

Supporting Information

Designed Synthesis of a sp^2 Carbon-Conjugated Fluorescent Covalent Organic Framework for Selective Detection of Fe^{3+}

Jia You, Qianqian Kong, Cuiling Zhang* and Yuezhong Xian*

Shanghai Engineering Research Center of Molecular Therapeutics and New Drug Development, Department of Chemistry, School of Chemistry and Molecular Engineering, East China Normal University, Shanghai 200241, China.

E-mails: clzhang@chem.ecnu.edu.cn, yzxian@chem.ecnu.edu.cn.

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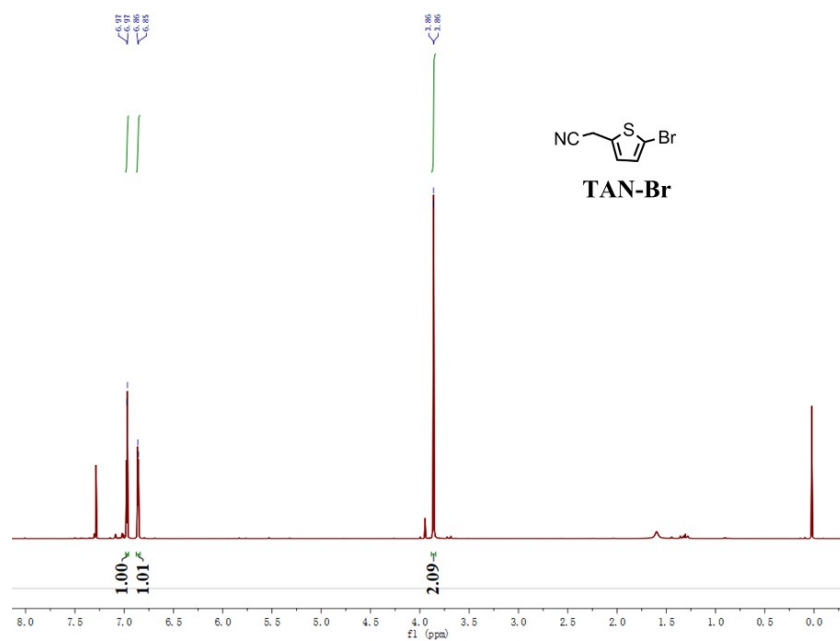


Figure S1. ^1H NMR of TAN-Br.

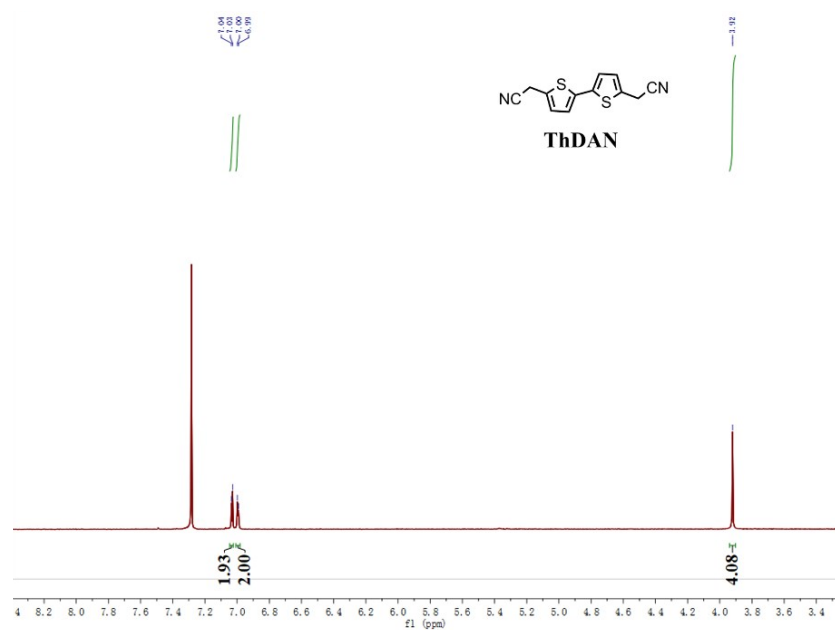


Figure S2. ¹H NMR of ThDAN.

