

Cross-linked polyelectrolyte microspheres: preparation and new insights into electro-surface properties

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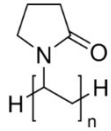
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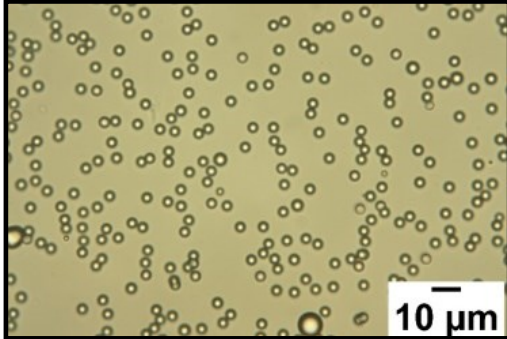
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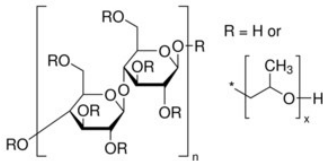
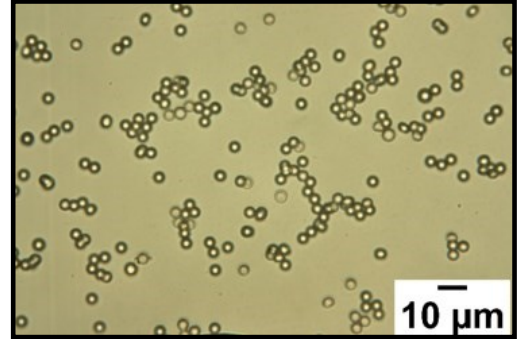


Polyvinylpyrrolidone (PVP). Mw ~29,000(Sigma Aldrich)

Before sulfonation. $[\text{SO}_3^-]=0 \text{ mol/m}^2$

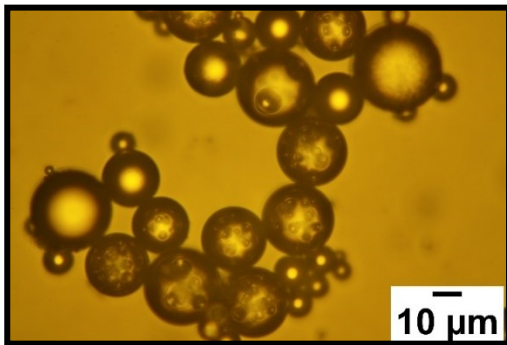


After sulfonation. $[\text{SO}_3^-]=0.85 \cdot 10^{-4} \text{ mol/m}^2$

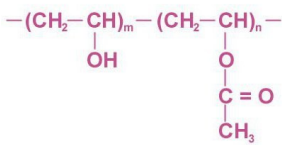
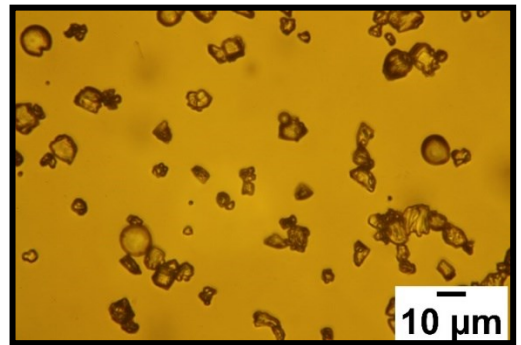


Hydroxypropyl cellulose. Mw ~100,000 (Acros Organics)

Before sulfonation. $[\text{SO}_3^-]=0 \text{ mol/m}^2$

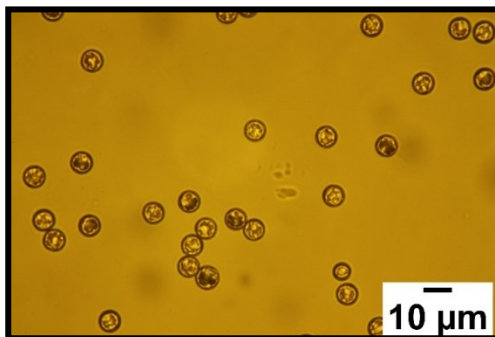


After sulfonation. $[\text{SO}_3^-]=1.64 \cdot 10^{-4} \text{ mol/m}^2$



PVA, Poly(vinyl alcohol). Mw ~12 0 000 (Sigma Aldrich)

