

## Electronic Supplementary Information

### **Tunable Third-Order Nonlinear Optical Effect via Modifying $\text{Ti}_4(\text{embonate})_6$ Cage-Based Ionic Pairs**

*Guang-Hui Chen, Yan-Ping He \*, Zi-Rui Wang, Qiao-Hong Li, Zhi-Zhou Ma, and Jian Zhang*

\*

G.-H. Chen, Y.-P. He, Z.-R. Wang, Q.-H. Li, Z.-Z. Ma, J. Zhang

State Key Laboratory of Structural Chemistry, Fujian Institute of Research on the Structure of Matter, Chinese Academy of Sciences, Fuzhou, Fujian 350002, P. R. China.

E-mail: hyp041@163.com; zhj@fjirsm.ac.cn

Y.-P. He

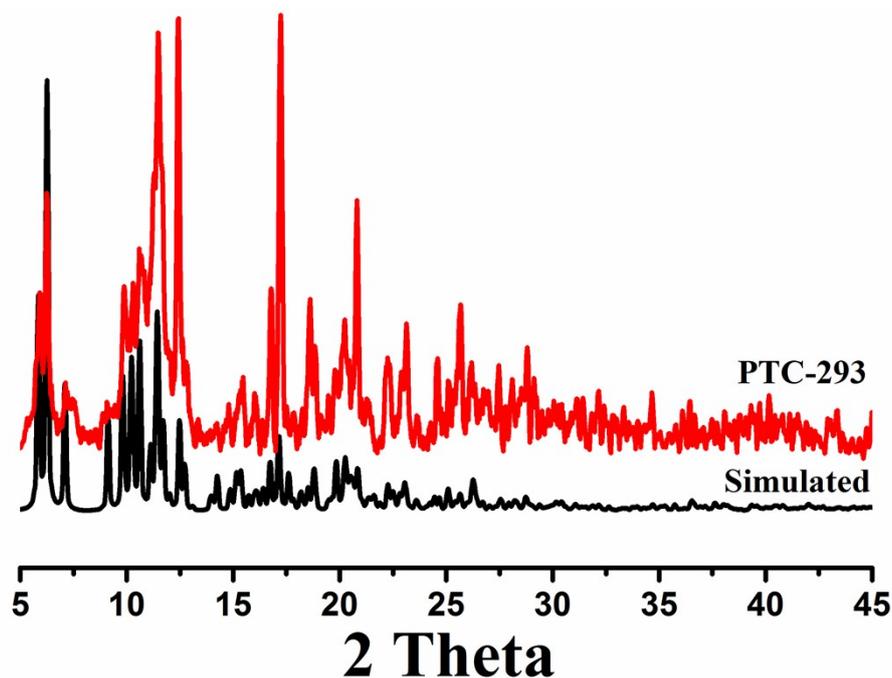
University of Chinese Academy of Sciences, Beijing 100049, P. R. China.

E-mail: hyp041@163.com

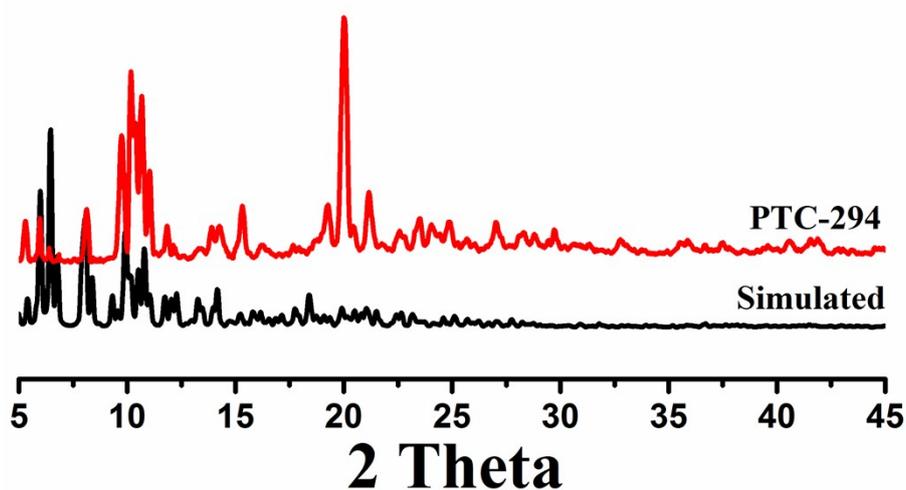
## Contents

<b>1. The PXRD Analyses for PTCs .....</b>	<b>3</b>
<b>2. Detailed Structure Information for PTCs.....</b>	<b>5</b>
<b>3. The SEM-Mapping Spectra of PTCs.....</b>	<b>17</b>
<b>4. The EDS Spectra of PTCs .....</b>	<b>18</b>
<b>5. The UV-Vis Spectra of PTCs.....</b>	<b>20</b>
<b>6. The FT-IR Spectra of PTCs .....</b>	<b>22</b>
<b>7. The TGA curve of PTCs .....</b>	<b>23</b>
<b>8. The XPS Spectra of PTCs.....</b>	<b>25</b>
<b>9. The NLO Property and NICS(0) Method Studies of PTCs.....</b>	<b>28</b>
<b>10. Crystallography Data.....</b>	<b>32</b>
<b>11. Reference.....</b>	<b>33</b>

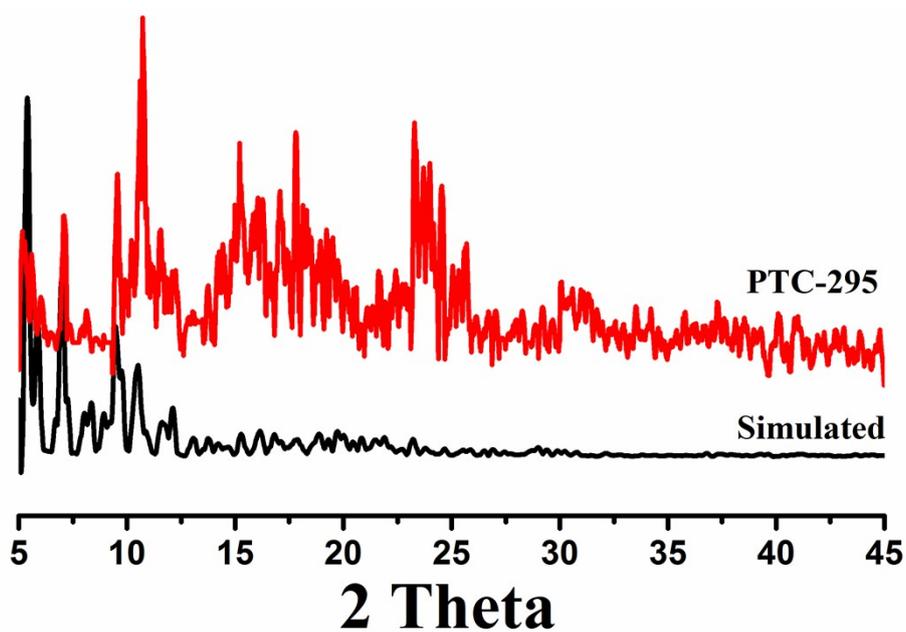
## 1. The PXRD Analyses for PTCs



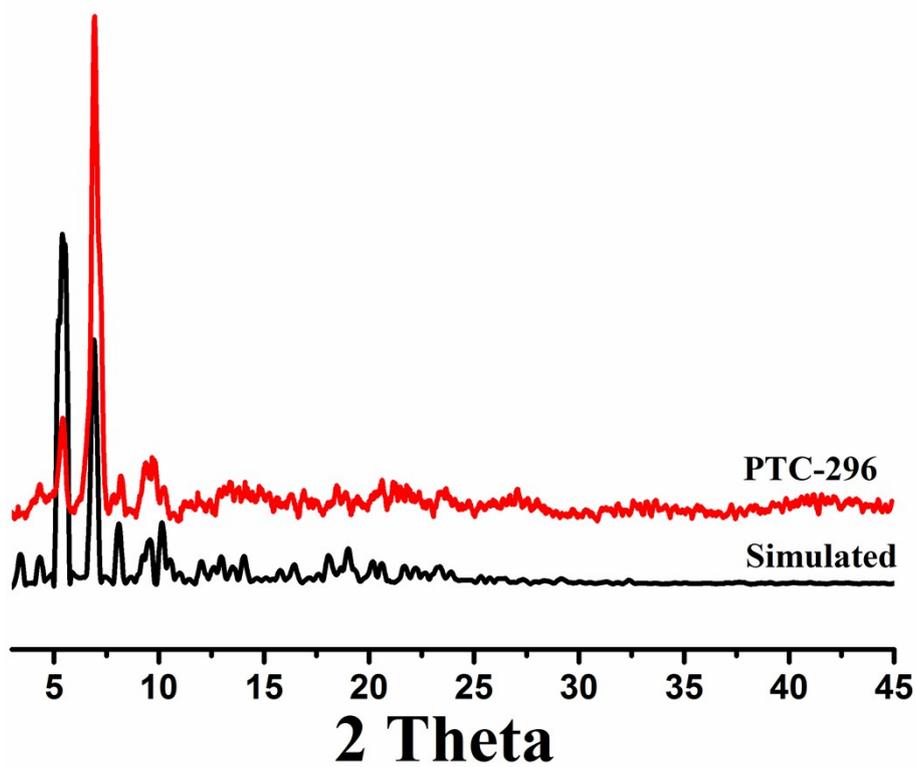
**Figure S1.** The PXRD patterns of simulated from the single-crystal data of **PTC-293** (black), as-synthesized **PTC-293** (red).



**Figure S2.** The PXRD patterns of simulated from the single-crystal data of **PTC-294** (black), as-synthesized **PTC-294** (red).

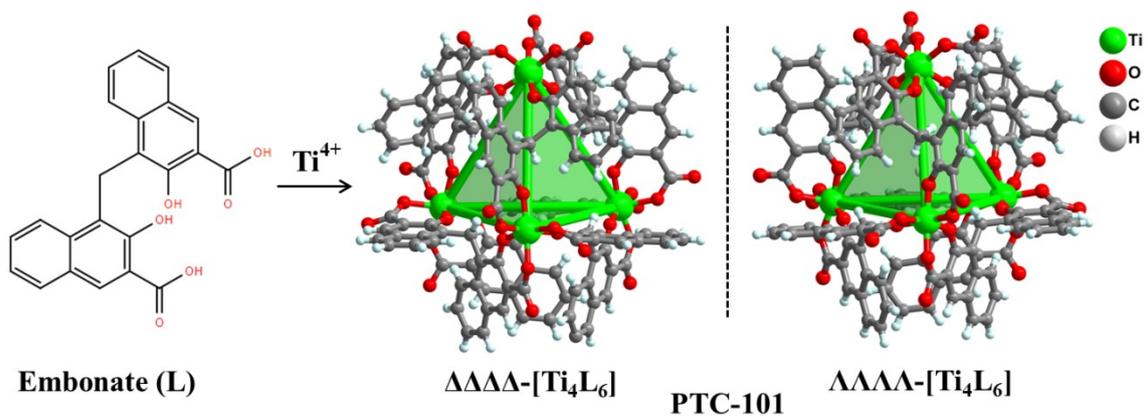


**Figure S3.** The PXRD patterns of simulated from the single-crystal data of **PTC-295** (black), as-synthesized **PTC-295** (red).

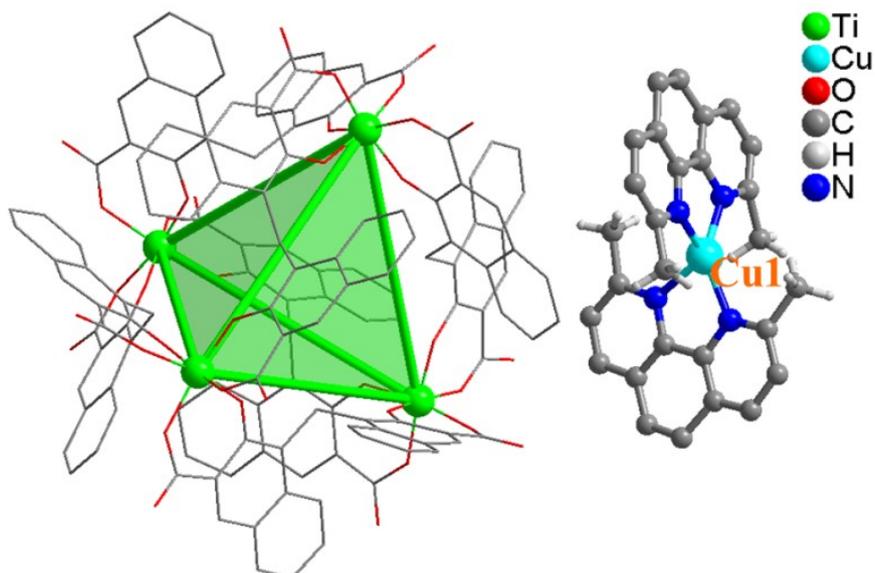


**Figure S4.** The PXRD patterns of simulated from the single-crystal data of **PTC-296** (black), as-synthesized **PTC-296** (red).

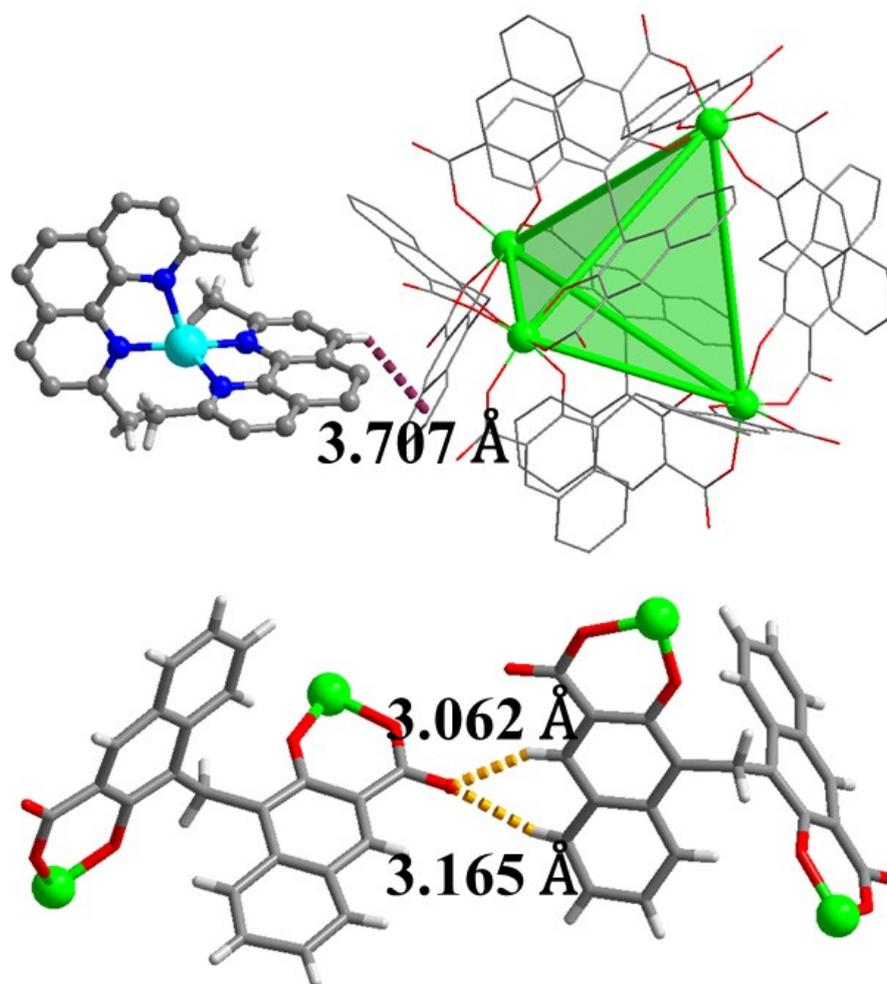
## 2. Detailed Structure Information for PTCs.



