

## ***Supplementary Information***

### **Reaction Kinetics of 1,4-Cyclohexadienes with OH radicals: Part I. An Experimental and Theoretical Study**

Binod Raj Giri,<sup>1</sup> Tam V.-T. Mai<sup>2,3,4</sup>, Mohamed Assali<sup>5</sup>, Thi T.-D. Nguyen<sup>4,6</sup>, Hieu T. Nguyen<sup>2</sup>, Milán Szőri<sup>7</sup>, Lam K. Huynh<sup>\*4,6</sup>, Christa Fittschen<sup>\*5</sup> and Aamir Farooq<sup>1</sup>

<sup>1</sup>King Abdullah University of Science and Technology (KAUST), Clean Combustion Research Center, Physical Sciences and Engineering Division, Thuwal 23955-6900, Saudi Arabia

<sup>2</sup> Molecular Science and Nano-Materials Lab, Institute for Computational Science and Technology, SBI Building, Quang Trung Software City, Tan Chanh Hiep Ward, District 12, Ho Chi Minh City, Vietnam.

<sup>3</sup> University of Science, Vietnam National University – HCMC, 227 Nguyen Van Cu, Ward 4, District 5, Ho Chi Minh City, Vietnam.

<sup>4</sup> Vietnam National University – HCMC, Quarter 6, Linh Trung Ward, Thu Duc District, Ho Chi Minh City, Vietnam.

<sup>5</sup> Université Lille, CNRS, UMR 8522 - PC2A - Physicochimie des Processus de Combustion et de l'Atmosphère, F-59000 Lille, France

<sup>6</sup> International University, Quarter 6, Linh Trung Ward, Thu Duc District, Ho Chi Minh City, Vietnam.

<sup>7</sup> Institute of Chemistry, Faculty of Materials Science and Engineering, University of Miskolc, Hungary

\*Corresponding authors' email: [hklam@hcmiu.edu.vn](mailto:hklam@hcmiu.edu.vn)

[christa.fittschen@univ-lille.fr](mailto:christa.fittschen@univ-lille.fr)

## Contents

<b>Table S1:</b> Optimized geometries, rotational constants (GHz), electronic energies at 0 K ( $E_{elec}^{0\text{K}}$ ), zero-point energy (ZPE) corrections and harmonic wavenumbers of the species involved with the lowest-energy conformer of a given species, calculated at M06-2X/aug-cc-pVTZ level of theory for the title reaction.....	4
<b>Table S2:</b> Calculated overall rate constants, $k_{tot}$ , of the $14\text{CHD} + \text{OH} \rightarrow \text{Products}$ over the range of temperature 200 – 2000 K at $P = 760$ Torr, including the HIR treatments, Eckart quantum tunneling effects and those collected from the literatures. Units are in $\text{cm}^3\cdot\text{molecule}^{-1}\cdot\text{s}^{-1}$ .....	10
<b>Table S3:</b> Calculated overall rate constants, $k_{tot}$ , of the $14\text{CHD} + \text{OH} \rightarrow \text{Products}$ over the range of temperature 200 – 2000 K at different pressures e.g., $P = 0.76 - 7600$ Torr, including the HIR treatments, Eckart quantum tunneling effects. Units are in $\text{cm}^3/\text{molecule/s}$ . The $k_{tot}(T, P)$ at different pressures are fitted as the double modified Arrhenius formats.....	11
<b>Table S4:</b> The calculated Eckart tunneling factor via tight transition state channels over the wide temperature range of 200 – 2000 K. ....	12
<b>Table S5:</b> Individual rate constants for $14\text{CHD} + \text{OH} \rightarrow \text{Products}$ ( $\text{cm}^3/\text{molecule/s}$ ) at different pressures (0.76, 7.6, 76, 760, and 7600 Torr). (“Unc.” stands for the uncertainty). Units are in $\text{cm}^3/\text{molecule/s}$ .....	13
<b>Table S6:</b> Calculated branching ratios (%) for each species of $14\text{CHD} + \text{OH} \rightarrow \text{Products}$ reactions at $P = 760$ Torr. ....	17
<b>Table S7:</b> Relative energies to that of the reactants ( $14\text{CHD} + \text{OH}$ ) of main TSs ( <b>TS1</b> , <b>TS4</b> and <b>TS5</b> ), calculated at M06-2X/aug-cc-pVTZ and CCSD(T)/CBS//M06-2X/aug-cc-pVTZ. Units are in kcal/mol. ....	17
<b>Table S8:</b> Ratios of $k_{\text{TST}}(T)/k_{\text{VTST}}(T)$ for the reaction channels $\text{RC} \rightarrow [\text{TS1}]^\ddagger \rightarrow \text{IM1}$ and $\text{RC} \rightarrow [\text{TS4}]^\ddagger \rightarrow \text{P3} + \text{H}_2\text{O}$ . The constant rate calculations were carried out with the minimum energy paths (MEPs), together with its properties (i.e., Hessian and Gradient), obtained at the M06-2X/aug-cc-pVTZ level.....	18
<b>Figure S1:</b> M06-2X/aug-cc-pVTZ optimized geometries for the species involved in the $14\text{CHD} + \text{OH}$ reaction. All structures were obtained for the lowest-energy conformer of a given species. Bond lengths and bond angles are in Å and degree (°), respectively. <sup>a,b,c</sup> obtained from Huber <i>et al.</i> <sup>1</sup> , Hoy <i>et al.</i> <sup>7</sup> , and Hellwege <i>et al.</i> <sup>8</sup> , respectively.....	20
<b>Figure S2:</b> Predicted rate coefficients, $k(T, P)$ , for the $14\text{CHD} + \text{OH} \rightarrow \text{I1}$ as functions of temperature at different pressures (e.g., 0.76, 7.6, 76, 760, and 7600 Torr). Literature data is from Peeters <i>et al.</i> <sup>9</sup> (“Calc. (Peeters 2007)”)......	20
<b>Figure S3:</b> Calculated overall rate coefficients, $k_{tot}(T, P)$ , for the $14\text{CHD} + \text{OH} \rightarrow \text{Products}$ as functions of temperature at different pressures (e.g., 0.76, 7.6, 76, 760, and 7600 Torr).....	21
<b>Figure S4:</b> Calculated overall rate constant, $k_{tot}$ , for the $14\text{CHD} + \text{OH} \rightarrow \text{products}$ reaction at $P = 760$ Torr as a function of temperature with (solid line) and without (dashed line) HIR treatment.....	21
<b>Figure S5:</b> Hindrance potentials for the species involved in the $14\text{CHD} + \text{OH}$ reaction, calculated at M06-2X/cc-pVDZ level of theory. ....	24
<b>Figure S6:</b> Comparison between the calculated and experimental global rate coefficients, $k(T, P)$ , for $14\text{CHD} + \text{OH} \rightarrow \text{Products}$ . Note that there is no energetic adjustment used in the calculations. ....	25
<b>Figure S7:</b> Comparison of the computed global rate constant, $k_{tot}$ , for $14\text{CHD} + \text{OH} \rightarrow \text{Products}$ reaction by using the $k^\circ(T)$ of $4.0 \times 10^{-10}$ (solid line), $4.0 \times 10^{-9}$ (dashed line) and $4.0 \times 10^{-11}$ (dotted line) $\text{cm}^3/\text{molecule/s}$ , at $T = 200 - 2000$ K & $P = 760$ Torr. ....	25

**Figure S8:** IRC data for **TS1** (a) and **TS4** (b) calculated at M06-2X/aug-cc-pVTZ level of theory.  
Distances are in Å..... 26

**Table S1:** Optimized geometries, rotational constants (GHz), electronic energies at 0 K ( $E_{elec}^{0\text{K}}$ ), zero-point energy (ZPE) corrections and harmonic wavenumbers of the species involved with the lowest-energy conformer of a given species, calculated at M06-2X/aug-cc-pVTZ level of theory for the title reaction.

