

Table S1. Fitting parameters for the kinetic equations that describe Cd (II) adsorption on NBCs

Adsorbent	pseudo-first-order			pseudo-second-order		
	$q_e(\text{mg/g})$	k_1	r^2	$q_e(\text{mg/g})$	$k_2 \times 10^{-3}$	r^2
Pr-NBC+20ppm	6.41	0.0049	0.787	7.22	0.193	0.853
Pd-NBC+20ppm	12.24	0.0034	0.959	14.97	0.239	0.962
Gp-NBC+20ppm	16.57	0.0123	0.944	18.43	0.896	0.971
Pr-NBC+50ppm	13.95	0.0027	0.983	17.86	0.140	0.984
Pd-NBC+50ppm	17.93	0.0026	0.983	24.33	0.101	0.985
Gp-NBC+50ppm	29.64	0.0038	0.993	36.51	0.106	0.993

Table S2. Fitting parameters for the Langmuir and Freundlich isotherms for Cd (II) adsorption by NBCs and pristine biochars

Adsorbent	Langmuir			Freundlich		
	$q_{max}(\text{mg/g})$	$b(\text{L/mg})$	r^2	$K_F(\text{mg}^{1+n}/\text{g/L}^n)$	n	r^2
Pr-NBC	12.68	0.085	0.968	3.02	0.277	0.957
Pd-NBC	24.04	0.357	0.990	9.10	0.249	0.988
Gp-NBC	35.95	0.468	0.973	13.94	0.261	0.962
700-BC	11.98	0.095	0.962	3.43	0.296	0.983
800-BC	15.82	0.110	0.924	3.80	0.325	0.982
900-BC	11.01	0.213	0.881	4.17	0.226	0.997

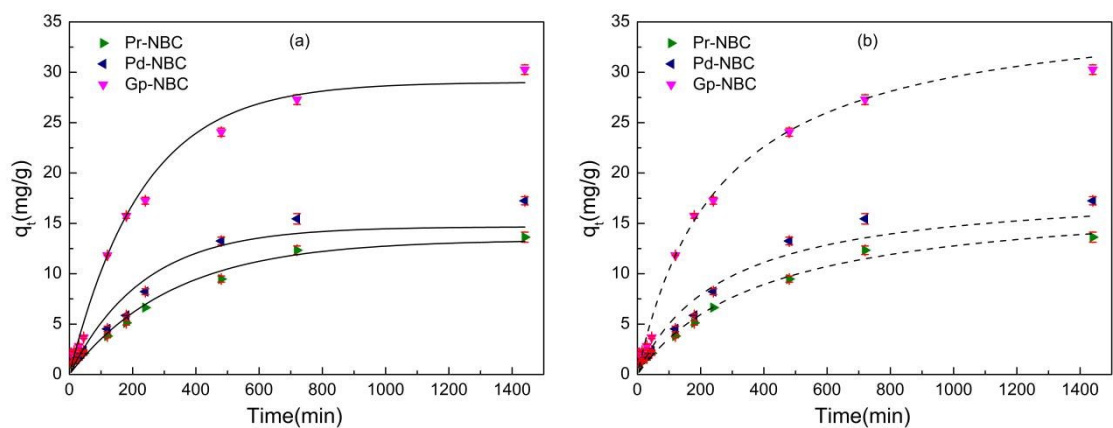


Fig.S1. Adsorption kinetics of 50 mg/L Cd(II) on NBCs (a): pseudo-first-order model; (b) : pseudo-second-order model

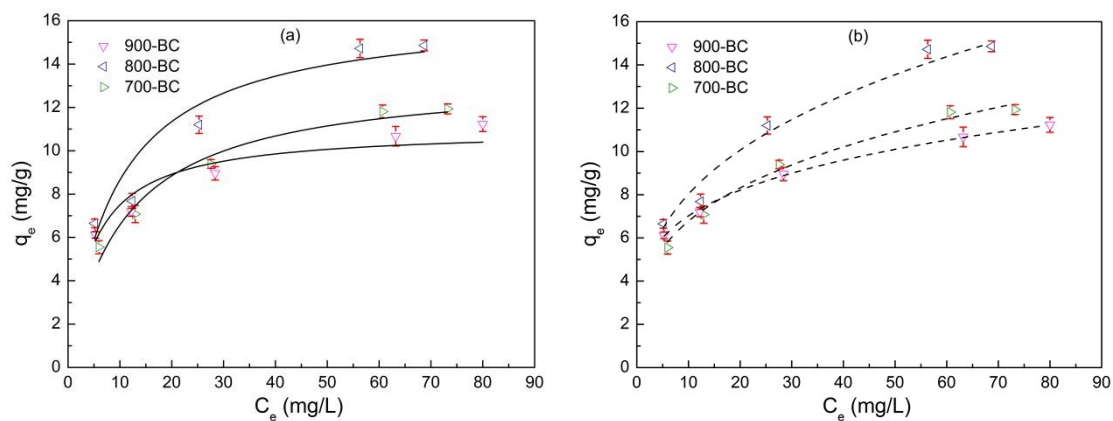


Fig.S2. Adsorption isotherm of Cd(II) on pristine BCs (a): Langmuir model; (b): Freundlich model.

