

## Supplementary Information

### **Accelerated Fenton degradation for azo dyes wastewater via a novel Z-scheme CoFeN-g-C<sub>3</sub>N<sub>4</sub> heterojunction photocatalyst with excellent charge transfer under visible light irradiation**

Shihao Miao<sup>a</sup>, Hongcheng Gao<sup>\*a</sup>, Hongyu Xia<sup>a</sup>, Xiaoxia Mao<sup>a</sup>, Lijuan Zhang<sup>a</sup>, Mengqin Shi<sup>a</sup>  
and Yuanguang Zhang<sup>\*a,b</sup>

<sup>a</sup>Key Laboratory of Aqueous Environment Protection and Pollution Control of Yangtze River in Anhui of Anhui, Provincial Education Department, College of Resources and Environment, Anqing Normal University, Anqing 246011, China

<sup>b</sup>College of Chemistry and Chemical Engineering, Anqing Normal University, Anqing 246011, China

\*Corresponding Authors

Email: hongchenggao@aqnu.edu.cn (H. Gao); zygaqtc@163.com (Y. Zhang)



**Fig. S1** Multifunctional photocatalytic reactor.

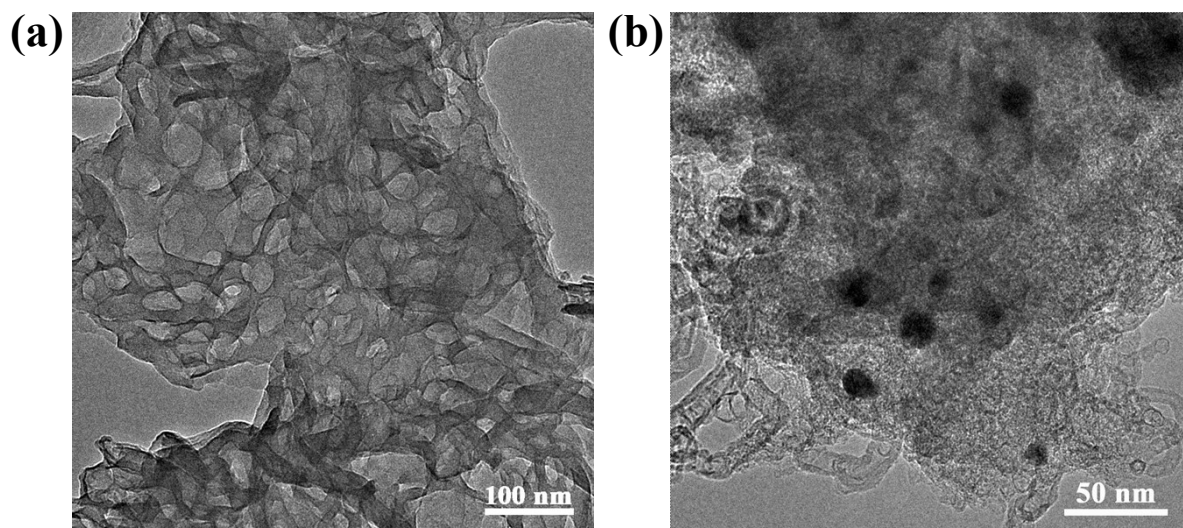


Fig. S2 TEM images of (a) CN and (b) CFN.

