

Modulation of Superficial Electronic Structure via Metal-Support Interaction for H₂ Evolution over Pd Catalysts

Jin Wang,^a Dan Cheng,^{b,c} Mengmeng Gao,^a Qian Li,^a Ying Xin,^a Nana Zhang,^a Zhaoliang Zhang,^{* a,d}
Xuehua Yu,^d Zhen Zhao,^{*d} and Kebin Zhou,^{*b}

a School of Chemistry and Chemical Engineering, Shandong Provincial Key Laboratory of Fluorine Chemistry and Chemical Materials, University of Jinan, Jinan 250022, P.R. China. Email: chm_zhangzl@ujn.edu.cn

b School of Chemical Sciences, University of Chinese Academy of Sciences, Beijing 100049, P.R. China. Email: kbzhou@ucas.ac.cn

c Department of Biological Sciences, National University of Singapore, 117543, Singapore.

d Institute of Catalysis for Energy and Environment, College of Chemistry and Chemical Engineering, Shenyang Normal University, Shenyang 110034, Liaoning, P. R. China. Email: zhenzhao@cup.edu.cn

1. Supplementary Results

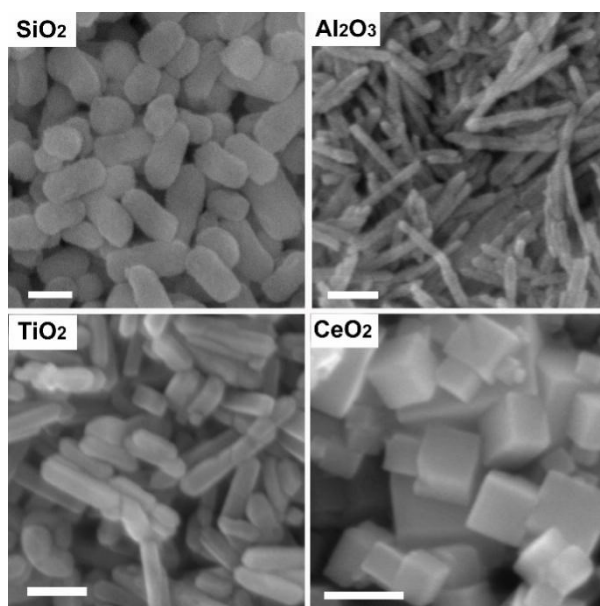


Figure S1. SEM images of the four metal oxide supports: SiO₂, Al₂O₃, TiO₂ and CeO₂.

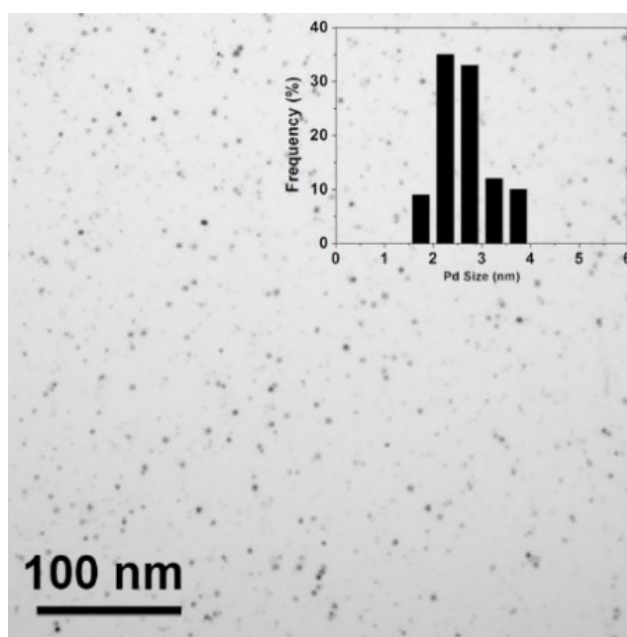


Figure S2. TEM images of colloidal Pd nanoparticles and the histogram of the particle-size distributions (inset).

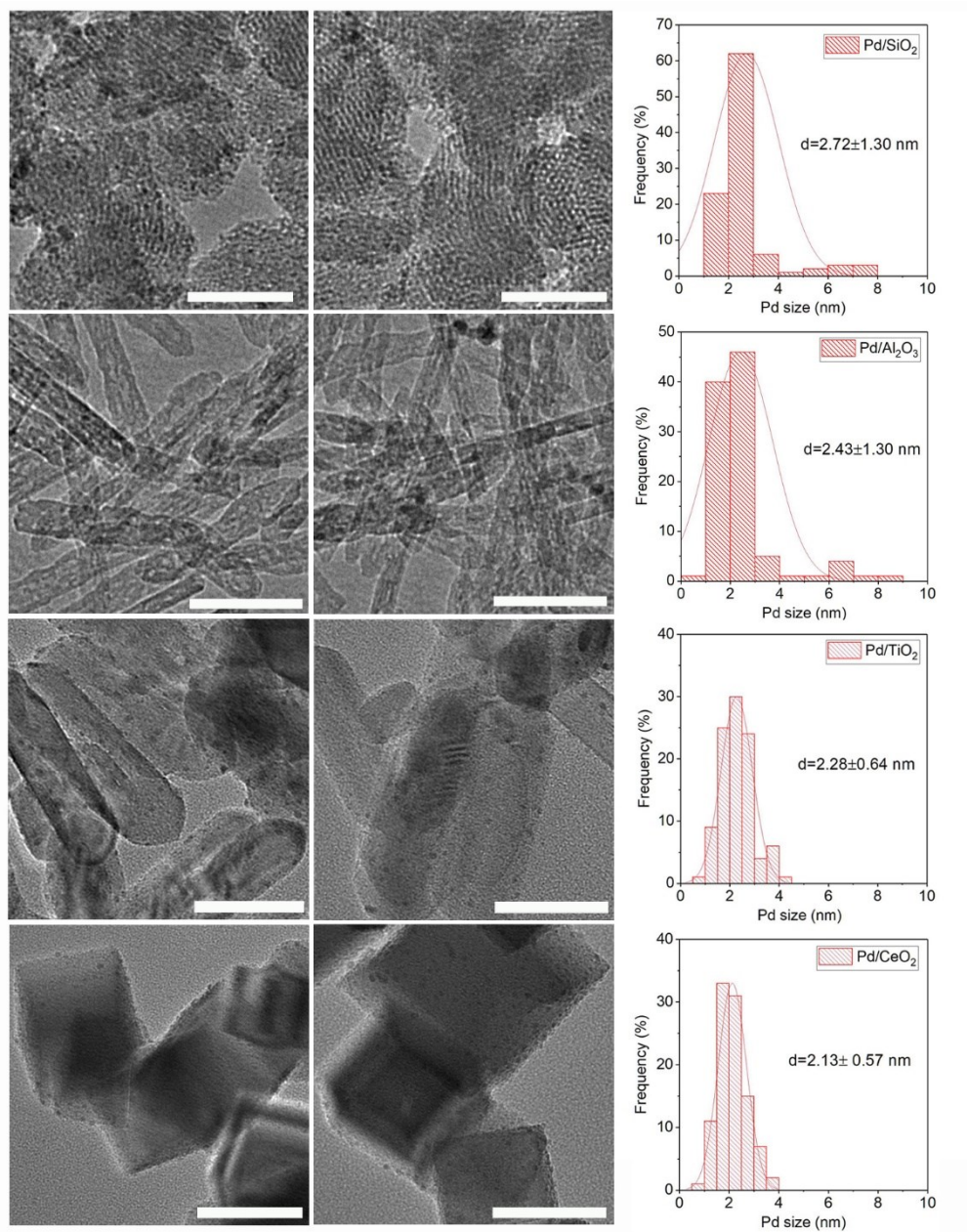


Figure S3. The TEM images of Pd/MO_x catalysts and the corresponding histogram of the particle size distribution. Scale bar is 20 nm.

