

Electronic Supplementary Information (ESI)

A series of isostructural lanthanide metal-organic frameworks: effective fluorescence sensing for Fe³⁺, 2,4-DNP and 4-NP

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S1. Calculation of quenching efficiency, Stern-Volmer (S-V) equation and LOD.

Quenching efficiency:

$$\text{quenching efficiency} = (I_0 - I)/I_0 \times 100\%$$

I_0 is the fluorescence intensity values of JLU-MOF201-Y and JLU-MOF201-Tb. I is the fluorescence intensity value after adding the Fe³⁺.

Stern-Volmer (S-V) equation:

$$I_0/I = 1 + K_{sv}[C]$$

I_0 is the fluorescence intensity values of JLU-MOF201-Y and JLU-MOF201-Tb. I is the fluorescence intensity value after adding the Fe³⁺. K_{sv} is Stern-Volmer quenching constant. $[C]$ is the concentration of the analyte.

LOD:

$$L = 3S_b/K$$

S_b is standard deviation for compounds and K is slope of the Stern-Volmer.

S1. Supporting Figures.

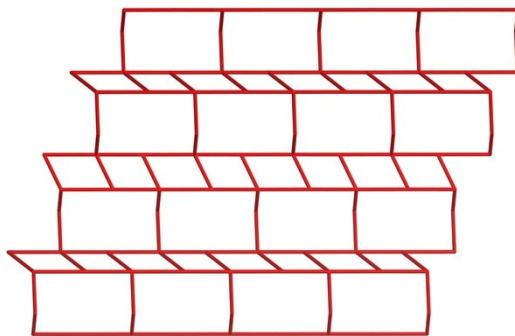


Fig. S1 3,4L83 topology of JLU-MOF201-Y.

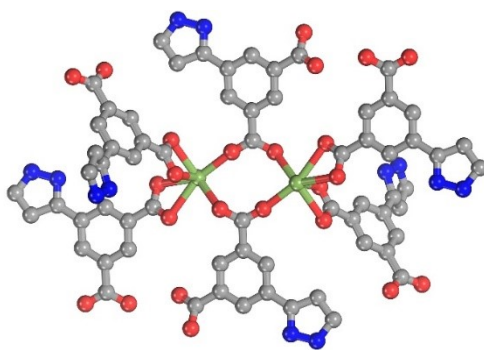


Fig. S2 Binuclear SBUs in JLU-MOF201-Y. Color scheme: carbon = gray, nitrogen = blue, oxygen = red, yttrium = green. (All H atoms and guest molecules are omitted for clarity)

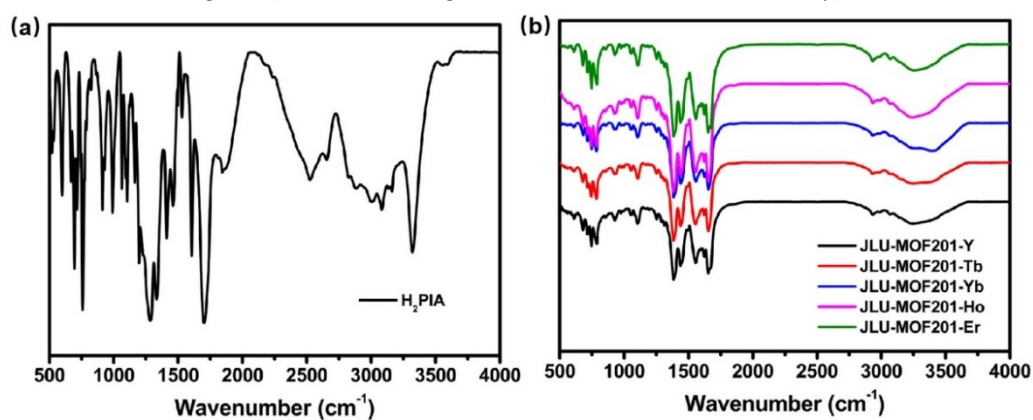


Fig. S3 The IR spectra of (a) H₂PIA and (b) JLU-MOF201-Ln.

