

Electronic Supplementary Information (ESI) II – structures, CV, UV-vis spectroscopy and DFT calculations

for

Isoelectronic Pt(II) Complexes of Cyclometalating C^NN Ligands with Phenyl/(Benzo)thiophenyl/Pyridyl/(Benzo)thiazole Moieties

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Supporting Tables

Table S1 Selected structure solution and refinement data for [Pt(th(ppy)py)Cl]·CH₂Cl₂ (**10**·CH₂Cl₂) and [Pt(ph(tbpppy)btz)(CCPh)] (**9**).

Table S2 Selected structural data for [Pt(th(ppy)py)Cl]·CH₂Cl₂ (**10**·CH₂Cl₂) and [Pt(ph(tbpppy)btz)(CCPh)] (**9**).

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Table S4 Redox potentials of the [Pt(C[^]N[^]N)X] complexes.

Table S5 UV-vis absorption maxima of HC[^]N[^]N protoligands.

Table S6 UV-vis absorption maxima of the [Pt(C[^]N[^]N)X] complexes.

Table S7 Complete emission data and PLYQs and excited state lifetime data in CH₂Cl₂ (298 K) and in frozen glassy CH₂Cl₂/MeOH matrix at 77 K.

Table S8 TDDFT calculated transitions contributing to the emissive *T*₁ state.

Photophysical characterisation

Steady-state excitation and emission spectra were recorded on a FluoTime300 spectrometer from PicoQuant equipped with a 300 W ozone-free Xe lamp (250-900 nm), a 10 W Xe flash-lamp (250-900 nm, pulse width < 10 μ s) with repetition rates of 0.1 to 300 Hz, an excitation monochromator (Czerny-Turner 2.7 nm/mm dispersion, 1200 grooves/mm, blazed at 300 nm), diode lasers (pulse width < 80 ps) operated by a computer-controlled laser driver PDL-820 (repetition rate up to 80 MHz, burst mode for slow and weak decays), two emission monochromators (Czerny-Turner, selectable gratings blazed at 500 nm with 2.7 nm/mm dispersion and 1200 grooves/mm, or blazed at 1250 nm with 5.4 nm/mm dispersion and 600 grooves/mm), Glan-Thompson polarisers for excitation (Xe-lamps) and emission, a Peltier-thermostated sample holder from Quantum Northwest (-40 $^{\circ}$ C to 105 $^{\circ}$ C), and two detectors, a PMA Hybrid 40 (transit time spread FWHM < 120 ps, 300 to 720 nm) and a R5509-42 NIR-photomultiplier tube (transit time spread FWHM 1.5 ns, 300-1400 nm) with external cooling (-80 $^{\circ}$ C) from Hamamatsu.

Steady-state and fluorescence lifetimes were recorded in TCSPC mode by a PicoHarp 300 (minimum base resolution 4 ps).

Phosphorescence lifetimes were recorded in MCS mode by a TimeHarp 300 (minimum base resolution 250 ps). Emission and excitation spectra were corrected for source intensity (lamp and grating) by standard correction curves.

Lifetime analysis was performed using the commercial FluoFit software. The quality of the fit was assessed by minimizing the reduced chi squared function (χ^2) and visual inspection of the weighted residuals and their autocorrelation.

Luminescence quantum yields (PLQY) were measured with a Hamamatsu Photonics absolute PL quantum yield measurement system (C9920-02) equipped with a L9799-01 CW Xenon light source (150 W), monochromator, C7473 photonic multi-channel analyser, integrating sphere and employing U6039-05 PLQY measurement software (Hamamatsu Photonics, Ltd., Shizuoka, Japan). The used dichloromethane was of spectroscopic grade (Uvasol[®]).

Supporting Figures

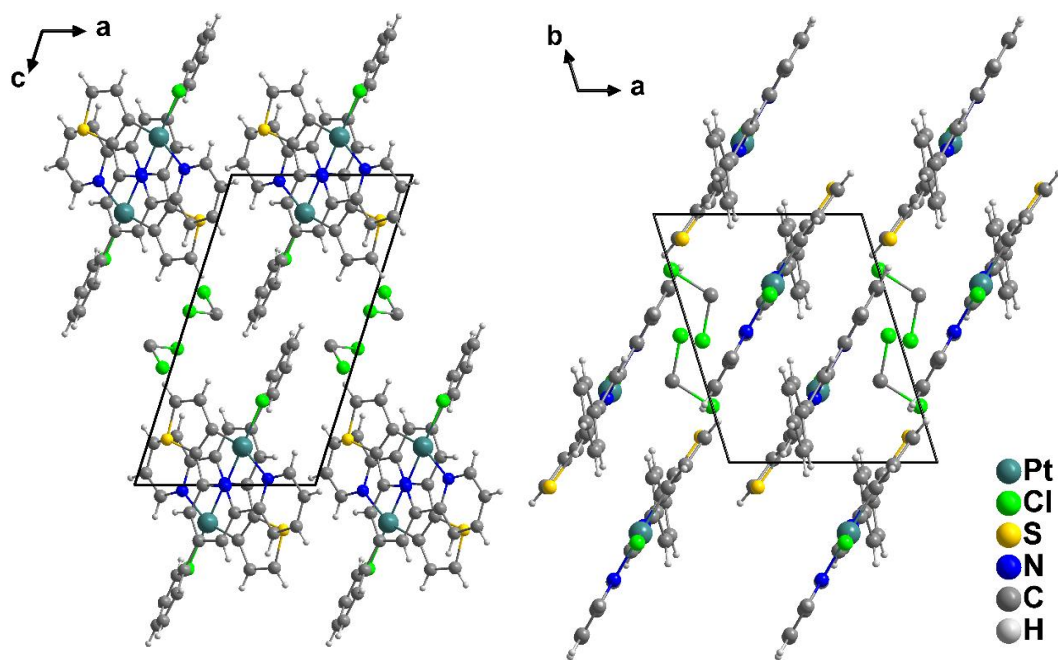


Fig. S1 Crystal structure of [Pt(th(ppy)py)Cl]·CD₂Cl₂ (10·CH₂Cl₂) viewed along the crystallographic *b* (left) and *c* (right) axes.

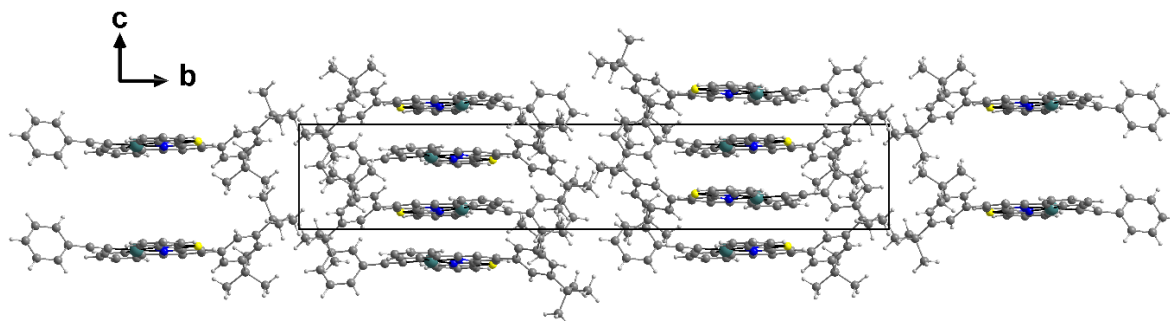


Fig. S2 Crystal structure of [Pt(ph(tbppy)bzt)CCPh] (9) viewed along the crystallographic *a* axis.

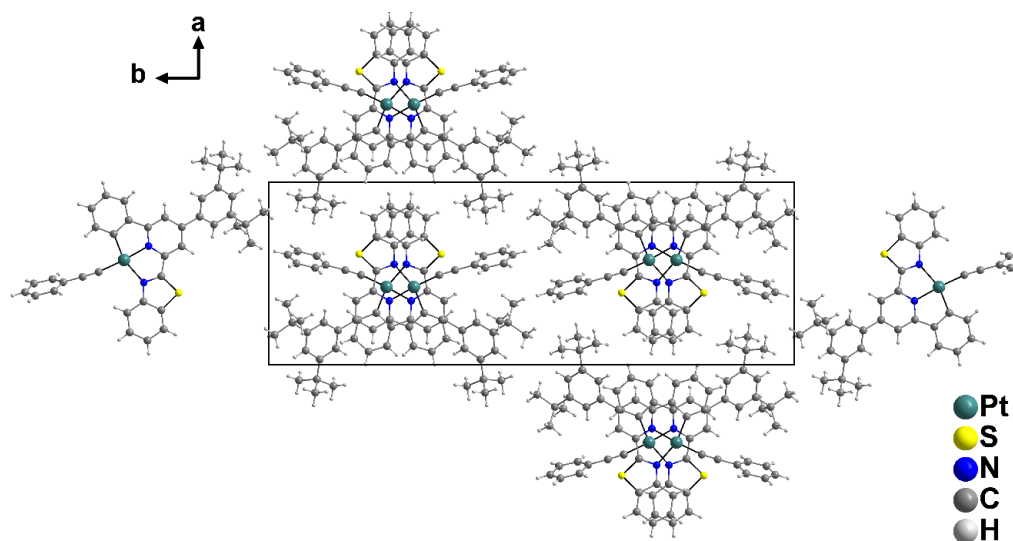


Fig. S3 Crystal structure of [Pt(ph(tbppy)bzt)CCPh] (9) viewed along the crystallographic *c* axis.

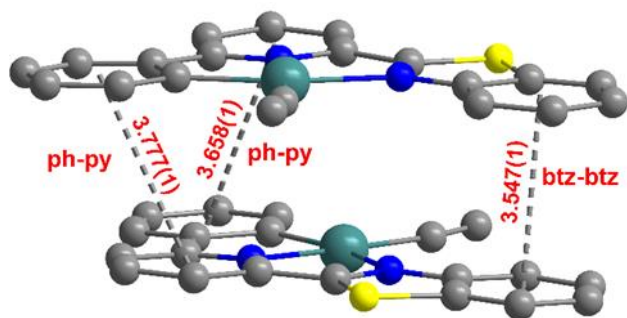


Fig. S4 Intermolecular π -stacking interactions in [Pt(ph(tbppy)bzt)CCPh] (9).

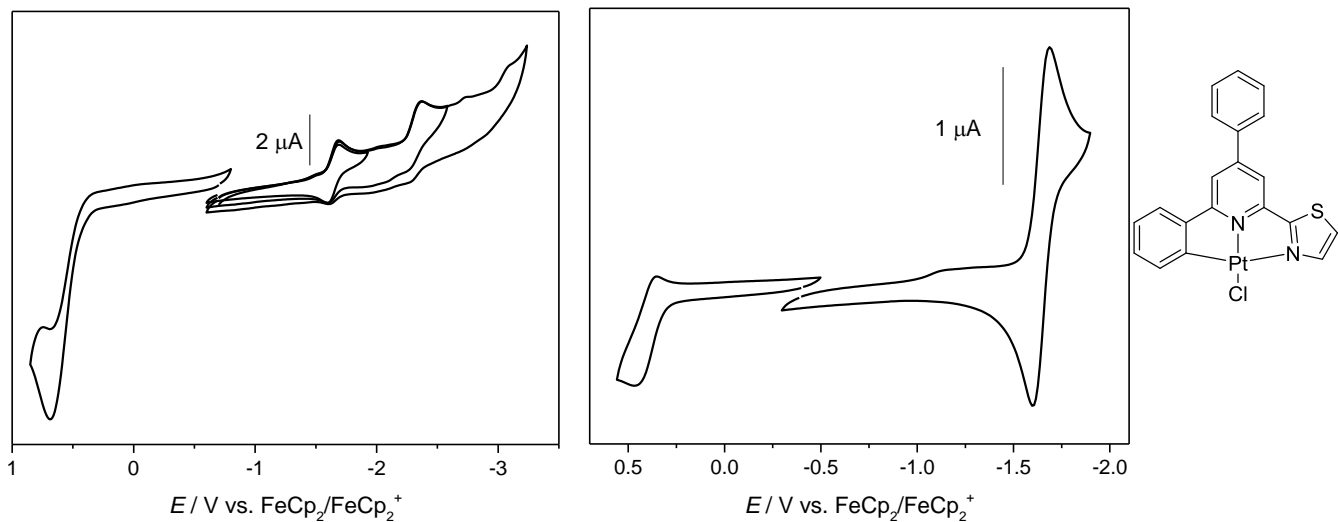


Fig. S5 Cyclic voltammograms of [Pt(ph(ppy)tz)Cl] (**1**) in 0.1 M *n*Bu₄NPF₆/THF (left) and 0.1 M /CH₂Cl₂ (right, Pt working electrode).

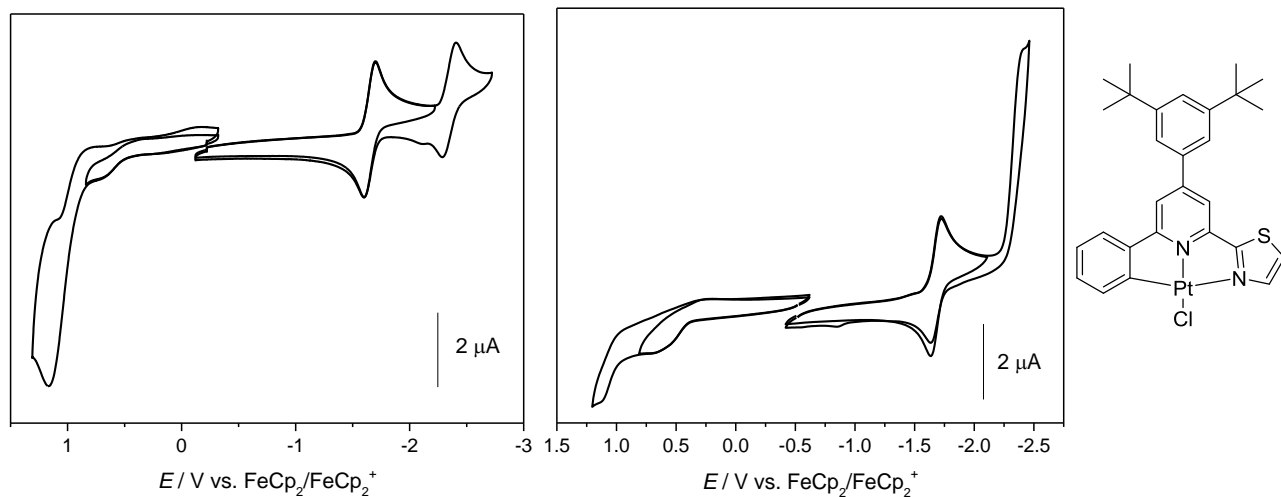


Fig. S6 Cyclic voltammograms of [Pt(ph(tbbpy)tz)Cl] (**2**) in 0.1 M *n*Bu₄NPF₆/THF (left) and /CH₂Cl₂ (right).

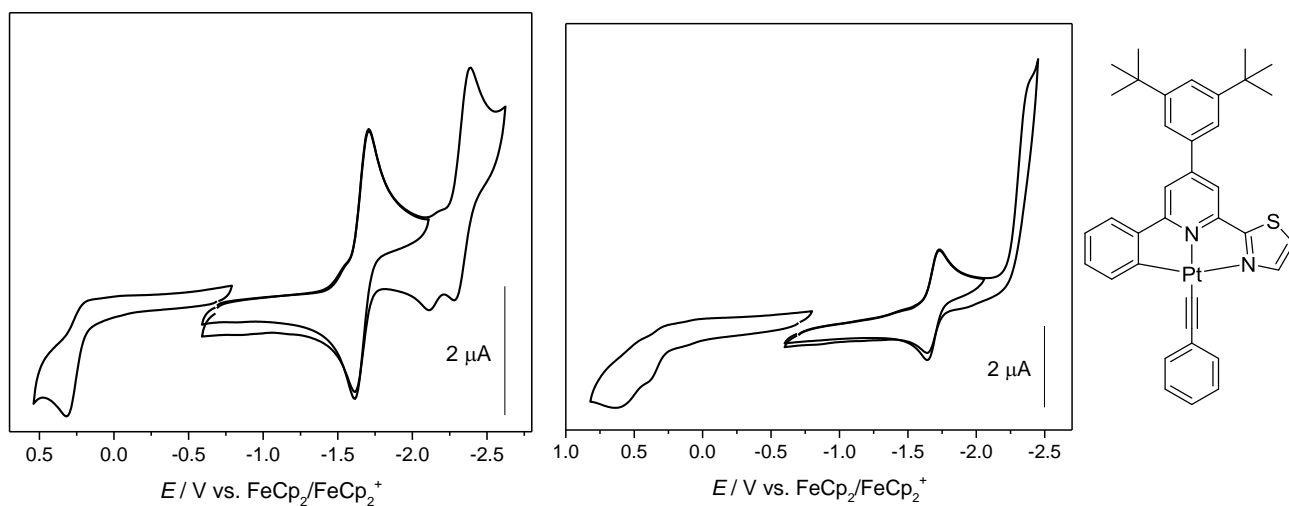


Fig. S7 Cyclic voltammograms of [Pt(ph(tbbpy)tz)(CCPh)] (**3**) in 0.1 M *n*Bu₄NPF₆/THF (left) and /CH₂Cl₂ (right).

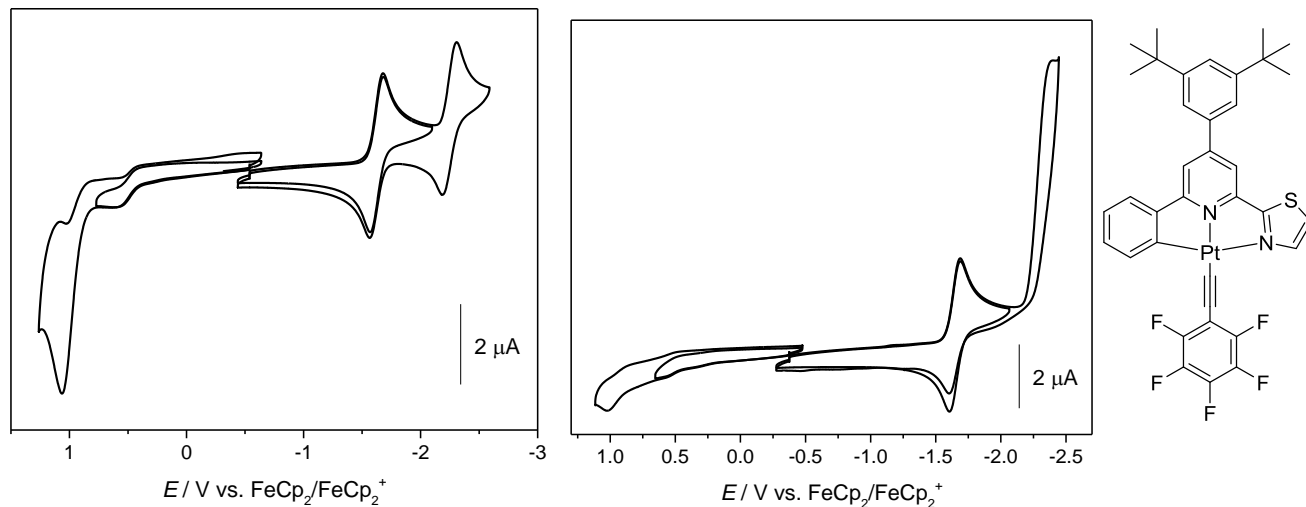


Fig. S8 Cyclic voltammograms of [Pt(ph(tbppy)tz)(CCCF₅)] (**4**) in 0.1 M *n*Bu₄NPF₆/THF (left) and /CH₂Cl₂ (right).

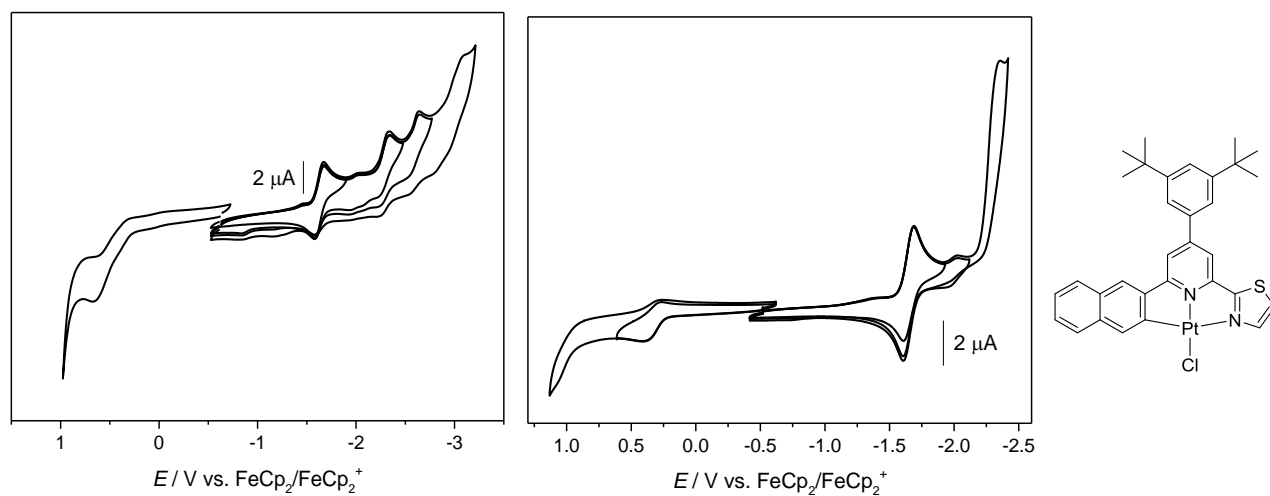


Fig. S9 Cyclic voltammograms of [Pt(na(tbppy)tz)Cl] (**5**) in 0.1 M *n*Bu₄NPF₆/THF (left) and /CH₂Cl₂ (right).

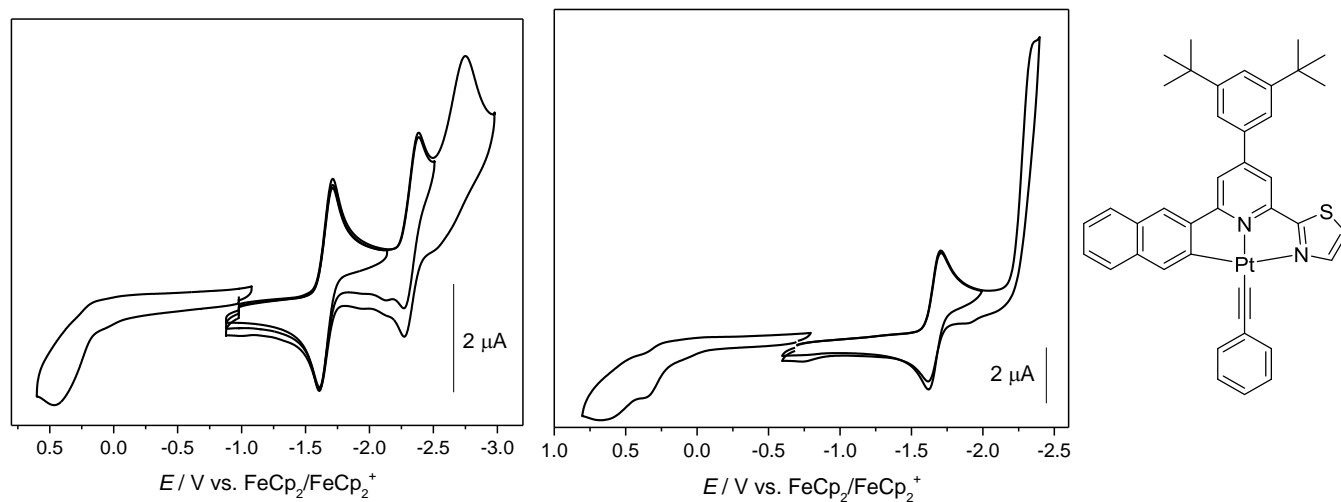


Fig. S10 Cyclic voltammograms of [Pt(na(tbppy)tz)(CCPh)] (**6**) in 0.1 M *n*Bu₄NPF₆/THF (left) and /CH₂Cl₂ (right).

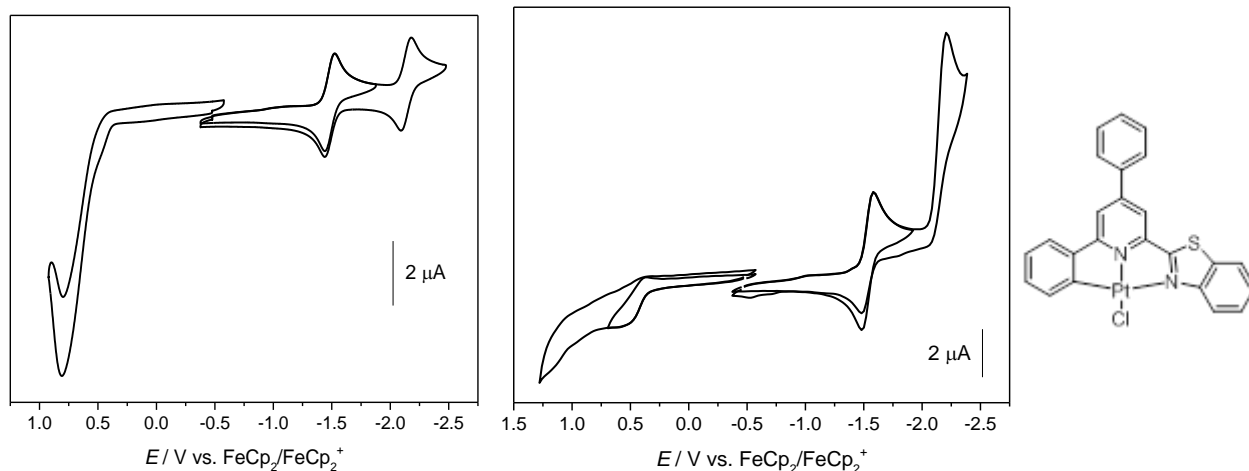


Fig. S11 Cyclic voltammograms of [Pt(ph(ppy)btz)Cl] (7) in 0.1 M *n*Bu₄NPF₆/THF (left) and /CH₂Cl₂ (right).

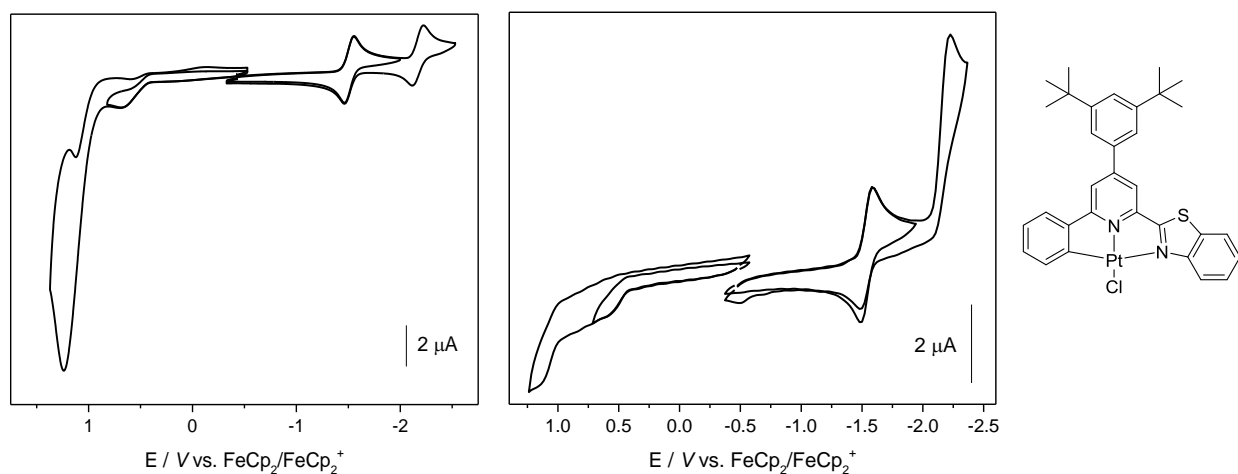


Fig. S12 Cyclic voltammograms of [Pt(ph(tbppy)btz)Cl] (8) in 0.1 M *n*Bu₄NPF₆/THF (left) and /CH₂Cl₂ (right).

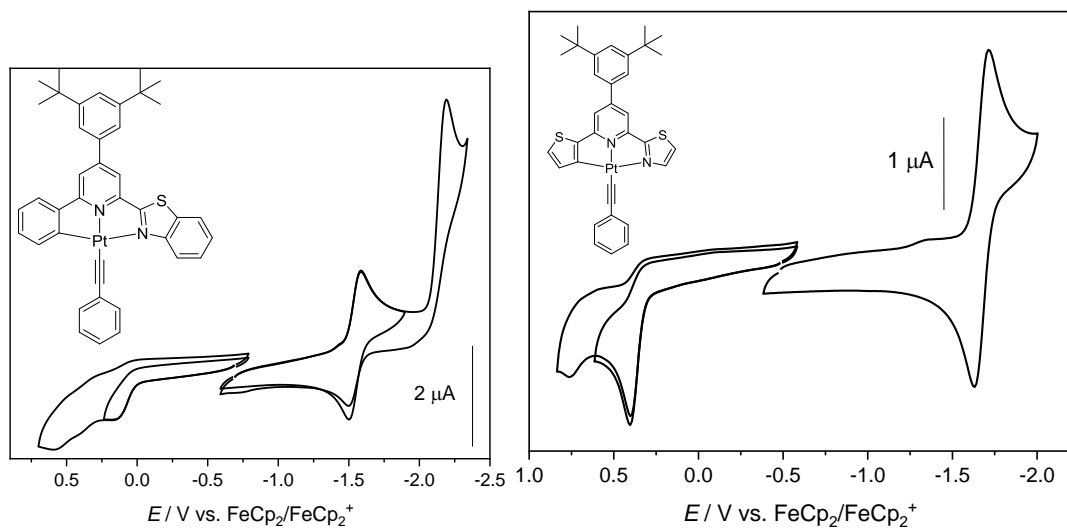


Fig. S13 Cyclic voltammograms of [Pt(ph(tbppy)btz)(CCPh)] (9) (left) and [Pt(th(tbppy)tz)(CCPh)] (18) (right) in 0.1 M *n*Bu₄NPF₆/CH₂Cl₂.

