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Mesoporous thin films onto graphene FETs: nanofiltrated, amplified and extended field-effect sensing

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Table S1

Layer	Material	Model and optical constants	Thickness (nm)
1	Si	Si from CompleteEASE library - $\eta_{632.8 \text{ nm}} = 3.874$;	-
		$k_{632.8 \text{ nm}} = 0.0146$	
2	SiO ₂	SiO ₂ from CompleteEASE library - $\eta_{632.8 \text{ nm}} =$	301.8
		1.4751; $k_{632.8 \text{ nm}} = 0$	
3	APTES	Cauchy - $\eta_{632.8 \text{ nm}} = 1.528$; $k_{632.8 \text{ nm}} = 0$; A = 1.504; B	1.4
		= 0.00962; C = 0	
4	rGO	Cauchy - $\eta_{632.8 \text{ nm}} = 1.8204$; $k_{632.8 \text{ nm}} = 0$; $A = 1.808$;	0.7
		B = 0.0048; C = 0	
5	Mesoporous	EMA Bruggeman with VOID and SiO ₂ - $\eta_{632.8 \text{ nm}} =$	135
	thin film at	1.205; $k_{632.8 \text{ nm}} = 0$; 53.4% SiO ₂	
	4000 rpm		
	- F		



Fig S1. (a) Side view SEM image of a MTF after the LbL deposition of four polyelectrolyte PEI/PSS layers (b) Adhesion test using a scotch-tape for MTF prepared onto a rGO-modified Si substrate. (c) Scanning electron micrograph of GFET modified with MTF made of non-optimized precursor formulation. It can be seen MTF fissures exposing the rGO flakes below the MTF.





Figure S2. Spectroscopic ellipsometric measurements (solid lines) and model fits (dashed lines) for Si wafer substrate modified with rGO (top plot) and silica MTF prepared at 4000 rpm onto rGOmodified Si wafer substrates (bottom plot).



Fig S3. Left: I_{DS} - V_G curves of MTF-GETs in the absence (blue) and the presence (light blue) of 1 mM H₂O₂. Right: Real-time measurements of H₂O₂ detection for a MTF-GFET by using a flow-cell. H₂O₂ concentrations values range from 0 to 1 mM. The arrows indicate the change of the solution to be eluted ($V_G = -400$ mV). All the experiments were carried out using an Ag/AgCl gate, 10 mM KCl and applying $V_{DS} = 100$ mV.



Fig S4. I_{DS} - V_G curves for GFET sensors without a MTF. The I_{DS} - V_G curves show the response before (red) and after (brown) lysozyme adsorption. The experiments were carried out using an Ag/AgCl gate, AcH/AcNa-NaCl buffer (pH = 5, ionic strength = 140 mM) and applying V_{DS} = 100 mV.



Figure S5. Plots of the ionic strength as a function of position (r, z) at 100 mM KCl and pH 7, as predicted by the thermodynamic model.



Figure S6. Plots of the concentration of bound potassium as a function of position (r,z) at different bulk salt concentrations and pH 7. The black dotted lines denote the charged interface between the silica (white region) and aqueous solution. These concentrations are defined as the local excess with respect to the concentration of chloride ions.