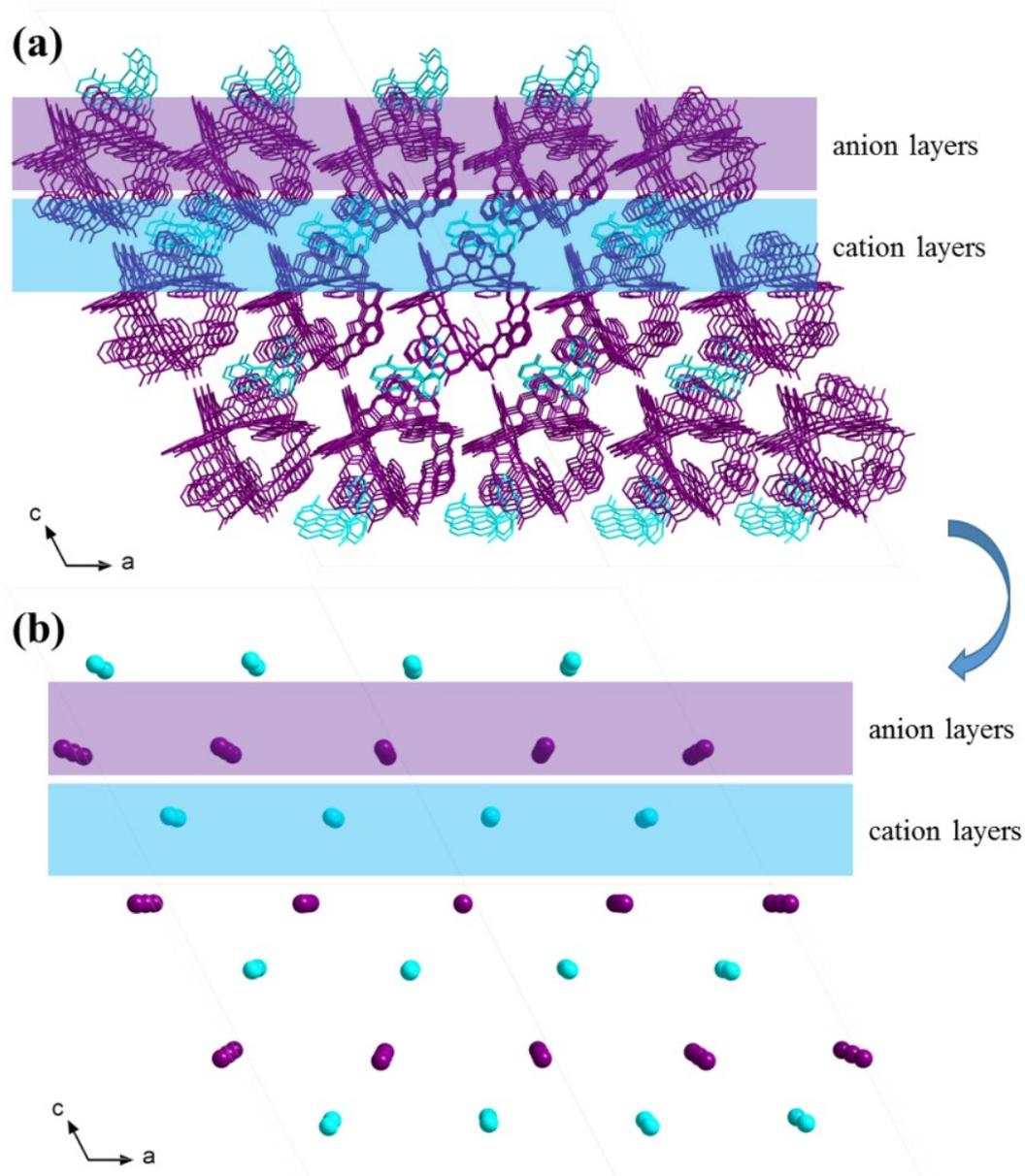
**Figure S5.** Illustration of the structures of the embonate (L) ligand, the  $\Delta\Delta\Delta\Delta\text{-}[\text{Ti}_4\text{L}_6]$  and  $\Delta\Delta\Delta\Delta\text{-}[\text{Ti}_4\text{L}_6]$  isomers in PTC-101.<sup>1</sup>



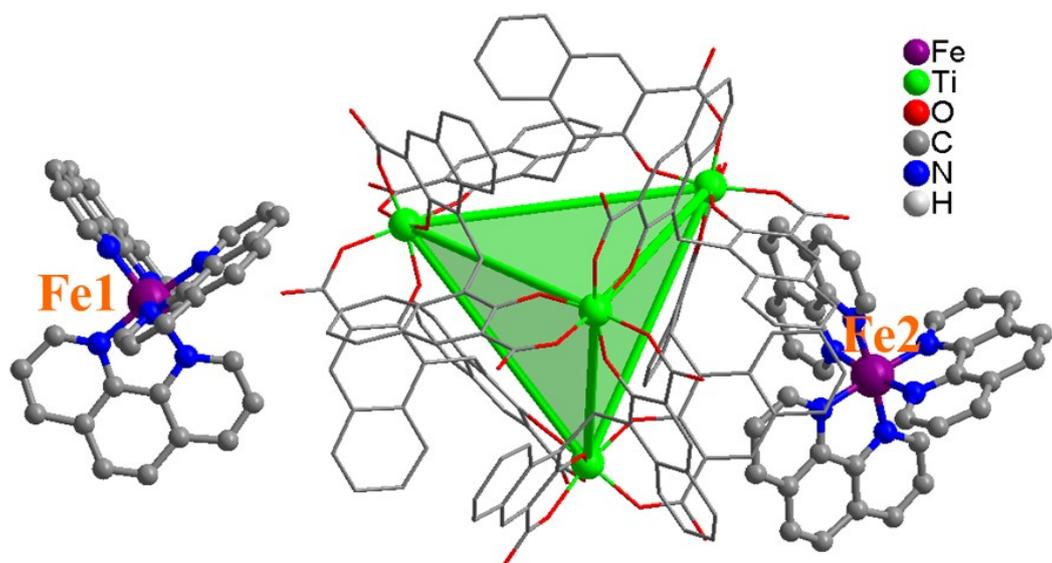
**Figure S6.** The asymmetric unit of **PTC-293**, showing one  $\text{Ti}_4\text{L}_6$  cage, one  $[\text{Cu}(\text{Me}_2\text{phen})_2]^+$  cation and some free  $(\text{NH}_4)^+$  cations which could not be located because of highly disorder.



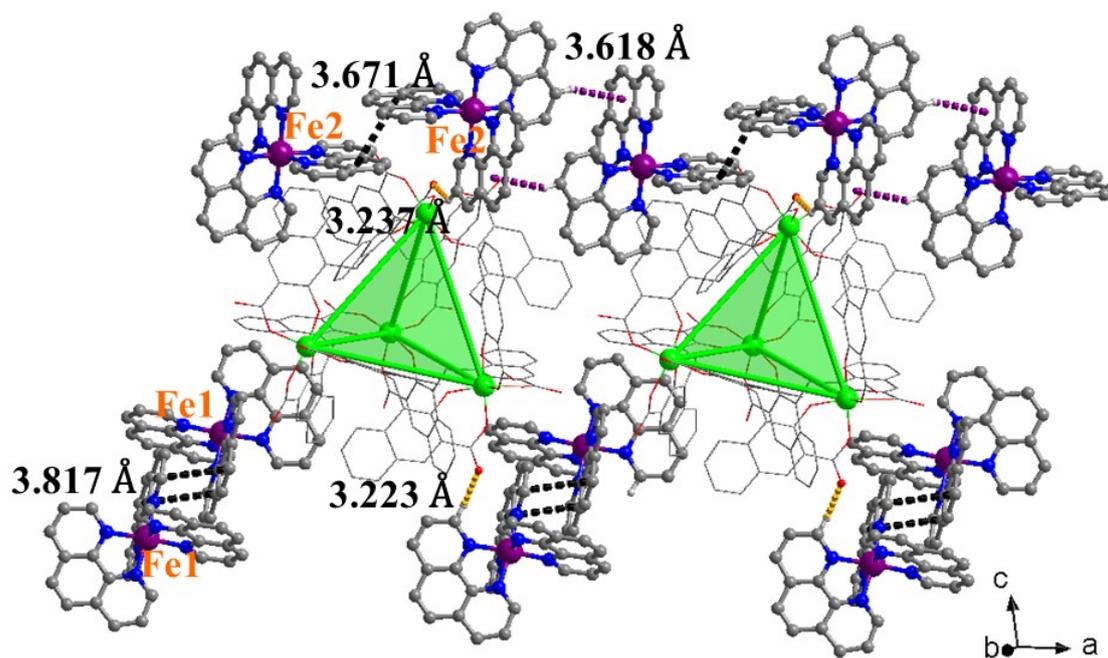
**Figure S7.** The highlighting of the C–H··· $\pi$  and the C–H···O weak interactions between adjacent Ti<sub>4</sub>L<sub>6</sub> cage and [Cu(Me<sub>2</sub>phen)<sub>2</sub>]<sup>+</sup> cation or two Ti<sub>4</sub>L<sub>6</sub> cages, respectively.



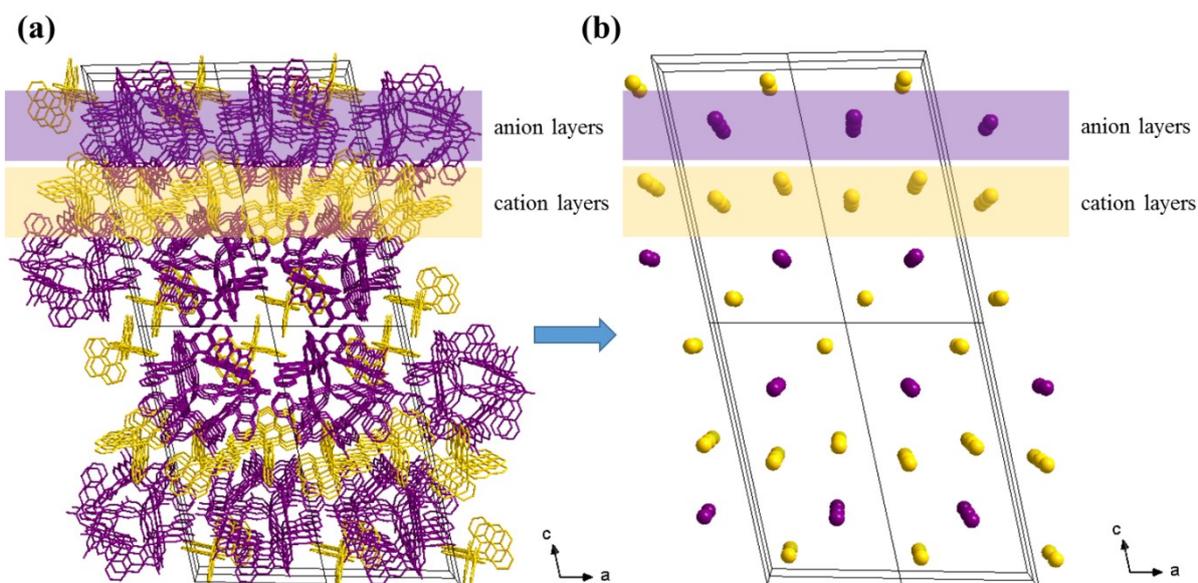
**Figure S8.** (a) Molecular packing diagram of PTC-293 co-crystals projected along the crystallographic b-axis. (b) Schematic illustration of  $\text{Ti}_4(\text{embonate})_6$ -based ionic pairs methodology with conventional parallel-type cation–anion assembly with alternating cation and anion layers.



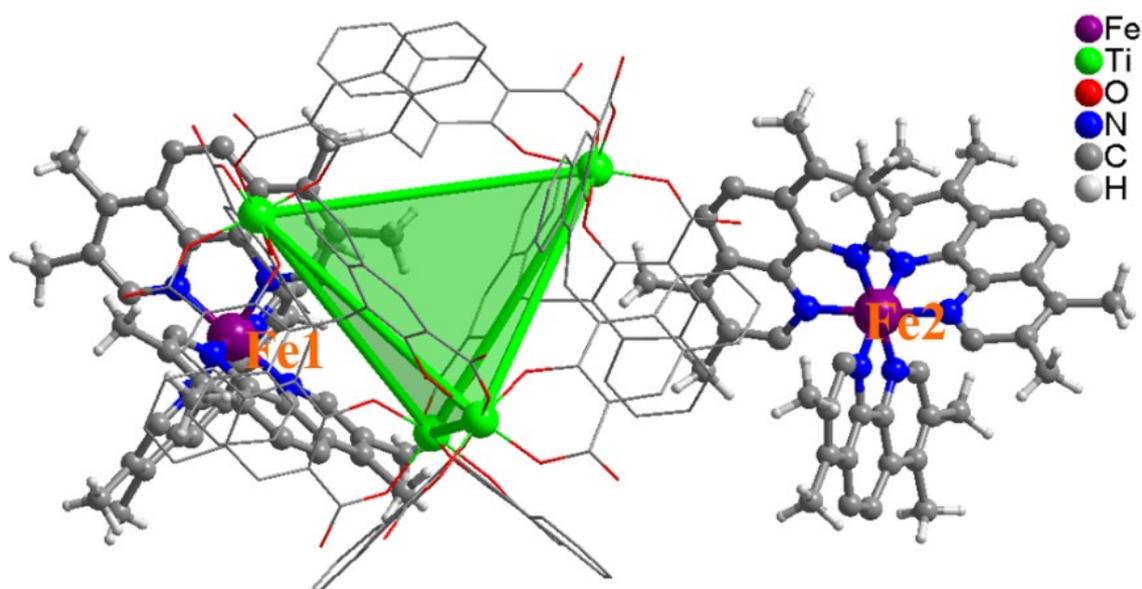
**Figure S9.** The asymmetric unit of PTC-294, showing one  $\text{Ti}_4\text{L}_6$  cage, two  $[\text{Fe}(\text{phen})_3]^{2+}$  cations, and some free  $(\text{NH}_4)^+$  cations.



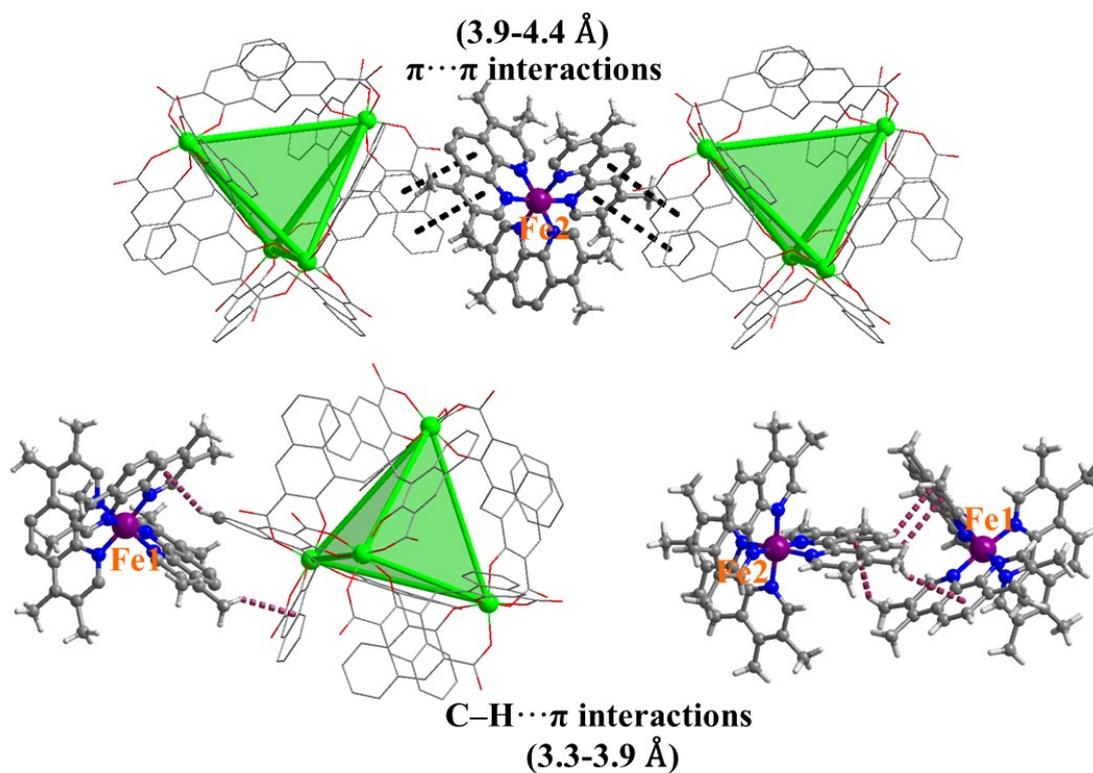
**Figure S10.** The highlighting of the  $\pi \cdots \pi$  stacking and  $\text{C}-\text{H} \cdots \pi$  interactions between adjacent  $[\text{Fe}(\text{phen})_3]^{2+}$  cations as well as the  $\text{C}-\text{H} \cdots \text{O}$  hydrogen bonds between adjacent  $\text{Ti}_4\text{L}_6$  cage and  $[\text{Fe}(\text{phen})_3]^{2+}$  cation, respectively.