Species	Cartesian coordinate (Å)				Rotational constants (GHz)	$E_{elec}^{0\text{K}}$ (Hartree)	ZPE (Hartree)	Unscaled vibrational frequencies (cm⁻¹)		
<b>OH</b> (C <sub>∞v</sub> )	8	0.0000000000	0.0000000000	0.107999000	564.211029	-75.733789	0.008530	3742.0378 (3737.8) <sup>1</sup>		
	1	0.0000000000	0.0000000000	-0.863995000						
<b>H<sub>2</sub>O</b> (C <sub>2v</sub> )	8	0.0000000000	0.0000000000	0.116332000	834.47311 431.04032 284.22579	-76.4300922	0.021565	1615.7190	3873.1681	3977.2474
	1	0.0000000000	0.762680000	-0.465326000				(1595.0; 3657.0 <sup>2</sup> ; 3756.0 <sup>1</sup> )		
	1	0.0000000000	-0.762680000	-0.465326000						
<b>H</b> (D <sub>∞h</sub> )	1	1.216198000	-1.008264000	0.000000000	/	-0.498207	0.00000	/		
<b>14CHD</b> (D <sub>2h</sub> )	6	0.662714000	1.248273000	0.000039000	5.19892 4.94389 2.61366	-233.391815	0.123092	120.2293	389.9917	415.4630
	6	-0.662714000	1.248273000	-0.000016000				541.8477	579.8579	635.9665
	6	-1.493271000	0.000000000	-0.000061000				731.4655	873.8115	905.1970
	6	-0.662714000	-1.248273000	-0.000040000				961.1601	967.1537	974.9870
	6	0.662714000	-1.248273000	0.000015000				992.8766	1021.5142	1039.9698
	6	1.493271000	0.000000000	0.000063000				1051.2751	1176.4052	1227.2242
	1	-2.161712000	-0.000008000	0.867632000				1230.6402	1230.8702	1366.3316
	1	-1.194524000	2.192590000	-0.000031000				1392.9875	1417.6612	1441.6111
	1	1.194524000	2.192590000	0.000070000				1474.0635	1477.1235	1734.1692
	1	-1.194524000	-2.192590000	-0.000073000				1780.3398	3029.4932	3031.3341
	1	1.194524000	-2.192590000	0.000027000				3043.7916	3045.1906	3171.0332
	1	2.161640000	-0.000008000	0.867812000				3171.3025	3192.5942	3194.3053
<b>RC</b> (C <sub>s</sub> )	6	-0.636695000	0.664414000	-1.036871000	2.89060 2.51017 2.24494	-309.135752	0.134257	78.7539	78.9420	138.1434
	6	-0.636650000	-0.664486000	-1.036856000				145.8412	336.0845	386.9745
	6	0.380957000	-1.494025000	-0.313428000				415.2247	513.1466	542.1583
	6	1.324438000	-0.663910000	0.506443000				580.2514	646.3938	741.8548
	6	1.324392000	0.664006000	0.506428000				872.9647	905.7025	954.0078
	6	0.380854000	1.494039000	-0.313461000				957.9271	983.9148	993.5833
	1	-0.130841000	-2.218237000	0.328460000				1022.1280	1044.1668	1050.3100
	1	-1.413770000	-1.196407000	-1.571543000				1177.4284	1227.4712	1227.7715
	1	-1.413853000	1.196270000	-1.571568000				1228.5405	1365.9759	1389.9499
	1	2.040658000	-1.195646000	1.121900000				1417.1172	1443.0305	1469.4024
	1	2.040577000	1.195805000	1.121872000				1472.5426	1717.3631	1767.5658









Species	Cartesian coordinate (Å)				Rotational constants (GHz)	$E_{elec}^{0K}$ (Hartree)	ZPE (Hartree)	Unscaled vibrational frequencies (cm <sup>-1</sup> )		
<b>P3</b> (C <sub>2v</sub> )	6	1.061963000	-1.181274000	0.134086000				807.2136	834.4487	902.6028
	6	1.667968000	-0.016676000	0.366564000				956.2380	977.8928	983.3483
	6	1.042862000	1.240409000	-0.051426000				1016.0390	1025.0994	1079.1969
	1	-0.845345000	-2.083904000	-0.290040000				1108.6206	1169.7765	1200.6063
	1	-0.708070000	2.217781000	-0.704108000				1214.9097	1278.0798	1344.2357
	1	1.528521000	-2.113675000	0.424679000				1361.0997	1409.5683	1425.0225
	1	2.634198000	0.019104000	0.851646000				1444.4834	1455.9879	1671.9528
	1	1.633105000	2.146992000	-0.013468000				1733.1617	3027.5568	3107.0441
	1	-0.045648000	-1.332724000	-1.662009000				3109.9988	3188.6080	3195.1505
	1	-1.860352000	0.076249000	-1.148795000				3210.8045	3218.0908	3850.6383
	8	-1.799993000	0.008970000	0.848397000						
	1	-1.165463000	0.138305000	1.560388000						
<b>P4</b> (C <sub>s</sub> )	6	-0.619666000	-1.248242000	-0.000001000	5.39110 5.30371 2.71671	-232.762553	0.109461	166.8983	388.5388	528.8322
	6	0.735115000	-1.221179000	-0.000002000				562.4582	589.7407	643.6157
	6	1.449138000	0.000000000	-0.000001000				745.4125	796.8784	887.9329
	6	0.735115000	1.221179000	-0.000002000				941.7741	982.1833	992.1747
	6	-0.619666000	1.248242000	-0.000001000				994.0679	997.0331	997.2256
	6	-1.441322000	0.000000000	0.000005000				1098.4352	1165.6252	1193.4141
	1	1.285917000	-2.152897000	-0.000004000				1199.0590	1312.8759	1367.3338
	1	-1.143252000	-2.195184000	-0.000003000				1422.0177	1439.3877	1458.0230
	1	1.285916000	2.152897000	-0.000004000				1567.3665	1634.0315	2971.3039
	1	-1.143252000	2.195184000	-0.000003000				2974.3542	3189.7098	3191.0994
	1	-2.123086000	0.000000000	0.863375000				3207.7366	3208.9364	3232.4877
	1	-2.123100000	0.000000000	-0.863353000						
	1	2.528574000	0.000000000	-0.000001000						

Frequencies in the parentheses (“( )”) are taken from experimental studies.

**Table S2:** Calculated overall rate constants,  $k_{\text{tot}}$ , of the  $14\text{CHD} + \text{OH} \rightarrow \text{products}$  over the range of temperature 200 – 2000 K at  $P = 760$  Torr, including the HIR treatments, Eckart quantum tunneling effects and those collected from the literatures.

This work (Calc.)		This work (Expt.)		Lui <i>et al.</i> (2020) <sup>3</sup>		Ohta <i>et al.</i> (1983) <sup>4</sup>		Atkinson <i>et al.</i> (1983) <sup>5</sup>		Grosjean <i>et al.</i> (1992) <sup>6</sup>	
T (K)	$k_{\text{tot}} (\times 10^{11})$ (cm <sup>-3</sup> ·molecule <sup>-1</sup> ·s <sup>-1</sup> )	T (K)	$k_{\text{tot}} (\times 10^{11})$ (cm <sup>-3</sup> ·molecule <sup>-1</sup> ·s <sup>-1</sup> )	T (K)	$k_{\text{tot}} (\times 10^{11})$ (cm <sup>-3</sup> ·molecule <sup>-1</sup> ·s <sup>-1</sup> )	T (K)	$k_{\text{tot}} (\times 10^{11})$ (cm <sup>-3</sup> ·molecule <sup>-1</sup> ·s <sup>-1</sup> )	T (K)	$k_{\text{tot}} (\times 10^{11})$ (cm <sup>-3</sup> ·molecule <sup>-1</sup> ·s <sup>-1</sup> )	T (K)	$k_{\text{tot}} (\times 10^{11})$ (cm <sup>-3</sup> ·molecule <sup>-1</sup> ·s <sup>-1</sup> )
200	23.1	<b>295.15</b>	<b>9.73</b>	921	3.11	<b>297</b>	<b>9.86</b>	<b>298</b>	<b>9.48</b>	<b>298</b>	<b>7.21</b>
250	17.2	322.15	9.1	923	3.11						
<b>300</b>	<b>11.9</b>	372.15	7.74	941	3.16						
400	5.39	399.15	7.36	956	4.10						
500	2.63	438.15	6.32	991	3.49						
600	1.57			1014	3.65						
700	1.22			1027	3.77						
800	1.22			1040	3.74						
900	1.42			1046	4.27						
1000	1.77			1089	4.27						
1100	2.23										
1200	2.80										
1300	3.48										
1400	4.27										
1500	5.17										
1600	6.18										
1700	7.30										
1800	8.53										
1900	9.89										
2000	11.3										

**Table S3:** Calculated overall rate constants,  $k_{\text{tot}}$ , of the  $14\text{CHD} + \text{OH} \rightarrow$  products over the range of temperature 200 – 2000 K at different pressures e.g.,  $P = 0.76 – 7600$  Torr, including the HIR treatments, Eckart quantum tunneling effects. Units are in  $\text{cm}^3/\text{molecule/s}$ . The  $k_{\text{tot}}(T, P)$  at different pressures are fitted as the double modified Arrhenius formats.