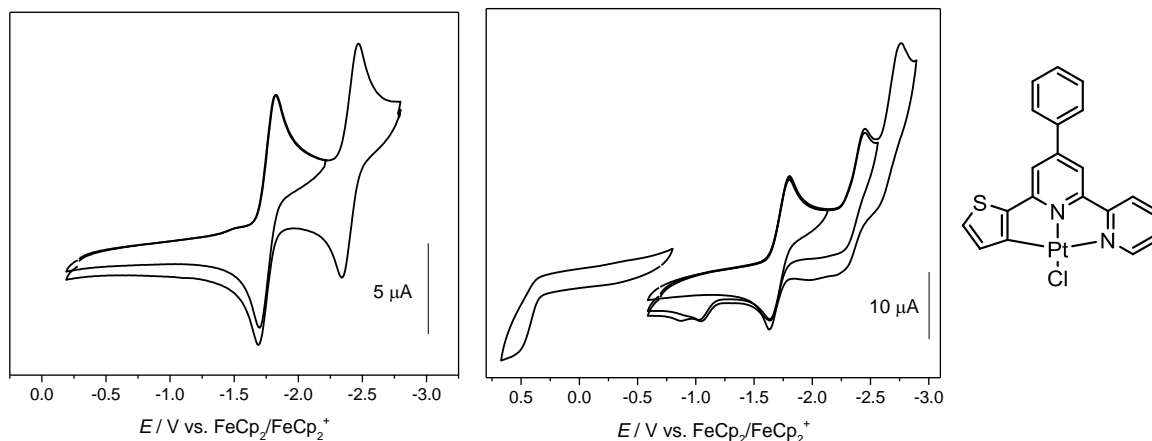


Fig. S14 Cyclic voltammograms of [Pt(th(tpy)py)Cl] (**10**) in 0.1 M *n*Bu₄NPF₆/THF (left) and /DMF (right).

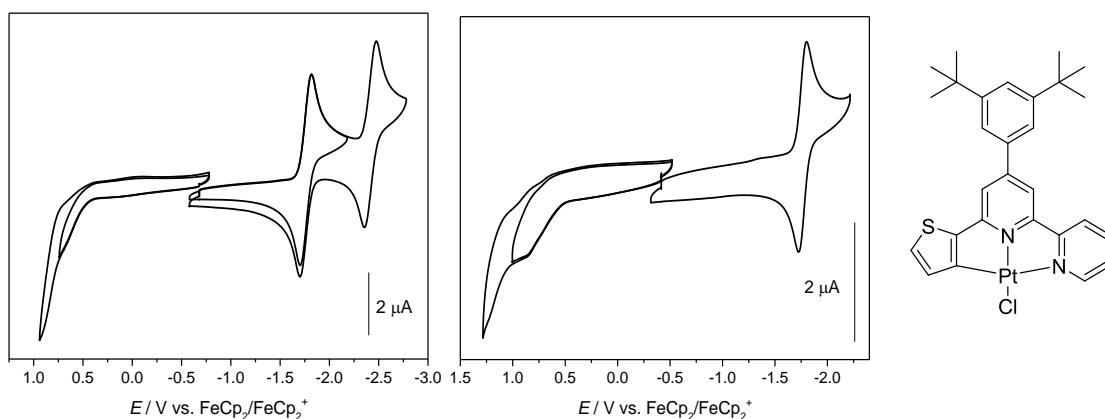


Fig. S15 Cyclic voltammograms of [Pt(th(tbbpy)py)Cl] (**11**) in 0.1 M *n*Bu₄NPF₆/THF (left) and 0.1 M /CH₂Cl₂ (right).

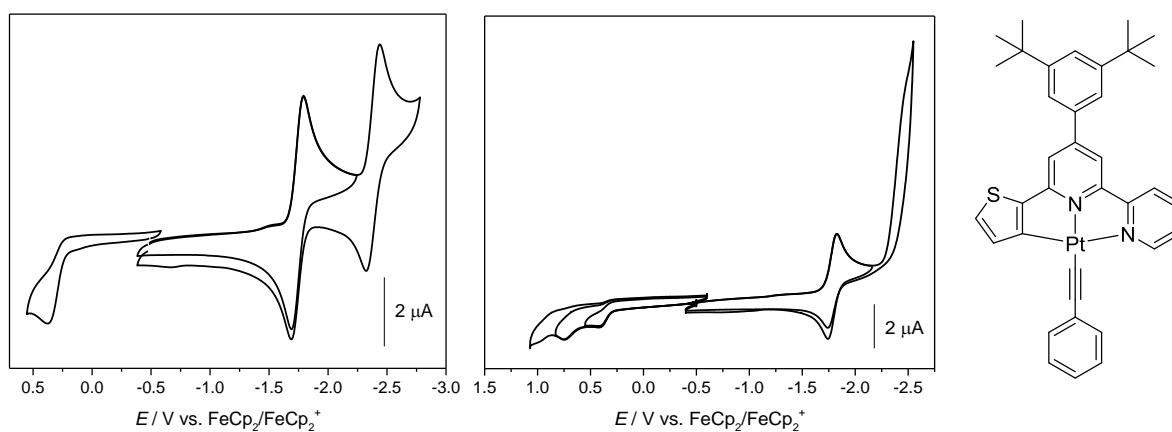


Fig. S16 Cyclic voltammograms of [Pt(th(tbbpy)py)CCPh] (**12**) in 0.1 M *n*Bu₄NPF₆/THF (left) and /CH₂Cl₂ (right).

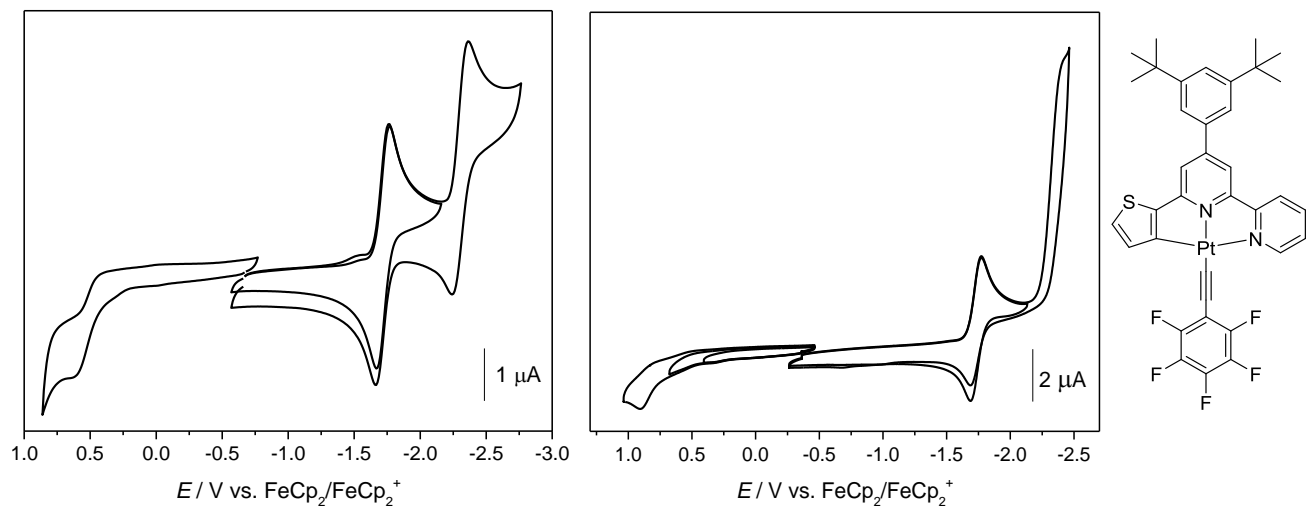


Fig. S17 Cyclic voltammograms of [Pt(th(tbbppy)py)CCCF₅] (13) in 0.1 M *n*Bu₄NPF₆/THF (left) and /CH₂Cl₂ (right).

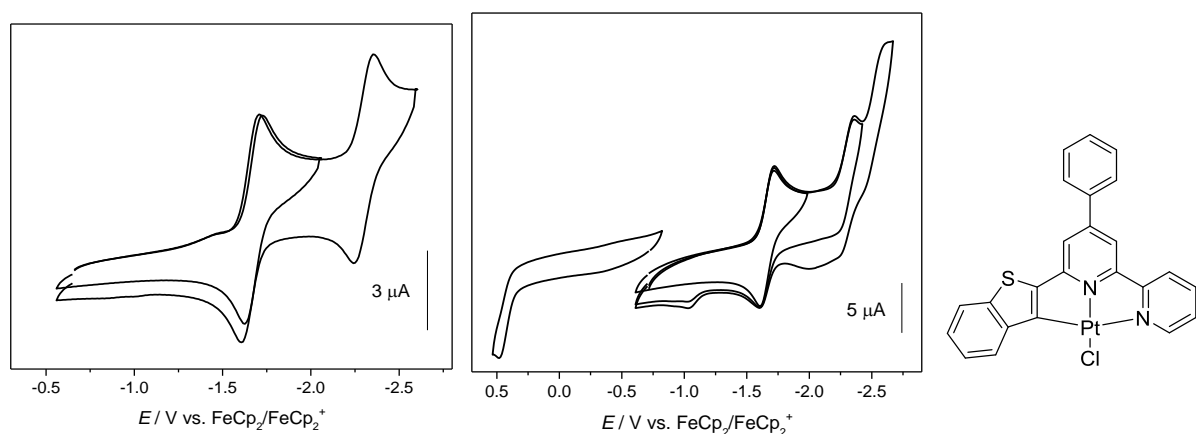


Fig. S18 Cyclic voltammograms of [Pt(bth(ppy)py)Cl] (14) in 0.1 M *n*Bu₄NPF₆/THF (left) and /DMF (right).

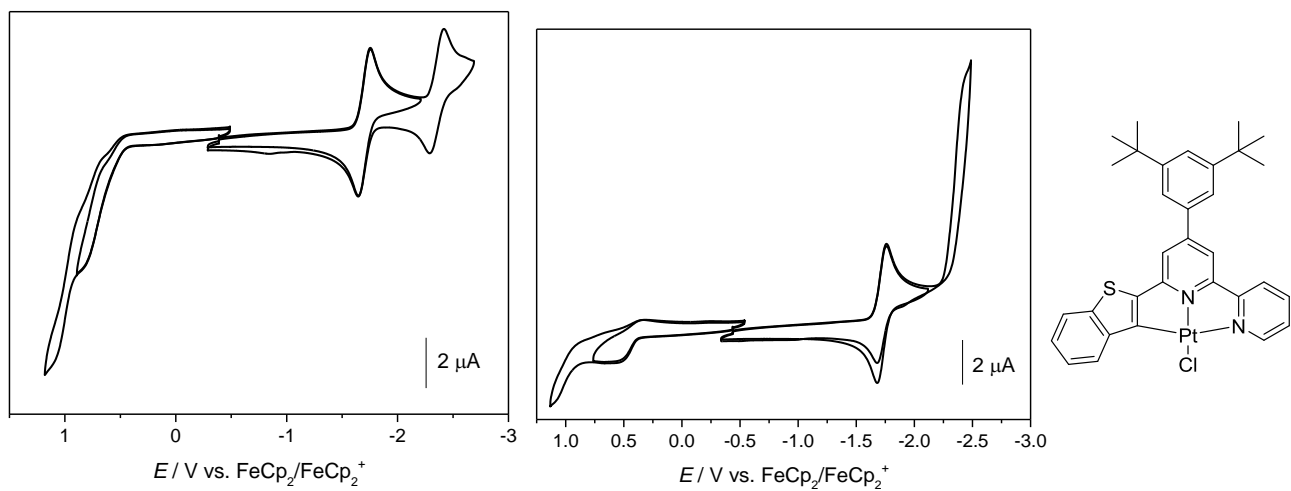


Fig. S19 Cyclic voltammograms of [Pt(bth(tbbppy)py)Cl] (15) in 0.1 M *n*Bu₄NPF₆/THF (left) and /CH₂Cl₂ (right).

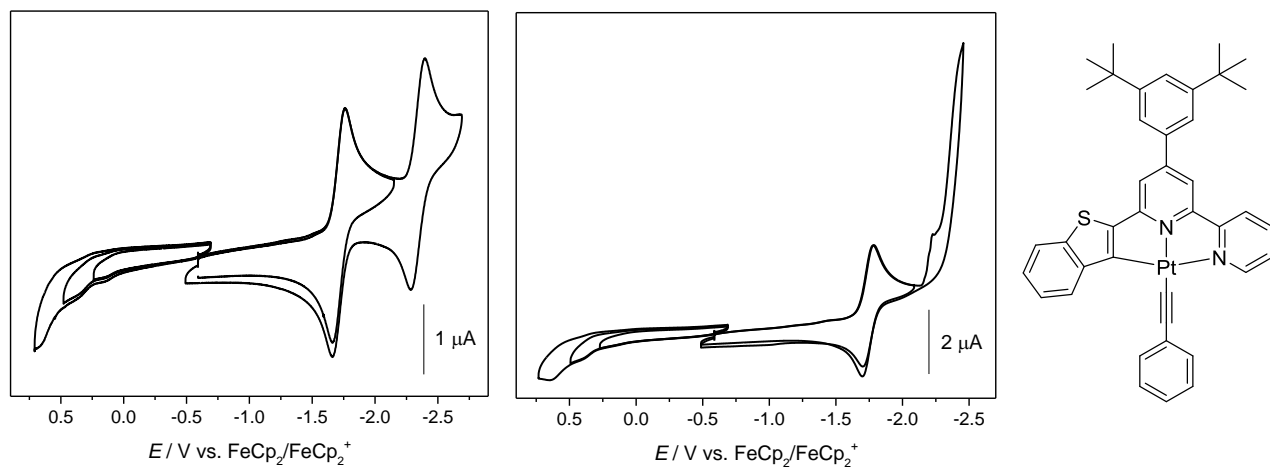


Fig. S20 Cyclic voltammograms of [Pt(bth(tbppy)py)CCPh] (**16**) in 0.1 M *n*Bu₄NPF₆/THF (left) and /CH₂Cl₂ (right).

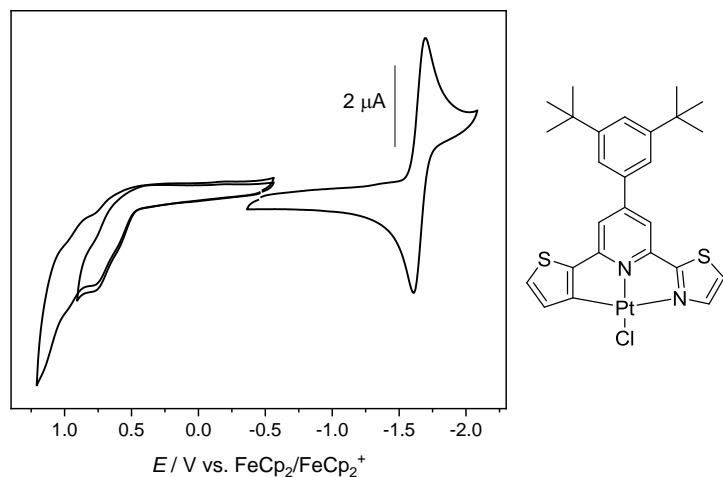


Fig. S21 Cyclic voltammograms of [Pt(th(tbppy)tz)Cl] (**17**) in 0.1 M *n*Bu₄NPF₆/CH₂Cl₂.

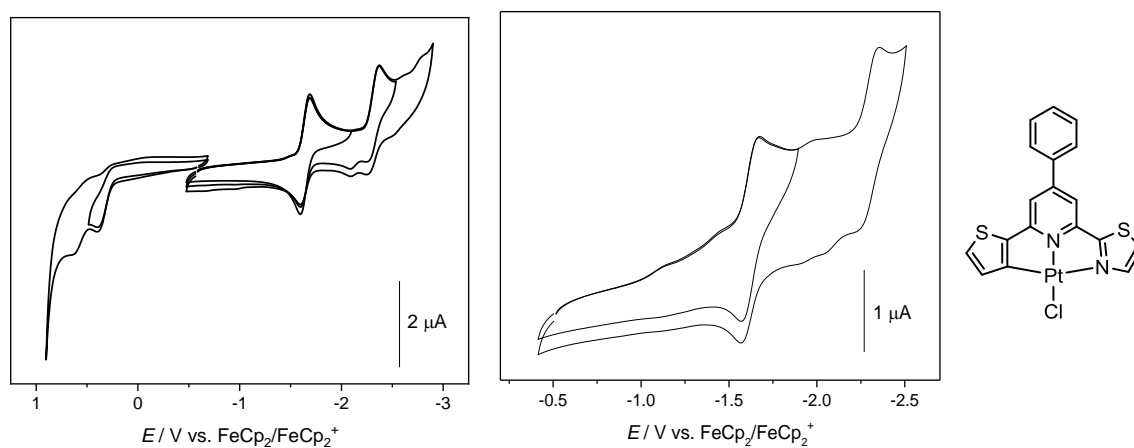


Fig. S22 Cyclic voltammograms of [Pt(th(ppy)tz)Cl] in 0.1 M *n*Bu₄NPF₆/THF (left) and /CH₂Cl₂ (right).

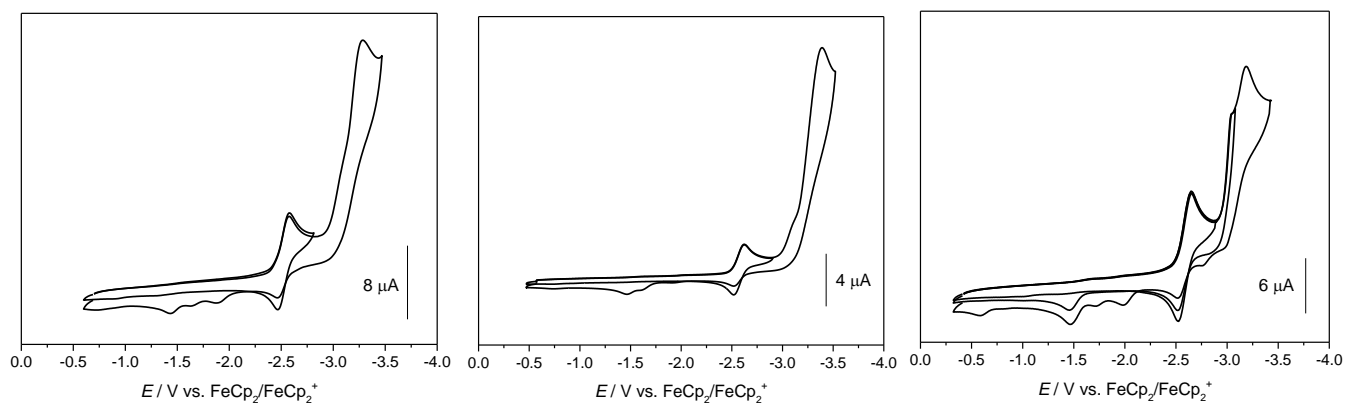


Fig. S23 Cyclic voltammograms of Hth(ppy)py (left), Hth(tbppy)py (middle), and Hbth(ppy)py (right) in 0.1 M *n*Bu₄NPF₆/THF.

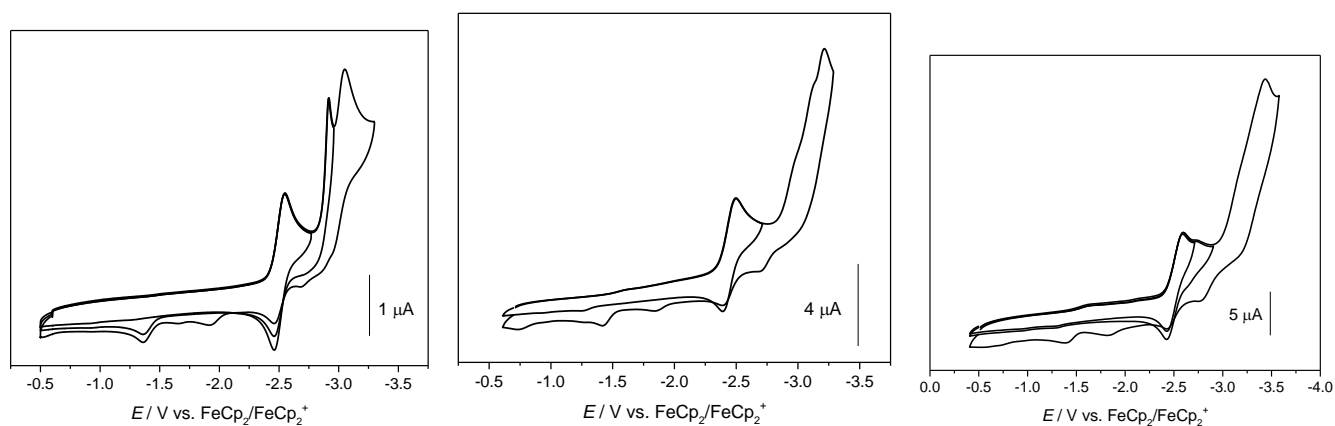


Fig. S24 Cyclic voltammograms of Hbth(tbppy)py (left), Hph(ppy)tz (middle), and Hph(tbppy)tz (right) in 0.1 M *n*Bu₄NPF₆/THF.

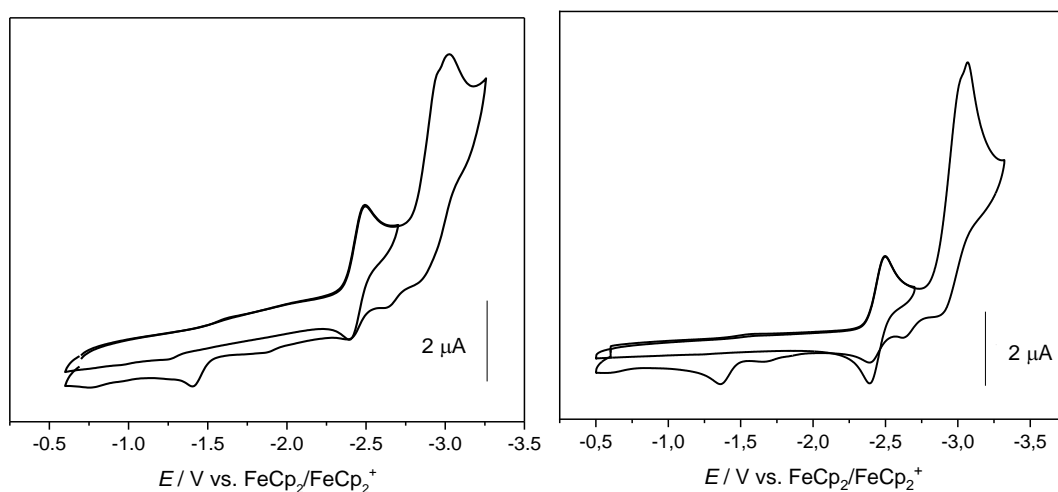


Fig. S25 Cyclic voltammograms of Hna(ppy)tz (left) and Hna(tbppy)tz (right) in 0.1 M *n*Bu₄NPF₆/THF.

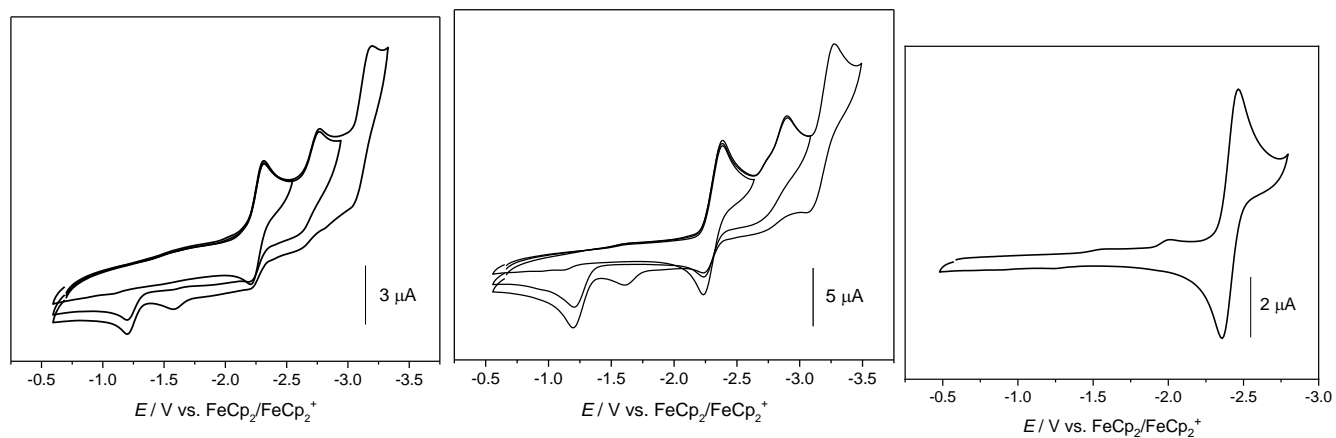


Fig. S26 Cyclic voltammograms of Hph(ppy)btz (left), Hph(tbppy)btz (middle), and Hth(tbppy)tz (right) in 0.1 M *n*Bu₄NPF₆/THF.

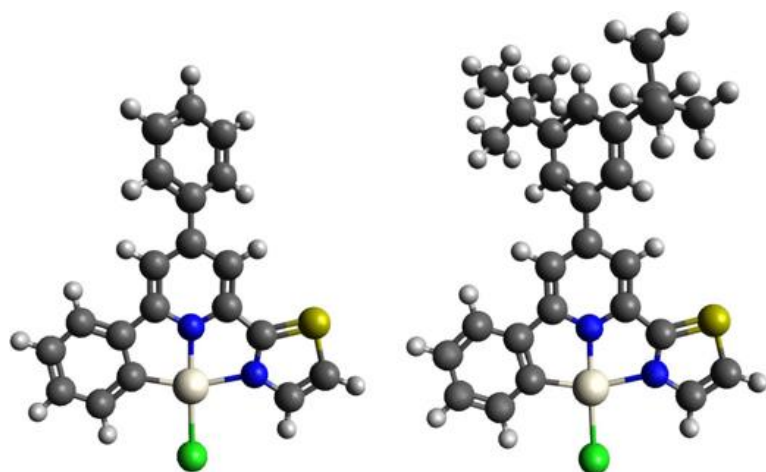


Fig. S27 DFT-optimised geometries of complexes from **class A** with coligand X = Cl and R¹ = H (left) and R¹ = 3,5-*t*Bu₂ (right).

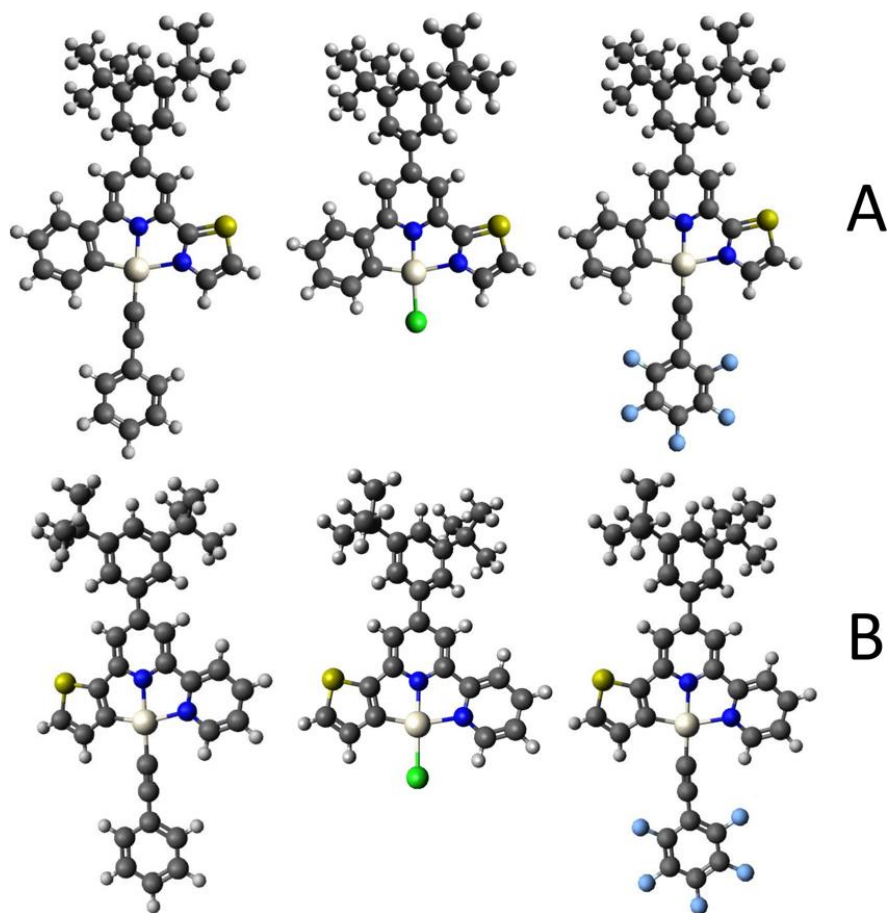


Fig. S28 DFT-optimised geometries of complexes from **classes A and B** with coligands $X = \text{CCPh}$, Cl , CCC_6F_5 (from left to right) and $R^1 = 3,5\text{-}t\text{Bu}$.

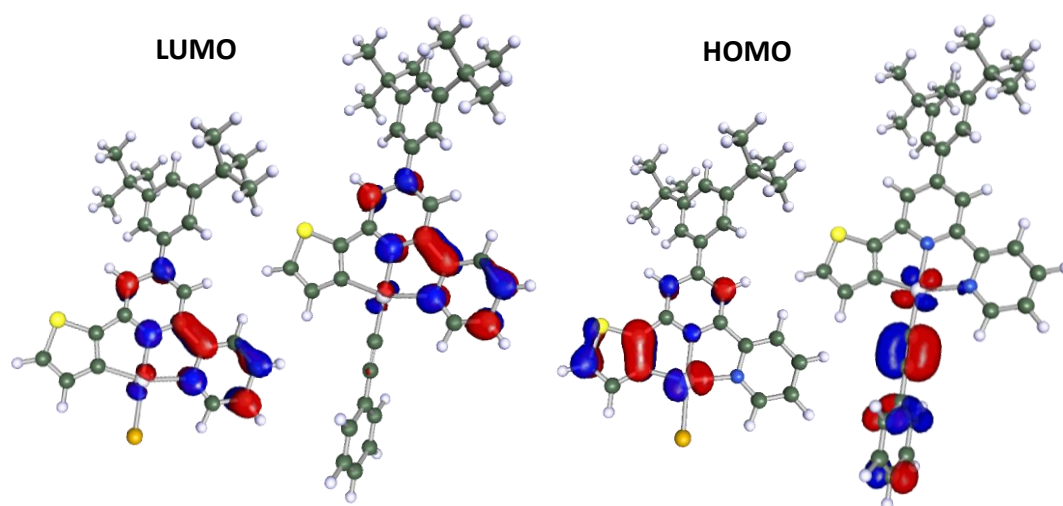


Fig. S29 DFT-calculated compositions of the lowest unoccupied molecular orbitals (LUMO) and highest occupied molecular orbitals (HOMO) of $[\text{Pt}(\text{th}(\text{tbpppy})\text{py})\text{Cl}]$ (**11**) and $[\text{Pt}(\text{th}(\text{tbpppy})\text{py})(\text{CCPh})]$ (**12**) calculated on B3LYP level using def2-TZV(P) for C, H, N, S und LAN-L2DZ basis sets for Pt (ECP: Hay/Wadt ($n = 1$)), isosurface at 0.05.

