

Supporting Information

Tunable fluorescent amino-functionalized $\text{Ti}_3\text{C}_2\text{T}_x$ MXene quantum dots for ultrasensitive Fe^{3+} ion sensing

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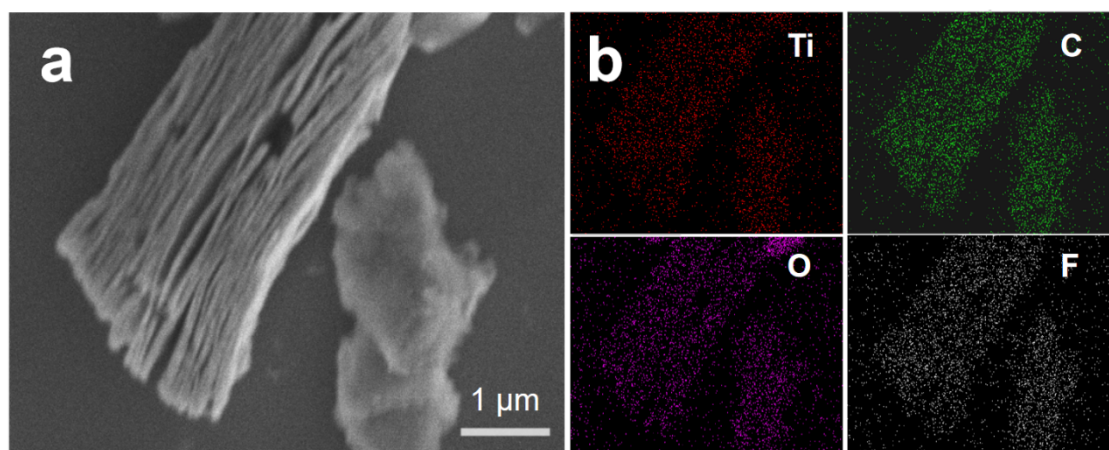


Figure S1. (a) SEM image of $\text{Ti}_3\text{C}_2\text{T}_x$ and (b) the corresponding EDS elemental mapping images.