T (K)	0.76 Torr	7.6 Torr	76 Torr	760 Torr	7600 Torr
200	2.29E-10	2.29E-10	2.29E-10	2.31E-10	2.48E-10
250	1.70E-10	1.70E-10	1.71E-10	1.72E-10	1.81E-10
300	1.19E-10	1.19E-10	1.19E-10	1.19E-10	1.24E-10
400	5.35E-11	5.37E-11	5.37E-11	5.39E-11	5.46E-11
500	2.57E-11	2.62E-11	2.62E-11	2.63E-11	2.64E-11
600	1.50E-11	1.56E-11	1.57E-11	1.57E-11	1.57E-11
700	1.14E-11	1.21E-11	1.22E-11	1.22E-11	1.23E-11
800	1.15E-11	1.20E-11	1.22E-11	1.22E-11	1.22E-11
900	1.37E-11	1.40E-11	1.42E-11	1.42E-11	1.42E-11
1000	1.74E-11	1.75E-11	1.76E-11	1.77E-11	1.77E-11
1100	2.21E-11	2.22E-11	2.22E-11	2.23E-11	2.23E-11
1200	2.79E-11	2.79E-11	2.80E-11	2.80E-11	2.80E-11
1300	3.48E-11	3.48E-11	3.49E-11	3.48E-11	3.48E-11
1400	4.27E-11	4.27E-11	4.27E-11	4.27E-11	4.27E-11
1500	5.17E-11	5.17E-11	5.17E-11	5.17E-11	5.17E-11
1600	6.18E-11	6.17E-11	6.18E-11	6.18E-11	6.18E-11
1700	7.30E-11	7.29E-11	7.30E-11	7.30E-11	7.30E-11
1800	8.53E-11	8.53E-11	8.53E-11	8.53E-11	8.54E-11
1900	9.88E-11	9.88E-11	9.88E-11	9.89E-11	9.90E-11
2000	1.14E-10	1.14E-10	1.14E-10	1.13E-10	1.13E-10

- $k(T) = 3.58 \times 10^5 \times T^{5.67} \times \exp[-986.0 \text{ K}/T] + 3.35 \times 10^{-18} \times T^{2.33} \times \exp[-705.2 \text{ K}/T]$   
( $\text{cm}^3/\text{molecule/s}$ ) ( $T = 200 – 2000 \text{ K}$  &  $P = \mathbf{0.76 \text{ Torr}}$ ; error = 0.4 %).
- $k(T) = 2.93 \times 10^5 \times T^{5.64} \times \exp[-977.2 \text{ K}/T] + 2.77 \times 10^{-19} \times T^{2.63} \times \exp[-254.8 \text{ K}/T]$   
( $\text{cm}^3/\text{molecule/s}$ ) ( $T = 200 – 2000 \text{ K}$  &  $P = \mathbf{7.6 \text{ Torr}}$ ; error = 0.3 %).
- $k(T) = 3.52 \times 10^5 \times T^{5.67} \times \exp[-983.8 \text{ K}/T] + 1.71 \times 10^{-19} \times T^{2.68} \times \exp[-155.6 \text{ K}/T]$   
( $\text{cm}^3/\text{molecule/s}$ ) ( $T = 200 – 2000 \text{ K}$  &  $P = \mathbf{76 \text{ Torr}}$ ; error = 0.2 %).
- $k(T) = 2.73 \times 10^5 \times T^{5.63} \times \exp[-971.4 \text{ K}/T] + 3.68 \times 10^{-19} \times T^{2.59} \times \exp[-267.4 \text{ K}/T]$   
( $\text{cm}^3/\text{molecule/s}$ ) ( $T = 200 – 2000 \text{ K}$  &  $P = \mathbf{760 \text{ Torr}}$ ; error = 0.2 %).
- $k(T) = 4.59 \times 10^8 \times T^{5.72} \times \exp[-971.6 \text{ K}/T] + 1.41 \times 10^{-19} \times T^{2.70} \times \exp[-92.1 \text{ K}/T]$   
( $\text{cm}^3/\text{molecule/s}$ ) ( $T = 200 – 2000 \text{ K}$  &  $P = \mathbf{7600 \text{ Torr}}$ ; error = 0.3 %).

**Table S4:** The calculated Eckart tunneling factor via tight transition state channels over the wide temperature range of 200 – 2000 K.

T (K)	via TS1	via TS2	via TS3	via TS4	via TS5
200	1.21	13.40	5.18	3.05	3.89
250	1.13	4.70	2.76	2.01	2.46
300	1.09	2.82	2.00	1.62	1.90
400	1.05	1.76	1.47	1.31	1.46
500	1.03	1.44	1.28	1.19	1.29
600	1.02	1.29	1.19	1.13	1.20
700	1.02	1.21	1.14	1.09	1.15
800	1.01	1.16	1.10	1.07	1.12
900	1.01	1.12	1.08	1.06	1.09
1000	1.01	1.10	1.07	1.05	1.08
1100	1.01	1.08	1.06	1.04	1.06
1200	1.01	1.07	1.05	1.03	1.06
1300	1.01	1.06	1.04	1.03	1.05
1400	1.00	1.05	1.03	1.02	1.04
1500	1.00	1.04	1.03	1.02	1.04
1600	1.00	1.04	1.03	1.02	1.03
1700	1.00	1.04	1.02	1.02	1.03
1800	1.00	1.03	1.02	1.01	1.03
1900	1.00	1.03	1.02	1.01	1.03
2000	1.00	1.03	1.02	1.01	1.02







200	4.57E-13	2.60E-03	1.66E-10	7.20E-06	1.20E-17	1.00E+02	1.20E-17	1.00E+02	8.14E-11	1.50E-05	1.06E-13	1.10E-02
250	7.32E-14	1.60E-02	1.18E-10	1.00E-05	1.20E-17	1.00E+02	1.20E-17	1.00E+02	6.28E-11	1.90E-05	2.18E-13	5.50E-03
300	1.98E-14	6.10E-02	7.80E-11	1.50E-05	1.20E-17	1.00E+02	1.20E-17	1.00E+02	4.53E-11	2.70E-05	3.97E-13	3.00E-03
400	2.66E-15	4.50E-01	3.13E-11	3.80E-05	1.20E-17	1.00E+02	1.20E-17	1.00E+02	2.23E-11	5.40E-05	9.94E-13	1.20E-03
500	4.22E-16	2.90E+00	1.29E-11	9.30E-05	1.21E-17	1.00E+02	1.21E-17	1.00E+02	1.15E-11	1.00E-04	2.01E-12	6.00E-04
600	3.63E-17	3.30E+01	5.82E-12	2.10E-04	1.21E-17	1.00E+02	1.21E-17	1.00E+02	6.43E-12	1.90E-04	3.49E-12	3.50E-04
700	1.22E-17	1.00E+02	2.87E-12	4.20E-04	3.65E-17	3.30E+01	1.22E-17	1.00E+02	3.82E-12	3.20E-04	5.57E-12	2.20E-04
800	1.23E-17	1.00E+02	1.53E-12	8.00E-04	1.23E-16	1.00E+01	2.45E-17	5.00E+01	2.40E-12	5.10E-04	8.31E-12	1.50E-04
900	1.24E-17	1.00E+02	8.67E-13	1.40E-03	2.97E-16	4.20E+00	4.94E-17	2.50E+01	1.59E-12	7.80E-04	1.18E-11	1.00E-04
1000	1.25E-17	1.00E+02	5.07E-13	2.50E-03	7.49E-16	1.70E+00	2.25E-16	5.60E+00	1.12E-12	1.10E-03	1.61E-11	7.80E-05
1100	1.26E-17	1.00E+02	2.79E-13	4.50E-03	1.66E-15	7.60E-01	4.93E-16	2.60E+00	8.09E-13	1.60E-03	2.12E-11	6.00E-05
1200	1.28E-17	1.00E+02	1.46E-13	8.80E-03	1.97E-15	6.50E-01	1.00E-15	1.30E+00	6.24E-13	2.10E-03	2.72E-11	4.70E-05
1300	1.30E-17	1.00E+02	6.51E-14	2.00E-02	2.91E-15	4.50E-01	1.17E-15	1.10E+00	4.99E-13	2.60E-03	3.43E-11	3.80E-05
1400	1.33E-17	1.00E+02	2.46E-14	5.40E-02	2.87E-15	4.60E-01	1.35E-15	9.80E-01	4.00E-13	3.30E-03	4.23E-11	3.10E-05
1500	1.35E-17	1.00E+02	7.57E-15	1.80E-01	2.90E-15	4.70E-01	1.30E-15	1.00E+00	3.19E-13	4.30E-03	5.14E-11	2.60E-05
1600	1.39E-17	1.00E+02	1.66E-15	8.30E-01	2.52E-15	5.50E-01	1.11E-15	1.20E+00	2.53E-13	5.50E-03	6.15E-11	2.30E-05
1700	1.42E-17	1.00E+02	2.84E-16	5.00E+00	2.18E-15	6.50E-01	1.04E-15	1.40E+00	2.04E-13	7.00E-03	7.28E-11	1.90E-05
1800	1.46E-17	1.00E+02	7.28E-17	2.00E+01	1.54E-15	9.40E-01	8.73E-16	1.70E+00	1.72E-13	8.50E-03	8.52E-11	1.70E-05
1900	1.50E-17	1.00E+02	1.50E-17	1.00E+02	1.54E-15	9.70E-01	7.63E-16	2.00E+00	1.48E-13	1.00E-02	9.88E-11	1.50E-05
2000	1.54E-17	1.00E+02	1.54E-17	1.00E+02	1.28E-15	1.20E+00	8.93E-16	1.70E+00	1.32E-13	1.20E-02	1.13E-10	1.40E-05

