Electronic Supplementary Material (ESI) for Environmental Science: Processes & Impacts. This journal is © The Royal Society of Chemistry 2022

Supporting Information for:

Crystal lattice defects in nanocrystalline metacinnabar in contaminated streambank soils indicate a role for biogenic sulfides in the formation of mercury sulfide phases.

Faye Koenigsmark¹, Michelle Chiu², Nelson Rivera¹, Alexander Johs², Jeremy Eskelsen², Donovan Leonard², Boakai K. Robertson³, Anna Szynkiewicz⁴, Christopher Derolph², Linduo Zhao², Baohua Gu², Heileen Hsu-Kim¹, Eric M. Pierce²*

¹Civil and Environmental Engineering, Box 90287, Duke University, Durham, North Carolina 27708, USA

²Environmental Sciences Division, Oak Ridge National Laboratory, Oak Ridge, TN 37831, USA

³Department of Biological Sciences, Alabama State University, Montgomery, AL 36104, USA

⁴Department of Earth and Planetary Sciences, University of Tennessee at Knoxville, Knoxville, TN 37996,

USA

Keywords: nanocrystalline metacinnabar, spectroscopy, mercury, speciation

*Corresponding author: Phone: 865-574-9968, Email: pierceem@ornl.gov

Notice: This manuscript has been authored by UT-Battelle, LLC, under contract DE-AC05-00OR22725 with the US Department of Energy (DOE). The US government retains and the publisher, by accepting the article for publication, acknowledges that the US government retains a nonexclusive, paid-up, irrevocable, worldwide license to publish or reproduce the published form of this manuscript, or allow others to do so, for US government purposes. DOE will provide public access to these results of federally sponsored research in accordance with the DOE Public Access Plan (https://energy.gov/downloads/doe-public-access-plan).

Table of Contents

Section 1: Methods	
Site Description Measurement	S3
Fig. SI1. Samples were collected from the historical release deposit (dark-colored layers) exposed streambanks along EFPC. Shown in this figure is sample HRD-43R, located 18.2 k	
upstream of the mouth of EFPC	
Sulfur Sequential Extraction and Isotope Measurement	S3
Bio-HgS Synthesis	S4
Section 2: Elemental Analysis	
Fig. SI2. Average THg concentrations for streambank soils, stream sediments, and floodp soils in EFPC from previous studies	
Section 3: Microscopy and Spectroscopy Analysis	
Fig. SI3. Size distribution of HgS aggregates (N = 45) measured by SEM from samples HRD-40R, -54L, -43R, -8R, -2L, -31L, -12R, and -22R	S6
Fig. SI4. Size distribution of HgS aggregates (N = 275) measured by SEM from one sectior HRD-2L	
Fig. SI5. EDS maps and spectra for HRD-54L streambank soil sample	S8
Fig. SI6. Secondary electron (SE) SEM imaging and multi-element EDS spectrum for a spo analysis on sample HRD-4L.	t
Fig. SI7. Secondary electron (SE) SEM imaging and multi-element EDS spectrum for a spo analysis on sample HRD-4L.	t
Fig. SI8. Secondary electron (SE) SEM imaging and multi-element EDS spectrum for a spo analysis on sample HRD-22L.	
Fig. SI9. Secondary electron (SE) SEM imaging and multi-element EDS spectrum for a spo analysis on sample HRD-43R.	
Fig. SI10. Secondary electron (SE) SEM imaging and multi-element EDS spectrum for a sp analysis on sample HRD-54L.	
Fig. SI11. HgS cluster generated from the fitted CNs in Table 2 with orientation along the crystal plane.	
Table SI1. Reference EXAFS parameters for fits to possible Hg compounds in EFPC soil environment	S15
Section 4: Metagenomic Sequencing and Analysis	
Microbiome Taxa Results	S16
Fig. SI12. Distribution of taxa (phylum level) in EFPC HRD soil	S16
Table SI13. Lineages known to harbor sulfate reducing microorganisms	S17
Section 5: SI References	S18

Section 1: Methods

Site Description



Fig. Sl1. Samples were collected from the historical release deposit (dark-colored layers) of the exposed streambanks along EFPC. Shown in this figure is sample HRD-43R, located 18.2 km upstream of the mouth of EFPC.