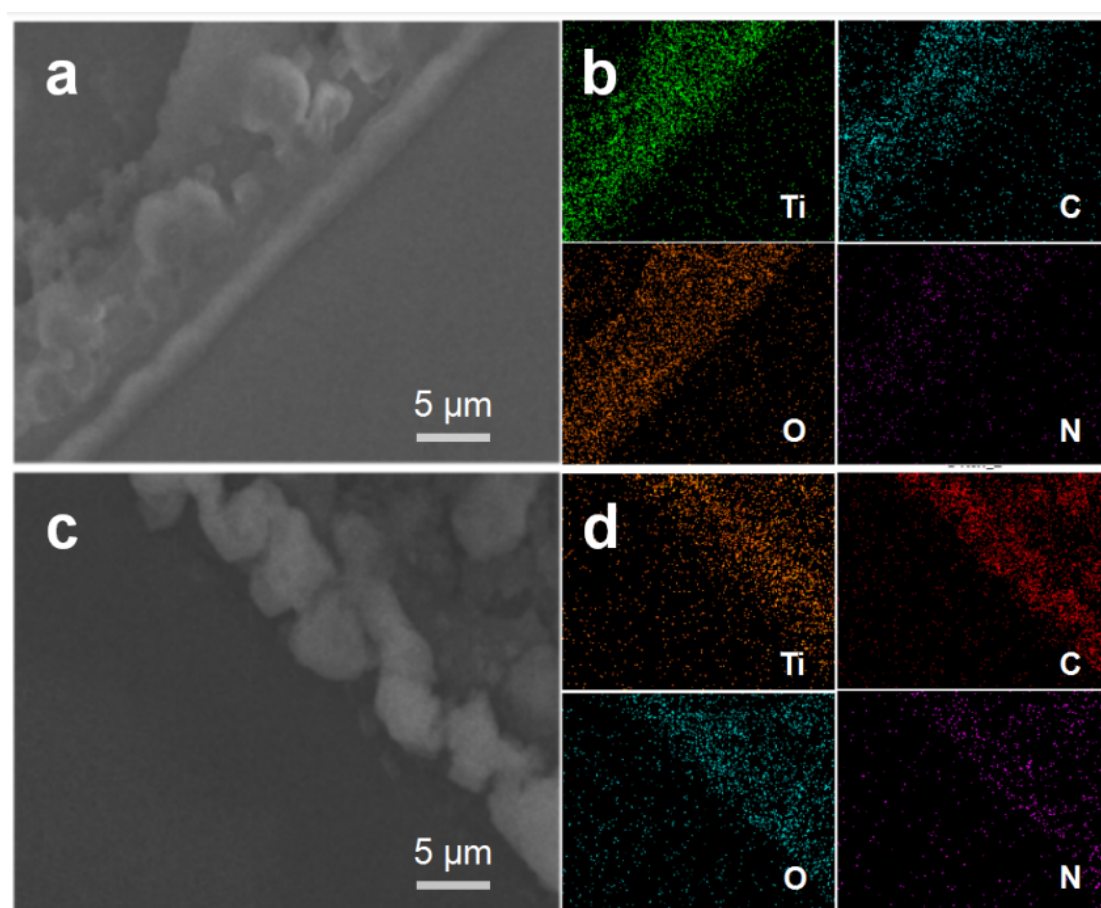


Figure S2. (a) SEM image of agglomerated MQDs and (b) the corresponding EDS elemental mapping images. (c) SEM image of agglomerated N-MQDs and (d) the corresponding EDS elemental mapping images.

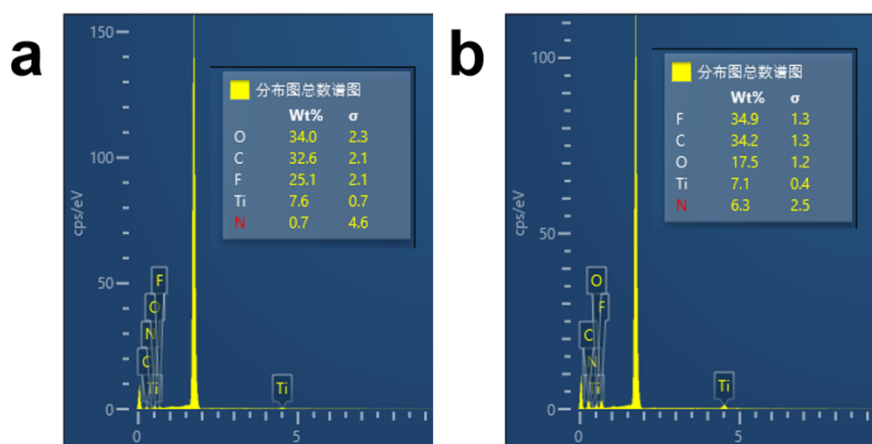


Figure S3. SEM-EDS of MQDs and N-MQDs.

