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Supporting Information

Deep-blue high-efficiency triplet-triplet annihilation organic light-emitting diodes using hydroxyl-substituted tetraphenylimidazole-functionalized anthracene fluorescent emitter

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Synthesis and characterization

9,10-Bis(4-bromophenyl)anthracene (2): A mixture of **1** (1.00 g, 2.32 mmol), 1-bromo-4-iodobenzene (1.51 g, 5.34 mmol), Pd(PPh₃)₄ (135 mg, 0.12 mmol), 2M sodium carbonate (4.90 g, 46.27 mmol) aqueous solution and anhydrous THF (40 mL)was degassed with nitrogen for 15 min while stirring. Then the stirring reaction was heated and refluxed under nitrogen atmosphere for 36 h. The mixture was cooled to room temperature after confirming consumption of the starting materials by TLC in 10% v/v dichloromethanehexane. Then the mixture was added with brine (20 mL) and extracted with a certain amount of dichloromethane, then dried over anhydrous sodium sulfate and filtered. The extract was concentrated under reduced pressure and finally purified by silica gel column chromatography eluting with 10% v/v dichloromethanehexane to get white solids (795 mg, 70%). ¹H NMR (600 MHz, CDCl₃) δ 7.75 (d, J = 8.2 Hz, 4H), 7.67 (dd, J = 6.8, 3.3 Hz, 4H), 7.39-7.33 (m, 8H). ¹³C NMR (151 MHz, CDCl₃) δ 137.88, 136.00, 133.00, 131.74, 129.74, 126.69, 125.39, 121.86. HSMS (APCI): m/z calcd for C₂₆H₁₆Br₂ (M+): 487.9598, found: 489.0103 (MH+).

3-(4-(10-(4-Bromophenyl)anthracen-9-yl)phenyl)-9-phenyl-9H-carbazole (3): A mixture of **2** (2.50 g, 5.12 mmol), (9-phenyl-9H-carbazol-3-yl)boronic acid (490 mg, 1.71 mmol), Pd(PPh₃)₄ (39 mg, 0.03 mmol), 2M aqueous solution of sodium carbonate (10 mL) and anhydrous THF (40 mL) was degassed with nitrogen for 15 min while stirring. Then the stirring reaction was heated and refluxed under nitrogen atmosphere for 12 h. The mixture was cooled to room temperature after confirming consumption of the starting materials by TLC in 10% v/v dichloromethane-hexane. Then the mixture was added with brine (20 mL) and extracted with a certain amount of dichloromethane, then dried over anhydrous sodium sulfate and filtered. The extract was concentrated under reduced pressure and finally purified by silica gel column chromatography eluting with 10% v/v dichloromethane-hexane to get white solids (800 mg, 72%). ¹H NMR (600 MHz, CDCl₃) δ 8.55 (s, 1H), 8.25 (d, J = 7.2 Hz, 1H), 7.97 (d, J = 8.4 Hz, 2H), 7.88-7.82 (m, 3H), 7.76 (d, J = 7.8 Hz, 2H), 7.70-7.67 (m, 2H), 7.67-7.63 (m, 4H), 7.59 (d, J = 7.8 Hz, 2H), 7.56 (d, J = 8.4 Hz, 1H), 7.53-7.50 (m, 1H), 7.48-

7.44 (m, 2H), 7.41-7.37 (m, 6H), 7.38 - 7.33 (m, 1H). ¹³C NMR (151 MHz, CDCl₃) δ 139.55, 139.28, 138.67, 136.21, 135.81, 135.58, 135.19, 133.62, 131.19, 129.86, 129.81, 128.09, 128.07, 127.92, 125.69, 125.37, 125.33, 125.24, 124.69, 124.32, 123.59, 123.45, 123.22, 122.14, 121.65, 119.86, 118.49, 118.27, 117.00, 108.28, 108.11. HSMS (APCI): m/z calcd for C₄₄H₂₈BrN (M+): 649.1405, found: 650.1385 (MH+).

9-Phenyl-3-(4-(10-(4-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)phenyl)anthracen-9-yl)phenyl)-9H-