Table S1. Synthesis of the COF_{TFPPy-ThDAN} under different conditions.

No.	Solvent	Ratio	Temperature	Catalyst	Product
1	<i>o</i> -DCB/ <i>n</i> -BuOH	1/1	120 °C	NaOH	Amorphous polymer
2	<i>o</i> -DCB/ <i>n</i> -BuOH	1/1	120 °C	DBU	Amorphous Polymer
3	<i>o</i> -DCB/ <i>n</i> -BuOH	1/1	120 °C	Cs ₂ CO ₃	Amorphous Polymer
4	1,4-dioxane/ Mesitylene	1/1	120 °C	NaOH	Amorphous Polymer
5	1,4-dioxane/ Mesitylene	1/1	120 °C	DBU	Amorphous Polymer
6	1,4-dioxane/ Mesitylene	1/1	120 °C	Cs ₂ CO ₃	Poor crystallinity
7	1,4-dioxane/ Mesitylene	1/1	120 °C	Cs ₂ CO ₃	Amorphous Polymer
8	1,4-dioxane/ Mesitylene	1/1	120 °C	Cs ₂ CO ₃	Medium crystallinity
9	1,4-dioxane/ Mesitylene	1/1	120 °C	Cs ₂ CO ₃	High crystallinity
10	1,4-dioxane/ Mesitylene	1/1	120 °C	Cs ₂ CO ₃	Medium crystallinity

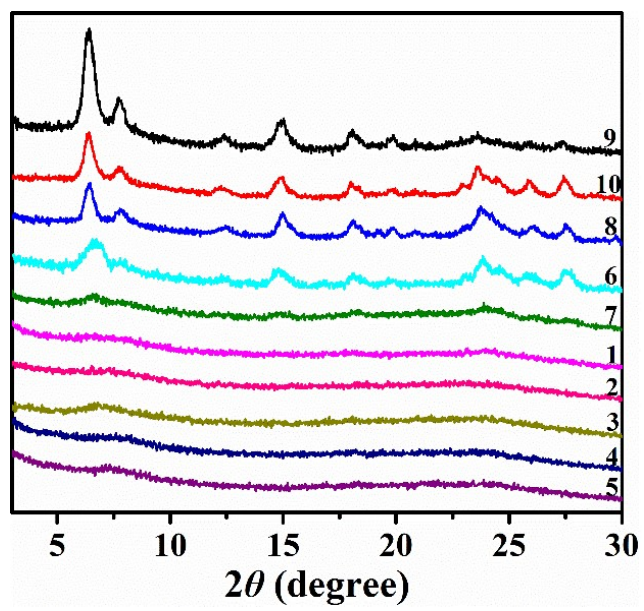


Figure S3. PXRD spectra of products under different synthesis conditions (the numbers correspond to the serial numbers in Table S1).

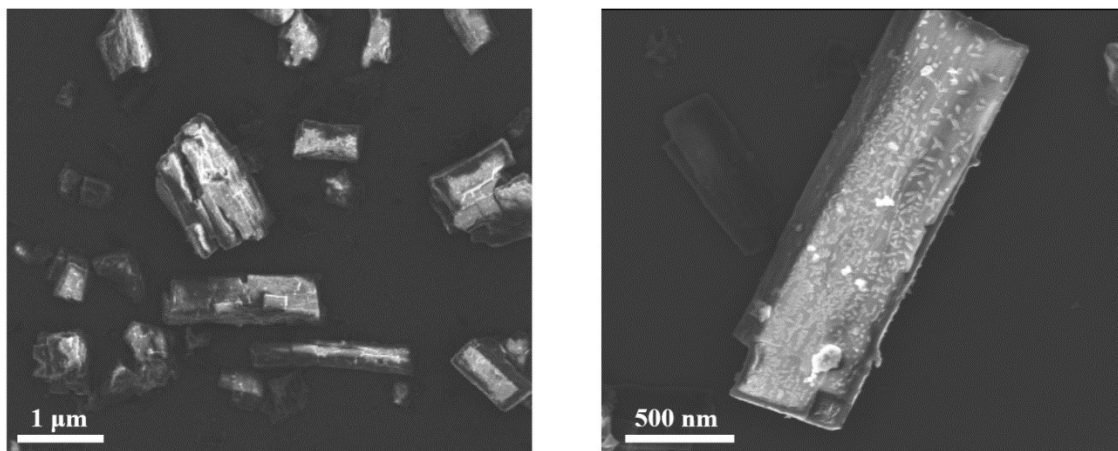


Figure S4. SEM images of COF_{TFPPy-ThDAN}.