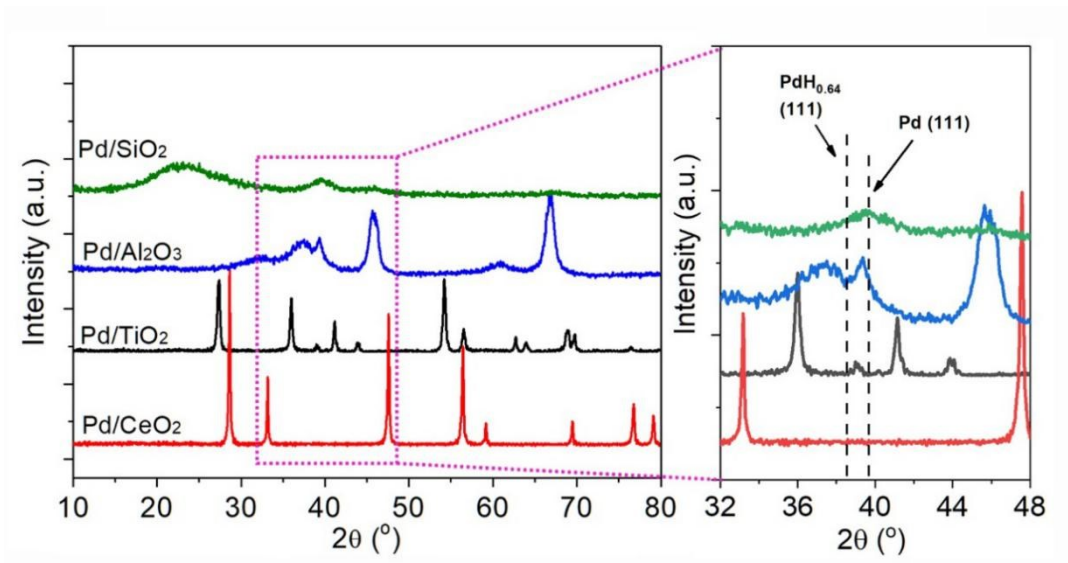


Figure S4. XRD patterns of Pd/SiO₂, Pd/Al₂O₃, Pd/TiO₂ and Pd/CeO₂, as well as enlarged patterns to analyze Pd crystal and PdH_{0.64} crystal phases.

Note: According to XRD patterns of Pd/MO_x, peaks of silica, alumina, rutile TiO₂ and ceria were observed. The enlarged patterns do not show Bragg reflections of Pd (JCPDS 87-0641) resulting from the small size of the Pd species except for Pd/SiO₂, due to the much weak diffraction peaks of SiO₂ nanocrystals. The peaks of Pd (111) and PdH_x (111) are located at 39.6° and 38.7°, respectively, indicating with dot lines.

Table S1. The Pd loading of various Pd/MO_x catalysts.

samples	Pd/SiO ₂	Pd/Al ₂ O ₃	Pd/TiO ₂	Pd/CeO ₂
Pd loading (wt.%)	1.71	1.70	1.75	1.73

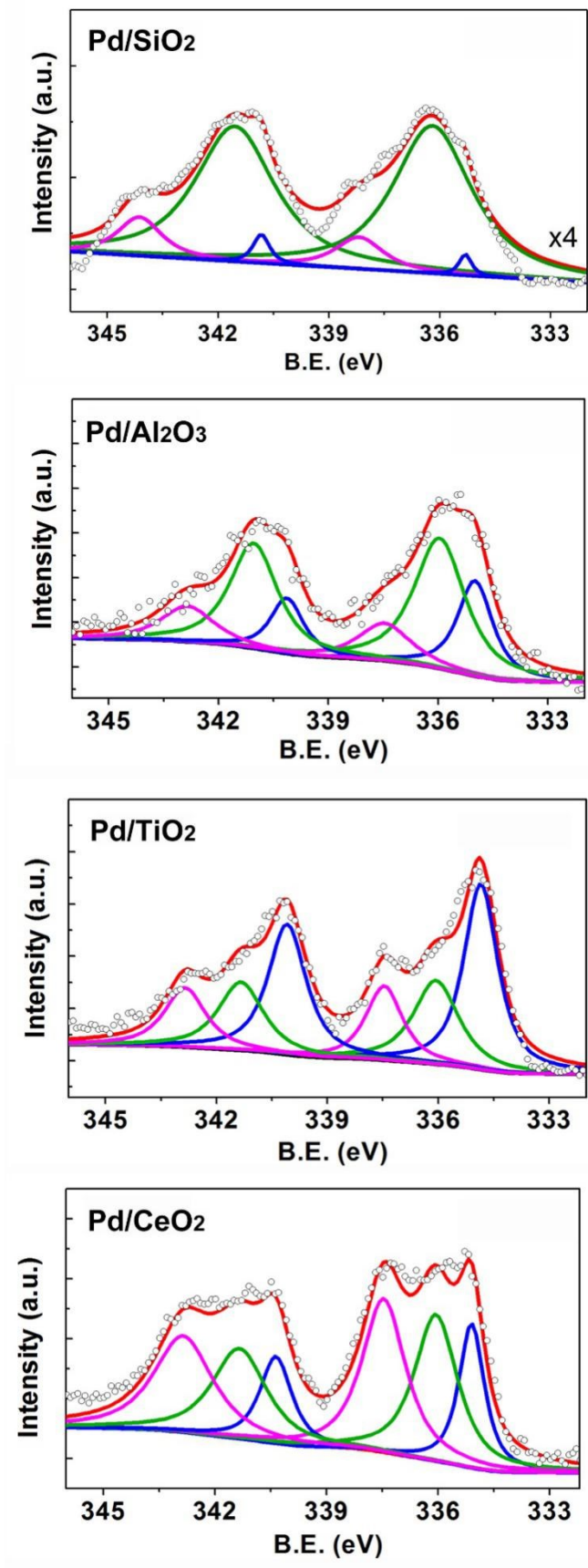


Figure S5. Pd 3d XPS profiles of Pd/MO_x catalysts.

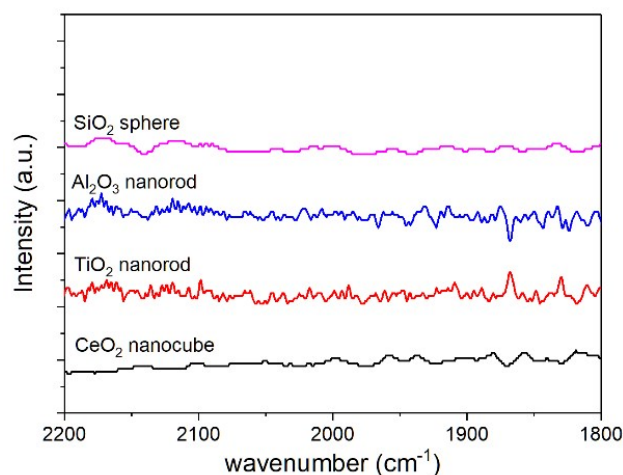


Figure S6. In situ DRIFTS spectra of CO chemisorption over MO_x at 25 °C. And no CO adsorption bands were observed on pure oxides.

