

Supporting Information

Discriminative 'turn-on' fluorescent sensing of volatile halogenated solvents using a cleft-shaped 4-amino-1,8-naphthalimide Tröger's base fluorophore

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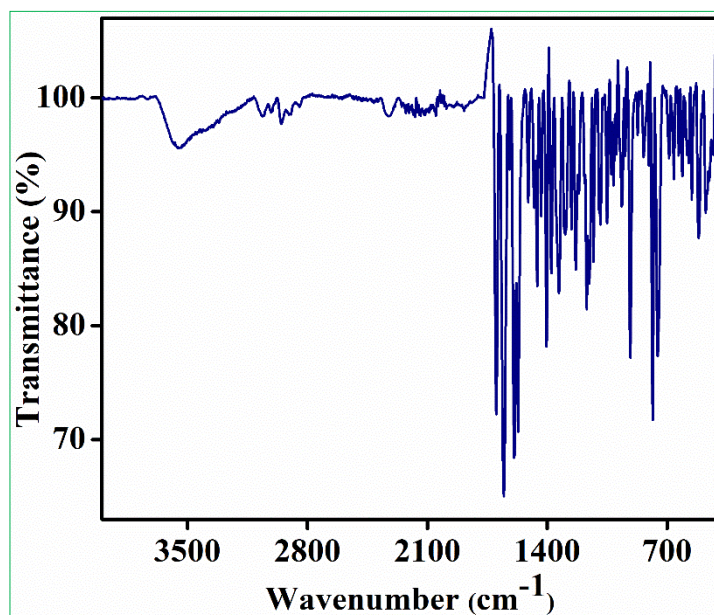


Fig. S1. FT-IR spectrum of TBNap.

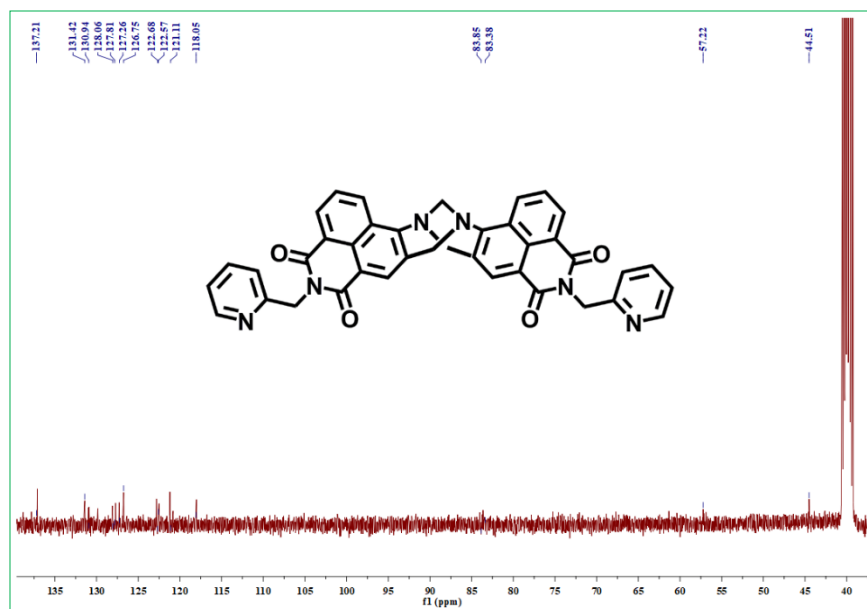


Fig. S2. ^{13}C NMR spectrum of **TBNap** (400 MHz, DMSO-d_6). Several attempts were made to obtain well resolved and high-intensity spectrum of the **TBNap** sensor. However, an increase in the concentration of **TBNap** in the DMSO medium was found to be precipitating/aggregating after a few minutes of preparing the NMR sample. Therefore, we were unable to produce a high-quality ^{13}C NMR spectrum. Nevertheless, ^1H NMR and HRMS unambiguously confirm the formation and purity of **TBNap**.

