

## Electronic Supplementary Information (ESI)

### **Facile dimerization strategy for producing narrowband green multi-resonance delayed fluorescence emitters**

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## 1. General Methods

All reagents and anhydrous solvents were purchased from Sigma-Aldrich, Tokyo Chemical Industry (TCI), or Fujifilm Wako Pure Chemical Corp., and were used without further purification unless otherwise noted. 1,3-bis(1,8-dimethylcarbazol-9-yl)benzene (mMCP),<sup>[S1]</sup> 3,3'-di(carbazol-9-yl)-5-cyano-1,1'-biphenyl (mCBP-CN),<sup>[S2]</sup> and 1,3-bis[3,5-di(pyridin-3-yl)phenyl]benzene (B3PyPB)<sup>[S3]</sup> were prepared according to the literature procedures and purified by vacuum sublimation before use. 2,8-Bis(diphenylphosphinyl)dibenzo[b,d]furan (PPF) was purchased from TCI and purified by vacuum sublimation before use. 2,3,6,7,10,11-Hexacyano-1,4,5,8,9,12-hexaazatriphenylene (HAT-CN), 1,1-bis[4-[*N,N*-di(*p*-tolyl)amino]phenyl]cyclohexane (TAPC), and 8-quinolinolato lithium (Liq) were purchased from LG Chem Ltd., Luminescence Technology Corp., and e-Ray Optoelectronics Technology Co., Ltd., respectively, and used as received.

NMR spectra were recorded on an Avance III 400 spectrometer (Bruker). <sup>1</sup>H and <sup>13</sup>C NMR chemical shifts were determined relative to the signals of tetramethylsilane ( $\delta = 0.00$ ) and CDCl<sub>3</sub> ( $\delta = 77.0$ ), respectively. Matrix-assisted laser desorption ionization time-of-flight (MALDI-TOF) mass spectra were collected on an Autoflex III spectrometer (Bruker Daltonics) using dithranol as a matrix. Elemental analysis was carried out using an MT-5 CHN corder (Yanaco). Vacuum sublimation was performed using a P-100 system (ALS Technology). Thermogravimetric analysis (TGA) was performed on a TG/DTA7300 analyzer (Hitachi High-Tech Science) under a N<sub>2</sub> atmosphere.

## 2. Computational Simulations

All quantum chemical calculations based on the time-dependent density functional theory (TD-DFT) were performed using the ADF2021 program package.<sup>[S4]</sup> The ground-state ( $S_0$ ) geometries were initially optimized using the B3LYP functional with the DZP basis set in the gas phase. The vertical excitation calculations were carried out using the optimized  $S_0$  geometries, and the geometry optimizations in the excited  $S_m$  ( $m = 1, 2$ ) and  $T_n$  ( $n = 1-4$ ) states were performed using TD-DFT at the same level of theory. For the  $S_0 \rightarrow S_m$  and  $S_0 \rightarrow T_n$  transitions, the natural transition orbitals (NTOs) with their adiabatic excitation energies were simulated using the optimized  $S_m$  and  $T_n$  geometries, respectively. Using the respective  $T_n$  geometries, spin-orbit coupling (SOC) matrix elements,  $\langle S_m | \hat{H}_{\text{SOC}} | T_n \rangle$ , were calculated using a scalar relativistic TD-DFT with the two-component zeroth-order relativistic approximation (ZORA)<sup>[S5]</sup> at the same level of theory. The contributions of the three degenerate triplet states ( $T_{n,x}$ ,  $T_{n,y}$ , and  $T_{n,z}$ ) were taken into account by calculating the root sum square of the real (Re) and imaginary (Im) parts of the matrix elements, as expressed by the following equation:<sup>[S6]</sup>

$$\langle S_m | \hat{H}_{\text{SOC}} | T_n \rangle = \left\{ \sum_{a=x,y,z} (\text{Re}^2 \langle S_m | \hat{H}_{\text{SOC}} | T_{n,a} \rangle + \text{Im}^2 \langle S_m | \hat{H}_{\text{SOC}} | T_{n,a} \rangle) \right\}^{1/2} \quad (\text{Eq. S1})$$

### 3. Photophysical Measurements

Organic thin-film samples for photophysical measurements were deposited onto quartz substrates via vacuum deposition ( $< 7 \times 10^{-5}$  Pa) using an E-200 vacuum evaporation system (ALS Technology). UV-vis absorption and photoluminescence (PL) spectra were measured using a V-670Y spectrometer (Jasco) and an FP-8600Y spectrophotometer (Jasco), respectively. The absolute PL quantum yields ( $\Phi_{\text{PL}}$ ) were determined using an ILF-835 integrating sphere system (Jasco) under a  $\text{N}_2$  atmosphere. Transient PL decay measurements were carried out using a C11367 Quantaaurus-tau fluorescence lifetime spectrometer (Hamamatsu Photonics) with an LED excitation source ( $\lambda_{\text{ex}} = 340$  nm, pulse width = 100 ps, repetition rate = 1 kHz) under a  $\text{N}_2$  atmosphere. The emission lifetimes were determined from the PL decay curves by performing exponential fitting and deconvolution with the instrument response function.

The photophysical rate constants for radiative decay ( $k_{\text{r}}$ ), non-radiative decay ( $k_{\text{nr}}$ ), intersystem crossing ( $k_{\text{ISC}}$ ), and reverse intersystem crossing ( $k_{\text{RISC}}$ ) were calculated according to the literature method:<sup>[S7]</sup>  $k_{\text{r}} = \Phi_{\text{p}}/\tau_{\text{p}}$ ;  $k_{\text{ISC}} = (1 - \Phi_{\text{p}})/\tau_{\text{p}}$ ;  $k_{\text{RISC}} = \Phi_{\text{d}} \cdot (k_{\text{ISC}} \cdot \tau_{\text{p}} \cdot \tau_{\text{d}} \cdot \Phi_{\text{p}})$ , where  $\Phi_{\text{p}}$  and  $\Phi_{\text{d}}$  are fractional quantum yields for prompt fluorescence and delayed fluorescence, and  $\tau_{\text{p}}$  and  $\tau_{\text{d}}$  are emission lifetimes for prompt fluorescence and delayed fluorescence, respectively.