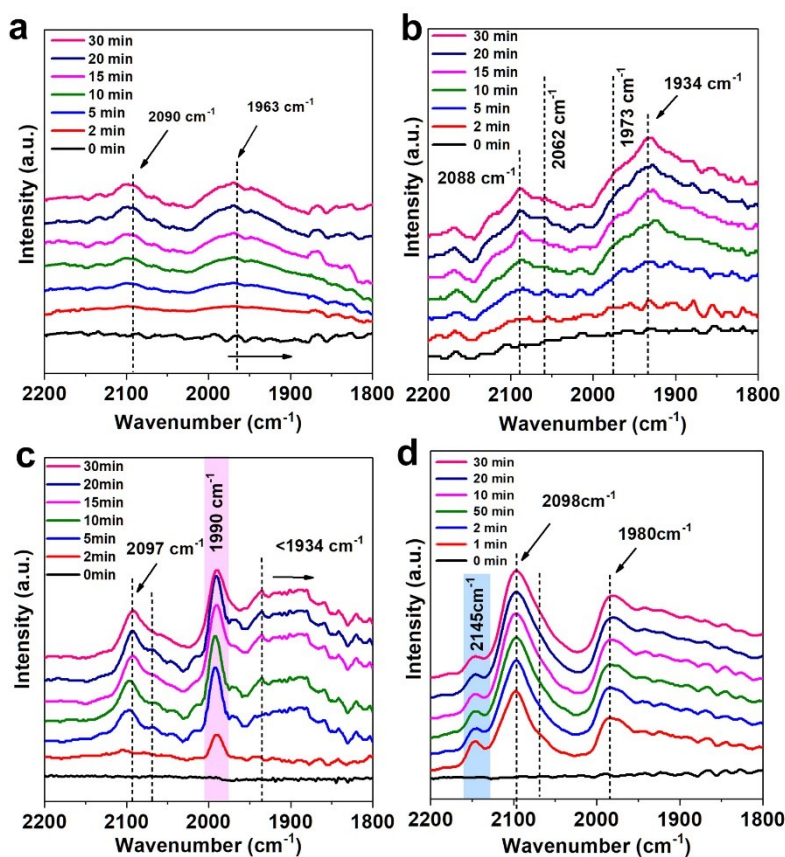


Figure S7. In situ DRIFTS spectra of CO chemisorption over Pd/MO_x at 25 °C: (a) Pd/SiO_2 ; (b) $\text{Pd}/\text{Al}_2\text{O}_3$; (c) Pd/TiO_2 ; (d) Pd/CeO_2 . The Pd/MO_x catalysts were pretreated under Ar at 25 °C for 60 min and then collect spectra once exposed to CO/Ar atmosphere. The decreasing of saturated

adsorption time suggests that the CO adsorption strength increase from Pd/SiO₂ (~20min), Pd/Al₂O₃ (~15min) and Pd/TiO₂ (5min) to Pd/CeO₂(<2min).

More details of the CO adsorption structure and additional explanations:

Both linear and multi-coordinated CO_Pd⁰ bands (CO adsorption on Pd⁰ species) present in Pd/SiO₂ and Pd/Al₂O₃ clearly, because of a large amount of Pd⁰ species exposing to surface (86.6% and 63.1%, respectively). As the surface Pd⁰ species decreases, the peak intensity declines obviously in Pd/TiO₂. Even for Pd/CeO₂, there is not multi-coordinated CO_Pd⁰, due to the dominant Pd²⁺ and Pd^{δ-} species on the surface.

For the CO_Pd^{δ-} bands (CO adsorption on Pd^{δ-} species), both linear and multi-coordinated CO_Pd^{δ-} present in Pd/Al₂O₃, Pd/TiO₂ and Pd/CeO₂, as the Pd^{δ-} species were primarily located on superficial, which is beneficial to the formation of multiple CO adsorption structures. It is generally accepted that the red-shift of the stretching wavenumber of adsorbed CO is indicative of an increasing electron donation from metals into the π* C=O orbital in the case of CO adsorption on electron-rich metal sites^{1,2}. By analogy, the peak at 2066 cm⁻¹ for Pd/TiO₂ also could be considered as a relative red-shifted peak from 2097 cm⁻¹. Furthermore, the red shift of 31 cm⁻¹ (from 2097 cm⁻¹ to 2066 cm⁻¹) is much larger than that of Ag@Pd NPs (11 cm⁻¹)², poly(allylamine)-stabilized Pd (~20 cm⁻¹)¹, and pyridinic-N-tuned Pd (<15 cm⁻¹)³, demonstrating that TiO₂ as a support promoted superficial Pd atoms to gather more electrons.

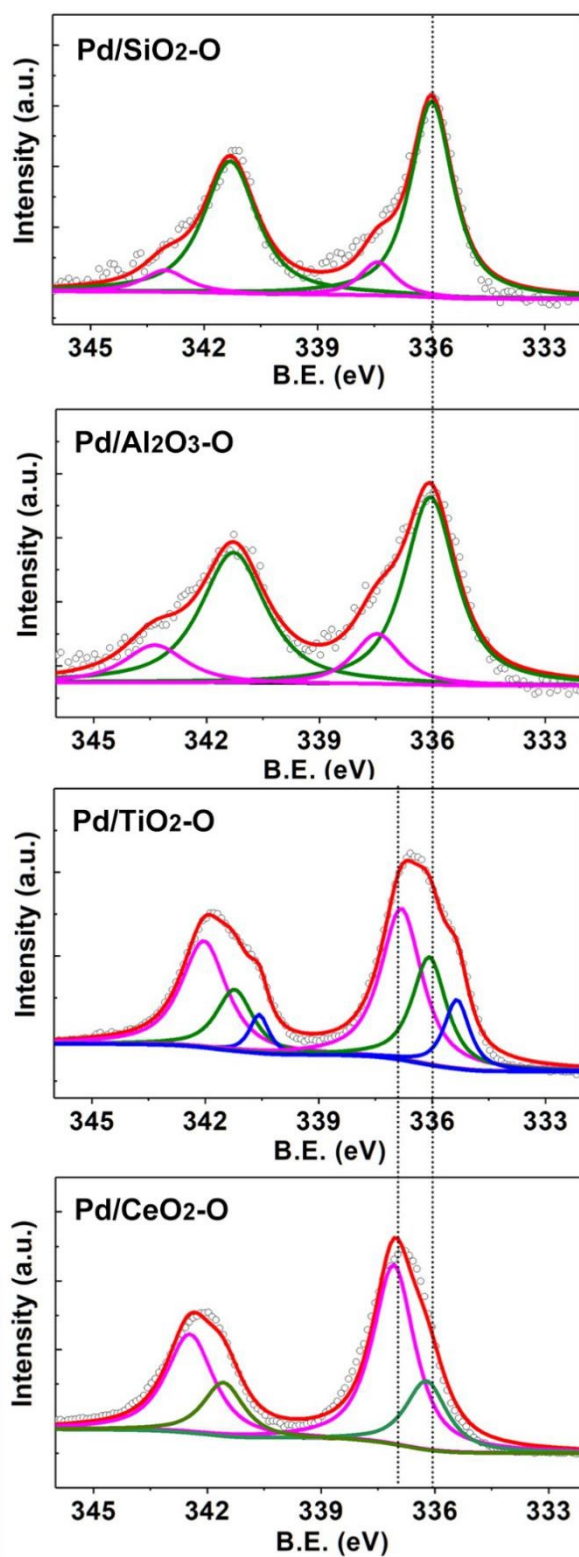


Figure S8. Pd 3d XPS profiles of Pd/MO_x after reoxidation treatment at 300°C under O₂/N₂ and followed by 5% H₂/Ar flushing for 30 min at 25 °C (denoted as Pd/MO_x-O).

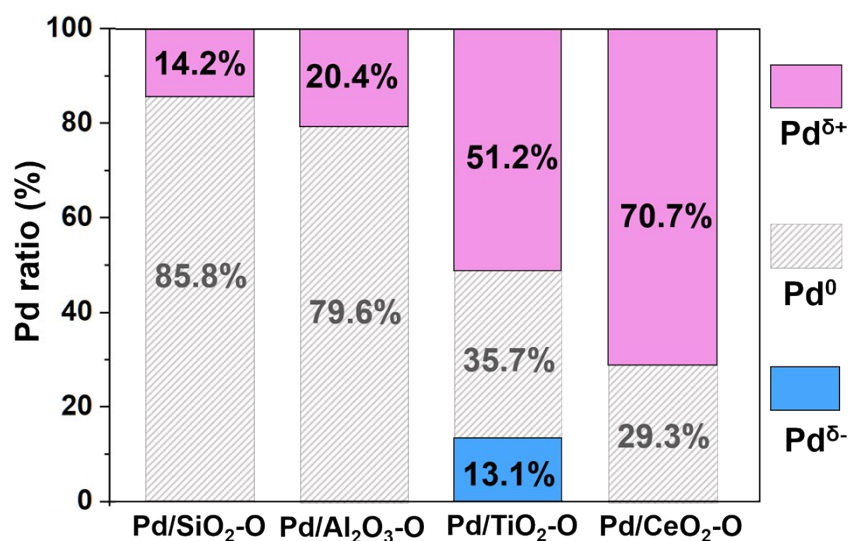


Figure S9. The relative proportions of Pd^{δ+}, Pd⁰ and Pd^{δ-} species in Pd/MO_x-O.