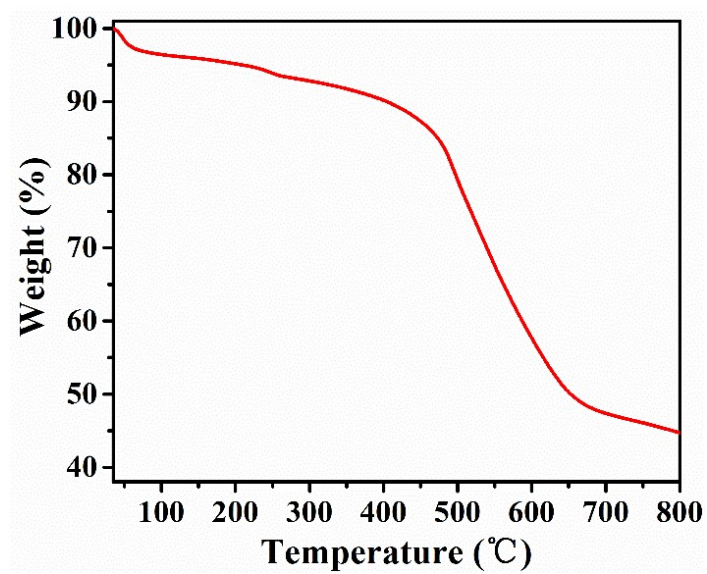


Figure S5. Thermogravimetric analysis curve of COF_{TFPPy-ThDAN}.

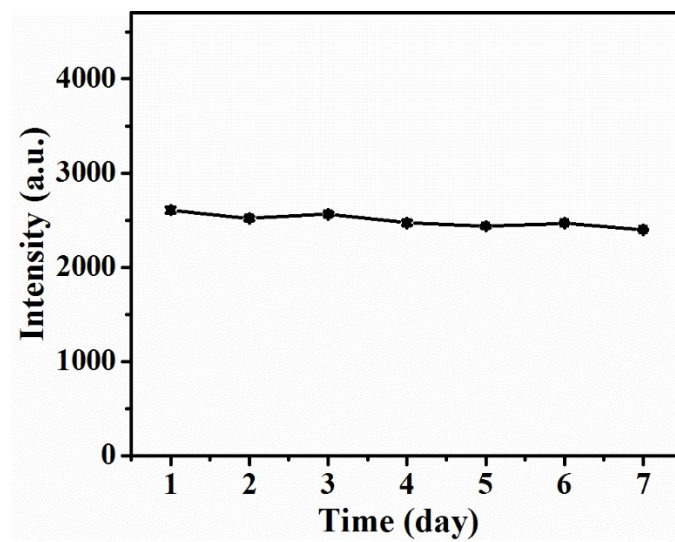


Figure S6. The fluorescence stability of COF_{TFPPy-ThDAN} within 7 days.