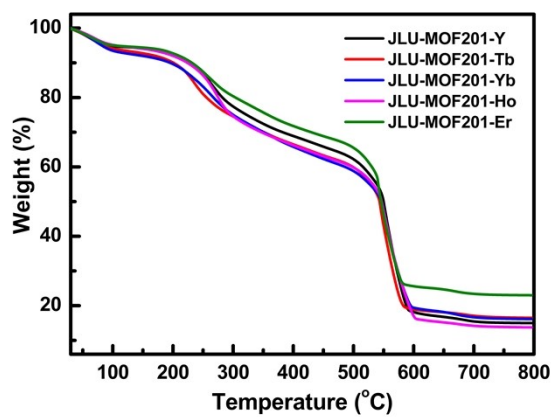


Fig. S4 Thermogravimetric analysis curves of JLU-MOF201-Ln for the as-synthesized samples.

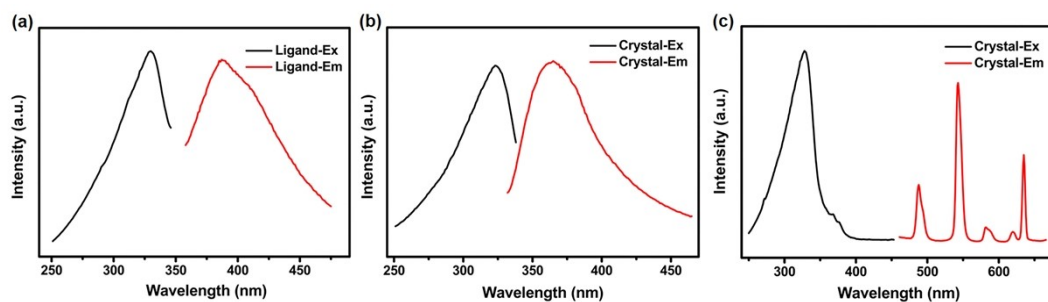


Fig. S5 Excitation and emission spectra of (a) H₂PIA, (b) JLU-MOF201-Y and (c) JLU-MOF201-Tb.

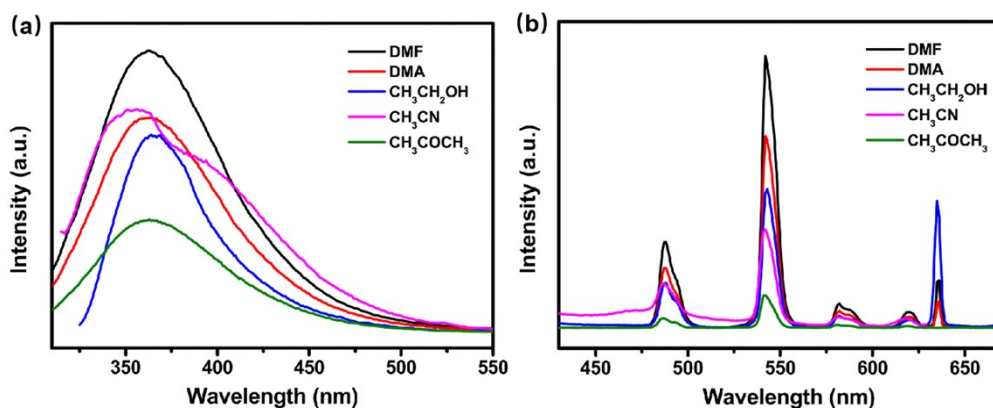


Fig. S6 Emission spectra of compounds (a) JLU-MOF201-Y and (b) JLU-MOF201-Tb in different solvents.

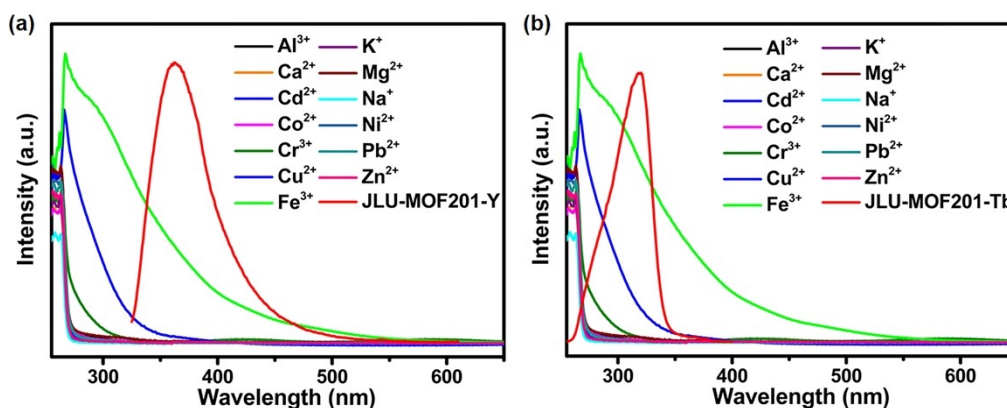


Fig. S7 UV-vis absorption spectra of different metal ions in DMF solution ($5 \times 10^{-4} \text{ mol}\cdot\text{L}^{-1}$), emission spectrum of (a) JLU-MOF201-Y and excitation spectrum of (b) JLU-MOF201-Tb in DMF solution ($2 \text{ mg}\cdot\text{mL}^{-1}$).