Note: The Pd/MO_x samples were re-oxidized at 300 °C under O₂/N₂ and followed by 5% H₂/Ar flushing for 30 min at 25 °C (denotated as Pd/MO_x-O). As shown in Figure S8 and S9, metallic Pd is the dominant species for Pd/SiO₂-O and Pd/Al₂O₃-O (almost 85.8% and 79.6%, respectively). In the case of Pd/TiO₂-O, the multi-peak fitting for the asymmetric peak suggests that Pd⁰, Pd^{δ-} and Pd^{δ+} species coexisted in the sample. Furthermore, the Pd²⁺ species in Pd/MO_x transform into Pd^{δ+} species in Pd/MO_x-O indicating that electrons partly transfer from the supports back to metals under oxidative conditions, and thus the *p*EMSI is weakened. Due to the strong interaction between Pd NPs and CeO₂, Pd^{δ+} species is dominant in Pd/CeO₂-O. However, in Pd/SiO₂-O and Pd/Al₂O₃-O, the absence of Pd^{δ-} species suggests that the *p*EMSI effect almost disappeared, due to the weak interaction between Pd NPs and SiO₂ or Al₂O₃.

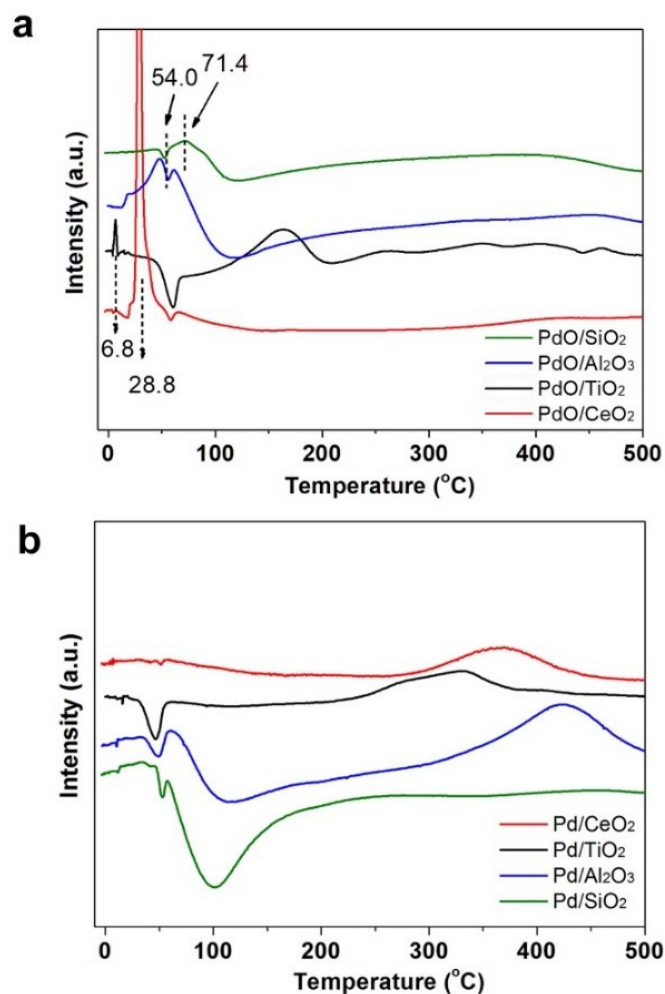


Figure S10. H₂-TPR profiles of (a) PdO/MO_x and (b) Pd/MO_x. Obviously, the PdO peaks appeared for PdO/MO_x, but those peaks were not detected in the Pd/MO_x catalysts.

Note: The PdO/MO_x samples were prepared by oxidation of the fresh Pd/MO_x samples under 20% O₂/N₂ atmosphere at 300 °C for 1 hour and cooling down to room temperature under the same atmosphere. H₂-TPR profiles were collected from ~0 to 500 °C. It is generally accepted that Pd species are relatively easy to be oxidized in air, so the origin of Pd²⁺ species were considered carefully. In comparison with PdO/MO_x (Figure S10a), the Pd/MO_x samples (Figure S10b) do not contain PdO species, and thus, all Pd²⁺ species in Pd/MO_x are derived from the charge transfer from Pd atoms to oxide supports.

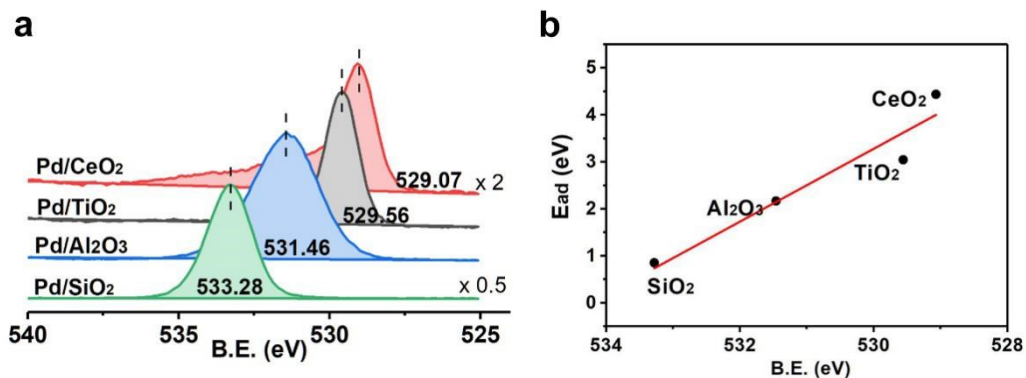


Figure S11. (a) O 1s XPS profiles of Pd/MO_x catalysts. (b) The correlation between adsorption energy (E_{ad}) and binding energy of O_{1s} (B.E.).

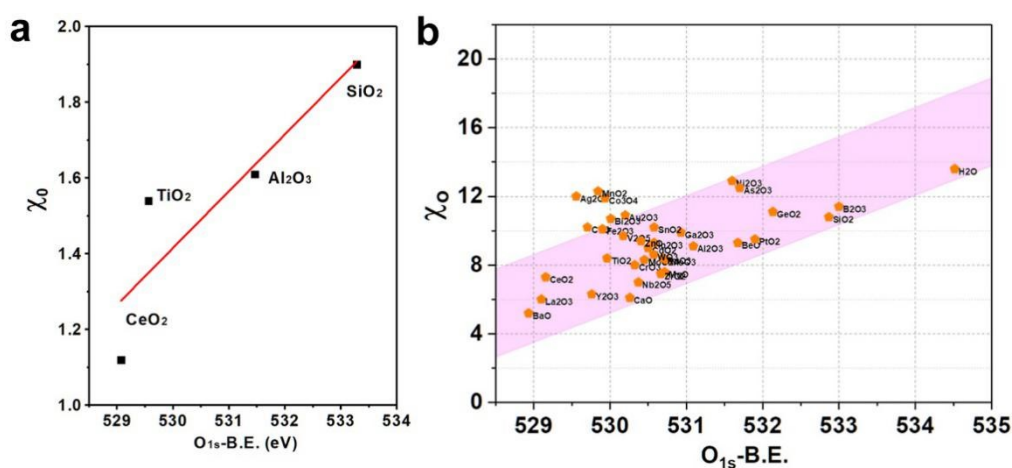


Figure S12. The correlation between binding energy of O_{1s} (BE) and the electronegativity of the central metal ion (χ_0) according to the table of the electronegativity of atoms coined by Roald Hoffmann^{4,5} (a) experimental oxides, the BE of O_{1s} obtained from O 1s XPS profiles of Pd/MO_x catalysts. (b) reference oxides, all data of O_{1s} BE were compiled from the NIST X-ray Photoelectron Spectroscopy Database⁵ by taking the mean value of O_{1s} BE as the input values.

