

Supplementary information

**Effect of crystal size on the acidity of nanometric Y zeolite: number of sites, strength, acid nature, and dehydration of 2-propanol**

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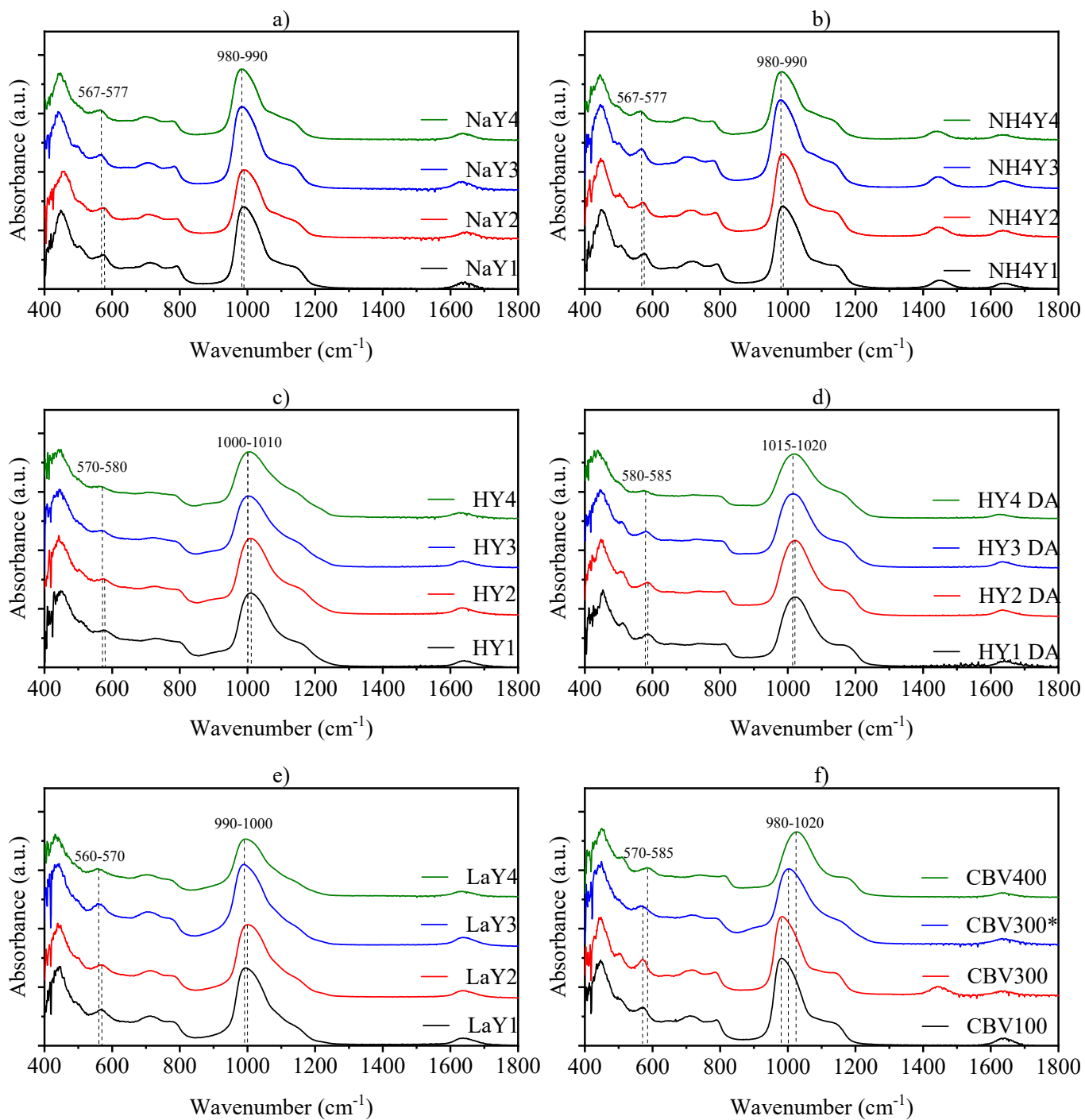
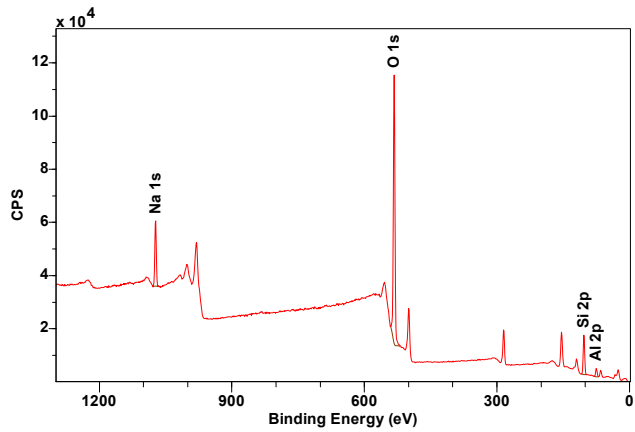
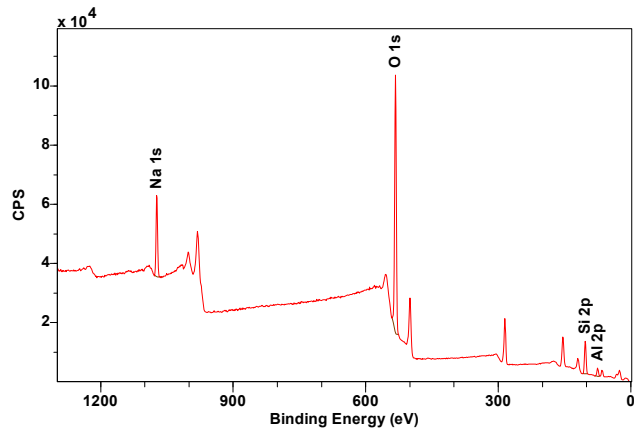