carbazole (4): A mixture of **3** (500 mg, 0.77 mmol), bis(pinacolato)diboron (587 mg, 2.31 mmol), KOAc (906 mg, 9.25 mmol), and Pd(dppf)₂Cl₂·CH₂Cl₂ (31 mg, 0.32 mmol) were mixed followed by adding 30 mL anhydrous Toluene. The mixture was degassed with nitrogen for 15 min while stirring. Then the stirring reaction was heated and refluxed under nitrogen atmosphere for 15 h. The mixture was cooled to room temperature after confirming consumption of the starting materials by TLC in 30% v/v dichloromethanehexane. Then the mixture was added with brine (20 mL) and extracted with a certain amount of dichloromethane, then dried over anhydrous sodium sulfate and filtered. The extract was concentrated under reduced pressure and finally purified by silica gel column chromatography eluting with 30% v/v dichloromethane-hexane to get white solids (419 mg, 78%). ¹H NMR (600 MHz, CDCl₃) δ 8.55 (s, 1H), 8.25 (d, J = 7.2 Hz, 1H), 8.07 (d, J = 7.2 Hz, 2H), 7.97 (d, J = 7.2 Hz, 2H), 7.88-7.82 (m, 3H), 7.71 (d, J = 8.4 Hz, 2H), 7.68-7.63 (m, 4H), 7.60 (d, J = 7.2 Hz, 2H), 7.57-7.49 (m, 4H), 7.48-7.43 (m, 2H), 7.35 (dt, J = 20.4, 6.6 Hz, 5H), 1.44 (s, 12H). ¹³C NMR (151 MHz, CDCl₃) δ 142.28, 141.46, 141.09, 140.56, 137.74, 137.28, 137.07, 137.01, 134.81, 133.16, 131.82, 130.84, 129.99, 129.96, 129.76, 127.57, 127.25, 127.14, 127.09, 126.97, 126.19, 125.51, 125.07, 125.04, 124.04, 123.57, 120.40, 120.16, 118.90, 110.17, 109.99, 83.97, 24.99. HSMS (APCI): m/z calcd for C₅₀H₄₀BNO₂ (M+): 697.3125, found: 698.2077 (MH+).

4-Bromo-2-(1,4,5-triphenyl-1H-imidazol-2-yl)phenol (6) & 2-(3-Bromophenyl)-1,4,5-triphenyl-1Himidazole (7): A mixture of **5** (500 mg, 2.38 mmol) and 5-bromosalicylaldehyde (478 mg, 2.38 mmol) or 3bromobenzaldehyde (440 mg, 2.38 mmol) were mixed at room temperature followed by adding 30 mL of acetic acid. Aniline of 1.022 g/mL (0.33 mL, 3.57 mmol or 0.33 mL, 3.60 mmol) was then added to this solution drop by drop, and ammonium acetate (917 mg, 11.9 mmol or 925 mg, 12 mmol) was added subsequently. The mixture was heated at 110 °C for 12 h under nitrogen atmosphere and reflux device. After termination of the reaction, the dark solution was poured into a copious amount of water in a 1000 mL beaker. Then a certain amount of sodium hydrogen carbonate was added to the mixture until neutralizing. The crude product was purified by recrystallization in dichloromethane.

6 as white solids (678 mg, 61%). ¹H NMR (600 MHz, CDCl₃) δ 7.57-7.53 (m, 2H), 7.47-7.40 (m, 3H), 7.32-7.22 (m, 7H), 7.22-7.14 (m, 5H), 6.98 (d, J = 8.4 Hz, 1H), 6.62 (s, 1H). ¹³C NMR (151 MHz, CDCl₃) δ 157.39, 143.44, 136.23, 134.83, 132.91, 132.05, 131.27, 130.80, 129.80, 129.67, 129.12, 129.01, 128.78, 128.60, 128.43, 128.42, 127.48, 127.10, 119.57, 114.22, 109.76. HSMS (APCI): m/z calcd for C₂₇H₁₉BrN₂O (M+): 466.0681, found: 467.0786 (MH+).

7 as white solids (848 mg, 79%). ¹H NMR (600 MHz, CDCl₃) δ 7.72 (s, 1H), 7.59 (d, J = 7.2 Hz, 2H), 7.39 (d, J = 7.8 Hz, 1H), 7.35-7.18 (m, 10H), 7.13 (d, J = 7.2 Hz, 2H), 7.08-7.02 (m, 3H). ¹³C NMR (151 MHz, CDCl₃) δ 145.27, 138.58, 136.81, 134.23, 132.49, 131.92, 131.33, 131.21, 131.10, 130.39, 129.47, 129.25, 128.58, 128.40, 128.38, 128.22, 128.12, 127.38, 127.17, 126.76, 122.29. HSMS (APCI): m/z calcd for C₂₇H₁₉BrN₂ (M+): 450.0732, found: 451.0560 (MH+).

All crystallographic information (CIF) data including CCDC numbers of 2017513 and 2051106 for HO-PIAC and PIAC, respectively, were deposited on and can be obtained with no cost at https://www.ccdc.cam.ac.uk/.

able S1. Crystallographic data					
Compound name	НО-РІАС	PIAC			
CCDC number	2017513	2051106			
Empirical formula	C ₇₂ H ₄₈ Cl ₃ N ₃ O	C ₇₁ H ₄₇ N ₃			
Formula weight	1077.48	942.11			
Temperature/K	100	100			
Crystal system	triclinic	monoclinic			
Space group	P-1	C2/c			
a/Å	9.8323(5)	41.3796(13)			
b/Å	15.2021(8)	8.3576(3)			
c/Å	19.9693(11)	30.6626(9)			
α/°	71.560(2)	90			
β/°	88.608(2)	100.697(2)			
$\gamma/^{\circ}$	74.215(2)	90			
Volume/Å ³	2718.6(3)	10419.9(6)			
Ζ	2	8			
$ ho_{calc}/g/cm^3$	1.316	1.201			
μ/mm ⁻¹	0.219	0.532			
F(000)	1120.0	3952.0			
Crystal size/mm ³	0.42 imes 0.05 imes 0.02	$0.57 \times 0.23 \times 0.09$			
Radiation	MoKα (λ = 0.71073)	CuKa ($\lambda = 1.54178$)			