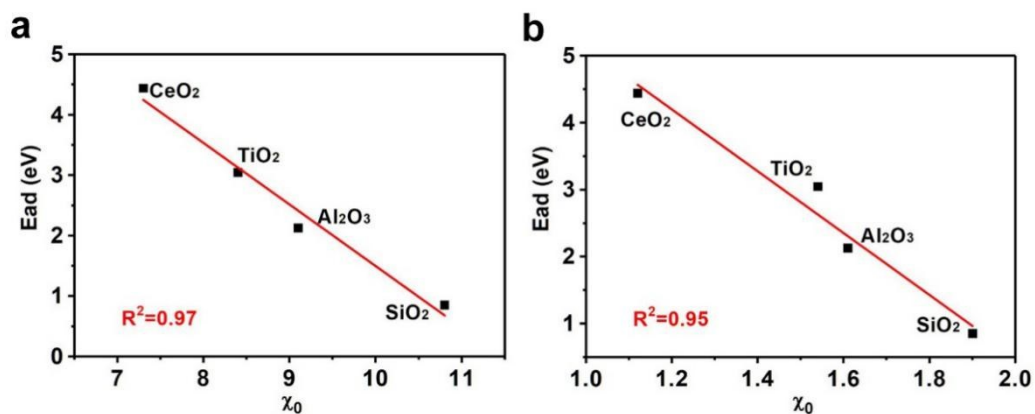


Figure S13. (a) The correlations between adsorption energy (E_{ad}) and the electronegativity of the central metal ion (χ_0) according to the table of the electronegativity of atoms coined by Roald Hoffmann; (b) The correlation between adsorption energy (E_{ad}) and the electronegativity of the central metal ion in Pauling scale (χ_{0-p}).

Note: The electronegativity of the central metal ion (χ_0) was introduced as a key factor, and the calculated E_{ad} was adopted as the descriptor of EMSI based on DFT calculations to better quantitatively describe and understand how to improve the electronic interaction. Surprisingly, the strong correlations among adsorption energy (E_{ad}), binding energy of O_{1s} and the electronegativity of the central metal ion (χ_0) were revealed (Figure S10, S11 and S12)⁴.

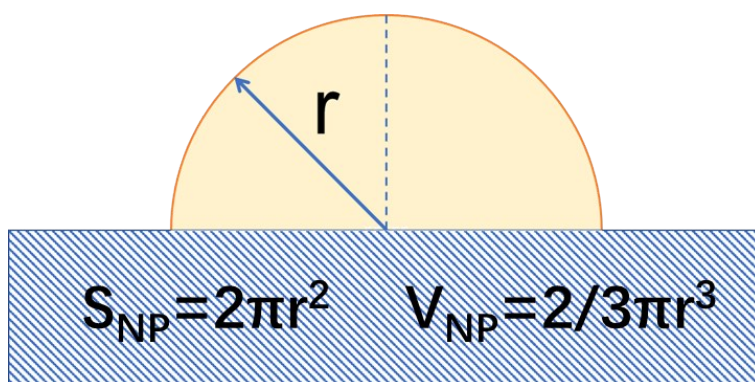


Figure S14. The model of supported Pd nanoparticle, the surface area and volume of nanoparticle in order to calculate metal dispersions for Pd/MO_x.

Note: The number of surface Pd atoms were calculated by calculated metal dispersion (D). The calculated dispersion can be expressed as $D = (S_{NP} \cdot N_s) / (V_{NP} \cdot N_v)^6$, where S_{NP} and V_{NP} represent the surface area and volume of supported Pd nanoparticles, respectively (Figure S14); N_s (15.25 atoms/nm²) and N_v (67.92 atoms/nm³) represent the number of atoms per unit volume and per unit area based on the bulk Pd crystalline structure, respectively. The radii of supported Pd NPs were derived from the histograms of the particle size distributions (Figure S3).

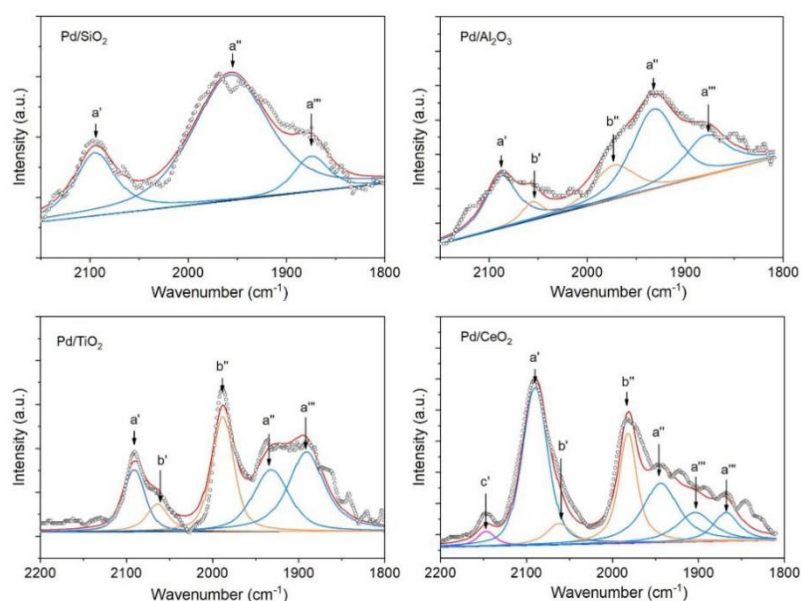


Figure S15. DRIFTS spectra of CO adsorption on Pd/MO_x catalysts at RT and corresponding peak-fitting results in order to calculate the number of surface Pd^{δ-} atoms for Pd/MO_x. The solid lines and scatter data represent the fitted curves and original data, respectively. a', a'' and a''' are the adsorption peaks of linear CO, bridge CO and three-fold CO on Pd⁰ atoms, respectively. b' and b'' are the adsorption peaks of linear CO, bridge CO on Pd^{δ-} atoms, respectively. c' is the adsorption peak of linear CO on Pd²⁺ atoms.

Note: Based on the detailed assignment of the C-O stretching vibration characteristic peaks of adsorbed CO (Figure 4), the deconvolution of the DRIFT spectra was performed for a quantitative analysis. The surface concentration (*C*) of the adsorbate is proportional to a simple function of *R* ($C \propto R$, $R = I_{\text{cat+ads}}/I_{\text{cat}}$), where the intensity $I_{\text{cat+ads}}$ and I_{cat} are measured over the catalyst in the presence of the adsorbate and free of adsorbates⁷. Here, we applied the peak area (*A*) as the intensity. As the baseline data, the I_{cat} for all peaks are constant in this system, i.e., $R = \text{constant} * I_{\text{cat+ads}}$. So, the surface concentration then is proportional to $I_{\text{cat+ads}}$ ($C \propto I_{\text{cat+ads}}$). Considering the adsorption structures, a'' and b'' are bridge CO on surface Pd atoms, which results in two surface Pd atoms occupied by a single CO molecule. Similarly a''' means that a single CO molecule occupies three Pd atoms. The areas of a'', b'' and a''' peaks should be multiplied by a coefficient of 2, 2 and 3, respectively, to measure the number of surface Pd atoms.

So, the ratio of surface Pd^{δ-} atoms (at%Pd^{δ-}) can be expressed as:

$$\text{at\%Pd}^{\delta-} = (A_b + A_{b''} * 2) / (A_{a'} + A_{b'} + A_c + A_{a''} * 2 + A_{b''} * 2 + A_{a'''} * 3 + A_{b'''} * 3)$$

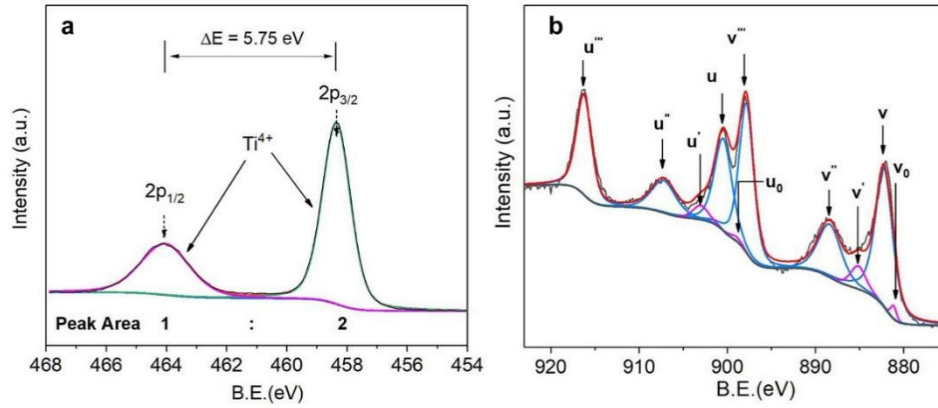


Figure S16. XPS spectra and corresponding peak fitting of (a) Ti2p in Pd/TiO₂ and (b) Ce3d in Pd/CeO₂.