### 4. OLED Fabrication and Evaluation

Indium-tin-oxide (ITO)-coated glass substrates were cleaned with detergent, deionized water, acetone, and isopropanol in this order. The substrates were then subjected to UV-ozone treatment for 30 min before loading into an E-200 vacuum evaporation system (ALS Technology). Organic layers and a cathode aluminum layer were thermally evaporated on the substrates under vacuum ( $< 6 \times 10^{-5}$  Pa) with a deposition rate of  $< 0.3$  nm  $\text{s}^{-1}$  through a shadow mask, defining a pixel size of  $0.04$  cm<sup>2</sup>. The layer thickness and deposition rate were monitored *in situ* during deposition by oscillating quartz thickness monitors. The  $J$ - $V$ - $L$  characteristics of the fabricated OLEDs were measured using a 2400 source meter (Keithley) and a CS-2000 spectroradiometer (Konica Minolta). The EL efficiencies were corrected by the Lambertian factors evaluated from the angle-resolved EL data. The device lifetime tests were performed using an EAS-31D OLED lifetime evaluation system (System Engineers). The luminance of each test device was measured under the constant current driving conditions with an initial luminance of  $100$  cd  $\text{m}^{-2}$ .

## 5. Synthesis and Characterization

All chemical reactions were performed under a dry N<sub>2</sub> atmosphere using anhydrous solvents unless otherwise noted.

**Synthesis of 1:** 3,6-Di-*tert*-butylcarbazole (13.97 g, 50.0 mmol) and *t*-BuOK (5.61 g, 50.0 mmol) were dissolved in DMF (200 mL) at room temperature. After stirring for 30 min, 5-bromo-1,3-difluoro-2-iodobenzene (6.38 g, 20.0 mmol) was added to the solution. The mixture was stirred at 100 °C for 12 h. After cooling to room temperature, the reaction mixture was added into a large amount of water and then extracted with CH<sub>2</sub>Cl<sub>2</sub>. The combined organic layer was washed with water and dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>. After filtration and evaporation, the crude product was purified by silica gel column chromatography (eluent: hexane/CH<sub>2</sub>Cl<sub>2</sub> = 4:1, v/v), followed by recrystallization from CH<sub>2</sub>Cl<sub>2</sub>/methanol to afford **1** as a white solid (yield = 12.57 g, 75%). <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>): δ 8.16 (s, 4H), 7.67 (s, 2H), 7.52 (d, *J* = 8.5 Hz, 4H), 7.11 (d, *J* = 8.5 Hz, 4H), 1.47 (s, 18H). <sup>13</sup>C NMR (100 MHz, CDCl<sub>3</sub>): δ 144.70, 143.45, 138.94, 133.37, 123.93, 123.55, 123.24, 116.57, 109.49, 104.14, 34.81, 32.03. MS (MALDI-TOF): *m/z* calcd 836.22 [*M*]<sup>+</sup>; found 836.43.

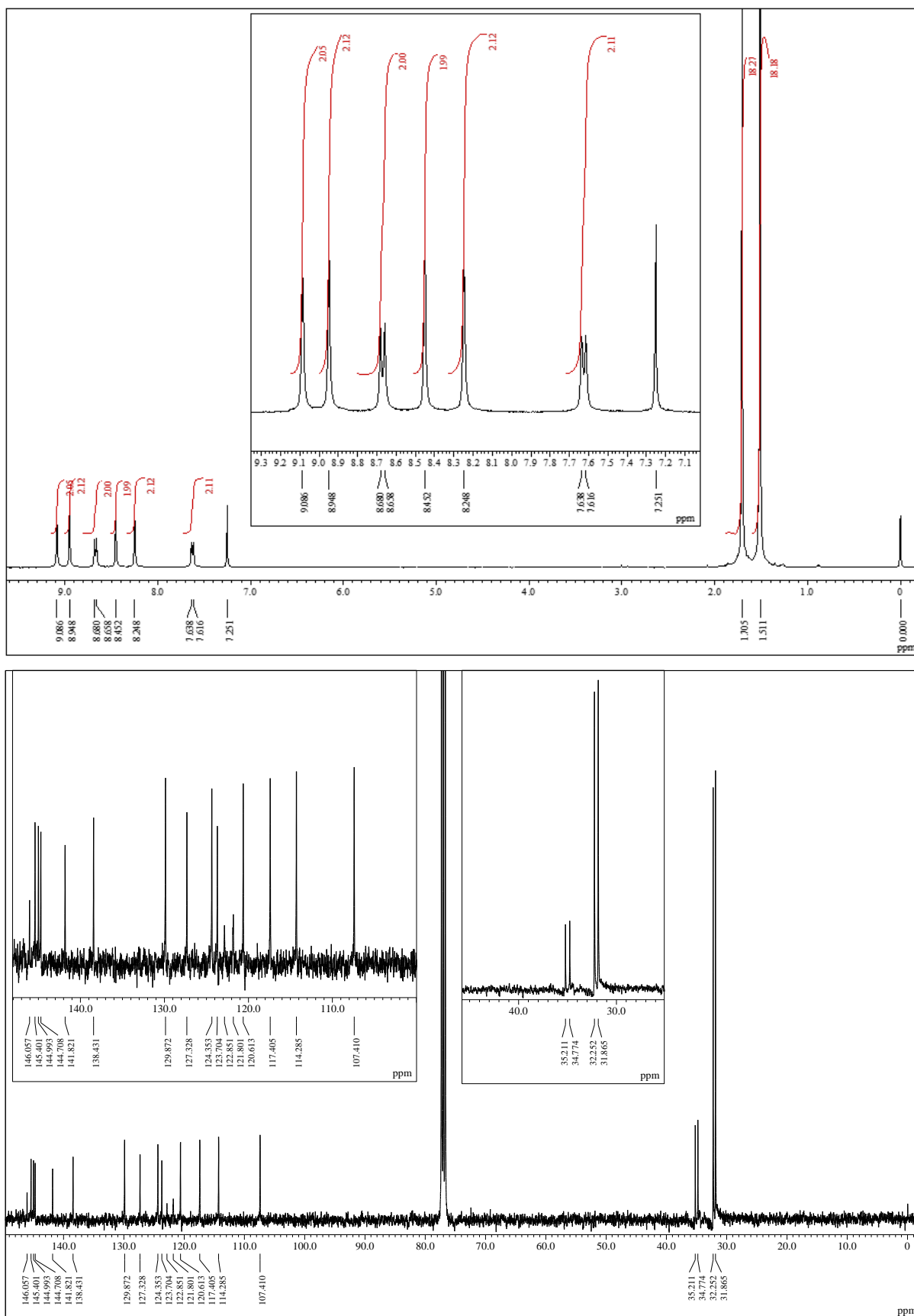
**Synthesis of 2:** To a solution of **1** (5.03 g, 6.0 mmol) in dry *t*-butylbenzene (120 mL) was slowly added *n*-BuLi (1.6 M in hexane, 4.1 mL, 6.6 mmol) at 0 °C. After reacting for 3 h at room temperature, BBr<sub>3</sub> (2.25 g, 9.0 mmol) was slowly added at 0 °C, and then the mixture was stirred at room temperature for 3 h. After addition of NEt(*i*-Pr)<sub>2</sub> (2.4 mL, 13.5 mmol) at 0 °C, the reaction mixture was further stirred at 160 °C for 24 h. After cooling to room temperature, the reaction mixture was carefully quenched by addition of water, and then extracted with toluene. The combined organic layer was dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>. After filtration and evaporation, the crude product was purified by silica gel column chromatography (eluent: hexane/CH<sub>2</sub>Cl<sub>2</sub> = 9:1, v/v), followed by recrystallization from CH<sub>2</sub>Cl<sub>2</sub>/methanol to afford **2** as a yellow solid (yield = 1.32 g, 31%). <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>): δ 9.00 (s, 2H), 8.37 (s, 2H), 8.20 (s, 2H), 8.16-8.12 (m, 4H), 7.60 (d, *J* = 8.8 Hz, 2H), 1.66 (s, 18H), 1.52 (s, 18H). <sup>13</sup>C NMR (100 MHz, CDCl<sub>3</sub>): δ 145.63, 144.88, 144.54, 141.36, 137.94, 129.70, 127.66, 127.12, 124.45, 123.67, 121.52, 120.87, 120.77, 117.22, 114.00, 110.77, 35.18, 34.80, 32.17, 31.81. MS (MALDI-TOF): *m/z* calcd 718.31 [*M*]<sup>+</sup>; found 718.38.

