

Boosting photocatalytic H₂O₂ production by coupling of sulfuric acid and 5-sulfosalicylic acid incorporated polyaniline with g-C₃N₄

Hao-Dong Yang^a, Jun-Hao Huang^a, Kengo Shibata^b, Daling Lu^c, Kazuhiko Maeda^b, Chechia Hu^{a*}

^a Department of Chemical Engineering, R&D center for Membrane Technology and Luh Hwa
Research Center for Circular Economy, Chung Yuan Christian University, Chungli District, Taoyuan
City, 32023 Taiwan

^b Department of Chemistry, School of Science, Tokyo Institute of Technology, 2-12-1-NE-2,
Ookayama, Meguro-ku, Tokyo 152-8550, Japan

^c Suzukakedai Materials Analysis Division, Technical Department, Tokyo Institute of Technology,
4259 Nagatsuta-cho, Midori-ku, Yokohama 226-8503, Japan

* To whom correspondence should be addressed. E-mail: chechiah@cyu.edu.tw

Tel: 886-3-2654152 Fax: 886-3-2654199

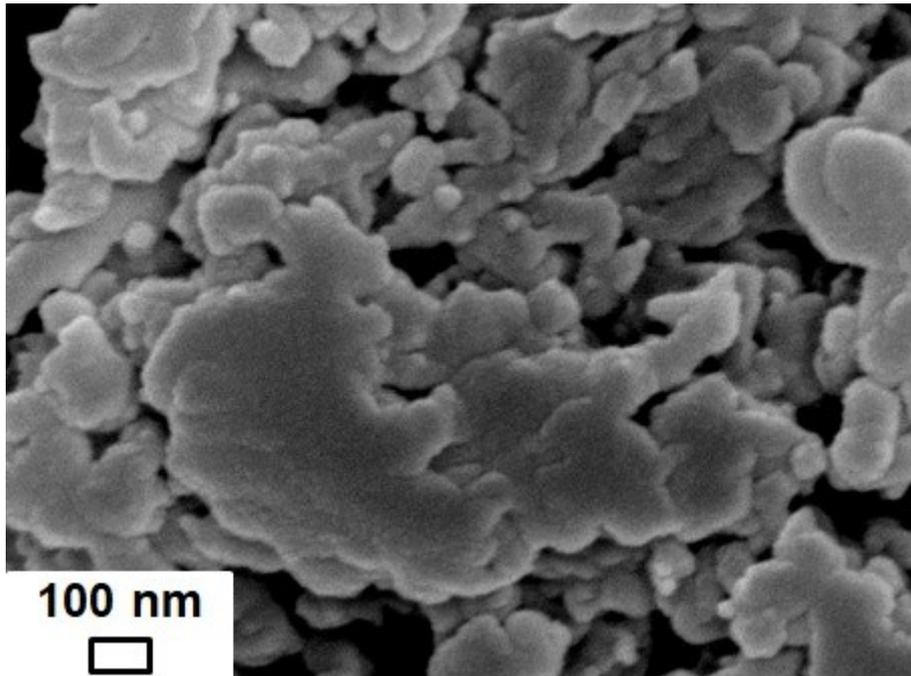


Figure S1 SEM image of undoped PANI.