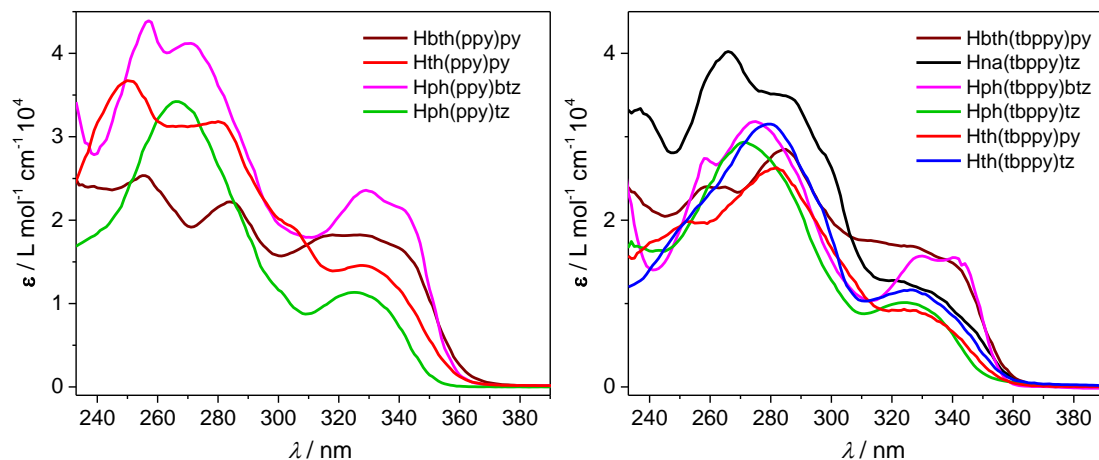


Fig. S30 UV-vis absorption spectra of HC^NN protoligands with 4-phenyl-pyridyl (Hppy) (left) and 3,5-*t*Bu₂-phenyl-pyridyl (Htbppy) (right) in CH₂Cl₂ (298 K).

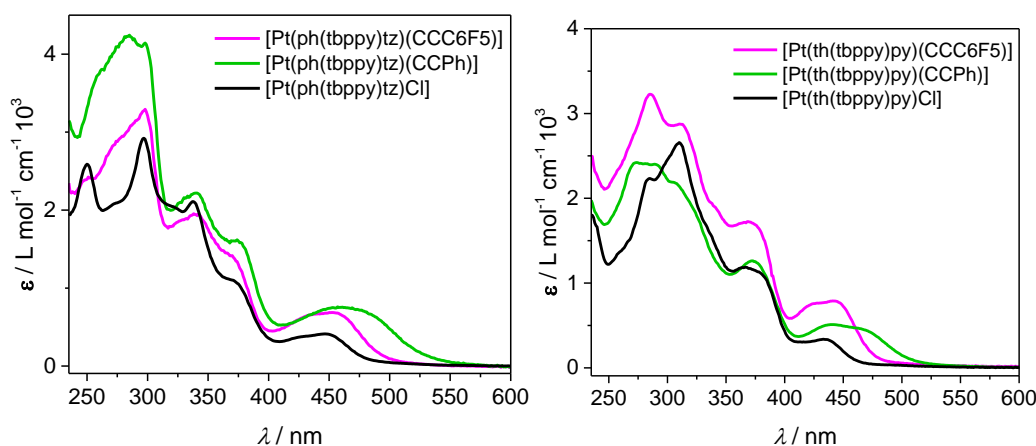


Fig. S31 UV-vis absorption spectra of [Pt(ph(tbppy)tz)(Cl/R)] (R = CCPh, CCC₆F₅) (**2**, **3**, **4**) (left) and [Pt(th(tbppy)py)(Cl/R)] (R = CCPh, CCC₆F₅) (**11**, **12**, **13**) (right) in CH₂Cl₂ (298 K).

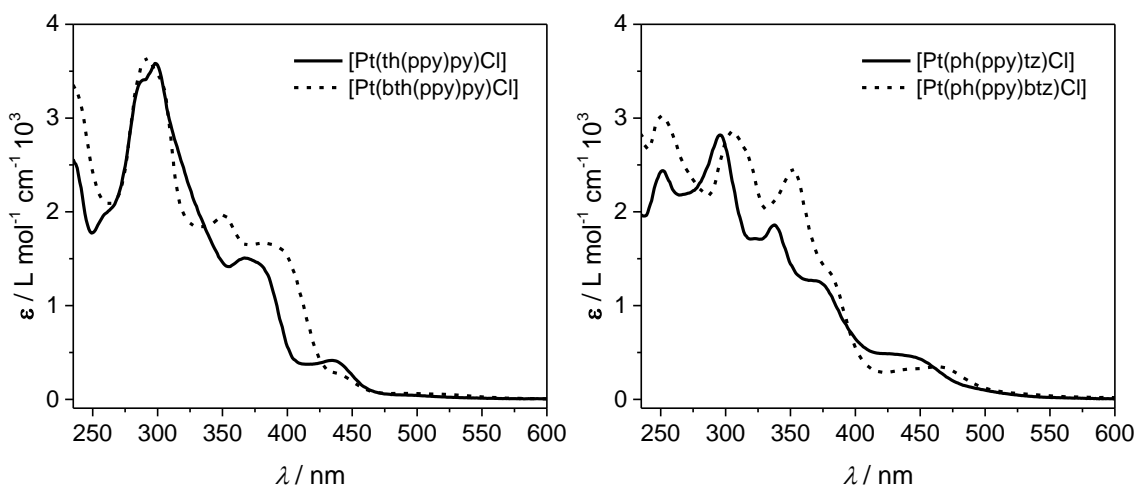


Fig. S32 UV-vis absorption spectra of [Pt(th(ppy)py)Cl] (**10**) and [Pt(bth(ppy)py)Cl] (**14**) (left) and [Pt(ph(ppy)tz)Cl] (**1**) and [Pt(ppy)btz)Cl] (**7**) (right) in CH₂Cl₂ (298 K).

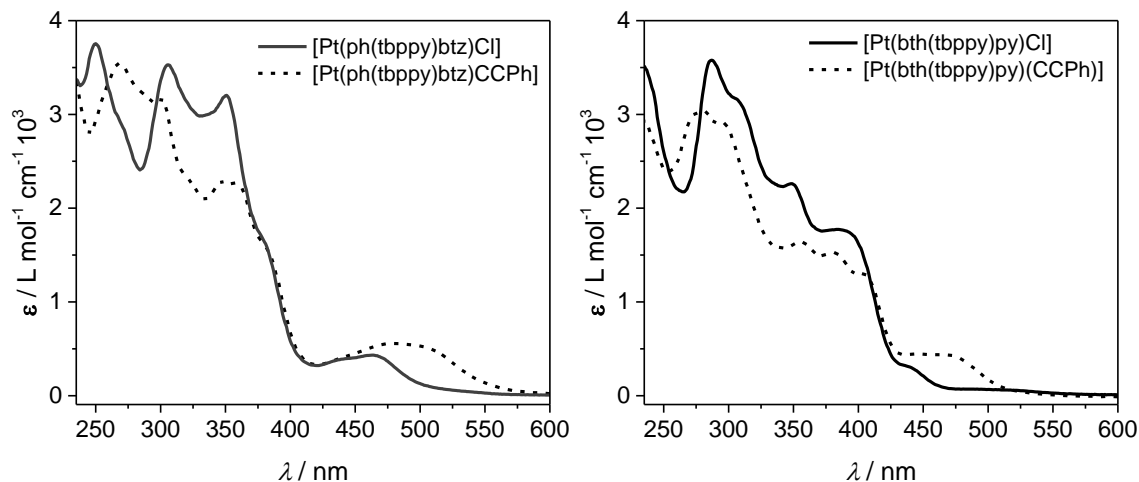


Fig. S33 UV-vis absorption spectra of [Pt(ph(tbppy)btz)Cl] (**8**) and [Pt(ph(tbppy)btz)(CCPh)] (**9**) (left) and [Pt(bth(tbppy)py)Cl] (**15**) and [Pt(bth(tbppy)py)(CCPh)] (**16**) (right) in CH_2Cl_2 (298 K).

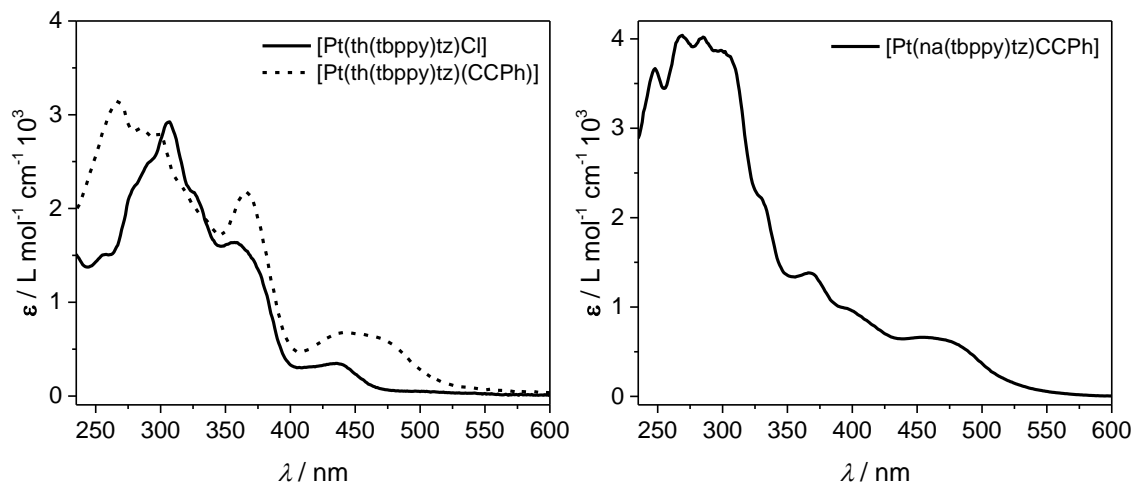


Fig. S34 UV-vis absorption spectra of [Pt(th(tbppy)tz)Cl] (**17**) and [Pt(th(tbppy)tz)(CCPh)] (**18**) (left) and [Pt(na(tbppy)tz)(CCPh)] (**6**) (right) in CH_2Cl_2 (298 K).

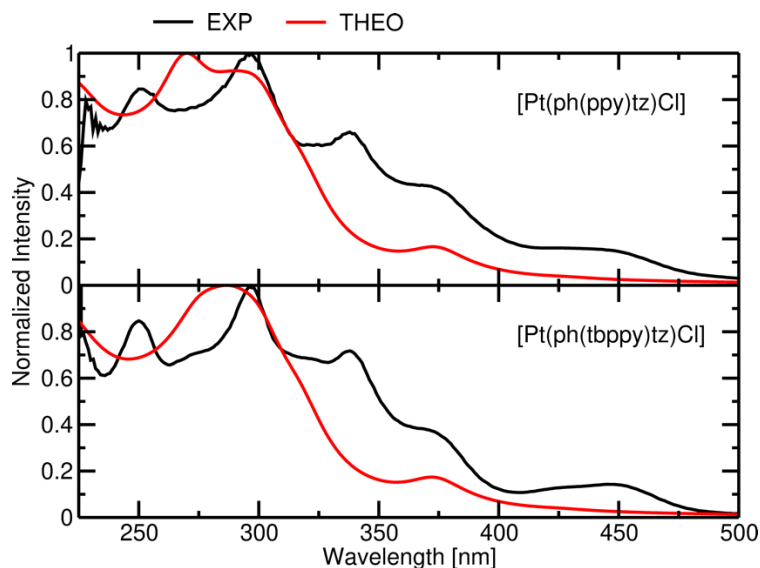


Fig. S35 TD-DFT-calculated UV-vis absorption spectra (THEO, red) compared to experimental spectra (EXP, black) of [Pt(ph(ppytz)Cl] (**1**) and [Pt(ph(tbppy)tz)Cl] (**2**).

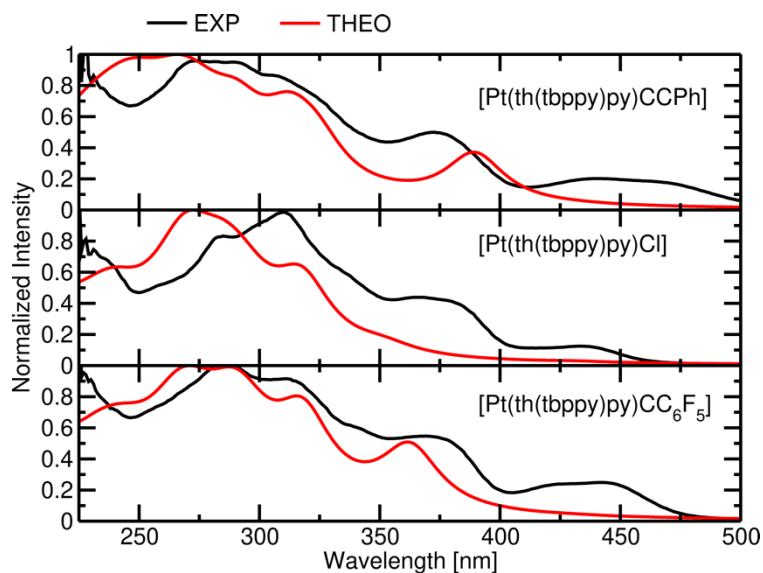


Fig. S36 TD-DFT-calculated UV-vis absorption spectra (THEO, red) compared to experimental spectra (EXP, black) of [Pt(th(tbppy)py)C≡CPh] (**12**), [Pt(ph(tbppy)tz)Cl] (**2**), and [Pt(th(tbppy)py)C≡CC₆F₅] (**13**).

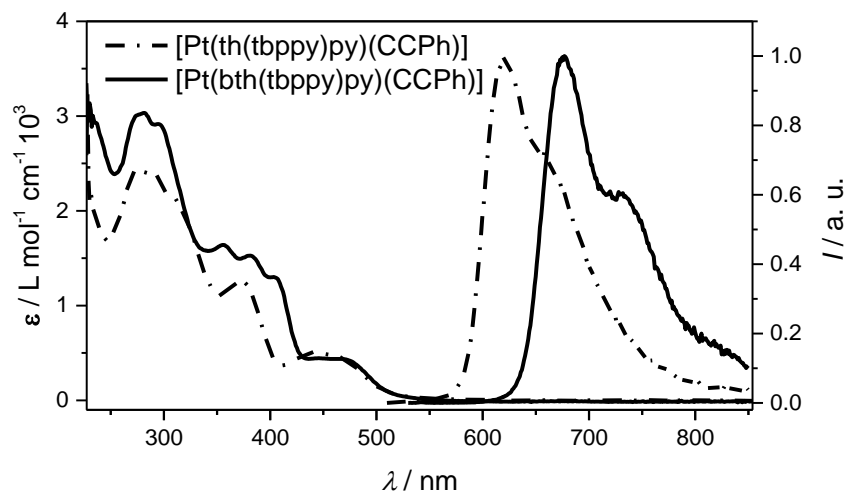


Fig. S37 UV-vis absorption spectra (left) and normalised emission spectra (right) of [Pt(th(tbppy)py)CCPh] (**12**) and [Pt(bth(tbppy)py)CCPh] (**16**) in CH₂Cl₂ at 298 K upon excitation at 350 nm.

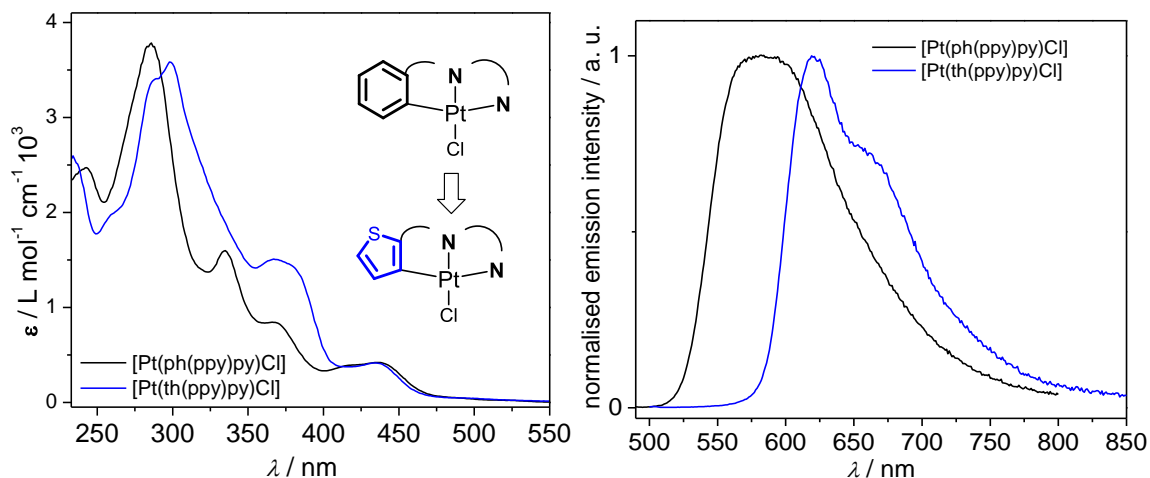


Fig. S38 UV-vis absorption spectra (left) and normalised emission spectra (right) of [Pt(th(ppy)py)Cl] (**10**) and [Pt(ph(ppy)py)Cl] in CH₂Cl₂ at 298 K upon excitation at 350 nm.

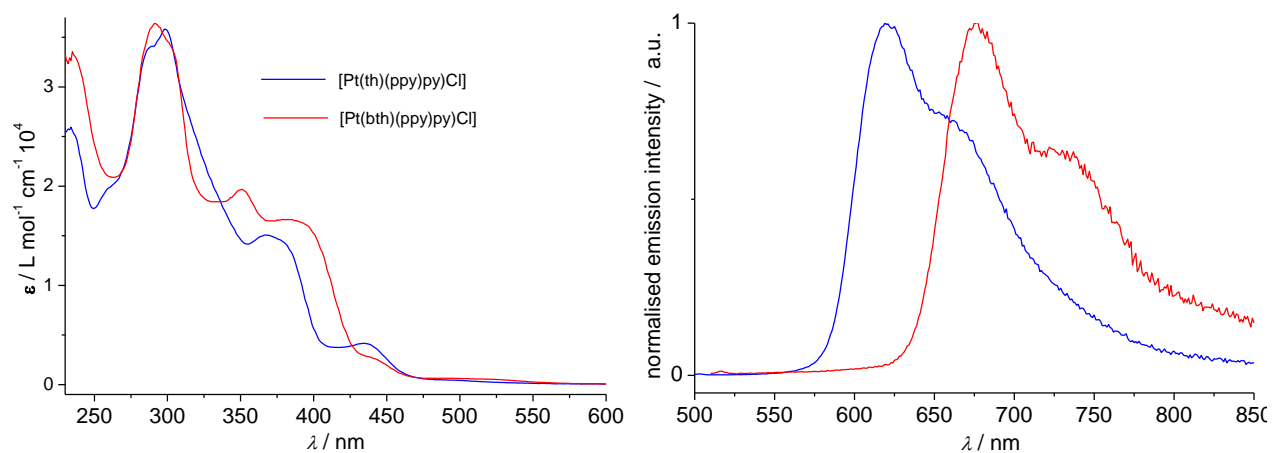


Fig. S39 UV-vis absorption spectra (left) and normalised emission spectra (right) of [Pt(th(tbppy)py)Cl] (**11**) and [Pt(bth(tbppy)py)Cl] (**15**) in CH₂Cl₂ at 298 K upon excitation at 350 nm.

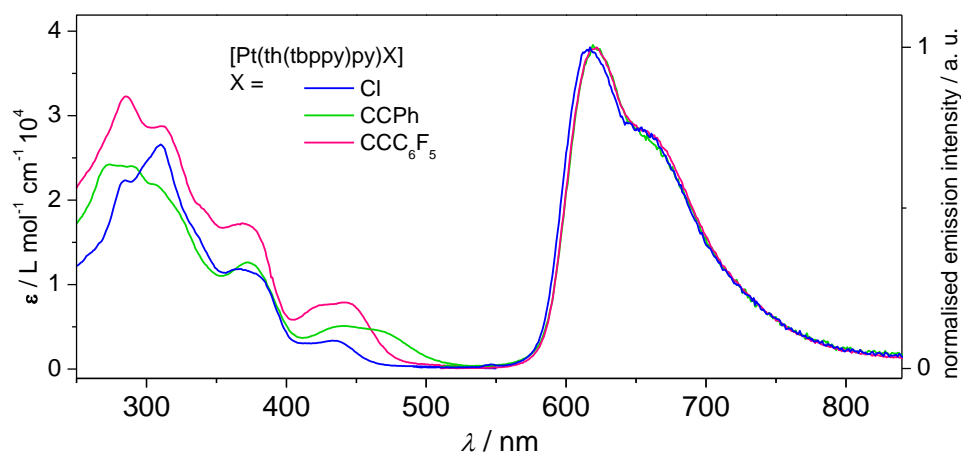


Fig. S40 UV-vis absorption spectra (left) and normalised emission spectra (right) of [Pt(th(tbppy)py)X] with X = Cl, CCPh, or CCC₆F₅ (**11**, **12**, **13**) in CH₂Cl₂ at 298 K upon excitation at 350 nm.

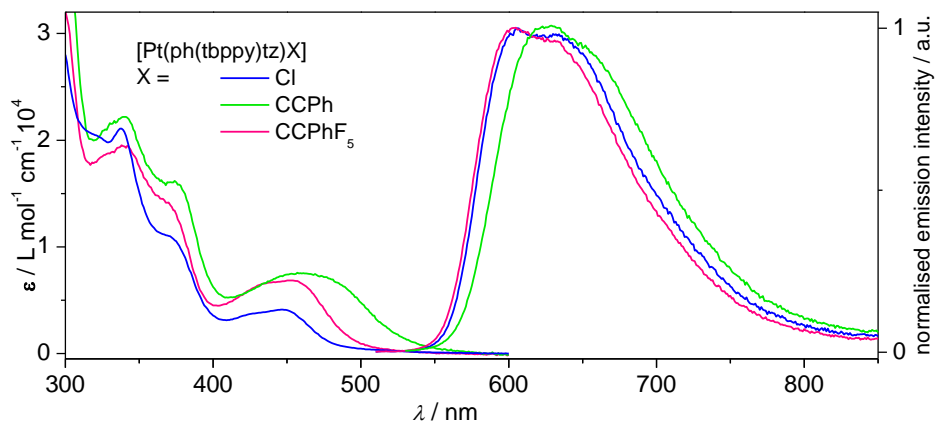


Fig. S41 UV-vis absorption spectra (left) and normalised emission spectra (right) of $[\text{Pt}(\text{ph}(\text{tbppy})\text{tz})\text{X}]$ with $\text{X} = \text{Cl}$, CCPh , or CCPhF_5 (**2**, **3**, **4**) in CH_2Cl_2 at 298 K upon excitation at 350 nm.

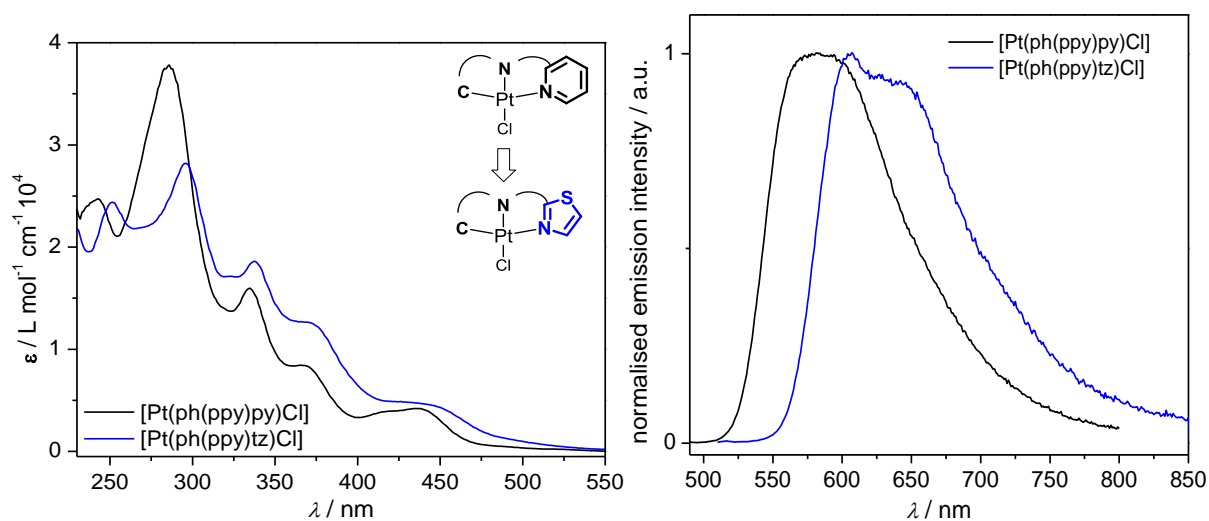


Fig. S42 UV-vis absorption spectra (left) and normalised emission spectra (right) of $[\text{Pt}(\text{ph}(\text{ppy})\text{py})\text{Cl}]$ and $[\text{Pt}(\text{ph}(\text{ppy})\text{tz})\text{Cl}]$ (**1**) in CH_2Cl_2 at 298 K upon excitation at 350 nm.

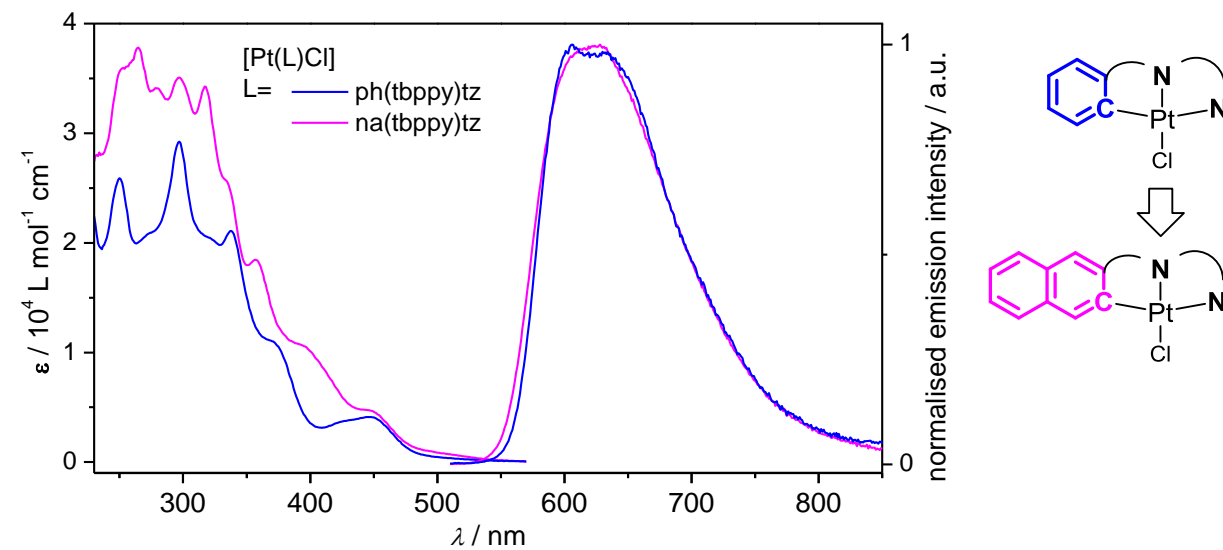


Fig. S43 UV-vis absorption spectra (left) and normalised emission spectra (right) of $[\text{Pt}(\text{ph}(\text{tbppy})\text{tz})\text{Cl}]$ (**2**) and $[\text{Pt}(\text{na}(\text{tbppy})\text{tz})\text{Cl}]$ (**5**) in CH_2Cl_2 at 298 K upon excitation at 350 nm.

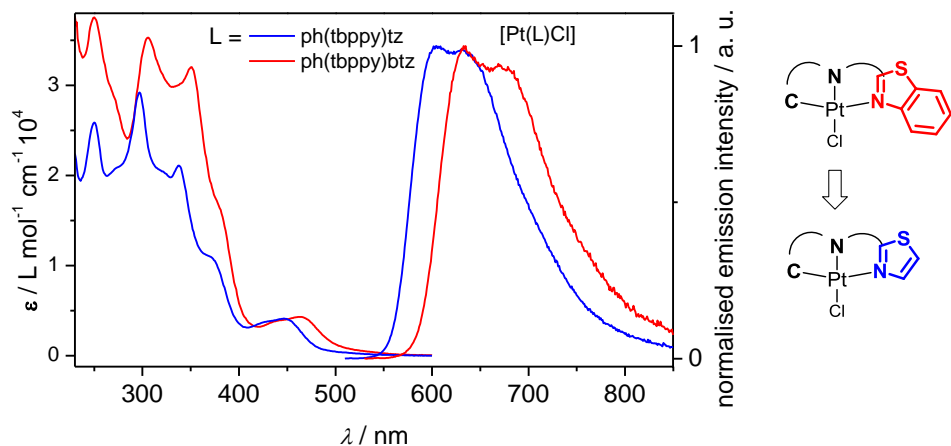


Fig. S44 UV-vis absorption spectra (left) and normalised emission spectra (right) of [Pt(ph(tbppy)tz)Cl] (**2**) and [Pt(ph(tbppy)btz)Cl] (**8**) in CH₂Cl₂ at 298 K upon excitation at 350 nm.