Sulfur sequential extraction method was used to extract various oxidation states of sulfur (SO_4^2 ,

Sulfur Sequential Extraction and Isotope Measurement

S⁰, S⁻, S²⁻) from the HRD soil samples selected for determination of the amount and isotope composition of sulfur (δ^{34} S). Five gram aliquots of HRD soil samples were prepared by grinding the samples in an agate mortar and pestle. The elemental sulfur (S⁰) was extracted from the soil using dichloromethane (DCM) in a Soxhlet apparatus held at 40°C for 12 hours. Activated copper granules were added to the sample flask to recover elemental sulfur as CuS. In a separate extraction apparatus, the copper granules were treated with 30 mL solution of hot deoxygenated 6 N HCl while bubbling the solution with N₂ gas, which released H₂S that reacted with AgNO₃ solution resulting in precipitation of Ag₂S.

Following DCM extraction, the residual soil was treated with 30 mL of 6 N HCl in the same way as the copper granules were. In this step, evolved H₂S from acid-volatile sulfur phase (S²⁻) was also precipitated as Ag₂S. Additionally, dissolved iron was removed from the acid leachate by precipitation as

iron oxides following the addition of NaOH pellets to adjust the pH to 9 – 10. After the NaOH treatment and removal of iron oxides, the remaining solution was acidified to pH<3 with 12 N HCl and acid-soluble sulfate (SO_4^{2-}) was precipitated from the remaining solution as $BaSO_4$ by adding 1-2 mL of 10% barium chloride solution. In the last step, the remaining sediment was treated with a mixture of 20 mL in 12 N HCl and 20 mL of 1 M CrCl₂·6H₂O under N₂. This released H₂S from chromium-reducible sulfide phase (S⁻) which was recovered as described above.

The amount (wt.% S) of SO₄²⁻, S⁰, S⁻, S²⁻ phases was calculated based on the air-dried masses of BaSO₄ and Ag₂S relative to the dry mass of the soil sample used for extraction. The S isotope composition (δ^{34} S) of BaSO₄ and Ag₂S was measured using a Costech elemental analyzer (EA) coupled with a Delta Plus XL isotope ratio mass spectrometer in the Stable Isotope Laboratory, University of Tennessee. Approximately 0.5 mg of BaSO₄/Ag₂S was packed into a tin capsule with 1–5 mg of V₂O₅ to allow for a complete combustion of the sample inside the EA. Isotope data are reported with respect to VCDT (Vienna Canon Diablo Troilite). The correction was based on laboratory standards calibrated to the IAEA SO5, SO6 and NBS-127 standards. Analytical reproducibility was better than ±0.2% for standards and duplicate samples.

Bio-HgS Synthesis

Biotic HgS, referred to as bio-HgS, was extracellularly synthesized using sulfate-reducing bacteria, *Desulfovibrio desulfuricans* ND132 (culture collection number DSM 101870)¹. *D. desulfuricans* ND132 was cultured anaerobically at 30 °C in a modified yeast extract medium (25 mL)^{2–4}, supplemented with 40 mM pyruvate and 20 mM Na₂SO₄ as the respective electron donor and electron acceptor. After 3 days incubation, the measured sulfide concentration was ~650 mg/L in the reactor. Sulfide concentrations were measured using colorimetric methods (or Hach Method 8131, Hach Inc., Loveland, CO). At this time, 0.05 mM anoxic HgCl₂ was added to form biogenic HgS precipitates. The HgS was then harvested from the reactor, rinsed with deoxygenated water, deposited on a TEM grid, and analyzed.

S4

Section 2: Elemental Analysis



Fig. SI2. Average THg concentrations for streambank soils,^{5,6} stream sediments,^{6,7} and floodplain soils^{6,8–}¹¹ in EFPC reported by previous studies. Purple (X) denote total Hg concentrations for streambank soils in this study.