Fig. 15. FT-IR spectra of a) NaY, b) NH4Y, c) HY, d) HY DA, e) LaY and f) Commercial zeolites. CBV300\*: calcination.

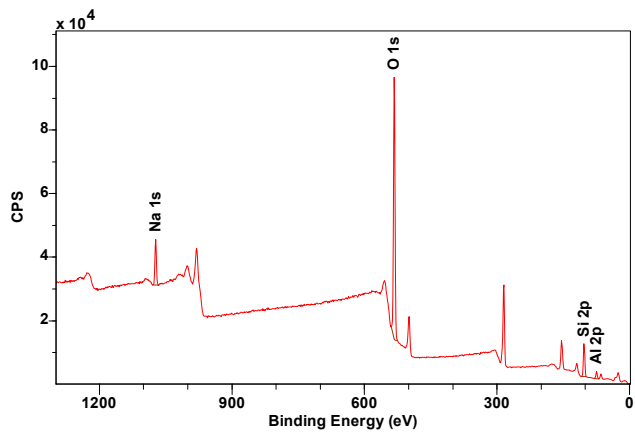
a)  
Survey



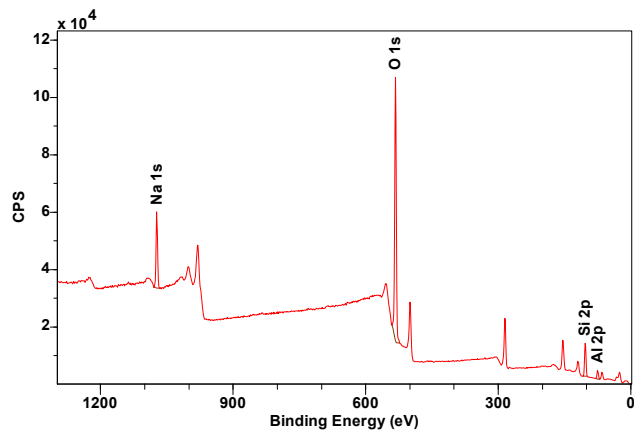
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Survey