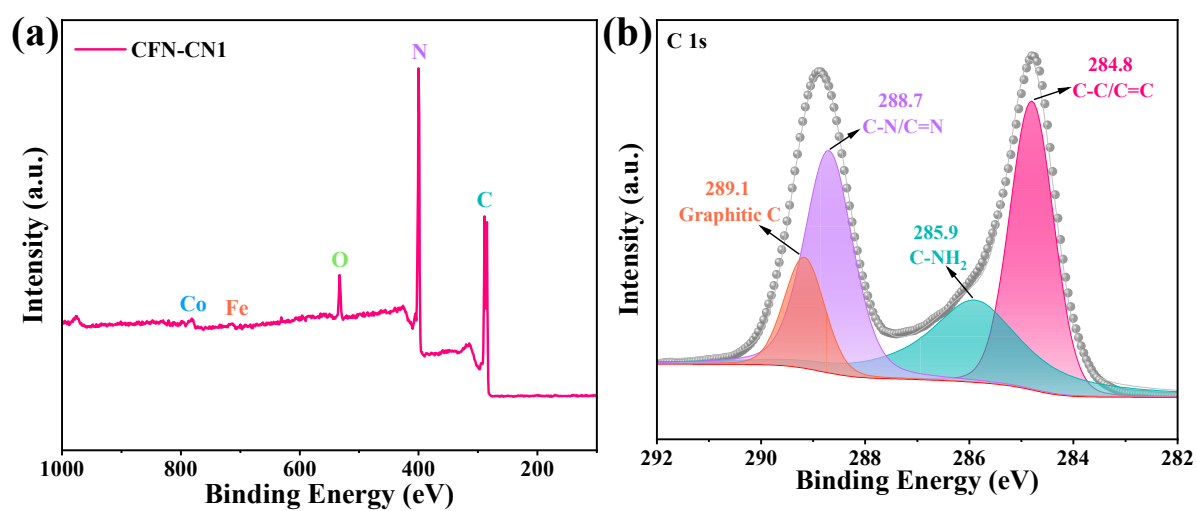
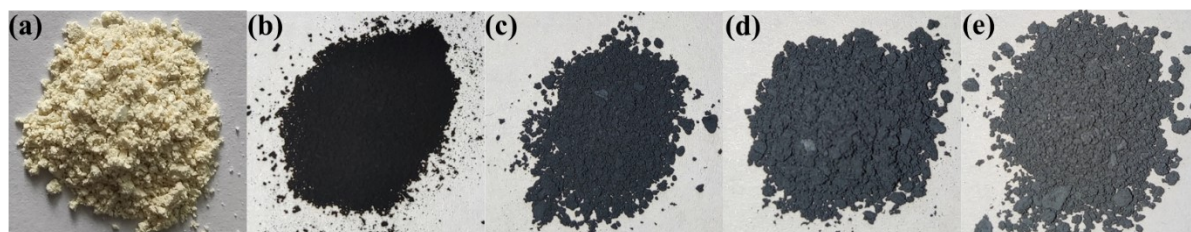
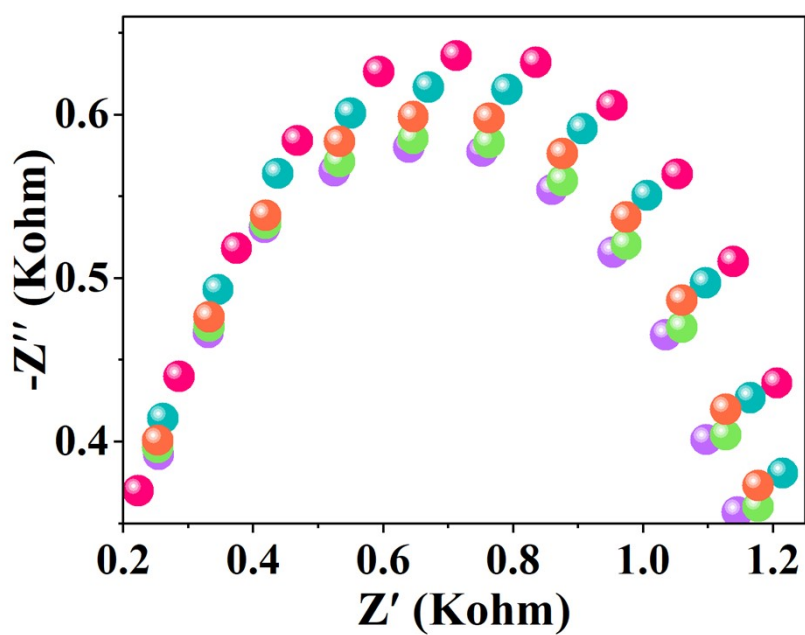


Fig. S3 (a) Full XPS spectra of CFN-CN1, (c) C 1s High-resolution XPS spectrum of CFN-CN1.