Section 3: Microscopy and Spectroscopy



Fig. SI3. Size distribution of HgS aggregates (N = 45) measured by SEM from samples HRD-40R, -54L, - 43R, -8R, -2L, -31L, -12R, and -22R.



Fig. SI4. Size distribution of HgS aggregates (N =275) measured by SEM from one area of streambank sample HRD-2L.



Fig. SI5. Left: EDS maps for streambank soil HRD-54 showing the presence of Al, Hg, Ca, S, and Si. Right: corresponding EDS average horizontal line for each element.



Fig. SI6. Streambank sample HRD-4L. Top: Secondary electron (SE) SEM imaging and EDS mapping indicating collocation of mercury (Hg) and sulfur (S). Bottom: Multi-element EDS spectrum for a spot analysis.



Fig. SI7. Streambank sample HRD-4L. Top: Secondary electron (SE) SEM imaging and EDS mapping indicating collocation of mercury (Hg) and sulfur (S). Bottom: Multi-element EDS spectrum for a spot analysis.



Fig. SI8. Streambank sample HRD-22L. Top: Secondary electron (SE) SEM imaging and EDS mapping indicating collocation of mercury (Hg) and sulfur (S). Bottom: Multi-element EDS spectrum for a spot analysis.



Fig. SI9. Streambank sample HRD-43R. Top: Secondary electron (SE) SEM imaging and EDS mapping indicating collocation of mercury (Hg) and sulfur (S). Bottom: Multi-element EDS spectrum for a spot analysis.



Fig. SI10. Streambank sample HRD-54L. Top: Secondary electron (SE) SEM imaging and EDS mapping indicating collocation of mercury (Hg) and sulfur (S). Bottom: Multi-element EDS spectrum for a spot analysis.



A He

Fig. SI11. A HgS cluster generated from the fitted CNs in Table 2 with orientation along the 111 crystal plane. The grey central Hg atom is surrounded by 6 Hg atoms (orange). The mercury atoms are tetrahedrally coordinated to sulfur (pink). The HgS cluster measures ~1 nm in size. *Images and video generated using CrystalMaker*: a crystal and molecular structures program for Mac and Windows. CrystalMaker Software Ltd, Oxford, England (<u>www.crystalmaker.com</u>)*

Table S11. Reference EXAFS parameters for fits to possible Hg compounds in EFPC soil environment. Where A-B is the absorber-backscatter atom pair; R is the interatomic distance between A and B; and N is the coordination number of backscatter atom.

Phase	A-B	Ν	R (Å)	Ref.
Nano-HgS (192 h old)	Hg-S	3.9 ± 0.4	2.53 ± 0.01	Pham 2014 [12]
Metacinnabar	Hg-S	4	2.54	Pham 2014 [12]
(β -HgS)	Hg-Hg	12	4.14	
	Hg-S	12	4.85	
	Hg-Hg	6	5.85	
Cinnabar ($lpha$ -HgS) ¹²	Hg-S	2	2.37	Pham 2014 [12]
		2	3.06	
		2	3.26	
	Hg-Hg	2	3.76	
		4	4.03	
		6	4.07	
Hg-thiol	Hg-S	2	2.33	Skyllberg 2006 [13]
0	Hg-C(-S)	2.2	3.29	
Hg-organic soil	Hg-S	0.7-2.0	2.33-2.36	Skyllberg 2006 [13]
Mercury Chloride	Hg-Cl	2	2.28	Serrano 2018 [¹⁴]
(HgCl ₂ (s))	Hg-Cl	6	3.36-3.46	
	Hg-Hg	4	4.33-4.41	
	Hg-Cl	6	4.58-4.84	
	Hg-Hg	4	4.86	
Montroydite (HgO)	Hg-O	2	2.05	Serrano 2018 [¹⁴]
,	Hg-O	4	2.83	
	Hg-Hg	12	3.31-3.74	
	Hg-O	6	4.08-4.48	
	Hg-Hg	4	4.83	
Schuetteite	Hg-O	2	2.07,2.12	Serrano 2018 [¹⁴]
(Hg ₃ (SO ₄)O ₂ (s))	Hg-O	4	2.46-2.75	
5 7 2	Hg-S	1	3.38	
	Hg-S	1	3.68	
	Hg-Hg	3	3.49-3.56	
	Hg-Hg	2	3.71,	
	Hg-Hg	1	3.83	

Section 4: Metagenomic Sequencing and Analysis

Soil Microbiome Results

The shotgun metagenome data of the HRD soil sample contained 62,502,761 sequences (14.2 Gbps), 73.5% of these passed dereplication and quality control. The number of sequences containing ribosomal RNA genes was 77,716. The full dataset is available on MG-RAST under sample ID mgs812418.