**Table S6:** Calculated branching ratios (%) for each species of  $14\text{CHD} + \text{OH} \rightarrow$  products reactions at  $P = 760$  Torr.

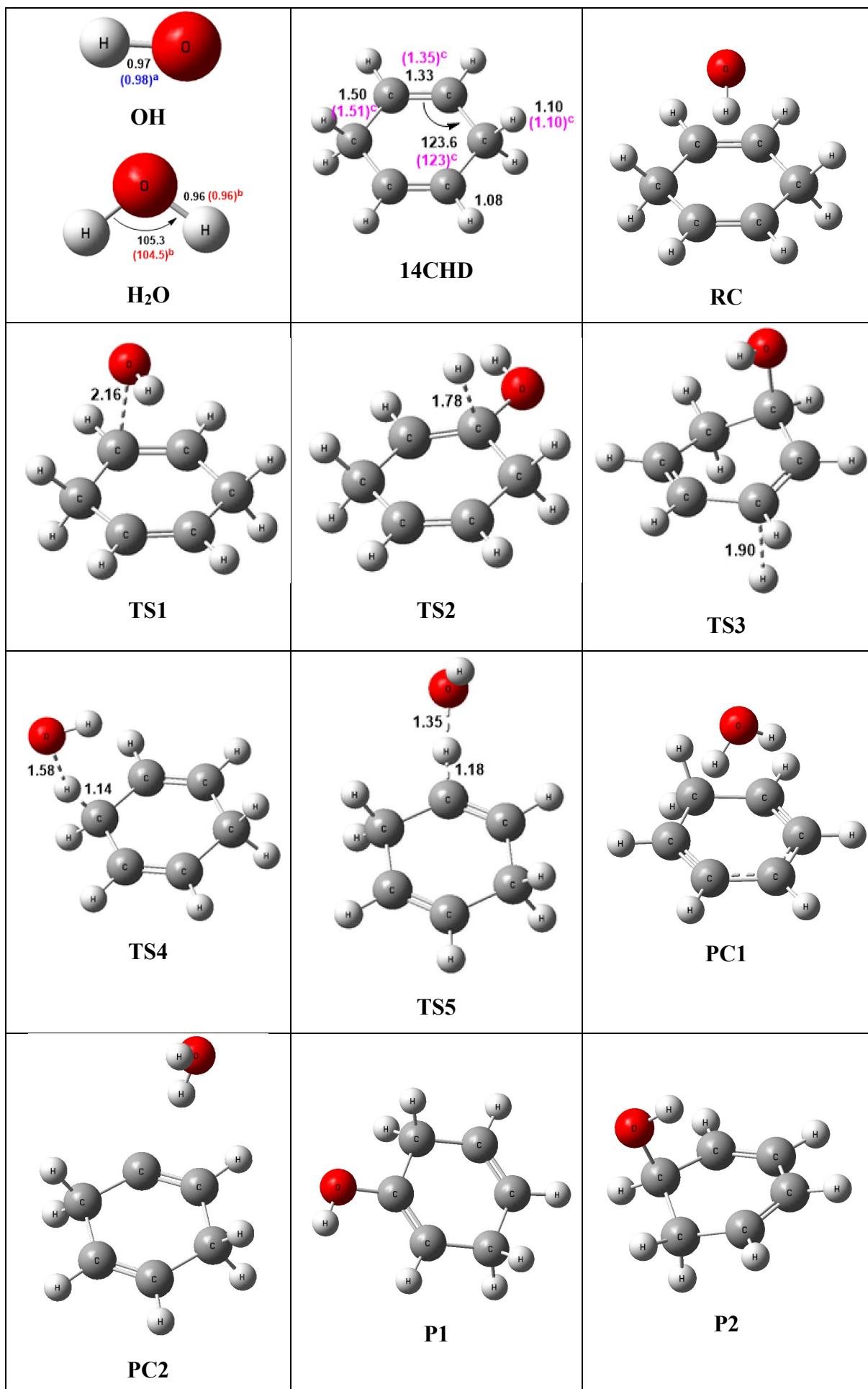
T (K)	RC	I1	P1 + H	P2 + H	P3 + H <sub>2</sub> O	P4 + H <sub>2</sub> O
200	0.00	65.77	0.00	0.00	34.19	0.04
250	0.00	64.36	0.00	0.00	35.51	0.13
300	0.00	62.52	0.00	0.00	37.15	0.33
400	0.00	57.12	0.00	0.00	41.04	1.84
500	0.00	48.63	0.00	0.00	43.76	7.60
600	0.00	36.84	0.00	0.00	40.90	22.26
700	0.00	23.26	0.00	0.00	31.36	45.38
800	0.00	12.10	0.00	0.00	19.98	67.91
900	0.00	5.37	0.01	0.00	11.81	82.81
1000	0.00	2.05	0.02	0.00	7.04	90.89
1100	0.00	0.64	0.01	0.00	4.19	95.15
1200	0.00	0.16	0.01	0.00	2.55	97.28
1300	0.00	0.03	0.01	0.00	1.55	98.41
1400	0.00	0.00	0.01	0.00	0.97	99.02
1500	0.00	0.00	0.00	0.00	0.61	99.38
1600	0.00	0.00	0.00	0.00	0.39	99.60
1700	0.00	0.00	0.00	0.00	0.27	99.73
1800	0.00	0.00	0.00	0.00	0.19	99.81
1900	0.00	0.00	0.00	0.00	0.14	99.86
2000	0.00	0.00	0.00	0.00	0.11	99.89

