

## Supplementary File

### **Role of Metal and Anions in Organo-Metal Halide Perovskites CH<sub>3</sub>NH<sub>3</sub>MX<sub>3</sub> (M: Cu, Zn, Ga, Ge, Sn, Pb; X: Cl, Br, I) on Structural and Optoelectronic Properties for Photovoltaic Applications**

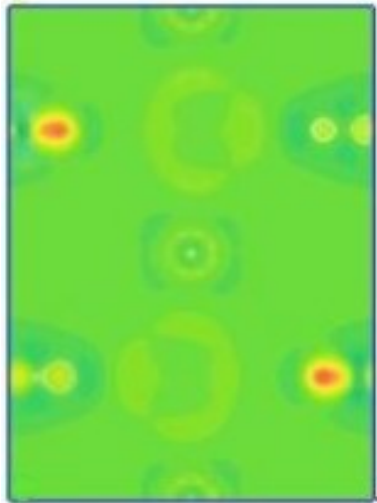
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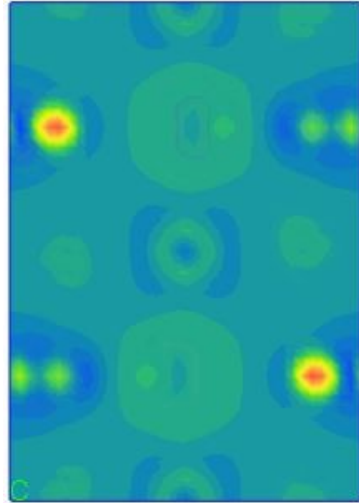
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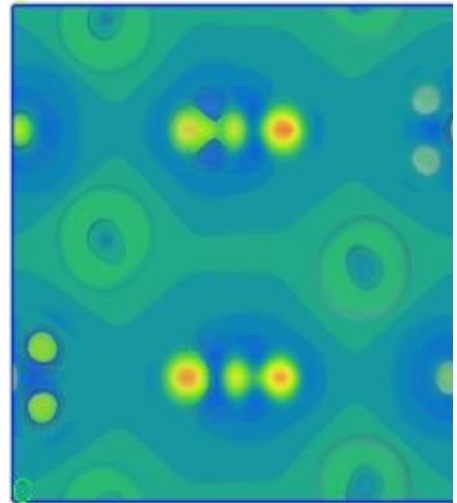
Corresponding author: Maliha Nishat ([malihanishat@pust.ac.bd](mailto:malihanishat@pust.ac.bd))



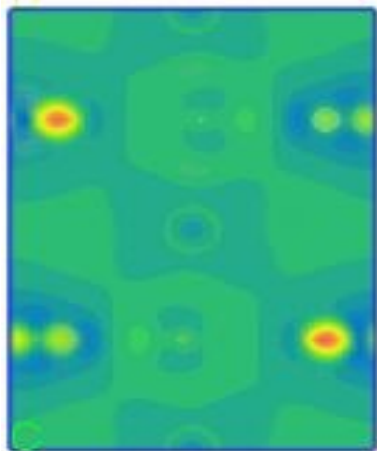
(a)  $\text{CH}_3\text{NH}_3\text{PbI}_3$



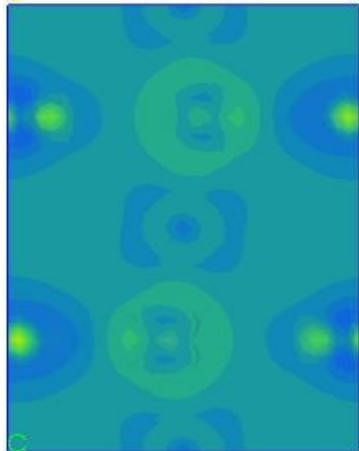
(b)  $\text{CH}_3\text{NH}_3\text{Br}_3$



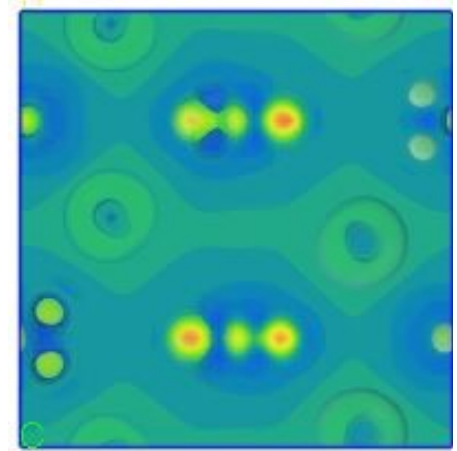
(c)  $\text{CH}_3\text{NH}_3\text{PbCl}_3$