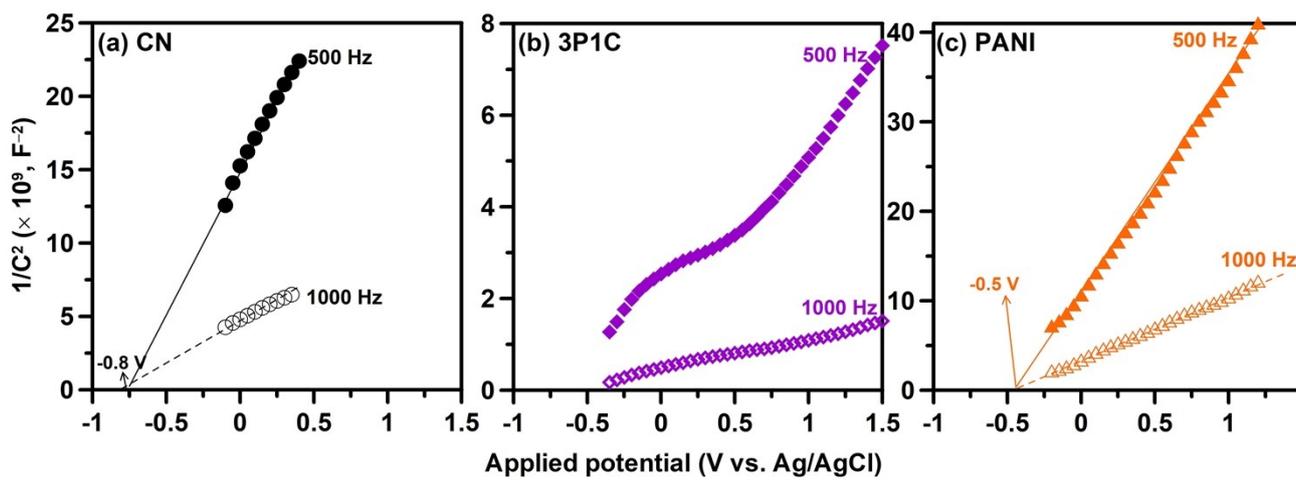


Figure S2 Mott-Schottky plot of (a) CN, (b) 3P1C, and (c) PANI samples measured at 500 and 1000 Hz in an aqueous Na_2SO_4 solution (0.5 M, pH = 6.6).

