# Synthetic Manifestation of Trinitro-Pyrazolo-2H-1,2,3Triazoles (TNPT) as Insensitive Energetic Materials 

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## SUPPORTING INFORMATION

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## General Experimental

All the reactions has beencarried out in an oven-dried round bottomed flask. Commercial grade solvents were distilled prior to use. Column chromatography has been performed using silica gel (100-200 Mesh) with hexanes and ethyl acetate mixture. Thin layer chromatography (TLC) has beenchecked on silica gel GF254 plates. The spots on TLC plate have been visualized with UV light ( 254 nm ) and/or staining over $\mathrm{I}_{2}$ chamber.
Proton and carbon nuclear magnetic resonance spectra ( ${ }^{1} \mathrm{H}$ NMR, ${ }^{13} \mathrm{C} \mathrm{NMR}$ ) are recorded on a 400 MHz $\left({ }^{13} \mathrm{C}\right.$ NMR, 101 MHz ), $500 \mathrm{MHz}\left({ }^{13} \mathrm{C}\right.$ NMR, 126 MHz$)$ spectrometer (spectra are recorded with a BRUKER-DMX-NMR) and $600 \mathrm{MHz}\left({ }^{1} \mathrm{H}\right.$ NMR, $600 \mathrm{MHz} ;{ }^{13} \mathrm{C}$ NMR, $151 \mathrm{MHz} ;{ }^{15} \mathrm{~N}$ NMR, 61 MHz ; spectra are recorded with a JEOL JNM-ECZ-600R/M1) spectrometer, having solvent resonance as internal standard ( ${ }^{1} \mathrm{H} \mathrm{NMR}: \mathrm{CDCl}_{3}$ at $7.26 \mathrm{ppm}, ~ D M S O-\mathrm{d}_{6}$ at 2.49 ppm , Acetone- $\mathrm{d}_{6}$ at 2.04 ppm , $\mathrm{CD}_{3} \mathrm{CN}$ at $1.93 \mathrm{ppm} ;{ }^{13} \mathrm{C} \mathrm{NMR:} \mathrm{CDCl}_{3}$ at $77.0 \mathrm{ppm}, ~ D M S O-d_{6}$ at 39.5 ppm , Acetone- $\mathrm{d}_{6}$ at $29.8 \& 206.5$ $\mathrm{ppm}, \mathrm{CD}_{3} \mathrm{CN}$ at $1.3 \& 118.3 \mathrm{ppm}$ ). Few cases tetramethylsilane (TMS) at 0.00 ppm has been used as reference standard. The chemical shift values (ppm) are expressed relative to the chemical shift of [deuterated] solvent or to the external standard liq. $\mathrm{NH}_{3}$ without correction $\left({ }^{15} \mathrm{~N}\right.$ NMR). Data for ${ }^{1} \mathrm{H}$ NMR are reported as follows: chemical shift ( ppm ), multiplicity ( $\mathrm{s}=$ singlet; $\mathrm{bs}=$ broad singlet; $\mathrm{d}=$ doublet; $\mathrm{bd}=$ broad doublet; $\mathrm{dd}=$ doublet of doublet; $\mathrm{dt}=$ doublet of triplet; $\mathrm{tt}=$ triplet of triplet; $\mathrm{t}=$ triplet; $\mathrm{bt}=$ broad triplet; $\mathrm{q}=$ quartet; pent $=$ pentet, $\mathrm{m}=$ multiplet), coupling constants $J$ in $(\mathrm{Hz})$, and integration. ${ }^{13} \mathrm{C}$ NMR is reported in terms of chemical shift (ppm). Melting points and decomposition temperatures are determined by DSC and TG-DSC measurements. IR spectra are recorded on FT/IR spectrometer and are reported in $\mathrm{cm}^{-1}$. High resolution mass spectra (HRMS) are obtained in ESI mode. X-ray data are collected on a 'Bruker D8 VENTURE Photon III detector' and Rigaku Oxford Hypix3000 diffractometer using Mo-K $\alpha$ radiation $(0.71073 \AA$ ) and $\mathrm{Cu}-\mathrm{K} \alpha$ radiation ( $1.54 \AA$ ).