c)  
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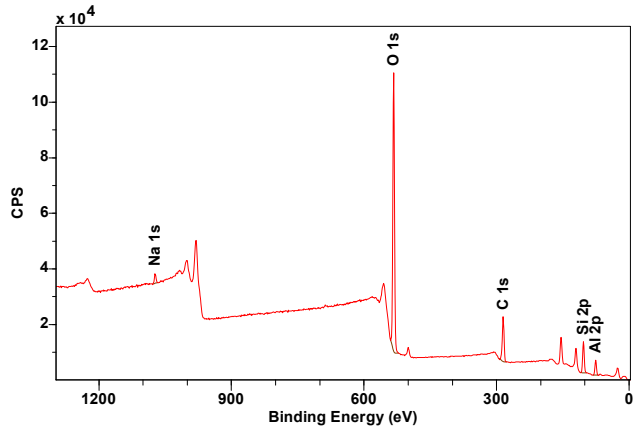


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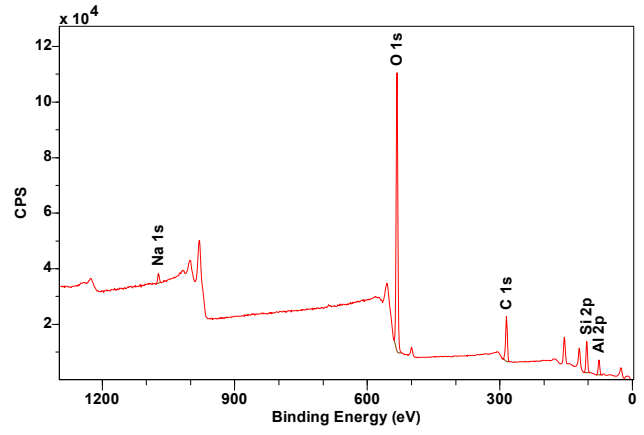
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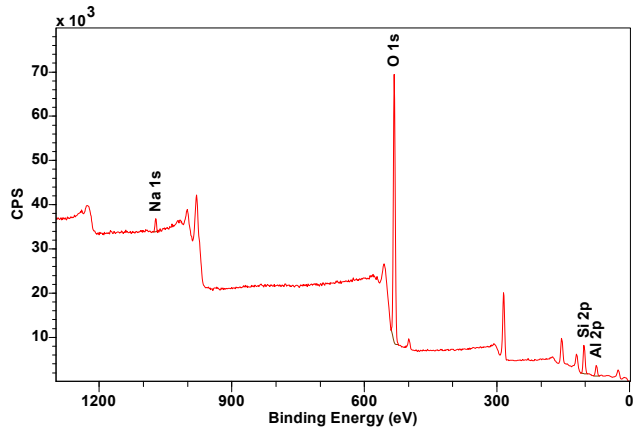
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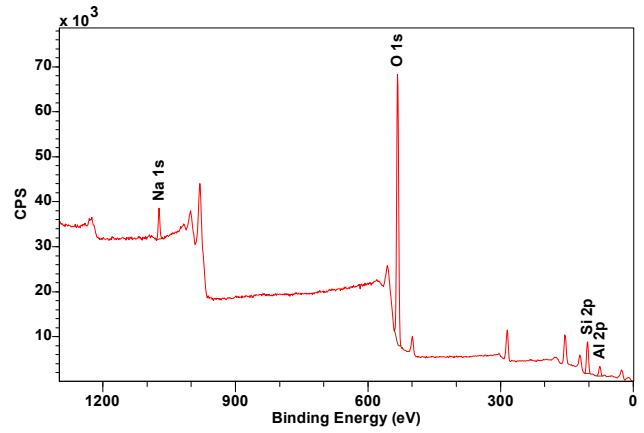
g)

