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Supporting Information

Switchable light vs acid-induced transformations of complex framework compounds at room temperature

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1. Proposed mechanisms of reactions.

Based on^[1], it can be suggested that, initially, under irradiation with UV at a wavelength of 365 nm, as a result of the cleavage of the C_3-C_{3a} bond, a biradical I is generated. Then it is isomerized into a biradical II, from which a cyclopropane ring is formed (Scheme S1).



Scheme S1. A proposed mechanism for the forward rearrangement of cyclopentenes 3 into vinylcyclopropanes 2.

In a trifluoroacetic acid medium at room temperature, 2-azabenzo[*a*]cyclopropa[*cd*]azulenes **2a,d,g,k** are converted back to (epiminomethano)cyclopenta[*a*]indenes **3a,d,g,k** in 69-73% yields (**Scheme S2**). We assume that the first step of this transformation is the addition of a proton to the enamine fragment of the molecules of the cyclopropane derivatives **2a,d,g,k**, which leads to the formation of cation **III**, which undergoes transformation into a cation-biradical **IV**. Intermediate **IV** is isomerized to **V**, the recombination of which and subsequent elimination of the proton leads to the target products.



Scheme S2. A proposed mechanism for the reverse transformation of vinylcyclopropanes 2 into cyclopentenes 3 at room temperature.

Further, the proposed mechanisms reported in Scheme S1 and S2 was investigated in detail by using DFT calculations.

2. Computational Details

DFT static calculations have been performed with Gaussian09 and Gaussian16 set of programs^[2] using the B3LYP functional of Becke and Perdew^[3] for the closed shell systems and UB3LYP for the open shell systems. The geometry optimization has been described with the 6-31G(d) basis set and stationary points were characterized using vibrational analysis, and this analysis has been also used to calculate zero-point energies and thermal (enthalpy and entropy) corrections (298.15 K, 1 bar). Improved electronic energies has been obtained from single-point energy calculations using a 6-311G++(d,p) basis set, a solvation contribute (PCM model,^[4] acetone for 3-2 and acetonitrile for 2-3 reactions, respectively) and dispersion corrections^[5] (EmpiricalDispersion=GD3BJ in Gaussian package). These energies added to the 6-31G(d)-level thermal corrections are named ΔG . It is well-known that open-shell systems place additional challenges to DFT, due to: (i) the spin-contamination problem; (ii) numerically instability in the exchange-correlation (XC) kernel; (iii) the single-determinant description of open-shell ground states can soon be inaccurate and energetically instable.^[6] Linear-response time-dependent density functional theory (TDDFT) along with multireference methods (i.e., complete active space selfconsistent field or CASSCF, multireference configuration interaction or MRCI, complete active space with second-order perturbation theory or CASPT2) can provide more accurate results for activation energies. TDDFT, indeed, can partially overcome these well-known limitations with a reduced computational cost with respect to post-Hartree-Fock methods, since low lying triplet states can be obtained as linear combination of one particle excitations. On the other hand, electronic correlation and orbital relaxations are treated in a more accurate way with TDDFT with respect to converging an unrestricted DFT ground state triplet.^[6,7] Thus, the electronic structures of triplet species for the transformation of 3a-2a (Figure S1) and 3s-2s (Figure S4) were corrected using the TDDFT within the linear-response framework.^[8,9,10] These additional analysis on these triplet states were also made possible at TDDFT level relying on the recent development and implementation of analytic expressions for higherorder TDDFT properties using the most recent version of Gaussian electronic structure package G16.C01,^[2] such as the TDDFT Hessians.^[11,12]

3. Supplementary Figures



Figure S1. Gibbs energies (kcal/mol) for 3a-2a light-transformation obtained by DFT and TDDFTcalculations.



Figure S2. Gibbs energies (kcal/mol) for reaction path acid-catalyzed 2a-3a obtained by DFT calculations.



Figure S3. Spin density (difference between α and β densities, green and blue as positive and negative, respectively) iso-surface of **TS2a**. Generalized spin density corresponding to the excited state energy calculated at TD-B3LYP/6-311++G(d,p) (implicit solvent) was employed.



Figure S4. Gibbs energies (kcal/mol) for 3s-2s light-transformation obtained by DFT and TDDFTcalculations.



Figure S5. Spin density (difference between α and β densities, green and blue as positive and negative, respectively) iso-surface of TS4' reported in Figure S2. Spin density corresponding to the energy calculated for unrestricted triplet B3LYP/6-311++G(d,p) (implicit solvent) was employed.