**Fig. S4** Photos of all samples (a) CN, (b) CFN, (c) CFN-CN1, (d) CFN-CN2 and (e) CFN-CN3.



**Fig. S5** EIS Nyquist plot of the catalysts.

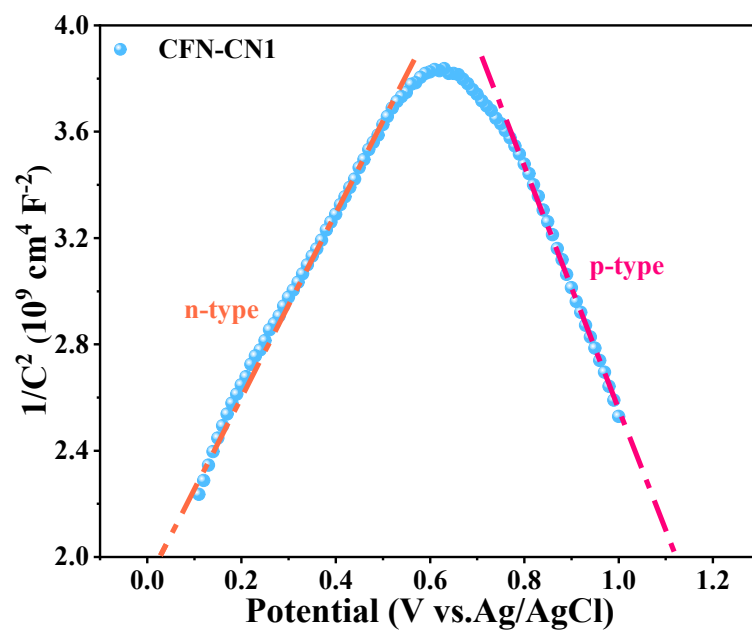


Fig. S6 M-S plots vs. Ag/AgCl CFN-CN1.