Survey



h)

Survey



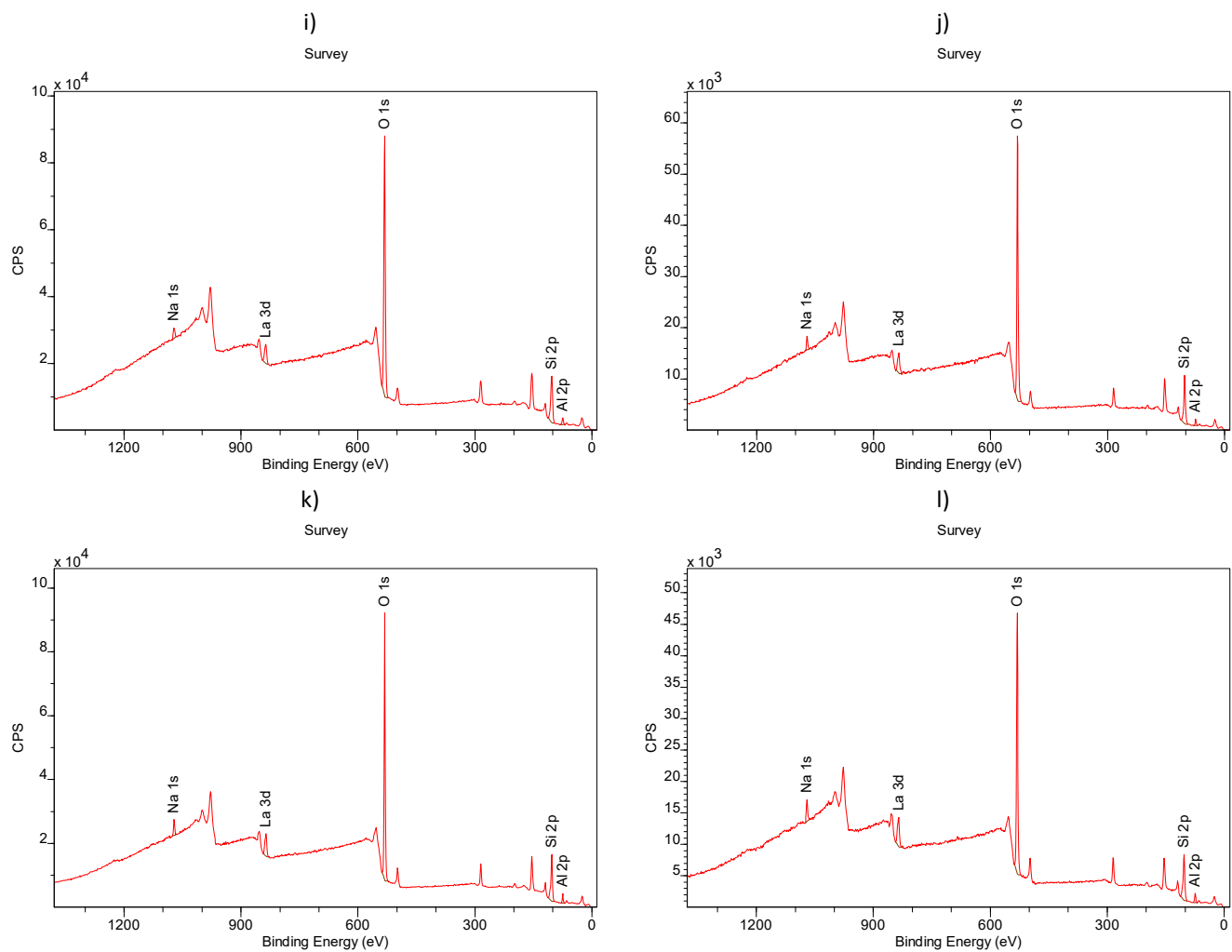


Fig. 2S. XPS spectra of a) NaY1, b) NaY2, c) NaY3, d) NaY4, e) HY1 DA, f) HY2 DA, g) HY3 DA and h) HY4 DA, i) LaY1, j) LaY2, k) LaY3 and l) LaY4