Table S2. Intensities (calculated peak areas), inter-component intensity ratios for peaks relevant to Ce³⁺ and Ce⁴⁺ species derived from the curve-fitting of the XPS Ce3d spectrum (Figure 5.) $I_{\text{tot}}\text{Ce}^{3+}$ and $I_{\text{tot}}\text{Ce}^{4+}$ represent the total intensities of Ce³⁺ and Ce⁴⁺ related peaks, respectively. Relative errors are reported, which referred to the theoretical intensity ratio of 1.5 in Ref.[^{8,9}].

Intensity, $I_{\text{tot}}\text{Ce}^{3+} = 2.53 * 10^4$			
I_v^o	I_v'	I_u^o	I_u'
$0.28 * 10^4$	$1.27 * 10^4$	$0.19 * 10^4$	$0.78 * 10^4$
Inter-component ratio			
I_v^o / I_u^o	I_v' / I_u'	$\sum I_v^i / I_u^i$	
1.49(0.67%)	1.63(8.67%)	1.60(6.67%)	

1.50		1.50		1.50	
Intensity, $I_{\text{tot}}\text{Ce}^{4+} = 34.6 \times 10^4$					
I_v	I_v''	I_v'''	I_u	I_u''	I_u'''
8.87×10^4	4.32×10^4	7.66×10^4	5.70×10^4	2.81×10^4	5.28×10^4
Inter-component ratio					
I_v/I_u	I_v''/I_u''	I_v'''/I_u'''	$\Sigma I_v^j/I_u^i$		
1.55(3.33%)	1.54(2.67%)	1.45(3.33%)	1.51(0.67%)		
1.50	1.50	1.50	1.50		

Note: Quantitative analysis of XPS spectra (Ti2p and Ce3d) was performed applying peak fitting (Figure S16 and Table S2). The formation of any surface oxygen defects triggers the reduction of Ti^{4+} to Ti^{3+} and Ce^{4+} to Ce^{3+} . And the relative concentration of Ti^{3+} and Ce^{3+} can be calculated by the area of XPS deconvolution peaks. For Pd/TiO₂, the peak area ratio (1:2) and doublet separation energy for Ti2p XPS peaks (5.75 eV) were determined to agree well with the theoretical values, 1:2 and 5.76 eV, respectively (Figure S16a)¹⁰, indicating that the only Ti^{4+} existed in the Pd/TiO₂ catalysts. Because of the spectral complexity of Ce3d XPS spectra for cerium oxides, inter-component peak intensity ratios are concerned in the peak fitting to evaluate the accurate Ce^{3+} percentage (Figure S16b). Theory suggests that each of the two spin-orbit split components of the Ce3d subshell contains 10 electrons, which distributed into n_{j+} (6-fold degenerate electrons) and n_{j-} (4-fold degenerate electrons) resulting in their intensity ratio of 6/4 (= 1.5) [8]. That is, each of the XPS Ce3d_{5/2} peaks is 1.5 times as intense as the corresponding Ce3d_{3/2} peak, and similarly the ratio of the summation of the intensities of Ce3d_{5/2} peaks to Ce3d_{3/2} peaks is 1.5. As shown in Table S2, the inter-component peak intensity ratios were calculated, and the relative errors are also shown in parenthesis, all of which are within 10%.

Therefore, the at%[Ce³⁺] percentage is:

$$\text{at}\%[\text{Ce}^{3+}] = \frac{I_{\text{tot}}\text{Ce}^{3+}}{I_{\text{tot}}\text{Ce}^{3+} + I_{\text{tot}}\text{Ce}^{4+}} \diamond 100 = \frac{2.53}{2.53 + 36.4} \diamond 100 = 6.5$$

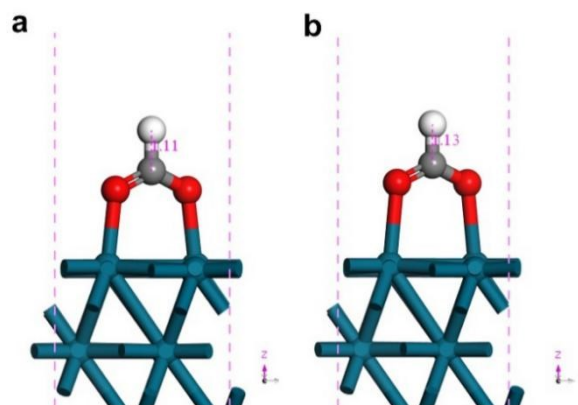


Figure S17. The interactions of bridging formate on the surface of Pd(111): (a) neutral HCOO*-Pd(111), (b) electron-rich HCOO*-Pd(111) with extra 2e-.

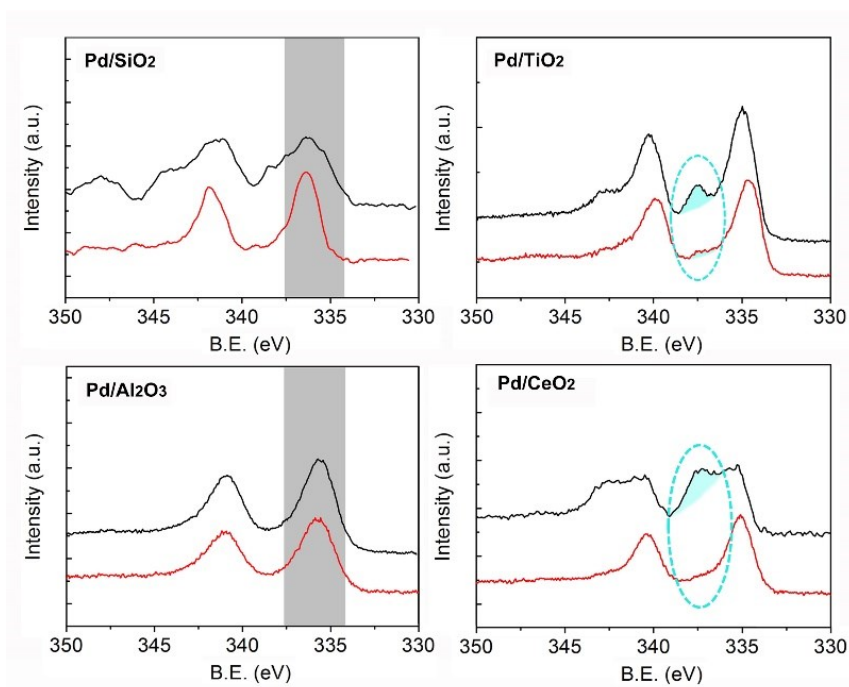


Figure S18. Pd 3d XPS profiles for Pd/MO_x catalysts before (black line) and after (red line) usage.

