

Electronic Supplementary Information

Au/TiO₂ nanobelts: Thermal enhancement vs plasmon enhancement for visible-light-driven photocatalytic selective oxidation of amines into imines

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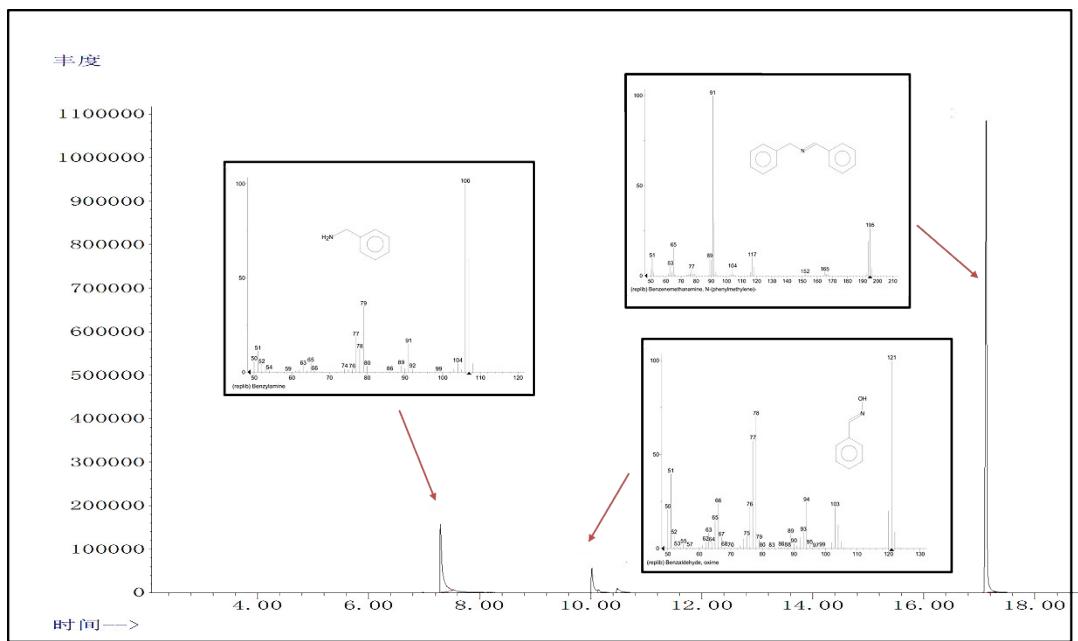


Figure S1. The GC-MS analysis of supernatant using 0.5 mol% Au/TiO₂ as photocatalyst for photocatalytic oxidation of benzylamine.

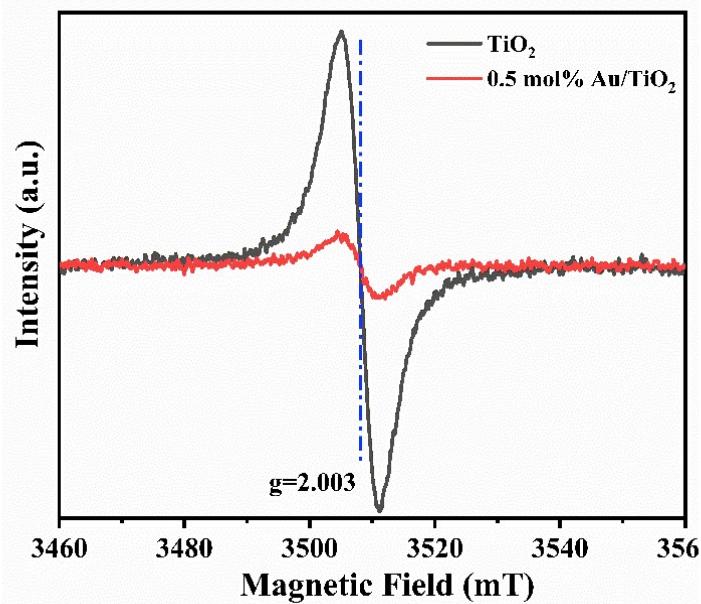


Figure S2. ESR spectrum of TiO₂ and 0.5 mol% Au/TiO₂ at room temperature.