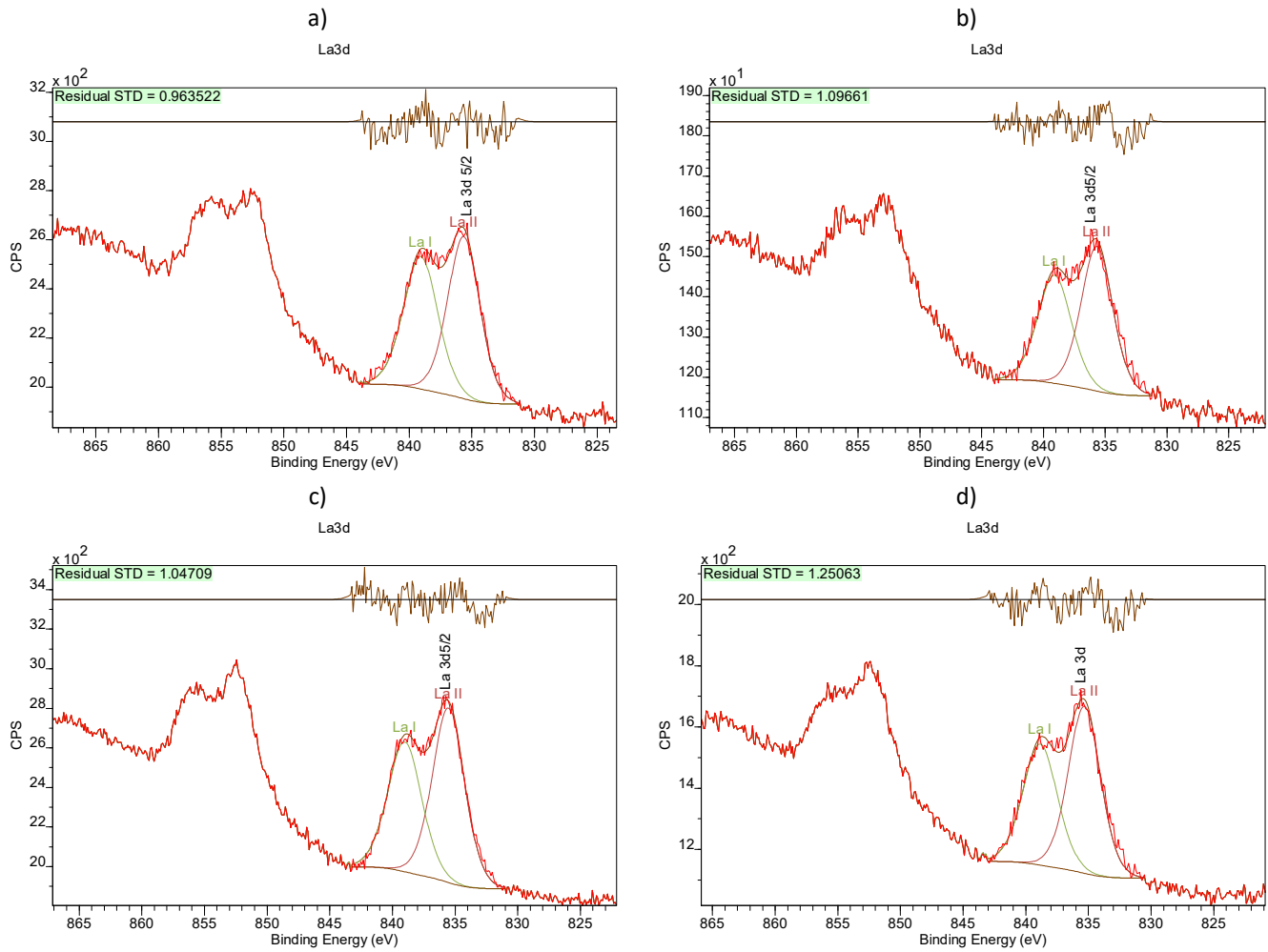


Fig. 3S. XPS spectra 3d5/2 deconvolution of a) LaY1, b) LaY2, c) LaY3 and d) LaY4