4.168 to 52.21

 $-12 \le h \le 12$,

 $-18 \le k \le 18$,

 $-24 \le l \le 24$

83699

 $10787 [R_{int} = 0.1341,$

 $R_{sigma} = 0.0750$]

10787/0/717

1.040

 $R_1 = 0.0690, wR_2 = 0.1461$

 $R_1 = 0.1352, wR_2 = 0.1719$

0.64/-0.60

4.346 to 136.472

 $-49 \le h \le 49$,

 $-9 \le k \le 10$,

 $-32 \le l \le 36$

61803

9533 [$R_{int} = 0.0856$,

 $R_{sigma} = 0.0432$]

9533/0/667

1.027

 $R_1 = 0.0531, wR_2 = 0.1340$

 $R_1 = 0.0687, wR_2 = 0.1432$

0.66/-0.35

Т

 2Θ range for

data collection/°

Index ranges

Reflections collected

Independent reflections

Data/restraints/parameters

Goodness-of-fit on F²

Final R indexes

 $[I \ge 2\sigma(I)]$ Final R indexes

[all data] Largest diff.

peak/hole / e Å-3



Fig .S1 Normalized PL spectra in varying solvents.



Fig. S2 S₀ optimized geometries of enol forms computed at B3LYP/6-31G(d,p) level.



Fig. S3 S₁ optimized geometries of enol forms computed at B3LYP/6-31G(d,p) level.

Table S2 Summary of relative proton transfer barriers and computed energy differences between the enol and keto forms ($\Delta E = E_{keto} - E_{enol}$) between S₀ and S₁ computed at B3LYP/ 6-31G(d,p) level.

Molecule	PT barrier (kcal/mol)		ΔE (kcal/mol)		
	S ₀	S ₁	S ₀	S ₁	
HO-PIAC	11.75	5.93	10.62	4.04	

Table S3 Selected bond lengths (Å) and torsional angles (°) at S_0 and S_1 optimized geometries computed at B3LYP/6-31G(d,p) and TD-B3LYP/6-31G(d,p) levels, respectively.

		Enol form				Keto form			
Molecule	State	Distance (Å)		Torsional angle (°)	Distance (Å)		Torsional angle (°)		
		01-	N1 II1	C3-	C2C2C4C5	01 111	N1-	C3-	C2C2C4C5
		H1	NI····HI	C4	02030405	01	H1	C4	02030403
HO-	S_0	0.996	1.697	1.483	36.06	1.552	1.035	1.480	33.46
PIAC	S_1	0.999	1.681	1.473	28.92	1.695	1.035	1.467	29.02
PIAC	S_0	-	-	1.485	36.48	-	-	-	-
	S_1	-	-	1.479	30.96	-	-	-	-

Table S4 Simulated enol absorption maximum wavelengths (λ_{abs}), enol and keto emission maximum wavelengths (λ_{em}), oscillator strength (f), molecular orbitals (MOs) contribution, and Stokes shift calculated by TD-B3LYP/6-31G(d,p) level.

	Absorption			Emis	sion	Stokes shift (nm)
	Enol			Enol	Keto	
	$\lambda_{abs} (nm)$	f	MOs Contribution	$\lambda_{em} (nm)$	$\lambda_{em} (nm)$	
HO-PIAC	399	0.3927	HOMO \rightarrow LUMO (97%)	499	549	150
PIAC	396	0.3582	HOMO \rightarrow LUMO (97%)	494	-	98



Fig. S4 DSC and TGA thermograms analyzed at a heating rate of 10°C min⁻¹ under N₂ flow.



Fig. S5 Cyclic voltammograms recorded in dry dichloromethane containing n-Bu₄NPF₆ at a scan rate of 50 mV s⁻¹ under argon atmosphere.



Fig. S6 a) and b) PL spectra and c) and d) transient PL decay spectra in neat films and doped (10, 20, 30 wt% doped CBP thin films.

	HO-PIAC (do	oped in CBP)	PIAC (doped in CBP)		
	neat	10%	neat	10%	
λ_{em} (nm)	451	441	450	437	
τ (ns)	0.70	2.62	1.21	2.54	
$\Phi_{ ext{PL}}$ (%)	19	56	40	57	

Table S5 Photoluminescent properties of doped films in various conditions.



Fig. S7 EL spectra of OLEDs at different applied voltages.



Fig. S8 Transient EL decay of OLED device based on PIAC at different voltages.



Fig. S9 Amplified transient EL decay plots of the delayed component at a driving voltage of 10 V.





Compound 3:



Compound 4:



Compound 6:



Compound 7:



210 200 190 180 170 160 150 140 130 120 110 100 90 80 70 60 50 40 30 20 10 0 ppm





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PIAC:

D:\Data\MSE\VIP_LAB\Wan Li\WL01-45_final\0_G19\1\1SRef

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