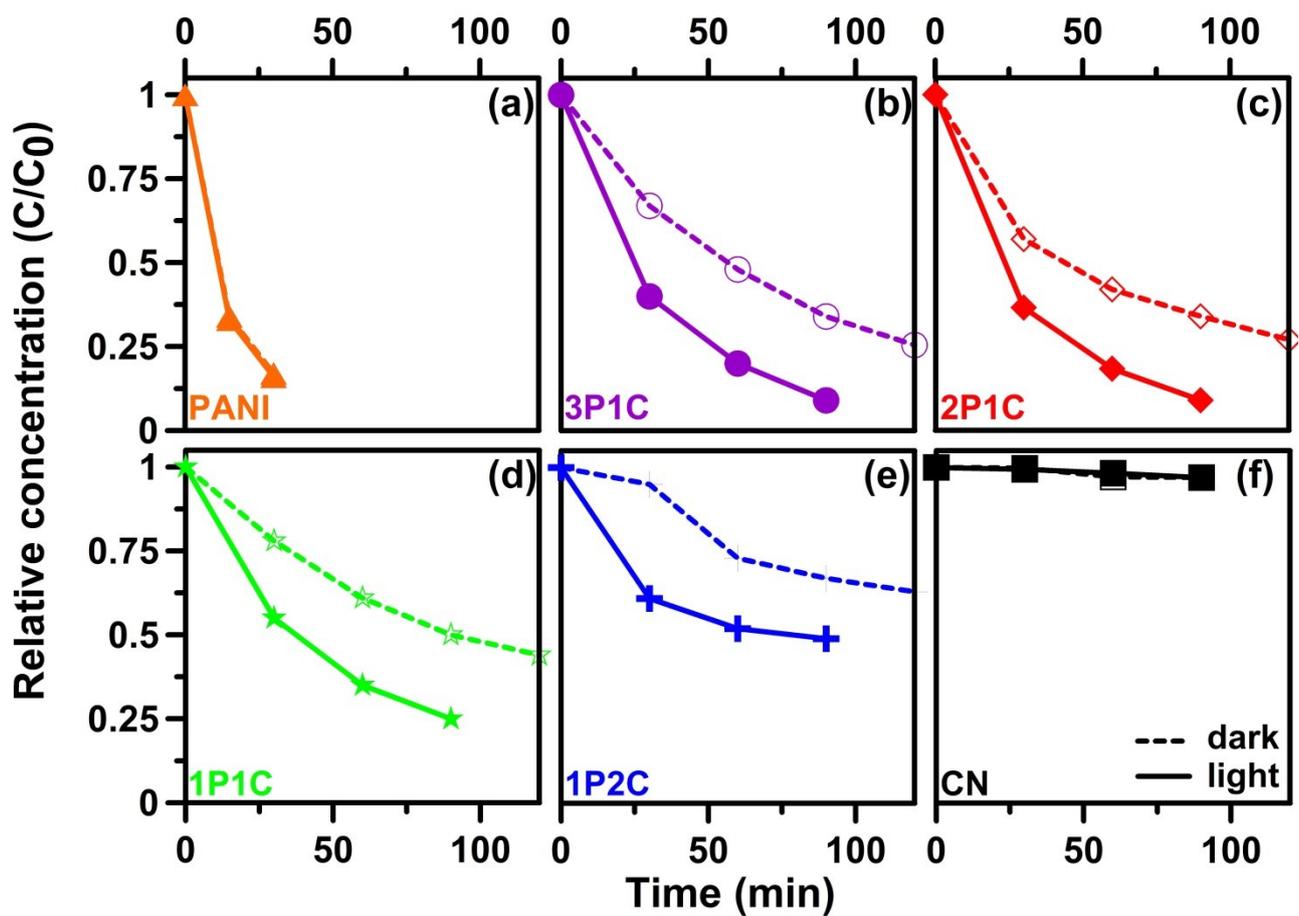


Figure S3 Photocatalytic decoloration of MB aqueous solution (40 ppm, 250 mL) using CN, PANI, and *mPnC* samples under visible-light irradiation or dark condition.