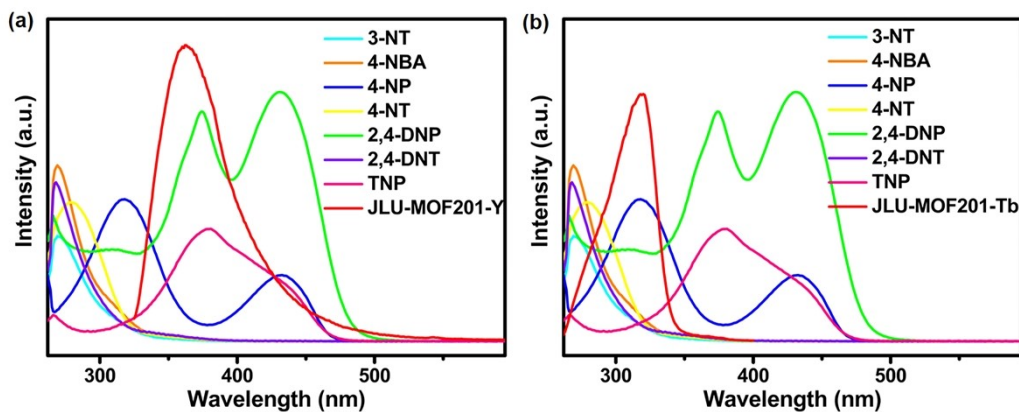


Fig. S8 UV-vis absorption spectra of different NAEs in DMF solution ($1 \times 10^{-4} \text{ M}$) and emission spectrum of (a) JLU-MOF201-Y and (b) JLU-MOF201-Tb in DMF solution ($2 \text{ mg}\cdot\text{mL}^{-1}$).