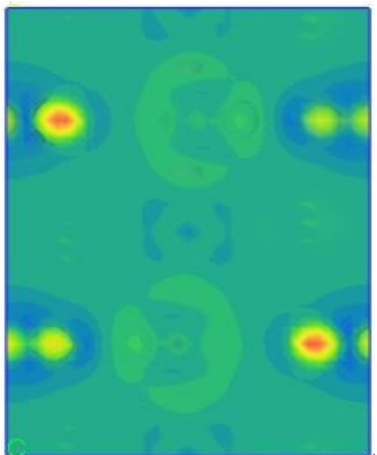
(d)  $\text{CH}_3\text{NH}_3\text{SnI}_3$



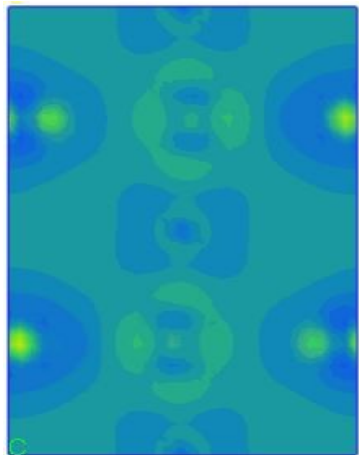
(e)  $\text{CH}_3\text{NH}_3\text{SnBr}_3$



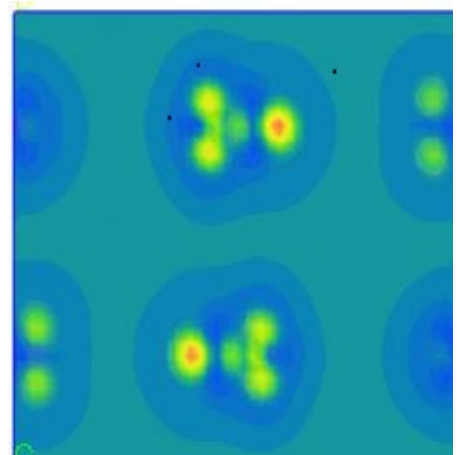
(f)  $\text{CH}_3\text{NH}_3\text{SnCl}_3$



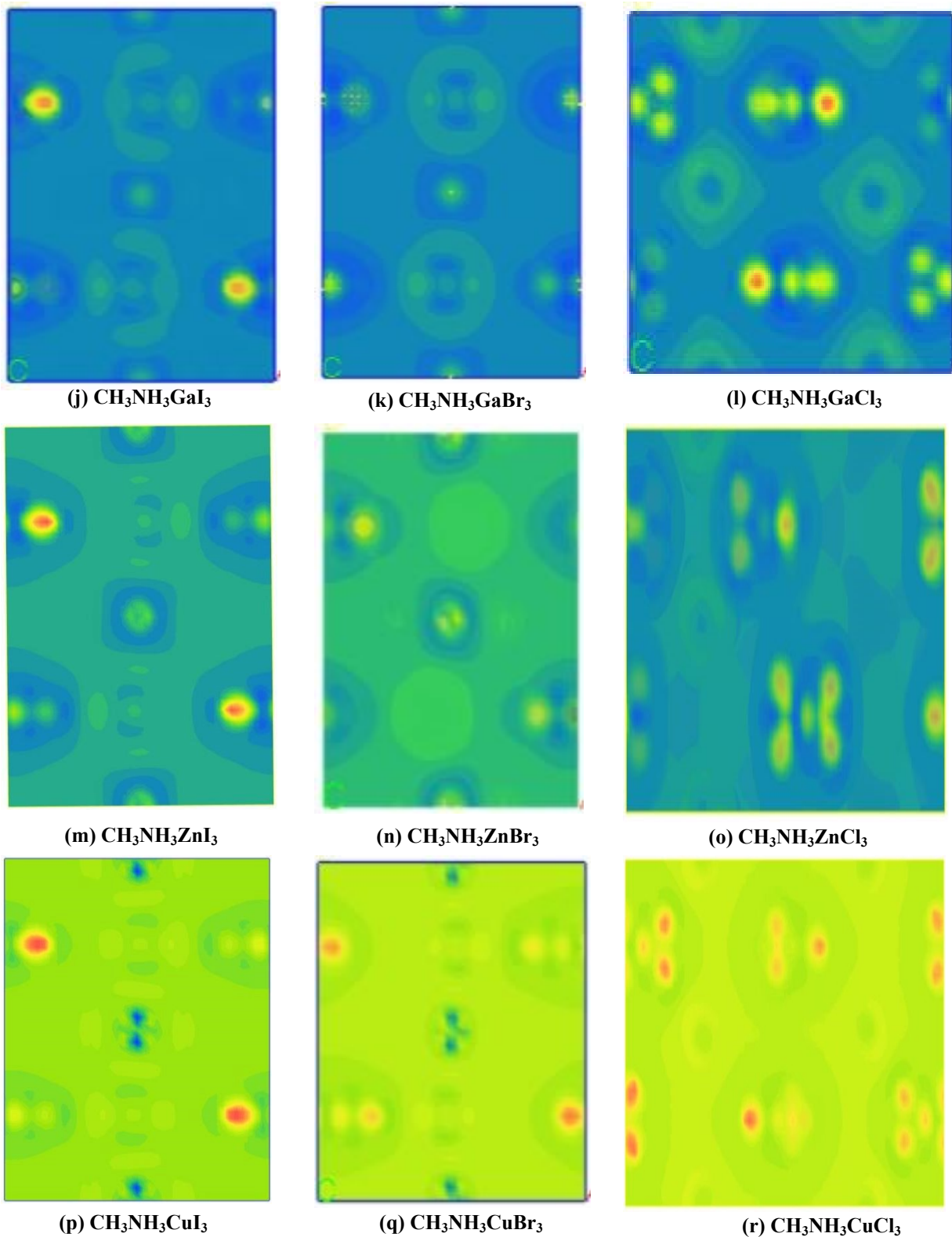
(g)  $\text{CH}_3\text{NH}_3\text{GeI}_3$