The MG-RAST functional gene sequence analysis showed that Bacteria were the dominant domain representing 98.4% of sequences, whereas the Archaea and Eukaryota represented 0.69% and 0.71% of total sequences, respectively. The bacterial sequences are distributed across 35 different phyla. The most abundant phyla were the Proteobacteria with 54.6%, the Actinobacteria with 10.8% and the Acidobacteria with 8.5% of total sequences (**Fig. SI12**). Among the Proteobacteria, the most abundant classes were the Alphaproteobacteria with 20.7%, the Betaproteobacteria with 10.2%, the Gammaproteobacteria with 9.1% and the Deltaproteobacteria with 8.79% of total sequences.





Domain	Phylum	Class	Order	Family	Genus
Archaea	Crenarchaeota	Thermoprotei	Thermoproteales	Thermoproteaceae	
Archaea	Euryarchaeota	Archaeoglobi	Archaeoglobales	Archaeoglobaceae	Archaeoglobus
Bacteria	Actinobacteria	Coriobacteriia	Eggerthellales	Eggerthellaceae	
Bacteria	Firmicutes	Clostridia	Clostridiales	Peptococcaceae	
Bacteria	Firmicutes	Clostridia	Thermoanaerobacterales	Thermoanaerobacteraceae	
Bacteria	Firmicutes	Clostridia	Thermoanaerobacterales	Thermodesulfobiaceae	
Bacteria	Firmicutes	Negativicutes	Selenomonadales	Sporomusaceae	
Bacteria	Nitrospirae	Thermodesulfovibrionia	Thermodesulfovibrionales	Thermodesulfovibrionaceae	Thermodesulfovibrie
Bacteria	Thermodesulfobacteria	Thermodesulfobacteria	Thermodesulfobacteriales	Thermodesulfobacteriaceae	
Bacteria	Proteobacteria	Deltaproteobacteria	Desulfarculales	Desulfarculaceae	
Bacteria	Proteobacteria	Deltaproteobacteria	Desulfobacterales	Desulfobacteraceae	
Bacteria	Proteobacteria	Deltaproteobacteria	Desulfobacterales	Desulfobulbaceae	
Bacteria	Proteobacteria	Deltaproteobacteria	Desulfovibrionales	Desulfohalobiaceae	
Bacteria	Proteobacteria	Deltaproteobacteria	Desulfovibrionales	Desulfomicrobiaceae	
Bacteria	Proteobacteria	Deltaproteobacteria	Desulfovibrionales	Desulfonatronumaceae	
Bacteria	Proteobacteria	Deltaproteobacteria	Desulfovibrionales	Desulfovibrionaceae	
Bacteria	Proteobacteria	Deltaproteobacteria	Syntrophobacterales	Syntrophaceae	
Bacteria	Proteobacteria	Deltaproteobacteria	Syntrophobacterales	Syntrophobacteraceae	

Table SI2. Lineages known to harbor sulfate-reducing microorganisms based on the distribution of dissimilatory (bi)sulfite reductase genes (reductive *dsrAB*) as a functional marker. Adapted from Müller et al.¹⁵