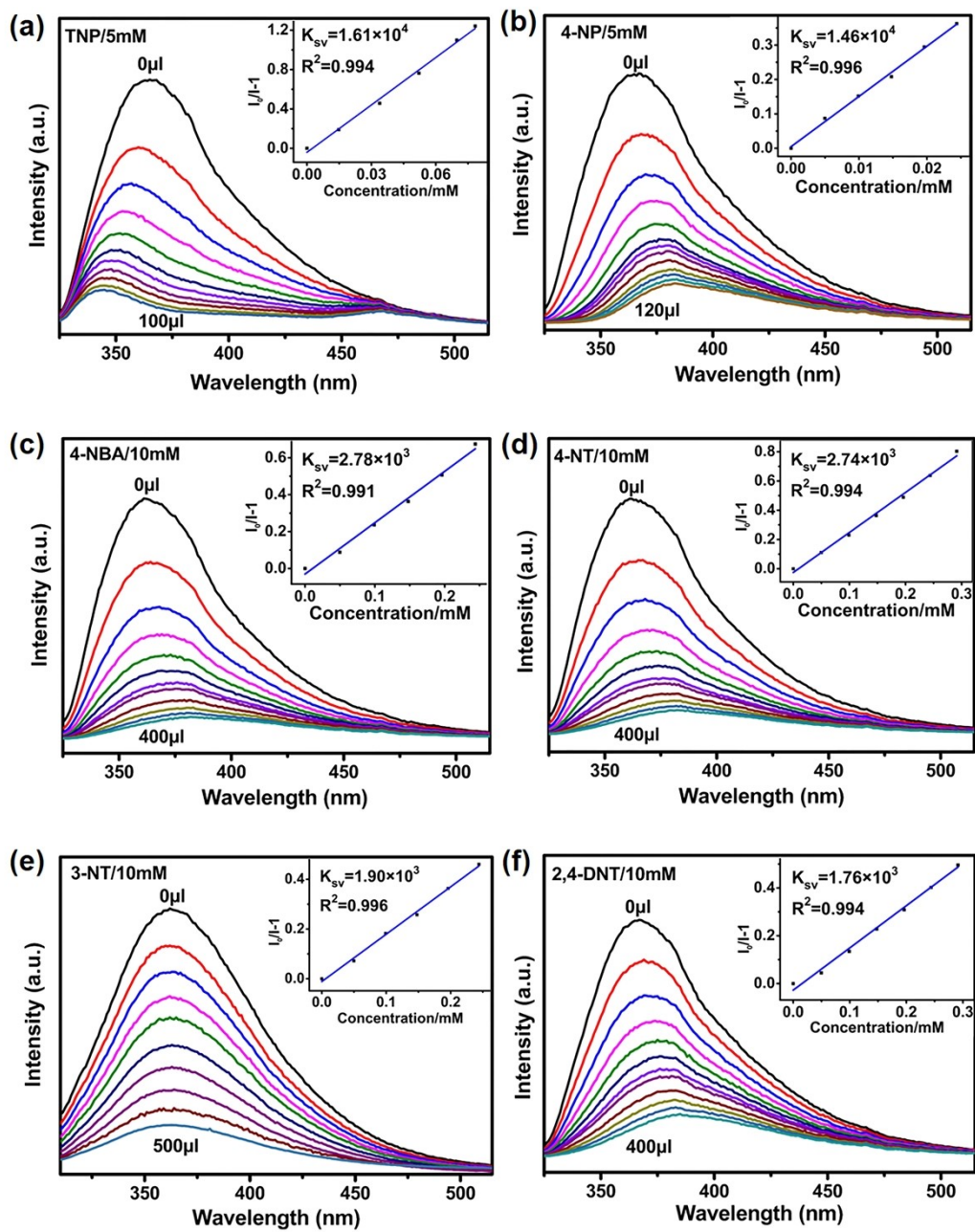


Fig. S9 Influence on the emission spectra of JLU-MOF201-Y decentralized in DMF by gradual addition of (a) TNP, (b) 4-NP, (c) 4-NBA, (d) 4-NT, (e) 3-NT and (f) 2,4-DNT solution.