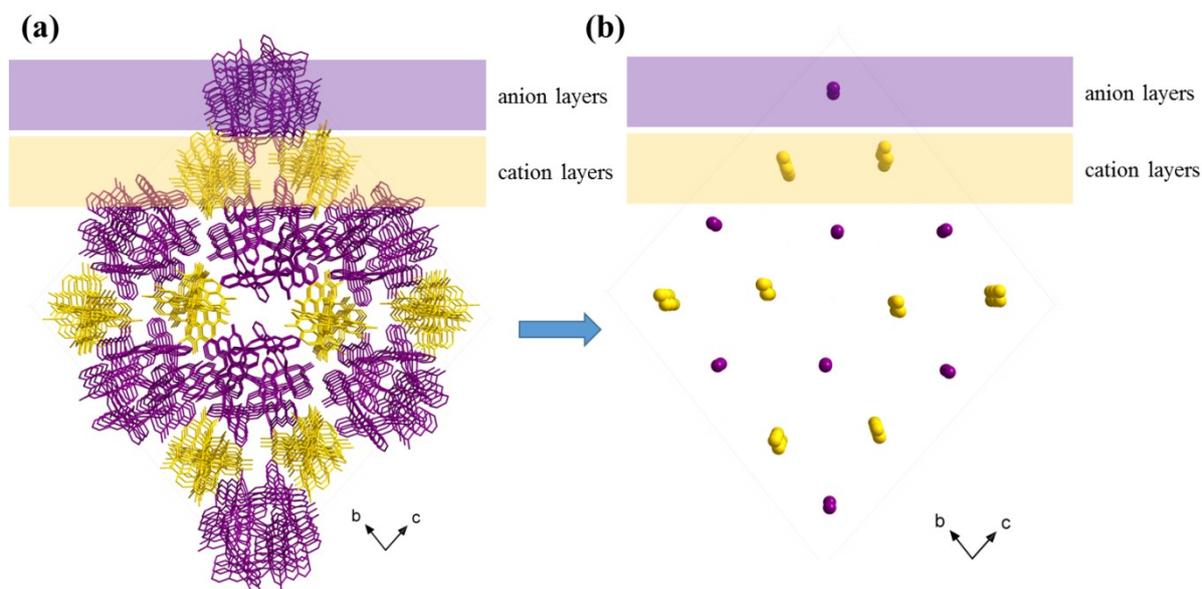
**Figure S11.** (a) Molecular packing diagram of PTC-294 co-crystals projected along the crystallographic b-axis. (b) Schematic illustration of  $Ti_4(embonate)_6$ -based ionic pairs methodology with conventional parallel-type cation–anion assembly with alternating cation and anion layers.



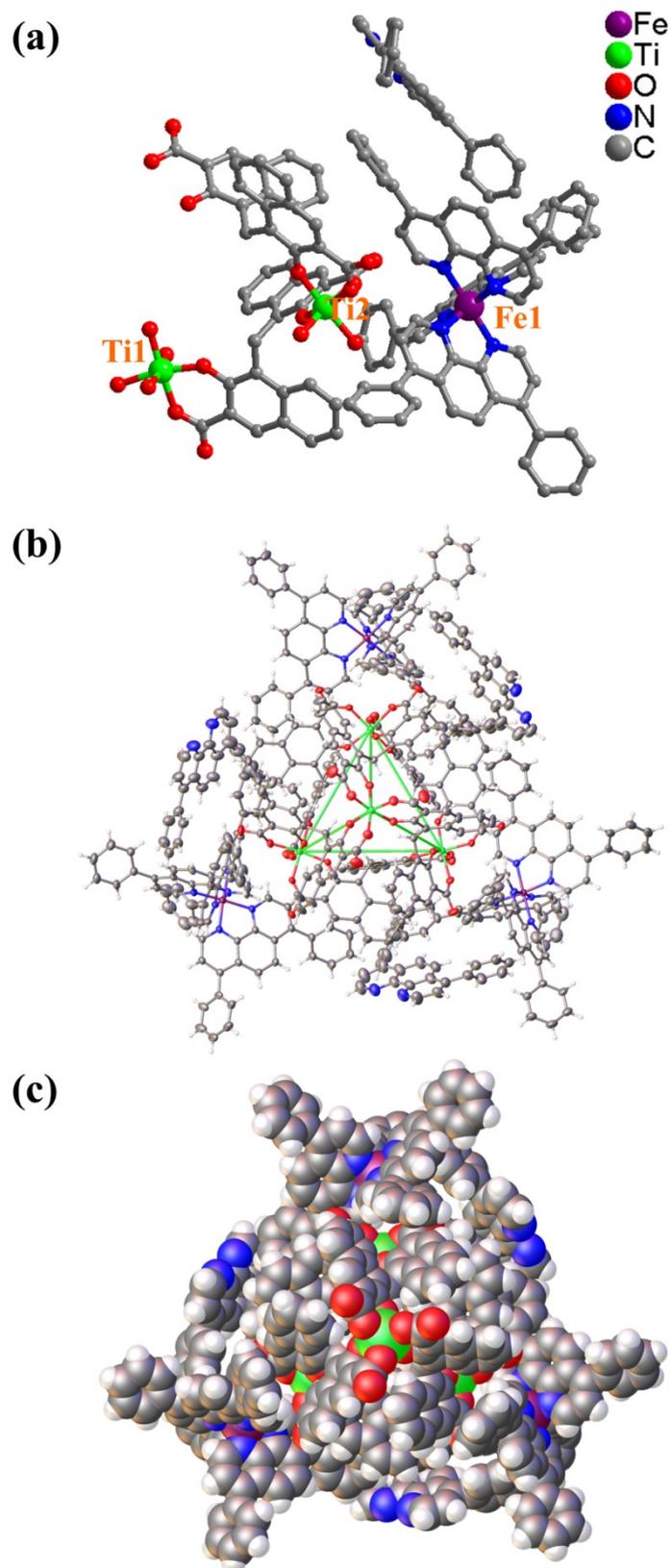
**Figure S12.** The asymmetric unit of PTC-295, showing one  $Ti_4L_6$  cage, two  $[Fe(Me_4phen)_3]^{3+}$  cations and some free  $(NH_4)^+$  cations.



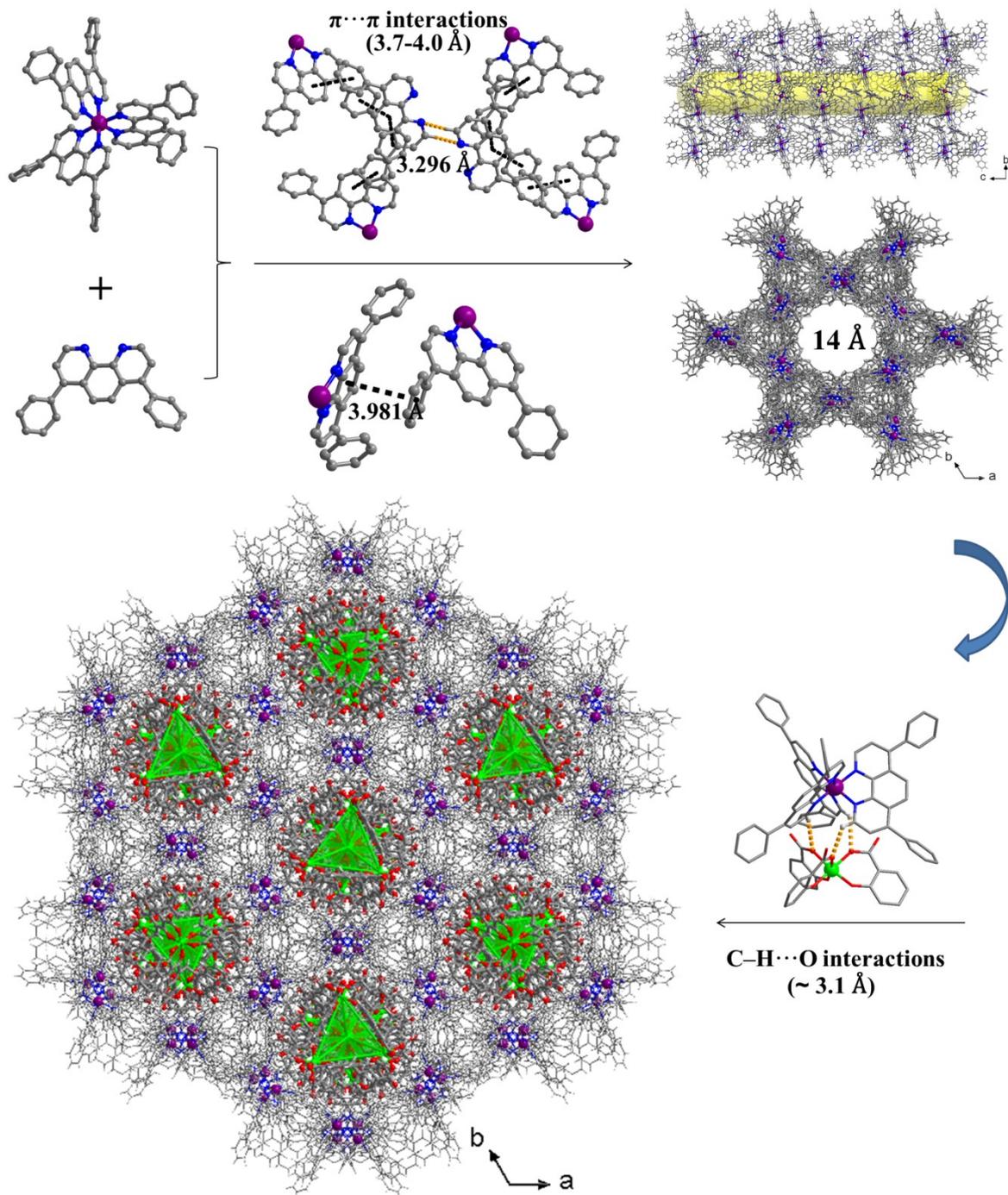
**Figure S13.** The highlighting of the  $\pi \cdots \pi$  stacking and C-H $\cdots\pi$  interactions between adjacent  $\text{Ti}_4\text{L}_6$  cage and  $[\text{Fe}(\text{Me}_4\text{phen})_3]^{3+}$  cation.



**Figure S14.** (a) Molecular packing diagram of PTC-295 co-crystals projected along the crystallographic a-axis. (b) Schematic illustration of  $\text{Ti}_4(\text{embonate})_6$ -based ionic pairs methodology with conventional parallel-type cation–anion assembly with alternating cation and anion layers.



**Figure S15.** (a) The asymmetric unit of **PTC-296**, showing one third of  $\text{Ti}_4\text{L}_6$  cage and one  $[\text{Fe}(\text{Ph}_2\text{phen})_3]^{3+}$  cation, one  $\text{Ph}_2\text{phen}$  ligand and some free  $(\text{NH}_4)^+$  cations. (b) ORTEP diagram for **PTC-296** with the ellipsoids drawn at the 30% probability level. (c) Sphere packing representations of the **PTC-296**.



**Figure S16.** Structural assembly process of PTC-296. Atom color code: green, Ti; cyan, Cu; modena, Fe; red, O; blue, N; gray, C; offwhite, H.

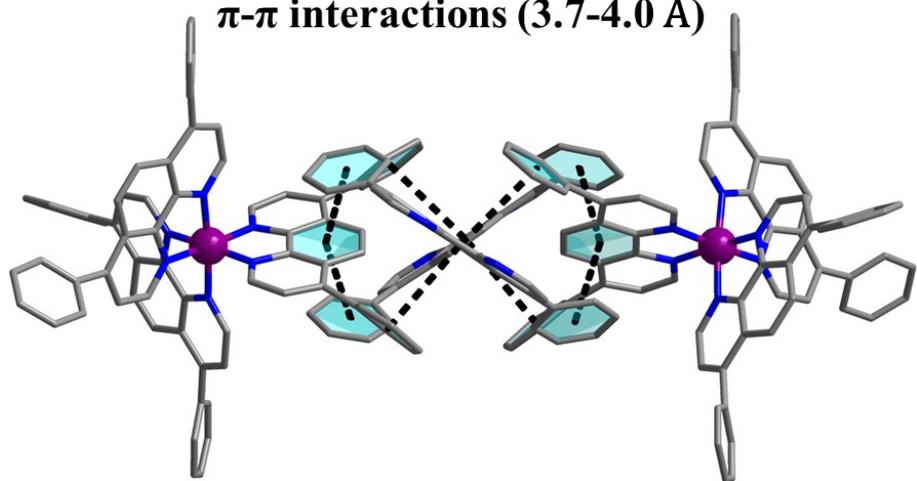
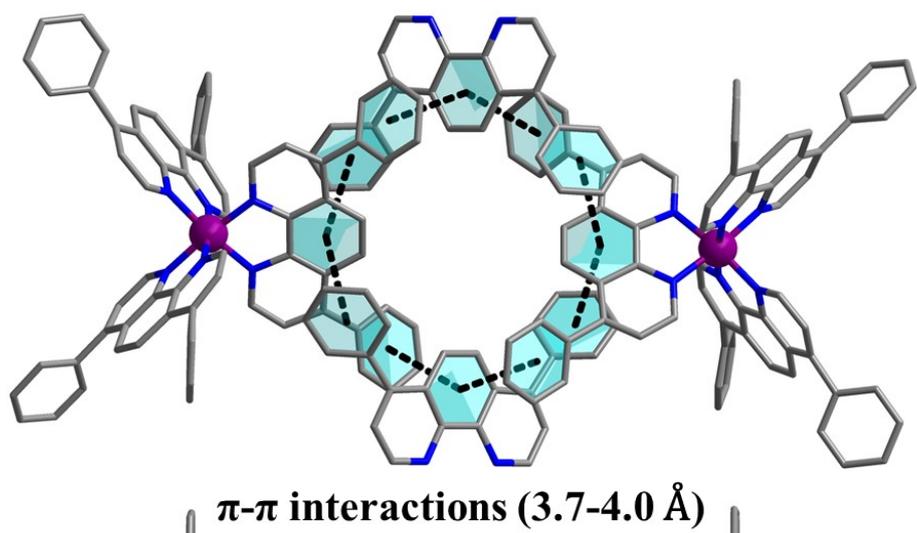


Figure S17. The schematic diagram of  $\pi \cdots \pi$  in PTC-296.