**Synthesis of 3:** A mixture of **2** (0.72 g, 1.0 mmol), bis(pinacolato)diboron (0.38 g, 1.5 mmol), PdCl<sub>2</sub>(dppf) (0.04 g, 0.05 mmol), and potassium acetate (0.29 g, 3.0 mmol) in dry 1,4-dioxane (15 mL) was stirred at 100 °C for 24 h. After cooling to room temperature, the reaction mixture was added into a large amount of water and then extracted with CH<sub>2</sub>Cl<sub>2</sub>. The combined organic layer was dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>. After filtration and evaporation, the crude product was purified by silica gel column chromatography (eluent: hexane/CH<sub>2</sub>Cl<sub>2</sub> = 2:1, v/v), followed by recrystallization from CH<sub>2</sub>Cl<sub>2</sub>/methanol to afford **3** as a yellow solid

(yield = 0.51 g, 67%) <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>): δ 9.14 (s, 2H), 8.79 (s, 2H), 8.53 (d, *J* = 8.8 Hz, 2H), 8.48 (s, 2H), 8.27 (s, 2H), 7.74 (d, *J* = 8.5 Hz, 2H), 1.67 (s, 18H), 1.54 (s, 18H), 1.50 (s, 12H). <sup>13</sup>C NMR (100 MHz, CDCl<sub>3</sub>): δ 145.18, 144.56, 143.69, 141.65, 138.51, 129.78, 126.97, 125.17, 124.56, 123.77, 121.68, 120.70, 119.07, 117.08, 114.45, 113.77, 84.35, 35.18, 34.82, 32.22, 31.87, 25.09. MS (MALDI-TOF): *m/z* calcd 766.48 [*M*]<sup>+</sup>; found 766.64.

**Synthesis of *p*-CzB:** A mixture of **2** (0.36 g, 0.5 mmol), **3** (0.38 g, 0.5 mmol), Pd(PPh<sub>3</sub>)<sub>4</sub> (0.008 g, 0.005 mmol), and Na<sub>2</sub>CO<sub>3</sub> (0.21 g, 2.0 mmol) in toluene (14 mL), ethanol (4 mL), and water (2 mL) was stirred at 85 °C for 48 h. After cooling to room temperature, the reaction mixture was added into a large amount of water and then extracted with CH<sub>2</sub>Cl<sub>2</sub>. The combined organic layer was dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>. After filtration and evaporation, the crude product was purified by silica gel column chromatography (eluent: hexane/CH<sub>2</sub>Cl<sub>2</sub> = 4:1, v/v), followed by recrystallization from CH<sub>2</sub>Cl<sub>2</sub>/methanol to afford ***p*-CzB** as an orange solid (yield = 0.36 g, 56%). <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>): δ 9.09 (s, 4H), 8.95 (s, 4H), 8.67 (d, *J* = 8.8 Hz, 4H), 8.45 (s, 4H), 8.25 (s, 4H), 7.63 (d, *J* = 8.8 Hz, 4H), 1.71 (s, 36H), 1.51 (s, 36H). <sup>13</sup>C NMR (100 MHz, CDCl<sub>3</sub>): δ 146.06, 145.40, 144.99, 144.71, 141.82, 138.43, 129.87, 127.33, 124.35, 123.70, 122.85, 121.80, 120.61, 117.41, 114.28, 107.41, 35.21, 34.77, 32.25, 31.87. MS (MALDI-TOF): *m/z* calcd 1278.78 [*M*]<sup>+</sup>; found 1278.87. Anal. calcd (%) for C<sub>92</sub>H<sub>96</sub>B<sub>2</sub>N<sub>4</sub>: C 86.37, H 7.56, N 4.38; found: C 86.11, H 7.56 N 4.29.