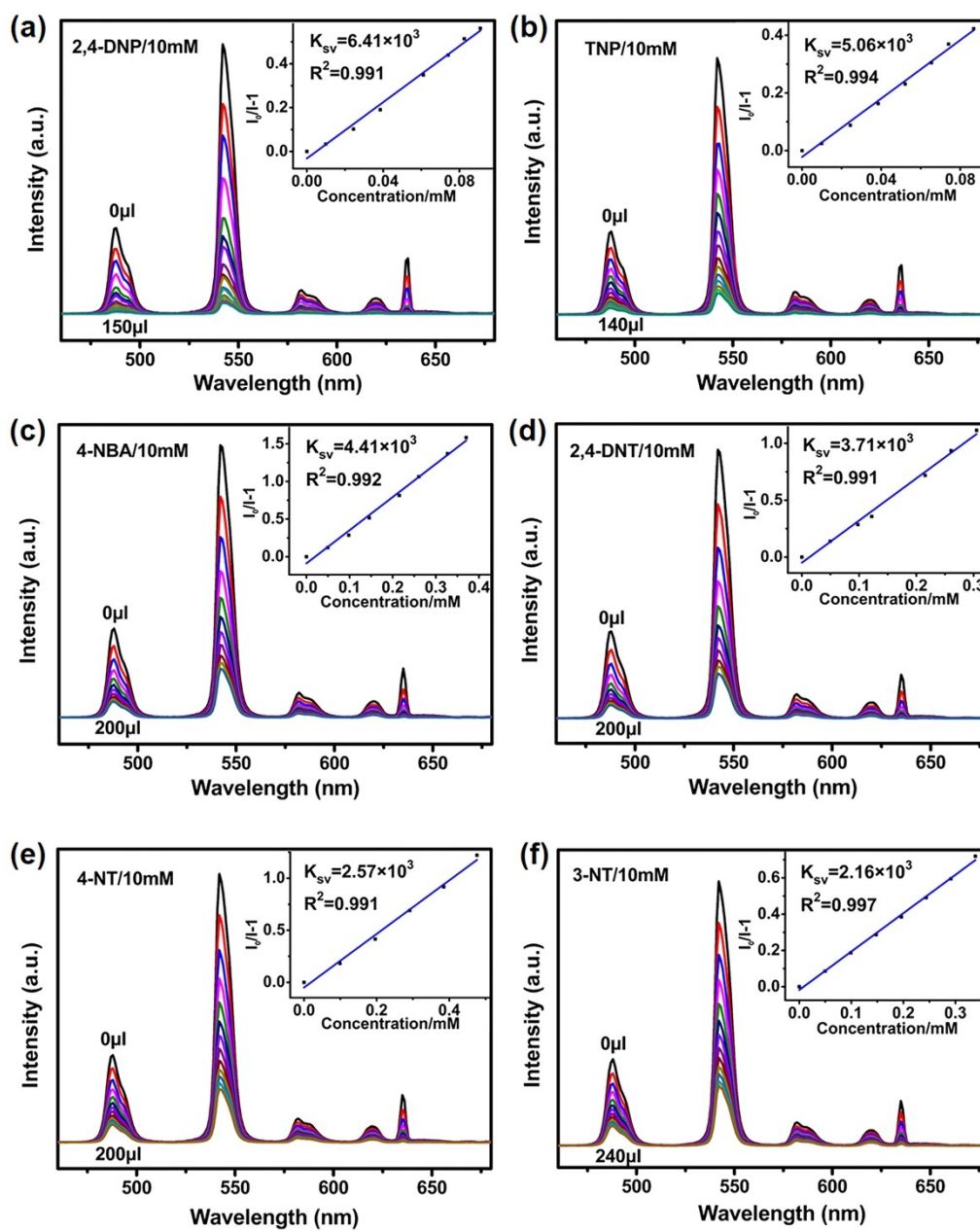


Fig. S10 Influence on the emission spectra of JLU-MOF201-Y decentralized in DMF by gradual addition of (a) 2,4-DNP, (b) TNP, (c) 4-NBA, (d) 2,4-DNT, (e) 4-NT and (f) 3-NT solution.

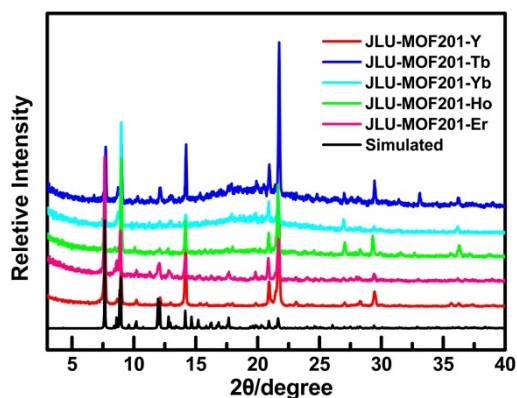


Fig. S11 PXRD patterns of JLU-MOF201-Y for simulated and as-synthesized samples of JLU-MOF201-Ln.

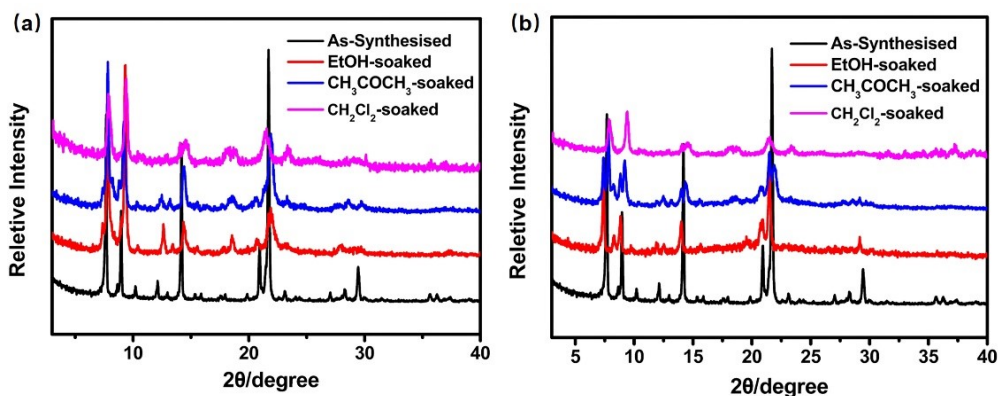


Fig. S12 PXRD patterns of (a) JLU-MOF201-Y and (b) JLU-MOF201-Tb as-synthesized and after soaking in different organic solvents.