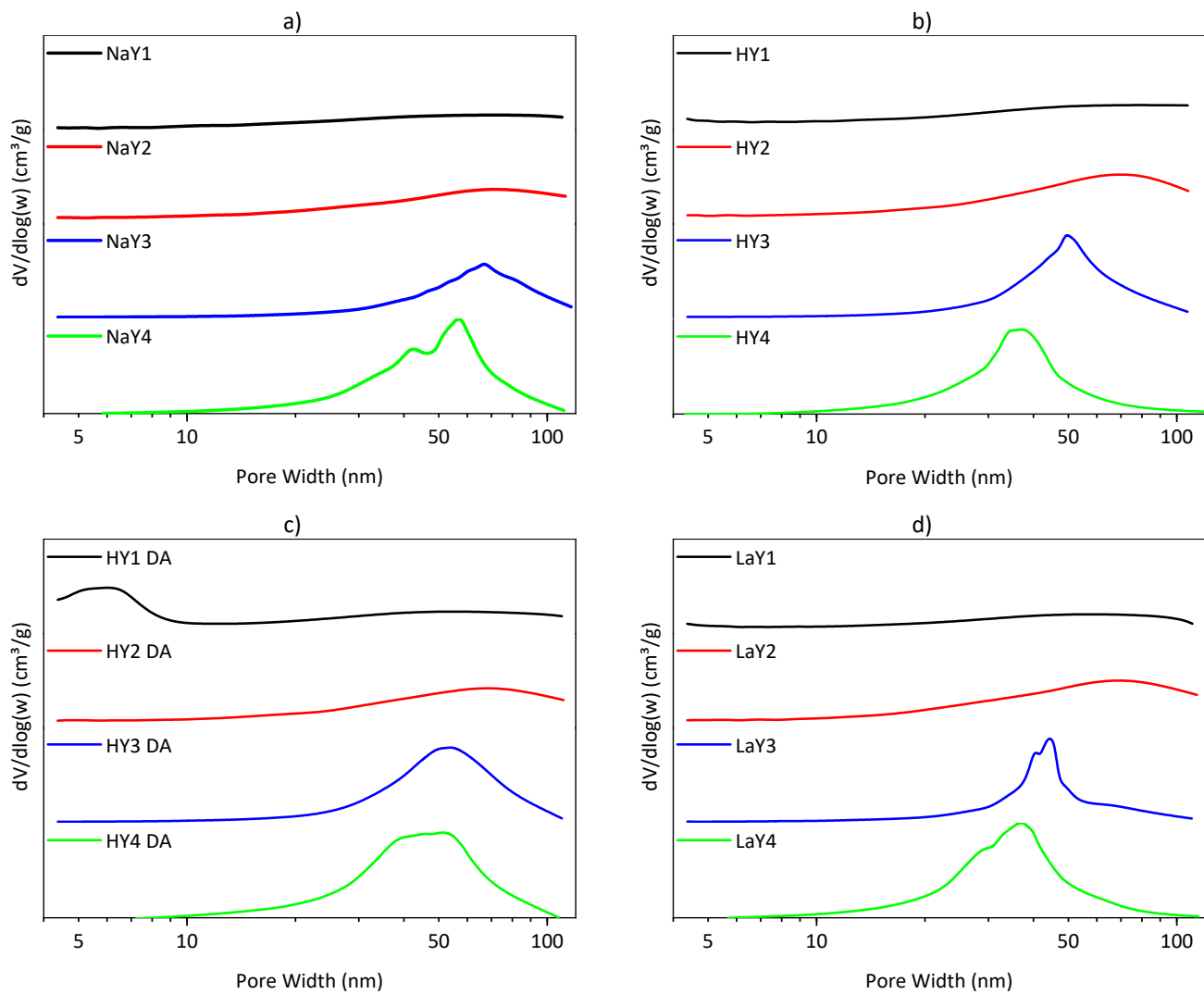


Fig. 4S. Pore size distribution of a) NaY, b) HY, c) HY DA and d) LaY

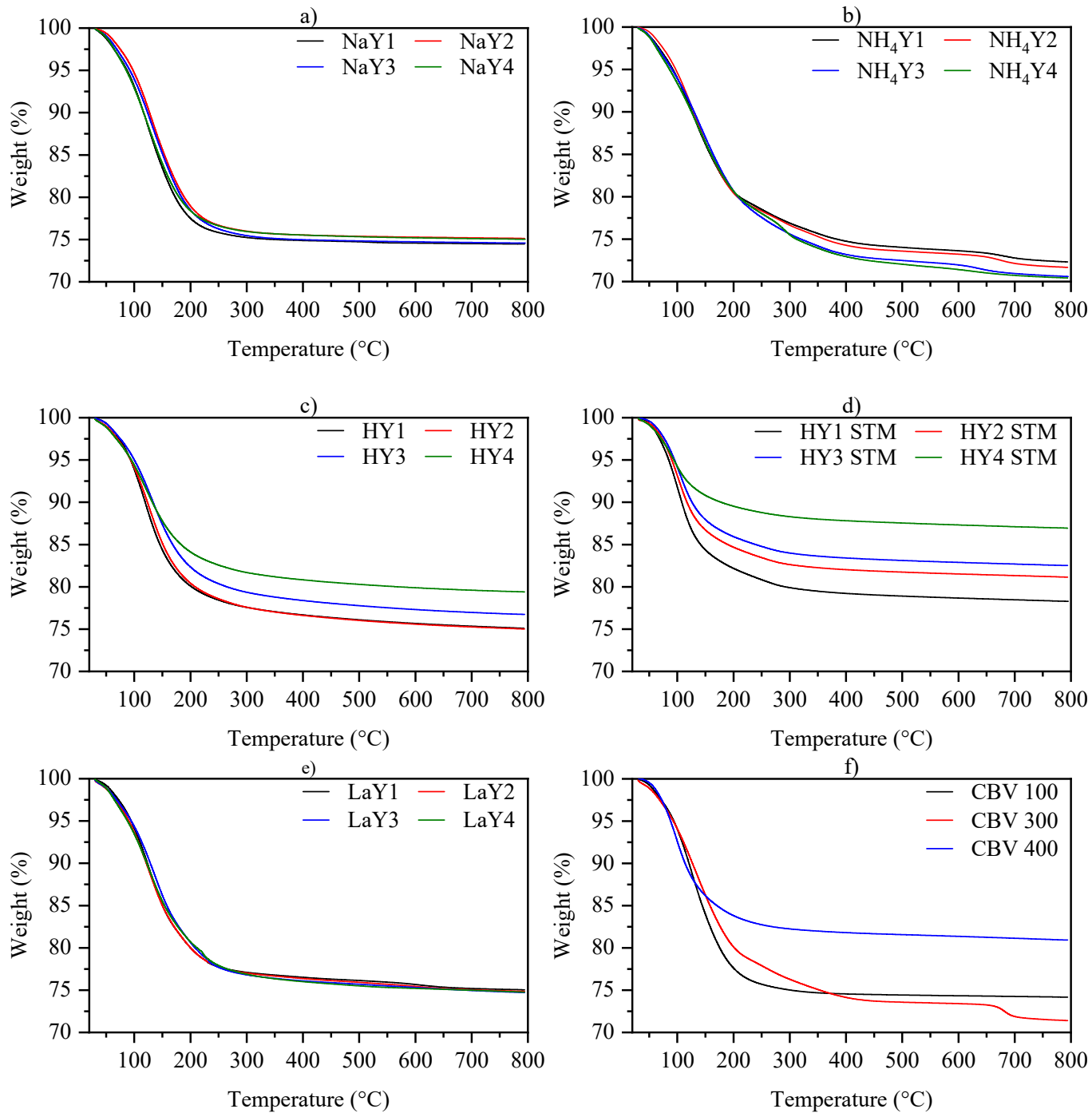


