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d)



Figure S1 IR spectra of dicarbonyl ruthenium(II) complexes a) 6, b) 7, c) 8, d) 9, e) 10.







Figure S2 <sup>1</sup>H NMR spectra of dicarbonyl ruthenium(II) complexes a) 6, b) 7, c) 8, d) 9, e) 10.



Figure S3 <sup>13</sup>C NMR spectra of dicarbonyl ruthenium(II) complexes a) 6, b) 7.



Figure S4 Hammett plots of the chemical shift of a) NH and b) axial proton of  $CH_2$  group as a function of  $\sigma$  and  $\sigma^+$ , respectively.

Data	7	10
CCDC number	2085341	2085344
Empirical formula	$C_{18}H_{15}CI_2N_3O_4Ru$	$C_{17}H_{15}Cl_2N_3O_2Ru$
Formula weight (g·mol <sup>−1</sup> )	509.30	465.29
Temperature (K)	100.03(17)	100.01(2)
Radiation, $\lambda$ (Å)	Cu-K <sub>α</sub> , 0.0395	Cu-K/ <sub>α</sub> 0.71073
Crystal size (mm <sup>3</sup> )	$0.04 \times 0.09 \times 0.18$	0.04×0.11×0.29
Crystal color, habit	Orange block	Clear yellow block
Crystal system	Monoclinic	Monoclinic
Space group	P21/n	P21/n
<i>a</i> (Å)	7.86247(9)	13.75386(13)
b (Å)	11.42684(13)	8.76188(7)
<i>c</i> (Å)	21.9614(2)	15.45738(15)
eta (°)	92.922(2)	101.9067
Volume (ų)	1970.52	1822.69
Z	4	4
$ ho_{ m calc}$ (g·cm <sup>-3</sup> )	1.717	1.696
F(000)	1016	928
hetarange (°)	4.0530 - 77.0920	2.9217 – 77.3646
Reflections collected	16184	3806
Parameters / restraints	254/0	227 / 0
<sup>i</sup> GooF on <i>F</i> <sup>2</sup>	1.11870	1.07429
<sup><i>ii</i></sup> R <sub>1</sub> [I>2σ(I)]	0.0269	0.028725
wR² (all data) <sup>iii</sup>	0.07138	0.082546
Max./min. residual electron density (e·Å <sup>-3</sup> )	0.448/-0.900	0.470 / -1.099

Table S1 Single-crystal X-ray diffraction data and structure refinements of compound 7 and 10.

<sup>i</sup> Least squares goodness\_of\_fit

<sup>ii</sup>R\_factor\_gt. (For each reflection class, the residual factor for significantly intense reflections <sup>iii</sup>wR\_factor\_ref. (For each reflection class, the weighed residual factors for all reflections included in the refinement)

	7	10
Ru–Cl1	2.369(6)	2.403(7)
Ru–Cl2	2.402(6)	2.389(6)
Ru-N4_1	2.102(2)	2.099(2)
Ru-N1_1	2.190(2)	2.192(2)
Ru-C1	1.868(3)	1.891(3)
Ru–C2	1.893(3)	1.872(3)
C1-01	1.131(3)	1.130(3)
C2-02	1.340(4)	1.133(3)
C1-Ru-Cl1	91.39(8)	89.86(8)
C1–Ru–Cl2	94.49(8)	93.72(8)
C1-Ru-C2	87.30(1)	88.20(1)
C1-Ru-N4_1	98.93(9)	99.61(9)
C1-Ru-N1_1	174.71(9)	173.86(9)
Cl1-Ru-Cl2	172.42(2)	174.48(2)
Cl1-Ru-C2	91.16(9)	88.57(7)
Cl1-Ru-N4_1	86.18(5)	86.44(5)
Cl1-Ru-N1_1	89.12(5)	91.16(5)
Cl2-Ru-C2	94.05(9)	95.74(8)
Cl2-Ru-N4_1	88.03(5)	89.63(5)
Cl2-Ru-N1_1	84.56(5)	84.10(5)
C2-Ru-N4_1	173.30(1)	96.64(9)
C2-Ru-N1_1	98.00(1)	172.23(9)
N4_1-Ru-N1_1	75.85(7)	75.60(6)

 Table S2 Selected experimental bond lengths (Å) and angles (°) of 7 and 10.



**Figure S5** an infinite chain structure of **10** along the axis "*c*" *via* the conventional hydrogen bond interaction.



Figure S6 Electronic absorption spectra of a) 7 and b) 9 in different solvents.



Figure S7 Local minmum structures of 6–10 obtained at B3LYP/LNAL2DZ level of theory.



Figure S8 TD-DFT calculated spectra of complexes 6–10 in the range of 265–415 nm.