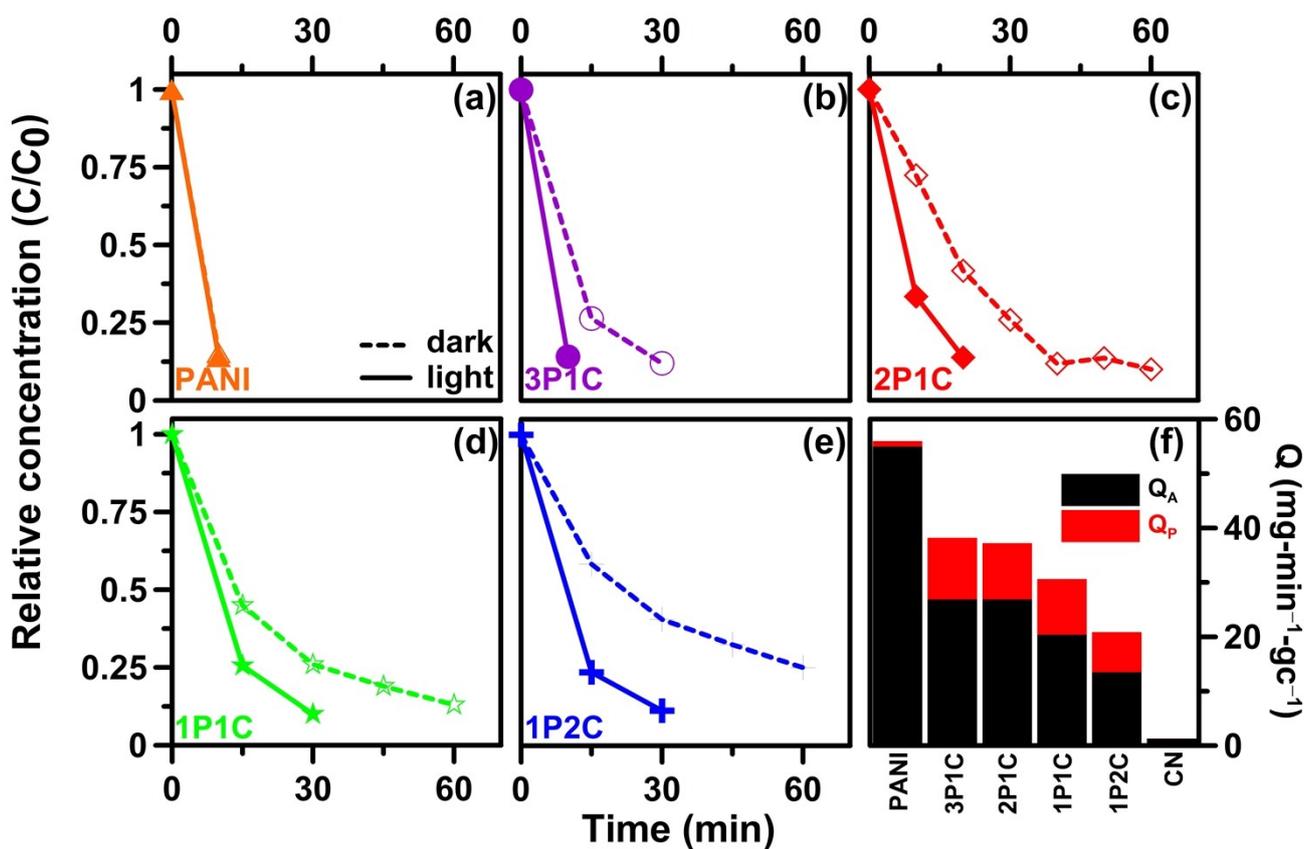


Figure S4 (a-e) Photocatalytic decoloration of MB aqueous solution (20 ppm, 250 mL) using CN, PANI, and $mPnC$ samples under visible-light irradiation or dark condition. (f) Photocatalytic (Q_P) and adsorption capacities (Q_A) per unit time per unit catalyst for these samples.

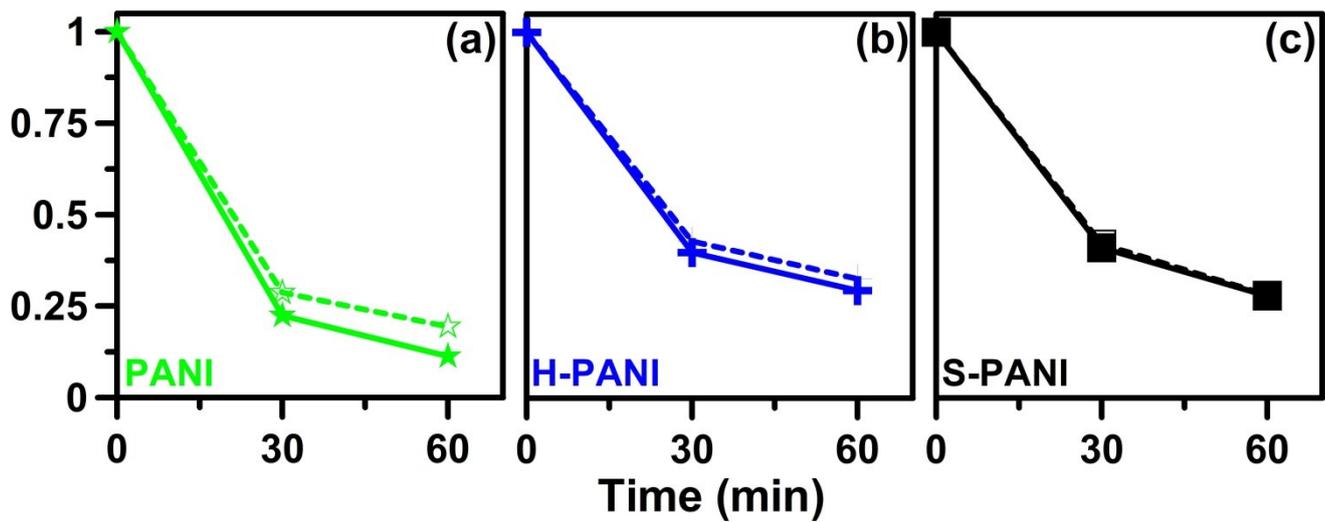


Figure S5 Photocatalytic decoloration of MB aqueous solution (80 ppm, 250 mL) using PANI, and H-PANI (H_2SO_4 mono-doped PANI), S-PANI (SSA mono-doped PANI) samples under visible-light irradiation or dark condition.