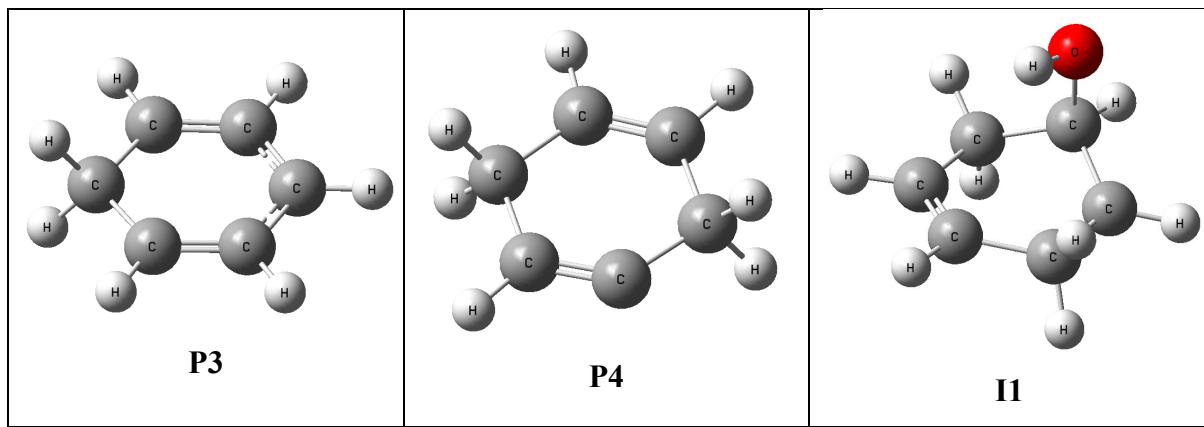
**Table S7:** Relative energies to that of the reactants ( $14\text{CHD} + \text{OH}$ ) of main TSs (**TS1**, **TS4** and **TS5**), calculated at M06-2X/aug-cc-pVTZ and CCSD(T)/CBS//M06-2X/aug-cc-pVTZ. Units are in kcal/mol.

Species	CCSD(T)/CBS//M06-2X/aug-cc-pVTZ	M06-2X/aug-cc-pVTZ
<b>TS1</b>	-2.6	-2.9
<b>TS4</b>	-0.9	-1.4
<b>TS5</b>	2.8	2.1

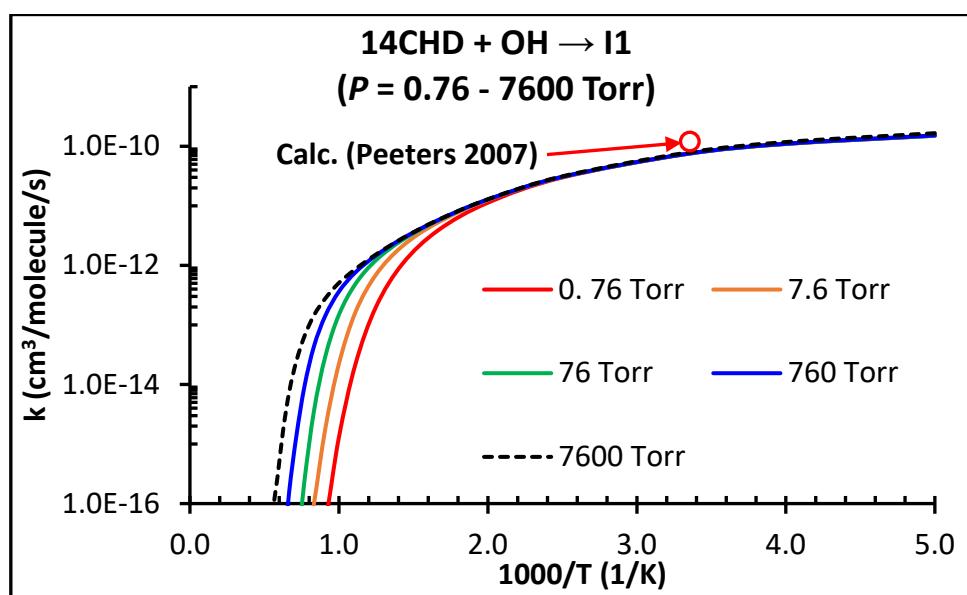
**Table S8:** Ratios of  $k_{\text{TST}}(T)/k_{\text{VTST}}(T)$  for the reaction channels of  $\text{RC} \rightarrow [\text{TS1}]^\ddagger \rightarrow \text{IM1}$  and  $\text{RC} \rightarrow [\text{TS4}]^\ddagger \rightarrow \text{P3} + \text{H}_2\text{O}$ . The constant rate calculations were carried out with the minimum energy paths (MEPs), together with its properties (i.e., Hessian and Gradient), obtained at the M06-2X/aug-cc-pVTZ level.

T (K)	$k_{\text{TST}}(T)/k_{\text{VTST}}(T) (P \sim \infty)$	
	$\text{RC} \rightarrow [\text{TS1}]^\ddagger \rightarrow \text{IM1}$	$\text{RC} \rightarrow [\text{TS4}]^\ddagger \rightarrow \text{P3} + \text{H}_2\text{O}$
200	1.00	1.00
250	1.00	1.00
300	1.00	1.31
400	1.00	1.87
500	1.00	1.58
600	1.00	1.42
700	1.00	1.32
800	1.00	1.26
900	1.00	1.22
1000	1.00	1.18
1100	1.00	1.16
1200	1.00	1.14
1300	1.00	1.13
1400	1.00	1.12
1500	1.00	1.11
1600	1.06	1.11
1700	1.21	1.10
1800	1.38	1.10
1900	1.56	1.09
2000	1.75	1.09

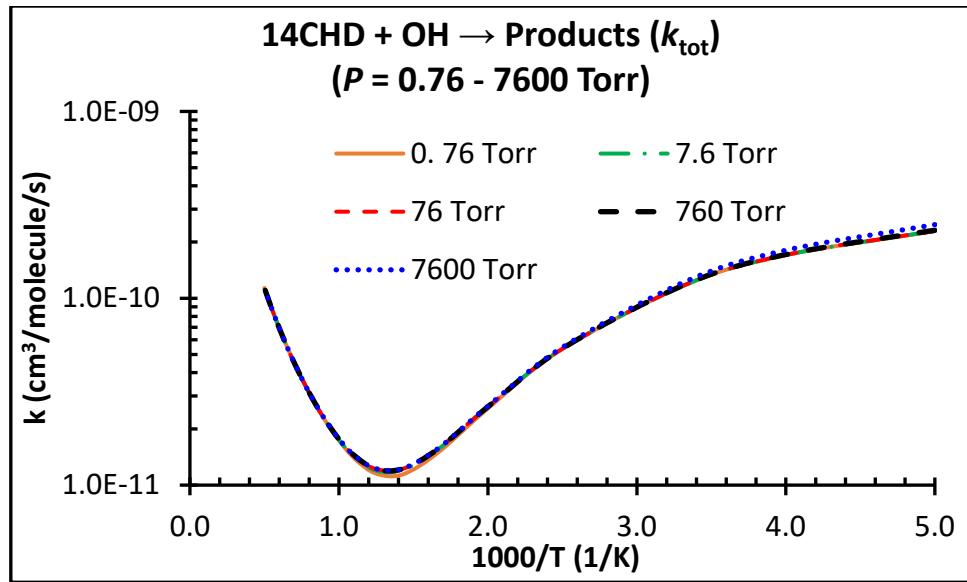




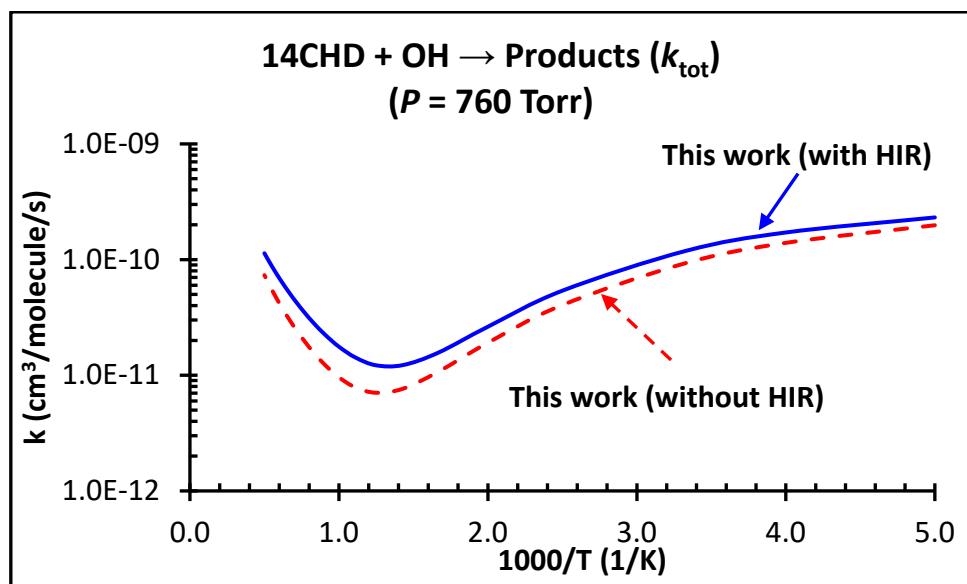
**Figure S1:** M06-2X/aug-cc-pVTZ optimized geometries for the species involved in the 14CHD + OH reaction. All structures were obtained for the lowest-energy conformer of a given species. Bond lengths and bond angles are in Å and degree ( $^{\circ}$ ), respectively. <sup>a, b, c</sup> obtained from Huber *et al.*<sup>1</sup>, Hoy *et al.*<sup>7</sup>, and Hellwege *et al.*<sup>8</sup>, respectively.