(h)  $\text{CH}_3\text{NH}_3\text{GeBr}_3$

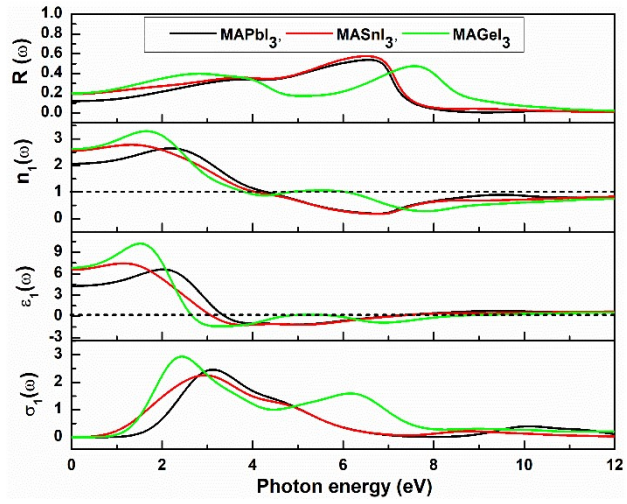


(i)  $\text{CH}_3\text{NH}_3\text{GeCl}_3$

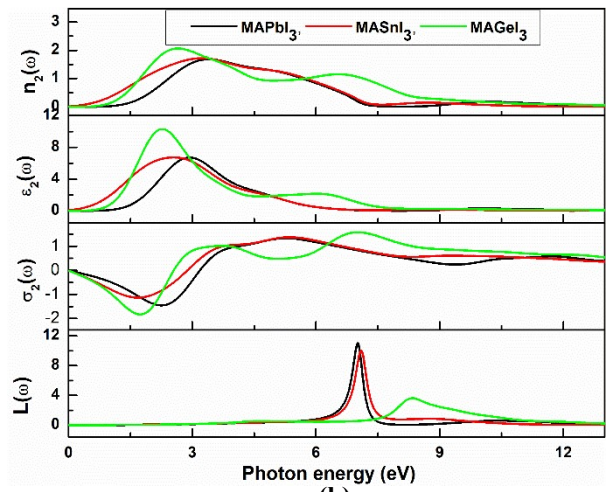


**Fig. S1.** Total electronic charge density for methylammonium metal halide; a field representation. The units are electrons  $\text{\AA}^{-3}$ . In field representation, from blue to orange color symbolize the change of electron density from low to high.

