

Using Siloxane-based Liquid Electrolytes with High Stability for Fluoride Shuttle Batteries

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Supporting information

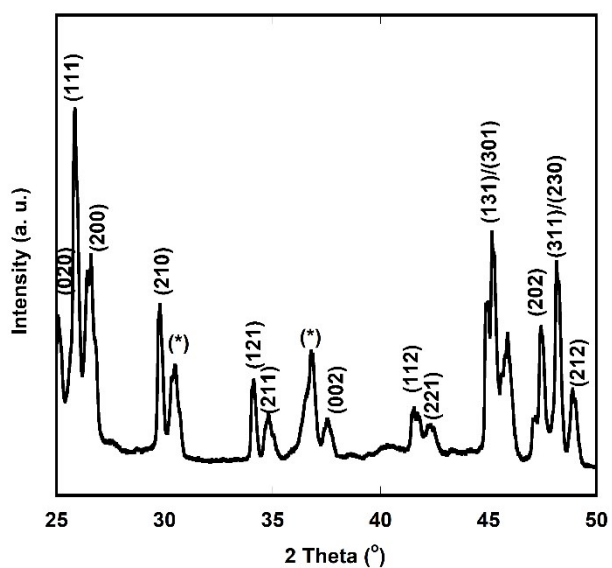


Fig. S1 X-ray diffraction patterns of BiF₃/C in the pristine state; * represents the peak related to the pocket.

a)

b)

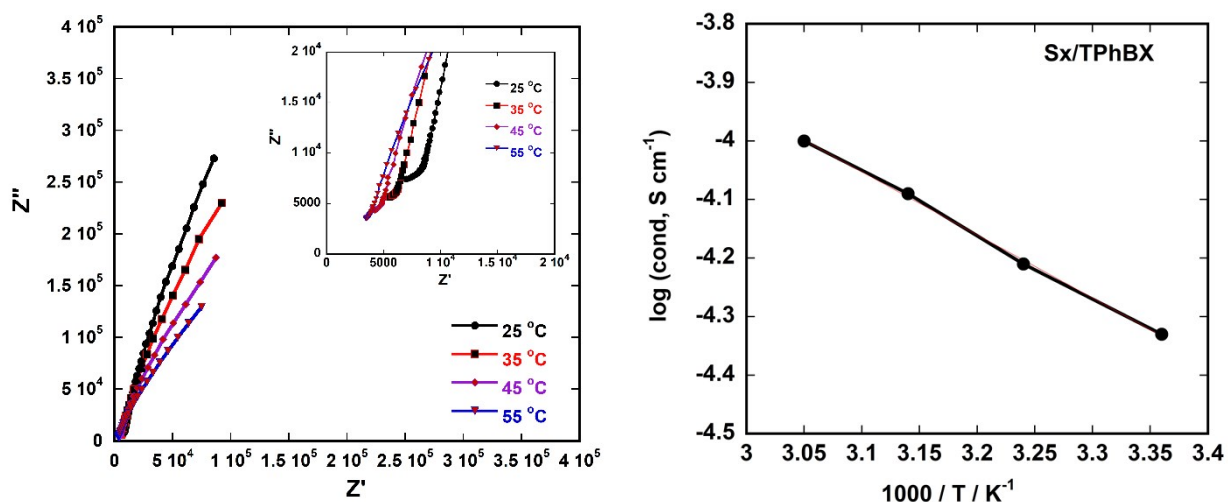


Fig. S2 (a) Temperature dependence of the ionic conductivity for the Sx/TPhBX electrolyte solution (b) Temperature dependence of the conductivity of Sx/TPhBX prepared using Sx; 2,2,4,4-Tetramethyl-3,8,11,14,17-pentaoxa-2,4-disilaooctadecane and TPhBX; 2,4,6-triphenylboroxin, with saturated CsF.

Table S1 Comparison between the temperature dependence of the conductivity of the Sx/TPhBX and Sx solutions (TPhBX; triphenyl boroxine, Sx; 2,2,4,4-tetramethyl-3,8,11,14,17-pentaoxa-2,4-disilaooctadecane, Ea; Activation Energy)

	25 °C	35 °C	45 °C	55 °C	Ea (kJ/mol)
Sx/TPhBX	4.62×10^{-5}	6.24×10^{-5}	8.06×10^{-5}	9.96×10^{-5}	20.61