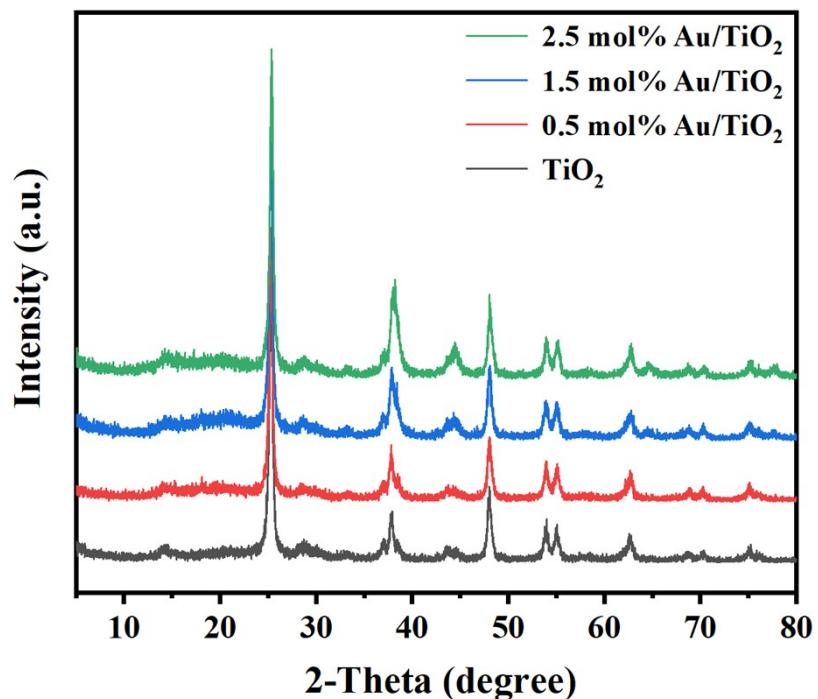


Figure S3. XRD patterns of as-prepared samples after the recycling experiment .