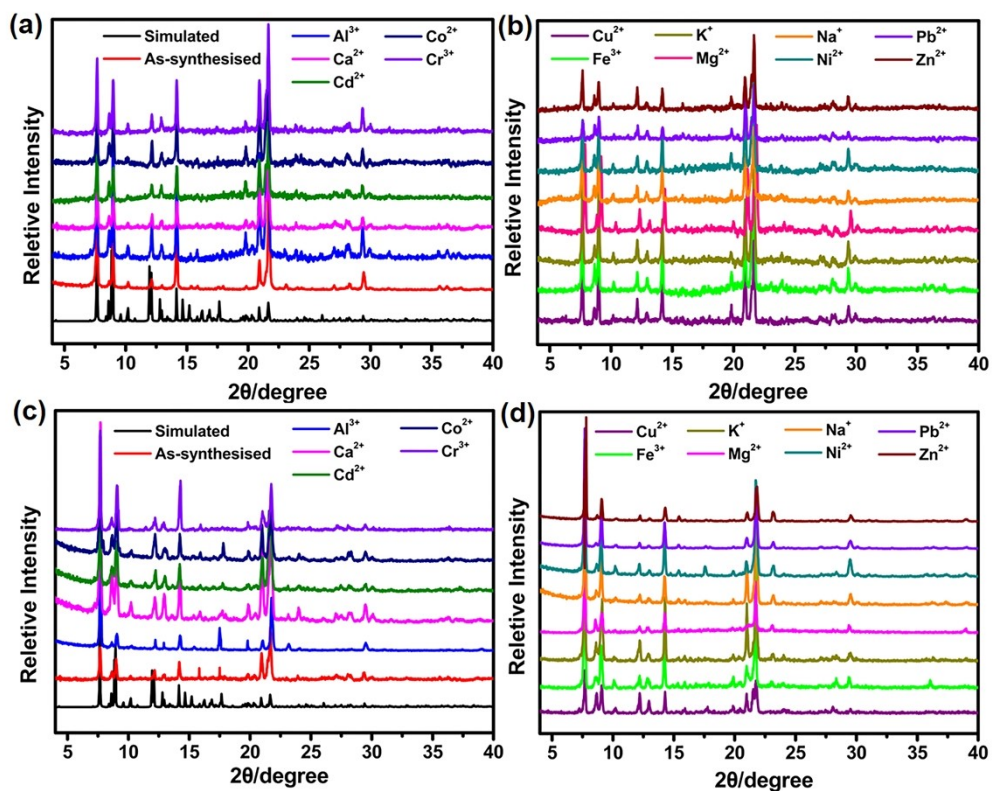


Fig. S13 PXRD patterns of JLU-MOF201-Y (a, b) and JLU-MOF201-Tb (c, d) as-synthesized and after soaking in

DMF solutions with different metal ions.

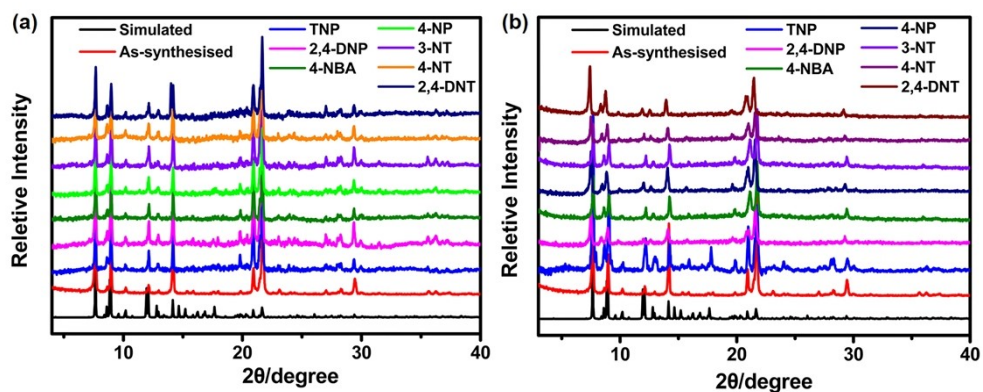


Fig. S14 PXRD patterns of as-synthesized JLU-MOF201-Y and samples after soaking in different NAEs DMF solution.