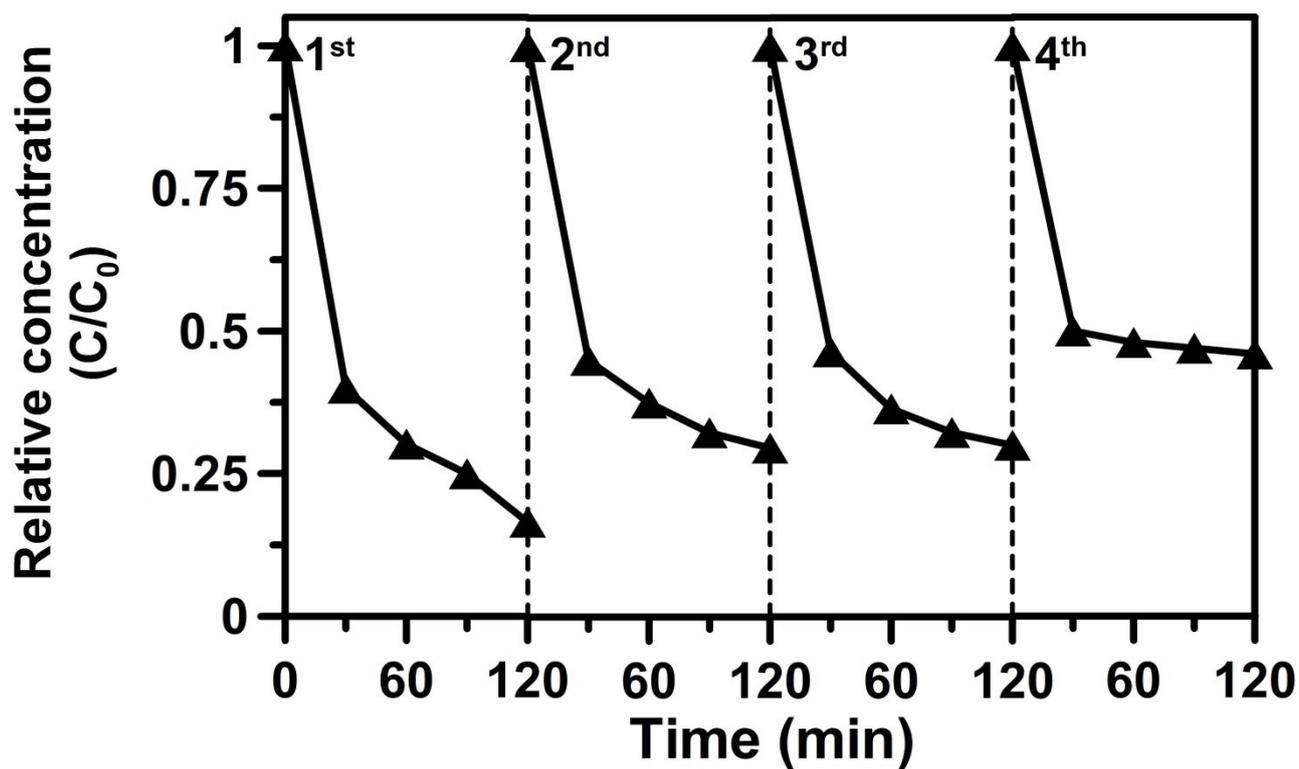


Figure S6 Cycle test for photocatalytic decoloration of MB aqueous solution (40 ppm, 250 mL) using 3P1C sample under visible-light irradiation.