Table S3 Com	nputed excitation of compounds 6–1	energies (e 1 <b>0</b>	V), electronic transition configurations and oscillator
Energy	Wavelength	.0.	
$(cm^{-1})$	(nm)	f	Major contributions
	(1111)		
• 6			
23860	419	0.0007	HOMO→LUMO (78%)
24749	404	0.0018	HOMO−1→LUMO (78%)
31713	315	0.0002	HOMO→LUMO+1 (71%)
32152	311	0.0042	HOMO−2→LUMO (69%)
32665	306	0.0484	HOMO–4→LUMO (23%), HOMO–3→LUMO (57%)
33771	296	0.0098	HOMO→L+2 (54%)
38077	263	0.0148	HOMO→L+3 (47%)
39668	252	0.0002	HOMO→L+5 (29%)
39922	250	0.0579	HOMO–4→LUMO+2 (49%), HOMO–3→LUMO+2 (25%)
• 7			
23898	418	0.0007	HOMO→LUMO (74%)
24766	404	0.0019	HOMO-2→LUMO (70%)
31101	322	0.0528	HOMO−1→LUMO (81%)
31795	315	0.0003	HOMO→LUMO+1 (67%)
32221	310	0.0042	HOMO-3→LUMO (68%)
33738	296	0.0069	HOMO→LUMO+2 (59%)
35827	279	0.0006	HOMO→LUMO+4 (54%)
38023	263	0.0126	HOMO→LUMO+3 (62%)
40321	248	0.0573	HOMO–4→LUMO+3 (25%), HOMO–2→LUMO+5 (25%)
• 8			
236623	423	0.0002	HOMO→LUMO (83%)
24528	408	0.001	HOMO–1→LUMO (85%)
301436	332	0.0015	HOMO→LUMO+2 (69%)
31331	319	0.0089	HOMO→LUMO+1 (79%)
31980	313	0.0004	HOMO–1→LUMO+1 (25%), HOMO–1→LUMO+2 (60%)
32488	308	0.0009	HOMO–1→LUMO+1 (65%), HOMO–1→LUMO+2 (26%)
33297	300	0.0025	HOMO→LUMO+3 (39%)
34493	289	0.0076	HOMO–5→LUMO (34%), HOMO–2→LUMO (43%)
39643	252	0.0147	HOMO-2→LUMO+2 (20%)
• 9			
23721	422	0.0007	HOMO <del>∡</del> LUMO (59%)
24634	406	0.0018	HOMO–1→LUMO (61%)
31186	321	0.0004	HOMO→L+2 (48%)

31692	316	0.004	HOMO−2→LUMO (57%)
33120	302	0.003	HOMO–3→LUMO (73%)
33207	301	0.001	HOMO−1→L+2 (48%)
34098	293	0.0601	HOMO–4→LUMO (44%), HOMO→LUMO+3 (29%)
38360	261	0.387	HOMO–4→LUMO+1 (72%)
• 10			
23887	419	0.0007	HOMO→LUMO (58%)
24754	404	0.0018	HOMO−1→LUMO (62%)
31299	319	0.0003	HOMO→LUMO+2 (42%)
32033	312	0.0035	HOMO−2→LUMO (57%)
33059	302	0.0002	HOMO−1→LUMO+1 (37%), HOMO−1→LUMO+2 (28%)
33139	302	0.0028	HOMO→LUMO+1 (43%), HOMO→LUMO+2 (27%)
33812	296	0.0223	HOMO→LUMO+3 (43%)
35918	278	0.0021	HOMO→LUMO+5 (60%)
38174	262	0.009	HOMO→LUMO+4 (56%)
39198	255	0.0067	HOMO–3→LUMO+2 (34%), HOMO–1→LUMO+4 (24%)



Figure S9 Selected FMO's orbitals of 6–10 calculated at CAM-B3LYP/LANL2DZ level of theory.



b)







e)

**Figure S10** Electronic absorption changes upon incubation of complexes a) **6** (0.4 mM), b) **7** (0.86 mM), c) **8** (0.70 mM), d) **9** (0.84 mM), and e) **10** (1.2 mM) in DMSO in the dark for 16 h.



**Figure S11** UV/Vis spectral changes of **9** (0.84 mM in DMSO) upon photolysis at 365 nm with increasing illumination time (0–230 min).



**Figure S12** UV/Vis spectral changes of a) **8** (0.70 mM in DMSO) and b) **10** (0.86 mM in DMSO) upon photolysis at 365 nm with increasing illumination time (0–420 min).





b)



**Figure S13** Time-dependent changes in the FTIR spectra of a) **7**, b) **8**, and d) **10** (8 mM in DMSO) upon exposure to light source at 365 nm with increasing illumination time (0–330, 0–355, and 0–355 min, respectively).



Figure S14 Negative mode ESI MS spectra of a) 7 and b) 10.