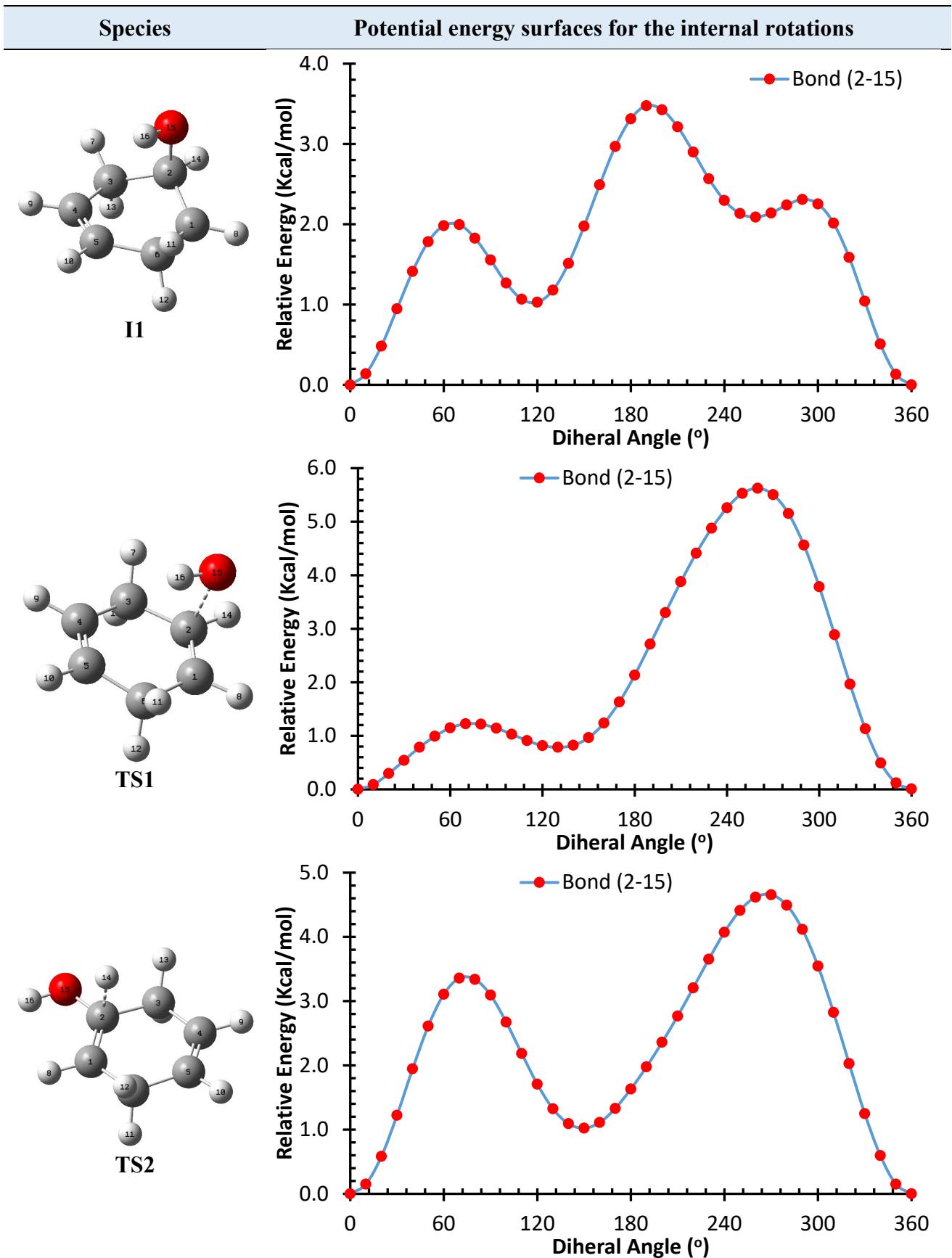
**Figure S2:** Predicted rate coefficients,  $k(T, P)$ , for the 14CHD + OH  $\rightarrow$  I1 as functions of temperature at different pressures (e.g., 0.76, 7.6, 76, 760, and 7600 Torr). Literature data is from Peeters *et al.*<sup>9</sup> (“Calc. (Peeters 2007)”).