**Synthesis of *m*-CzB:** To a solution of **BBCz-SB**<sup>[S8]</sup> (1.41 g, 2.2 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (88 mL) was slowly added Cu(ClO<sub>4</sub>)<sub>2</sub>·6H<sub>2</sub>O (0.075 M in CH<sub>3</sub>CN, 44 mL, 3.3 mmol) at room temperature. The mixture was stirred for 5 h at room temperature. After quenching by the addition of aqueous K<sub>2</sub>CO<sub>3</sub> (3.3 M, 4.0 mL), the product was extracted with CHCl<sub>3</sub>. The combined organic layer was washed with water and dried over anhydrous Na<sub>2</sub>SO<sub>4</sub>. After filtration and evaporation, the crude product was purified by silica gel column chromatography (eluent: hexane/CHCl<sub>3</sub> = 1:1, v/v), preparative gel permeation chromatography (eluent: CHCl<sub>3</sub>), followed by recrystallization from CHCl<sub>3</sub>/methanol to afford ***m*-CzB** as an orange solid (yield = 0.75 g, 54%). <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>): δ 8.94 (d, *J* = 1.8 Hz, 2H), 8.76 (m, 4H), 8.72 (d, *J* = 1.8 Hz, 2H), 8.57 (d, *J* = 8.8 Hz, 2H), 8.44 (d, *J* = 1.8 Hz, 2H), 8.31 (d, *J* = 2.0 Hz, 2H), 7.75 (dd, *J* = 8.8, 2.0 Hz, 2H), 7.73 (d, *J* = 1.8 Hz, 2H), 7.09 (d, *J* = 1.5 Hz, 2H), 6.03 (d, *J* = 8.8 Hz, 2H), 5.97 (dd, *J* = 8.8, 2.0 Hz, 2H), 1.73 (s, 18H), 1.65 (s, 18H), 1.58 (s, 18H), 0.78 (s, 18H). <sup>13</sup>C NMR (100 MHz, CDCl<sub>3</sub>): δ 145.24, 144.35, 143.52, 143.41, 142.98, 142.27, 141.22, 140.56, 138.26, 137.07, 135.67, 130.45, 128.64, 127.23, 126.42, 125.71, 124.25, 124.06, 123.26, 121.75, 120.73, 120.17, 119.28, 117.36, 115.41, 113.88, 111.48, 110.60, 35.15, 35.08, 34.87, 33.93, 32.46, 32.22, 31.91, 31.22 (Two aromatic signals could not be observed probably because of the overlapping). MS (MALDI-TOF): *m/z* calcd 1278.78 [*M*]<sup>+</sup>; found 1278.72. Anal. calcd (%) for C<sub>92</sub>H<sub>96</sub>B<sub>2</sub>N<sub>4</sub>: C 86.37, H 7.56, N 4.38; found: C 86.43, H 7.60 N 4.35.