Before sulfonation. $[\text{SO}_3^-]=0 \text{ mol/m}^2$



After sulfonation. $[-\text{H}_2\text{C}-\text{CO}-\text{CH}_2-]+[\text{SO}_3^-]$

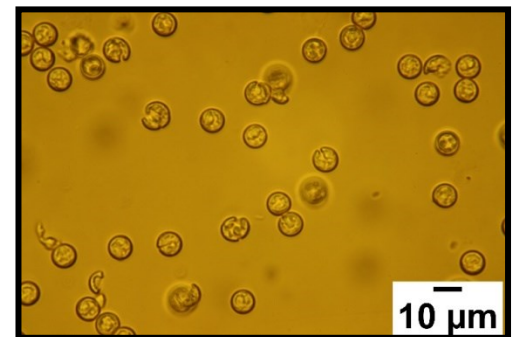


Fig. S1. Optical microscopy of PSt and sulfonated PSt-SO₃ microspheres.

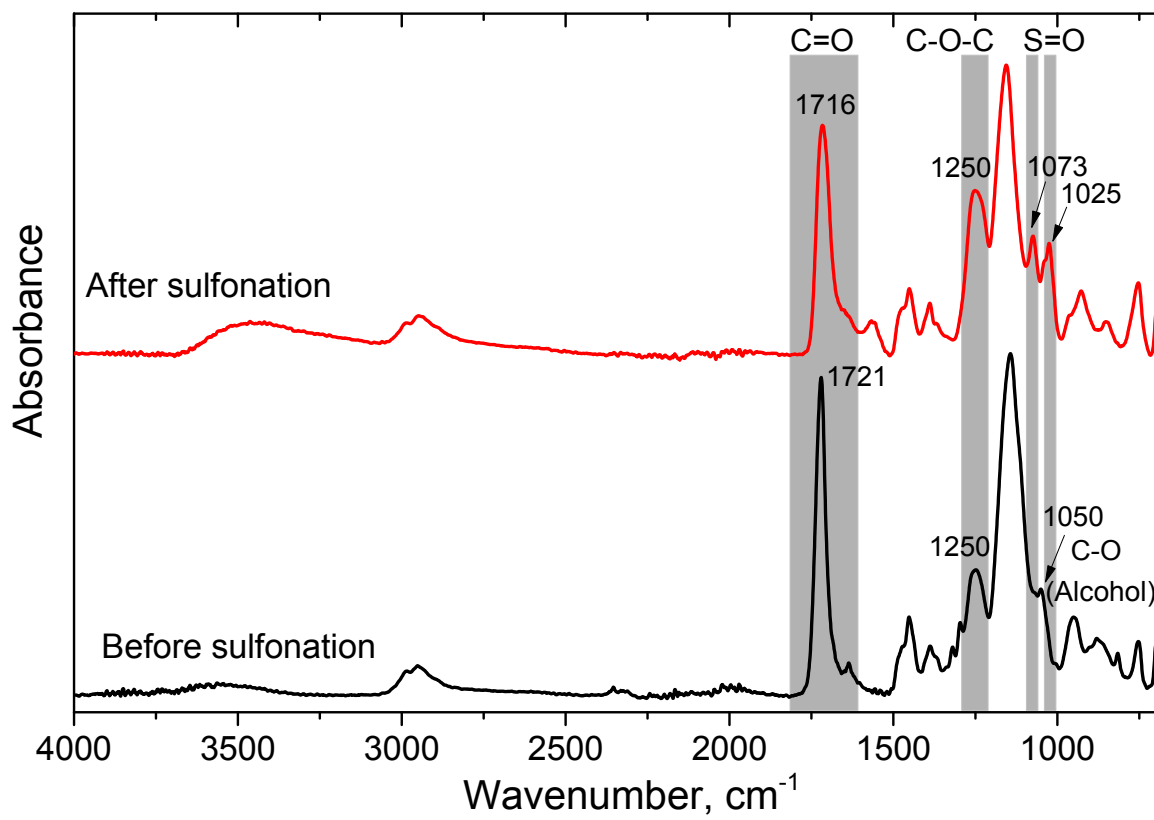


Fig. S3. FTIR-ATR spectra of polyelectrolytes microspheres with PVA in the surface layer.