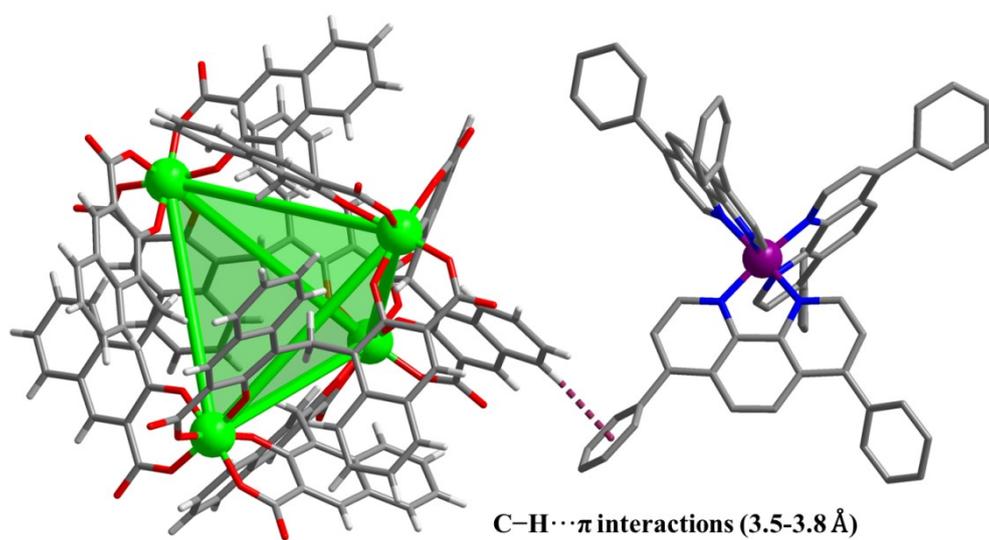
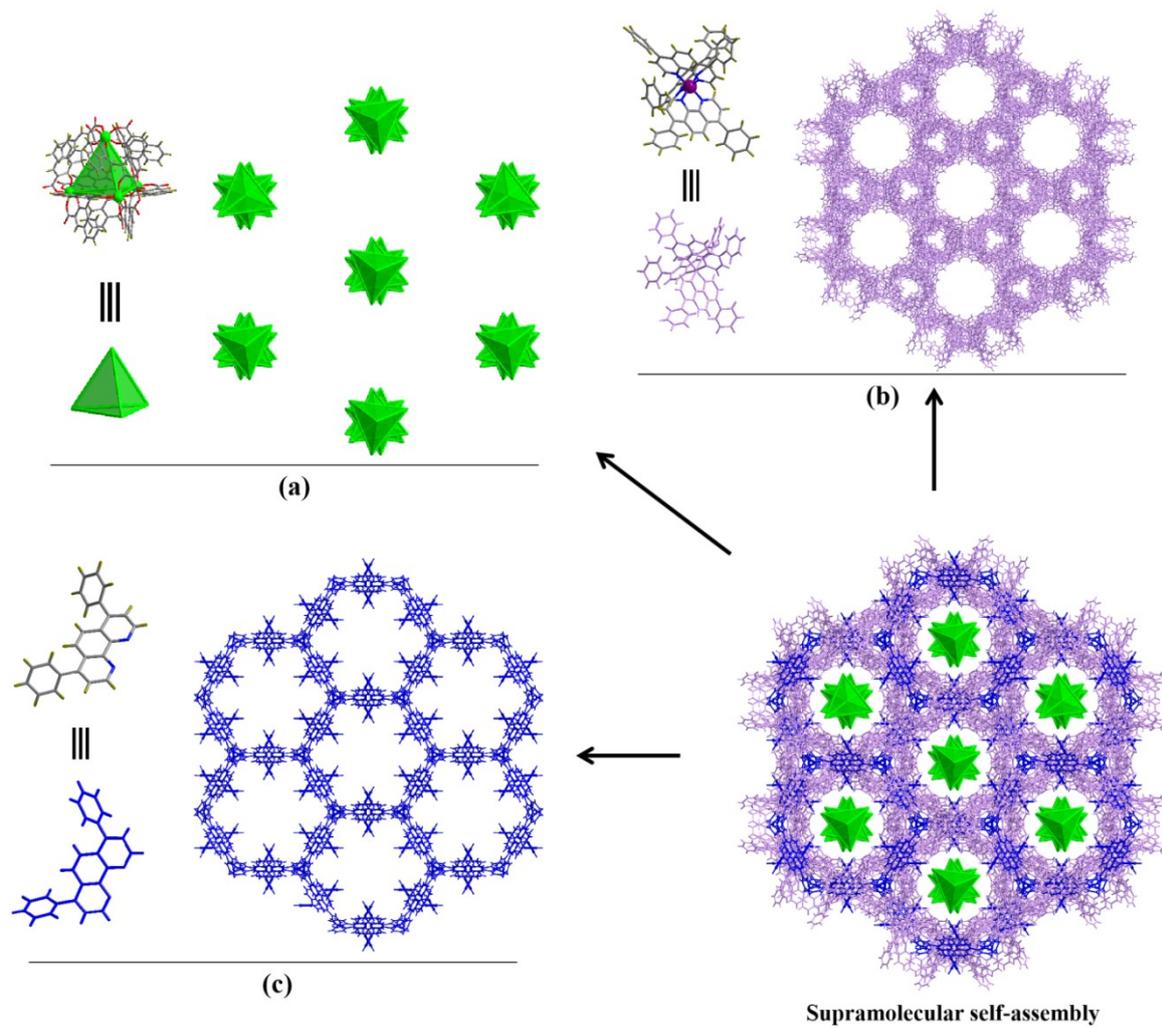
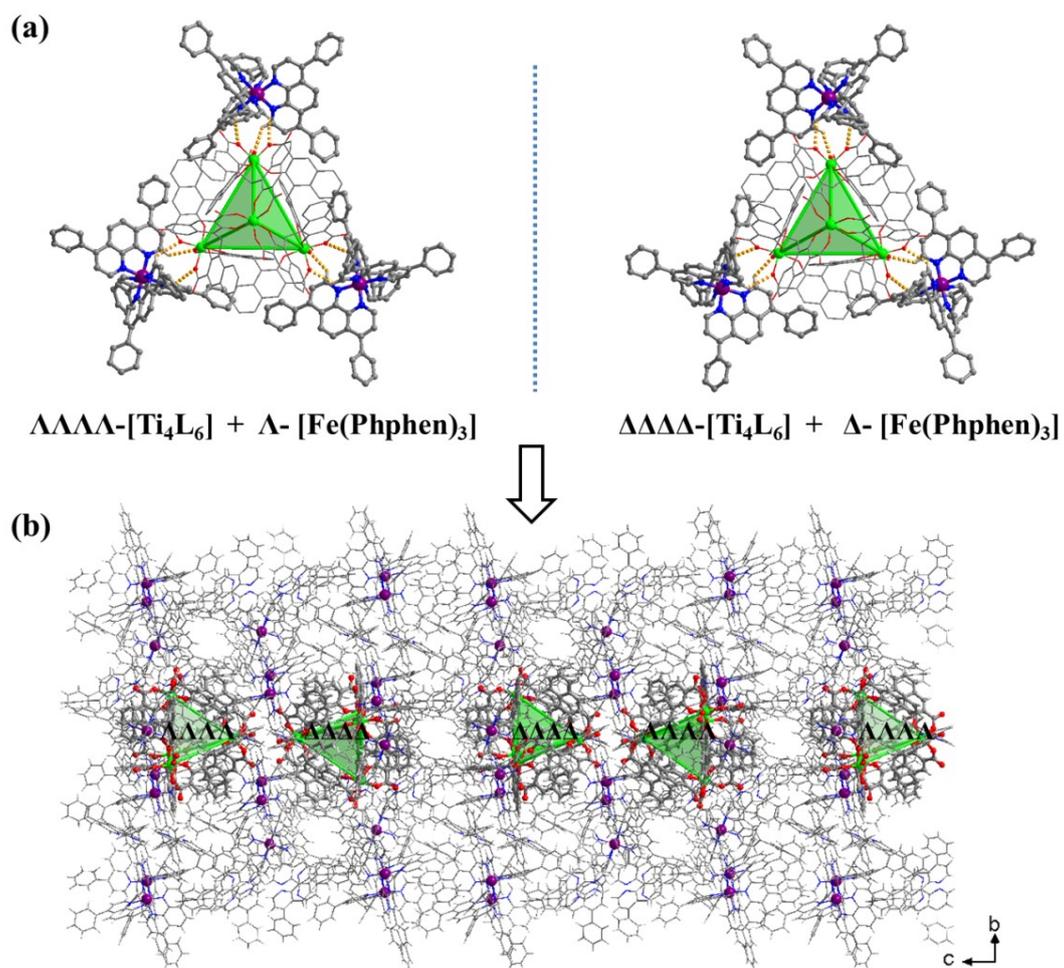


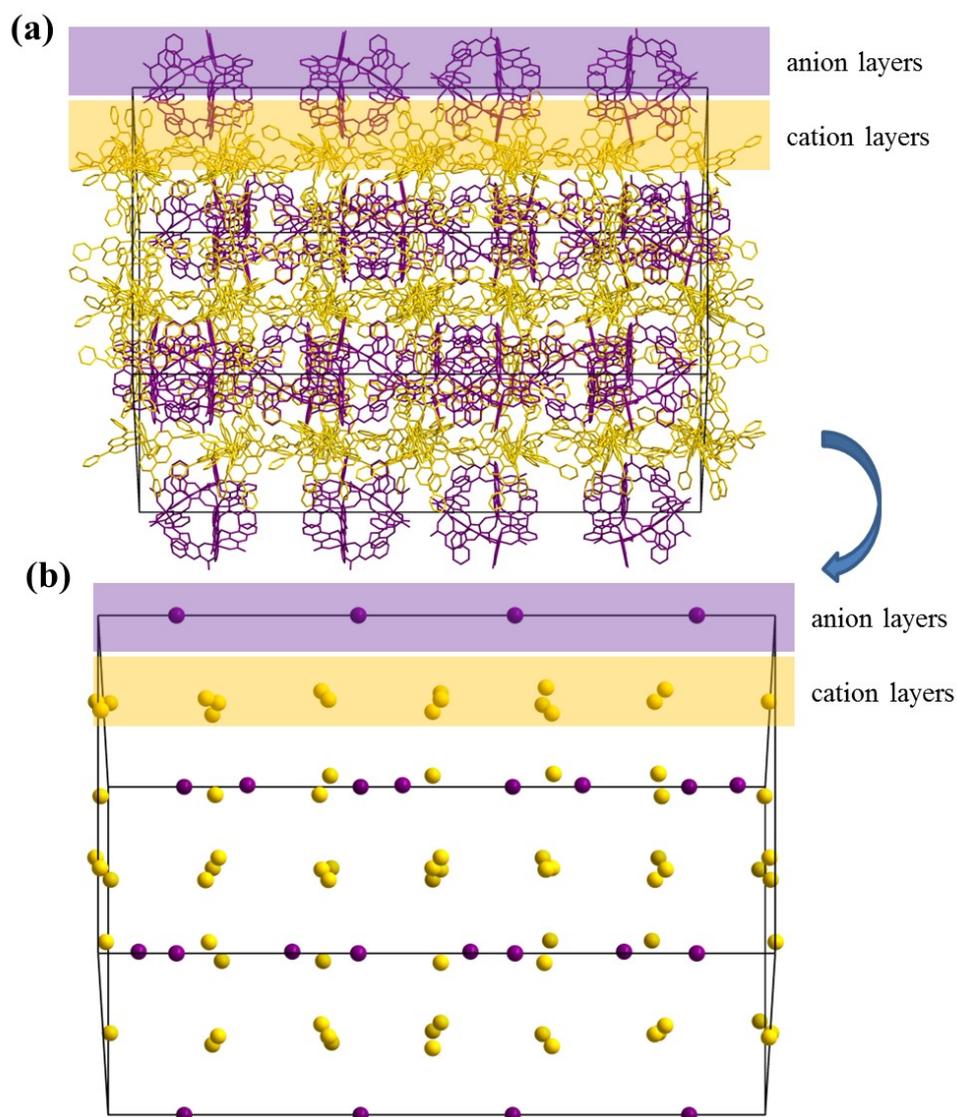
Figure S18. The schematic diagram of C-H $\cdots$  $\pi$  in PTC-296.



**Figure S19.** Illustration the intergration of  $Ti_4L_6$  cage,  $Ph_2phen$  ligand and  $[Fe(Ph_2phen)_3]^{3+}$  unit into a 3D supramolecular architecture of **PTC-296**.

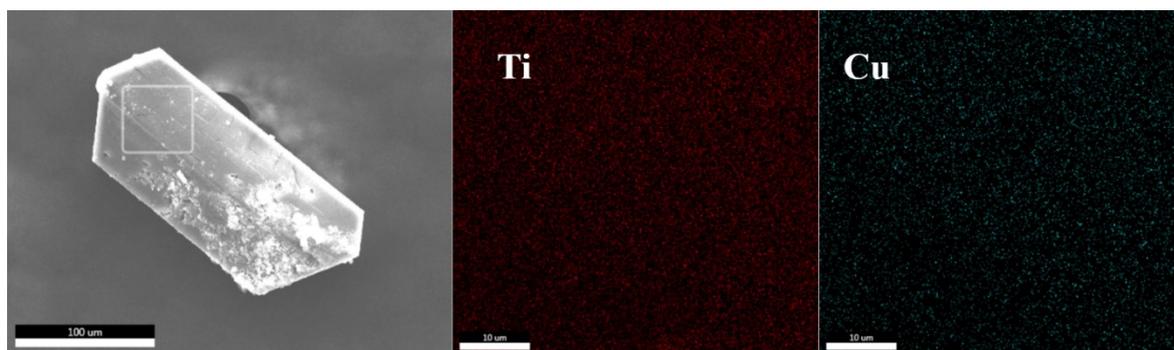


**Figure S20.** (a) Matching between the calixarene-like oxygen vertices of the Ti<sub>4</sub>L<sub>6</sub> cage and the [Fe(Ph<sub>2</sub>phen)<sub>3</sub>]<sup>3+</sup> unit through C–H···O hydrogen bonds in **PTC-296**; (b) the 3D packing superstructure of **PTC-296** along the a-axis direction.

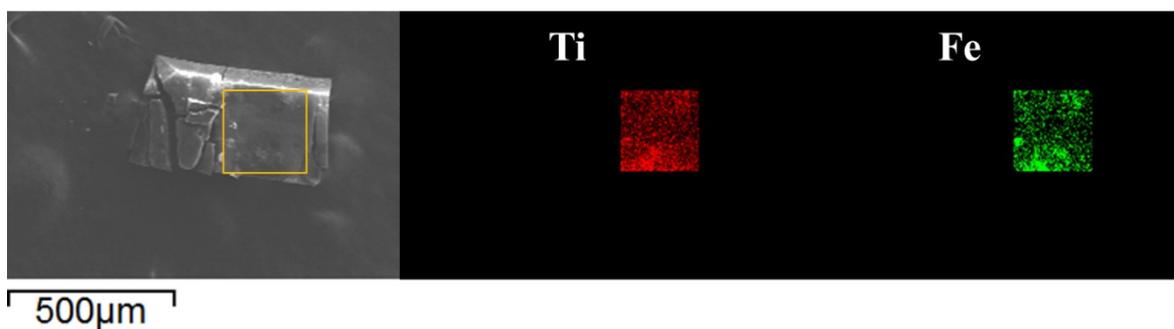


**Figure S21.** (a) Molecular packing diagram of PTC-296 co-crystals projected. (b) Schematic illustration of  $\text{Ti}_4(\text{embonate})_6$ -based ionic pairs methodology with conventional parallel-type cation–anion assembly with alternating cation and anion layers.