**Fig. S1** <sup>1</sup>H and <sup>13</sup>C NMR spectra of *p*-CzB in CDCl<sub>3</sub>.

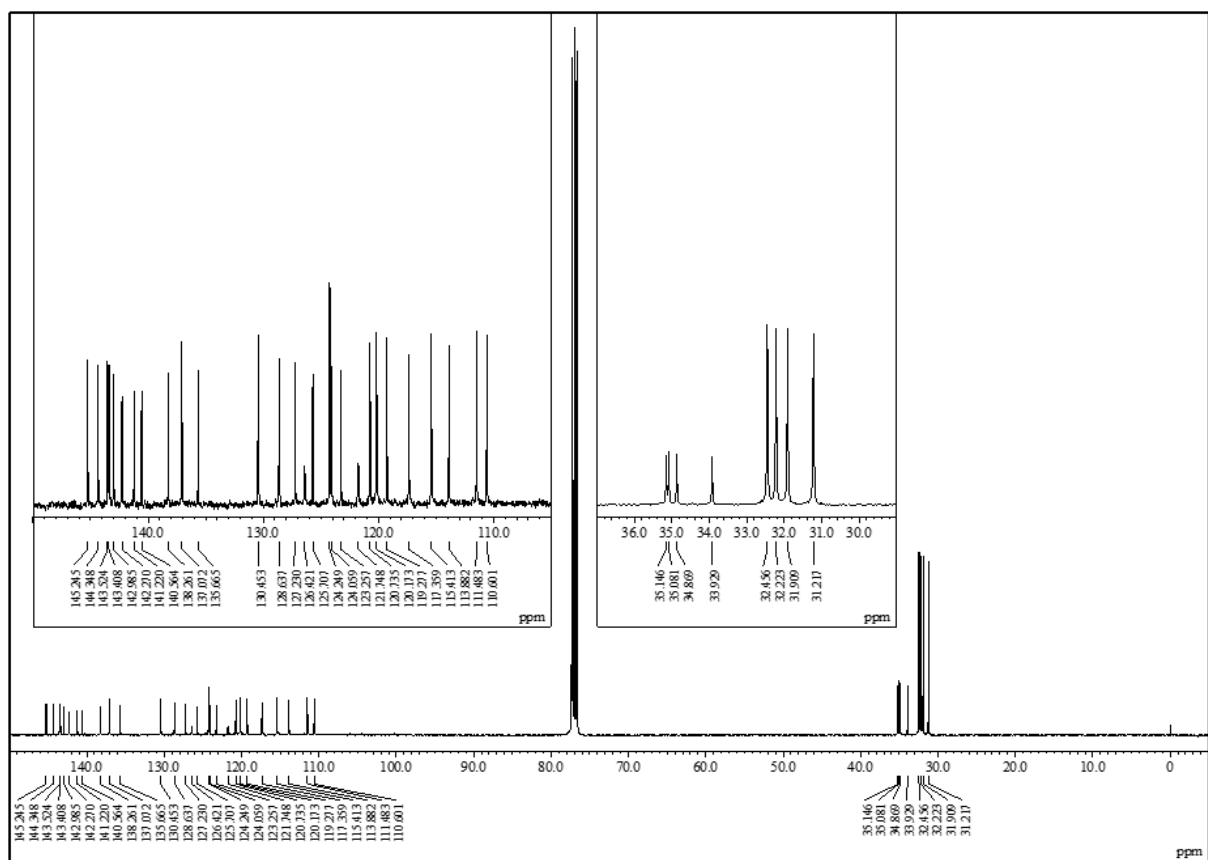
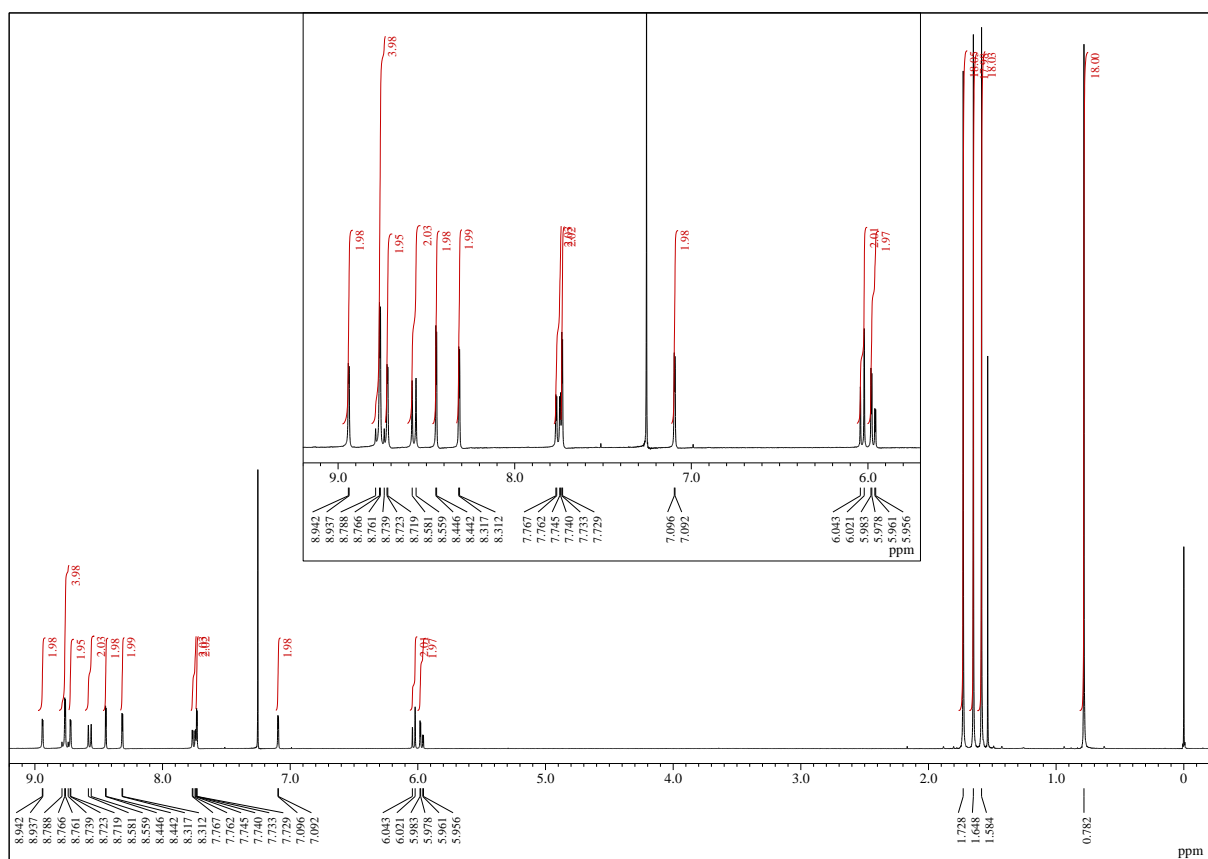
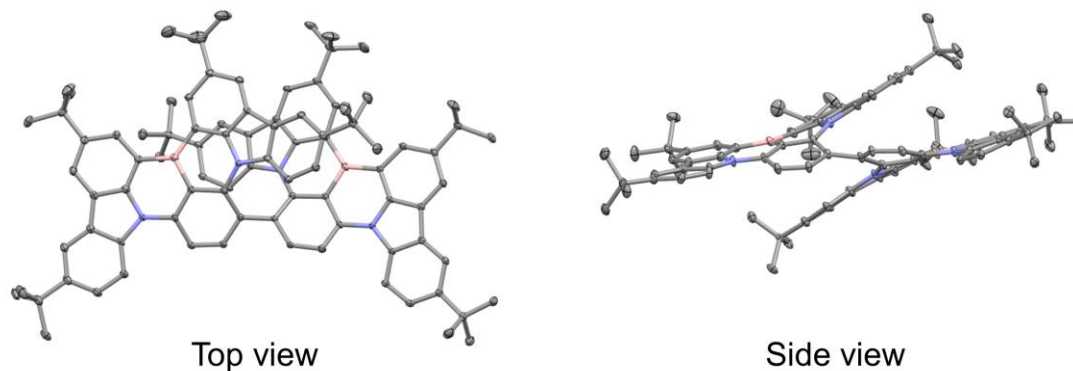
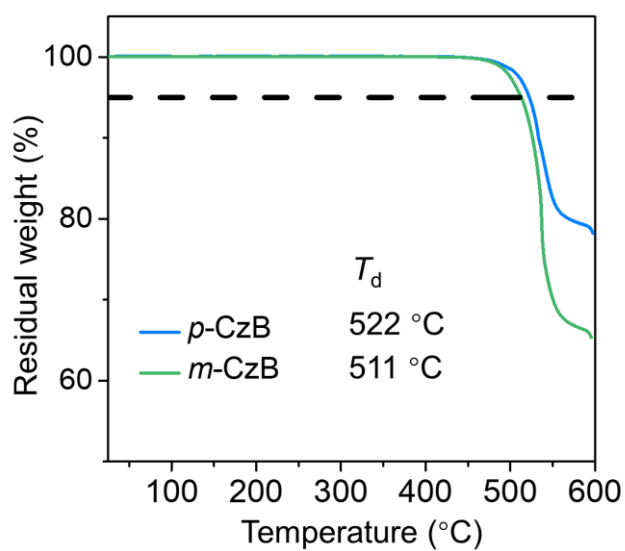


Fig. S2 <sup>1</sup>H and <sup>13</sup>C NMR spectra of *m*-CzB in CDCl<sub>3</sub>.