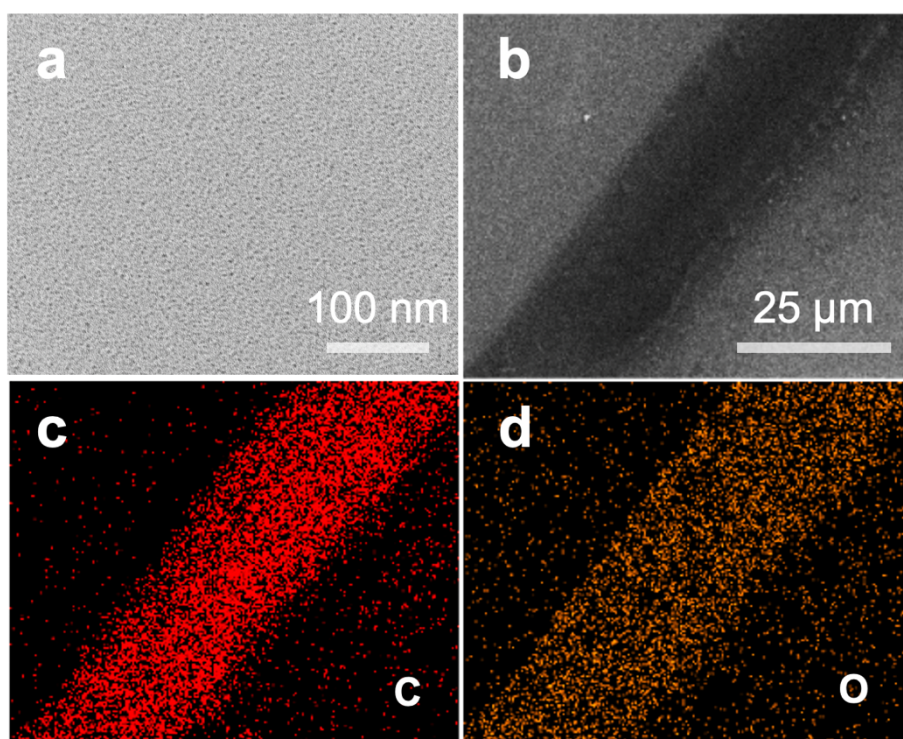


Figure S4. (a) TEM image of CQDs and (b) SEM image of agglomerated CQDs and (c), (d) the corresponding EDS elemental mapping images.

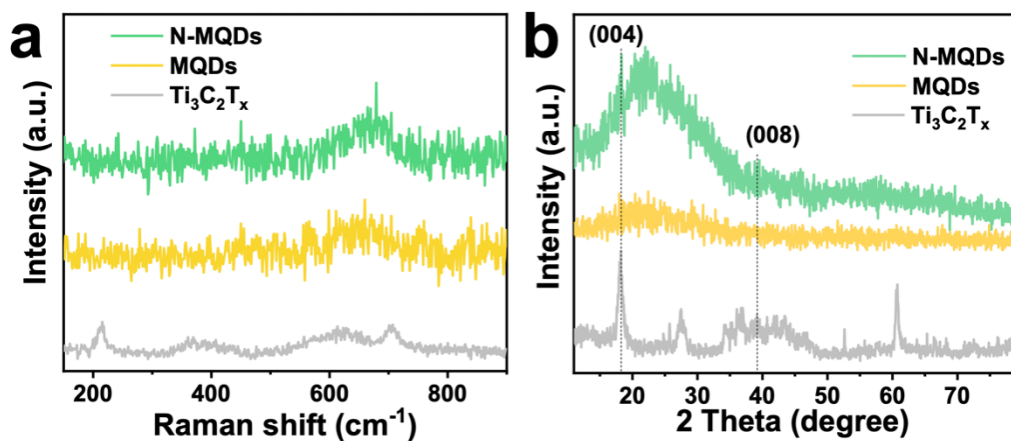


Figure S5. (a) Raman spectra and (b) XRD patterns of Ti₃C₂T_x, MQDs and N-MQDs.