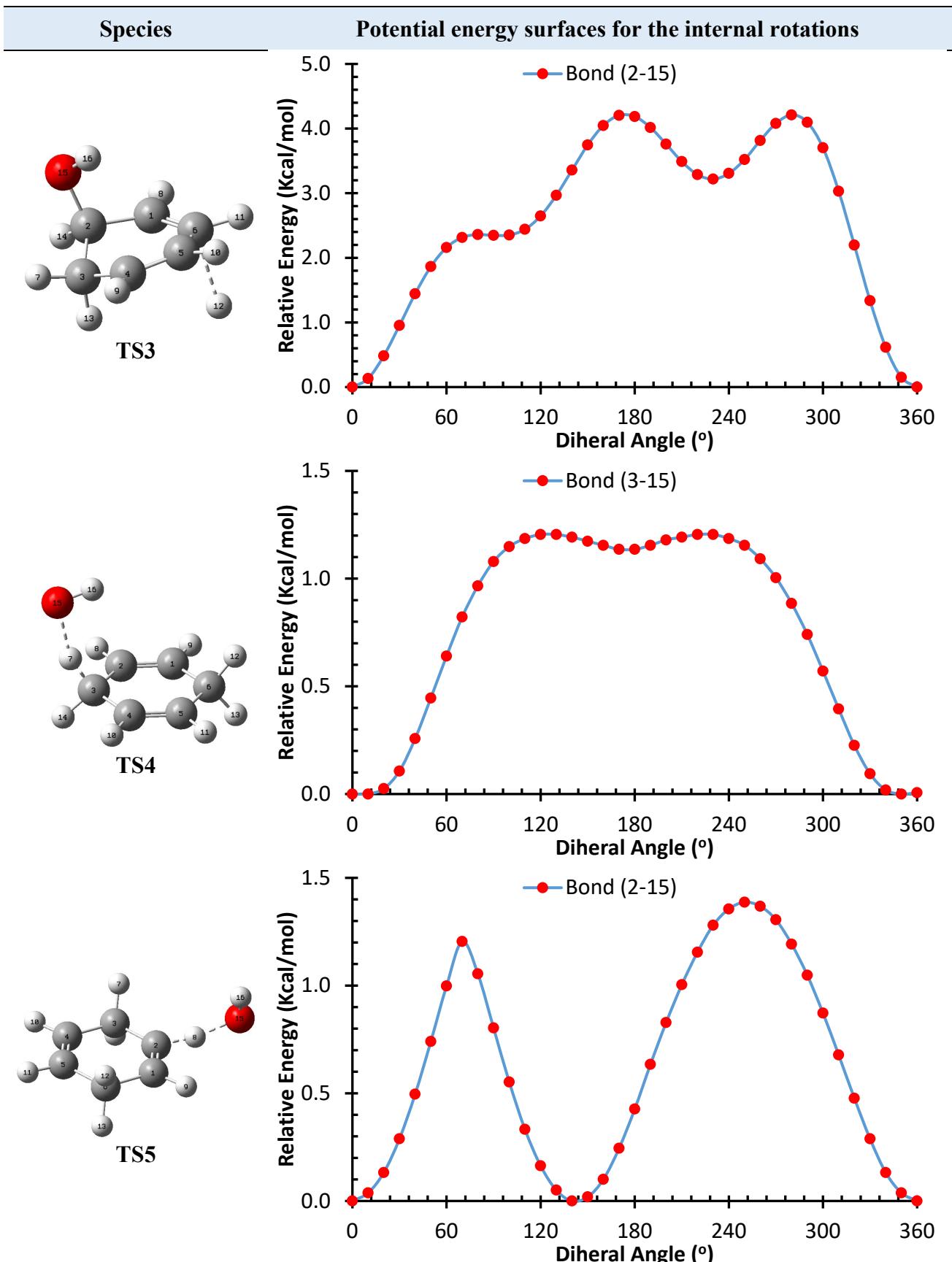


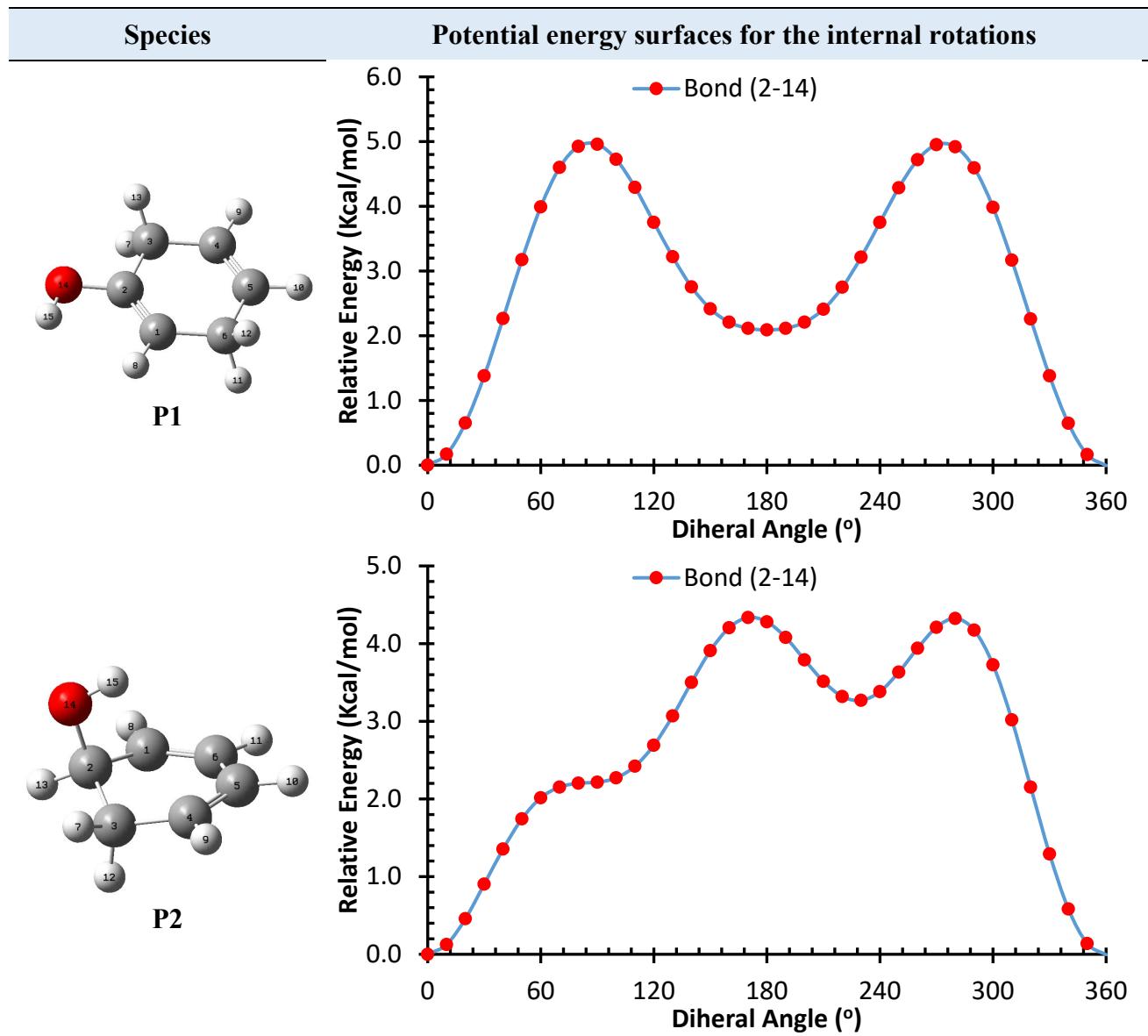
**Figure S3:** Calculated overall rate coefficients,  $k_{tot}(T, P)$ , for the  $14\text{CHD} + \text{OH} \rightarrow \text{Products}$  as functions of temperature at different pressures (e.g., 0.76, 7.6, 76, 760, and 7600 Torr).



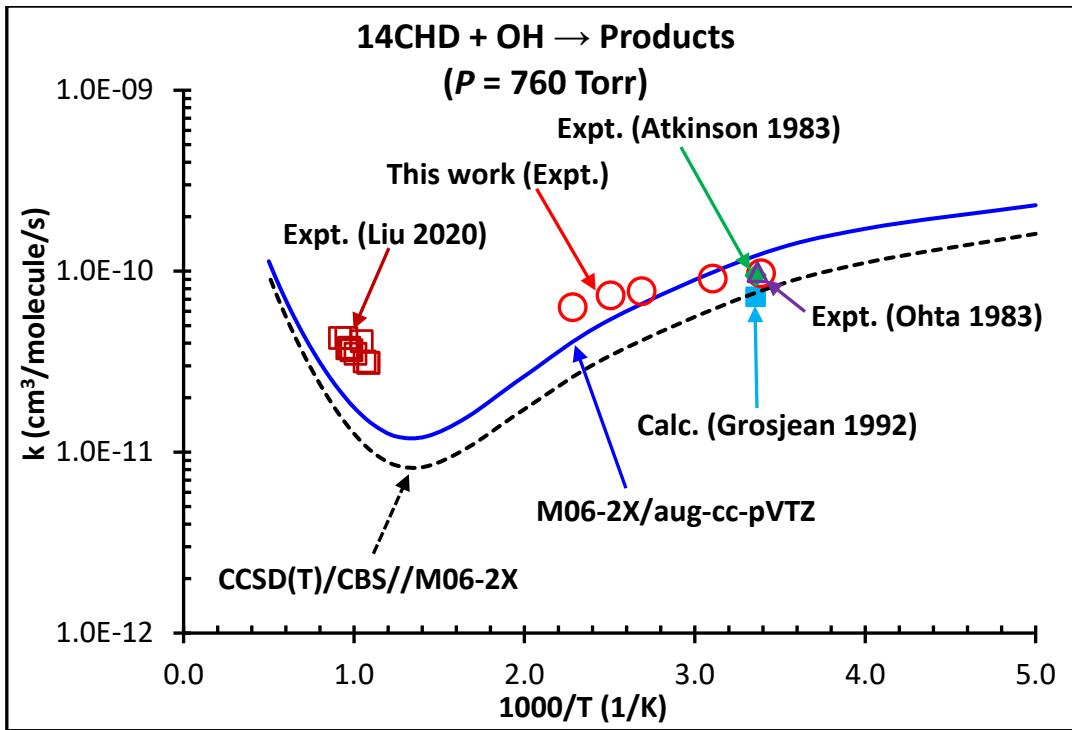
**Figure S4:** Calculated overall rate constant,  $k_{tot}$ , for the  $14\text{CHD} + \text{OH} \rightarrow \text{products}$  reaction at  $P = 760$  Torr as a function of temperature with (solid line) and without (dashed line) HIR treatment.