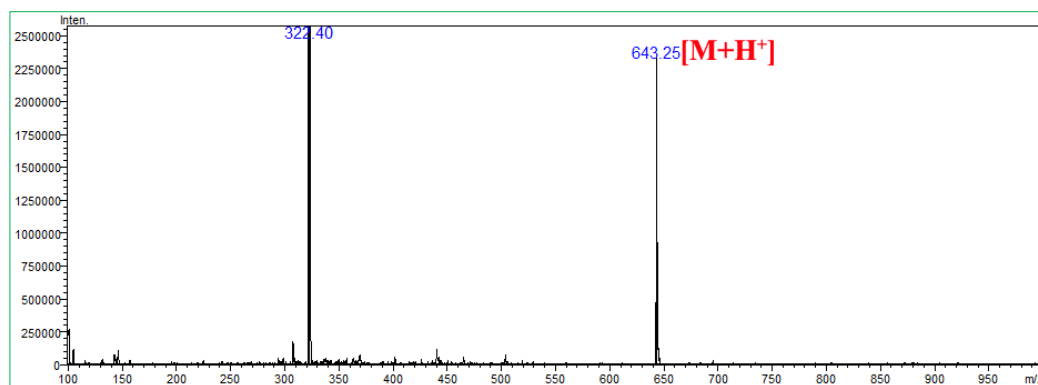


Fig. S3. HRMS Spectrum of **TBNap**.

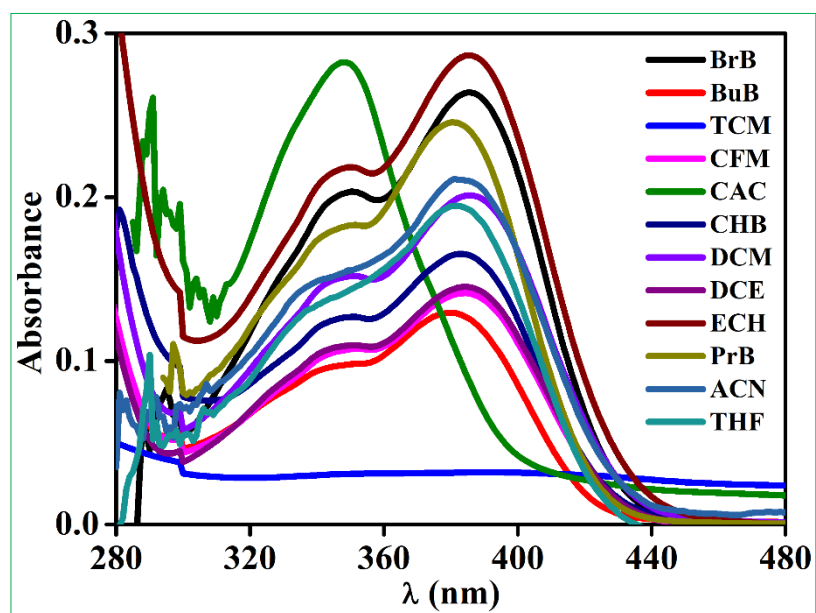


Fig. S4. UV-visible absorption spectra of TBNap in different organic solvents.

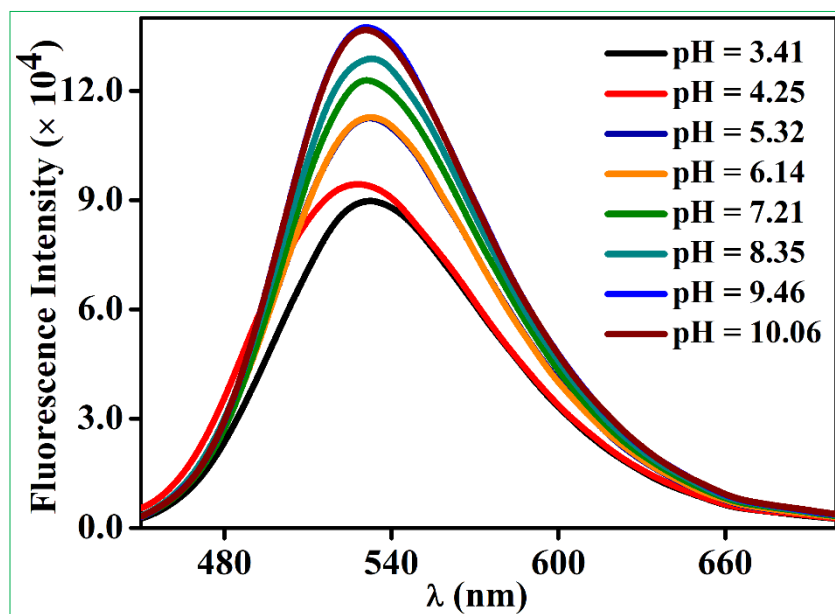


Fig. S5. The fluorescence emission spectra of TBNap were measured at different pH.