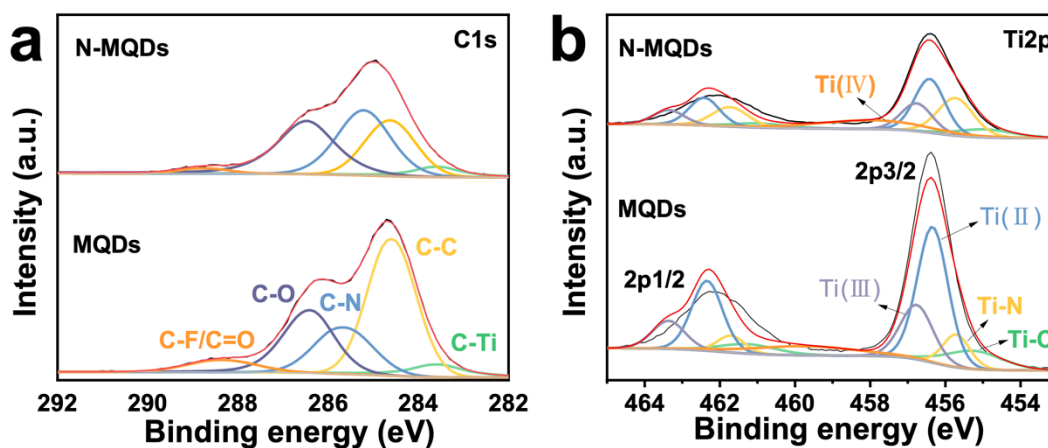


Figure S6. High-resolution XPS spectra of (a) C 1s and (b) Ti 2p for MQDs and N-MQDs.

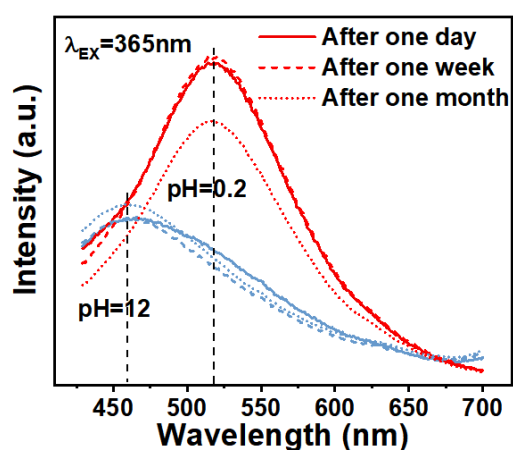


Figure S7. PL stability of N-MQDs at the pH of 12 (blue) and 0.2 (red) under the excitation of 365 nm.

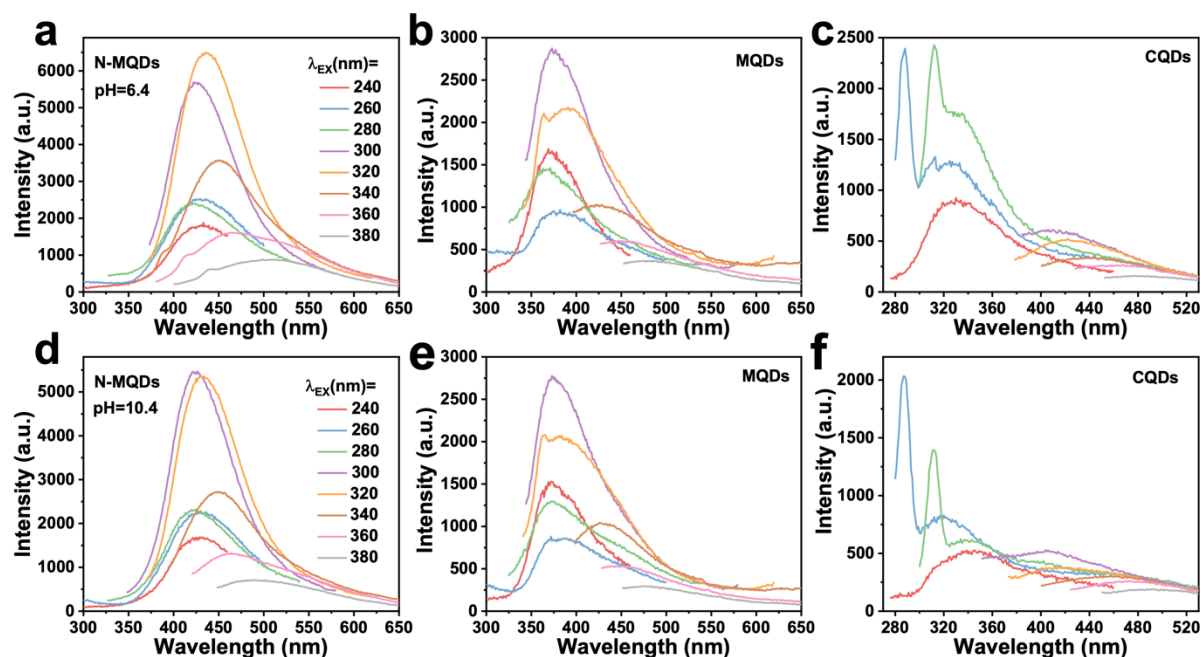


Figure S8. PL spectra of different QDs under varying excitation wavelength at the pH of (a-c) 6.4 and (d-f) 10.4.