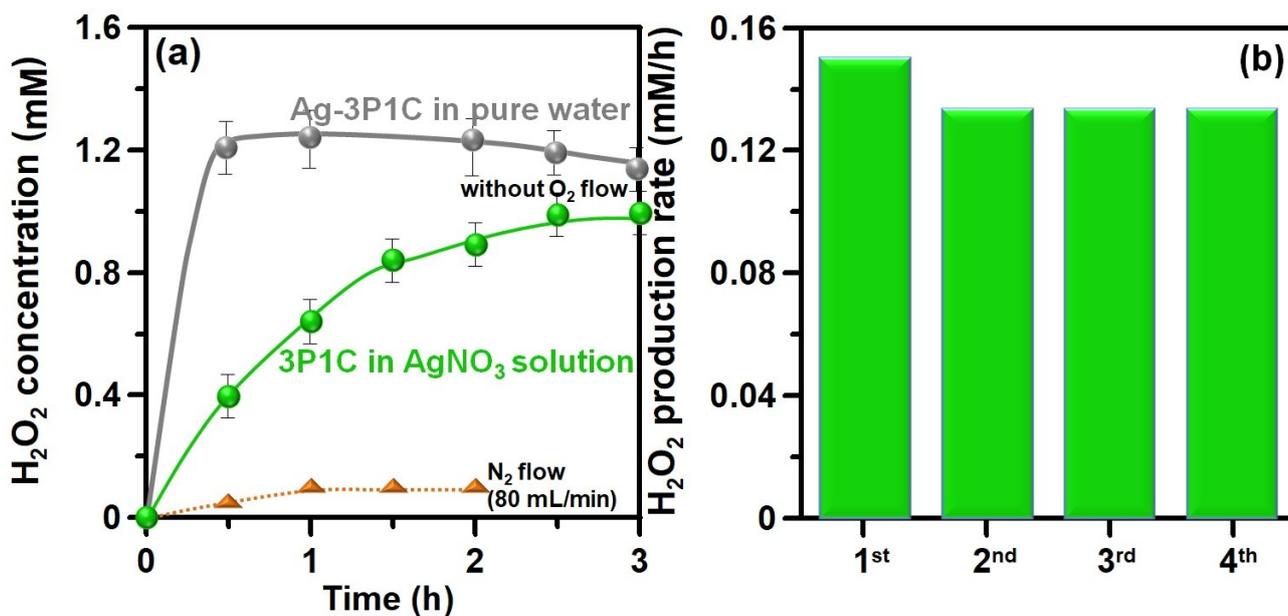


Figure S7 Time-dependent change in H₂O₂ concentration during photocatalytic reaction using 3P1C or Ag-deposited 3P1C sample in the solution containing AgNO₃ (400 ppm) as sacrificial reagent or pure water without O₂ flow, and 3P1C sample in the pure water with N₂ flow (80 mL/min); (b) H₂O₂ production rate in the repeated reaction in the presence of IPA (20 vol%) with O₂ flow rate of 80 mL·min⁻¹ using 3P1C sample (0.25 g) under simulated solar irradiation for 3 h.