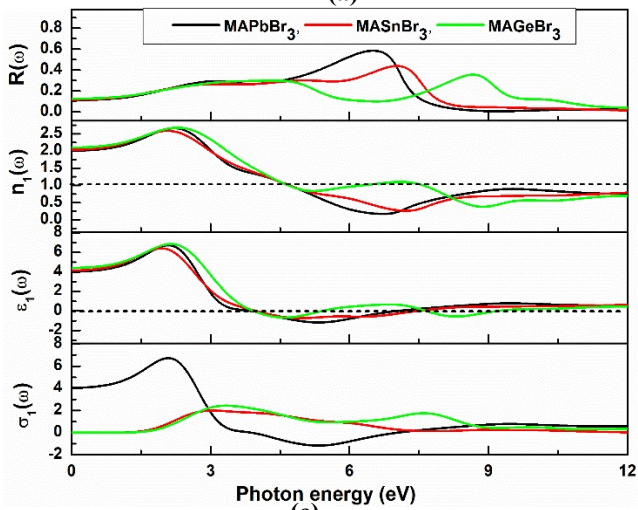




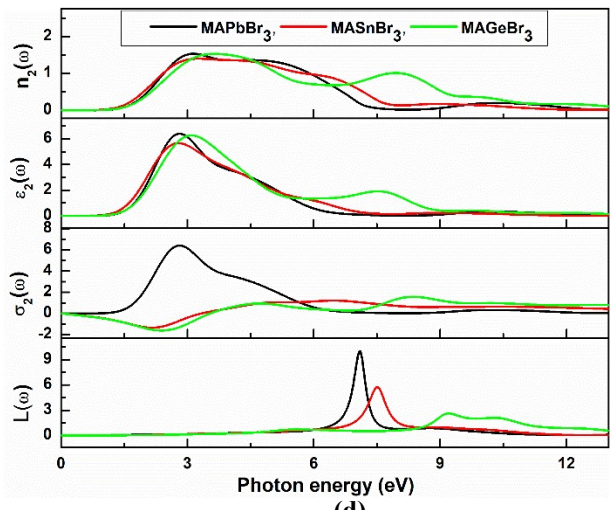
(a)



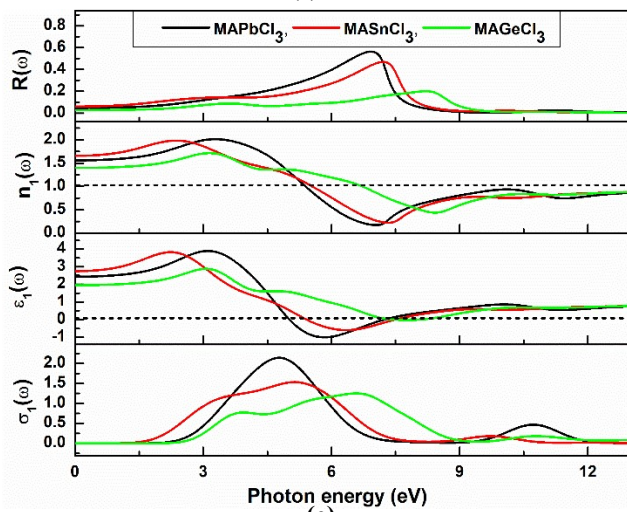
(b)



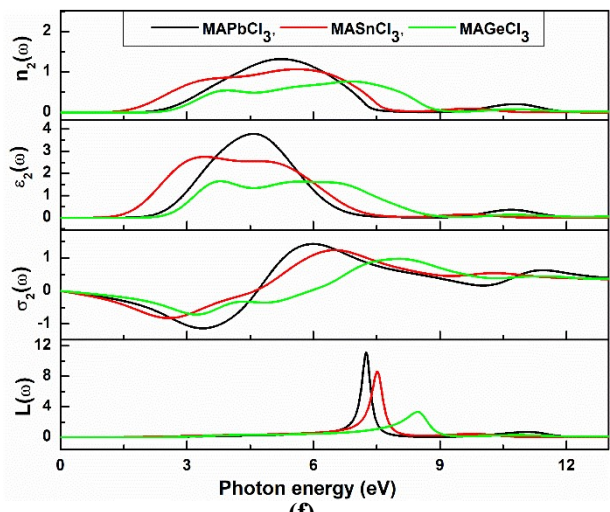
(c)