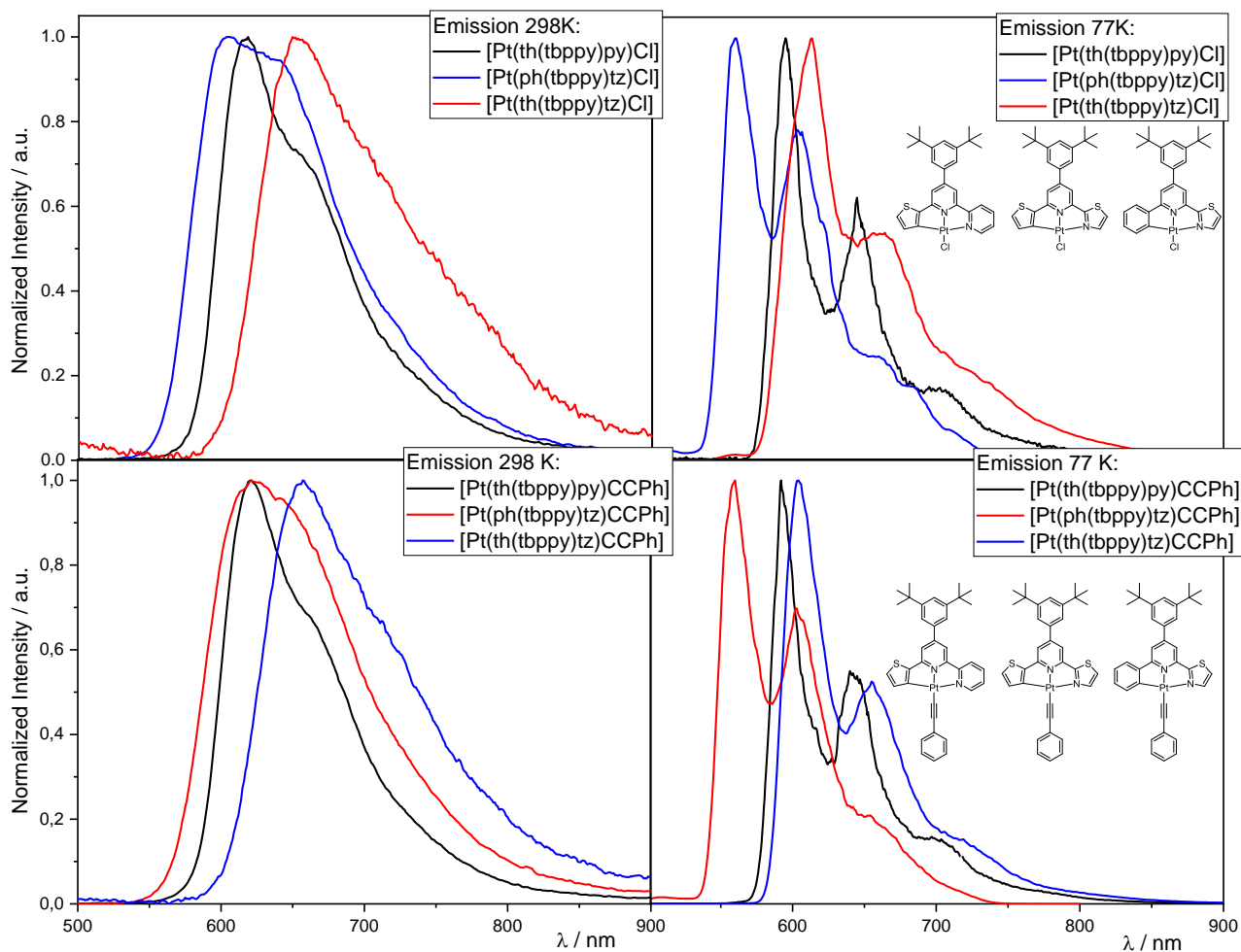


Fig. S45 Normalised photoluminescence spectra ($\lambda_{\text{exc}} = 350 \text{ nm}$) with focus on the influence of the peripheral ring; top: [Pt(th(tbppy)py)Cl] (**11**) (black), [Pt(ph(tbppy)tz)Cl] (**2**) (blue) and [Pt(th(tbppy)tz)Cl] (**17**) (red); bottom: [Pt(th(tbppy)py)CCPh] (**12**) (black), [Pt(ph(tbppy)tz)CCPh] (**3**) (red) and [Pt(th(tbppy)tz)CCPh] (**18**) (blue). Measured in CH₂Cl₂ at 298 K (left) or frozen glassy CH₂Cl₂/MeOH matrices at 77 K (right).

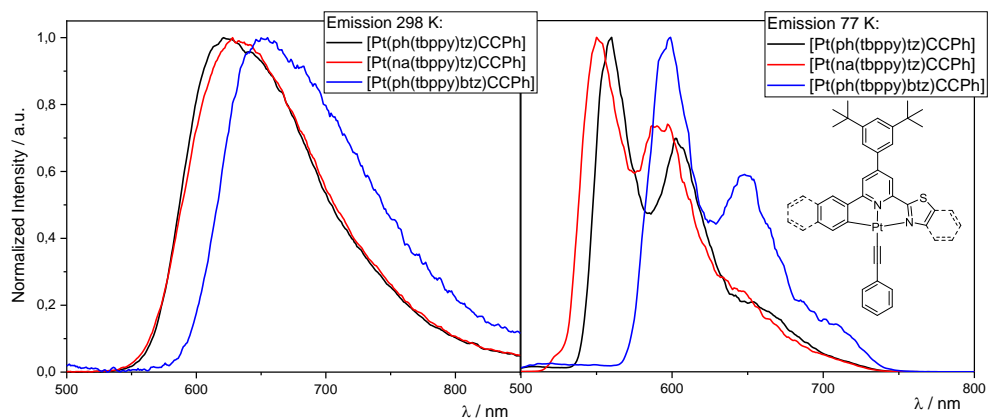


Fig. S46 Normalised photoluminescence spectra ($\lambda_{exc} = 350$ nm) involving the main ligands used for **class A** complexes (see Table 1); [Pt(ph(tbppytz)CCPh) (**3**) (black), [Pt(na(tbppytz)CCPh) (**6**) (red) and [Pt(ph(tbppytz)CCPh) (**9**) (blue) in CH_2Cl_2 at 298 K (left) and in frozen glassy $\text{CH}_2\text{Cl}_2/\text{MeOH}$ matrix (V:V = 1:1) at 77 K (right).

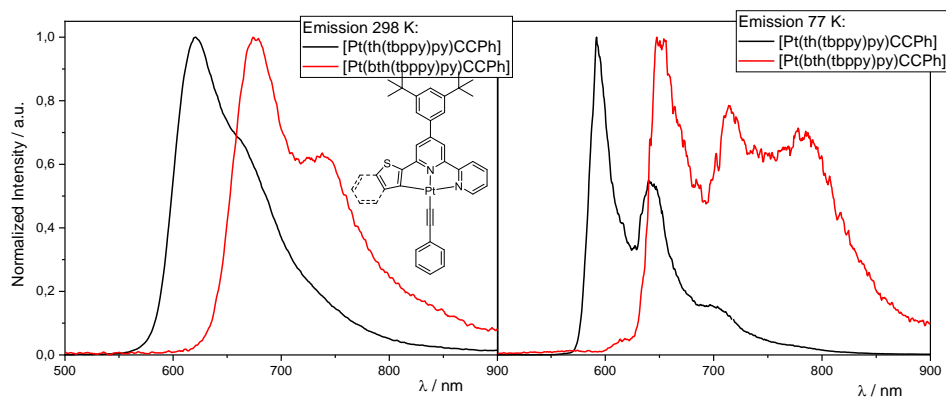


Fig. S47 Normalised photoluminescence spectra ($\lambda_{exc} = 350$ nm) involving the main ligands used for **class B** complexes (see Table 1) [Pt(th(tbppypy)CCPh) (**12**) (black) and [Pt(bth(tbppypy)CCPh) (**16**) (red) in fluid CH_2Cl_2 at 298 K (left) and in frozen glassy $\text{CH}_2\text{Cl}_2/\text{MeOH}$ matrices (V:V = 1:1) at 77 K (right).

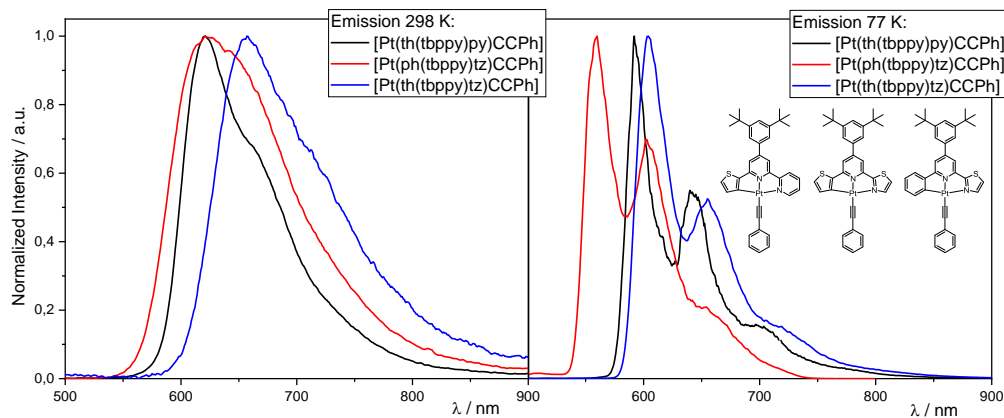


Fig. S48 Normalised photoluminescence spectra ($\lambda_{exc} = 350$ nm) of [Pt(th(tbppypy)CCPh) (**12**) (black), [Pt(ph(tbppytz)CCPh) (**3**) (red) and [Pt(th(tbppytz)CCPh) (**18**) (blue) in CH_2Cl_2 at 298 K (left) and in frozen glassy $\text{CH}_2\text{Cl}_2/\text{MeOH}$ matrices (V:V = 1:1) (right) at 77 K.

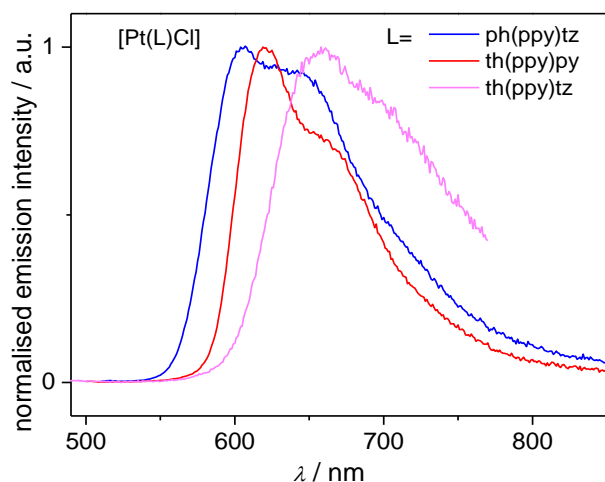


Fig. S49 Normalised emission spectra of [Pt(ph(ppy)tz)Cl] (**1**) (blue), [Pt(th(ppy)py)Cl] (**10**) (red) and [Pt(th(ppy)tz)Cl] (violet) in CH₂Cl₂ at 298 K upon excitation at 350 nm.

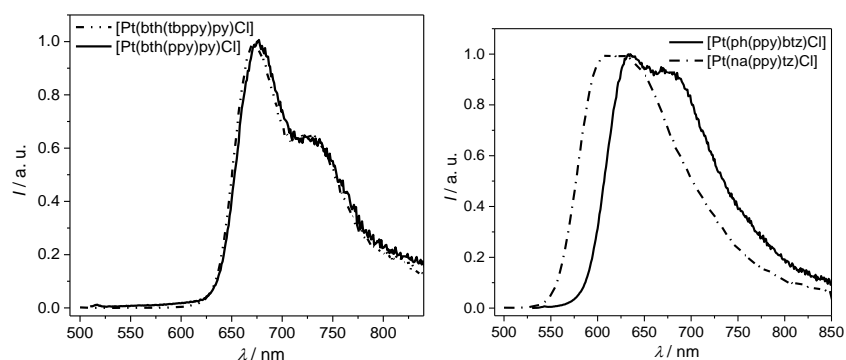


Fig. S50 Normalised emission spectra of, [Pt(bth(tbppy)py)Cl] (**15**), [Pt(bth(ppy)py)Cl] (**14**) (left), and [Pt(ph(ppy)btz)Cl] (**7**), [Pt(na(ppy)tz)Cl] (right), in CH₂Cl₂ at 298 K upon excitation with 350 nm.

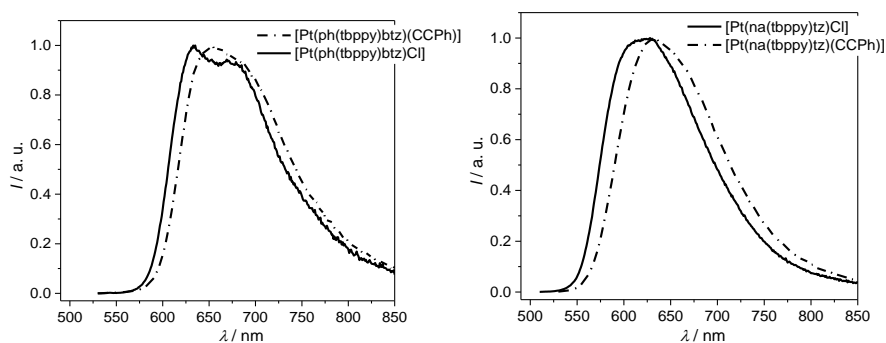


Fig. S51 Normalised emission spectra of (left), [Pt(ph(tbppy)btz)(CCPh)] (**9**), [Pt(ph(tbppy)btz)Cl] (**8**) (left), and [Pt(na(tbppy)tz)Cl] (**5**), [Pt(na(tbppy)tz)(CCPh)] (**6**) (right), in CH₂Cl₂ at 298 K upon excitation with 350 nm.

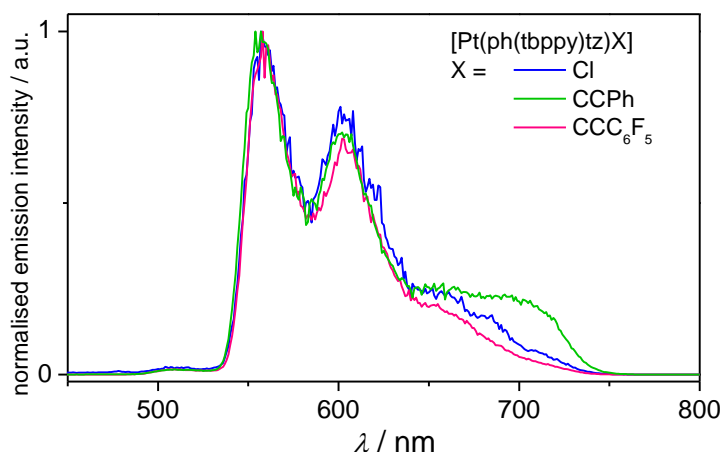


Fig. S52 Normalised emission spectra of [Pt(ph(tbppy)tz)X] with X = Cl, CCPh, or CCC₆F₅ (**2**), (**3**) and (**4**) in glassy frozen CH₂Cl₂/MeOH (1:1) matrix at 77 K upon excitation at 350 nm.

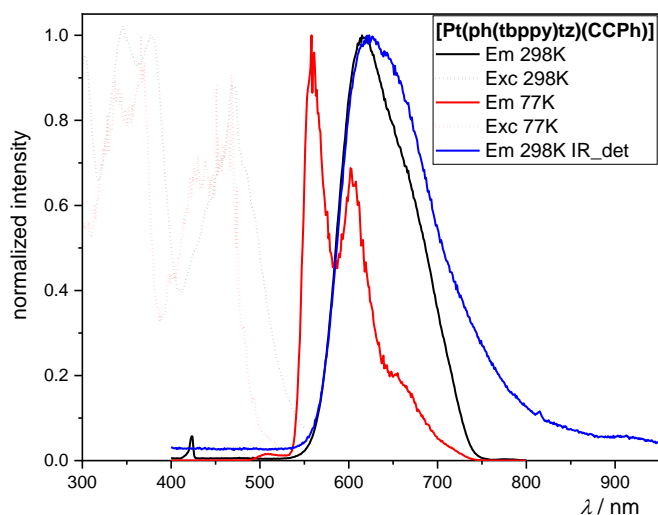


Fig. S53 Excitation (dotted line) and emission spectra (solid line) of [Pt(ph(tbppy)tz)(CCPh)] (**3**) ($\lambda_{exc} = 350$ nm; $\lambda_{em} = 590$ nm at 298 K, $\lambda_{em} = 560$ nm at 77 K) at 298 K with the UV/Vis-detector (black) and with IR-detector (blue) in CH₂Cl₂ and at 77 K (red) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1). All solutions were optically diluted ($A < 0.1$). Spectra normalised to the highest intensity.

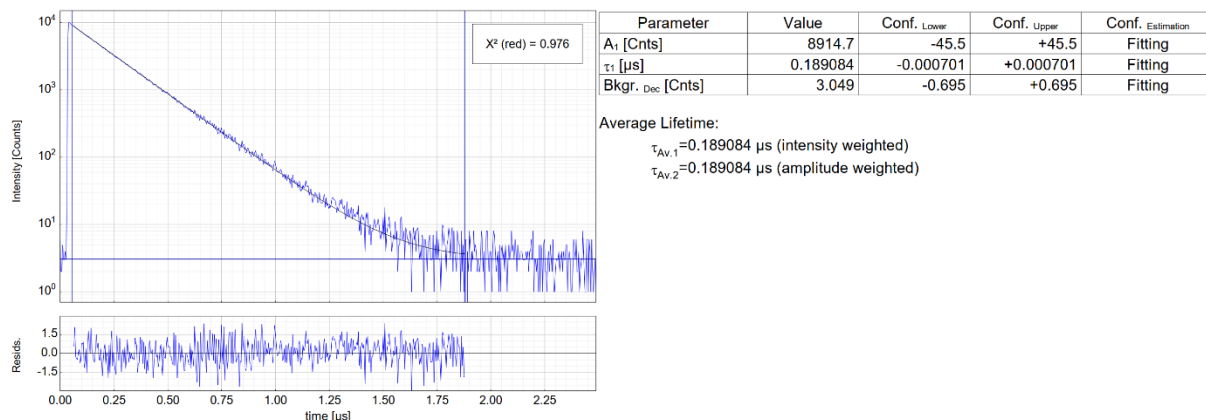


Fig. S54 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(ph(tbppy)tz)CCPh] (**3**) in CH₂Cl₂ at 298 K (air-equilibrated), including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 630$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

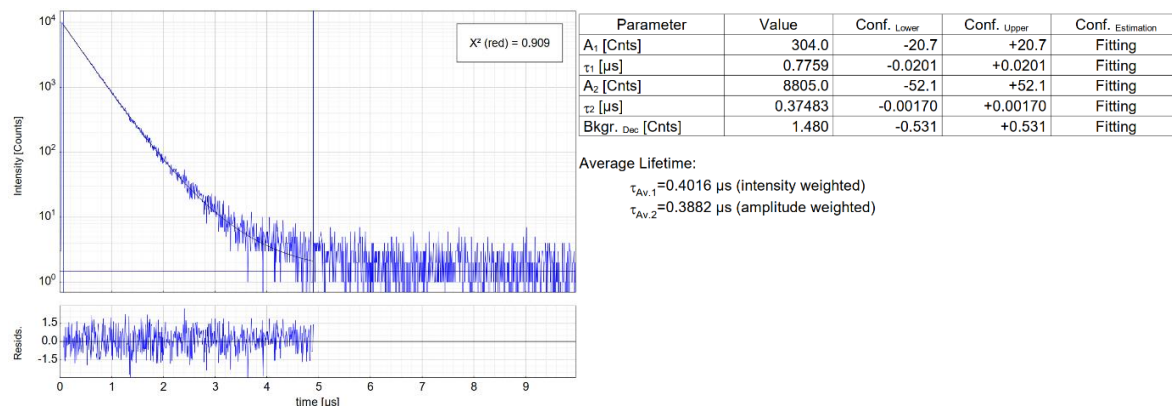


Fig. S55 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(ph(tbppy)tz)CCPh] (**3**) in CH₂Cl₂ at 298 K (Ar-purged), including the residuals ($\lambda_{\text{exc}} = 376.7$ nm, $\lambda_{\text{em}} = 630$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

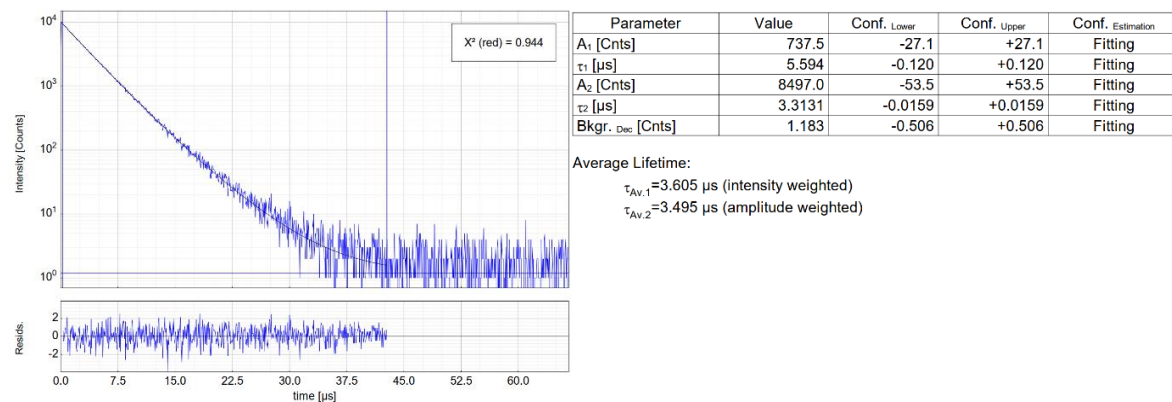


Fig. S56 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(ph(tbppy)tz)CCPh] (**3**) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1) at 77 K, including the residuals ($\lambda_{\text{exc}} = 376.7$ nm, $\lambda_{\text{em}} = 560$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

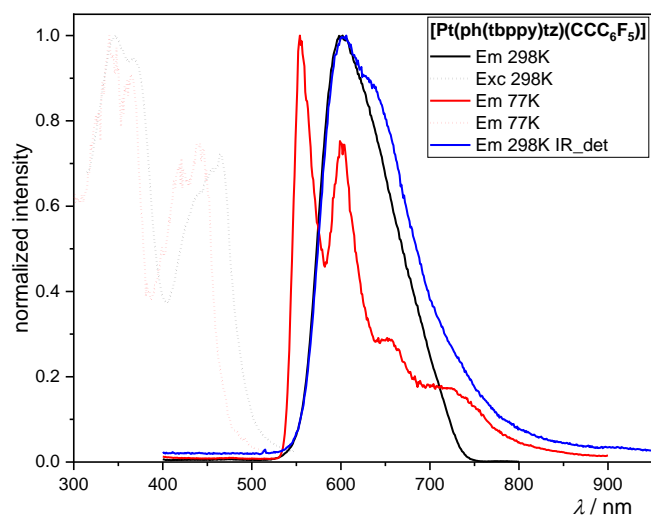


Fig. S57 Excitation (dotted line) and emission spectra (solid line) of [Pt(ph(tbppy)tz)(CCC₆F₅)] (**4**) ($\lambda_{\text{exc}} = 350$ nm; $\lambda_{\text{em}} = 580$ nm at 298 K, $\lambda_{\text{em}} = 560$ nm at 77 K) at 298 K with the UV/Vis-detector (black) and with IR-detector (blue) in CH₂Cl₂ and at 77 K (red) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1). All solutions were optically diluted ($A < 0.1$). Spectra normalised to the highest intensity.

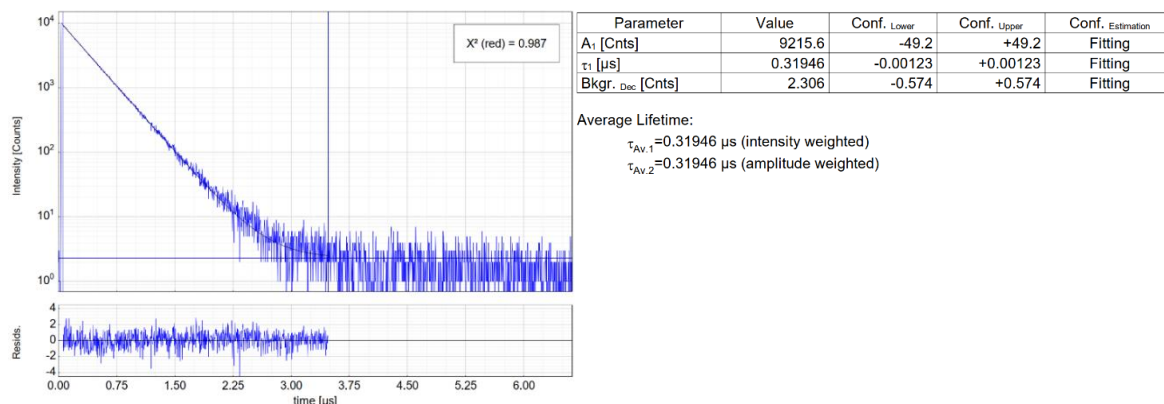


Fig. S58 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(ph(tbppy)tz)CC₆F₅] (**4**) in CH₂Cl₂ at 298 K (air-equilibrated), including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 600$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

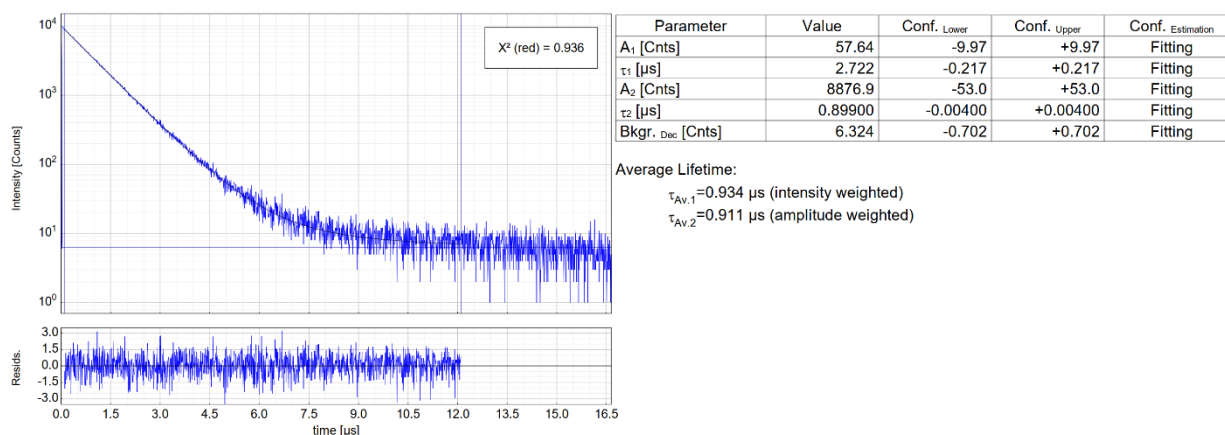


Fig. S59 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(ph(tbppy)tz)CC₆F₅] (**4**) in CH₂Cl₂ at 298 K (Ar-purged), including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 600$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

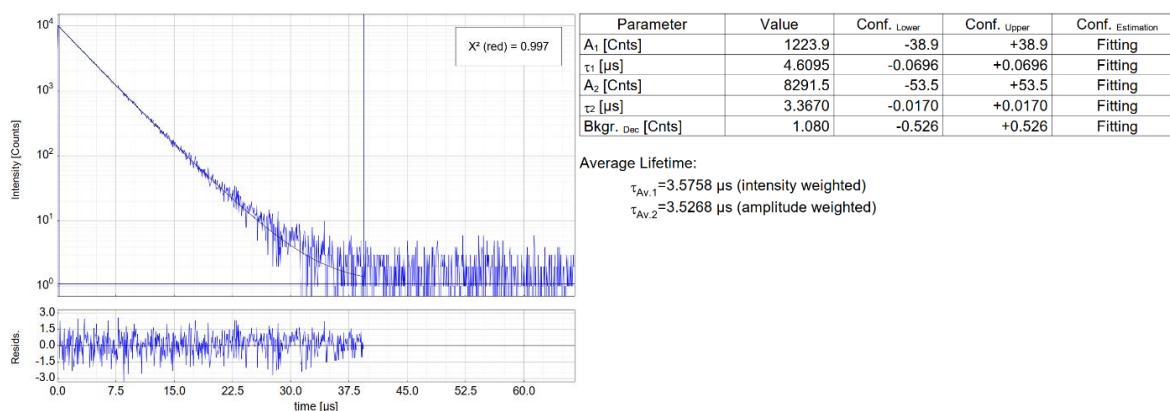


Fig. S60 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(ph(tbppy)tz)CC₆F₅] (**4**) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1) at 77 K, including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 560$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

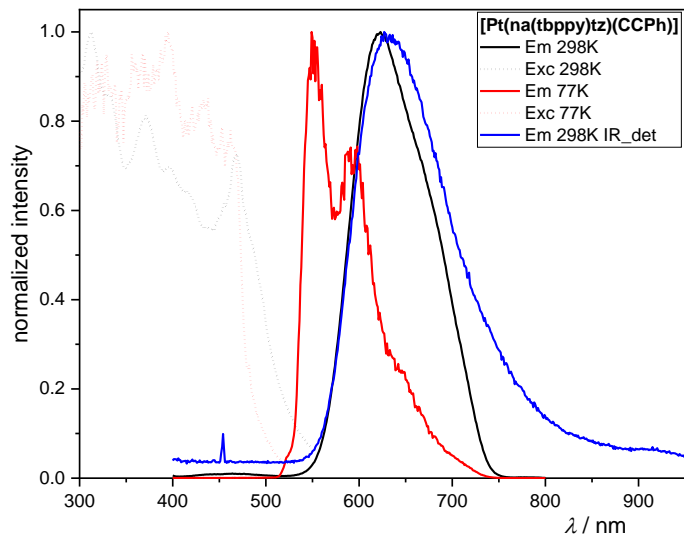


Fig. S61 Excitation (dotted line) and emission spectra (solid line) of [Pt(na(tbppy)tz)(CCPh)] (**6**) ($\lambda_{exc} = 350$ nm; $\lambda_{em} = 590$ nm at 298 K, $\lambda_{em} = 550$ nm at 77 K) at 298 K with the UV/Vis-detector (black) and with IR-detector (blue) in CH_2Cl_2 and at 77 K (red) in a frozen glassy $\text{CH}_2\text{Cl}_2/\text{MeOH}$ matrix ($V:V = 1:1$). All solutions were optically diluted ($A < 0.1$). Spectra normalised to the highest intensity.