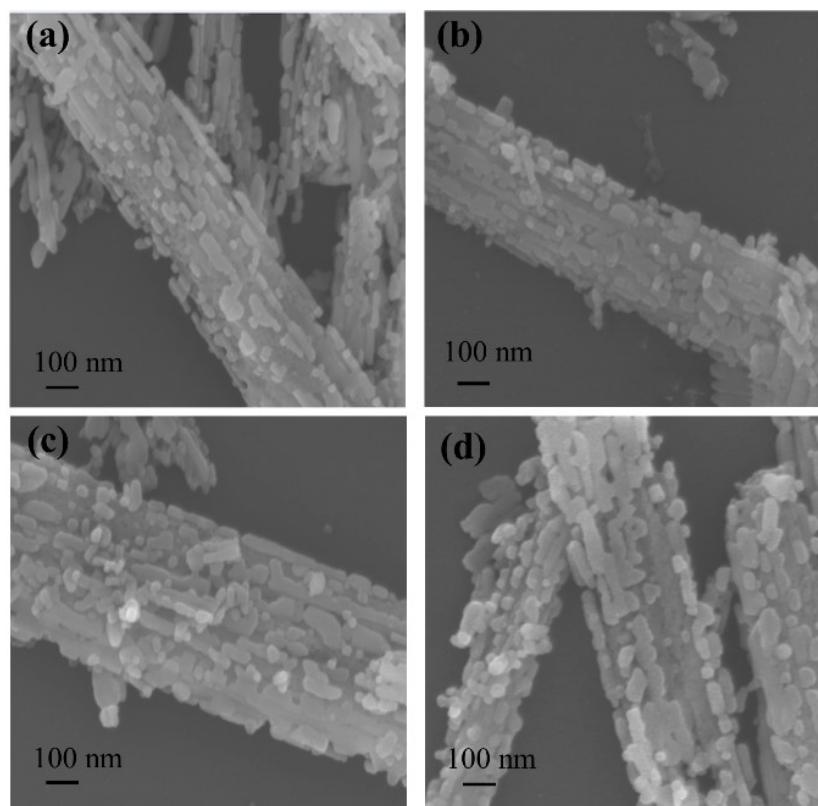


Figure S4. SEM images of (a) TiO₂ nanobelts; (b) 0.5 mol% Au-TiO₂ ;(c) 1.5 mol% Au-TiO₂ samples; (d) 2.5 mol% Au-TiO₂ samples after cycle experiments.

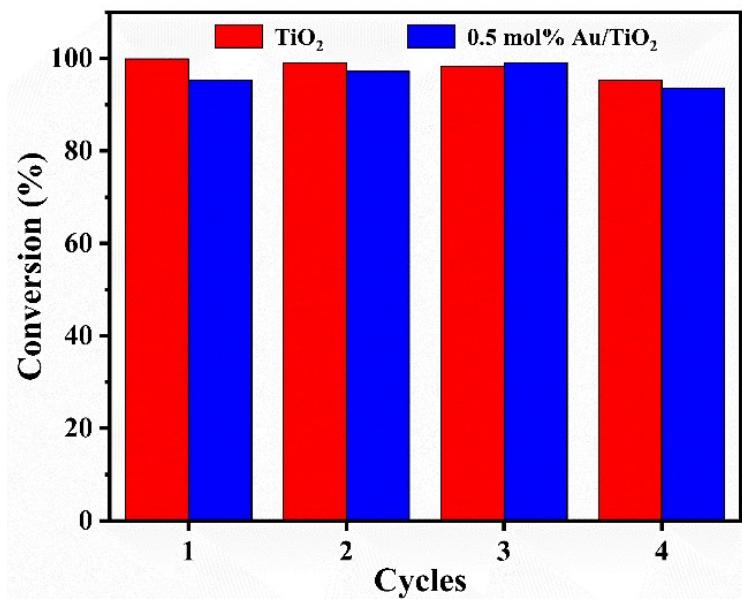


Figure S5. The recyclability of TiO_2 and $0.5 \text{ mol\% Au/TiO}_2$ catalyst in the photocatalytic oxidation of benzylamine at 60°C , 12 h.

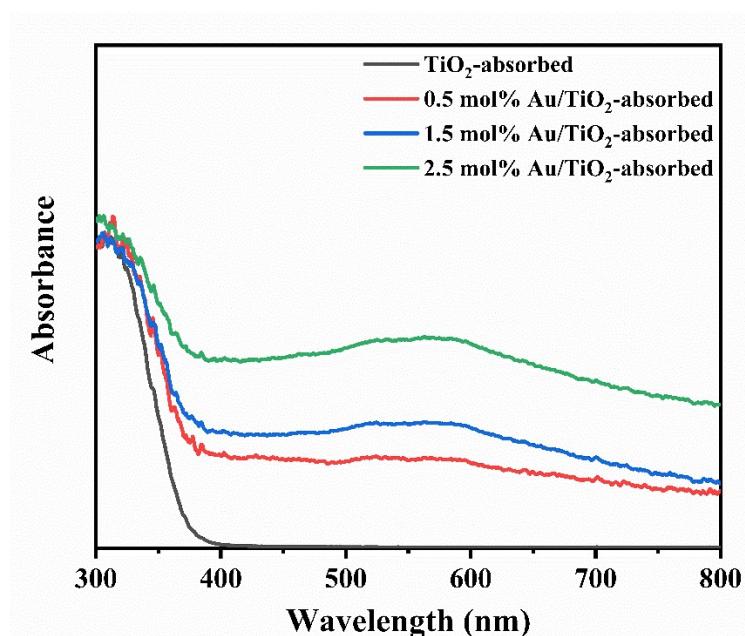


Figure S6. UV-vis adsorption spectra of the corresponding benzylamine-adsorbed catalyst.

