and a cut at the constant emission energy, corresponding to the maximum of the RIXS intensity was made along the incident photon energy axis. The RIXS maps were calculated in a manner described in literature^{1,2} by taking into account the full multiplet structure due to intra-atomic and crystal field interactions. The Slater integrals $F^k(5f,5f)$, $F^k(3d,5f)$, $F^k(4f,5f)$ as well as $G^k(3d,5f)$ and $G^k(4f,5f)$ calculated for the U(IV) ion were scaled down to 80 % of their *ab-initio* Hartree-Fock values. The ground, intermediate and final states of the spectroscopic process were represented by the $3d^{10}5f^2$, $3d^95f^3$ and $4f^{13}5f^3$ configurations, respectively. To simulate different U(IV) environment in terms of varying U coordination number (CN), the calculations were done in tetrahedral (T_d), octahedral (O_h⁶), cubic (O_h⁸) and icosahedral (I_h) symmetry for CN equal to 4, 6, 8 and 12, respectively. The Wybourne's crystal field parameters in the calculations were set to Td: $B^4_0 = -0.30 \text{ eV}$, $B^6_0 = -0.70 \text{ eV}$; O_h^6 : B^4_0 = 0.93 eV, $B^6_0 = 0.35 \text{ eV}$; O_h^8 : $B^4_0 = -0.93 \text{ eV}$, $B^6_0 = 0.35 \text{ eV}$; I_h : $B^6_0 = 0.65$ (-0.65) eV. For I_h symmetry both positive and negative values of the B^6_0 parameter were used which produce different results. The abovementioned choice of values was based on already established values for UO₂³ and relative changes of parameter values for different symmetries discussed in literature.^{4,5}





Fig. S6. M₄ HERFD spectra calculated for the U(IV) ion in environments with different coordination number (CN).





Fig. S7. U L3-EXAFS results. a) U L3-EXAFS spectra $\chi(R)$ fit results, Fourier transform (FT) magnitude of experimental EXAFS data (black) and shell fit (red) with U-O, U-U shells.

Table S3. Metric p	arameters extracted by least-squares fit anal	ysis of U L ₃ EXAFS spectra with U-O, U-U shells,	, (k range of 2.	0-11.5 Å ⁻¹).

Table 55. Mictile p								
First coordination shell			on shell	Second coordination shell			AF 1-341	R-factor,
Sample	CN	R [Å]	σ² [Ų]	CN	R [Å]	σ² [Ų]	ΔE₀ [eV]	%
0.01M U(IV)	4.9(0.6)	2.33(0.01)	0.011	3.6(1.0)	3.855(0.008)	0.005	EO	3.1
pH 8	0	2.33(0.01)	0.011	U	3.855(0.008)	0.005	5.9	3.1
0.01M U(IV)	4.2 O	2 22/0 01	0.011	2.9	2 861(0.007)	0.004	7.0	26
pH >11	(0.5)	2.33(0.01)	0.011	U(0.8)	3.861(0.007)	0.004	7.0	2.6
UO ₂ structure	80	2.3677		12 U	3.8665 U			

Reactivity of the NPs



Fig. S8. HRTEM images of the one-year-old 0.1 M U(IV) pH >11 sample at different scales: a), b) scale bar 20 nm, different analysis regions, c) scale bar 10 nm.



Fig. S9. X-ray diffraction patterns measured from fresh and 1-year old samples.

References:

- 1 S. M. Butorin, K. O. Kvashnina, A. L. Smith, K. Popa and P. M. Martin, Crystal-Field and Covalency Effects in Uranates : An X-ray Spectroscopic Study, *Chem. Eur. J.*, 2016, **22**, 9693–9698.
- 2 S. M. Butorin, K. O. Kvashnina, J. R. Vegelius, D. Meyer and D. K. Shuh, High-resolution X-ray absorption spectroscopy as a probe of crystal-field and covalency effects in actinide compounds, *Proc. Natl. Acad. Sci.*, 2016, **113**, 8093–8097.
- S. M. Butorin, A. Modin, J. R. Vegelius, K. O. Kvashnina and D. K. Shuh, Probing Chemical Bonding in
 Uranium Dioxide by Means of High- Resolution X-ray Absorption Spectroscopy, *J. Phys. Chem. C*, 2016, **120**, 29397–29404.
- 4 N. Edelstein, Optical and magnetic properties of tetravalent actinide ions and compounds, *J. Less Common Met.*, 1987, **133**, 39–51.
- 5 J. C. Krupa, Spectroscopic properties of tetravalent actinide ions in solids, *Inorganica Chim. Acta*, 1987, **139**, 223–241.