Fig. 5S. Thermograms of a) NaY, b) NH<sub>4</sub>Y, c) HY, d) HY DA, e) LaY and f) commercial zeolites



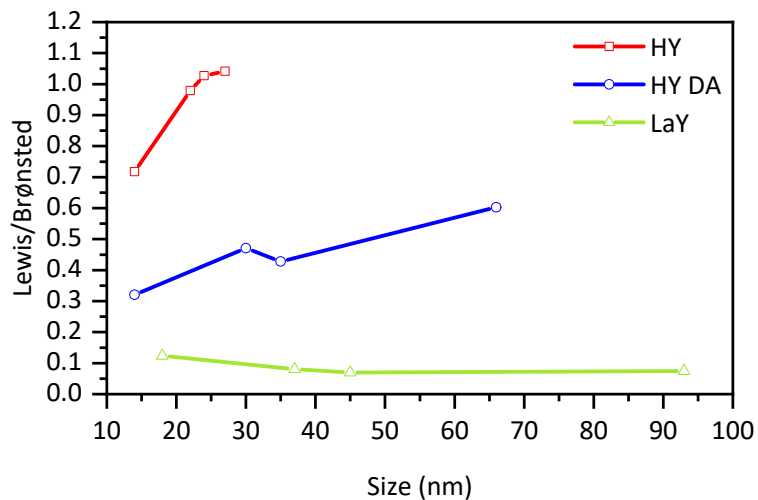