The peak at 1025 cm^{-1} represents the symmetric stretching vibration of the SO_3 -groups which is significantly lower than that of the ketone groups.

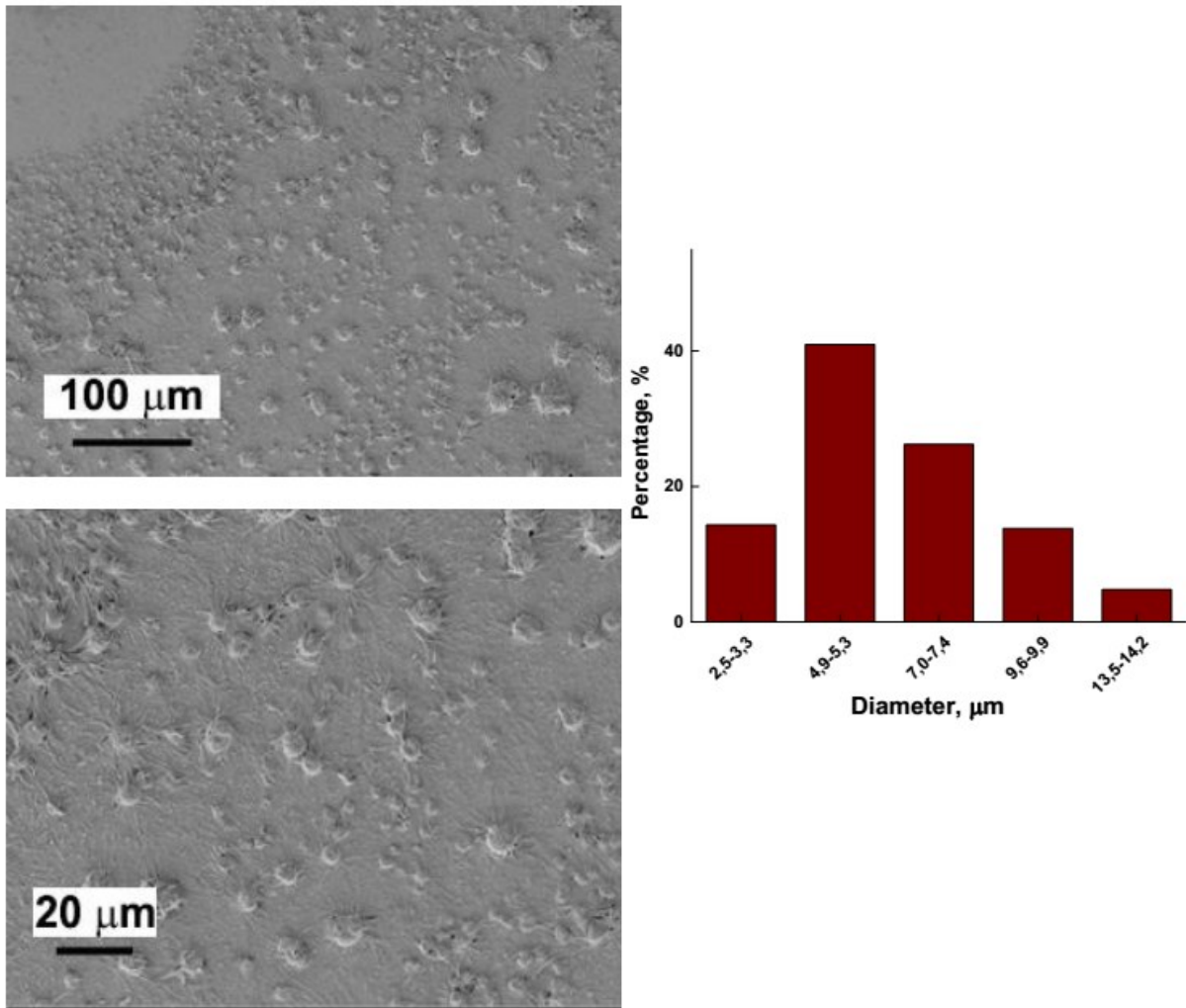


Fig. S3. SEM images of polyelectrolytes PNaSS6 microspheres and their respective particle size distribution. (C.V. 52%). C.V.= (Mean/ Standard Deviation)*100%.

