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Supporting information

An electrochemical aptasensor for lead ion detection based on catalytic hairpin

assembly and porous carbon supported platinum as signal amplification

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Experimental

Apparatus and electrochemical measurement

Electrochemical impedance spectroscopy (EIS) and differential pulse voltammetry (DPV) measurements were carried out with a CHI 660D electrochemical workstation (Shanghai Chenhua Instrument, China). A conventional three-electrode system: a platinum wire as counter electrode, a saturated calomel as reference electrode and the modified Au electrode (AuE, Φ =3 mm)) as working electrode. The prepared electrochemical biosensor was performed by DPV in PBS (pH 7.4) containing 10 mM HQ and 10 mM H₂O₂. EIS measurement were carried out in 5 mL PBS solution (0.1 M, pH 7.4) containing 5.0 mM [Fe(CN)₆]^{3-/4-} and 0.5 M KCl as redox probe with a potential range of 0.1Hz to 10kHz and a scan rate of 100 mV s⁻¹.

Analysis of the crystal structure of carbon materials are carried out using X-ray powder diffractometer (XRD, Bruker D8 Advance) from Bruker, Germany. Scanning electron microcopy (SEM, JSM-7610F JEOL, Japan) and Transmission electron microscopy (TEM, JEOL 2010F) are used to characterize the prepared samples.

Results and discussion



Fig.S1. DPV results of (a) bare electrode only containing HQ; (b) bare electrode containing HQ and H₂O₂; (c) PtNPs@PCs modified electrode only containing HQ; (d)

PtNPs@PCs modified electrode containing HQ and H2O2.

Method	Strategy	Linear range (nM)	LOD (nM)	Reference
SWASV	Au@PANI	20 to 720	3	1
DPASV	Au-CFs	100 to 2000	5	2
DPASV	MWCNT-PARS	24.1 to 724	2.27	3
SWASV	Fe ₃ O ₄ /MWCNTs/LSG	4.82 to 965.3	0.337	4
SWASV	AgNPs@p-1,8-DAN	1000 to 1200	0.15	5
DPV	ssDNAzyme	0.5 to 5000	0.25	6
EIS	OMC-GNPs	0.5 to 50000	0.2	7
EIS	AQMS	1 to 100	0.33	8
DPV	TdTase enzyme	0.05 to 500	0.043	9
DPV	AuNPs	0.6 to 50	0.312	10
DPV	СНА	0.04 to 3000	0.027	11
DPV	Fe-MOF&AuNPs	0.03 to 1000	0.02	12
DPV	PtNPs@PCs, CHA	0.05 to 1000	0.018	This work

Table S1 Analytical performances of different electrochemical methods for Pb²⁺

detection

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