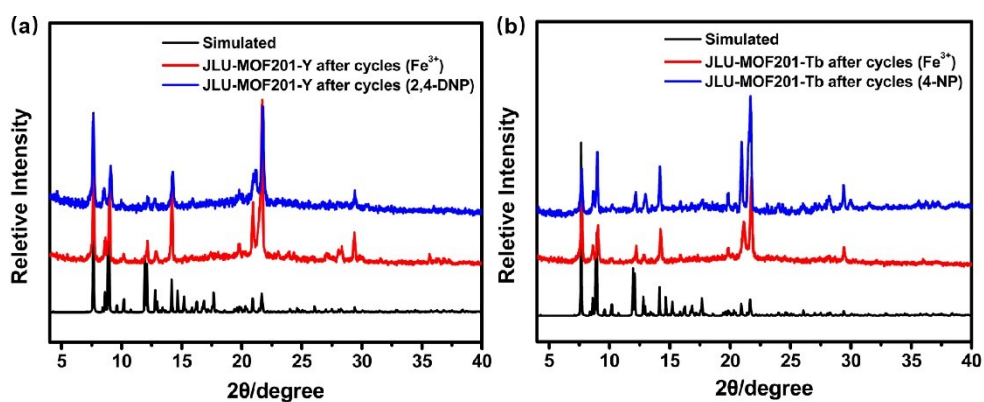


Fig. S15 PXRD patterns of JLU-MOF201-Y (a) and JLU-MOF201-Tb (b) as-synthesized and after four quenching-recovery cycles.

S1. Supporting Tables.**Table S1.** Crystal data and structure refinements for the JLU-MOF201-Y.

Parameters	JLU-MOF201-Y
Formula	C ₄₃ H ₄₂ N ₉ O ₁₆ Y ₂
Formula weight	1118.67
Temperature(K)	297.8
Wavelength (Å)	0.71073
Crystal system	Monoclinic
Space group	<i>C2/c</i>
<i>a</i> (Å)	44.809(4)
<i>b</i> (Å)	10.8964(9)
<i>c</i> (Å)	25.255(2)
α (°)	90.00
<i>B</i> (°)	113.735(2)
γ (°)	90.00
volume (Å ³)	11287.9(17)
Z, D _x (g/cm ³)	8, 1.317
<i>F</i> (000)	4552.0
ϑ range (deg)	2.39-24.26
reflns collected/unique	4927/10254
<i>R</i> _{int}	0.1071
data/restraints/params	10254/779/638
GOF on <i>F</i> ²	1.069
<i>R</i> ₁ , <i>wR</i> ₂ (<i>I</i> > 2σ(<i>I</i>))	0.0837/0.1975
<i>R</i> ₁ , <i>wR</i> ₂ (all data)	0.1466/0.2258

Table S2. Unit cell data for the JLU-MOF201-Ln.

	JLU-MOF201-Y	JLU-MOF201-Tb	JLU-MOF201-Yb	JLU-MOF201-Ho	JLU-MOF201-Er
Crystal system	Monoclinic	Monoclinic	Monoclinic	Monoclinic	Monoclinic
<i>a</i> (Å)	44.81	45.21	45.77	44.93	45.02
<i>b</i> (Å)	10.90	10.98	11.06	10.90	10.91
<i>c</i> (Å)	25.26	25.34	25.73	25.23	25.26
α (°)	90.00	90.00	90.00	90.00	90.00
β (°)	113.74	113.76	113.44	113.73	113.45
γ (°)	90.00	90.00	90.00	90.00	90.00
volume (Å ³)	11287	11514	11950	11309	11387

Table S3. A comparison of K_{sv} and LODs values of luminescent MOFs for Fe^{3+} detection.

MOFs	K_{sv} (M^{-1})	LOD (μM)	Ref.
$[Zr_6O_6(OH)_2(CF_3COO)_2(C_{11}H_5NO_4)_4(H_2O)_4]$	2.25×10^7	1.70×10^{-3}	1
$[Zn_2Na_2(TPHC)(4,4-Bipy)(DMF)] \cdot 8H_2O$	5.57×10^4	6.40	2
$[EuDTTA(DMF)_3]NO_3$	3.63×10^4	4.14	3
$\{[Cd(tptb)(H_2DOBD)(H_2O)] \cdot DMF\}_n$	3.00×10^4	1.67×10^{-5}	4
$[Tb(L_2)(H_2O)(DMF)]_n$	2.89×10^4	0.91	5
$\{[Eu(L)(H_2O)] \cdot 4H_2O\}_n$	1.88×10^4	0.57	6
$\{[Tb(L)(H_2O)] \cdot 7H_2O\}_n$	1.48×10^4	1.26	7
$[Tb_2(PIA)_3(DMF)_3(CH_3OH)]$	8.83×10^4	2.49	This work
FJI-C8 (Zn)	8.25×10^3	23.3	8
$[Y_2(PIA)_3(DMF)_3(CH_3OH)]$	7.67×10^3	2.21	This work
534-MOF-Tb(L11)	5.51×10^3	130	9
$[Zn_2(cptpy)(btc)(H_2O)]_n$	5.46×10^3	4.33	10
$[Zr_6O_4(OH)_4(C_8H_2O_4S_2)_6] \cdot DMF \cdot 18H_2O$	4.41×10^3	1.26	11
$[Zn(3-bpmh)(HEA) \cdot H_2O]_n$	4.14×10^3	0.89	12
$\{[Eu(BIPATC)_{0.5}(DMA)_2(NO_3)] \cdot DMA \cdot H_2O\}_n$	3.89×10^3	-	13
$[Zn(HBCPBA)(tpim)] \cdot H_2O$	3.57×10^3	-	14
$\{[Eu(TMCA)(DEF)(H_2O)] \cdot (CH_3CN)\}_n$	1.84×10^3	30.10	15