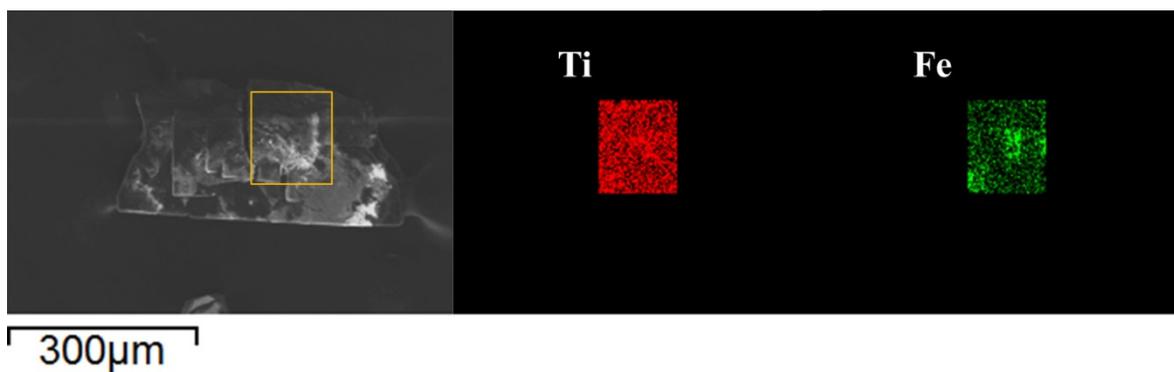
### 3. The SEM-Mapping Spectra of PTCs



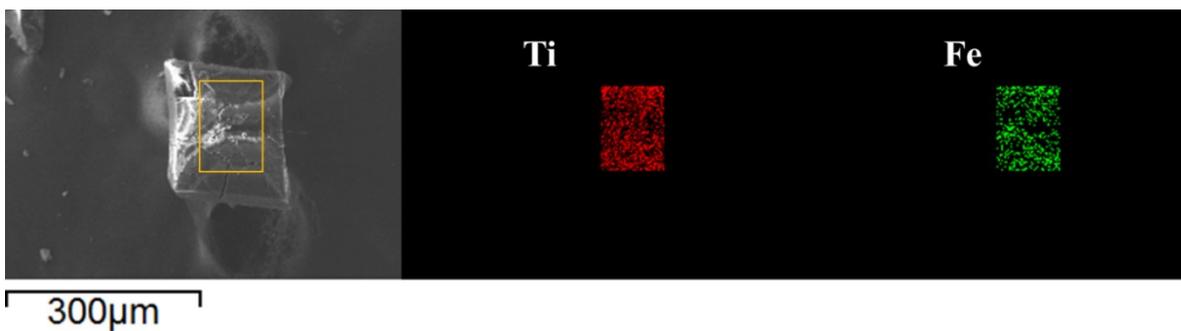
**Figure S22.** The SEM-mapping spectra of PTC-293. From left to right are crystal appearance, titanium, and copper elements.



**Figure S23.** The SEM-mapping spectra of PTC-294. From left to right are crystal appearance, titanium, and iron elements.

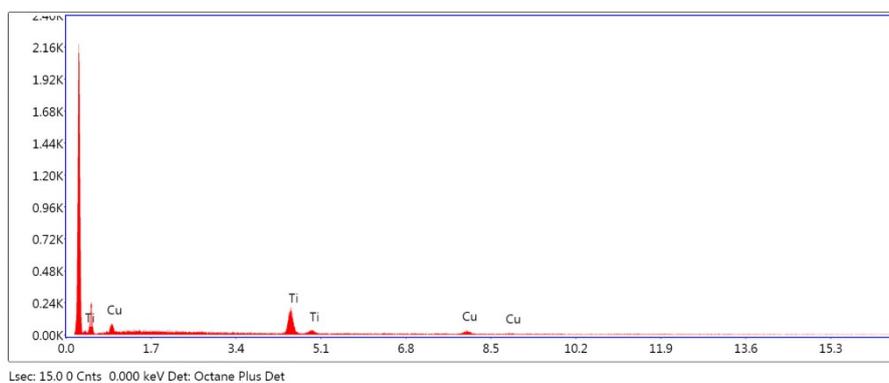


**Figure S24.** The SEM-mapping spectra of PTC-295. From left to right are crystal appearance, titanium, and iron elements.

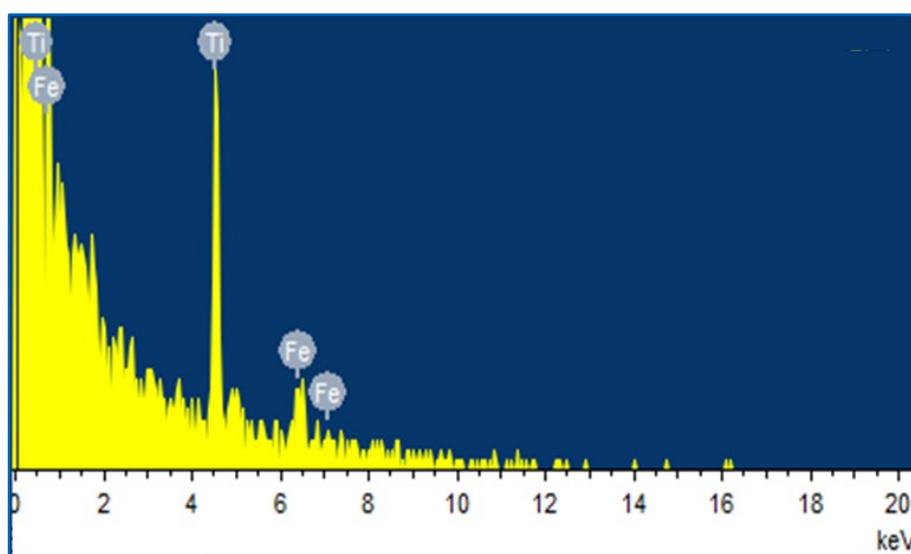


**Figure S25.** The SEM-mapping spectra of PTC-296. From left to right are crystal appearance, titanium, and iron elements.

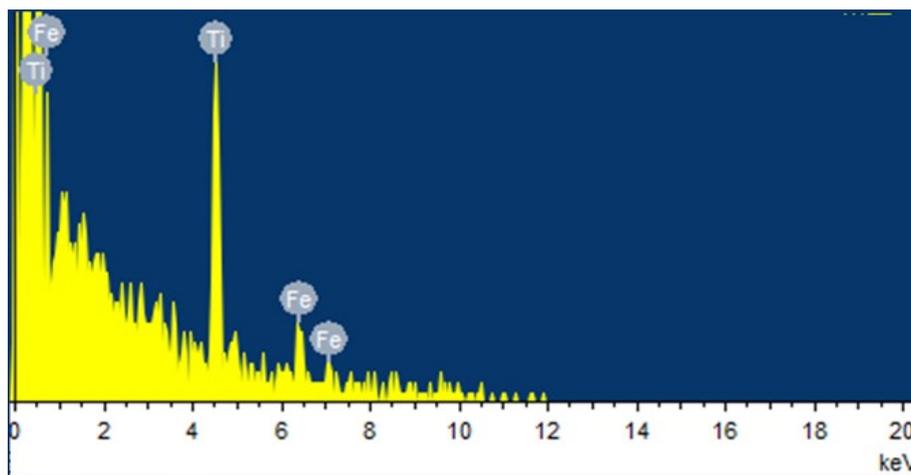
#### 4. The EDS Spectra of PTCs



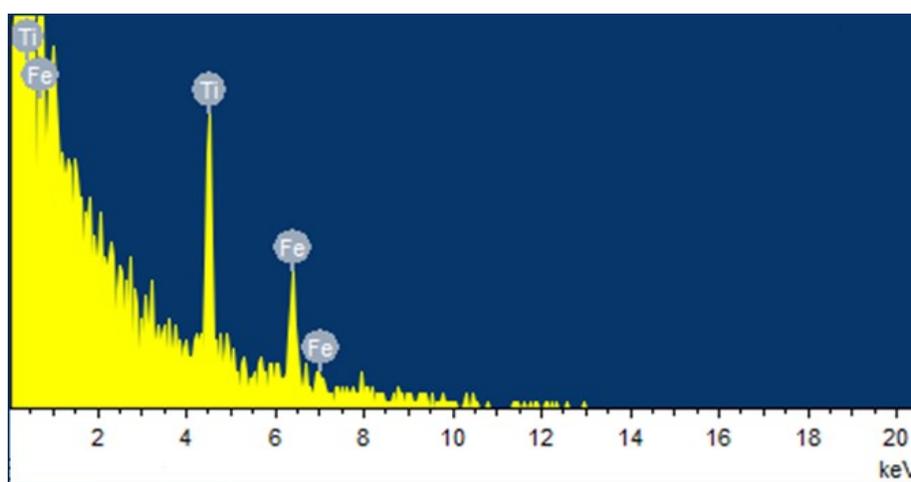
**Figure S26.** The EDS spectrum of compound PTC-293.



**Figure S27.** The EDS spectrum of compound PTC-294.



**Figure S28.** The EDS spectrum of compound **PTC-295**.



**Figure S29.** The EDS spectrum of compound **PTC-296**.

## 5. The UV-Vis Spectra of PTCs

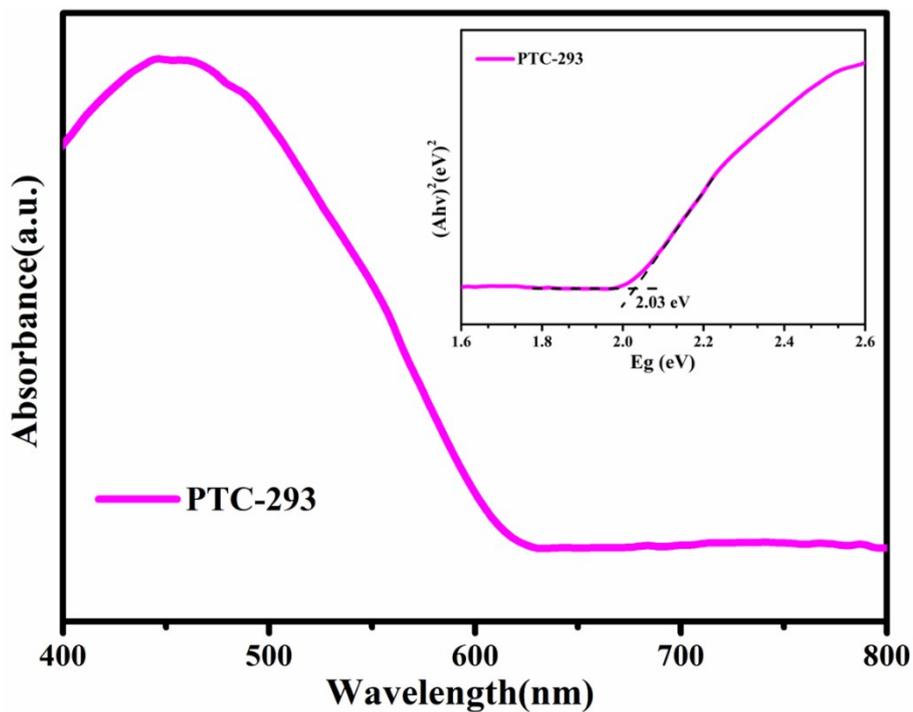


Figure S30. The solid-state UV absorption spectra of PTC-293.

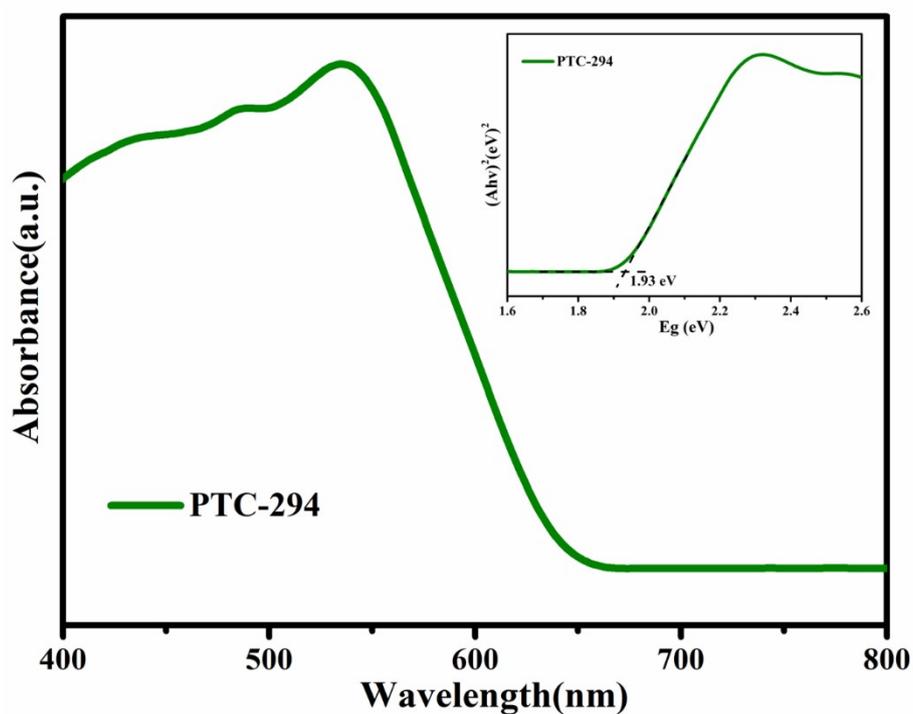


Figure S31. The solid-state UV absorption spectra of PTC-294.

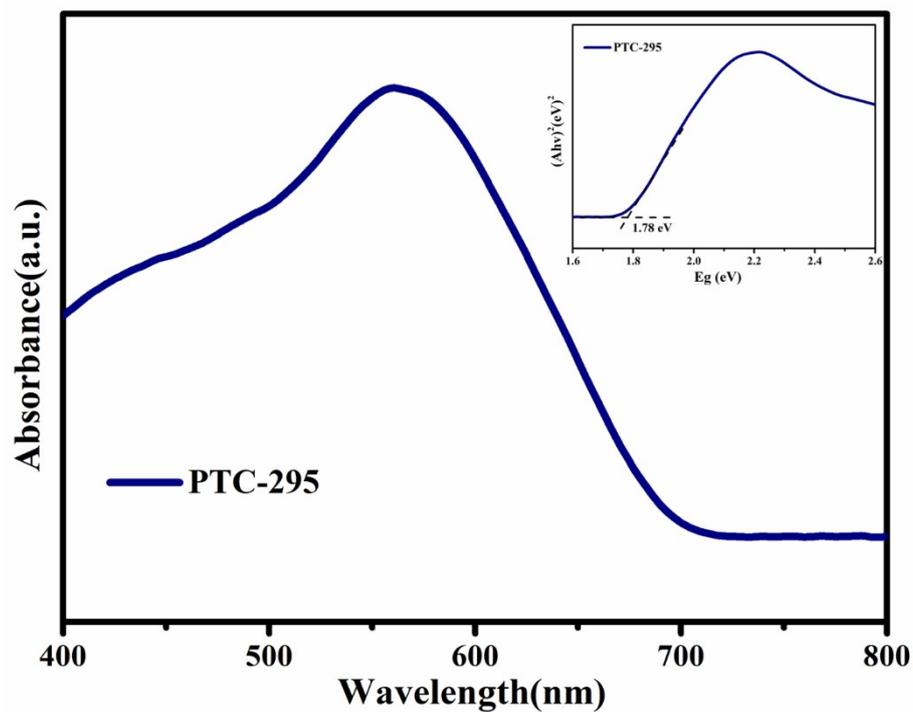


Figure S32. The solid-state UV absorption spectra of PTC-295.