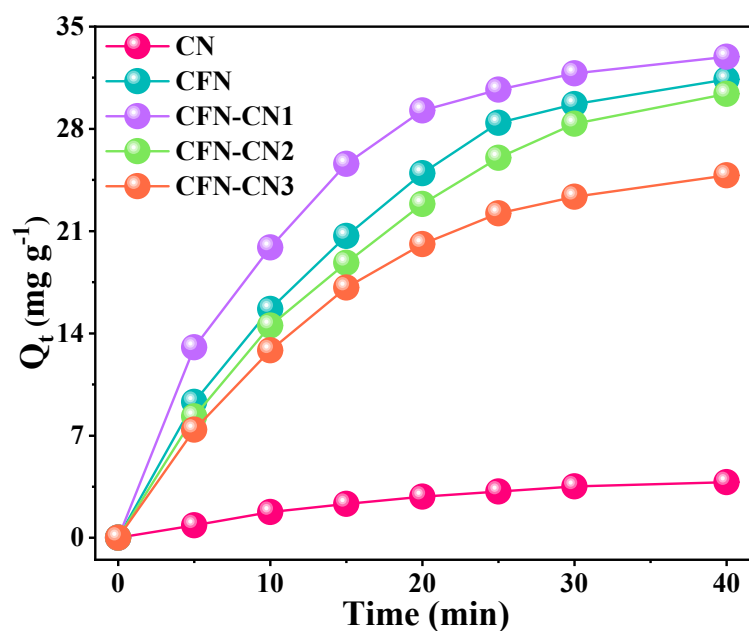
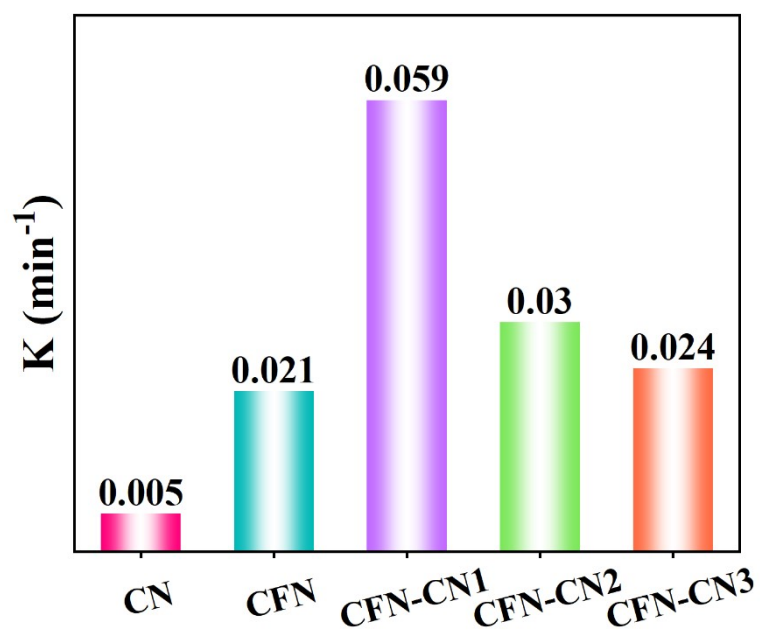
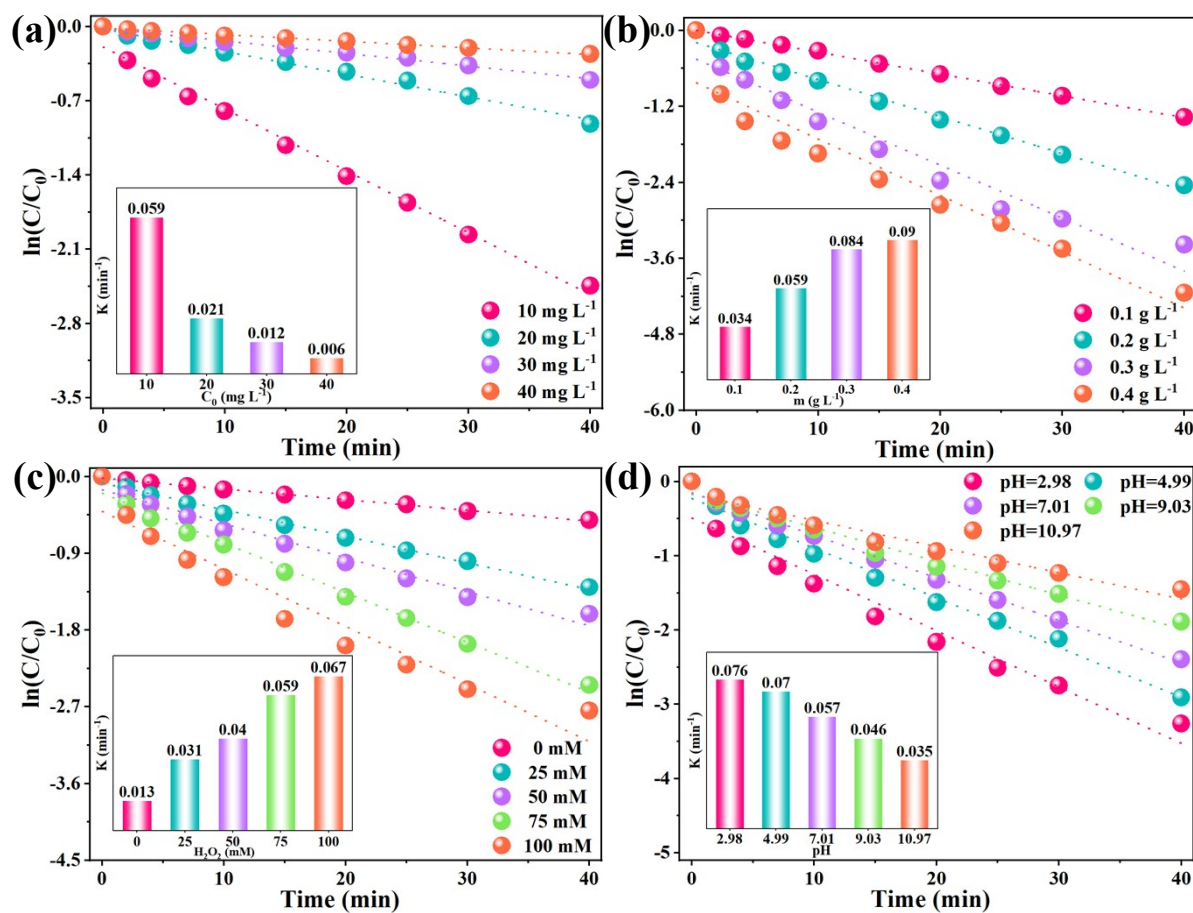