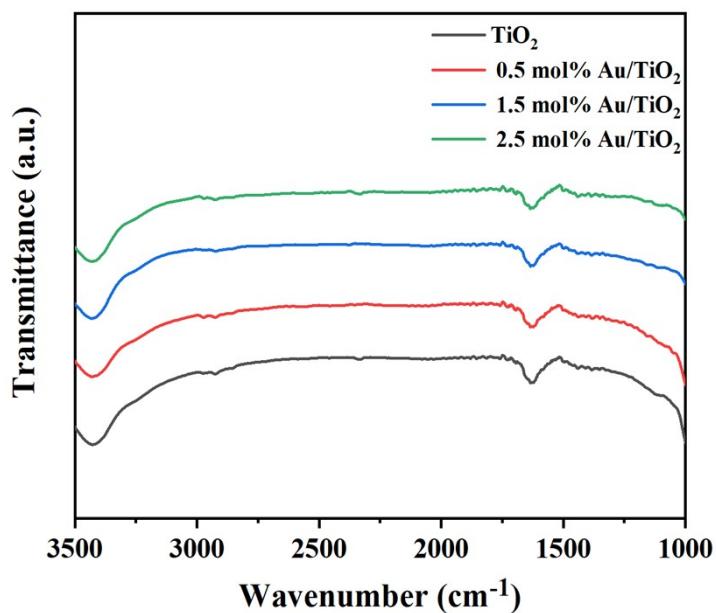


Figure S7. FT-IR image of TiO₂ and 0.5, 1.5, 2.5 mol% Au-TiO₂ nanobelts.

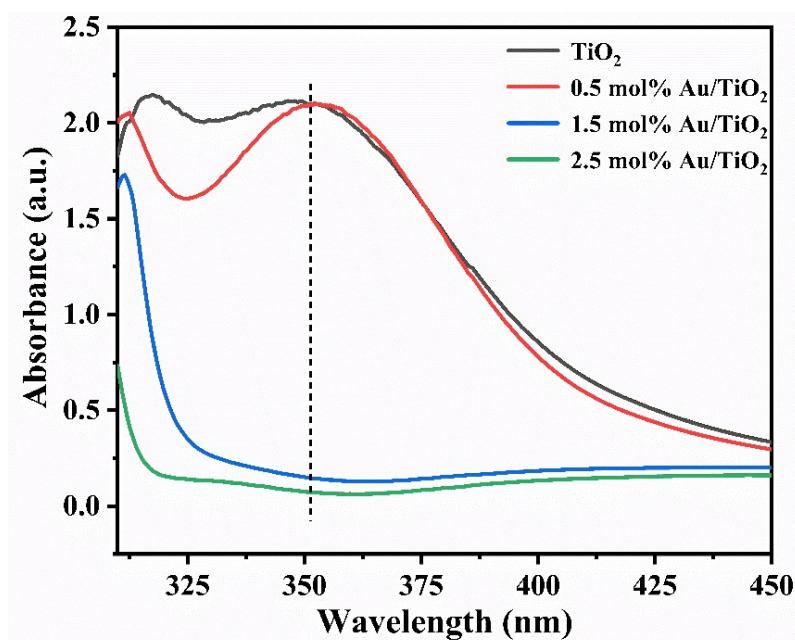


Figure S8. The corresponding UV-vis absorption spectra for detecting H₂O₂ produced in the reaction suspensions for as-prepared samples.