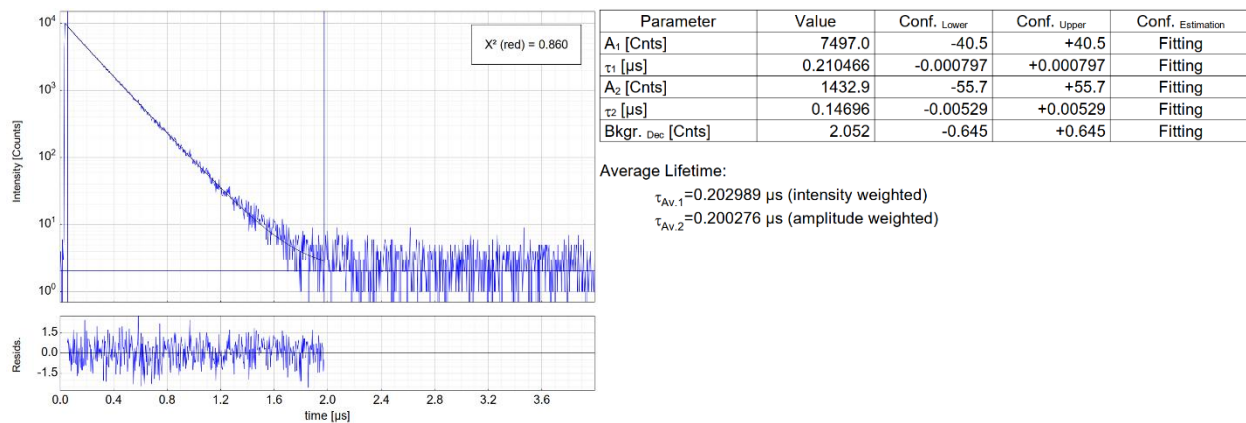


Fig. S62 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(na(tbppy)tz)CCPh] (**6**) in CH_2Cl_2 at 298 K (air-equilibrated), including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 620$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

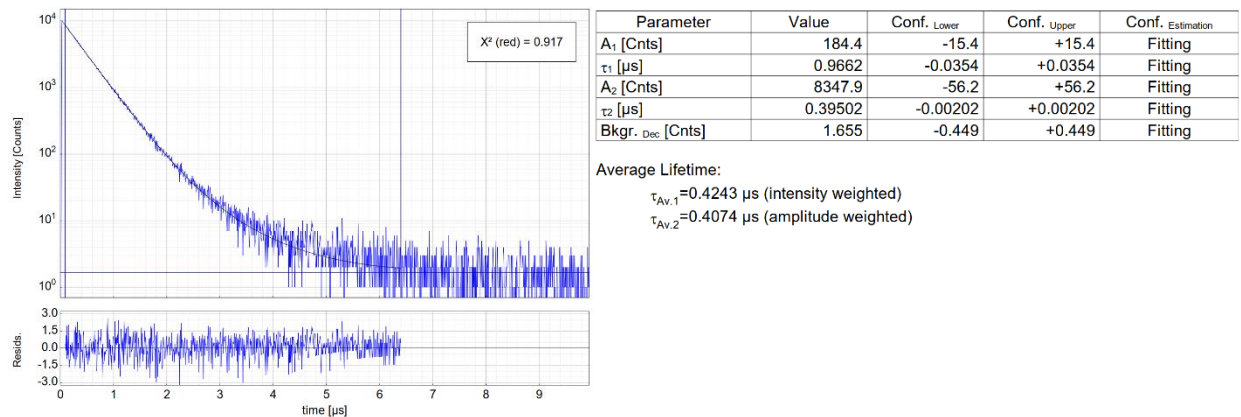


Fig. S63 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(na(tbppy)tz)CCPh] (**6**) in CH_2Cl_2 at 298 K (Ar-purged), including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 620$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

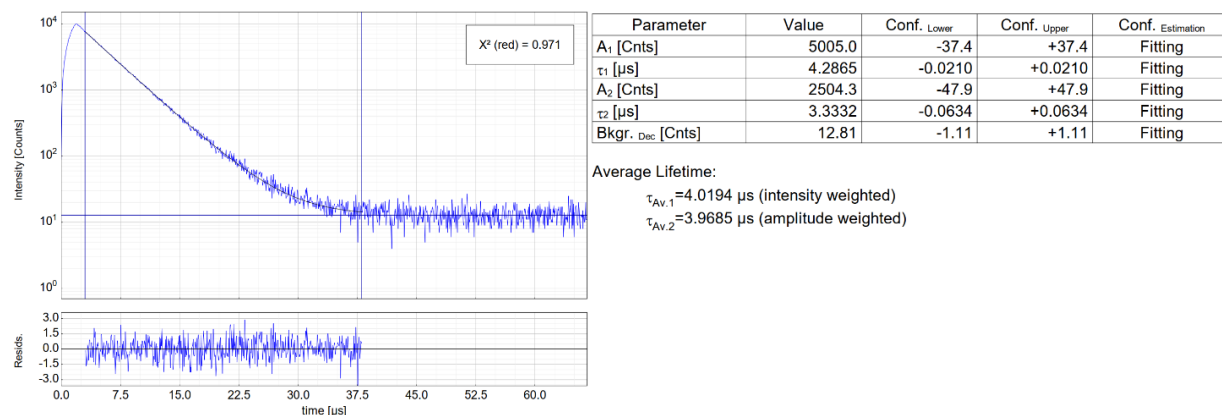


Fig. S64 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(na(tbppy)tz)CCPh] (**6**) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1) at 77 K, including the residuals ($\lambda_{\text{exc}} = 376.7$ nm, $\lambda_{\text{em}} = 550$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

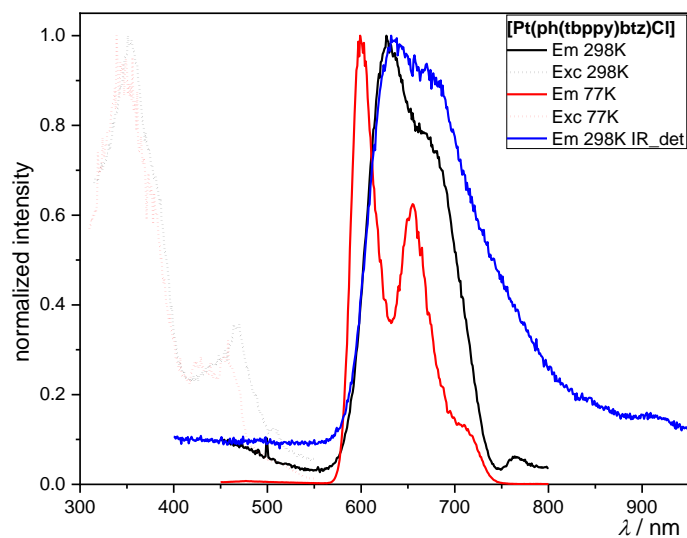


Fig. S65 Excitation (dotted line) and emission spectra (solid line) of [Pt(ph(tbppy)btz)Cl] (**8**) ($\lambda_{\text{exc}} = 350$ nm; $\lambda_{\text{em}} = 600$ nm) at 298 K with the UV/Vis-detector (black) and with IR-detector (blue) in CH₂Cl₂ and at 77 K (red) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1). All solutions were optically diluted ($A < 0.1$). Spectra normalised to the highest intensity.

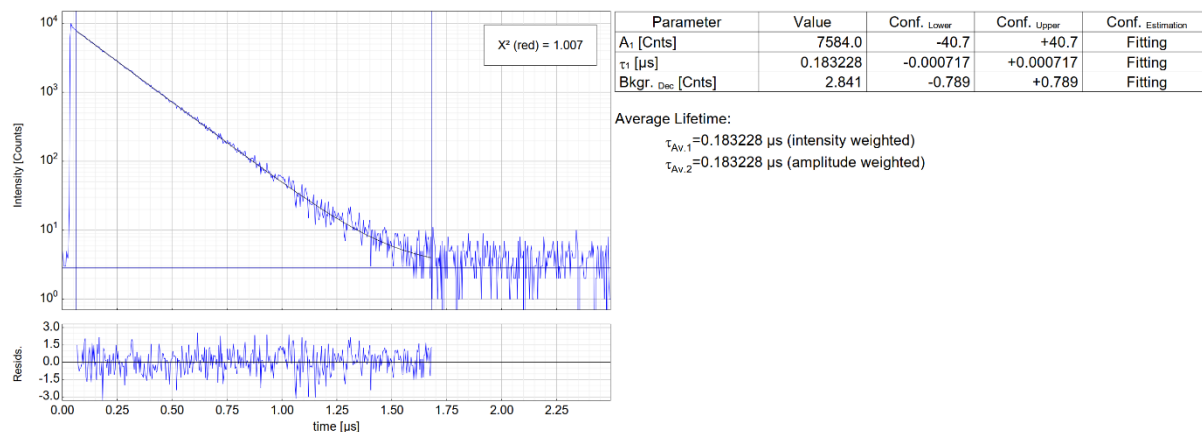


Fig. S66 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(ph(tbppy)btz)Cl] (**8**) in CH₂Cl₂ at 298 K (air-equilibrated), including the residuals ($\lambda_{\text{exc}} = 376.7$ nm, $\lambda_{\text{em}} = 630$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

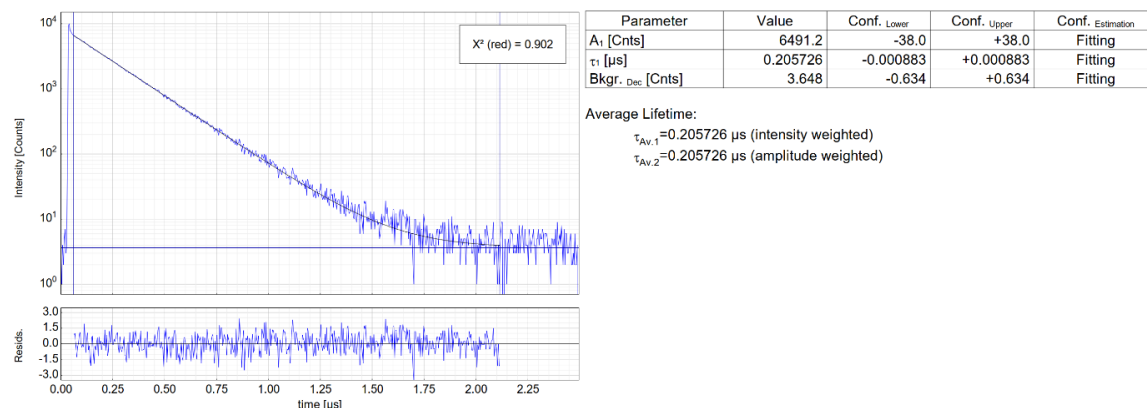


Fig. S67 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(ph(tbppy)btz)Cl] (**8**) in CH₂Cl₂ at 298 K (Ar-purged), including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 630$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

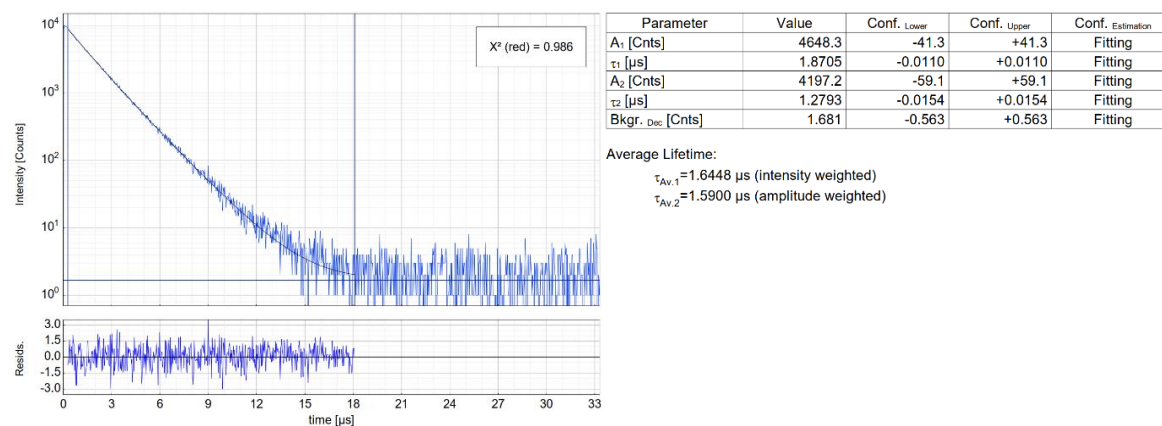


Fig. S68 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(ph(tbppy)btz)Cl] (**8**) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1) at 77 K, including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = \text{nm}$). Right: Fitting parameters including pre-exponential factors and confidence limits.

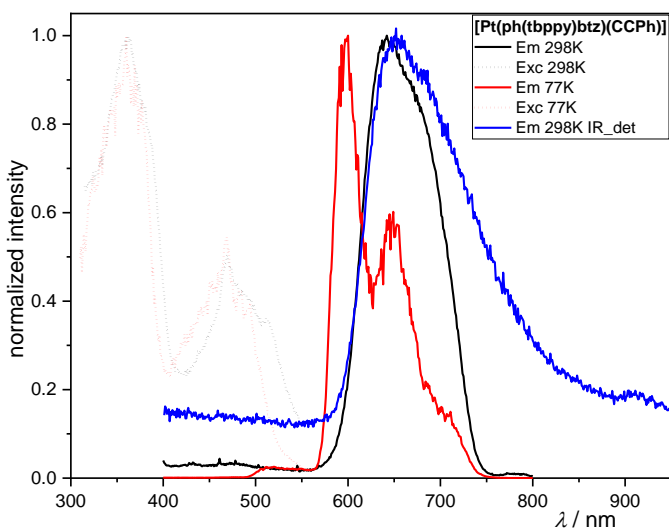


Fig. S69 Excitation (dotted line) and emission spectra (solid line) of [Pt(ph(tbppy)btz)CCPh] (**9**) ($\lambda_{exc} = 350$ nm; $\lambda_{em} = 610$ nm at 298 K, $\lambda_{em} = 600$ nm at 77 K) at 298 K with the UV/Vis-detector (black) and with IR-detector (blue) in CH₂Cl₂ and at 77 K (red) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1). All solutions were optically diluted ($A < 0.1$). Spectra normalised to the highest intensity.

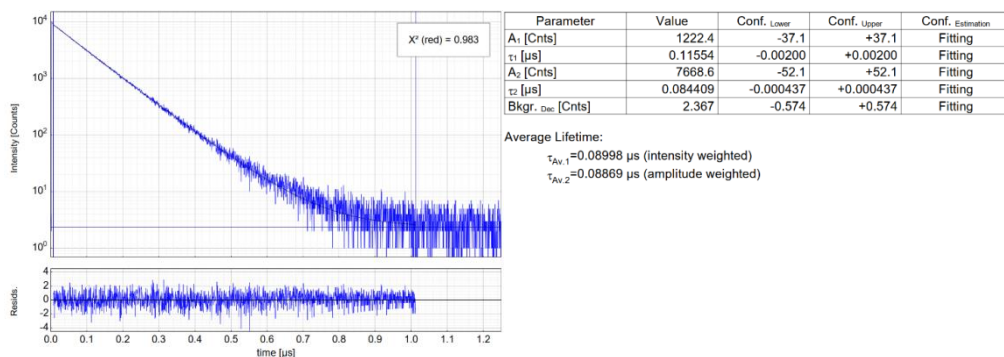


Fig. S70 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(ph(tbppy)btz)CCPh] (9) in CH₂Cl₂ at 298 K (air-equilibrated), including the residuals ($\lambda_{exc} = 376.7 \text{ nm}$, $\lambda_{em} = 640 \text{ nm}$). Right: Fitting parameters including pre-exponential factors and confidence limits.

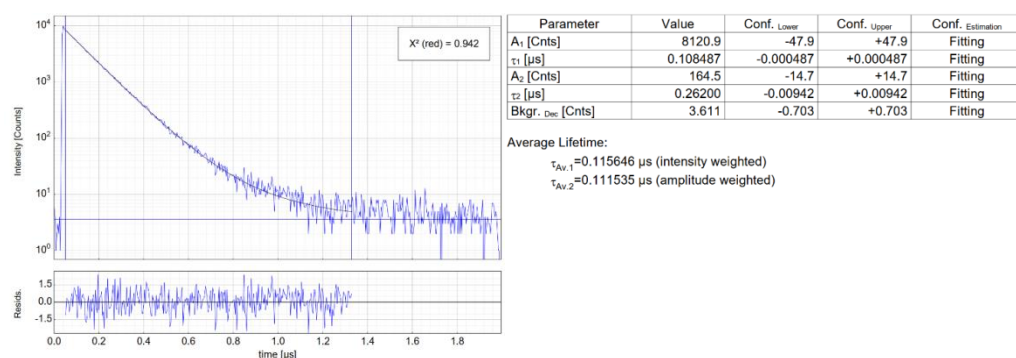


Fig. S71 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(ph(tbppy)btz)CCPh] (9) in CH₂Cl₂ at 298 K (Ar-purged), including the residuals ($\lambda_{exc} = 376.7 \text{ nm}$, $\lambda_{em} = 640 \text{ nm}$). Right: Fitting parameters including pre-exponential factors and confidence limits.

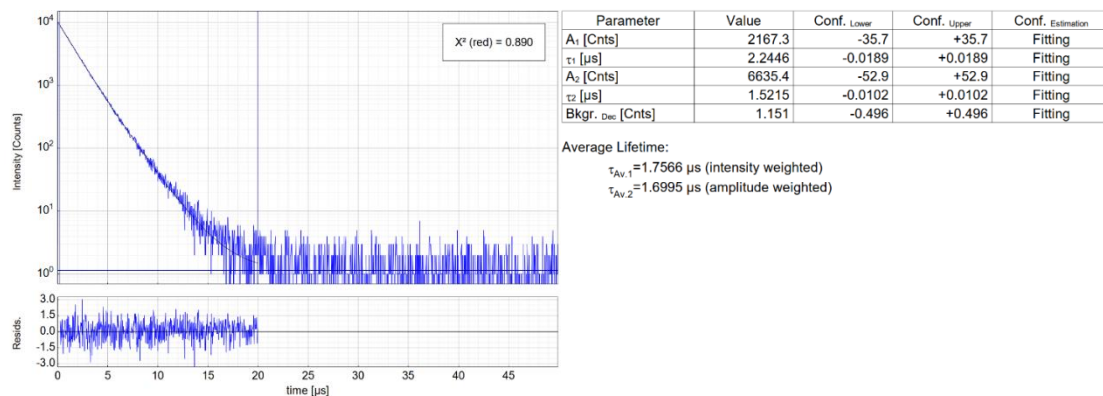


Fig. S72 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(ph(tbppy)btz)CCPh] (9) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1) at 77 K, including the residuals ($\lambda_{exc} = 376.7 \text{ nm}$, $\lambda_{em} = 600 \text{ nm}$). Right: Fitting parameters including pre-exponential factors and confidence limits.

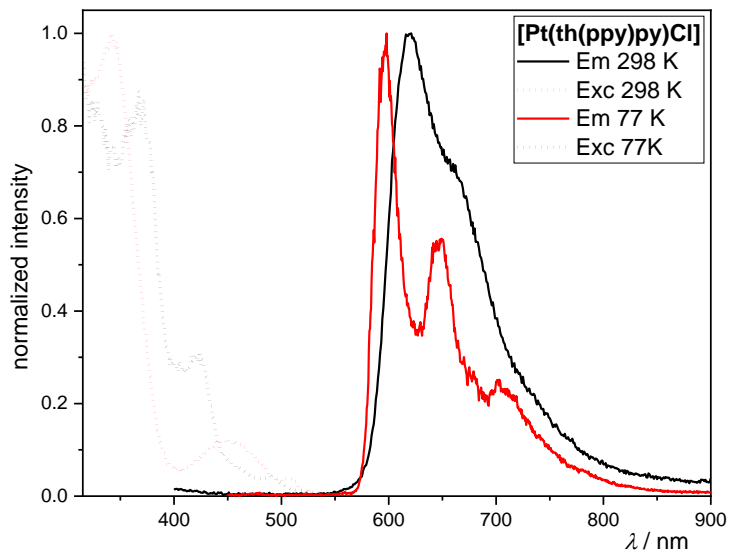


Fig. S73 Excitation (dotted line) and emission spectra (solid line) of [Pt(th(ppy)py)Cl] (**10**) ($\lambda_{\text{exc}} = 350 \text{ nm}$; $\lambda_{\text{em}} = 600 \text{ nm}$) at 298 K (black) in CH_2Cl_2 and at 77 K (red) in a frozen glassy $\text{CH}_2\text{Cl}_2/\text{MeOH}$ matrix (V:V = 1:1). All solutions were optically diluted ($A < 0.1$). Normalised to the highest intensity.

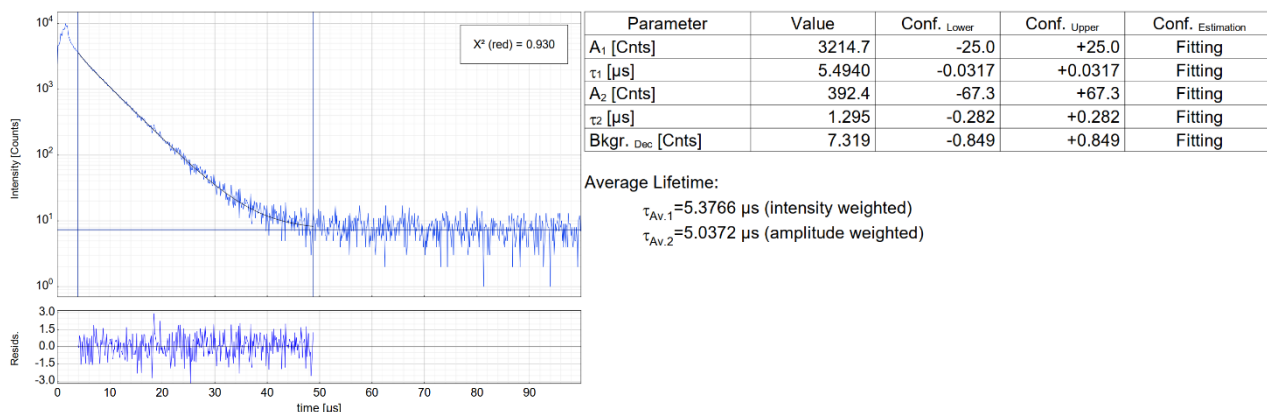


Fig. S74 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(ppy)py)Cl] (**10**) in CH_2Cl_2 at 298 K (Ar-purged), including the residuals ($\lambda_{\text{exc}} = 376.7 \text{ nm}$, $\lambda_{\text{em}} = 620 \text{ nm}$). Right: Fitting parameters including pre-exponential factors and confidence limits.

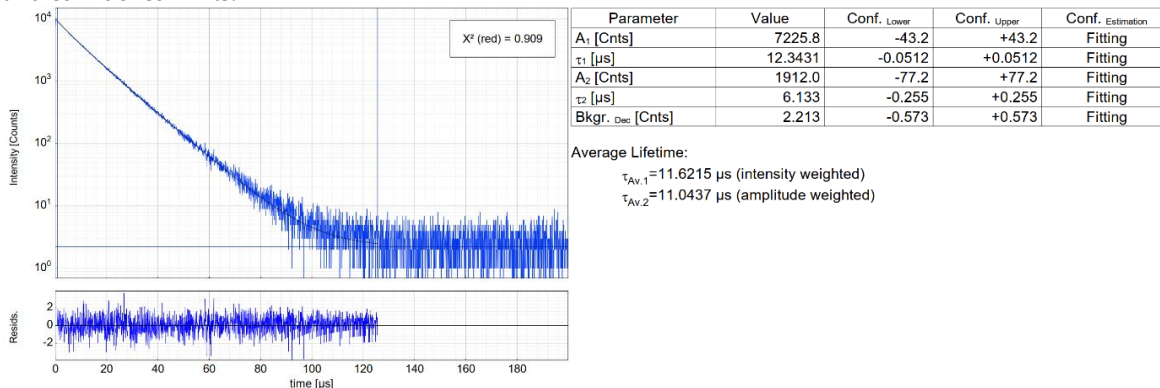


Fig. S75 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(ppy)py)Cl] (**10**) in a frozen glassy $\text{CH}_2\text{Cl}_2/\text{MeOH}$ matrix (V:V = 1:1) at 77 K, including the residuals ($\lambda_{\text{exc}} = 376.7 \text{ nm}$, $\lambda_{\text{em}} = 600 \text{ nm}$). Right: Fitting parameters including pre-exponential factors and confidence limits.

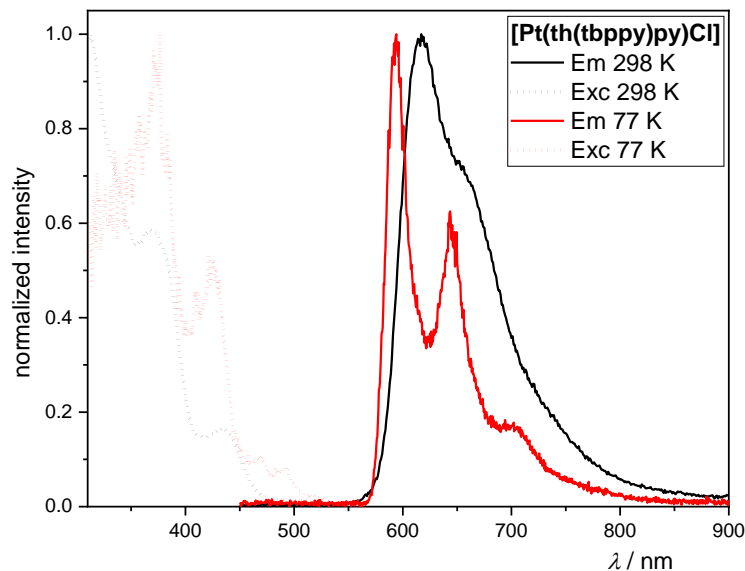


Fig. S76 Excitation (dotted line) and emission spectra (solid line) of [Pt(th(tbpp)py)Cl] (**11**) ($\lambda_{exc} = 350$ nm; $\lambda_{em} = 600$ nm) at 298 K (black) in CH_2Cl_2 and at 77 K (red) in a frozen glassy $\text{CH}_2\text{Cl}_2/\text{MeOH}$ matrix (V:V = 1:1). All solutions were optically diluted ($A < 0.1$). Spectra normalised to the highest intensity.

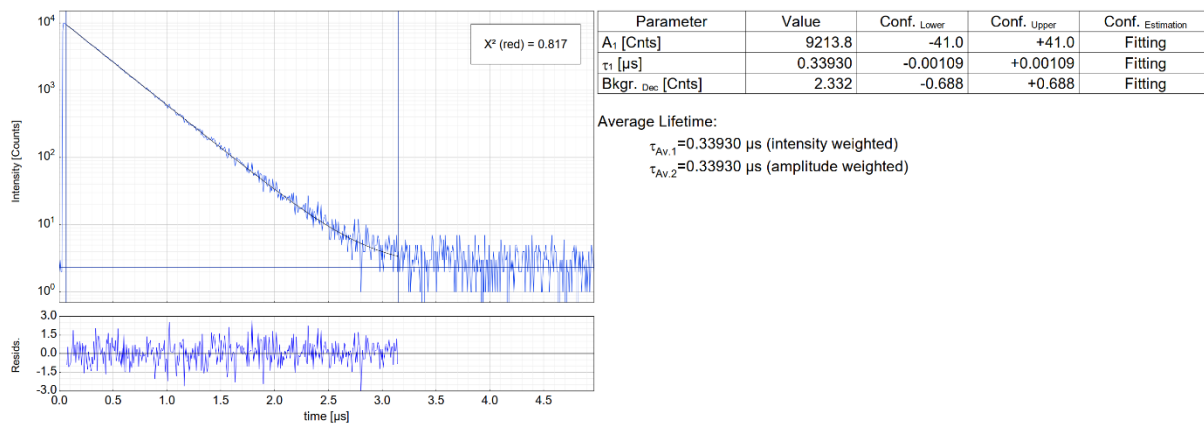


Fig. S77 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbpp)py)Cl] (**11**) in CH_2Cl_2 at 298 K (air-equilibrated), including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 620$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

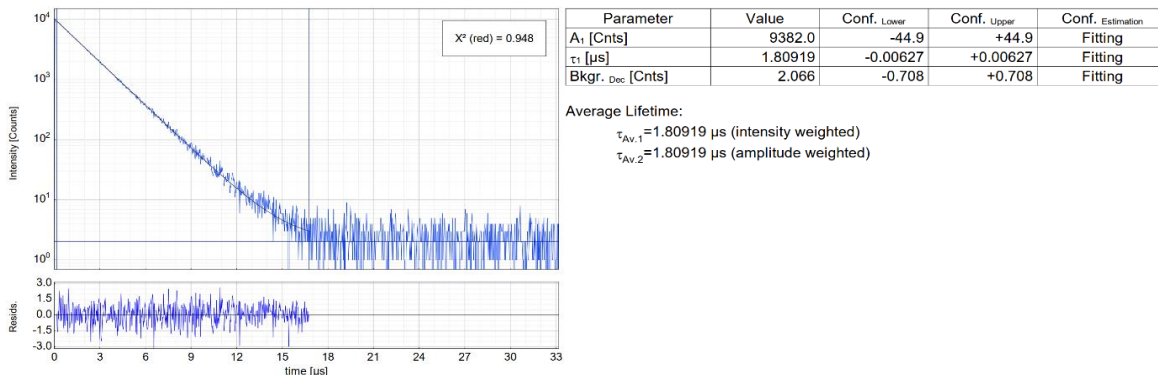


Fig. S78 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbpp)py)Cl] (**11**) in CH_2Cl_2 at 298 K (Ar-purged), including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 620$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

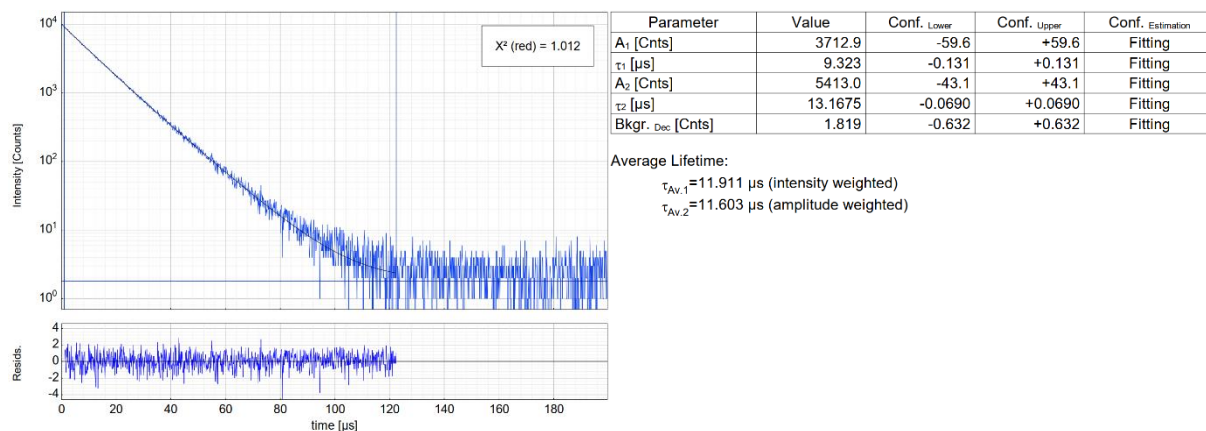


Fig. S79 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbpp)py)Cl] (**11**) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1) at 77 K, including the residuals ($\lambda_{\text{exc}} = 376.7$ nm, $\lambda_{\text{em}} = 600$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

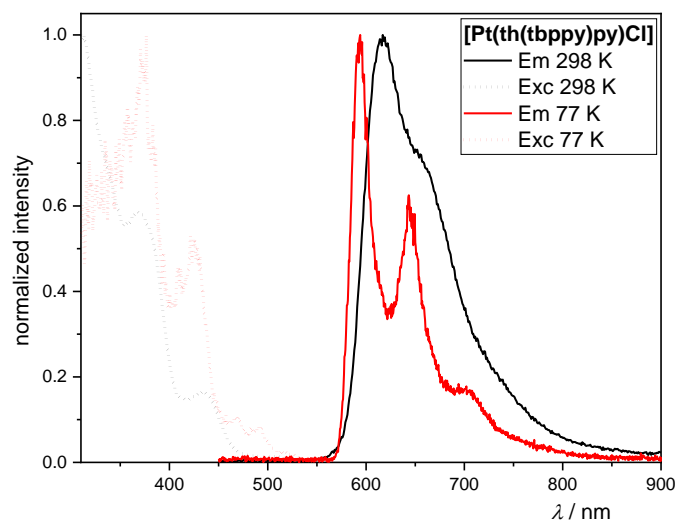


Fig. S80 Excitation (dotted line) and emission spectra (solid line) of [Pt(th(tbpp)py)CCPh] (**12**) ($\lambda_{\text{exc}} = 350$ nm; $\lambda_{\text{em}} = 600$ nm) at 298 K (black) in CH₂Cl₂ and at 77 K (red) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1). All solutions were optically diluted ($A < 0.1$). Spectra normalised to the highest intensity.