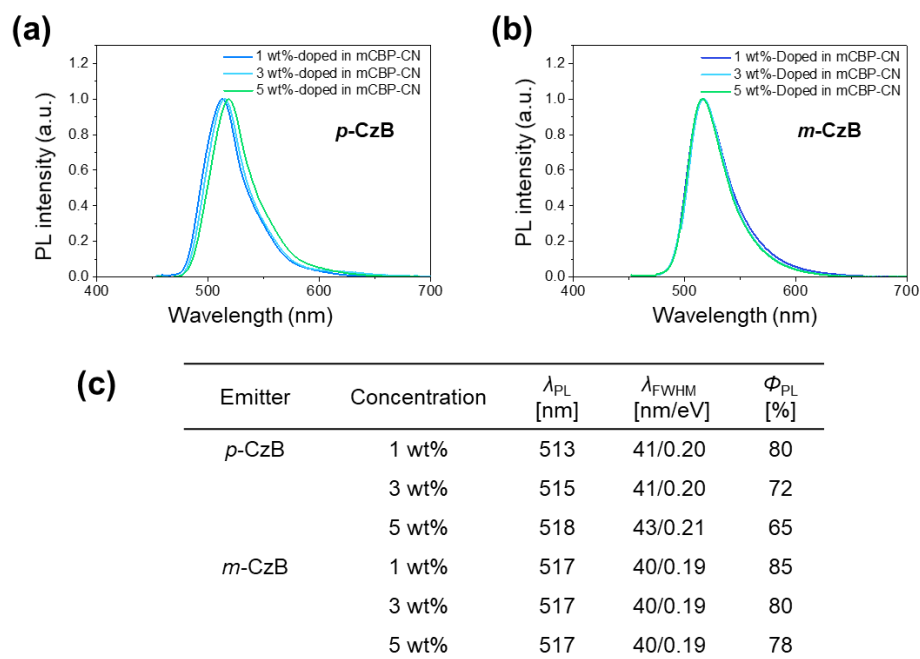


**Fig. S3** X-ray structure of *m*-CzB with thermal ellipsoids at a 50% probability. Hydrogen atoms are omitted for clarity.

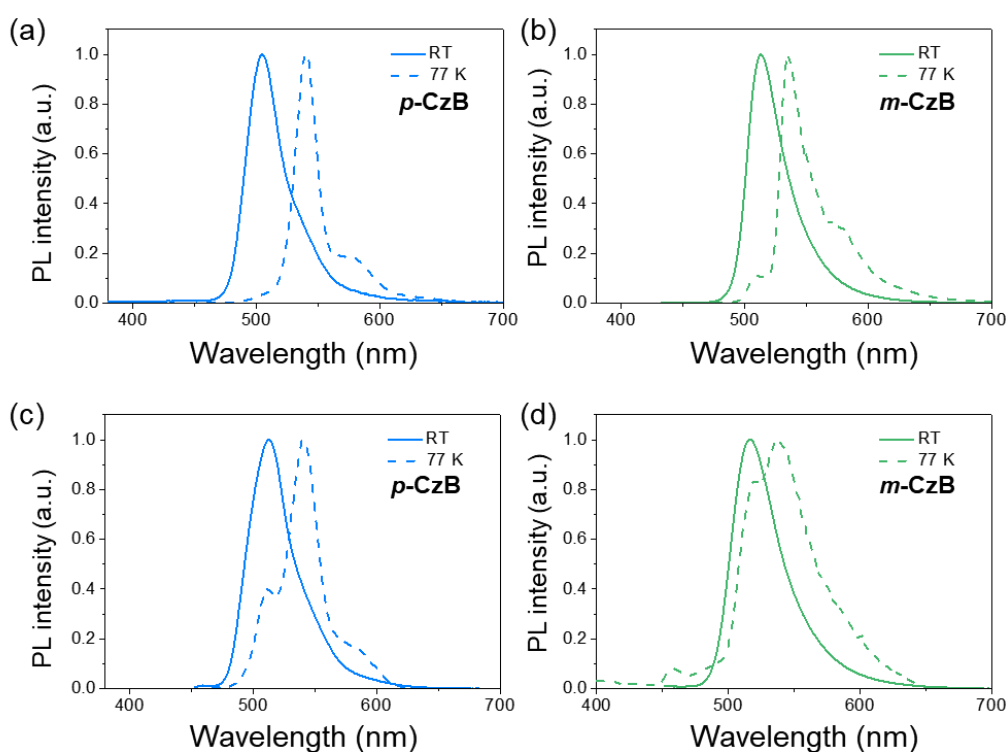


**Fig. S4** TGA thermograms of *p*-CzB and *m*-CzB recorded at a heating rate of 10 °C min<sup>-1</sup> under N<sub>2</sub> ( $T_d$  = 5% weight-loss decomposition temperature).