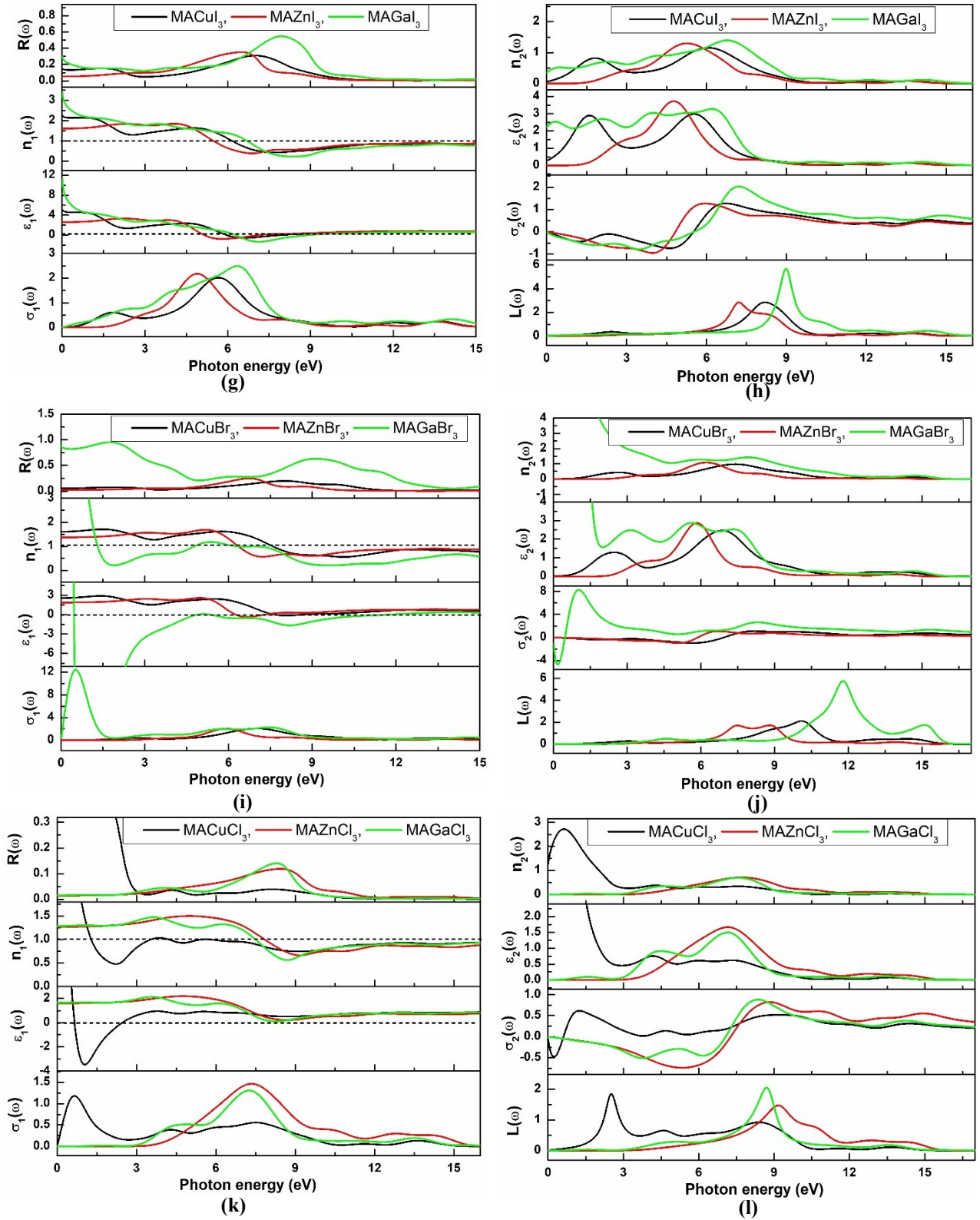
(d)



(e)



(f)



**Fig. S2.** Representation of different optical properties of orthorhombic  $\text{CH}_3\text{NH}_3\text{MX}_3$  (where  $\text{CH}_3\text{NH}_3 = \text{MA}$ ,  $\text{M} = \text{Cu, Zn, Ga, Ge, Sn, Pb}$ ;  $\text{X} = \text{Cl, Br, I}$ ). Reflectivity  $R(\omega)$ , Real part of the refractive index  $n_1(\omega)$ , dielectric constant  $\epsilon_1(\omega)$ , and optical conductivity  $\sigma_1(\omega)$  as a function of energy in eV are illustrated in (a), (c), (e), (g), (i) and (K); Imaginary

part of the refractive index  $n_2(\omega)$ , dielectric constant  $\epsilon_2(\omega)$ , and optical conductivity  $\sigma_2(\omega)$  and energy loss function  $L(\omega)$  are shown in (b), (d), (f), (h), (j) and (l).