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5. Appendix

- Copies of ¹H and ¹³C NMR spectra



1-[5,6-Dimethoxy-3a,10-dimethyl-1-phenyl-3,3a,8,8a-tetrahydro-3,8-(epiminomethano)cyclopenta[a]inden-2-yl]ethanone (3b)

¹H NMR spectra for 3b.











S14

¹³C NMR spectra for 3b.















¹³C NMR spectra for 3h.











¹H NMR spectra for 3k.

MeO

¹H NMR spectra for 3k.







¹H NMR spectra for 3k.

N-Me ∕_Ph

MeO



¹³C NMR spectra for 3k.



¹³C NMR spectra for 3k.







¹³C NMR spectra for 3k.

1-(6,7-Dimethoxy-2,4b-dimethyl-4a-phenyl-1,2,4a,4b,8b,8c-hexahydro-2-azabenzo[a]cyclopropa[cd]azulen-4-yl)ethanone (2b)







Ó



¹H NMR spectra for 2b.





¹H NMR spectra for 2b.









Ме



Methyl 4b-ethyl-6,7-dimethoxy-2-methyl-4a-phenyl-1,2,4a,4b,8b,8c-hexahydro-2-azabenzo[a]cyclopropa[cd]azulene-4-carboxylate (2c)




¹H NMR spectra for 2c.











Methyl 6,7-dimethoxy-2-methyl-4b-(2-methylpropyl)-4a-phenyl-1,2,4a,4b,8b,8c-hexahydro-2-azabenzo[a]cyclopropa[cd]azulene-4-carboxylate (2e)

















¹H NMR spectra for 2f.







¹³C NMR spectra for 2f.















∕−128.58 - ---127.46 1.0 0.9 0.8 Normalized Intensity -131.23 125.35 -56.06 42.51 -38.00 -35.04 0.4 -55.91 --52.57 -25.41 _____149.02 __147.36 0.3 -77.21 77.00 76.79 -47.14 -133.95 -44.08 194.30 0.2 38.42 113.29 0.1 192 184 176 168 160 152 144 136 128 120 112 104 96 88 80 72 64 56 48 40 32 24 16 8 0 Chemical Shift (ppm)

¹³C NMR spectra for 2h.







Methyl 2-methyl-4a,4b-diphenyl-1,2,4a,4b,8b,8c-hexahydro-2-azabenzo[a]cyclopropa[cd]azulene-4-carboxylate (2i)



¹H NMR spectra for 2i.

Me

¹H NMR spectra for 2i.







¹H NMR spectra for 2i.





¹³C NMR spectra for 2i.









¹³C NMR spectra for 2i.



1-(2-Methyl-4a,4b-diphenyl-1,2,4a,4b,8b,8c-hexahydro-2-azabenzo[a]cyclopropa[cd]azulen-4-yl)ethanone (2j)









¹H NMR spectra for 2j.





¹H NMR spectra for 2j.



¹³C NMR spectra for 2j.





¹³C NMR spectra for 2j.











¹H NMR spectra for 2k.





¹H NMR spectra for 2k.




¹³C NMR spectra for 2k.















¹H NMR spectra for 2|.





¹H NMR spectra for 2|.





¹³C NMR spectra for 2|.





¹³C NMR spectra for 2|.





¹³C NMR spectra for 2|.





Methyl 4b-(4-fluorophenyl)-6,7-dimethoxy-2-methyl-4a-phenyl-1,2,4a,4b,8b,8c-hexahydro-2-azabenzo[a]cyclopropa[cd]azulene-4-carboxylate (2m)















1-[4b-(4-Fluorophenyl)-6,7-dimethoxy-2-methyl-4a-phenyl-1,2,4a,4b,8b,8c-hexahydro-2-azabenzo[a]cyclopropa[cd]azulen-4-yl]ethanone (2n)

















Methyl 6,7-dimethoxy-4b-(4-methoxyphenyl)-2-methyl-4a-phenyl-1,2,4a,4b,8b,8c-hexahydro-2-azabenzo[a]cyclopropa[cd]azulene-4-carboxylate (20)













1-[6,7-Dimethoxy-4b-(4-methoxyphenyl)-2-methyl-4a-phenyl-1,2,4a,4b,8b,8c-hexahydro-2-azabenzo[a]cyclopropa[cd]azulen-4-yl]ethanone (2p)






































Methyl 6,7-dimethoxy-2-methyl-4b-(4-nitrophenyl)-4a-phenyl-1,2,4a,4b,8b,8c-hexahydro-2-azabenzo[a]cyclopropa[cd]azulene-4-carboxylate (2s)





