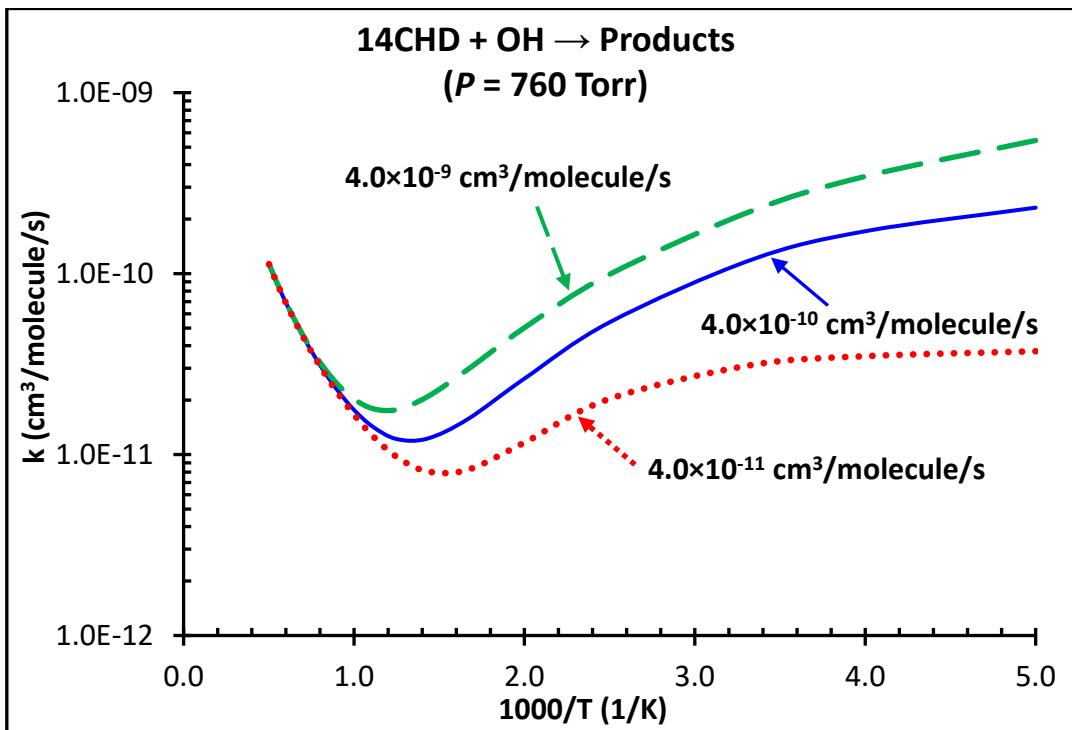




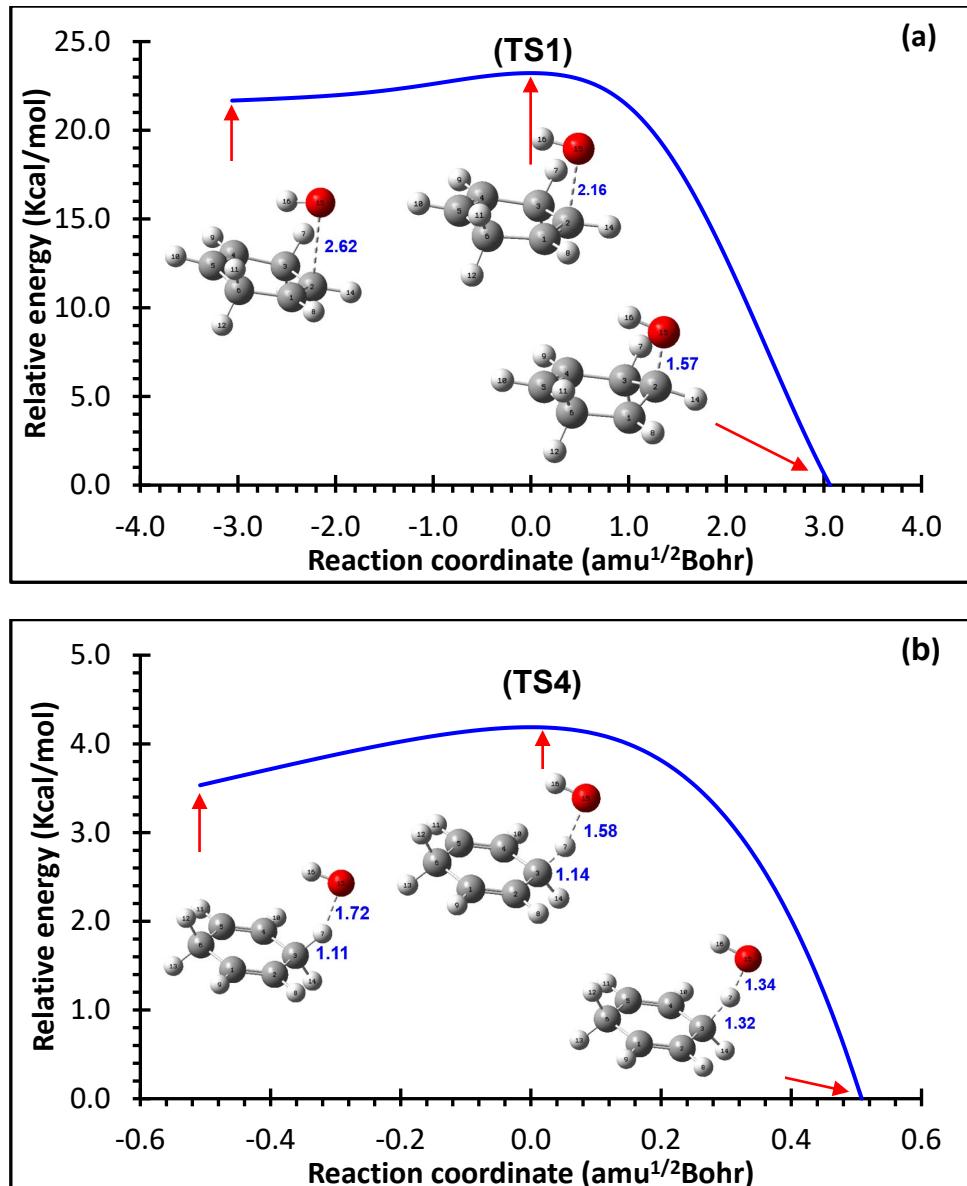
**Figure S5:** Hindrance potentials for the species involved in the  $^{14}\text{CHD} + \text{OH}$  reaction, calculated at M06-2X/cc-pVDZ level of theory.



**Figure S6:** Comparison between the calculated and experimental global rate coefficients,  $k(T, P)$ , for  $14\text{CHD} + \text{OH} \rightarrow \text{Products}$ . Note that there is no energetic adjustment used in the calculations.



**Figure S7:** Comparison of the computed global rate constant,  $k_{\text{tot}}$ , for  $14\text{CHD} + \text{OH} \rightarrow \text{Products}$  reaction by using the  $k^\circ(T)$  of  $4.0 \times 10^{-10}$  (solid line),  $4.0 \times 10^{-9}$  (dashed line) and  $4.0 \times 10^{-11}$  (dotted line)  $\text{cm}^3/\text{molecule/s}$ , at  $T = 200 - 2000 \text{ K}$  &  $P = 760 \text{ Torr}$ .



**Figure S8:** IRC data for **TS1** (a) and **TS4** (b) calculated at M06-2X/aug-cc-pVTZ level of theory. Distances are in Å.

## References:

1. K. P. Huber and G. Herzberg, *Molecular Spectra and Molecular Structure. IV. Constants of Diatomic Molecules*, Van Nostrand Reinhold Co, 1979.
2. T. Shimanouchi, *Tables of Molecular Vibrational Frequencies, Consolidated Volume 1, NSRDS NBS-39*.
3. D. Liu, B. R. Giri, M. Szőri, B. Viskolcz, L. K. Huynh and A. Farooq, *Proc. Combust. Inst.*, 2020, **38**, 947-955. (10.1016/j.proci.2020.06.259)
4. T. Ohta, *J. Phys. Chem.*, 1983, **87**, 1209-1213. (10.1021/j100230a023)
5. R. Atkinson, S. M. Aschmann and W. P. L. Carter, *Int. J. Chem. Kin.*, 1983, **15**, 1161-1177. (10.1002/kin.550151105)
6. D. Grosjean and E. L. Williams, *Atmos. Environ. Part A*, 1992, **26**, 1395-1405. (10.1016/0960-1686(92)90124-4)
7. A. R. Hoy and P. R. Bunker, *J. Mol. Spectrosc.*, 1979, **74**, 1-8. (10.1016/0022-2852(79)90019-5)
8. K. Hellwege and A. Hellwege, *Landolt-Bornstein: Group II: Atomic and Molecular Physics Structure Data of Free Polyatomic Molecules*, Springer-Verlag, Berlin, 1976.
9. J. Peeters, W. Boullart, V. Pultau, S. Vandenbergk and L. Vereecken, *J. Phys. Chem. A*, 2007, **111**, 1618-1631. (10.1021/jp066973o)