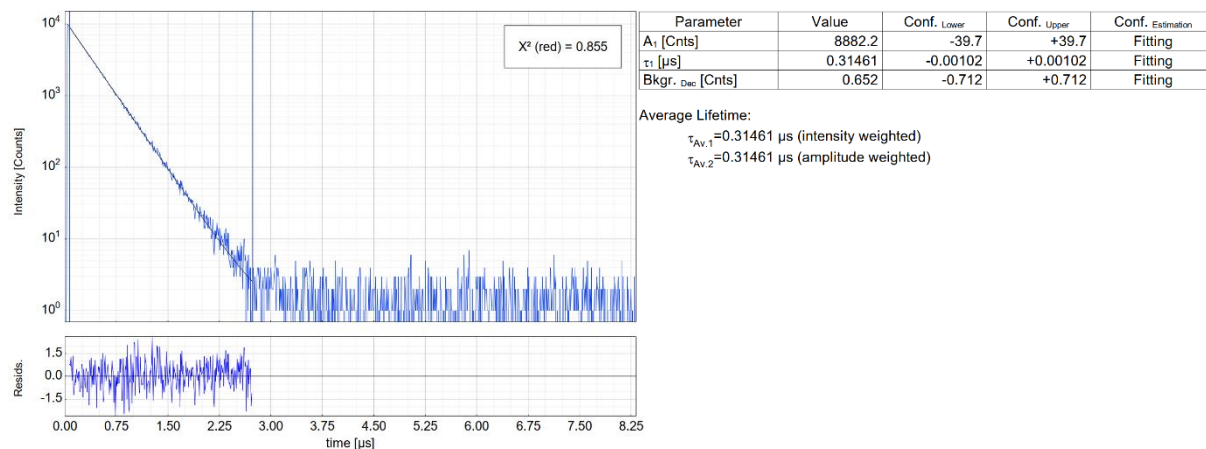


Fig. S81 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbpp)py)CCPh] (**12**) in CH₂Cl₂ at 298 K (air-equilibrated), including the residuals ($\lambda_{\text{exc}} = 376.7$ nm, $\lambda_{\text{em}} = 620$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

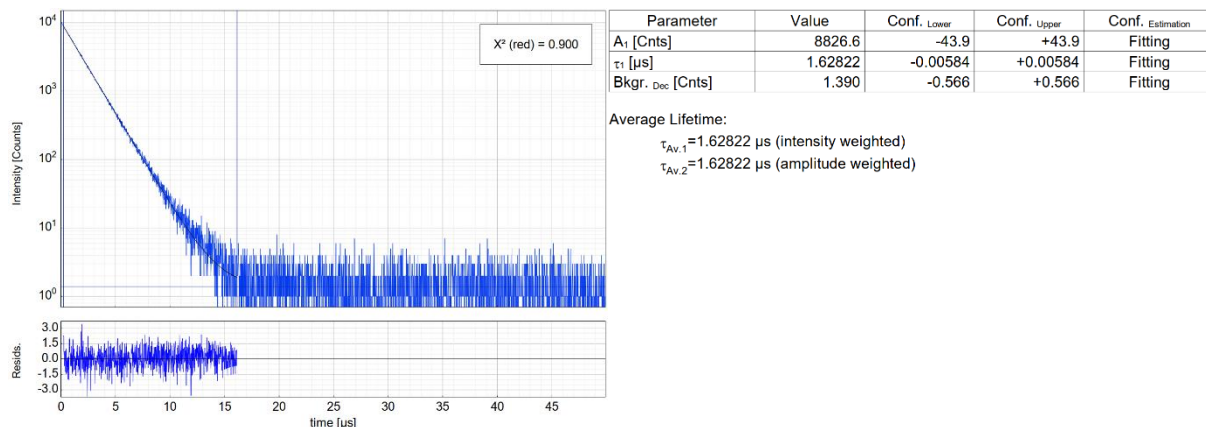


Fig. S82 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbpp)py)CCPh] (**12**) in CH₂Cl₂ at 298 K (Ar-purged), including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 620$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

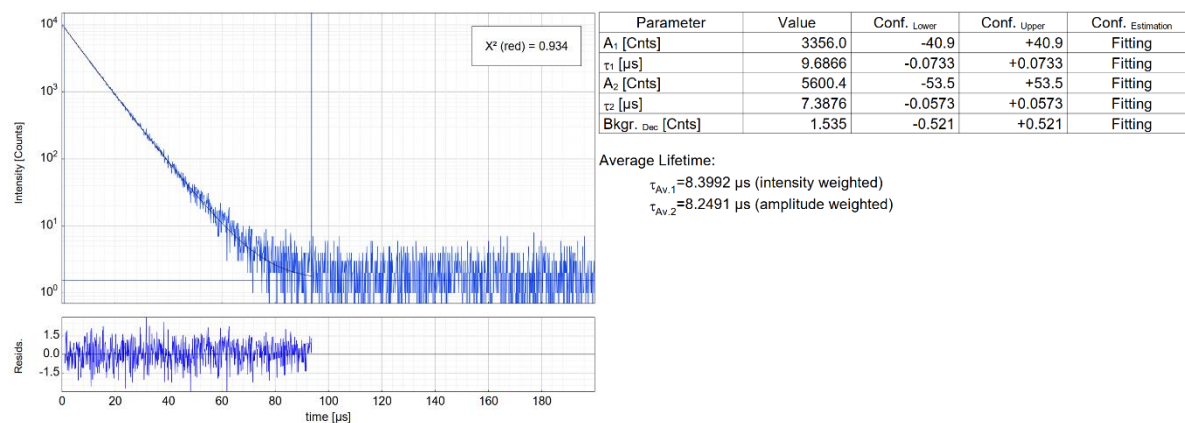


Fig. S83 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbpp)py)CCPh] (**12**) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1) at 77 K, including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 600$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

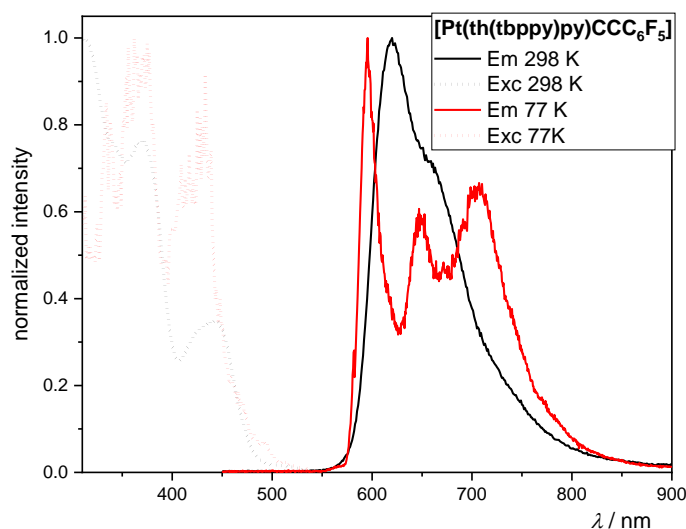


Fig. S84 Excitation (dotted line) and emission spectra (solid line) of [Pt(th(tbpp)py)CCC₆F₅] (**13**) ($\lambda_{exc} = 350$ nm; $\lambda_{em} = 600$ nm) at 298 K (black) in CH₂Cl₂ and at 77 K (red) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1). All solutions were optically diluted ($A < 0.1$). Spectra normalised to the highest intensity.

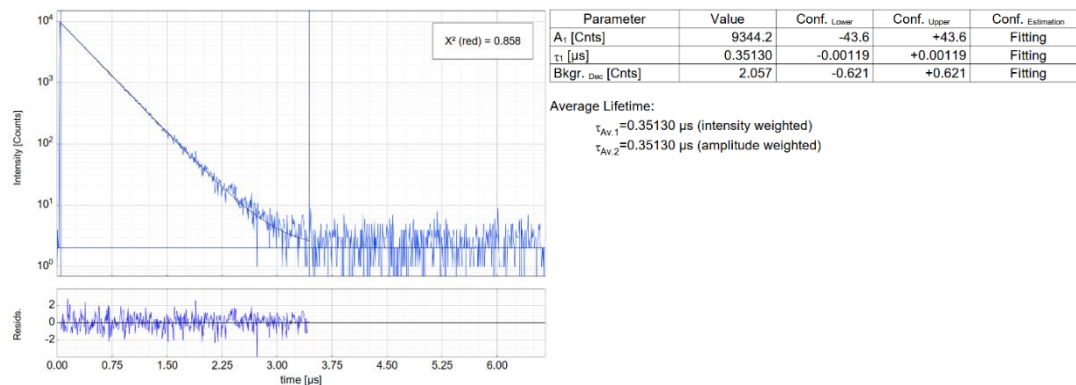


Fig. S85 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbppy)py)CCF₅] (**13**) in CH₂Cl₂ at 298 K (air-equilibrated), including the residuals ($\lambda_{exc} = 376.7 \text{ nm}$, $\lambda_{em} = 620 \text{ nm}$). Right: Fitting parameters including pre-exponential factors and confidence limits.

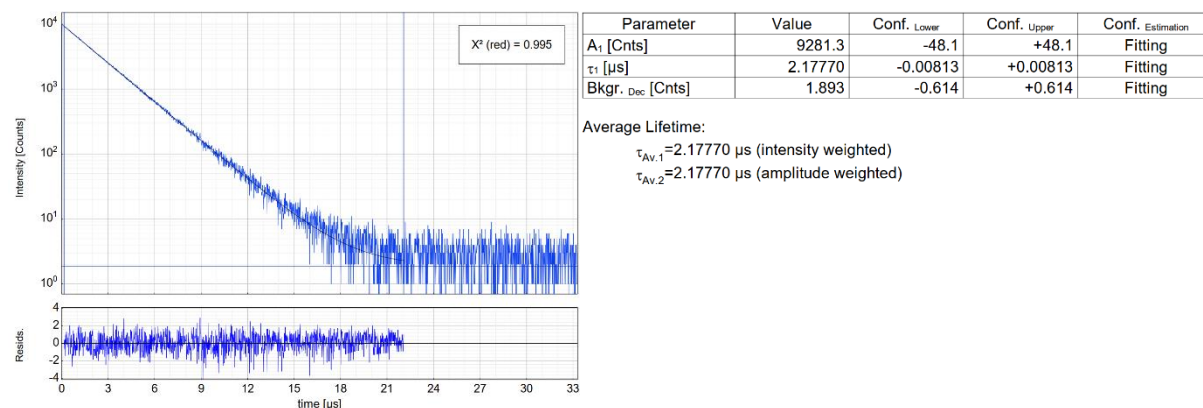


Fig. S86 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbppy)py)CCF₅] (**13**) in CH₂Cl₂ at 298 K (Ar-purged), including the residuals ($\lambda_{exc} = 376.7 \text{ nm}$, $\lambda_{em} = 620 \text{ nm}$). Right: Fitting parameters including pre-exponential factors and confidence limits.

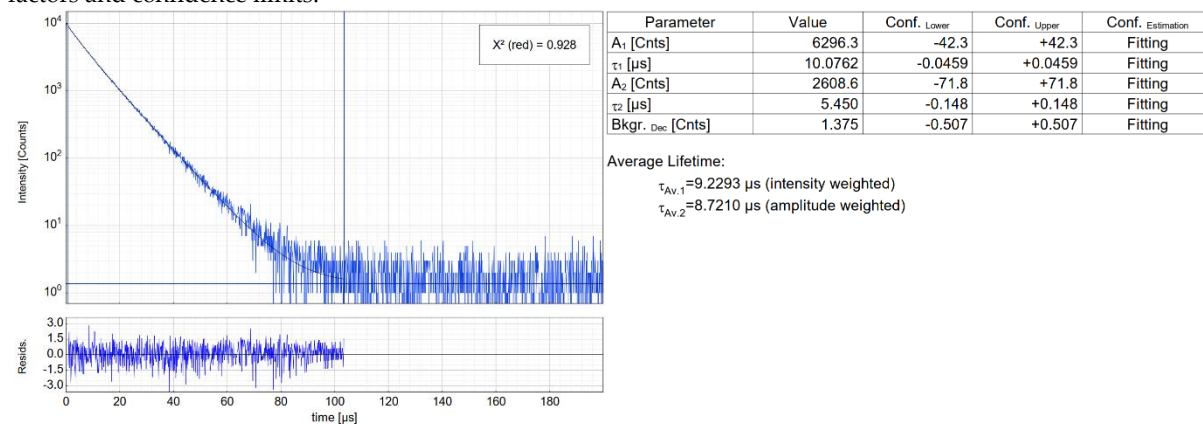


Fig. S87 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbppy)py)CCF₅] (**13**) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1) at 77 K, including the residuals ($\lambda_{exc} = 376.7 \text{ nm}$, $\lambda_{em} = 600 \text{ nm}$). Right: Fitting parameters including pre-exponential factors and confidence limits.

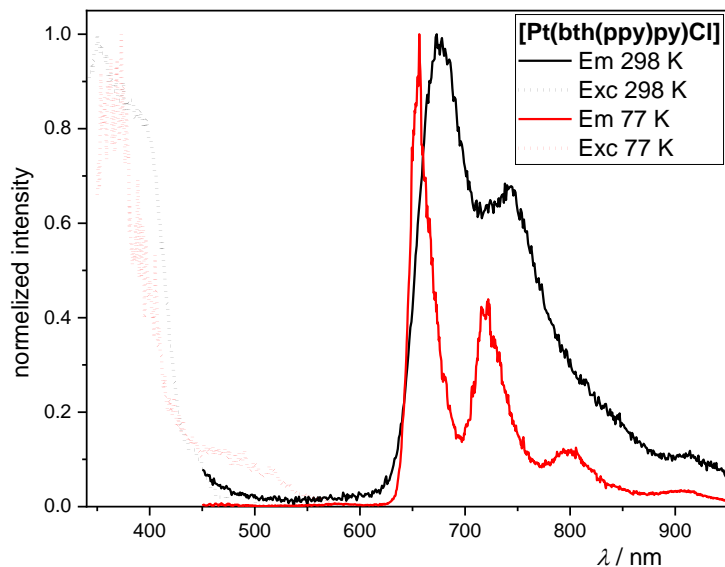


Fig. S88 Excitation (dotted line) and emission spectra (solid line) of [Pt(bth(ppy)py)Cl] (**14**) ($\lambda_{\text{exc}} = 350 \text{ nm}$; $\lambda_{\text{em}} = 650 \text{ nm}$ at 298 K; $\lambda_{\text{em}} = 660 \text{ nm}$ at 77 K) at 298 K (black) in CH_2Cl_2 and at 77 K (red) in a frozen glassy $\text{CH}_2\text{Cl}_2/\text{MeOH}$ matrix (V:V = 1:1). All solutions were optically diluted ($A < 0.1$). Spectra normalised to the highest intensity.

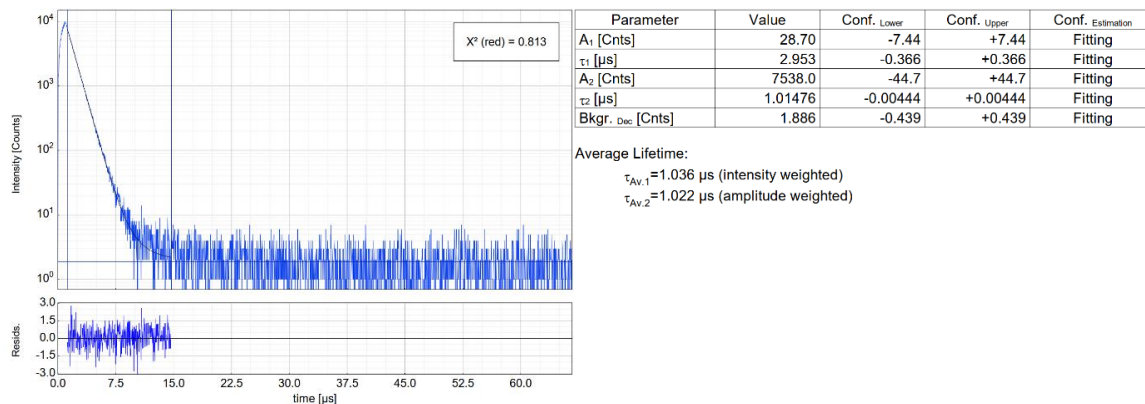


Fig. S89 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(bth(ppy)py)Cl] (**14**) in CH_2Cl_2 at 298 K (Ar-purged), including the residuals ($\lambda_{\text{exc}} = 376.7 \text{ nm}$, $\lambda_{\text{em}} = 670 \text{ nm}$). Right: Fitting parameters including pre-exponential factors and confidence limits.

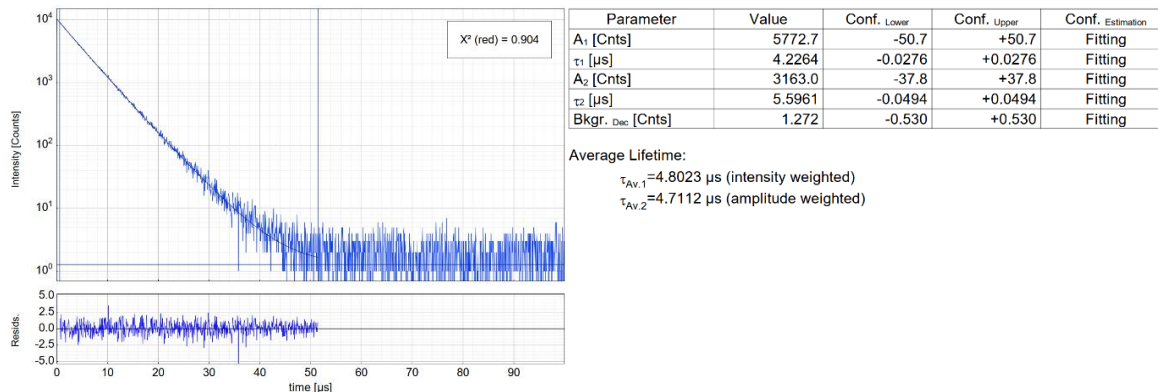


Fig. S90 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(bth(ppy)py)Cl] (**14**) in a frozen glassy $\text{CH}_2\text{Cl}_2/\text{MeOH}$ matrix (V:V = 1:1) at 77 K, including the residuals ($\lambda_{\text{exc}} = 376.7 \text{ nm}$, $\lambda_{\text{em}} = 660 \text{ nm}$). Right: Fitting parameters including pre-exponential factors and confidence limits.

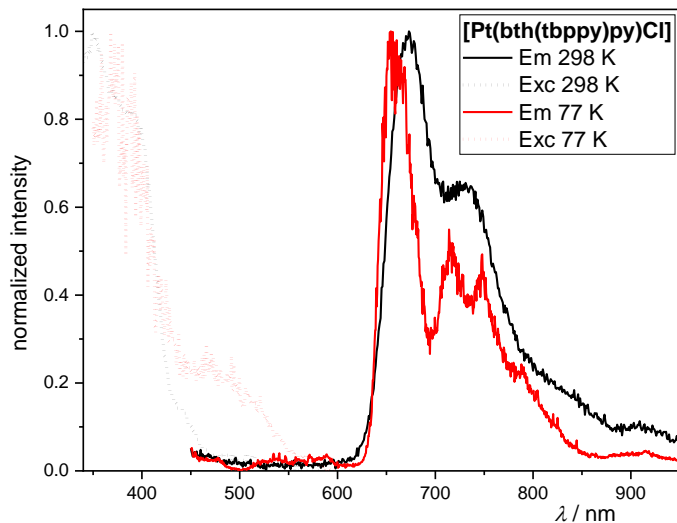


Fig. S91 Excitation (dotted line) and emission spectra (solid line) of [Pt(bth(tbppy)py)Cl] (**15**) ($\lambda_{\text{exc}} = 350 \text{ nm}$; $\lambda_{\text{em}} = 650 \text{ nm}$ at 298 K; $\lambda_{\text{em}} = 660 \text{ nm}$ at 77 K) at 298 K (black) in CH_2Cl_2 and at 77 K (red) in a frozen glassy $\text{CH}_2\text{Cl}_2/\text{MeOH}$ matrix (V:V = 1:1). All solutions were optically diluted ($A < 0.1$). Spectra normalised to the highest intensity.

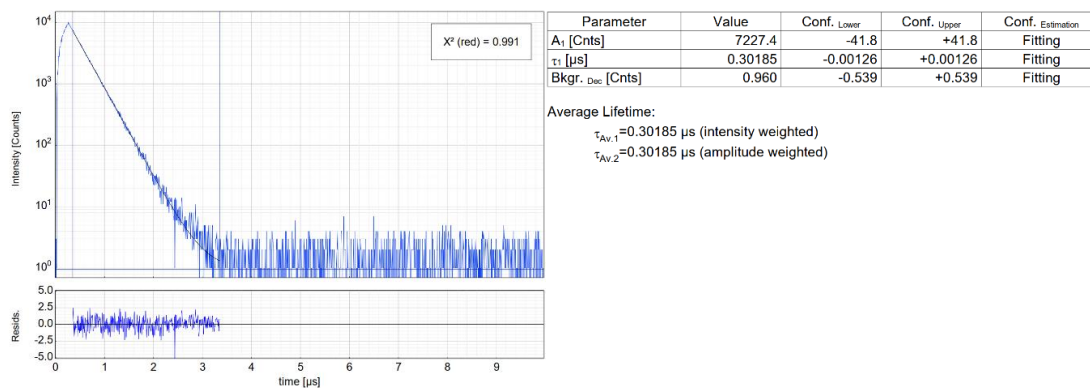


Fig. S92 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(bth(tbppy)py)Cl] (**15**) in CH_2Cl_2 at 298 K (air-equilibrated), including the residuals ($\lambda_{\text{exc}} = 376.7 \text{ nm}$, $\lambda_{\text{em}} = 670 \text{ nm}$). Right: Fitting parameters including pre-exponential factors and confidence limits.

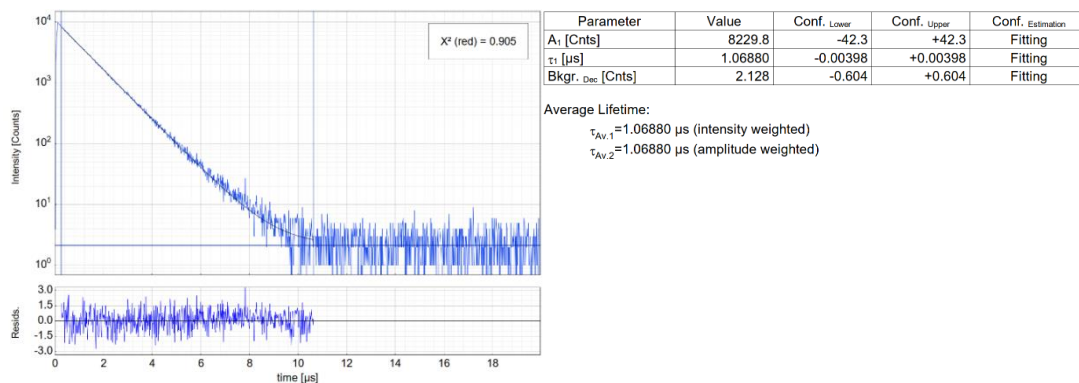


Fig. S93 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(bth(tbppy)py)Cl] (**15**) in CH_2Cl_2 at 298 K (Air-purged), including the residuals ($\lambda_{\text{exc}} = 376.7 \text{ nm}$, $\lambda_{\text{em}} = 670 \text{ nm}$). Right: Fitting parameters including pre-exponential factors and confidence limits.

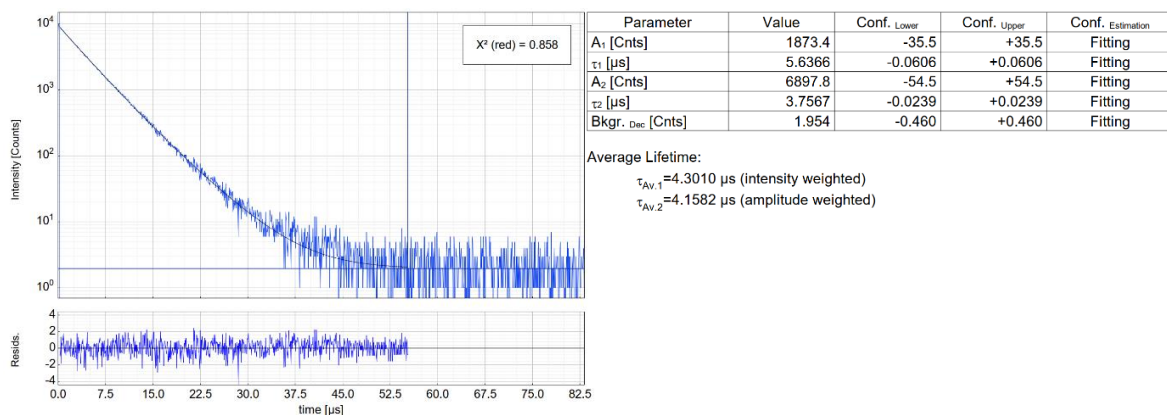


Fig. S94 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(bth(tbppy)py)Cl] (**15**) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1) at 77 K, including the residuals ($\lambda_{\text{exc}} = 376.7$ nm, $\lambda_{\text{em}} = 660$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

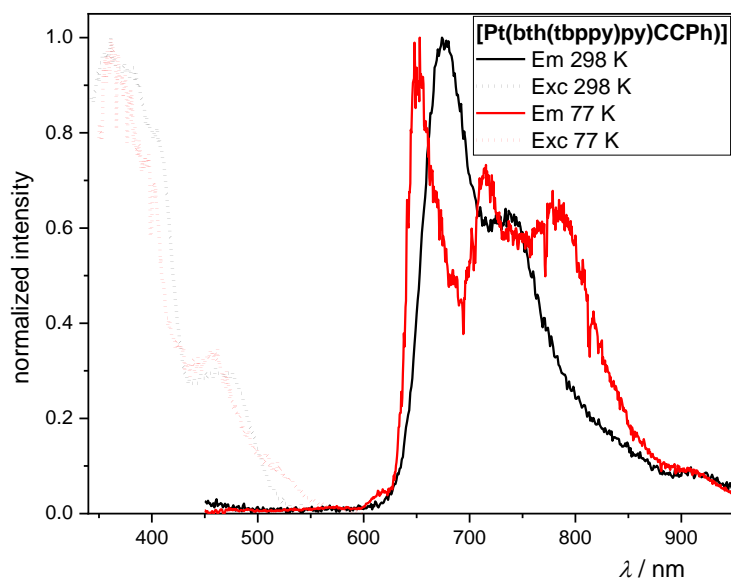


Fig. S95 Excitation (dotted line) and emission spectra (solid line) of [Pt(bth(tbppy)py)CCPh] (**16**) ($\lambda_{\text{exc}} = 350$ nm; $\lambda_{\text{em}} = 650$ nm at 298 K; $\lambda_{\text{em}} = 660$ nm at 77 K) at 298 K (black) in CH₂Cl₂ and at 77 K (red) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1). All solutions were optically diluted ($A < 0.1$). Spectra normalised to the highest intensity.

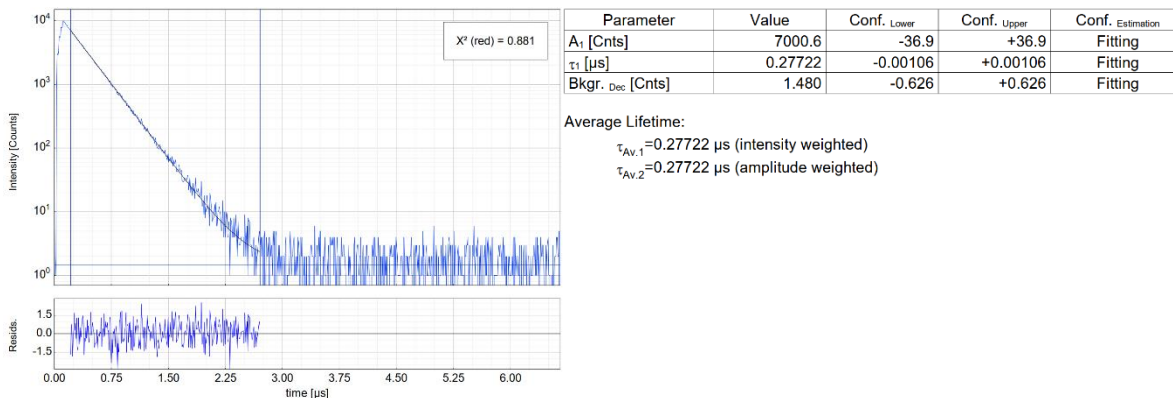


Fig. S96 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(bth(tbppy)py)CCPh] (**16**) in CH₂Cl₂ at 298 K (air-equilibrated), including the residuals ($\lambda_{\text{exc}} = 376.7$ nm, $\lambda_{\text{em}} = 675$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

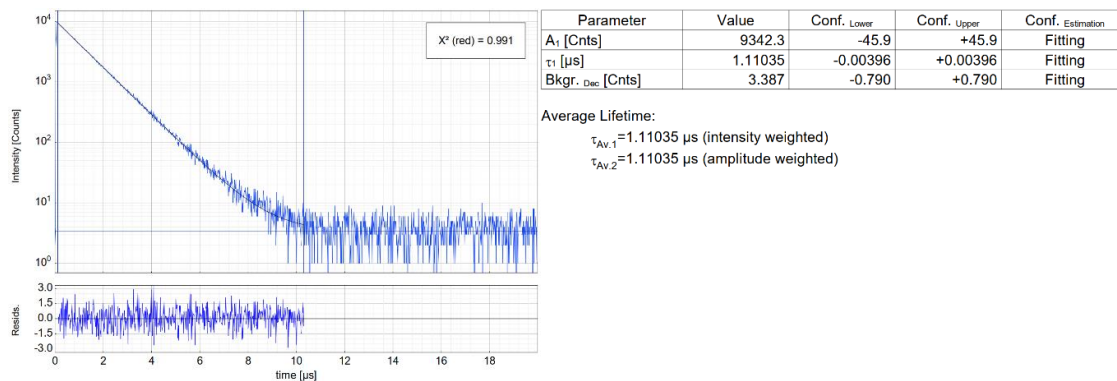


Fig. S97 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(bth(tbppy)py)CCPh] (**16**) in CH₂Cl₂ at 298 K (Ar-purged), including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 675$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

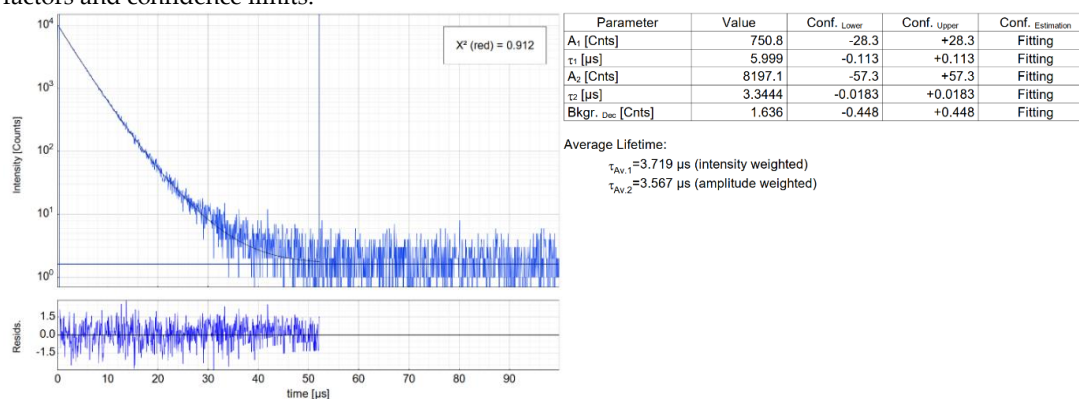


Fig. S98 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(bth(tbppy)py)CCPh] (**16**) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1) at 77 K, including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 660$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

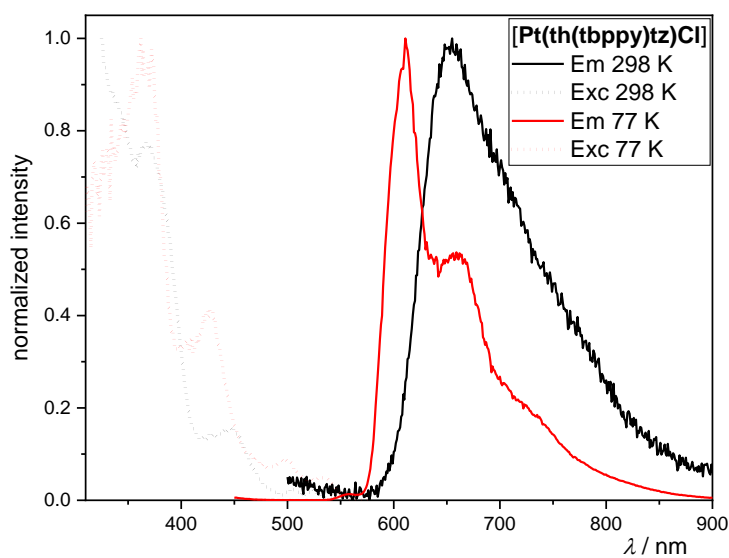


Fig. S99 Excitation (dotted line) and emission spectra (solid line) of [Pt(th(tbppy)tz)Cl] (**17**) ($\lambda_{exc} = 350$ nm; $\lambda_{em} = 620$ nm at 298 K, $\lambda_{em} = 600$ nm at 77 K) at 298 K (black) in CH₂Cl₂ and at 77 K (red) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1). All solutions were optically diluted ($A < 0.1$). Spectra normalised to the highest intensity.