Table S4. A comparison of K_{sv} and LODs values of luminescent MOFs for 2,4-DNP detection.

MOFs	K_{sv} (M^{-1})	LOD (μM)	Ref.
$[Cd_2(TPA)_2(BIYB)_2]_n$	1.60×10^5	0.13	16
$[Zn_5(DpImDC)_2(DMF)_4(H_2O)_3] \cdot H_2O \cdot DMF$	8.70×10^4	-	17
$[Cd(TTPBA-4)_2(OH)_2 \cdot H_2O]_n$	4.10×10^4	9.50×10^{-3}	18
$[Y_2(PIA)_3(DMF)_3(CH_3OH)]$	3.63×10^4	0.46	This work
Sc-EBTC	2.85×10^4	4.30×10^{-5}	19
$[Zn_2(TCPE)(tta)_2] \cdot 2DMF \cdot 4H_2O \cdot 2Me_2NH_2^+$	2.60×10^4	6.12×10^{-5}	20
$\{[(CH_3)_2NH_2]_6[Cd_5(L)_4] \cdot H_2O \cdot 3DMF\}_n$	2.37×10^4	4.57×10^{-3}	21
$\{(NH_2(CH_3)_2)[Zn_4(ddn)_2(COO)(H_2O)_4] \cdot solvent\}_n$	0.89×10^4	8.75×10^{-3}	22

Table S5. A comparison of the values of K_{sv} and LODs of luminescent MOFs for the detection of 4-NP.

MOFs	K_{sv} (M^{-1})	LOD (μM)	Ref.
$[Eu_{0.5}Tb_{0.5}(L)(H_2O)_3]_n$	7.50×10^6	-	23
$\{[Tb(HL)] \cdot 3DMF \cdot 3H_2O\}_n$	3.60×10^5	2.33×10^{-8}	24
$\{[Cd_{3.5}(\mu_3-OH)(TCPB)_2(bimb)(H_2O)_3] \cdot 0.5dioxane \cdot 0.5H_2O\}$	3.80×10^4	0.49×10^{-3}	25
$[Tb_2(PIA)_3(DMF)_3(CH_3OH)]$	1.89×10^4	1.01	This work
Eu-CP	1.38×10^4	3.00×10^{-9}	26
$[Cd(L)(phen)_2] \cdot 5H_2O$	5.13×10^3	1.15×10^{-5}	27

Table S6. The K_{sv} and LODs values of JLU-MOF201-Y and JLU-MOF201-Tb for the detection of NAEs.

NAEs	JLU-MOF201-Y		JLU-MOF201-Tb	
	K_{sv} (M^{-1})	LOD (μM)	K_{sv} (M^{-1})	LOD (μM)
2,4-DNP	3.36×10^4	0.46	6.41×10^3	2.99
4-NP	1.46×10^4	1.15	1.89×10^4	1.01
TNP	1.16×10^4	1.04	5.06×10^3	3.79
4-NBA	2.78×10^3	6.05	4.41×10^3	4.35
3-NT	1.90×10^3	8.85	2.16×10^3	8.87
4-NT	2.74×10^3	6.14	2.57×10^3	7.46
2,4-DNT	1.76×10^3	9.60	3.71×10^3	5.17

Table S7. The ICP data of JLU-MOF201-Y and Fe^{3+} @JLU-MOF201-Y.

Compounds	Y^{3+} (ppm)	Fe^{3+} (ppm)
JLU-MOF201-Y	17.96	-
Fe^{3+} @ JLU-MOF201-Y	23.95	0.11

Table S8. The ICP data of JLU-MOF201-Tb and Fe^{3+} @JLU-MOF201-Tb.

Compounds	Tb^{3+}	Fe^{3+}
JLU-MOF201-Tb	18.24	-
Fe^{3+} @ JLU-MOF201-Tb	25.83	0.19

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