Supplementary Information for

High molecular weight PE elastomers through 4,4-difluorobenzhydryl substitution in

symmetrical α -diimino-nickel ethylene polymerization catalysts

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¹H NMR (400 MHz, CDCl₃, 25 °C) spectra



Figure S2 ¹H NMR spectrum of L2



Figure S4 ¹H NMR spectrum of L4; * refers to δ H for water



Figure S5 ¹H NMR spectrum of L5; * refers to δ H for water, n-hexane and grease.



Figure S7 ¹³C NMR spectrum of L2



Figure S9 ¹³C NMR spectrum of L4



Figure S10 ¹³C NMR spectrum of L5

FT-IR spectra



Figure S11 FT-IR spectrum of L1



Figure S12 FT-IR spectrum of L2



Figure S13 FT-IR spectrum of L3



Figure S14 FT-IR spectrum of L4



Figure S15 FT-IR spectrum of L5



Figure S16 FT-IR spectrum of Ni1



Figure S17 FT-IR spectrum of Ni2



Figure S18 FT-IR spectrum of Ni3



Figure S19 FT-IR spectrum of Ni4



Figure S20 FT-IR spectrum of Ni5

¹H NMR (470 MHz, CDCl₃, 25 °C) spectra



Figure S21 ¹H NMR spectrum of Ni1 along with an expansion of the δ 10.0 to 0.0 ppm region



Figure S22 ¹H NMR spectrum of Ni2 along with an expansion of the δ 10.0 to 0.0 ppm region



Figure S23 ¹H NMR spectrum of Ni3 along with an expansion of the δ 10.0 to 0.0 ppm region



Figure S24 1H NMR spectrum of Ni4 along with an expansion of the δ 10.0 to 0.0 ppm region



Figure S25 ¹H NMR spectrum of Ni5 along with an expansion of the δ 10.0 to 0.0 ppm region

¹⁹F NMR (470 MHz, CDCl₃, 25 °C) spectra



Figure S26 ¹⁹F NMR spectrum of Ni1



Figure S27 ¹⁹F NMR spectrum of Ni2







Figure S29 ¹⁹F NMR spectrum of Ni4



Figure S31 ¹⁹F NMR spectra of a) L1 and its nickel complex Ni1 along with that for b) L3 and Ni3.

GPC traces for the polyethylenes



Figure S32 a) GPC traces and b) plots of catalytic activity and molecular weight of the polyethylene produced using Ni2/MAO at different Al:Ni molar ratios (entries 1 - 5, Table 3).



Figure S33 a) GPC traces and b) plots of catalytic activity and molecular weight of the polyethylene produced using Ni2/MAO at various run temperatures (entries 2 and 6 - 9, Table 3).



Figure S34 a) GPC traces and b) plots of catalytic activity and molecular weight of the polyethylene produced using **Ni2**/MAO at various time (entries 2 and 10-13, Table 3).



Figure S35 a) GPC traces and b) plots of catalytic activity and molecular weight of the polyethylene produced using Ni1 – Ni5 in combination with MAO (entries 2 and 16 - 19, Table 3).



Figure S36 a) GPC traces and b) plots of catalytic activity and molecular weight of the polyethylene produced using Ni2/ EtAlCl₂ at different Al:Ni molar ratios (entries 1-6, Table 4)



Figure S37 a) GPC traces and b) plots of catalytic activity and molecular weight of the polyethylene produced using Ni2/ EtAlCl₂ at different run temperatures (entries 4 and 7-10, Table4)



Figure S38 a) GPC traces and b) plots of catalytic activity and molecular weight of the polyethylene produced using Ni2/ $EtAlCl_2$ at different run time (entries 4 and 7-10, Table4)



Figure S39 a) GPC traces and b) plots of the catalytic activity and molecular weight of the polyethylene produced using Ni1 – Ni5 in combination with $EtAlCl_2$ (entries 2 and 17 – 20, Table 4).