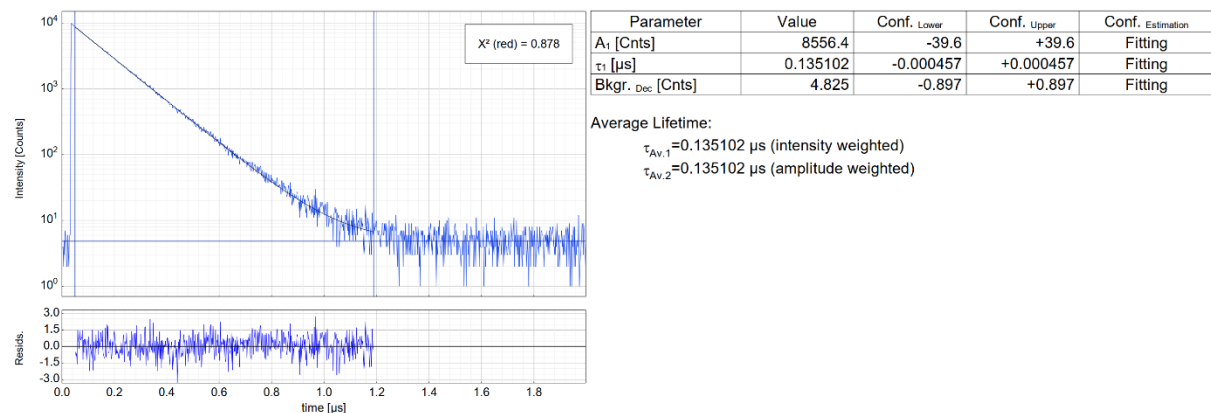


Fig. S100 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbpp)tz)Cl] (**17**) in CH₂Cl₂ at 298 K (air-equilibrated), including the residuals ($\lambda_{\text{exc}} = 376.7$ nm, $\lambda_{\text{em}} = 650$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

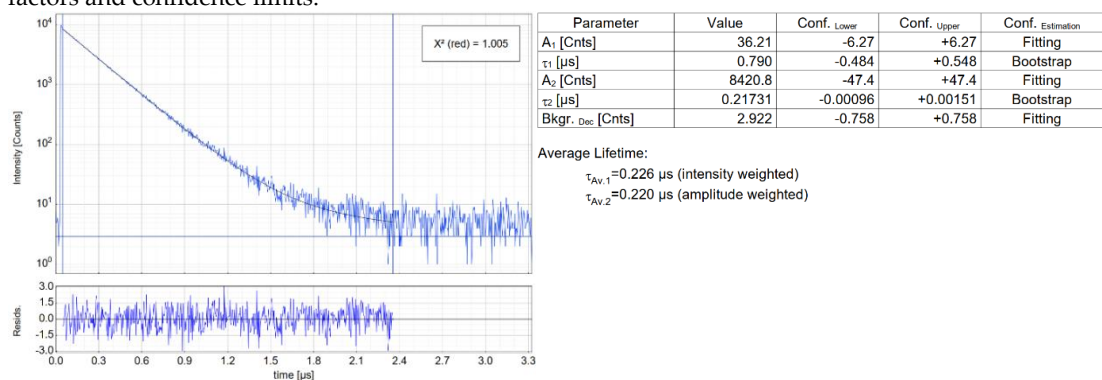


Fig. S101 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbpp)tz)Cl] (**17**) in CH₂Cl₂ at 298 K (Ar-purged), including the residuals ($\lambda_{\text{exc}} = 376.7$ nm, $\lambda_{\text{em}} = 650$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

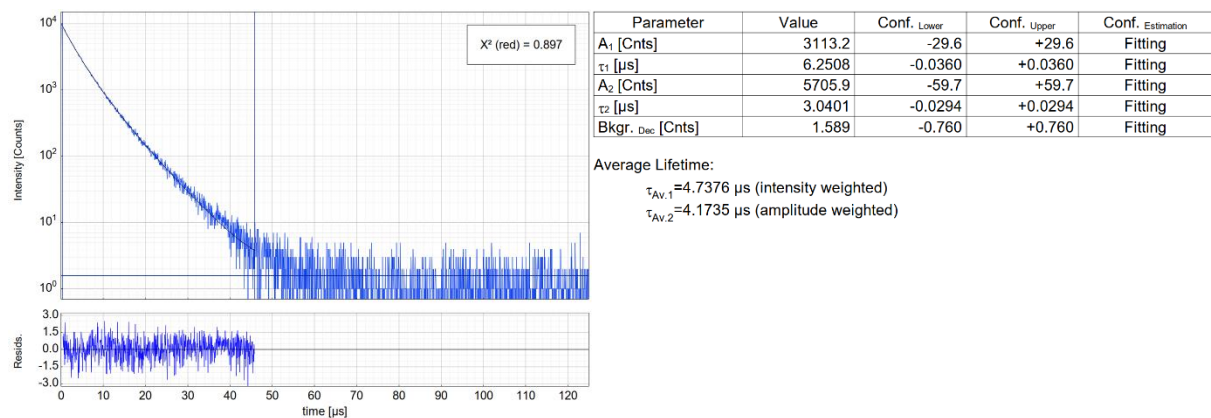


Fig. S102 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbpp)tz)Cl] (**17**) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1) at 77 K, including the residuals ($\lambda_{\text{exc}} = 376.7$ nm, $\lambda_{\text{em}} = 600$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

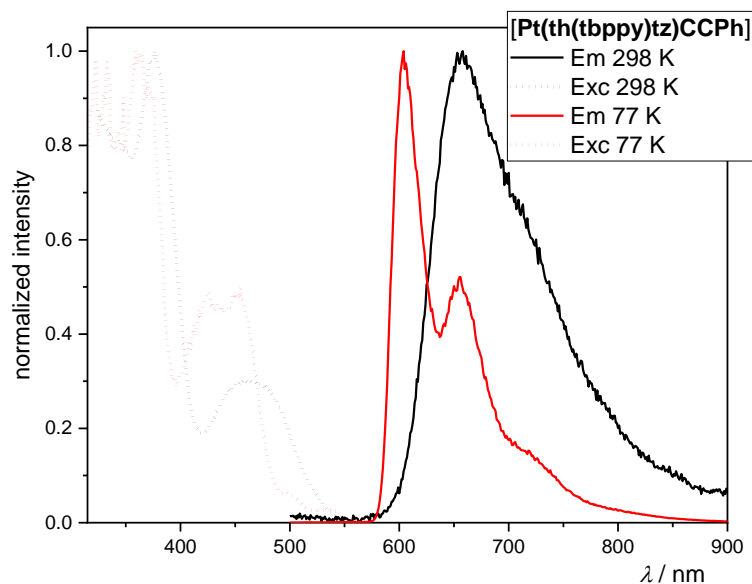


Fig. S103 Excitation (dotted line) and emission spectra (solid line) of [Pt(th(tbppytz)CCPh) (**18**) ($\lambda_{exc} = 350$ nm; $\lambda_{em} = 620$ nm at 298 K, $\lambda_{em} = 600$ nm at 77 K) at 298 K (black) in CH_2Cl_2 and at 77 K (red) in a frozen glassy $\text{CH}_2\text{Cl}_2/\text{MeOH}$ matrix (V:V = 1:1). All solutions were optically diluted ($A < 0.1$). Spectra normalised to the highest intensity.

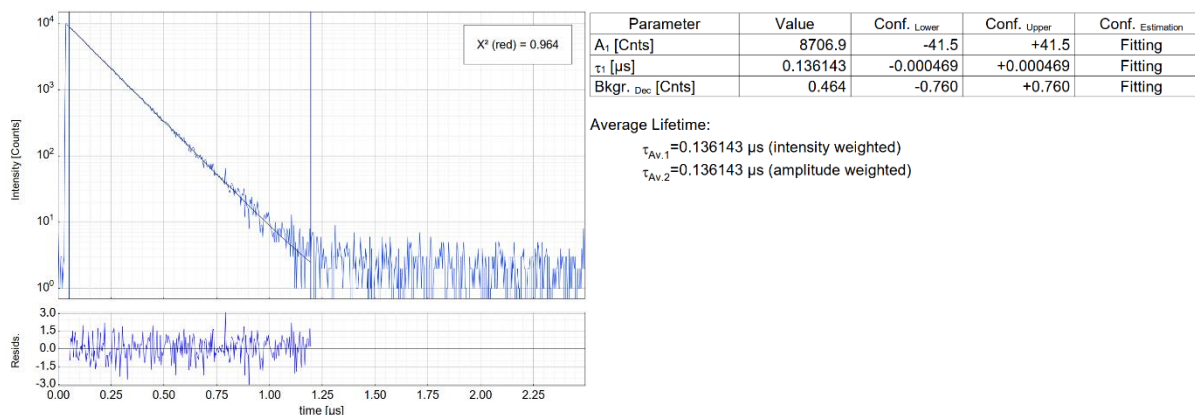


Fig. S104 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbppytz)CCPh) (**18**) in CH_2Cl_2 at 298 K (air-equilibrated), including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 650$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

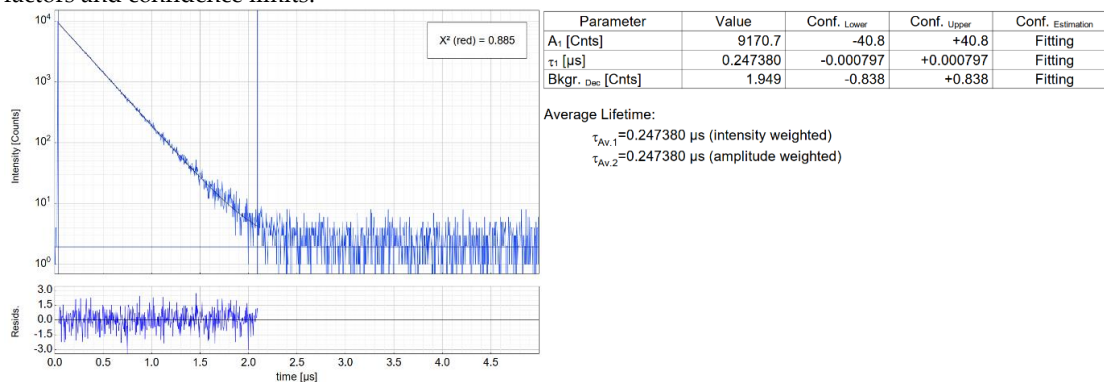


Fig. S105 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbppytz)CCPh) (**18**) in CH_2Cl_2 at 298 K (Ar-purged), including the residuals ($\lambda_{exc} = 376.7$ nm, $\lambda_{em} = 650$ nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

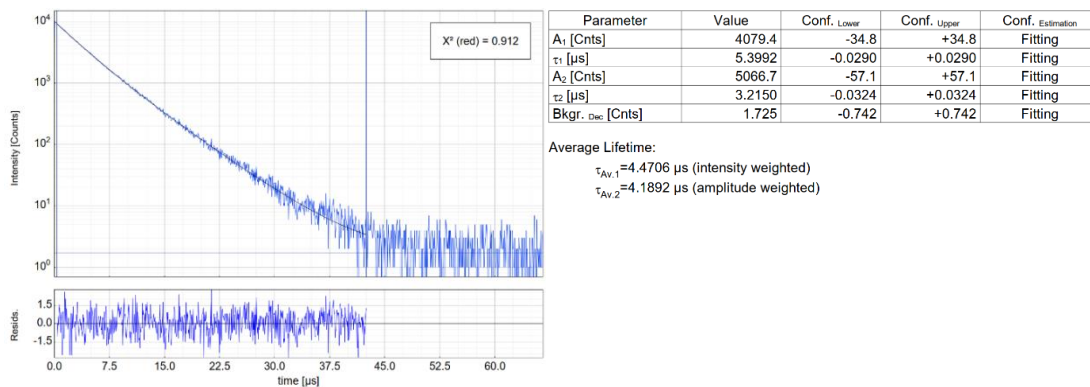


Fig. S106 Left: Raw (experimental) time-resolved photoluminescence decay of [Pt(th(tbppy)tz)CCPh] (**18**) in a frozen glassy CH₂Cl₂/MeOH matrix (V:V = 1:1) at 77 K, including the residuals (λ_{exc} = 376.7 nm, λ_{em} = 600 nm). Right: Fitting parameters including pre-exponential factors and confidence limits.

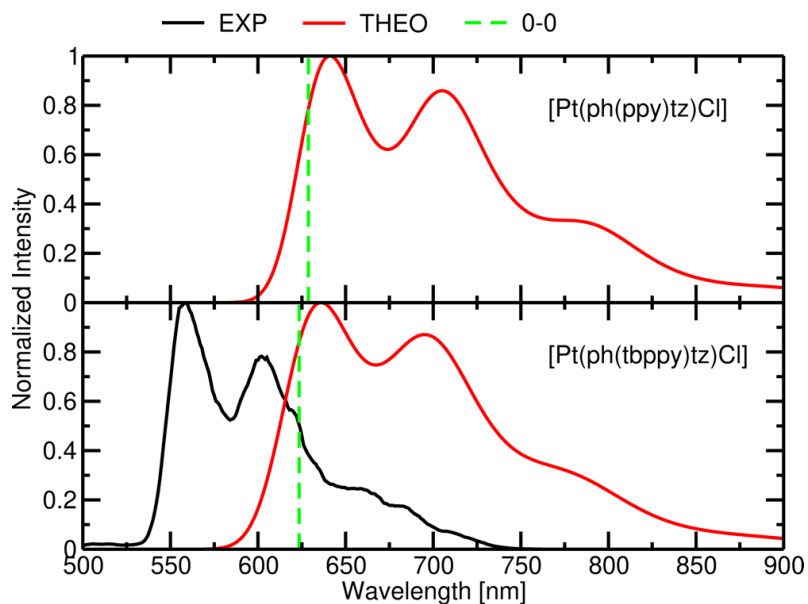


Fig. S107 TD-DFT-calculated emission spectra (THEO, red) compared to experimental spectra at 77 K (EXP, black). Theoretical 0-0 transitions are shown as vertical dashed green lines. Upper panel: complex **1** [Pt(ph(ppy)tz)Cl]. Lower panel complex **2** [Pt(ph(tbppy)tz)Cl].

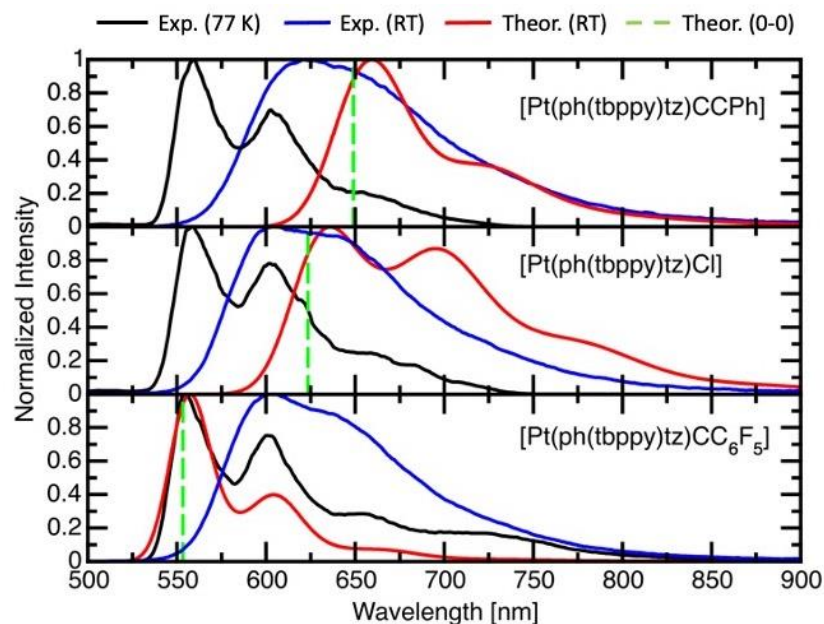


Fig. S108 TD-DFT-calculated emission spectra (Theor, red) compared to experimental spectra (EXP, black) for complexes of **class A** complexes at 77 K. Theoretical 0–0 transitions are shown as vertical dashed green lines. Upper panel: complex 3 [Pt(ph(tbppy)tz)C≡CPh]. Middle panel: complex 2 [Pt(ph(tbppy)tz)Cl]. Lower panel: complex 4 [Pt(ph(tbppy)tz)C≡CC₆F₅].

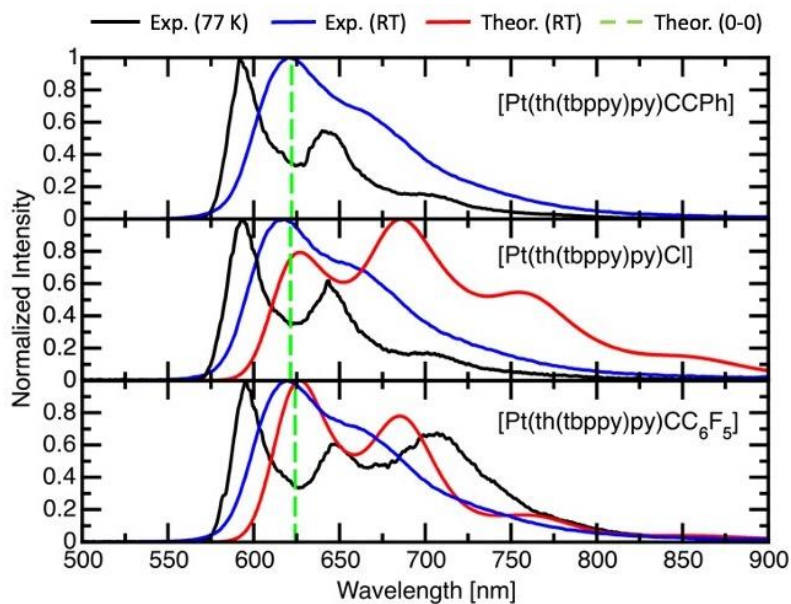


Fig. S109 TD-DFT-calculated emission spectra (Theor, red) of **class B** complexes compared to experimental spectra (EXP, black) at 77 K. Calculated 0–0 transitions are shown as vertical dashed green lines. Upper panel: complex 12 [Pt(th(tbppy)py)C≡CPh]. Middle panel: complex 11 [Pt(th(tbppy)py)Cl]. Lower panel: complex 13 [Pt(th(tbppy)py)C≡CC₆F₅].

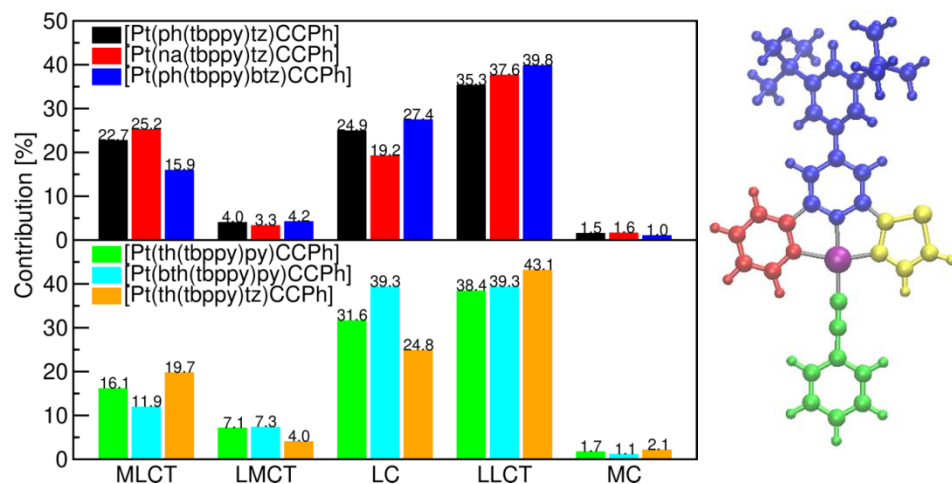


Fig. S110 Left: Decomposition of the emissive T_1 states into MLCT, LMCT, LC, LLCT, and MC contributions of complexes **3**, **6**, **9** (top), and **12**, **16**, **18** (bottom). Right: Molecular partitioning used for the example of complex **3**, all other complexes were partitioned in an analogous manner.

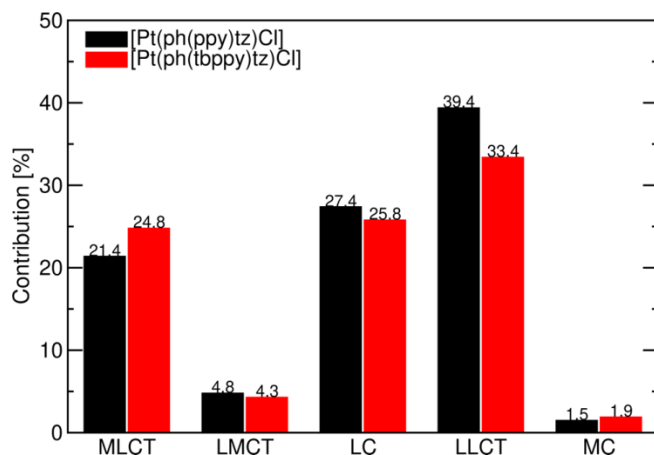


Fig. S111 Decomposition of the emissive T_1 states into MLCT, LMCT, LC, LLCT, and MC contributions of [Pt(ph(ppy)tz)Cl] (**1**) and [Pt(ph(tbppy)tz)Cl] (**2**).

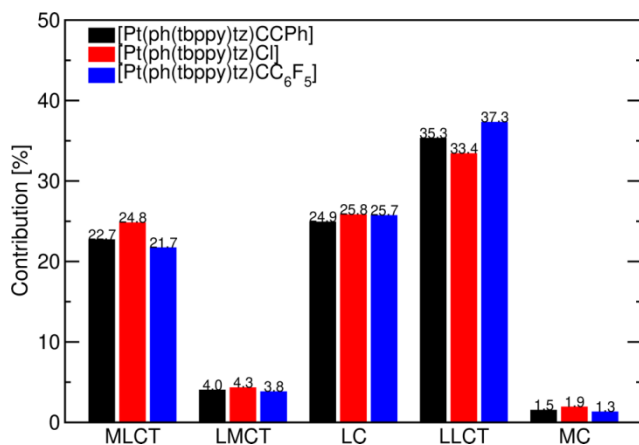


Fig. S112 Decomposition of the emissive T_1 states into MLCT, LMCT, LC, LLCT, and MC contributions for **class A** complexes (black: complex **3**) [Pt(ph(tbppy)tz)CCPh], red: complex **2**) [Pt(ph(tbppy)tz)Cl], blue: complex **4**) [Pt(ph(tbppy)tz)CC₆F₅]).

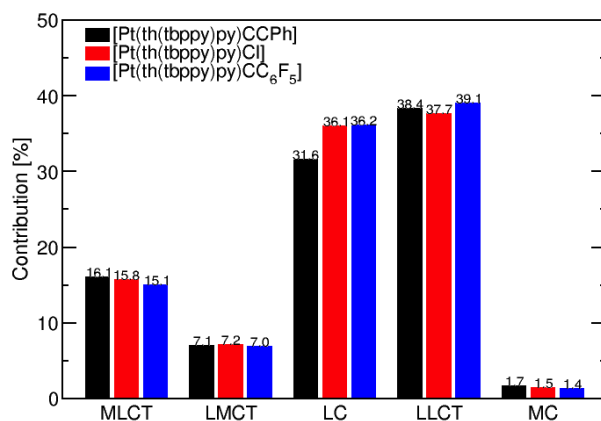


Fig. S113 Decomposition of the emissive T_1 states into MLCT, LMCT, LC, LLCT, and MC contributions for **class B** complexes (black: [Pt(th(tbbpy)py)(C≡CPh)] (**12**), red: [Pt(th(tbbpy)py)Cl] (**11**), blue: [Pt(th(tbbpy)py)(C≡CC₆F₅)] (**13**)).

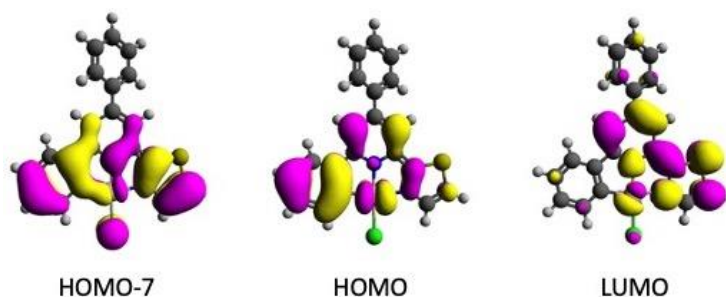


Fig. S114 Molecular orbitals of complex (**1**) with the largest contributions to the emissive T_1 state.

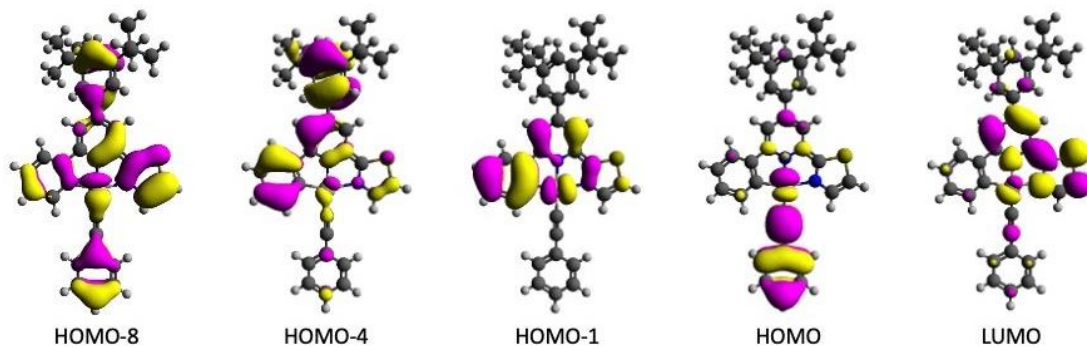


Fig. S115 Molecular orbitals of complex (**3**) with the largest contributions to the emissive T_1 state.

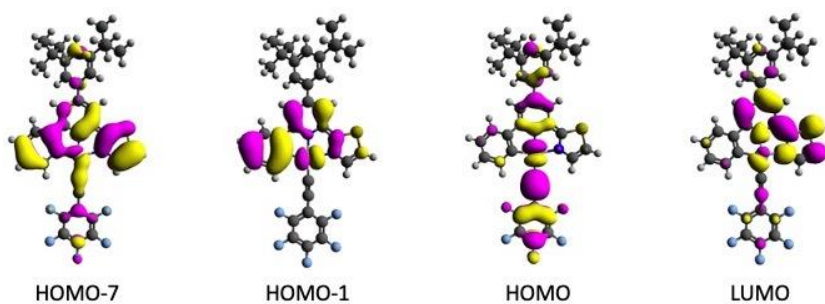


Fig. S116 Molecular orbitals of complex (**4**) with the largest contributions to the emissive T_1 state.

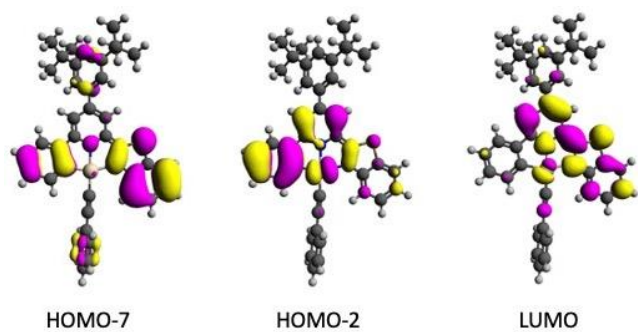


Fig. S117 Molecular orbitals of complex (9) with the largest contributors to the emissive T_1 state.

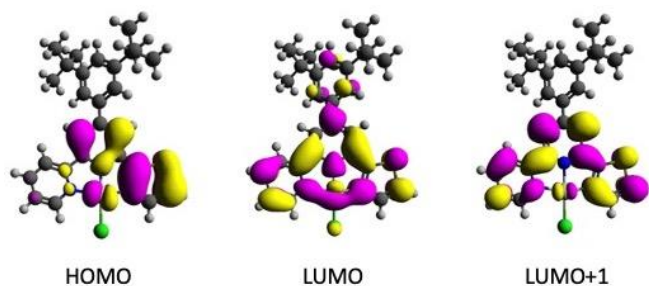


Fig. S118 Molecular orbitals of complex (11) with the largest contributions to the emissive T_1 state.