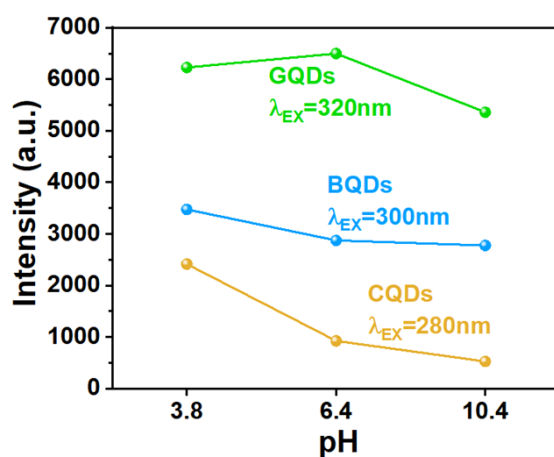


Figure S9. The plot of the PL emission peak intensity against the pH for the three different QDs.

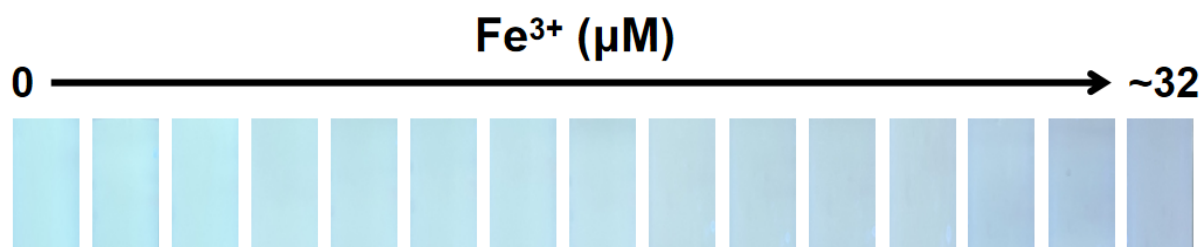


Figure S10. Fluorescent photos of the N-MQDs in the presence of Fe^{3+} with different concentration at the pH of 3.8. The photos were taken under a 365 nm UV lamp.

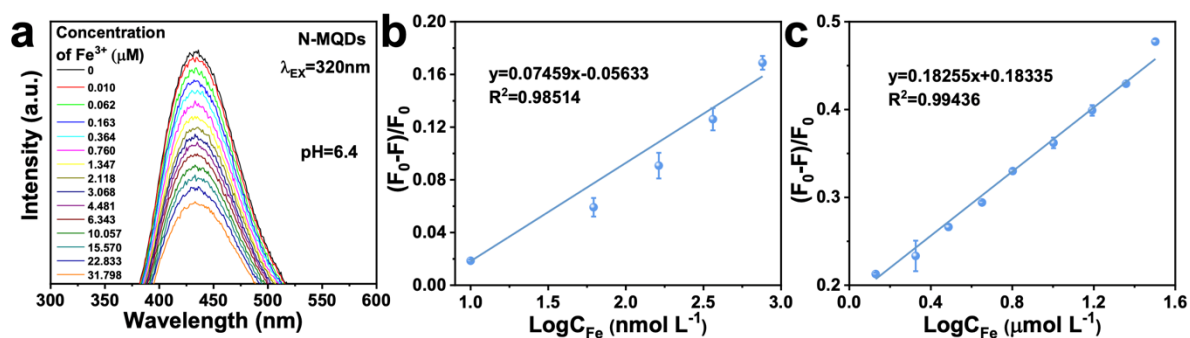


Figure S11. (a) PL spectra of the N-MQDs in the presence of Fe^{3+} (0–32 μM) at the pH of 6.4. Linear relationship between F_0-F/F_0 and the logarithm of Fe^{3+} concentration in the (b) nanomolar range (<1 μM) and (c) micromolar range (<32 μM).