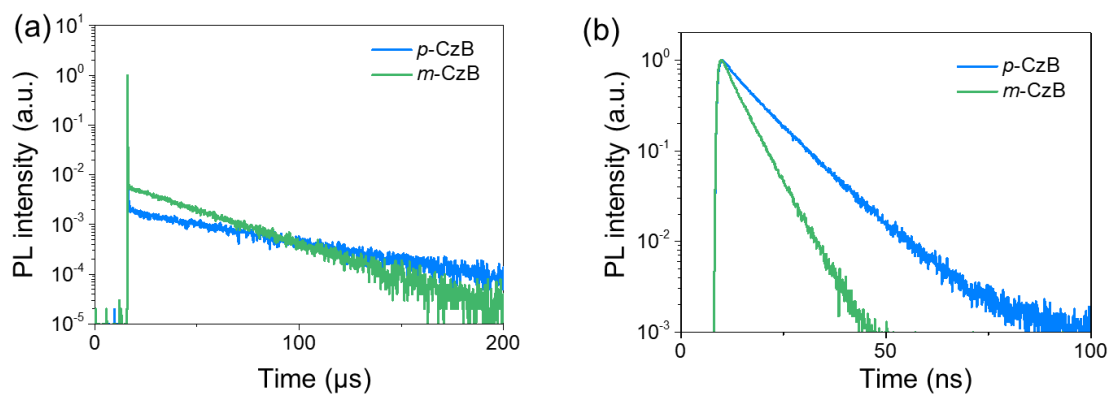




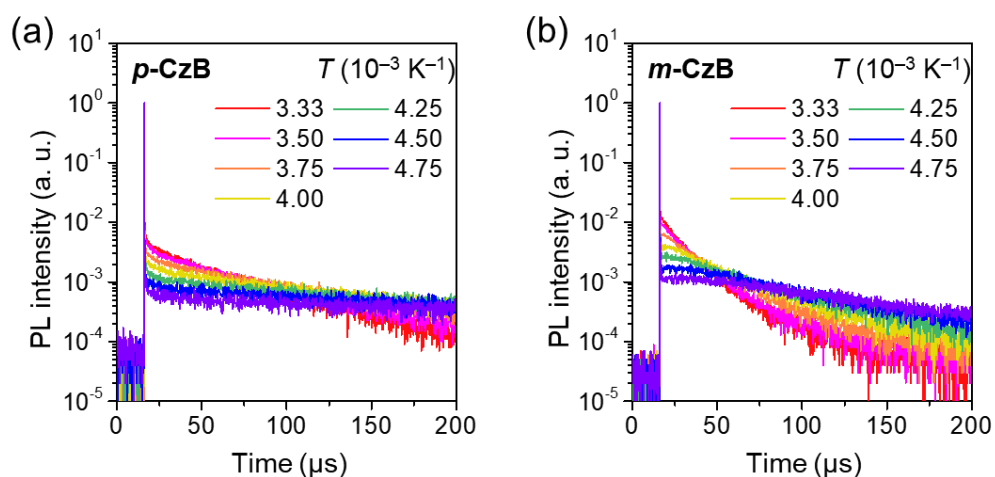
**Fig. S5** Doping concentration dependence of PL spectra for (a) *p*-CzB and (b) *m*-CzB in mCBP-CN host matrices. (c) PL data of the doped films with different doping concentrations.



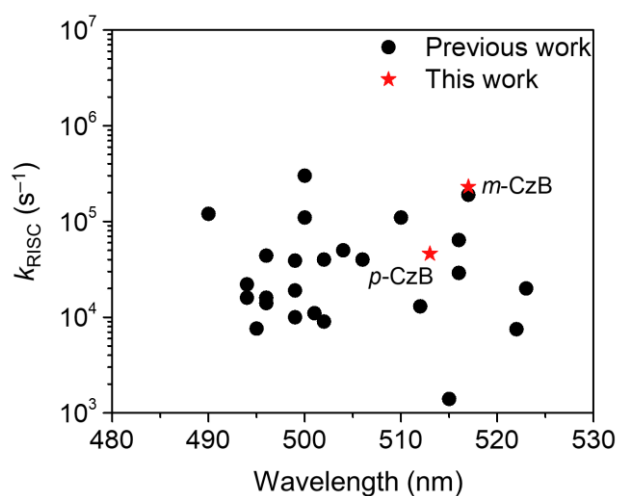
**Fig. S6** Fluorescence (300 K) and phosphorescence (77 K) spectra of (a,c) *p*-CzB and (b,d) *m*-CzB in toluene solutions ( $10^{-5}$  M, top) and in the 1 wt%-doped films in mCBP-CN (bottom).



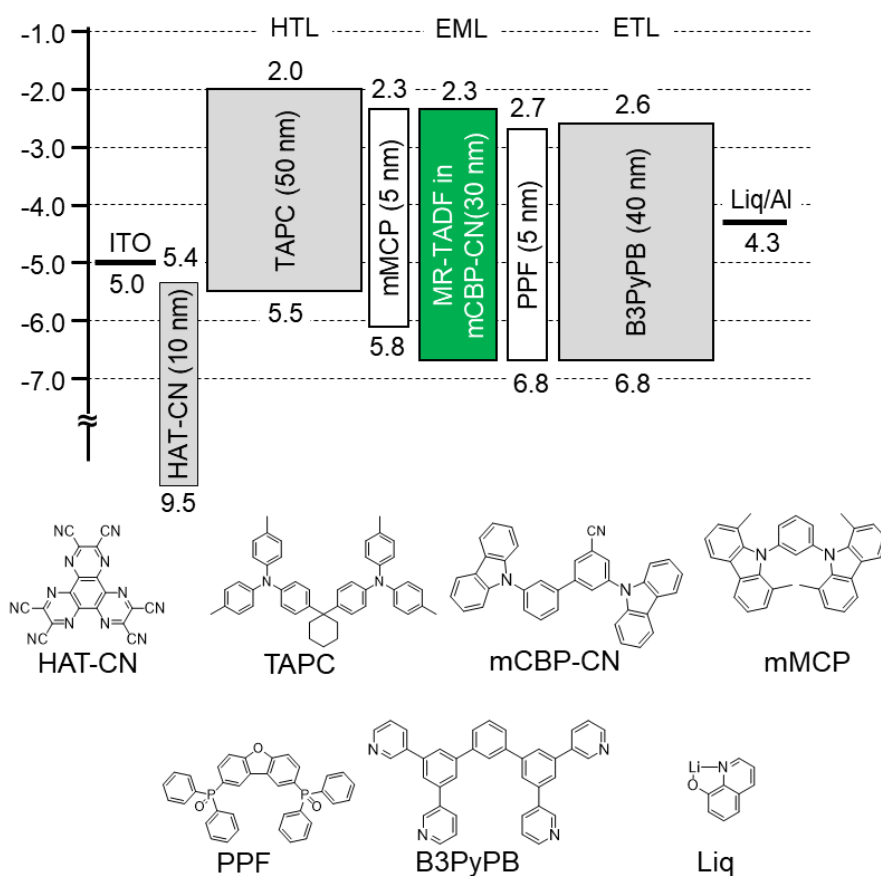
**Fig. S7** Transient PL decay curves of *p*-CzB and *m*-CzB in deoxygenated toluene solutions ( $10^{-5}$  M) measured at 300 K in (a) 0–200  $\mu$ s and (b) 0–100 ns regions.