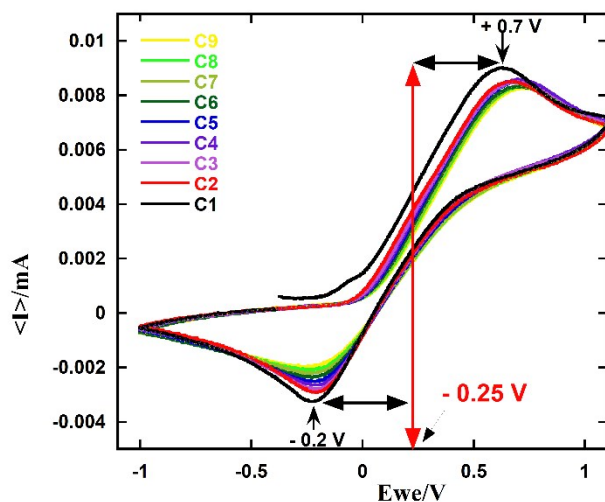


Fig. S3 Determination of the redox potential of Ag^+/Ag in the Sx/TPhBX electrolyte against SHE by adding 0.004 M ferrocene to Sx/TPhBX. The anodic (E_{pa}) and cathodic (E_{pc}) peak potentials were obtained at a scan rate of 5 mV s^{-1} . The halfwave potential ($E_{1/2}$) of the $\text{Fe}(\text{C}_2\text{H}_5)_2^+/\text{Fe}(\text{C}_2\text{H}_5)_2$ (Fc^+/Fc) couple in the Sx/TPhBX electrolyte was found to be 0.25 V.

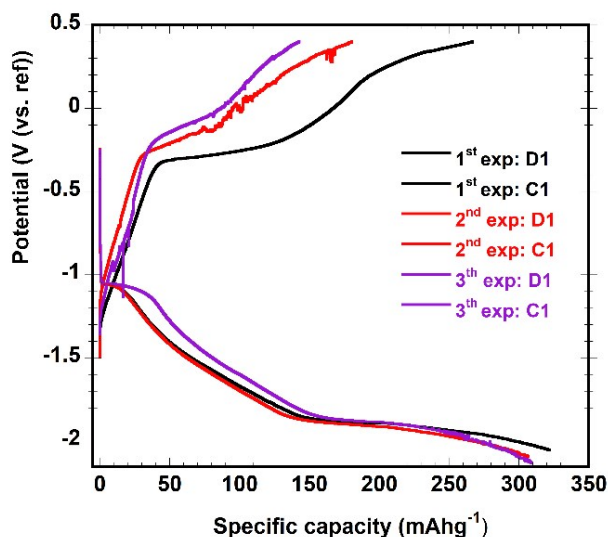


Fig. S4 The potential profile of the BiF_3/C nanocomposite electrode in Sx/TPhBX was investigated three times (1st exp: 1st experiment, 2nd exp: 2nd experiment, 3th exp: 3th experiment).

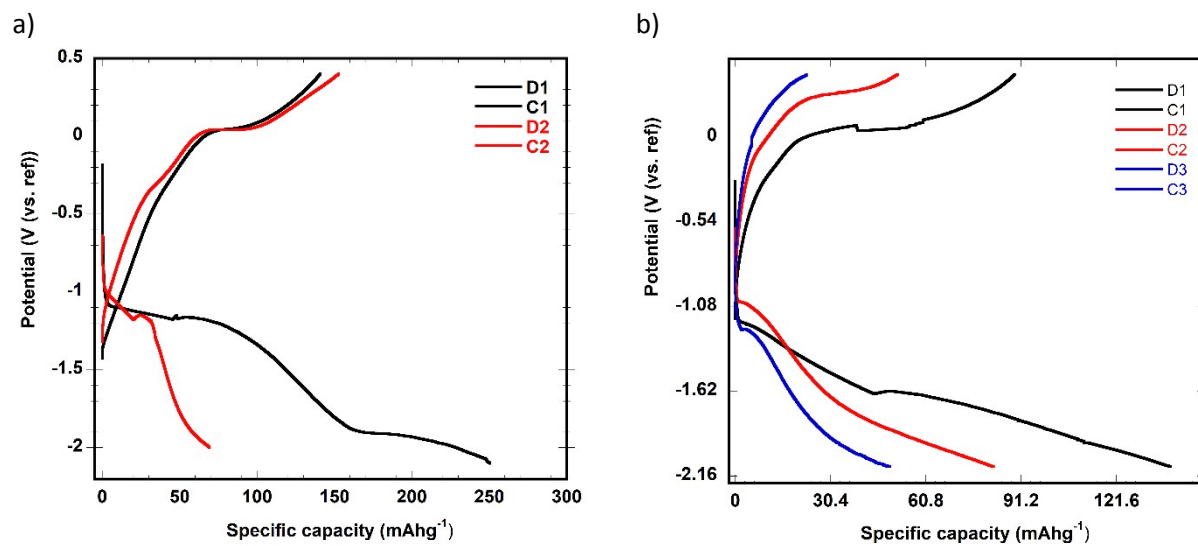


Fig. S5 Initial two or three discharge-charge curves of BiF₃/C nanocomposite electrodes cycled at room temperature and at (a) C/20, and (b) C/10 rates in Sx/TPhBX (discharging cutoff voltage: -2.05 V; charging cutoff voltage: 0.4 V).