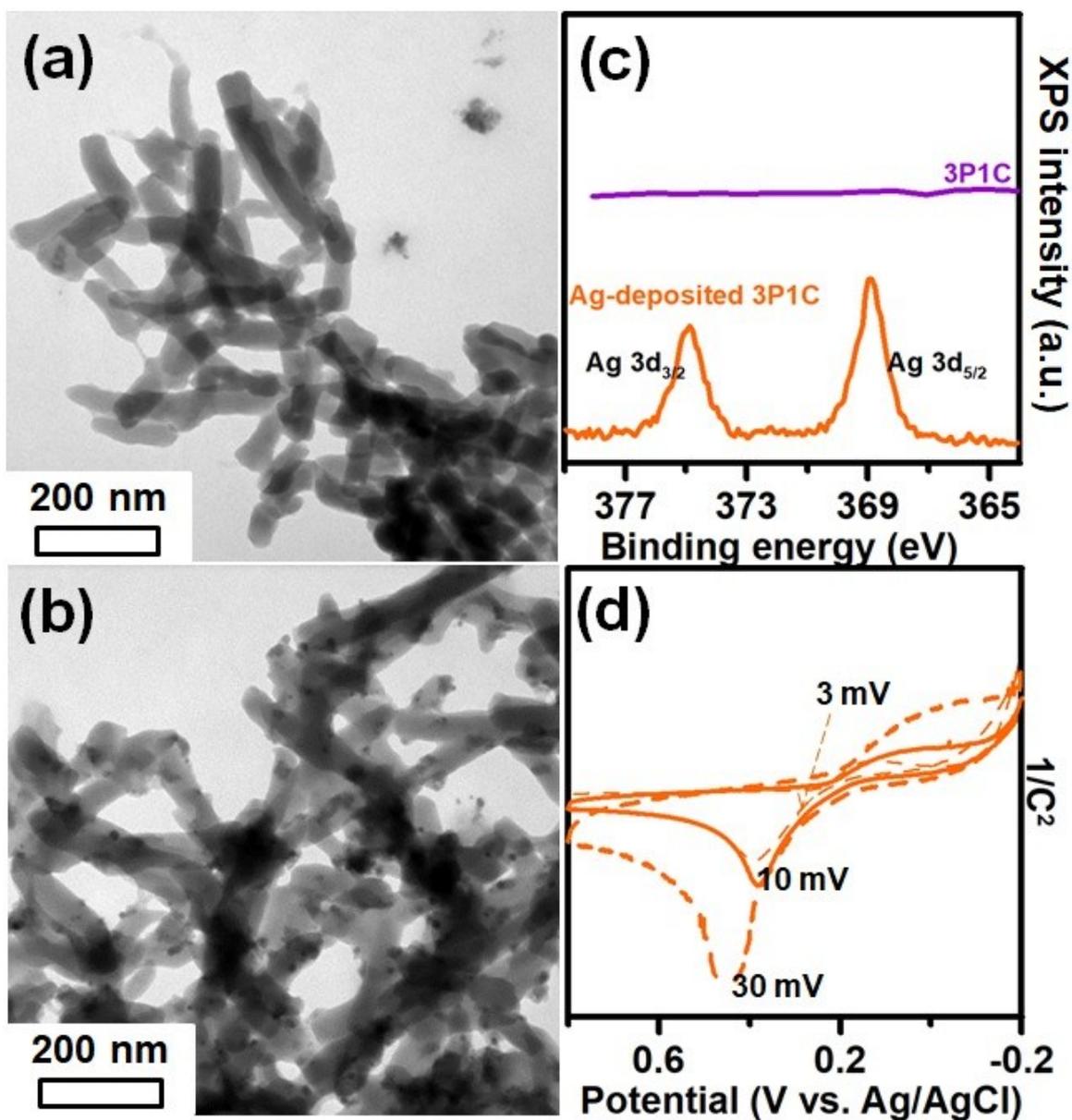


Figure S8 TEM images of (a) 3P1C and (b) Ag-deposited 3P1C samples. (c) High resolution XPS spectra of Ag 3d level, and cyclic voltammetry (CV) of the Ag-deposited 3P1C sample in 0.1 M of H_2SO_4 aqueous solution (pH = 0.8) at a scan rate of 3, 10, and 30 mV s^{-1} .