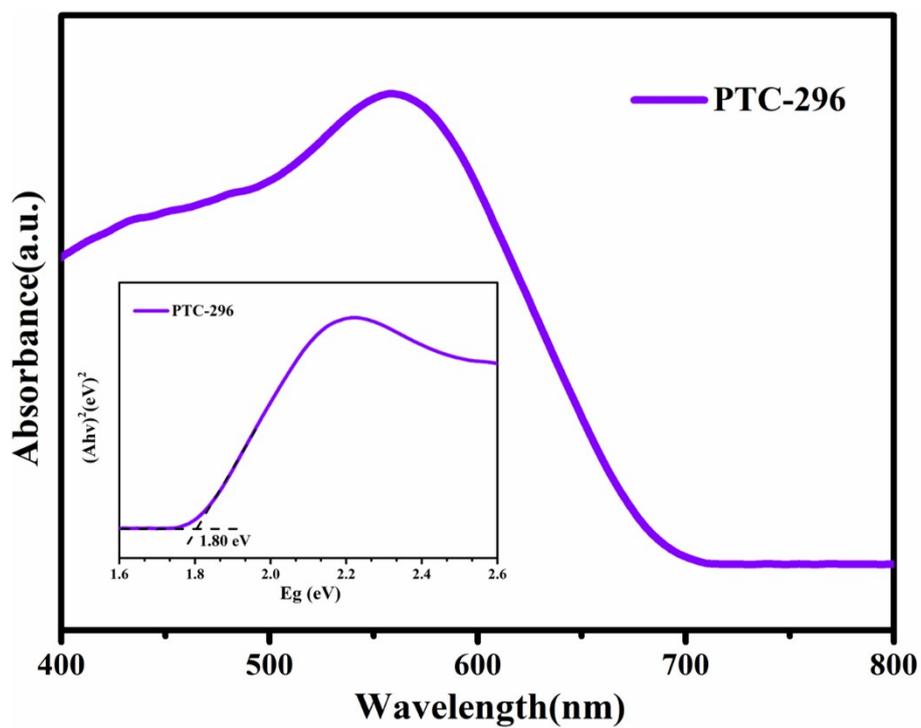


Figure S33. The solid-state UV absorption spectra of PTC-296.

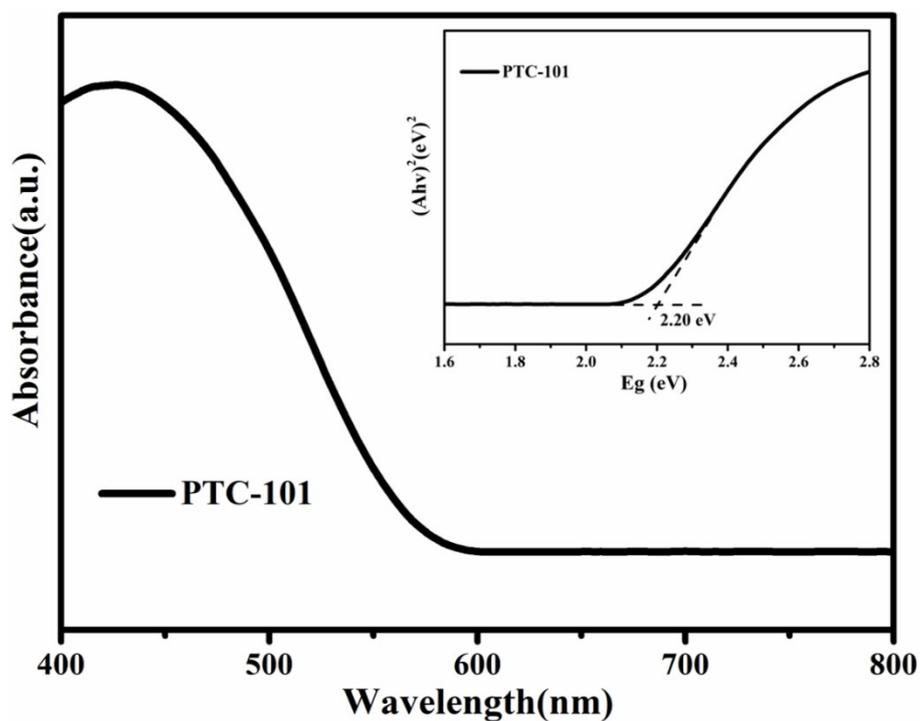


Figure S34. The solid-state UV absorption spectra of PTC-101.

## 6. The FT-IR Spectra of PTCs

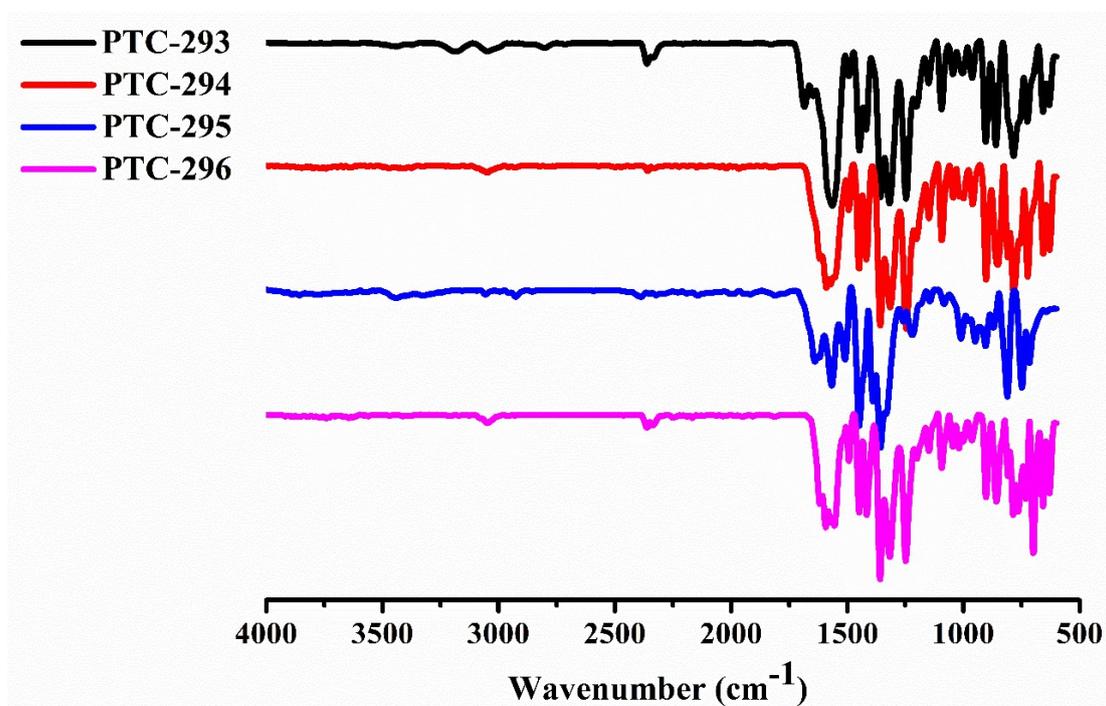


Figure S35. The FT-IR spectra of compounds PTC-293 to PTC-296.

## 7. The TGA curve of PTCs

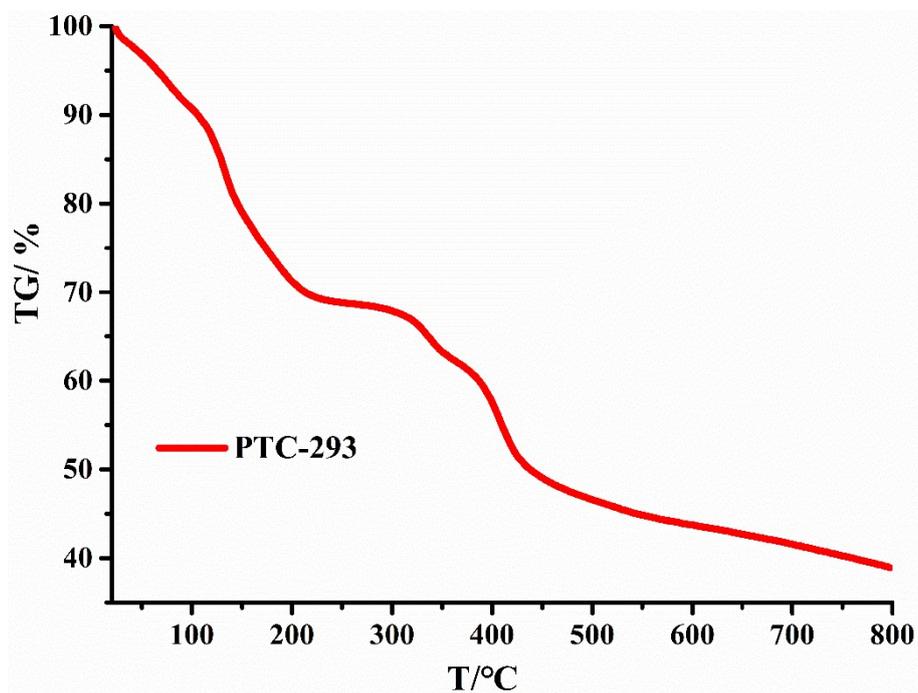


Figure S36. The TGA curve of PTC-293.

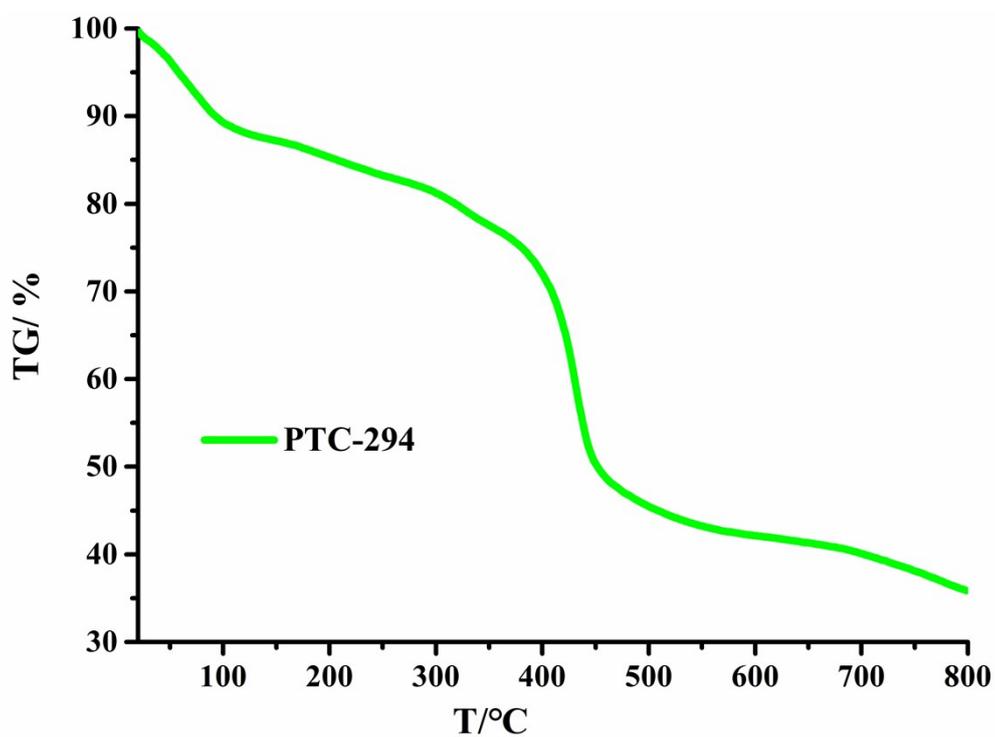


Figure S37. The TGA curve of PTC-294.

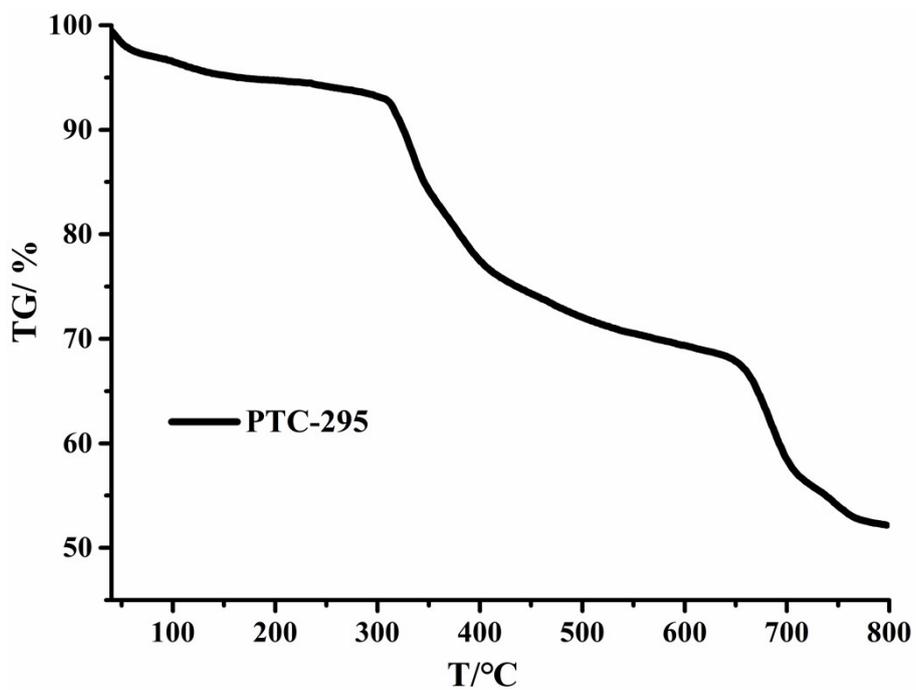


Figure S38. The TGA curve of PTC-295.

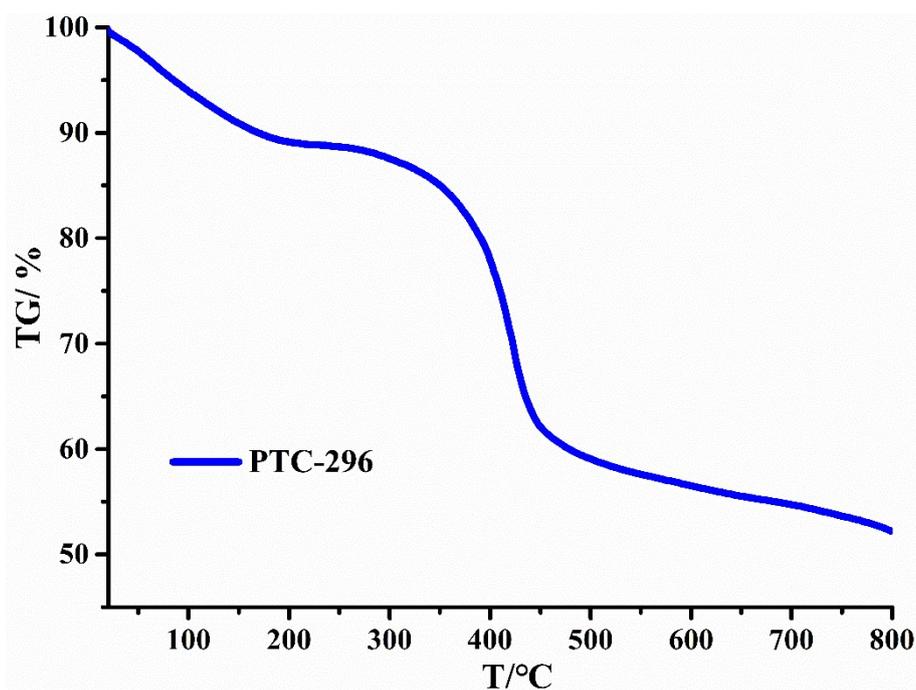


Figure S39. The TGA curve of PTC-296.