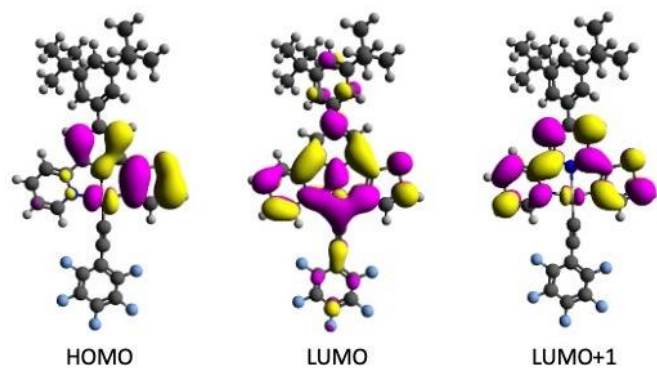


Fig. S119 Molecular orbitals of complex (12) with the largest contributions to the emissive T_1 state.

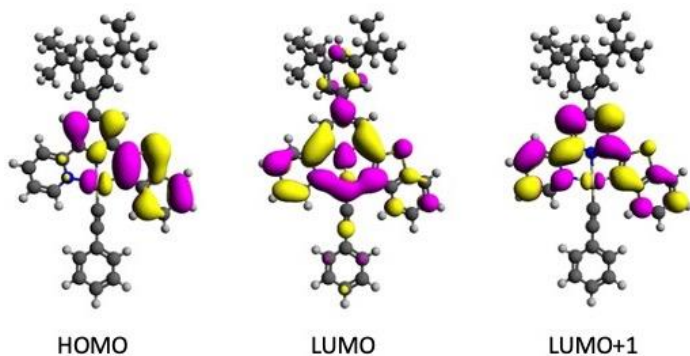


Fig. S120 Molecular orbitals of complex (16) with the largest contributions to the emissive T_1 state.

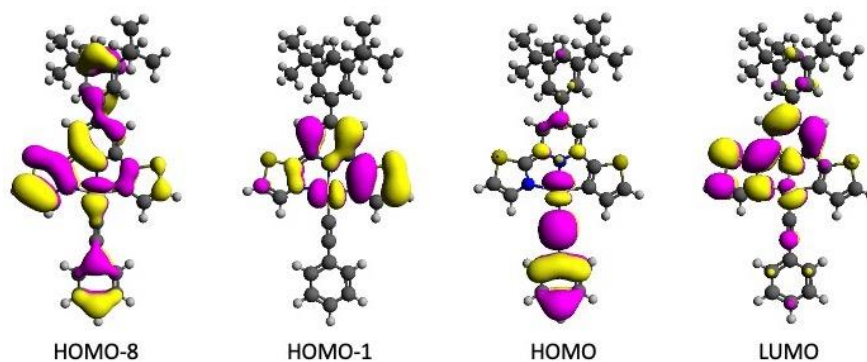


Fig. S121 Molecular orbitals of complex (18) with the largest contributions to the emissive T_1 state.

Supporting Tables

Table S1 Selected structure solution and refinement data for $[\text{Pt}(\text{th}(\text{ppy})\text{py})\text{Cl}]\text{CD}_2\text{Cl}_2$ (**10-CH₂Cl₂**) and $[\text{Pt}(\text{ph}(\text{tbbpy})\text{btz})(\text{CCPh})]$ (**9**).

Complex	$[\text{Pt}(\text{th}(\text{ppy})\text{py})\text{Cl}]\text{CH}_2\text{Cl}_2$ (10-CH₂Cl₂)	$[\text{Pt}(\text{ph}(\text{tbbpy})\text{btz})(\text{CCPh})]$ (9) ^a
Formula	$\text{C}_{20}\text{H}_{13}\text{N}_2\text{SPtCl}\text{CD}_2\text{Cl}_2$	$\text{C}_{40}\text{H}_{36}\text{N}_2\text{PtS}$
Molecular weight	1253.67	771.86
Temperature	170(2) K	100(2) K
Radiation wavelength	0.71073 Å	0.56076 Å (synchrotron)
Crystal system	triclinic	Monoclinic
Space group	$P\bar{1}$ (No. 2)	$P 2_1/c$ (No. 14)
Cell parameters	$a = 7.9533(7)$ Å	$a = 13.1805(4)$ Å
	$b = 9.8707(8)$ Å	$b = 37.3146(19)$ Å
	$c = 14.3097(11)$ Å	$c = 6.7184(2)$ Å
	$\alpha = 102.887(6)^\circ$	
	$\beta = 103.515(6)^\circ$	$\beta = 100.192(3)^\circ$
	$\gamma = 102.960(6)^\circ$	
Volume	1018.48(15) Å ³	3252.1(2) Å ³
Z	2	4
Density calculated	2.052 g/cm ³	1.576 g/cm ³
Absorption coefficient	7.397 mm ⁻¹	2.400 mm ⁻¹
F(000)	601	1536
Theta range for data collection	2.213 to 29.293°	1.729 to 25.650°.
Index ranges	$-10 \leq h \leq 10, -12 \leq k \leq 13, -19 \leq l \leq 19$	$-20 \leq h \leq 20, -53 \leq k \leq 53, -8 \leq l \leq 8$
Reflections collected	13115	45728
Independent reflections	5443 [$R_{\text{int}} = 0.0794$]	10542 [$R_{\text{int}} = 0.0982$]
Completeness to theta	theta = 19.661°: 99.8%	theta = 19.661°: 97.0%
Refinement method	Full-matrix least-squares on F^2	Full-matrix least-squares on F^2
Data / restraints / parameters	5443 / 0 / 254	10542 / 0 / 404
Goodness-of-fit on F^2	0.982	1.064
Final R indices [$I > 2\sigma(I)$]	$R_1 = 0.0493, wR_2 = 0.1175$	$R_1 = 0.00736, wR_2 = 0.1626$
R indices (all data)	$R_1 = 0.0709, wR_2 = 0.1300$	$R_1 = 0.1287, wR_2 = 0.1854$
Extinction coefficient	0.0082(9)	0.0015(3)
Largest diff. peak and hole	1.96 and -2.29 e.Å^{-3}	3.90 and -4.21 e.Å^{-3}
CCDC	2081667	2084095

The dataset for 9 is generally poor and A alerts appeared in the checkcif. The problem is due to poor quality of the crystal.

Table S2 Selected structural data for $[\text{Pt}(\text{th}(\text{ppy})\text{py})\text{Cl}]\text{CH}_2\text{Cl}_2$ (**10-CH₂Cl₂**) and $[\text{Pt}(\text{ph}(\text{tbbpy})\text{btz})(\text{CCPh})]$ (**9**).

distances (Å)	[Pt(th(ppy)py)Cl]CD ₂ Cl ₂ (10 CH ₂ Cl ₂)	distances (Å)	[Pt(ph(tbppy)btz)(CCPh)] (9)
Pt–C1	1.983(7)	Pt–C1	1.99(1)
Pt–N1	1.957(6)	Pt–N1	2.00(7)
Pt–N2	2.109(5)	Pt–N2	2.13(7)
Pt–Cl	2.293(3)	Pt–C33	1.98(1)
C4–C20	1.45(1)	C12–C11	1.46(1)
C30–C24	1.49(1)	C6–C7	1.47(1)
angles (°)		angles (°)	
Cl–Pt–N1	178.8(2)	C33–Pt–N1	175.3(4)
C1–Pt–N2	161.3(3)	C1–Pt–N2	159.1(3)
C1–Pt–N1	82.2(3)	C1–Pt–N1	81.4(3)
N1–Pt–N2	79.1(2)	N1–Pt–N2	77.6(3)
Cl–Pt–C1	98.6(3)	C33–Pt–C1	93.9(4)
Cl–Pt–N2	100.0(2)	C33–Pt–N2	107.1(4)
C3–S1–C4	90.4(4)	C14–S1–C12	89.3(5)
C1–C4–C20–C21	176.7(9)	C1–C6–C7–C8	175.8(9)
N2–C30–C24–C23	178.0(8)	N2–C12–C11–C10	179.5(9)
C2–C1–C4–C20	175.0(8)	C13–N2–C12–C11	179.6(8)
C21–C22–C40–C45	38.8(11)	C8–C9–C19–C24	33.2(14)
Sum of angles around Pt	360		360

Table S3 Redox potentials of the HC[^]N[^]N protoligands.^a

	$E_{1/2}$ (Red1)	E_{pc} (Red2)	E_{pc} (Red3)
Hth(ppy)py	-2.53	-3.28	
Hth(tbppy)py	-2.57	-3.39	
Hbth(ppy)py	-2.58	-3.04	-3.18
Hbth(tbppy)py	-2.51	-2.92	-3.05
Hph(ppy)tz	-2.44	-3.22	
Hph(tbppy)tz	-2.50	-2.75	-3.43
Hna(ppy)tz	-2.44	-3.02	
Hna(tbppy)tz	-2.45	-3.07	
Hth(tbppy)tz	-2.41		
Hph(ppy)btz	-2.26	-2.76	-3.19
Hph(tbppy)btz	-2.31	-2.89	-3.27

^a Electrochemical potentials [V], $E_{1/2}$ for reversible redox waves, E_{pc} for irreversible reduction waves, measured in 0.1 M *n*-Bu₄NPF₆/THF at 298 K, scan rate 100 mV/s.

Table S4 Selected redox potentials of [Pt(C[^]N[^]N)(Cl/R)] complexes.^a

C [^] N [^] N	X	Number	Ox (in CH ₂ Cl ₂)		Red (in THF)			ΔE
			E_{pa} ox 2	E_{pa} ox 1	$E_{1/2}$ red 1	$E_{1/2}$ red 2	red 2–red 1	
ph(ppy)tz	Cl	(1)	-	0.46	-1.64	-2.32	0.68	2.10
ph(tbppy)tz	Cl	(2)	1.14	0.67	-1.64	-2.34	0.70	2.31
ph(tbppy)tz	CCPh	(3)	0.60	0.39	-1.66	-2.33	0.67	2.05
ph(tbppy)tz	CCCF ₅	(4)	1.02	0.55	-1.62	-2.25	0.63	2.17
na(ppy)tz	Cl	(5)	1.08	0.40	-1.63	-2.28	0.65	2.03
na(tbppy)tz	CCPh	(6)	0.62	0.34	-1.66	-2.33	0.67	2.00
ph(ppy)btz	Cl	(7)	1.16	0.53	-1.48	-2.13	0.65	2.01
ph(tbppy)btz	Cl	(8)	1.16	0.61	-1.51	-2.17	0.66	2.12
ph(tbppy)btz	CCPh	(9)	0.57	0.17	-1.54 ^b	-2.20 ^b	0.66	1.71
th(ppy)py	Cl	(10)	-	0.55	-1.76	-2.40	0.64	2.28
th(tbppy)py	Cl	(11)	1.22	0.86	-1.76	-2.42	0.66	2.62
th(tbppy)py	CCPh	(12)	0.74	0.42	-1.74	-2.38	0.64	2.16

th(tbppy)py	CCF ₅	(13)	0.91	0.32	-1.72	-2.30	0.58	2.04
bth(ppy)py	Cl	(14)	-	0.48	-1.65	-2.30	0.62	2.16
bth(tbppy)py	Cl	(15)	1.09	0.55	-1.70	-2.35	0.65	2.25
bth(tbppy)py	CCPh	(16)	0.42	0.19	-1.71	-2.33	0.62	1.90
th(tbppy)tz	Cl	(17)	0.75	0.59	-1.64	-2.33	0.69	2.23
th(tbppy)tz	CCPh	(18)	0.76	0.40	-1.64	-2.32	0.68	2.04

^a Plots in Figs S5-S26, this ESI. Electrochemical potentials [V] with an accuracy of ± 0.003 V, vs ferrocene/ferrocenium; $E_{1/2}$ for reversible redox waves, E_{pc} for irreversible reduction waves and E_{pa} for irreversible oxidation waves; measured in 0.1 M *n*Bu₄NPF₆/CH₂Cl₂ or THF at 298 K, scan rate 100 mV/s, oxidations measured in CH₂Cl₂, reductions in THF. ^b Measured in CH₂Cl₂.

Table S5 UV-vis absorption maxima of HC^{^N^N} protoligands.^a

HC ^{^N^N}	λ_1 (ϵ)	λ_2 (ϵ)	λ_3 (ϵ)	λ_4 (ϵ)	λ_5 (ϵ)	λ_6 (ϵ)
Hph(ppy)tz	257sh (2.96)	266 (3.42)	-	299sh (1.20)	325 (1.13)	344sh (0.46)
Hph(tbppy)tz	-	273 (3.65)	-	-	325 (1.21)	340sh (0.74)
Hna(tbppy)tz	-	266 (4.02)	286 (3.48)	299sh (2.71)	333 (1.14)	347sh (0.69)
Hph(ppy)btz	257 (4.25)	271 (3.99)	-	286sh (3.22)	329 (2.29)	342 (2.05)
Hph(tbppy)btz	258 (2.88)	275 (3.19)	-	-	329 (1.70)	341 (1.60)
Hth(ppy)py	250 (3.67)	280 (3.18)	-	306sh (1.88)	328 (1.46)	350sh (0.56)
Hth(tbppy)py	252 (1.98)	282 (2.63)	-	-	328 (0.92)	350sh (0.27)
Hbth(ppy)py	255 (2.53)	284 (2.22)	-	317sh (1.82)	327 (1.82)	353sh (0.75)
Hbth(tbppy)py	261 (2.39)	284 (2.85)	-	318sh (1.72)	327 (1.69)	354sh (0.44)
Hth(tbppy)tz	-	280 (3.15)	-	-	326 (1.16)	349 (0.45)

^a Measured in CH₂Cl₂ (298 K), sh = shoulder, wavelength λ in nm, extinction coefficients ϵ in 10⁴ M⁻¹cm⁻¹.

Table S6 UV-vis absorption maxima of the [Pt(C^{^N^N})X] complexes.^a

C ^{^N^N}	X		λ_1 (ϵ)	λ_2 (ϵ)	λ_3 (ϵ)	λ_4 (ϵ)	λ_5 (ϵ)	λ_6 (ϵ)	λ_7 (ϵ)	λ_8 (ϵ)	λ_9 (ϵ)	λ_{10} (ϵ)
ph(ppy)tz	Cl	(1)		-	296 (2.82)	321sh (1.71)	337 (1.86)	373 (1.25)	-	426sh (0.49)	447 (0.44)	501sh (0.09)
ph(tbppy)tz	Cl	(2)	271sh (1.99)		297 (2.81)		338 (2.02)	373 (1.04)	-	425sh (0.36)	448 (0.40)	499sh (0.05)
	CCPh	(3)	273sh (4.04)	285 (4.24)	298 (4.14)		339 (2.22)	377 (1.58)	-	444sh (0.67)	466 (0.75)	545sh (0.07)
	CCF ₅	(4)		274sh (2.88)	298 (3.29)		338 (1.95)	371 (1.39)	-	435sh (0.66)	454 (0.68)	509sh (0.06)
na(tbppy)tz	Cl	(5)	250 (2.61)		297 (2.96)		337 (2.19)	369 (1.22)		430sh (0.44)	445 (0.49)	500sh (0.02)
	CCPh	(6)		284 (5.09)	300 (5.29)	308sh (5.12)	331sh (3.02)	369 (1.88)	398sh (1.32)	413sh (0.81)	463 (0.88)	527sh (0.18)
ph(ppy)btz	Cl	(7)		-	305 (2.86)	320sh (2.48)	352 (2.46)	384 (1.26)	-	440sh (0.32)	466 (0.35)	515sh (0.08)
ph(tbppy)btz	Cl	(8)		271sh (3.58)	306 (4.34)	323sh (3.07)	350 (3.87)	382sh (1.91)	-	438sh (0.46)	466 (0.51)	514sh (0.10)
	CCPh	(9)		270 (3.54)	301 (3.16)	323sh (2.28)	348 (2.28)	384 (1.51)	-	476 (0.56)	505 (0.52)	574sh (0.05)
th(ppy)py	Cl	(10)		-	287sh (3.39)	298 (3.58)	335sh (1.86)	367 (1.51)	384sh (1.35)	-	435 (0.41)	491sh (0.04)
th(tbppy)py	Cl	(11)		-	284 (2.23)	310 (2.66)	336sh (1.59)	366 (1.19)	380sh (1.11)	-	435 (0.33)	489sh (0.04)
	CCPh	(12)		273 (2.42)	290 (2.40)	306 (2.19)	323sh (1.85)	373 (1.26)	-	440 (0.51)	467sh (0.45)	517sh (0.04)
	CCF ₅	(13)		270sh (2.71)	285 (3.23)	313 (2.87)	341sh (1.89)	371 (1.72)	-	424sh (0.76)	443 (0.79)	501sh (0.05)

bth(ppy)py	Cl	(14)	-	291 (3.64)		351 (1.97)	385 (1.66)	-	-	439 (0.27)	495sh (0.06)
bth(tbppy)py	Cl	(15)	-	287 (3.96)	310 (3.43)	349 (2.50)	387 (1.96)	-	-	440 (0.34)	496 (0.07)
	CCPh	(16)	272sh (2.98)	287 (3.02)	296 (2.91)	355 (1.64)	381 (1.53)	-	407 (1.28)	469 (0.43)	541sh (0.02)
th(tbppy)tz	Cl	(17)	-	290sh (2.48)	306 (2.92)	327sh (2.15)	359 (1.64)	380sh (1.15)	-	436 (0.35)	505sh (0.05)
	CCPh	(18)	286 (2.86)	299 (2.80)	300 (2.79)	318sh (2.21)	367 (2.18)	381sh (1.55)	441 (0.68)	468 (0.63)	560sh (0.06)

^a Measured in CH₂Cl₂ (298 K), sh = shoulder, wavelength λ in nm, extinction coefficients ϵ in 10⁴ M⁻¹cm⁻¹.

Table S7 Complete emission data and PLYQs, as well as excited state lifetime data for each complex in CH₂Cl₂ at 298 K and in frozen glassy matrix of CH₂Cl₂/MeOH (V:V = 1:1) at 77 K. For multiexponential decays, the amplitude-weighted average lifetimes are given as well as the different components in square brackets with relative amplitudes as percentages in parentheses.

class	Complex	Medium (T/K)	λ_{em}	λ_{exc}	τ_{av}^a	$\Phi \pm 0.02/\pm 0.05$
B	[Pt(th(ppy)py)Cl] (10)	CH ₂ Cl ₂ , air (298)	621, 660 sh	341, 450	n.d.	< 0.02
		CH ₂ Cl ₂ , Ar (298)			5.04 ± 0.24 [5.49 ± 0.06 (89); 1.3 ± 0.6 (11)]	0.05
		glassy matrix (77)	598, 650, 710	320, 367, 426	11.0 ± 0.4 [12.34 ± 0.10 (79); 6.1 ± 0.5 (21)]	0.39
	[Pt(bth(ppy)py)Cl] (14)	CH ₂ Cl ₂ , air (298)	675, 745	350, 400, 450sh	n.d.	< 0.02
		CH ₂ Cl ₂ , Ar (298)			1.022 ± 0.029 [3.0 ± 0.7 (1); 1.015 ± 0.009 (99)]	0.02
		glassy matrix (77)	655, 720, 800	373, 400	4.71 ± 0.17 [5.60 ± 0.10 (35); 4.23 ± 0.06 (65)]	0.23
	[Pt(th(tbppy)py)Cl] (11)	CH ₂ Cl ₂ , air (298)	617, 660 sh	315, 370, 440	0.3393 ± 0.0022	< 0.02
		CH ₂ Cl ₂ , Ar (298)			1.809 ± 0.013	0.07
		glassy matrix (77)	595, 645, 700	375, 425, 470sh, 490sh	11.6 ± 0.9 [13.17 ± 0.14 (59); 9.32 ± 0.26 (41)]	0.58
	[Pt(bth(tbppy)py)Cl] (15)	CH ₂ Cl ₂ , air (298)	675, 730	350, 395, 440sh	0.3019 ± 0.0025	< 0.02
		CH ₂ Cl ₂ , Ar (298)			1.069 ± 0.008	0.03
		glassy matrix (77)	655, 720, 745	375, 480	4.16 ± 0.16 [5.64 ± 0.12 (21); 3.76 ± 0.05 (79)]	0.13
	[Pt(th(tbppy)py)CCPh] (12)	CH ₂ Cl ₂ , air (298)	620, 660sh	300, 373, 440, 475sh	0.3146 ± 0.0020	0.04
		CH ₂ Cl ₂ , Ar (298)			1.628 ± 0.012	0.14
		glassy matrix (77)	590, 640, 700sh	360, 425sh, 460	8.2 ± 0.3 [9.69 ± 0.15 (37); 7.39 ± 0.11 (63)]	0.79
	[Pt(bth(tbppy)py)CCPh] (16)	CH ₂ Cl ₂ , air (298)	675, 740	360, 380sh, 405sh, 475	0.2772 ± 0.0021	< 0.02
		CH ₂ Cl ₂ , Ar (298)			1.110 ± 0.008	0.04
		glassy	650, 720, 785	365, 400sh, 460	3.57 ± 0.13 [6.00 ± 0.23 (8);	0.24

		matrix (77)				3.34 ± 0.03 (92)]	
	[Pt(th(tbpppy)py)CC ₆ F ₅] (13)	CH ₂ Cl ₂ , air (298)	620, 660 sh	312, 370, 425sh, 445	0.3513 ± 0.0023		0.03
CH ₂ Cl ₂ , Ar (298)		2.178 ± 0.016			0.15		
glassy matrix (77)		595, 650, 705	350-375, 430	8.7 ± 0.3 [10.08 ± 0.09 (71); 5.5 ± 0.3 (29)]		0.79	
A	[Pt(ph(tbpppy)btz)Cl] (8)	CH ₂ Cl ₂ M, air (298)	635, 675sh	353, 464	0.1832 ± 0.0014		< 0.02
		CH ₂ Cl ₂ , Ar (298)			0.2057 ± 0.0018		< 0.02
		glassy matrix (77)	600, 656, 710sh	346, 425, 452	1.59 ± 0.06 [1.871 ± 0.022 (53); 1.28 ± 0.03 (47)]		0.07
	[Pt(ph(tbpppy)tz)CCPh] (3)	CH ₂ Cl ₂ , air (298)	626, 645sh	302, 346, 380, 443sh, 473, 495sh	0.1891 ± 0.0014		0.03
		CH ₂ Cl ₂ , Ar (298)			0.39 ± 0.04 [0.78 ± 0.04(3); 0.375 ± 0.003 (97)]		0.06
		glassy matrix (77)	558, 605, 660sh	336, 370, 425, 438, 468	3.50 ± 0.12 [5.59 ± 0.24 (8); 3.31 ± 0.03 (92)]		0.62
	[Pt(na(tbpppy)tz)CCPh] (6)	CH ₂ Cl ₂ , air (298)	635	311, 335sh, 370, 398, 420sh, 468	0.200 ± 0.007 [0.2105 ± 0.0016 (84); 0.147 ± 0.011 (16)]		0.05
		CH ₂ Cl ₂ , Ar (298)			0.407 ± 0.014 [0.97 ± 0.07 (2); 0.395 ± 0.004 (98)]		0.10
		glassy matrix (77)	555, 595, (650sh)	330, 395, 437, 463	3.97 ± 0.16 [4.29 ± 0.04 (67); 3.33 ± 0.12 (33)]		0.82
	[Pt(ph(tbpppy)btz)CCPh] (9)	CH ₂ Cl ₂ , air (298)	655	360, 470, 500sh	0.089 ± 0.003 [0.116 ± 0.004 (14); 0.0844 ± 0.0008 (86)]		< 0.02
		CH ₂ Cl ₂ , Ar (298)			0.112 ± 0.004 [0.262 ± 0.019 (2); 0.1085 ± 0.0010 (98)]		< 0.02
		glassy matrix (77)	600, 650, 700sh	360, 470, 500sh	1.70 ± 0.06 [2.25 ± 0.04 (25); 1.52 ± 0.02 (75)]		0.18
[Pt(ph(tbpppy)tz)CC ₆ F ₅] (4)	CH ₂ Cl ₂ , air (298)	605, 635sh	345, 370, 445sh, 465	0.3195 ± 0.0025		0.05	
	CH ₂ Cl ₂ , Ar (298)			0.91 ± 0.03 [2.7 ± 0.4 (1); 0.899 ± 0.008 (99)]		0.12	
	glassy matrix (77)	555, 600, 655sh, 725sh	340, 365, 420, 445	3.53 ± 0.12 [4.61 ± 0.14 (13); 3.37 ± 0.03 (87)]		0.49	
C	[Pt(th(tbpppy)tz)Cl] (17)	CH ₂ Cl ₂ , air (298)	655	368, 450	0.1351 ± 0.0009		< 0.02
		CH ₂ Cl ₂ , Ar (298)			0.220 ± 0.010 [0.8 ± 0.8 (0.4); 0.2173 ± 0.0020 (99.6)]		< 0.02
		glassy matrix (77)	610, 660, 725sh	364, 430, 503sh	4.17 ± 0.15 [6.25 ± 0.07 (35); 3.04 ± 0.06 (65)]		0.13
	[Pt(th(tbpppy)tz)CCPh] (18)	CH ₂ Cl ₂ , air (298)	658	375, 470	0.1361 ± 0.0009		< 0.02
		CH ₂ Cl ₂ , Ar (298)			0.2474 ± 0.0016		< 0.02
		glassy matrix (77)	604, 657, 725sh	365, 425, 454	4.19 ± 0.14 [5.40 ± 0.06 (45); 3.22 ± 0.06 (55)]		0.35

^a We attribute the biexponential lifetimes in fluid solution to different conformations that can be thermally accessed while possessing comparable excited-state character yet different deactivation rates. Moreover, diverse emitting states could coexist in thermal equilibrium, which might be influenced by the microenvironment affecting the excited states' character (particularly in frozen glassy matrices or in solid samples).

Table S8. Selected photophysical properties of previously reported Pt(II)-complexes.^a

class	complex	λ_{em} (nm)	τ (μ s)	Φ_L ^b
	[Pt(ph(tbppy)2-quin)Cl]	624	<0.2 μ s	<0.001
D	[Pt(ph(tbppy)1-isoquin)Cl]	621	<0.2	<0.001
	[Pt(ph(tbppy)3-isoquin)Cl]	530, 566(sh)	7.6	0.78
	[Pt(ph(tbppy)3-isoquin)(CCC ₆ F ₅)]	521, 556(sh)	7.6	0.99
	[Pt(2,4-F ₂ ph(tbppy)3-isoquin)Cl] ^c	528, 565(sh)	2.0	0.14
	[Pt(2-MeOph(tbppy)3-isoquin)Cl] ^c	585	3.3	0.23
	[Pt(2-CF ₃ ph(tbppy)3-isoquin)Cl] ^c	534, 569(sh)	8.9	0.63
E	[Pt(2,3F ₂ ph(tbppy)3-isoquin)Cl] ^c	529, 565(sh)	8.9	0.65
	[Pt(1,3-F ₂ ph(tbppy)3-isoquin)Cl] ^c	529, 561(sh)	10.1	0.69
	[Pt(1-F,3-MeOph(tbppy)3-isoquin)Cl] ^c	529, 561(sh)	9.5	0.73
	[Pt(2-NO ₂ ph(tbppy)3-isoquin)Cl] ^c	529, 565, 622 (sh)	6.7	0.79
	[Pt(th(tbppy)3-isoquin)Cl]	588, 632, 687(sh)	10.1	0.25
	[Pt(th(tbppy)3-isoquin)(CCC ₆ F ₅)]	591, 636, 680(sh)	8.7	0.66
	[Pt(bth(tbppy)3-isoquin)Cl]	644, 701, 780(sh)	4.5	0.033
	[Pt(bth(tbppy)3-isoquin)(CCC ₆ F ₅)]	653, 710(sh)	4.6	0.082
F	[Pt(na(tbppy)py]	561, 597(sh)	0.7	0.10
	[Pt(ph(tbppy)3-isoquin)(CC4-tolyl)]	573	3.5	0.27

^a From ref. 1, measured in CH₂Cl₂ solution at 298 K; quin = 2-quinolinyl, 3-isoquin = 3-isoquinolinyl, 1-isoquin = 1-isoquinolyl.

^b Φ_L measured by the optical dilute method with [Ru(bpy)₃](PF₆)₂ (bpy = 2,2'-bipyridine) in degassed CH₃CN as standard ($\Phi_L = 0.062$) and calculated by: $\Phi_{Ls} = \Phi_{Lr} (B_r/B_s)(n_s/n_r)^2(D_s/D_r)$, s and r refer to sample and reference, n is the refractive index of the solvents, D is the integrated intensity and Φ_L is the emission quantum yield. The excitation intensity B is calculated by: $B = 1 - 10^{-ad}$, where a is the absorbance at the excitation wavelength and d is the optical path length ($d = 1$ cm in all cases). Errors for Φ_L values (+/-10%) are estimated. ^c Nomenclature: 1 represents the position *ortho* to the Pt bound C atom.

Table S9 TDDFT calculated transitions contributing to the emissive T_1 state.^a

complex	transitions	contribution [%]
1	HOMO-7 \rightarrow LUMO	22.6
	HOMO \rightarrow LUMO	61.3
2	HOMO-7 \rightarrow LUMO	19.9
	HOMO-1 \rightarrow LUMO	28.5
	HOMO \rightarrow LUMO	27.3
3	HOMO-8 \rightarrow LUMO	20.5
	HOMO-4 \rightarrow LUMO	10.4
	HOMO-1 \rightarrow LUMO	22.7
	HOMO \rightarrow LUMO	31.2
4	HOMO-7 \rightarrow LUMO	28.0
	HOMO-1 \rightarrow LUMO	29.2
	HOMO \rightarrow LUMO	21.4
6	HOMO-8 \rightarrow LUMO	25.4
	HOMO-3 \rightarrow LUMO	17.4
	HOMO \rightarrow LUMO	35.2
9	HOMO-7 \rightarrow LUMO	15.2
	HOMO-2 \rightarrow LUMO	57.6
11	HOMO \rightarrow LUMO	56.1
	HOMO \rightarrow LUMO+1	37.9
12	HOMO \rightarrow LUMO	53.7

	HOMO → LUMO+1	35.0
13	HOMO → LUMO	56.7
	HOMO → LUMO+1	35.1
16	HOMO → LUMO	63.9
	HOMO → LUMO+1	31.9
18	HOMO-8 → LUMO	21.6
	HOMO-1 → LUMO	44.5
	HOMO → LUMO	14.6

^a Based on the optimised T_1 geometry. Only contributions with a weight > 10% are listed.

References

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