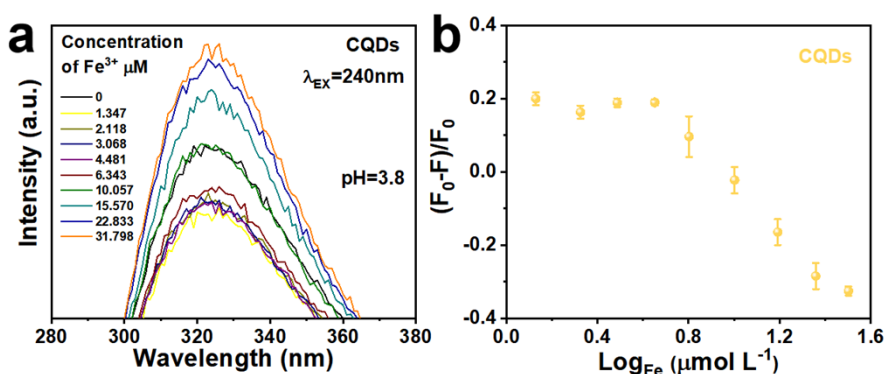


Figure S12. (a) CQDs in the presence of Fe^{3+} at the pH of 3.8. (b) Linear relationship between F_0-F/F_0 and the concentration of Fe^{3+} in the micromolar range (<32 μM) for CQDs.

Table S1. Performance comparison of different fluorescent nanoprobe for the detection of Fe³⁺.

Probe	Detection method	LOD (nM)	Linear range (μM)	Reference
T ₃ C ₂ T _x QDs	Fluorometry	310	5–1000	1
N-T ₃ C ₂ T _x QDs	Fluorometry	2000	2–5000	2
Graphene QDs	Fluorometry	20	0–1	3
Graphene QDs	Fluorometry	230	0.1–20	4
S-Graphene QDs	Fluorometry	4.2	0.01–0.7	5
Carbon QDs	Fluorometry	63.4	0.37–1500	6
Graphene QDs	Fluorometry	450	0–60	7
P-Carbon QDs	Fluorometry	9.5	0.02–3	8
N, S-Carbon QDs	Fluorometry	190	0.3–70	9
B, N, S-Carbon QDs	Fluorometry; colorimetry	90	0.3–546	10
N-Carbon QDs	Fluorometry	520	1–250	11
N, P-Carbon QDs	Fluorometry	330	1–150	12
B, N-Carbon QDs	Fluorometry	80	2–160	13
N-Graphene QDs	Fluorometry	17	0–0.125	14
N-Carbon QDs	Fluorometry	100	2–20	15
F-Carbon QDs	Fluorometry	10	1–100	16
N-Graphene QDs	Fluorometry	63	0–80	17
Carbon QDs	Fluorometry	355.4	0–1000	18
N-Carbon QDs	Fluorometry	500	0–70	19
N-Carbon QDs	Fluorometry	73	0.1–4	20
Carbon QDs	Fluorometry	770	0–300	21

N, P-Carbon QDs	Fluorometry	447	0–250	22
Carbon QDs	Fluorometry	5230	0–500	23
N-MQDs	Fluorometry	2	0–32	This work

Table S2. Results for the determination of Fe^{3+} in tap water, pond water and mineral water samples.

Samples	Concentration from ICP (nM)	Concentration from the present method (nM)	Recovery (%)
Tap water	54	60 ± 7	111
Pond water	84	89 ± 9	106
Mineral water	63	70 ± 8	111

Table S3. Results for the determination of Fe^{3+} with spiked concentration in tap water.

Spiked concentration of Fe^{3+} (nM)	Total Fe^{3+} concentration after addition (nM)	Found (nM)	Recovery (%)
20	74	84 ± 4	115
50	103	109 ± 8	106
100	153	161 ± 6	106
200	252	256 ± 3	102
500	549	554 ± 4	101

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