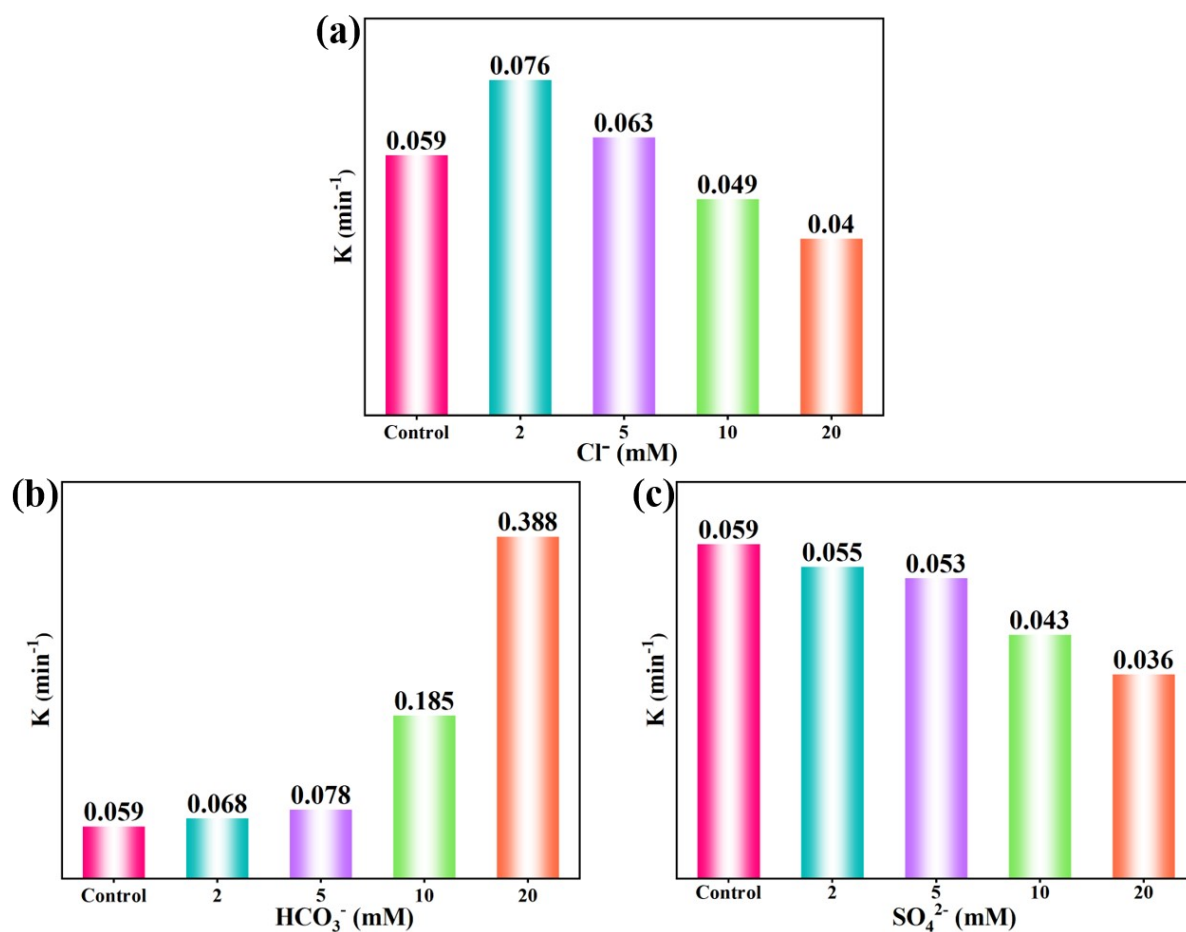
Fig. S7 Comparison of adsorption performance of different catalysts. Experimental conditions:  $[\text{MO}]_0 = 10 \text{ mg L}^{-1}$ ,  $[\text{catalyst}] = 0.2 \text{ g L}^{-1}$ , natural pH.