2. References

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3. Appendix

Calculated Cartesian coordinates of all calculated structures:

SiO2_001_Pd4

1.0

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5.0219998360	8.6983588715	0.0000000000
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O Si Pd

80 40 4

Cartesian

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4.949224472	3.339672565	15.426607132
2.406993866	0.313488573	3.652678967
2.529806137	0.448217005	9.210737228
2.525983334	1.870824337	15.438261986
7.696952820	1.748691797	1.049994946
8.917885780	2.559456587	7.349452972
8.876910210	2.358324051	12.835175514
10.160825729	3.259196758	1.040679097
9.948991776	0.263559848	5.522803307
10.080779076	0.165233165	11.005273819
9.989403725	3.311924934	15.409619331
7.428993702	0.313488573	3.652678967
7.551805973	0.448217005	9.210737228
7.536664009	1.874982119	15.426973343
5.185953140	6.097871304	1.049994946

6.406886101	6.908636093	7.349452972
6.366055489	6.707355499	12.834372520
7.649827003	7.608376026	1.040679097
7.437992573	4.612739086	5.522803307
7.568357944	4.512577057	11.004798889
7.458827972	7.690661430	15.419118881
4.917994022	4.662667751	3.652678967
5.040805817	4.797396660	9.210737228
4.987747669	6.228371620	15.417949677
10.207952499	6.097871304	1.049994946
11.428885460	6.908636093	7.349452972
11.373637199	6.694020748	12.848730087
12.671826363	7.608376026	1.040679097
12.459992409	4.612739086	5.522803307
12.591406822	4.510254383	11.011667252
12.463893890	7.693966866	15.429968834
9.939992905	4.662667751	3.652678967
10.062806129	4.797396660	9.210737228
10.042330742	6.221691608	15.454445839
9.002478600	4.703731537	18.426149368
6.696613312	3.372317314	18.550107956
6.726112843	6.027577877	18.404924393
7.513707161	4.808059692	20.517080307

Al2O3_110_Pd4

1.0

8.4130001068	0.0000000000	0.0000000000
-0.0830810040	8.0675720611	0.0000000000

-0.0000083582

-0.0000084447

31.3845996857

O Al Pd

54 36 4

Cartesian

3.586626768	7.386300564	3.084462643
3.475185156	7.316686630	8.594333649
7.764607906	5.055858135	1.363234043
7.858466625	5.042865276	7.026563168
3.634816647	6.819906712	31.046518326
3.509095430	6.979537487	5.639062405
3.609799862	6.915192604	11.127192497
7.847551346	5.230297565	4.375734806
7.739969730	5.141276360	10.135182381
7.791420937	0.997646034	4.283862591
7.814534187	0.918891370	9.714748383
3.707256317	3.454153299	31.380043030
3.643911123	3.221771717	5.814198017
3.675211668	3.180218220	11.615566254
7.826588631	1.263450623	1.678285122
7.779827118	1.263678312	7.152942657
3.569523096	2.988148928	2.545871735
3.651081562	2.987327576	8.448247910
5.184673309	0.989071906	4.284983158
5.133564949	0.914365470	9.714937210
0.795745194	3.459688187	31.368173599
0.833327234	3.218340635	5.813548088
0.859009624	3.193271637	11.635623932
5.133209705	1.249322653	1.679446340

5.150142193	1.257222652	7.151037216
0.992791414	2.983229160	2.514618874
0.852518022	2.985610485	8.452035904
0.817676127	7.383386612	3.089047909
0.904438376	7.311212540	8.598106384
5.164542198	5.049097061	1.354506016
5.019817352	5.038904667	7.021933556
0.870002568	6.810054779	31.035957336
0.920474291	6.976758003	5.644272327
0.785411656	6.893731117	11.136332512
5.003284454	5.221014500	4.367151260
5.080533028	5.140638351	10.147761345
6.417251587	7.201321602	2.986731052
6.403200626	7.104136467	8.556019783
6.445204735	7.382108212	0.063161507
6.416224003	7.264731407	5.644582748
6.402105331	7.448391438	11.192539215
2.273103237	1.209564328	4.184901237
2.261308908	1.084185481	9.899195671
2.279778957	1.026760340	1.215170860
2.261082172	0.979569912	7.164696217
2.258147240	5.143971443	1.436096549
2.228295088	5.235914230	7.052637577
2.222683907	5.145849705	4.394016743
2.216467381	5.024976730	9.997702599
6.465786457	2.743383408	0.164634198
6.450847626	3.024502039	5.761845112
6.474888802	2.805603027	11.386395454

6.469654560	3.118436098	2.831421137
6.449110985	3.081588984	8.448310852
0.755135894	5.255458832	0.351209372
0.741503894	5.061497688	5.692746162
0.734768033	4.954124451	11.196961403
4.972984314	7.064179897	4.276961327
5.022888660	6.993644714	9.951696396
5.085659027	3.198811769	1.342102766
4.998538971	3.207844496	7.145146370
0.797528028	1.136452198	2.954630852
0.810383499	1.141574502	8.497085571
7.842706203	3.205554008	1.344999552
7.907613754	3.214948654	7.163873672
3.760489702	1.144398689	2.973822594
3.716668129	1.146310925	8.500176430
3.757345676	5.257863045	0.350562841
3.708260298	5.070638180	5.705045700
3.685755730	4.916461468	11.244527817
7.854796410	7.071671963	4.290629387
7.779335022	7.005092621	9.954621315
6.479490757	0.963710845	0.412776530
6.478236198	1.127420545	5.713233471
6.470347881	1.160296082	10.993563652
6.485505581	1.180179238	3.001732826
6.466584206	1.047617078	8.416771889
6.461200714	6.267337322	1.385357022
6.435180664	6.089598179	7.071716309
2.247492313	7.603466988	0.264113963

2.201456070	7.117156506	7.123779774
2.207314491	7.121204376	4.245165825
2.195491314	7.181792736	10.009132385
2.246194601	2.768836260	0.762008667
2.241756678	2.144264698	5.745889187
2.273975611	2.375746012	11.102028847
2.275768995	4.177021503	2.956074476
2.250195265	4.086599350	8.436374664
6.433866501	4.158025265	4.309102058
6.411046505	3.955878735	10.088169098
3.465080976	6.445006847	13.261552811
2.217772722	5.780623913	15.410660744
1.840300083	4.239013672	13.386887550
0.856233895	6.662642956	13.380261421

TiO2_110_Pd4

1.0

8.8769998550	0.0000000000	0.0000000000
-0.0000034605	12.9937896729	0.0000000000
-0.0000077351	-0.0000077351	29.0448207855

O Ti Pd

72 36 4

Cartesian

0.000011795	-0.000007734	29.041769028
8.876978874	6.496884346	29.041770935
2.958992243	-0.000007734	29.041753769
2.958998442	6.496886253	29.041782379
5.917977810	-0.000007734	29.041769028

5.917997837	6.496884346	29.041769028
0.005273876	3.281103611	3.196702003
0.005771920	9.756334305	3.221056223
2.988711596	3.270154953	3.218035460
2.959222317	9.759776115	3.219990015
5.892809868	3.270953178	3.218322992
5.927707672	9.760282516	3.219740391
8.841572762	0.059989296	6.491363525
8.871250153	6.608335018	6.518944263
2.915585995	12.955182076	6.484014988
2.953367472	6.652069569	6.511393070
5.914141178	12.951030731	6.468560219
5.907949448	6.669470310	6.489243031
-0.000000684	-0.000000684	2.569206238
-0.000002414	6.496893406	2.569194317
2.958996058	-0.000000684	2.569182873
2.958998919	6.496894360	2.569206238
5.917990685	-0.000000684	2.569208860
5.917998791	6.496894360	2.569185734
0.007151583	3.221618891	5.759959698
0.009202179	9.755033493	5.855810642
2.918358326	3.256949186	5.818537235
2.970867395	9.752405167	5.845304966
5.977407455	3.253937006	5.818212032
5.924060822	9.747634888	5.845793247
0.054008797	0.032410596	9.070581436
0.048728794	6.527718544	9.095097542
3.039794922	0.121275119	9.068589211

2.988826752	6.386767387	9.087148666
5.932924271	0.112951294	9.052605629
5.944863796	6.383657932	9.062378883
1.479498863	4.456752300	1.050080657
1.479498148	10.953649521	1.050077796
4.438499928	4.456756115	1.050080657
4.438505650	10.953674316	1.050127029
7.397498131	4.456749439	1.050071955
7.397496700	10.953687668	1.050179362
1.481321096	1.313689113	4.499327660
1.477838755	7.844652176	4.508900642
4.438537598	1.311802506	4.553513527
4.436897755	7.846281528	4.491291046
7.393132687	1.314663768	4.491679668
7.401738644	7.844736576	4.504354954
1.430698276	4.534568310	7.861828804
1.509473681	10.998265266	8.026260376
4.462301731	4.561931610	7.607636929
4.465259075	10.978797913	8.026571274
7.493919849	4.530585766	7.867994785
7.414839745	10.995962143	8.022087097
1.479498506	2.040136576	1.050074816
1.479499459	8.537034988	1.050077796
4.438501835	2.040138960	1.050074816
4.438497066	8.537038803	1.050074816
7.397498608	2.040098667	1.050176501
7.397496700	8.537038803	1.050068974
1.480198026	5.220508575	4.553951740