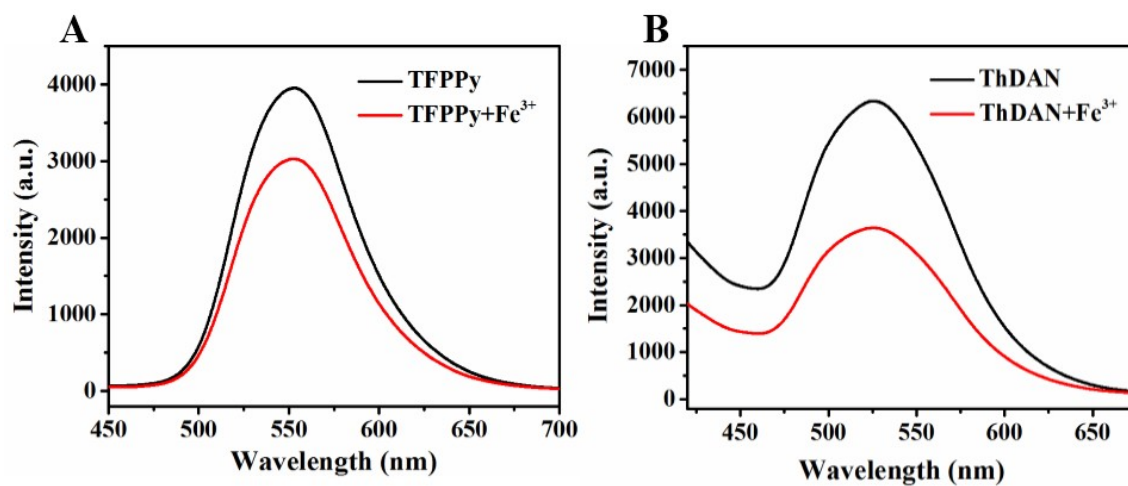


Figure S7. Fluorescence responses of TFPPy and ThDAN towards Fe^{3+} . The concentration of Fe^{3+} is $200 \mu\text{M}$.

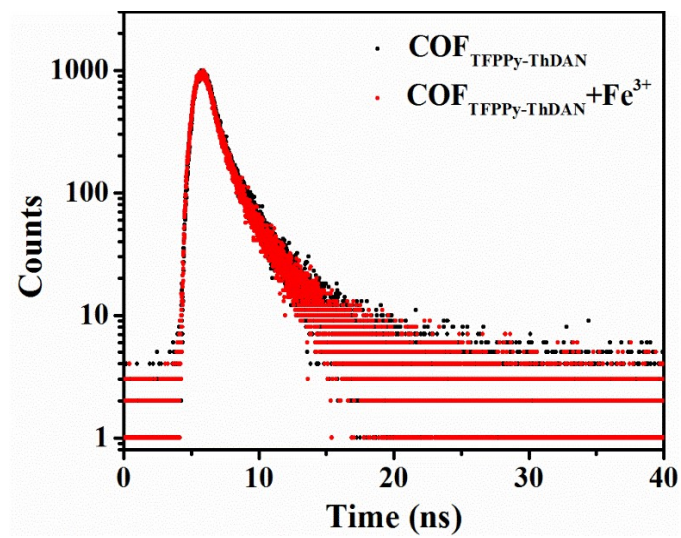


Figure S8. Fluorescence decay curves of $\text{COF}_{\text{TFPPy-ThDAN}}$ before and after adding 200 μM Fe^{3+} .

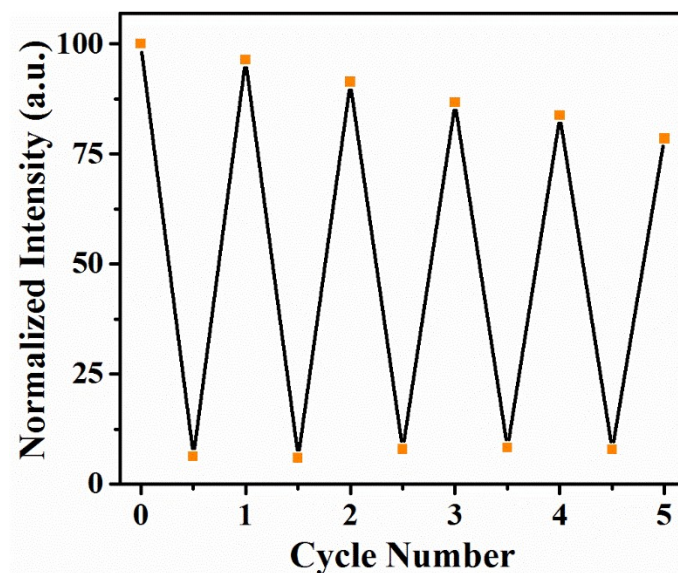


Figure S9. The reusable performance of COF_{TFPPy-ThDAN} for Fe³⁺ detection. The concentration of Fe³⁺ is 200 μ M.