¹³C NMR (470 MHz, chlorobenzene-d₅, 100 °C) spectra



Figure S41 ¹³C NMR spectrum of PE-EtAlCl₂40Ni5 produced using Ni5/EtAlCl₂ (entry 21, Table 4)

¹³C NMR (470 MHz, o-dichlorobenzene-d₄, 100 °C) spectra



Figure S42 ¹³C NMR spectrum of PE-Et₂AlCl40Ni5 produced using Ni5/Et₂AlCl (entry 23, Table 4)



Figure S43 ¹³C NMR spectrum of PE-EASC40Ni5 produced using Ni5/EASC (entry 24, Table4)

Stress-strain recovery test (performed at 30 °C)



Figure S44Stress-strain recovery tests for samplesPE-EtAlCl240Ni5, PE-EtAlCl250Ni5, PE-M40Ni5,PE-Et2AlCl40Ni5andPE-EASC40Ni5

	Ni1	Ni5	L1		
Empirical formula	$C_{210}H_{156}F_{24}N_6$	$C_{92}H_{81}Br_4F_8N_4Ni_2$	$C_{74}H_{60}Br_2F_8N_2Ni$		
Formula weight	3219.40	1831.67	1347.77		
Temperature/K	170.00(10)	170.00(13)	169.99(12)		
Crystal system	triclinic	triclinic	monoclinic		
Space group	P-1	P-1	$P2_1/n$		
a/Å	11.08370(10)	15.1623(12)	10.8443(4)		
b/Å	19.80860(10)	15.6426(9)	36.4718(12)		
c/Å	37.8410(3)	18.6622(9)	16.3540(7)		
α/\circ	89.0030(10)	78.343(4)	90		
β/°	84.9140(10)	78.089(5)	106.635(4)		
γ/°	83.4140(10)	78.104(6)	90		
Volume/Å ³	8220.55(11)	4178.7(5)	6197.5(4)		
Ζ	2	2	4		
$ ho_{calc}g/cm^3$	1.301	1.456	1.444		
µ/mm⁻¹	0.778	3.337	2.542		
F(000)	3348.0	1858.0	2752.0		
Crystal size/mm ³	$0.2\times0.15\times0.1$	$0.1\times0.05\times0.02$	$0.3\times0.1\times0.02$		
Radiation	Cu Ka (λ = 1.54184)	Cu Ka ($\lambda = 1.54184$)	Cu Ka (λ = 1.54184)		
2⊖ range for data collection/°	4.49 to 150.728	4.908 to 133.2	4.846 to 151.324		
Index ranges	$-13 \le h \le 13, -24 \le k$ $\le 24, -47 \le 1 \le 41$	$-18 \le h \le 18, -18 \le k$ $\le 18, -22 \le 1 \le 22$	$-12 \le h \le 13, -32 \le k$ $\le 44, -20 \le 1 \le 20$		
Reflections collected	120556	55023	46177		
Independent reflections	$32478 [R_{int} = 0.0293, R_{sigma} = 0.0299]$	14739 [$R_{int} = 0.1001$, $R_{sigma} = 0.0840$]	12209 [$R_{int} = 0.1010$, $R_{sigma} = 0.0654$]		
Data/restraints/para meters	32478/0/2173	14739/1068/1200	12209/0/790		
$\begin{array}{c} \text{Goodness-of-fit on} \\ F^2 \end{array}$	1.281	1.344	0.998		
Final R indexes	$R_1 = 0.1097, wR_2 =$	$R_1 = 0.1352,$	$R_1 = 0.1343,$		
$[1 \ge 2\sigma(1)]$ Final D indexes [all	0.289/	$WK_2 = 0.3446$ $P_1 = 0.1802$	$WR_2 = 0.3939$ $P_1 = 0.1624$		
rmar K muexes [all data]	$K_1 = 0.1223, WK_2 = 0.3090$	$K_1 = 0.1602,$ wR ₂ = 0.3816	$m_1 = 0.1034,$ wR ₂ = 0.4189		
Largest diff. peak/hole / e Å ⁻³	1.37/-0.55	2.88/-0.96	3.39/-1.02		

Table S1 Crystal data and structure refinements for L3, Ni1 and Ni5