1.471672654	11.671679497	4.530718803
4.440903664	5.222813129	4.579154491
4.437708855	11.671064377	4.514177799
7.400137424	5.221979141	4.553255081
7.405288696	11.670433998	4.527117252
1.511591434	2.092144728	8.008754730
1.508498788	8.583774567	8.074910164
4.460468292	2.040304899	7.756601810
4.458480358	8.565526009	8.080768585
7.421572208	2.090573788	8.016591072
7.418392181	8.579376221	8.079536438
-0.000001229	3.248436451	1.366698265
-0.000002959	9.745351791	1.366695404
2.958998442	3.248441696	1.366695404
2.958999395	9.745344162	1.366695404
5.918002129	3.248434067	1.366695404
5.918002129	9.745353699	1.366698265
0.015036535	12.960426331	4.510135174
0.007154610	6.460458755	4.483770847
2.978718996	12.970965385	4.504064560
2.960331917	6.471531868	4.481348515
5.928092957	12.963075638	4.488107681
5.929165840	6.464658737	4.464970589
0.025692483	3.228867769	7.581778526
0.020959258	9.811133385	7.708048344
2.974247217	3.393791199	7.930763721
2.976663589	9.789230347	7.695776939
5.949520588	3.377269745	7.924521923

5.945818901	9.770830154	7.697246552
1.479500055	-0.000000291	1.092294455
1.479494691	6.496891499	1.092276931
4.438509941	-0.000000291	1.092294455
4.438498497	6.496891499	1.092279911
7.397504330	-0.000000291	1.092271090
7.397509098	6.496894360	1.092303038
1.539546728	3.193821669	4.508502483
1.488355756	9.720085144	4.532830715
4.448624134	3.213570833	4.714166164
4.448429585	9.713326454	4.523495674
7.357898712	3.188255072	4.503500938
7.409200191	9.715070724	4.534050465
1.608223557	0.075383946	7.969324112
1.553978801	6.563497543	7.933455944
4.563691139	0.012053546	7.858645439
4.546891212	6.576210976	7.790886879
7.520062447	0.068807788	7.959501266
7.547746181	6.549422741	7.938919544
5.754799843	4.650426388	10.313879967
4.462348461	2.290805101	9.895799637
3.193104267	4.656729698	10.330072403
4.481328011	3.309069157	12.209053993

CeO2_100_Pd4

1.0

11.6406602859	0.0000000000	0.0000000000
-0.0000062002	11.6406602859	0.0000000000

-0.0000150368

-0.0000150368

28.2311897278

O Pd Ce

54 4 27

Cartesian

11.640538216	0.000934399	2.834597826
0.000335165	3.883974791	2.834857702
0.000505381	7.761735916	2.835230112
3.879990101	0.000582851	2.835746765
3.880111456	3.879341841	2.835187912
3.880350351	7.762204647	2.834075451
7.761276245	0.001398863	2.832048416
7.760268688	3.881731749	2.832471848
7.759448528	7.761775494	2.832017422
0.055436693	0.103070334	7.926372051
0.061633270	3.990561724	7.929443836
0.068911277	7.869951248	7.934516907
3.915889263	11.374185562	8.026640892
4.172091484	3.973444462	8.056794167
3.687893629	7.743387222	8.033198357
7.759673595	11.612403870	7.949361324
7.514908791	3.741704464	8.002122879
8.083173752	7.815333843	7.978286743
1.918997288	1.942014098	5.596918106
1.920999765	5.833542824	5.614638805
1.912726998	9.695086479	5.591486454
5.853621006	1.925180554	5.557594776
5.844498158	5.834518433	5.578113079
5.852842808	9.688963890	5.609232903

9.693275452	1.944695115	5.623859406
9.699575424	5.823544502	5.632709503
9.702073097	9.710083008	5.607067108
0.001403722	0.001060323	0.264396399
0.001994165	3.881085634	0.264433086
0.001968816	7.762451172	0.265141666
3.877348900	0.001283823	0.265745848
3.878720284	3.881487370	0.264147967
3.878720760	7.761273861	0.264060438
7.760569572	0.000853121	0.262310088
7.760439873	3.881053925	0.261875361
7.760780811	7.761268139	0.262151986
11.624558449	11.640566826	5.361577511
11.624709129	3.881902218	5.360574722
11.621519089	7.757256508	5.361111164
3.883294106	11.595368385	5.365554810
3.871594667	3.934543848	5.388278008
3.894294262	7.747628689	5.397442341
7.773256302	11.633890152	5.371449947
7.781994343	3.909987926	5.403401375
7.752559662	7.743330479	5.401264668
1.939564466	1.941963434	2.561661482
1.940604210	5.821379662	2.561655760
1.939992309	9.701933861	2.562861204
5.820340157	1.941529274	2.557440996
5.819457054	5.820868492	2.555202246
5.820469856	9.702482224	2.557870150
9.701293945	1.943530321	2.556173325

9.700654984	5.821850300	2.557443619
9.700472832	9.702004433	2.559784174
7.258338451	6.974469185	9.670647621
5.899278164	4.740690231	9.062799454
4.664913654	7.003130913	9.689100266
5.970375061	5.697406769	11.487663269
1.940760136	0.002769458	4.098869801
1.940165639	3.882886887	4.098313332
1.939546466	7.763078213	4.098381042
5.820614338	0.002243302	4.095010281
5.820399284	3.882557392	4.094654560
5.820059299	7.762651443	4.094202995
9.700977325	0.003473720	4.094982147
9.700627327	3.883723736	4.094614983
9.700263023	7.764190197	4.095287323
0.001347410	1.941884875	1.329203486
0.001188194	5.821822166	1.330420136
0.000927704	9.702592850	1.330575466
3.878973722	1.942084074	1.330310225
3.878939152	5.821653366	1.328689575
3.878876209	9.701571465	1.330479503
7.760471821	1.941922188	1.327139735
7.760471821	5.821761608	1.326826453
7.760714531	9.702084541	1.327475786
0.014505394	1.970054746	6.831778526
0.013107588	5.868490219	6.861155987
0.012918116	9.731803894	6.860182285
3.897530794	1.858934045	6.923213959

3.903353453	5.799028873	6.833986282
3.870040417	9.704613686	6.846562862
7.740246296	1.886386275	6.828636169
7.748097420	5.812539577	6.814670086
7.782177448	9.749215126	6.831515789

TiO2_110_Pd10

1.0

11.8360004425	0.0000000000	0.0000000000
0.0000000000	12.9938001633	0.0000000000
0.0000000000	0.0000000000	19.0447998047

O Ti Pd

96 48 10

Cartesian

2.894967318	3.248450041	19.042133331
2.894967318	9.745349884	19.042133331
5.853967667	3.248450041	19.042133331
5.853967667	9.745349884	19.042133331
8.812967300	3.248450041	19.042133331
8.812967300	9.745349884	19.042133331
11.771967888	3.248450041	19.042133331
11.771967888	9.745349884	19.042133331
2.893310308	0.0000000000	3.217809439
2.893310308	6.496900082	3.217809439
5.852310658	0.0000000000	3.217809439
5.852428436	6.496900082	3.217809439
8.811309814	0.0000000000	3.217809439
8.811309814	6.496900082	3.217809439

11.770310402	0.000000000	3.217809439
11.770310402	6.496900082	3.217809439
2.900317192	3.146367550	6.512630939
2.933832169	9.726785660	6.517506599
5.873107433	3.086225271	6.473622799
5.885790825	9.718057632	6.523687840
8.809265137	3.104153633	6.508354187
8.813790321	9.717393875	6.556962490
11.792618752	3.155513763	6.522076607
11.776374817	9.730433464	6.533092499
2.899346590	3.248450041	2.571429014
2.899346590	9.745349884	2.571429014
5.858346939	3.248450041	2.571429014
5.858346939	9.745349884	2.571429014
8.817346573	3.248450041	2.571429014
8.817346573	9.745349884	2.571429014
11.776347160	3.248450041	2.571429014
11.776347160	9.745349884	2.571429014
2.888254404	0.007934487	5.865884781
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144 72 22

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