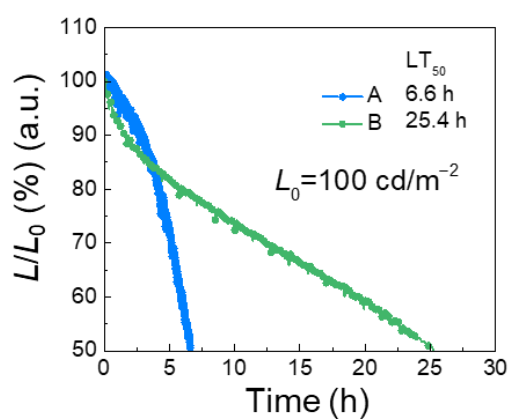
**Fig. S8** Temperature dependence of transient PL decay curves for (a) *p*-CzB and (b) *m*-CzB in 1wt% doped films in mCBP-CN host.



**Fig. S9** Comparison of  $k_{\text{RISC}}$  values for the present and reported green MR-TADF emitters with CIE- $y \geq 0.50$ .



**Fig. S10** Energy-level diagram and chemical structures of the materials used for TADF-OLEDs based on *p*-CzB and *m*-CzB.



**Fig. S11** Temporal changes of EL intensity for the TADF-OLEDs based on *p*-CzB and *m*-CzB with initial luminance at  $100 \text{ cd m}^{-2}$ .

**Table S1.** PL and EL data for green MR-TADF emitters

emitter	$\lambda_{\text{PL}}$ (nm)	$\Phi_{\text{PL}}$ (%)	$\tau_{\text{d}}$ ( $\mu\text{s}$ )	$k_{\text{RISC}}$ ( $\text{s}^{-1}$ )	$\text{EQE}_{\text{max}}$ (%)	$\text{EQE}_{1000}$ (%)	roll-off (%)	CIE(x, y)	ref
<b><i>p</i>-CzB</b>	513	80	41	$4.6 \times 10^4$	20.2	9.3	54	0.16, 0.66	this work
<b><i>m</i>-CzB</b>	517	85	15	$2.3 \times 10^5$	23.5	15.0	36	0.20, 0.70	this work
<b>2F-BN</b>	494	89	26	$2.2 \times 10^4$	22.0	15.0	32	0.16, 0.60	[S9]
<b>3F-BN</b>	499	83	17	$3.9 \times 10^4$	22.7	21.1	7	0.20, 0.58	[S9]
<b>4F-BN</b>	496	91	19	$4.4 \times 10^4$	20.9	16.4	22	0.12, 0.48	[S9]
<b>AZA-BN</b>	522	94	164	$7.5 \times 10^3$	25.7	9.0	64	0.27, 0.69	[S10]
<b>DtBuPhCzB</b>	496	90	61	-	23.4	5.7	75	0.15, 0.61	[S11]
<b>m-Cz-BNCz</b>	519	97	-	$7.5 \times 10^6$	31.4	17.5	44	0.23, 0.69	[S12]
<b>BN-CP1</b>	496	93	79	$1.6 \times 10^4$	40.0	18.5	47	0.09, 0.50	[S13]
<b>BN-CP2</b>	496	91	83	$1.4 \times 10^4$	36.4	19.2	47	0.10, 0.53	[S13]
<b>BBCz-G</b>	517	99	13	$1.9 \times 10^5$	31.8	-	-	0.26, 0.68	[S8]
<b>TCz-B</b>	512	89	71	$1.3 \times 10^4$	29.2	9.4	68	0.16, 0.71	[S14]
<b>DPTRZ</b>	521	87	787	$1.0 \times 10^3$	24.6	4.7	81	0.27, 0.60	[S15]
<b>TPTRZ</b>	501	95	83	$1.1 \times 10^4$	29.8	12.4	58	0.23, 0.68	[S15]
<b>PPm</b>	499	96	87	$1.0 \times 10^4$	28.6	10.2	64	0.17, 0.67	[S15]
<b>CNPm</b>	515	87	524	$1.4 \times 10^3$	25.0	5.7	77	0.36, 0.62	[S15]
<b>NBNP</b>	500	93	3.8	$3.0 \times 10^5$	28.0	25.6	8.0	0.09, 0.41	[S16]
<b>BN-TP</b>	523	96	44	$2.0 \times 10^4$	35.1	20.8	40	0.26, 0.70	[S17]
<b>Tph-BN</b>	495	94	62	$7.6 \times 10^3$	27.8	10.7	46	0.10, 0.46	[S18]
<b>pCz-BN</b>	496	95	89	$1.4 \times 10^4$	28.9	15.6	55	0.13, 0.54	[S18]
<b>mCz-BN</b>	494	88	95	$1.6 \times 10^4$	25.9	14.0	46	0.15, 0.55	[S18]
<b>BN1</b>	499	93	69	$1.9 \times 10^4$	17.0	8.5	50	0.15, 0.63	[S19]
<b>BN2</b>	538	89	107	$1.5 \times 10^5$	20.7	3.3	84	0.38, 0.61	[S19]
<b>BN-ICz-1</b>	516	99	342	$2.9 \times 10^4$	24.1	10.6	54	0.23, 0.72	[S20]
<b>BN-ICz-2</b>	516	98	48.6	$6.4 \times 10^4$	22.2	14.4	35	0.24, 0.73	[S20]
<b>OAB-ABP-1</b>	506	90	32	$4.0 \times 10^4$	-	-	-		[S21]
<b>PXZBN</b>	502	90	91	$0.9 \times 10^4$	23.3	11.5	51	0.22, 0.67	[S22]
<b>TPXZBN</b>	502	91	27	$0.4 \times 10^5$	21.3	17.4	18	0.16, 0.65	[S23]
<b>DPXCZBN</b>	500	90	15	$1.1 \times 10^5$	19.2	17.2	11	0.15, 0.64	[S23]
<b>BN-DPAC</b>	490	86	11.6	$1.2 \times 10^5$	23.5	16.3	31	0.14, 0.56	[S24]
<b>2PXZBN</b>	504	78	25.3	$0.5 \times 10^5$	17.1	7.4	58	0.28, 0.64	[S25]
<b>2PTZBN</b>	510	58	16.1	$1.1 \times 10^5$	25.1	17.2	33	0.28, 0.65	[S25]

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