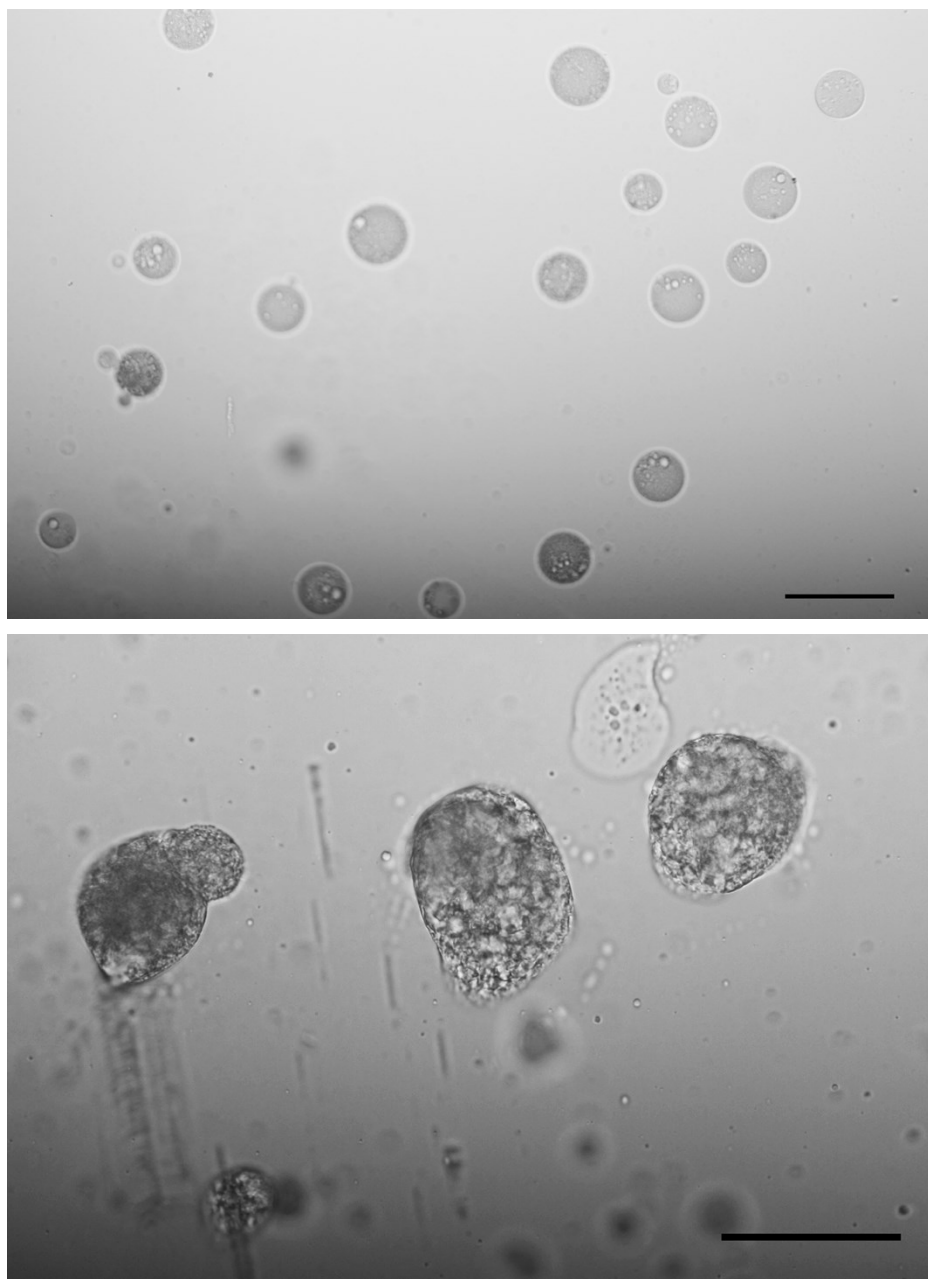


Fig. S4. Optical microscopy of polyelectrolyte PNaSS6 microspheres after removal of cyclohexane (scale bars 10 μm)