## 8. The XPS Spectra of PTCs

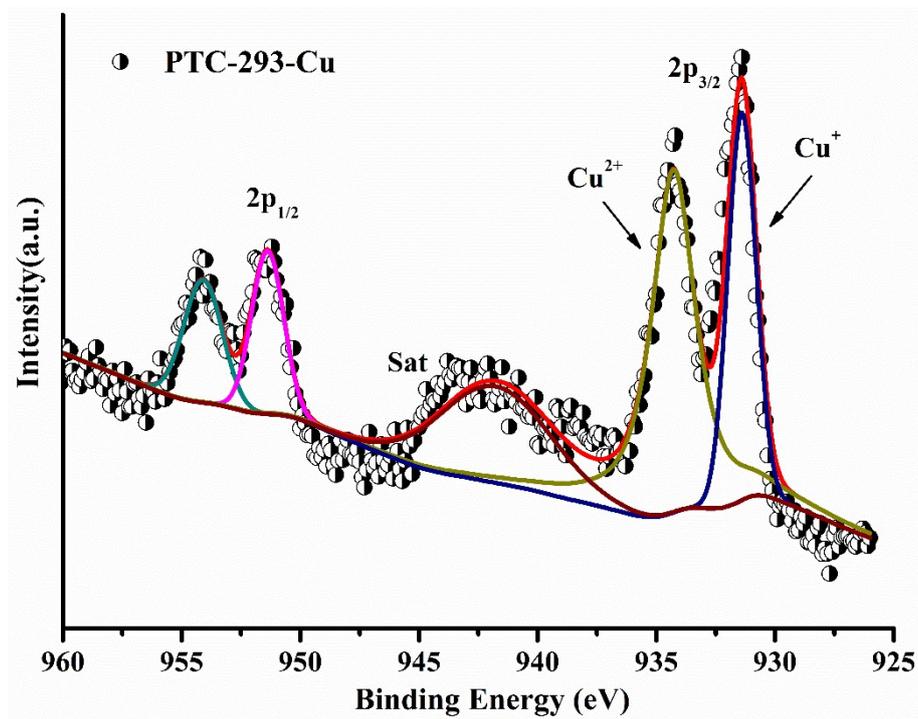


Figure S40. The XPS spectra of PTC-293-Cu.

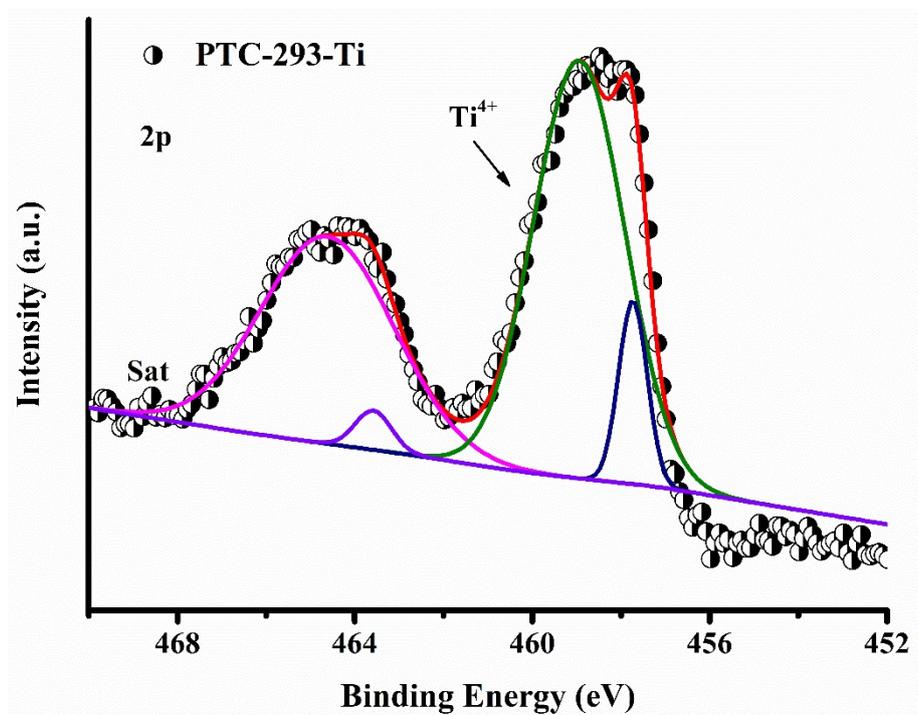


Figure S41. The XPS spectra of PTC-293-Ti.

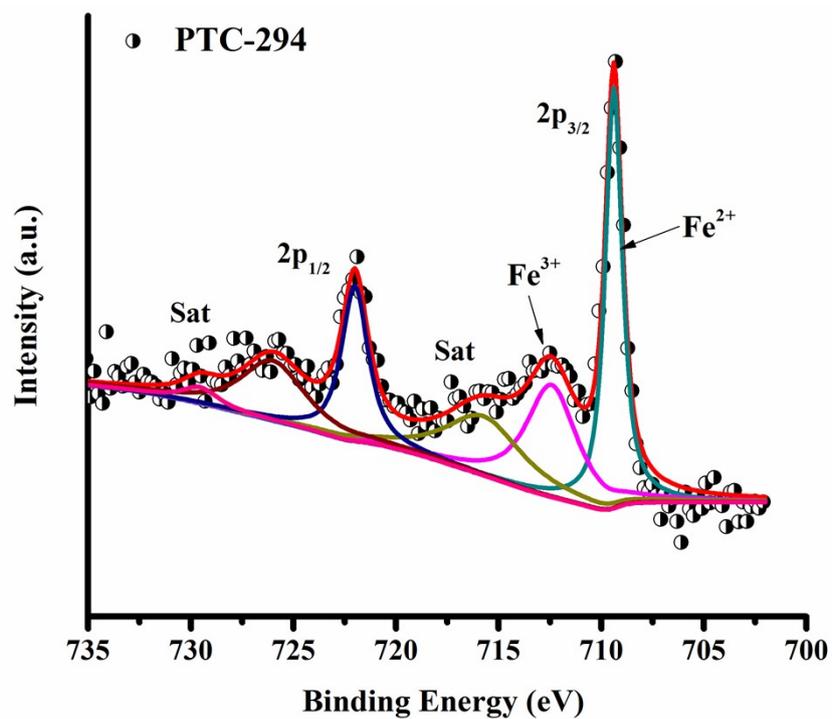


Figure S42. The XPS spectra of PTC-294.

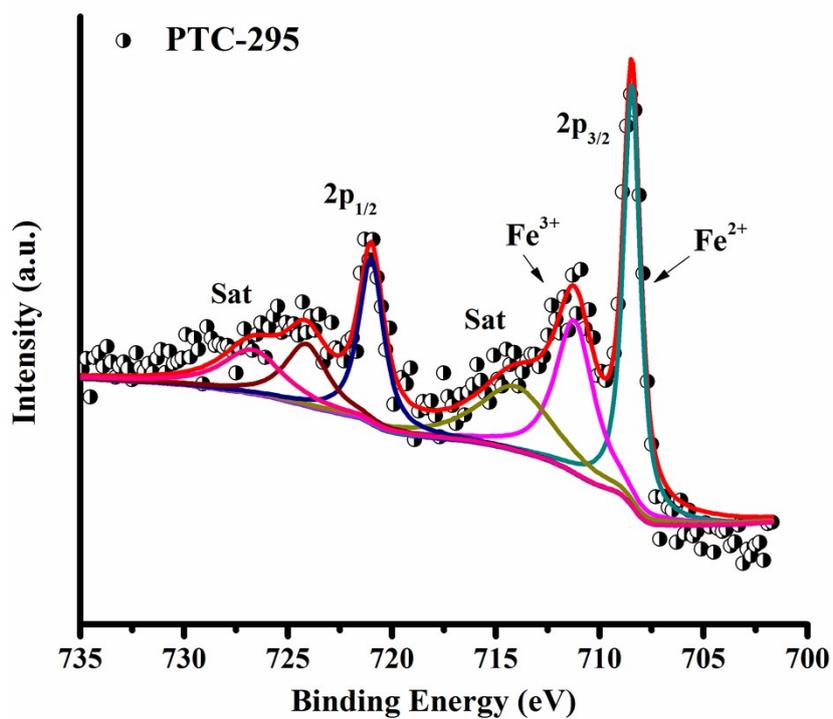


Figure S43. The XPS spectra of PTC-295.

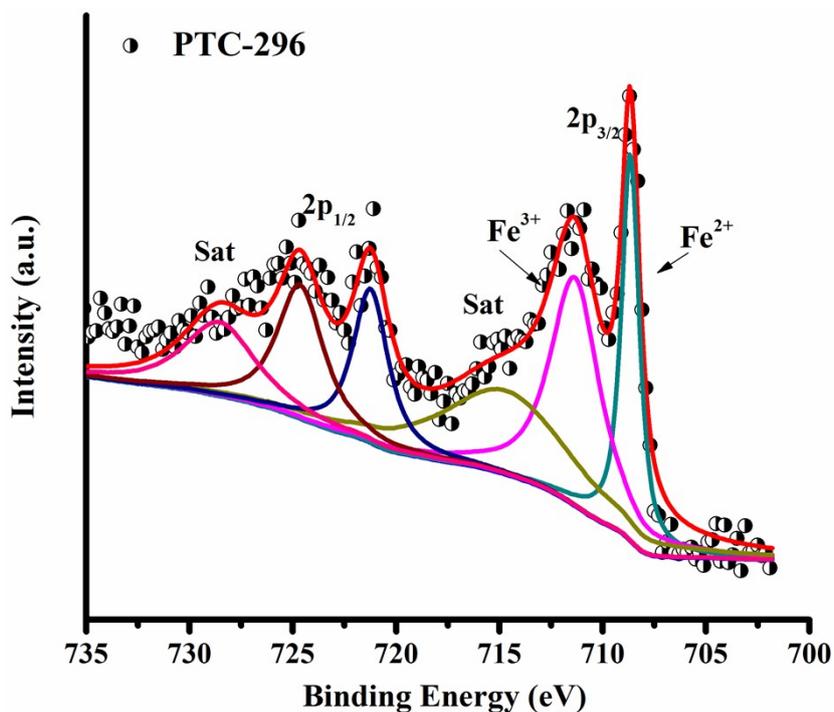


Figure S44. The XPS spectra of PTC-296.

X-ray photoelectron spectroscopy (XPS) was used to characterize the Fe oxidation state. It has been shown in previous studies that the peak positions of Fe 2p<sub>1/2</sub> and Fe 2p<sub>3/2</sub> depend on the ionic states of Fe. In **PTC-294**, all Fe is in the Fe<sup>2+</sup> state, while in **PTC-295** and **PTC-296**, iron is in mixed price state, but most of them are Fe<sup>3+</sup> valency. This result is consistent with the raw material of iron.

## 9. The NLO Property and NICS(0) Method Studies of PTCs

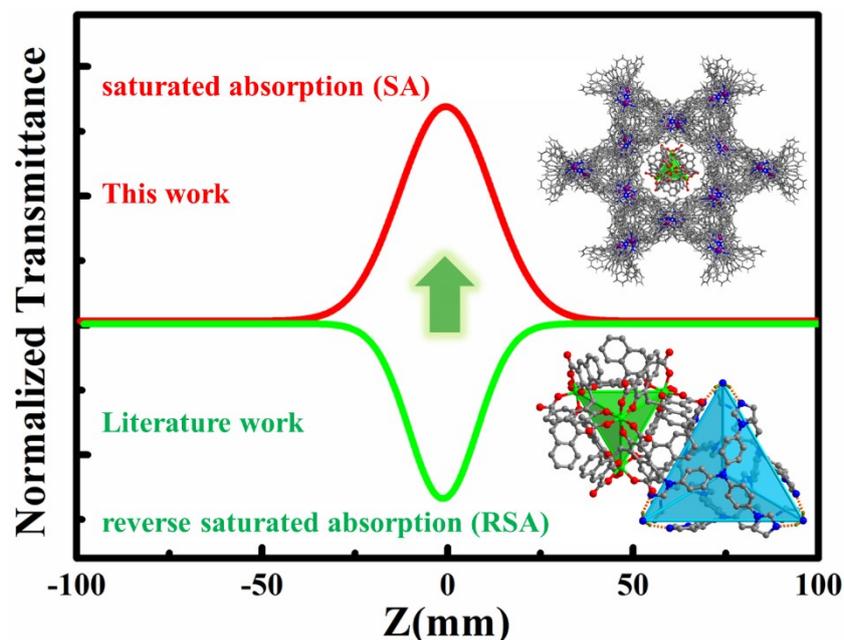


Figure S45.  $\text{Ti}_4(\text{embonate})_6$ -based ionic pairs with tunable nonlinear optics in our reported work and this work.<sup>2</sup>

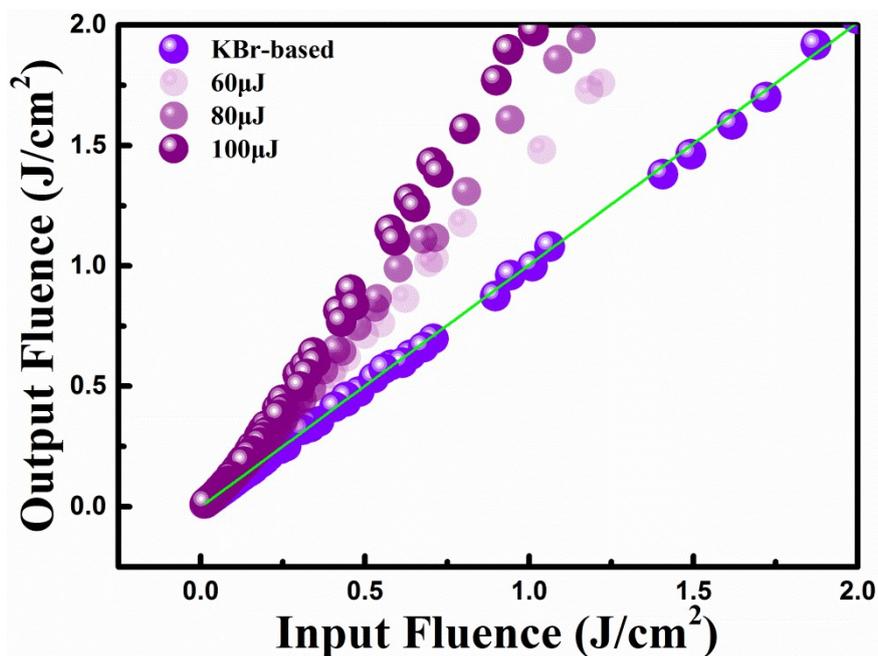


Figure S46. The curves of output fluence versus input fluence for two thin films.

The curves of output fluence versus input fluence of KBr-based thin film and KBr-based **PTC-296** thin film reveal that the output fluence linearly increased at low incident fluence (Figure S46). But at high-incident fluence, the output fluence deviated from linearity represent typical behavior of NLO response for KBr-based **PTC-296** thin film.

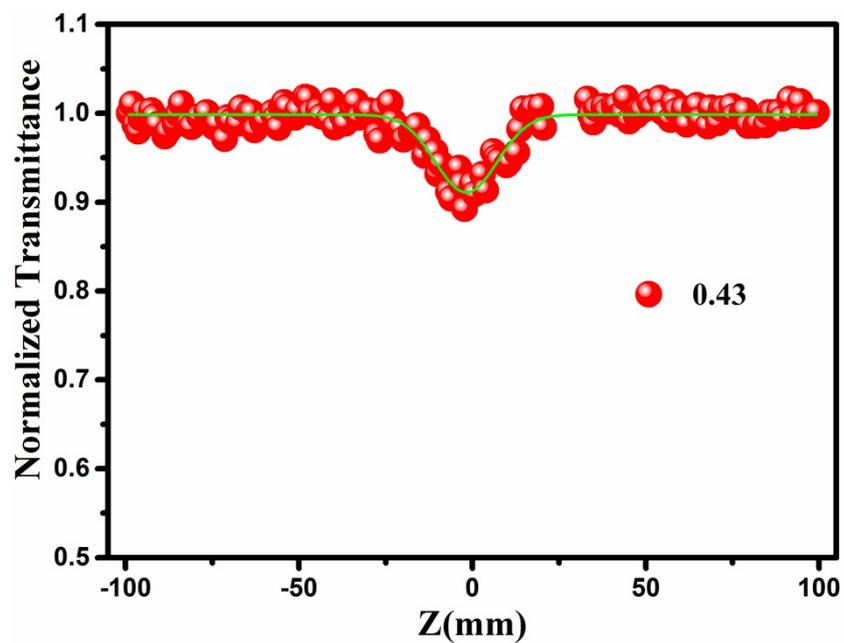


Figure S47. The open-aperture Z-scan results at 532 nm for KBr-based PTC-293 thin film.