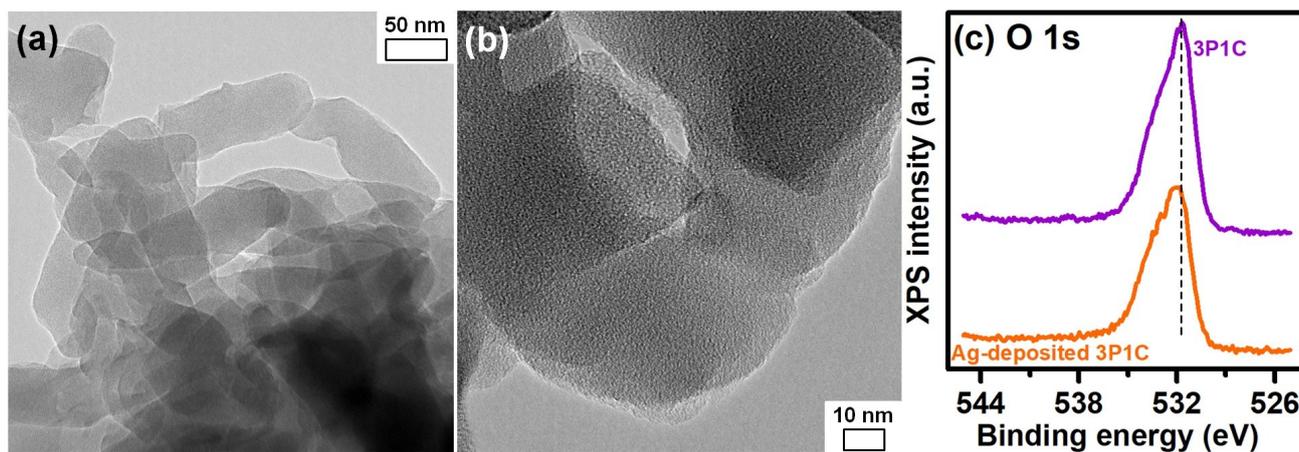


Figure S9 (a) TEM and (b) HRTEM images of 3P1C after the photocatalytic H_2O_2 production with the addition of IPA. (c) XPS spectra of O 1s level for 3P1C sample before (3P1C) and after (Ag-deposited 3P1C) the photocatalytic H_2O_2 production in a solution containing AgNO_3 as a sacrificial agent.

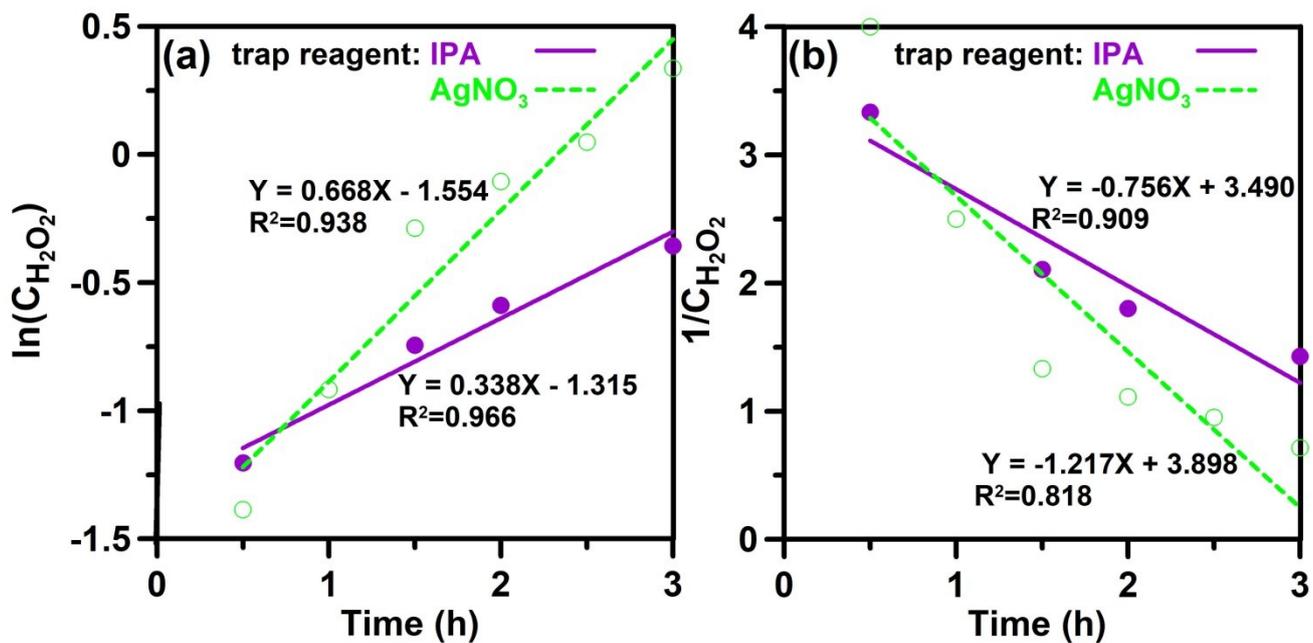


Figure S10 (a) Pseudo-first-order and (b) second-order reaction kinetic plot of the photocatalytic H_2O_2 production using 3P1C in the presence of IPA and AgNO_3 as trap reagents.