**Fig. S8** Degradation rate constants of MO with different catalysts. Experimental conditions:  $[\text{MO}]_0 = 10 \text{ mg L}^{-1}$ ,  $[\text{catalyst}] = 0.2 \text{ g L}^{-1}$ ,  $[\text{H}_2\text{O}_2] = 75 \text{ mM}$ , natural pH.

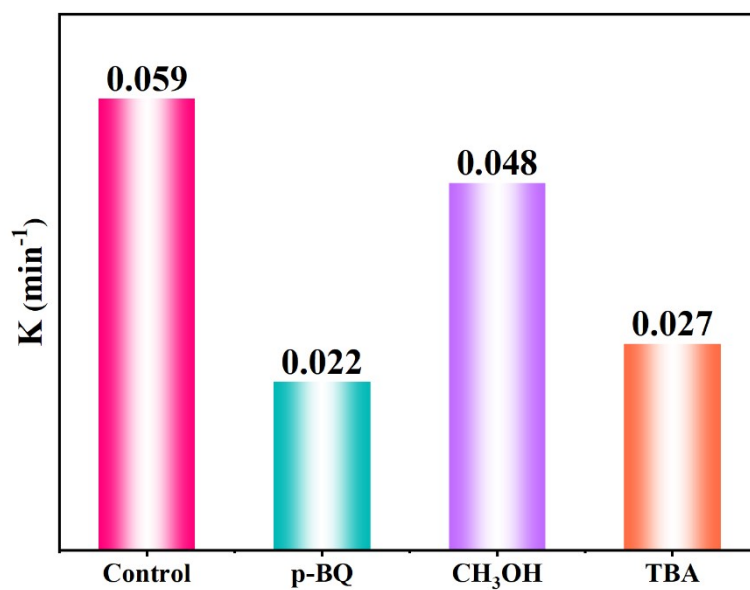


**Fig. S9** Kinetic analysis curves (a) original MO concentration, (b) catalyst dosage, (c) H<sub>2</sub>O<sub>2</sub> concentration, (d) pH, the illustration in (a), (b), (c) and (d) are the corresponding degradation rate constant. Experimental conditions: [MO]<sub>0</sub> = 10 mg L<sup>-1</sup>, [catalyst] = 0.2 g L<sup>-1</sup>, [H<sub>2</sub>O<sub>2</sub>] = 75 mM, natural pH.



**Fig. S10** Degradation rate constants of different inorganic anions (a)  $\text{Cl}^-$ , (b)  $\text{HCO}_3^-$  and (c)  $\text{SO}_4^{2-}$ . Experimental conditions:  $[\text{MO}]_0 = 10 \text{ mg L}^{-1}$ ,  $[\text{catalyst}] = 0.2 \text{ g L}^{-1}$ ,  $[\text{H}_2\text{O}_2] = 75 \text{ mM}$ , natural pH.





**Fig. S11** Degradation rate constants for the addition of different active species traps. Experimental conditions:  $[\text{MO}]_0 = 10 \text{ mg L}^{-1}$ ,  $[\text{catalyst}] = 0.2 \text{ g L}^{-1}$ ,  $[\text{H}_2\text{O}_2] = 75 \text{ mM}$ , natural pH.

**Table S1**

BET specific surface area and pore volume of catalyst

	BET Surface Area ( $\text{m}^2 \text{ g}^{-1}$ )	Pore Volume ( $\text{cm}^3 \text{ g}^{-1}$ )
CN	36.203	0.109
CFN	357.47	0.201
CFN-CN1	249.45	0.212
CFN-CN2	194.467	0.191
CFN-CN3	139.638	0.155