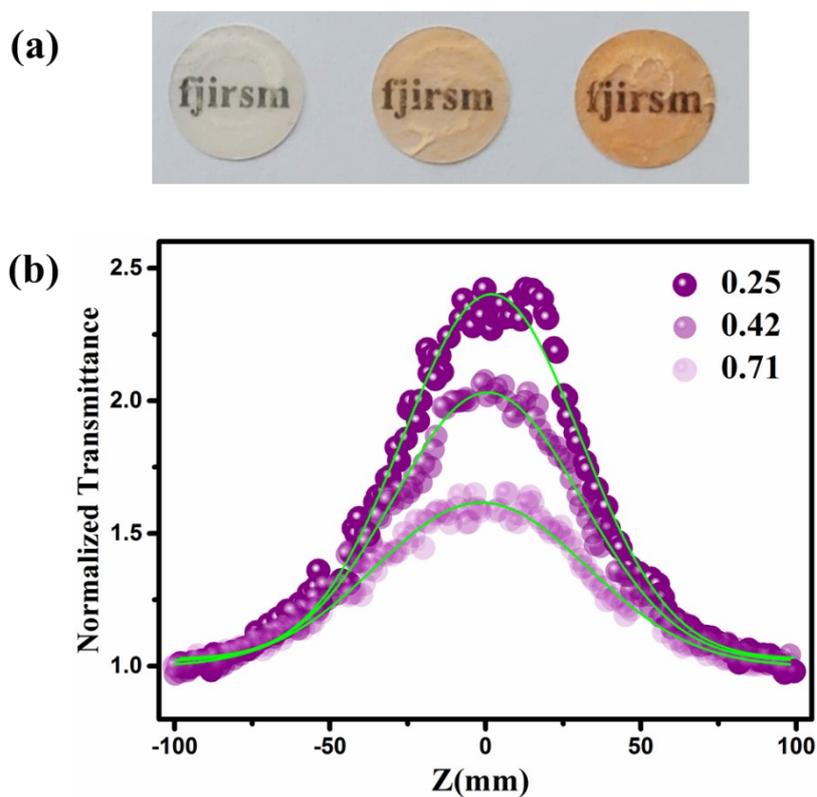
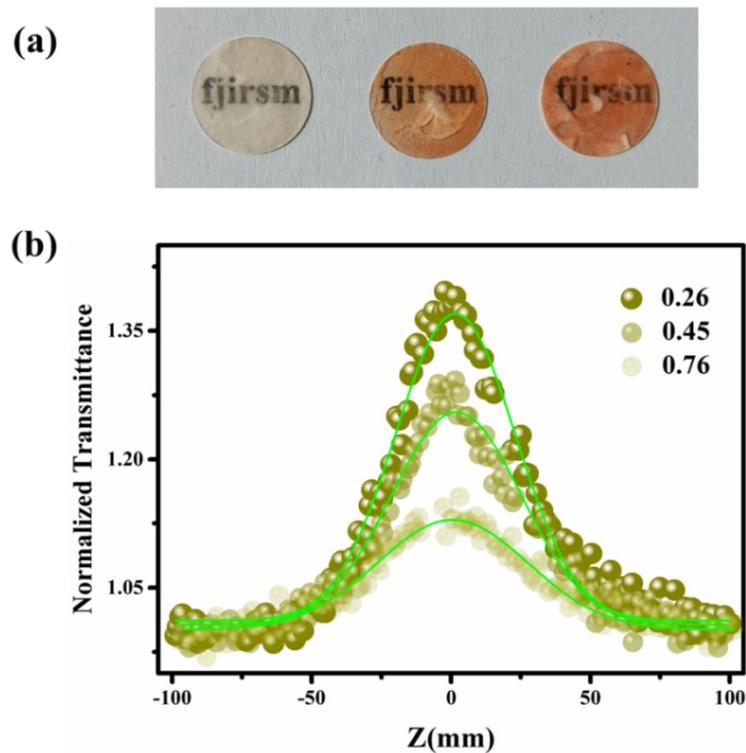
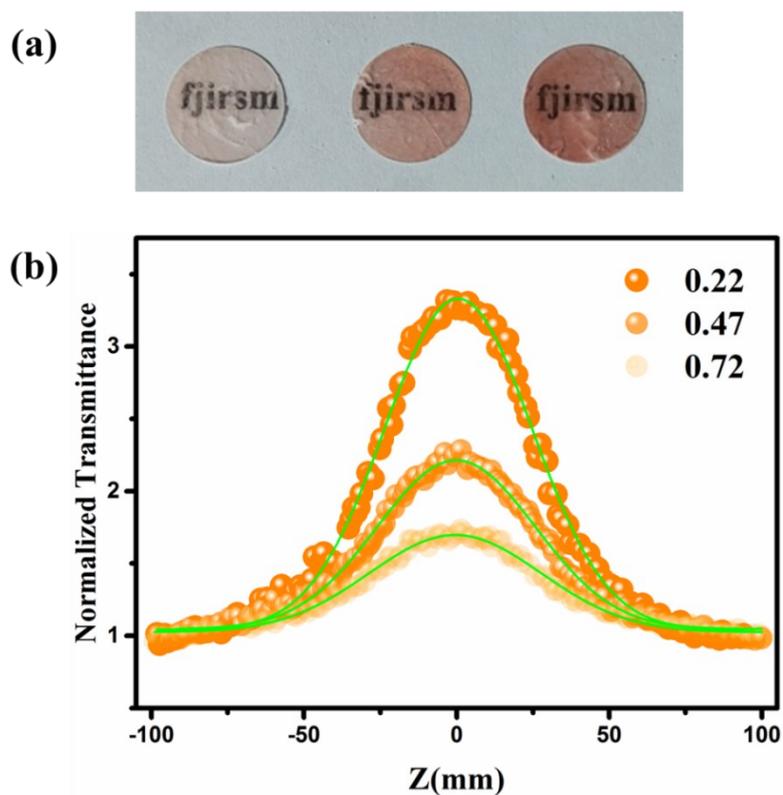


Figure S48. The open-aperture Z-scan results at 532 nm for KBr-based PTC-294 thin film and the linear transmittance ( $T_0$ ) was set.



**Figure S49.** The open-aperture Z-scan results at 532 nm for KBr-based PTC-295 thin film and the linear transmittance ( $T_0$ ) was set.

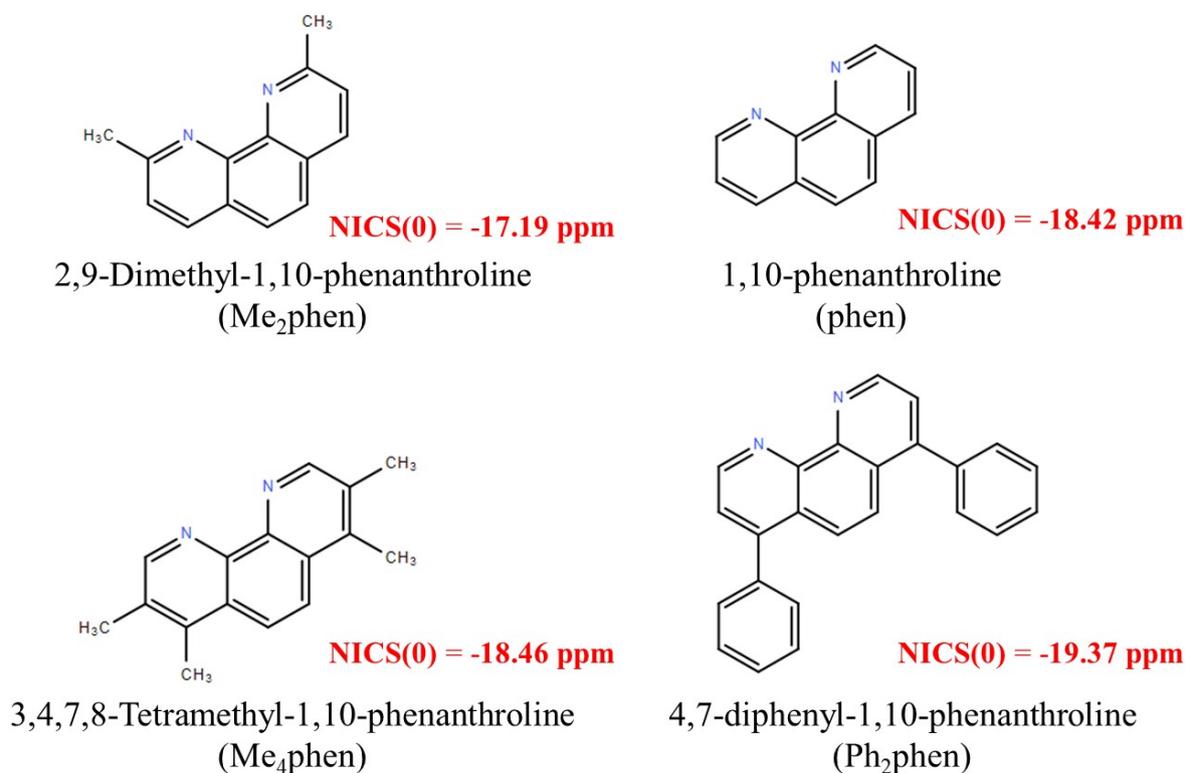


**Figure S50.** The open-aperture Z-scan results at 532 nm for KBr-based PTC-296 thin film and the linear transmittance ( $T_0$ ) was set.

**Table S1.** Linear, NLO data and NICS(0) values of the samples.

	<b>Linear Transmittance (T<sub>0</sub>)</b>	<b>Normalized Transmittance</b>	<b>Nonlinear Absorption Coefficients (β(m/W*10<sup>-10</sup>))</b>	<b>Nucleus-Independent Chemical Shift (NICS(0))</b>
<b>PTC-293</b>	0.43	0.91	6.00	-17.19
<b>PTC-294</b>	0.71	1.71	-11.50	-18.42
<b>PTC-295</b>	0.45	1.29	-8.30	-18.46
<b>PTC-296</b>	0.47	2.28	-16.10	-19.37

To test the aromaticity of the selected series of 1,10-phenanthroline type ligands we calculated the Nucleus-Independent Chemical Shift (NICS) values, which has become the most widely applied index to assess aromaticity due to its simplicity and efficiency.<sup>3-5</sup> NICS(0) values are defined as the value of absolute shielding observed at a ring center. Positive NICS(0) values denote antiaromaticity and negative NICS(0) values indicate the presence of aromatic character.

**Figure S51.** The NICS(0) values for selected 1,10-phenanthroline type ligands.

## 10. Crystallography Data

**Table S2.** Crystallographic Data and Structure Refinement Details for **PTC-293** to **PTC-296**.

Compound	PTC-293	PTC-294	PTC-295	PTC-296
CCDC	2107093	2107094	2107095	2107096
Formula	$C_{166}H_{96}CuN_4O_{36}Ti_4$	$C_{225}H_{157}Fe_2N_{17}O_{42}Ti$	$C_{250}H_{204}Fe_2N_{16}O_{40}Ti$	$C_{446}H_{294}Fe_3N_{34}O_{36}Ti$
$D_{calc.}/g\ cm^{-3}$	1.124	4 1.232	4 1.125	4 1.209
$\mu/mm^{-1}$	2.176	2.835	2.432	1.426
Formula	2977.60	4073.97	4375.58	7064.36
Weight				
Colour	orange	crimson	crimson	black
Shape	block	block	block	cube
Size/mm <sup>3</sup>	0.30×0.20×0.10	0.30×0.20×0.10	0.30×0.20×0.10	0.30×0.20×0.10
T/K	100.00(19)	99.98(12)	100.00(16)	100.00(11)
Crystal System	monoclinic	triclinic	triclinic	trigonal
Space Group	<i>Cc</i>	<i>P</i> -1	<i>P</i> -1	<i>R</i> -3 <i>c</i>
<i>a</i> /Å	31.8045(8)	17.6307(5)	19.8755(10)	40.7336(3)
<i>b</i> /Å	17.8553(3)	19.9692(6)	25.7801(10)	40.7336(3)
<i>c</i> /Å	34.7510(8)	34.0458(9)	25.8788(10)	81.0157(7)
$\alpha/^\circ$	90	99.561(2)	79.355(3)	90
$\beta/^\circ$	116.935(3)	98.330(2)	83.215(4)	90
$\gamma/^\circ$	90	108.172(3)	85.683(4)	120
V/Å <sup>3</sup>	17593.6(8)	10980.3(6)	12921.2(10)	116413.8(17)
<i>Z</i>	4	2	2	11.99988
<i>Z'</i>	1	1	1	0.333333
Wavelength/Å	1.54184	1.54184	1.54184	1.3405
Radiation type	CuK $\alpha$	CuK $\alpha$	CuK $\alpha$	GaK $\alpha$
$\theta_{min}/^\circ$	2.779	2.385	2.240	2.187
$\theta_{max}/^\circ$	50.436	66.399	64.865	56.909
Measured	18376	74929	120694	156225
Refl's.				
Indep't Refl's	11000	34861	42767	26453
Refl's $I \geq 2\ \sigma(I)$	10225	20925	21067	20850
$R_{int}$	0.0265	0.0747	0.0986	0.0333
Parameters	1883	2620	2835	1546
Restraints	712	249	444	132
Largest Peak	1.235	0.75	0.79	1.44
Deepest Hole	-0.397	-0.62	-0.78	-0.89
GooF	1.026	1.030	1.075	1.052
$wR_2$ (all data)	0.1627	0.1835	0.4488	0.3480
$wR_2$	0.1569	0.1572	0.3807	0.3249
$R_1$ (all data)	0.0620	0.1136	0.2012	0.1123
$R_1$	0.0577	0.0647	0.1358	0.0984

## 11. Reference

- 1 He, Y.-P.; Yuan, L.-B.; Chen, G.-H.; Lin, Q.-P.; Wang, F.; Zhang, L.; Zhang, J. Water-Soluble and Ultrastable  $Ti_4L_6$  Tetrahedron with Coordination Assembly Function. *J. Am. Chem. Soc.* **2017**, *139*, 16845-16851.
- 2 He, Y.-P.; Chen, G.-H.; Li, D.-J.; Li, Q.-H.; Zhang, L.; Zhang, J. Combining a Titanium–Organic Cage and a Hydrogen-Bonded Organic Cage for Highly Effective Third-Order Nonlinear Optics. *Angew. Chem. Int. Ed.* **2021**, *60*, 2920-2923.
- 3 Williams, R. V.; Armantrout, J. R.; Twamley, B.; Mitchell, R. H.; Ward, T. R.; Bandyopadhyay, S. A Theoretical and Experimental Scale of Aromaticity. The First Nucleus-Independent Chemical Shifts (NICS) Study of the Dimethyldihdropyrene Nucleus. *J. Am. Chem. Soc.* **2002**, *124*, 13495-13505.
- 4 Ozer, M. S.; Menges, N.; Keskin, S.; Sahin, E.; Balci, M. Synthesis of Pyrrole-Fused C,N-Cyclic Azomethine Imines and Pyrazolopyrrolopyrazines: Analysis of Their Aromaticity Using Nucleus-Independent Chemical Shifts Values. *Org. Lett.* **2016**, *18*, 408-411.
- 5 Pagano, J. K.; Xie, J.; Erickson, K. A.; Cope, S. K.; Scott, B. L.; Wu, R.; Waterman, R.; Morris, D. E.; Yang, P.; Gagliardi, L.; Kiplinger, J. L. Actinide 2-metallabiphenylenes that satisfy Huckel's rule. *Nature* **2020**, *578*, 563-567.