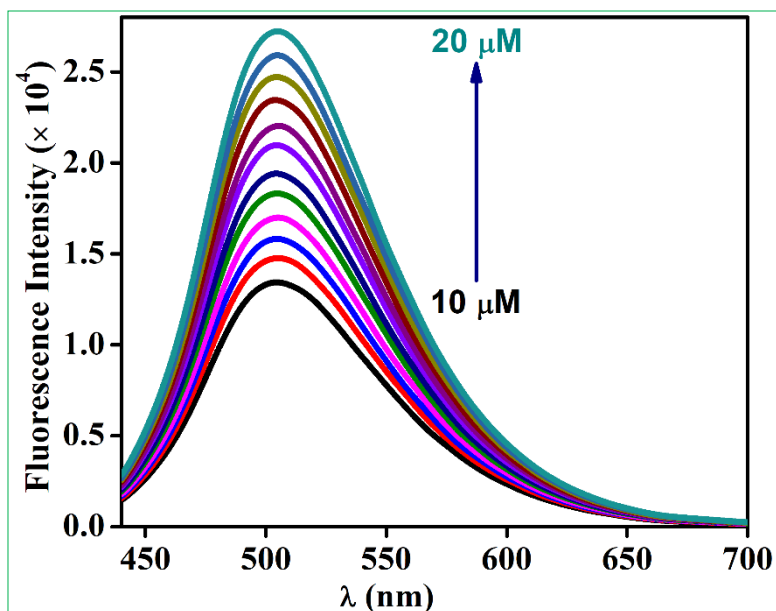


Fig. S6. Graph showing the concentration-dependent fluorescence emission studies.

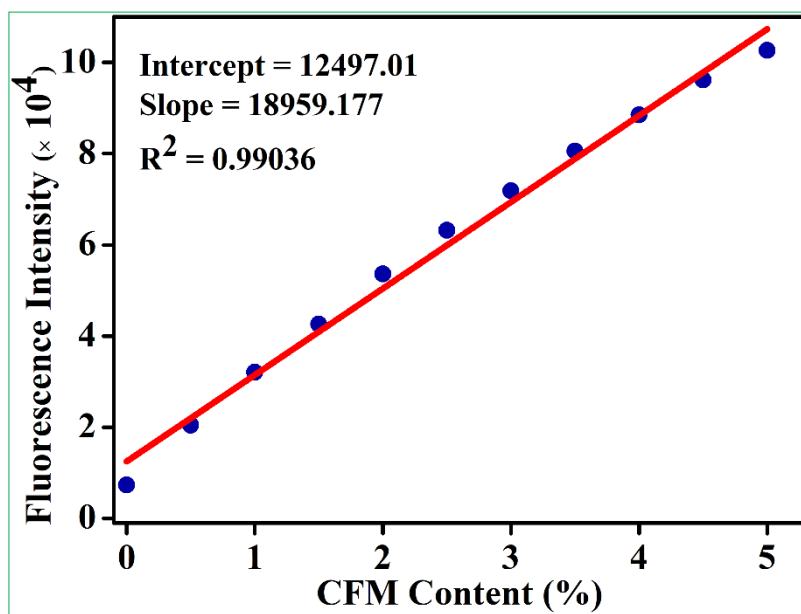


Fig. S7. A plot of fluorescence emission intensity vs. volume percentage of chloroform showing a linear curve.

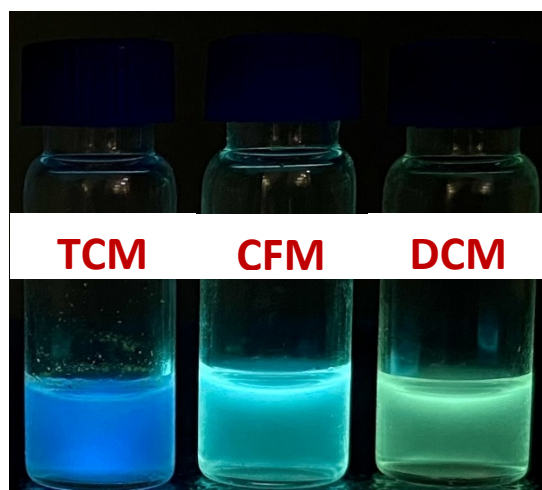


Fig. S8. The photograph of the visual color change for **TBNap** was imaged in TCE, CFM, and DCM solvents under UV-light irradiation.

Table S1. The measured emission maxima (λ_{max}) and quantum yield of **TBNap** in the selected different organic solvents and the dielectric constant for different solvents.

Solvents	Emission Maxima	Quantum Yield (Φ)	Dielectric Constant
CAC	436	0.0459	-
TCM	480	0.576	2.24
CFM	488	4.15	4.81
CHB	489	2.125	7.92
DCM	501	2.728	8.93
DCE	506	1.823	10.36
ECH	519	0.976	22.6
BrB	492	2.89	5.39
BuBr	486	2.338	6.93
PrBr	490	3.412	8.08
THF	507	0.978	7.58
ACN	530	0.262	37.5