Table S1. ICP tests of Au/TiO₂ catalyst

Sample name	Weight percent of Au	Actual test value	Theoretical value
0.5 mol% Au/TiO ₂	1.50wt%	0.6 mol%	0.5 mol%
1.5 mol% Au/TiO ₂	4.36wt%	1.7 mol%	1.5 mol%
2.5 mol% Au/TiO ₂	7.39wt%	3.0 mol%	2.5 mol%

Table S2. XPS peak details for different catalysts

Catalyst	Ti 2p (eV)	Assignments	O 1s (eV)	Assignments	Peak area	Au 4f (eV)	Assignments
TiO ₂	458.9	Ti ⁴⁺ 2p _{3/2}	530.2	Lattice oxygen	54.6%	\	\
	464.6	Ti ⁴⁺ 2p _{1/2}	530.5	Oxygen vacancy	24.1%	83.5	Au 4f _{7/2}
			532.1	Surface absorbed oxygen	21.3%	87.1	Au 4f _{5/2}
0.5 Au/TiO ₂	459.0	Ti ⁴⁺ 2p _{3/2}	530.2	Lattice oxygen	60.0%		
	464.7	Ti ⁴⁺ 2p _{1/2}	530.5	Oxygen vacancy	21.0%		
			532.3	Surface absorbed oxygen	19.0%		
1.5 Au/TiO ₂	459.0	Ti ⁴⁺ 2p _{3/2}	530.2	Lattice oxygen	67.4%	83.5	Au 4f _{7/2}
	464.7	Ti ⁴⁺ 2p _{1/2}	530.5	Oxygen vacancy	17.1%	87.1	Au 4f _{5/2}
			532.2	Surface absorbed oxygen	15.5%		
2.5 Au/TiO ₂	459.0	Ti ⁴⁺ 2p _{3/2}	530.2	Lattice oxygen	70.3%	83.5	Au 4f _{7/2}
	464.7	Ti ⁴⁺ 2p _{1/2}	530.5	Oxygen vacancy	14.5%	87.1	Au 4f _{5/2}
			532.2	Surface absorbed oxygen	15.2%		
TiO ₂ -absorbed	458.7	Ti ⁴⁺ 2p _{3/2}	529.9	Lattice oxygen	30.5%	\	\
	464.4	Ti ⁴⁺ 2p _{1/2}	532.0	Surface absorbed oxygen	69.5%		
0.5 Au/TiO ₂ -absorbed	458.7	Ti ⁴⁺ 2p _{3/2}	529.8	Lattice oxygen	53.6%	83.8	Au 4f _{7/2}
	464.3	Ti ⁴⁺ 2p _{1/2}	530.3	Oxygen vacancy	22.2%	87.4	Au 4f _{5/2}
			532.1	Surface absorbed oxygen	24.2%		

Table S3. Activity of TiO₂ and M-Au/TiO₂ Catalysts in oxidation of Benzyl Alcohol into Benzaldehyde

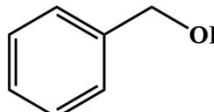
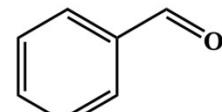
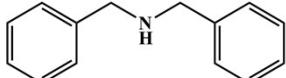
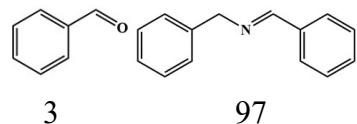
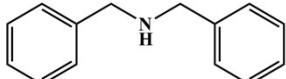
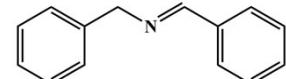
	Visible light (>420 nm), 1 atm O ₂ , 20 h 10 mg catalyst, 0.25 mmol benzyl alcohol, 5 ml CH ₃ CN			
<hr/>				
Temperatur	0.5 mol% Au-TiO ₂	1.5 mol% Au-TiO ₂	2.5 mol% Au-TiO ₂	
e	TiO ₂	TiO ₂	TiO ₂	
20 °C	9.2	16.4	19.2	22.6
60 °C	20.1	25.2	29.3	32.8

Table S4. Aerobic oxidation of secondary benzylic amines by TiO₂ and 0.5 mol% Au/TiO₂ at 25 °C, 12h visible-light irradiation.

Catalyst	Substrate	Conversion (%)	Product selectivity (%)
TiO ₂		42.4	 3 97
0.5 mol% Au/TiO ₂		23.6	 100