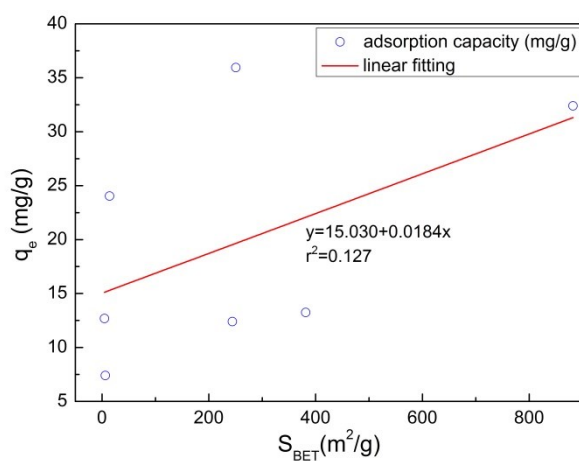


Fig.S3. The correlations of adsorption capacity (q_e , mg/g) with S_{BET} (m^2/g)

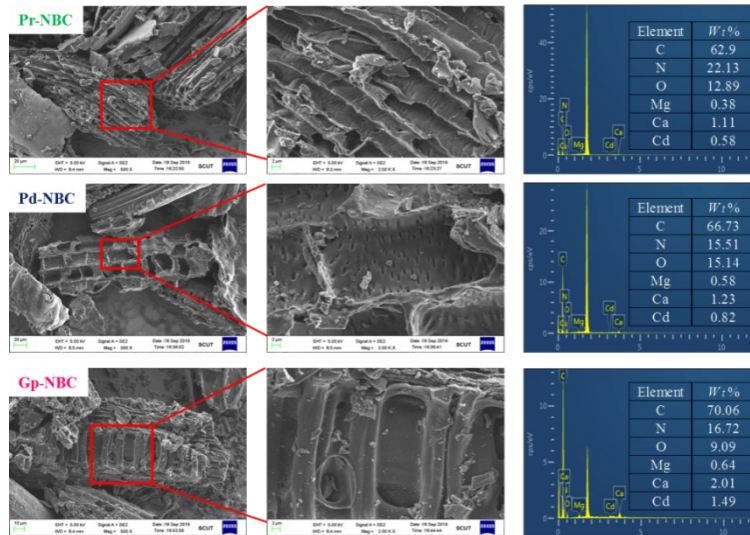


Fig.S4. SEM images and element relative content of NBCs' surface

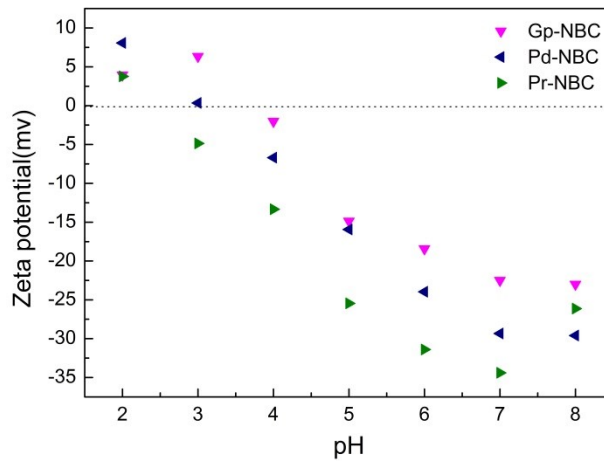


Fig. S5. Zeta potential of NBCs (pH from 2 to 8)

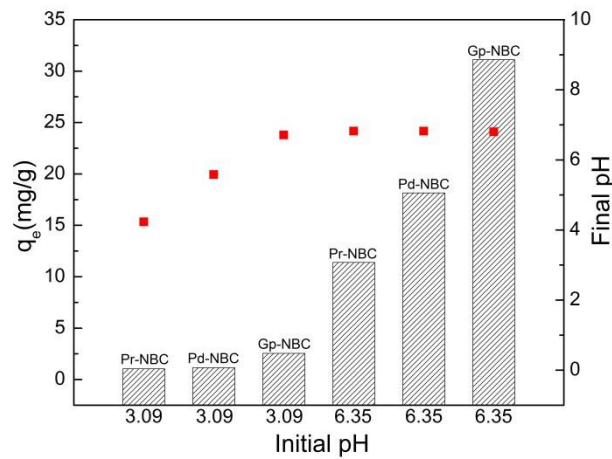


Fig. S6. Influence of initial pH on adsorption (red points represent pH after adsorption, and the columns represent adsorption capacities)

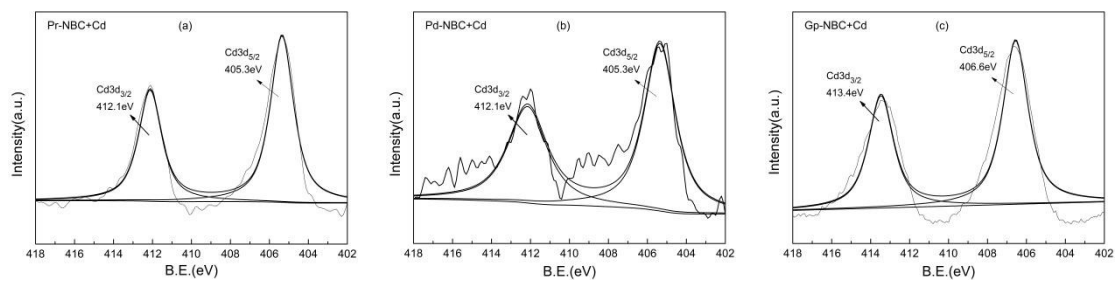


Fig. S7. High resolution XPS spectra of Cd3d on NBCs after adsorption