Fig. 6S. Lewis/Bronsted vs crystallite size of acid zeolites

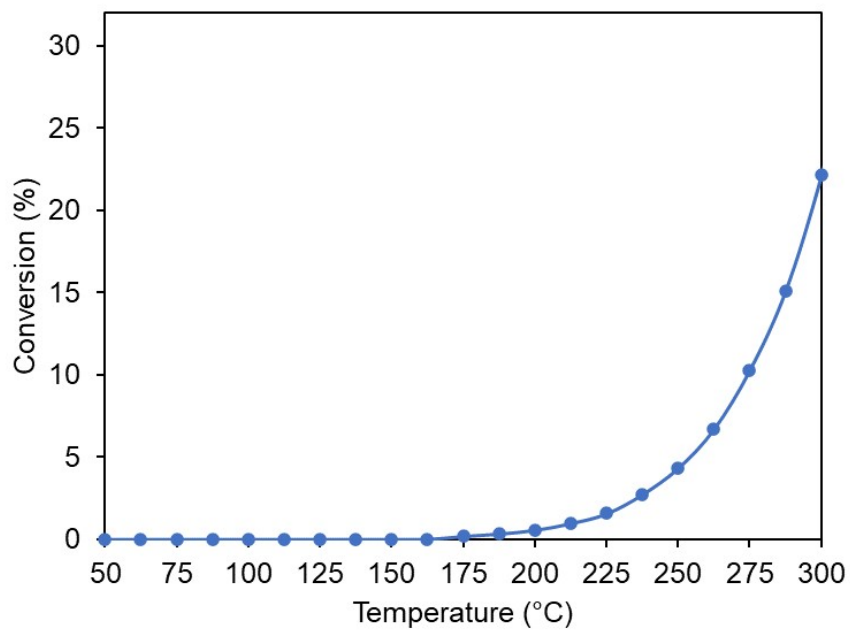


Fig. 7S. Thermal decomposition of 2-propanol.