Section 5: SI References

- 1. Gilmour, C. C., Soren, A. B., Gionfriddo, C. M., Podar, M., Wall, J. D., Brown, S. D., Michener, J. K., Urriza, M. S. G. & Elias, D. A. Pseudodesulfovibrio mercurii sp. nov., a mercury-methylating bacterium isolated from sediment. *Int. J. Syst. Evol. Microbiol.* **71**, (2019).
- 2. Zhao, L., Chen, H., Lu, X., Lin, H., Christensen, G. A., Pierce, E. M. & Gu, B. Contrasting Effects of Dissolved Organic Matter on Mercury Methylation by Geobacter sulfurreducens PCA and Desulfovibrio desulfuricans ND132. *Environ. Sci. Technol.* **51**, 10468–10475 (2017).
- 3. Liu, Y. R., Lu, X., Zhao, L., An, J., He, J. Z., Pierce, E. M., Johs, A. & Gu, B. Effects of cellular sorption on mercury bioavailability and methylmercury production by desulfovibrio desulfuricans ND132. *Environ. Sci. Technol.* **50**, 13335–13341 (2016).
- Hu, H., Lin, H., Zheng, W., Tomanicek, S. J., Johs, A., Feng, X., Elias, D. A., Liang, L. & Gu, B.
 Oxidation and methylation of dissolved elemental mercury by anaerobic bacteria. *Nat. Geosci.* 6, 751–754 (2013).
- 5. Southworth, G. & Ketelle, R. Sources of Mercury to East Fork Poplar Creek Downstream from the Y-12 National Security Complex : Inventories and Export Rates. Security (2010).
- 6. Watson, D., Bevelhimer, M., Brandt, C., Derolph, C., Brooks, S., Mayes, M., Olsen, T., Dickson, J., Peterson, M. & Ketelle, R. *Evaluation of Lower East Fork Poplar Creek Mercury Sources–Model Update*. (2016). doi:ORNL/SR-2016/503.
- Brooks, S., Eller, V., Dickson, J., Earles, J., Lowe, K., Mehlhorn, T., Olsen, T., Derolph, C., Watson, D., Phillips, D. & Peterson, M. *Mercury Content of Sediments in East Fork Poplar Creek: Current Assessment and Past Trends*. (2017). doi:ORNL/TM-2016/578.
- 8. Han, F. X., Shiyab, S., Chen, J., Su, Y., Monts, D. L., Waggoner, C. A. & Matta, F. B. Extractability and bioavailability of mercury from a mercury sulfide contaminated soil in Oak Ridge, Tennessee, USA. *Water. Air. Soil Pollut.* **194**, 67–75 (2008).
- 9. Poulin, B. A., Aiken, G. R., Nagy, K. L., Manceau, A., Krabbenhoft, D. P. & Ryan, J. N. Mercury transformation and release differs with depth and time in a contaminated riparian soil during simulated flooding. *Geochim. Cosmochim. Acta* **176**, 118–138 (2016).
- 10. Pant, P., Allen, M. & Tansel, B. Mercury contamination in the riparian zones along the East Fork Poplar Creek at Oak Ridge. *Ecotoxicol. Environ. Saf.* **74**, 467–472 (2011).
- 11. Revis, N. W., Osborne, T. R., Holdsworth, G. & Hadden, C. Distribution of mercury species in soil from a mercury-contaminated site. *Water. Air. Soil Pollut.* **45**, 105–113 (1989).
- 12. Pham, A. L. T., Morris, A., Zhang, T., Ticknor, J., Levard, C. & Hsu-Kim, H. Precipitation of nanoscale mercuric sulfides in the presence of natural organic matter: Structural properties, aggregation, and biotransformation. *Geochim. Cosmochim. Acta* **133**, 204–215 (2014).
- 13. Skyllberg, U., Bloom, P. R., Qian, J., Lin, C. M. & Bleam, W. F. Complexation of mercury(II) in soil organic matter: EXAFS evidence for linear two-coordination with reduced sulfur groups. *Environ. Sci. Technol.* **40**, 4174–4180 (2006).
- 14. Serrano, S., Vlassopoulos, D., Bessinger, B. & O'Day, P. A. Immobilization of Hg(II) by coprecipitation in sulfate-cement systems. *Environ. Sci. Technol.* **46**, 6767–6775 (2012).
- 15. Müller, A. L., Kjeldsen, K. U., Rattei, T., Pester, M. & Loy, A. Phylogenetic and environmental diversity of DsrAB-type dissimilatory (bi)sulfite reductases. *ISME J.* **9**, 1152–1165 (2015).