

Supplementary Information

**Structural effect of Ni/TiO₂ on CO methanation: improved activity
and enhanced stability**

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Table S1 Textural properties of the TiO₂ support and the Ni/TiO₂ catalysts.

Sample	S _{BET} (m ² /g)	V _{pore} (cm ³ /g)	D _{pore} (nm)
TiO ₂	13.2	0.03	9.1
Ni/TiO ₂ -C	14.1	0.04	10.2
Ni/TiO ₂ -P	15.1	0.03	8.5

The Ni $2p_{3/2}$ spectra of the Ni/TiO₂ catalysts are shown in Fig. S1. The Ni species exists in both elemental state (Ni⁰) centered at 851.8 eV and oxidized state (Ni²⁺) centered at 855.0 eV, along with a satellite located at 860.4 eV.^{S1, S2} The oxidic Ni²⁺ can be assigned to NiO from the air exposure. Compared to the standard binding energy (BE) of Ni⁰ (852.9 eV), the BE of Ni⁰ shifted negatively in catalysts, suggesting an electron transfer from TiO₂ to Ni.^{S1} The Ni $2p_{3/2}$ spectra of Ni/TiO₂-P catalyst shows more intensive peaks, indicating a better dispersion on the surface.^{S3} This is consistent with our conclusion that Ni/TiO₂-P catalyst possesses smaller Ni particle size and higher Ni dispersion.

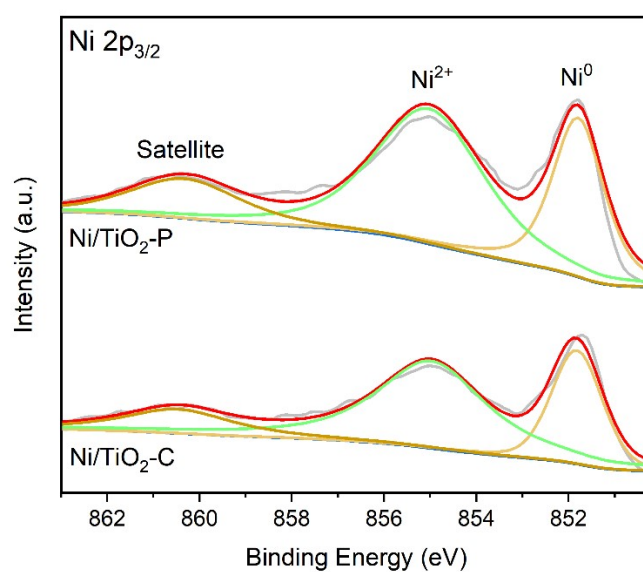


Fig. S1 Ni $2p_{3/2}$ XPS spectra of Ni/TiO₂-P and Ni/TiO₂-C.

The dispersion of Ni nanoparticles (D_{Ni}) was estimated by assuming Ni nanoparticle as a sphere, following the equations: ^{S4}

$$D_{Ni} = \frac{n_s}{n} = \frac{4\pi \cdot R^2 \cdot a_m}{\frac{4\pi}{3} \cdot R^3 \cdot \rho_0 \cdot \frac{N_A}{M}}$$

$$R = \frac{d}{2}$$

Where n_s is the number of Ni atoms on the surface of the sphere; n is the total number of Ni atoms in the sphere; d is the mean diameter of Ni nanoparticles obtained from XRD; a_m is the number of surface Ni atoms per unit m^2 , which is $1.54 \times 10^{19} m^{-2}$ calculated for fcc Ni using the proportions of low index planes fcc (111): (100): (110) = 1: 1: 1;^{S4} ρ_0 is the density of Ni, $8.902 \times 10^6 g \cdot m^{-3}$; N_A is $6.02 \times 10^{23} mol^{-1}$; M is the atomic weight of Ni, $58.69 g \cdot mol^{-1}$.

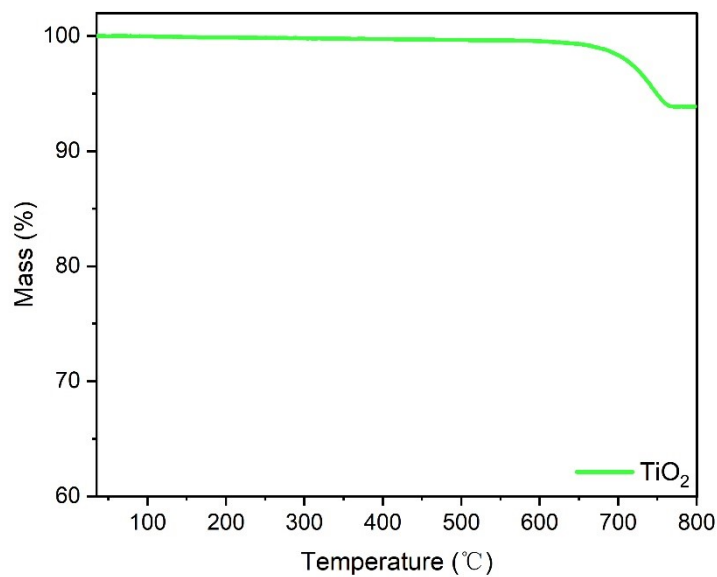


Fig. S2 TG curve of TiO₂.

References

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