Caution! All the TNPT and its derivatives are energetic materials and they tend to explode under certain conditions unpredictably. However, none of the compounds described herein has exploded or detonated in the course of this research. Caution should be exercised at all times during the synthesis, characterization, and handling of any of these materials, and mechanical actions involving scratching or scraping must be avoided. Ignoring safety precautions can lead to serious injuries.
Materials:Unless otherwise noted, all the reagents and intermediates are obtained commercially and used without purification. Pyrazole, 1,2,3-triazole, iodine, chloromethyl methyl ether, silver nitrate, $30 \% \mathrm{H}_{2} \mathrm{O}_{2}$, $O$-p-toluenesulfonylhydroxylamine, dimethyl sulfate, $70 \%$ perchloric acid, sodium hydride, sodium bicarbonate $\left(\mathrm{NaHCO}_{3}\right)$, tripotassium phosphate $\left(\mathrm{K}_{3} \mathrm{PO}_{4}\right)$, copper $(\mathrm{I})$ oxide $\left(\mathrm{Cu}_{2} \mathrm{O}\right)$, aqueous ammonia, THF, methanol, acetonitrile, dichloromethane, and $N, N^{\prime}$-dimethylformamide (DMF)are commercially available and used as received. Commercially available $\mathrm{H}_{2} \mathrm{SO}_{4}$ and $\mathrm{HNO}_{3}$ are used for nitration reactions.

## ExperimentalProcedures

## General procedure for the preparation of precursors 2 and 3 (GP-1):

Physical characterization data are exactly matching with the reported values of the respective compounds 2 and 3.


## 4-Iodo pyrazole (2): ${ }^{1}$



To a solution of pyrazole ( $1 ; 15.0 \mathrm{~g}, 220.58 \mathrm{mmol})$ and iodine $(27.9 \mathrm{~g}, 110.29 \mathrm{mmol})$ in water $(85.0 \mathrm{~mL})$ was added $30 \% \mathrm{H}_{2} \mathrm{O}_{2}(16.4 \mathrm{~mL}, 132.34 \mathrm{mmol})$. The reaction mixture was stirred for 12 h at room temperature. Upon reaction completion, a cold solution of $5 \%$ $\mathrm{NaHSO}_{3}(50 \mathrm{~mL})$ was added to the mixture. The compound strated precipitated providing off-white slurry. The product was filtered and washed with water to provide $2(36.5 \mathrm{~g}, 85 \%)$ as an off-white crystalline solid.
${ }^{1} \mathrm{H}$ NMR ( $600 \mathrm{MHz}, \mathrm{CDCl}_{3}$ ): $\delta 11.25(\mathrm{bs}, 1 \mathrm{H}), 7.65(\mathrm{~s}, 2 \mathrm{H}) ;{ }^{13} \mathrm{C}$ NMR ( $151 \mathrm{MHz}, \mathrm{CDCl}_{3}$ ): $\delta 138.6,56.6$ ppm.

## 4-Iodo-1-(methoxymethyl)- $\mathbf{1 H}$-pyrazole (3): ${ }^{2}$

 Sodium hydride ( $928 \mathrm{mg}, 23.2 \mathrm{mmol}, 60 \% \mathrm{wt}$ in mineral oil) was suspended in 25 mL dry THF and stirred for 15 minutes. Then a solution of 4-iodopyrazole (2; $3.0 \mathrm{~g}, 15.54$ mmol ) in 25 mL dry THF was added drop wise at $0^{\circ} \mathrm{C}$ and the resulting mixture was stirred at $40^{\circ} \mathrm{C}$ for 1 h . Chloromethyl methyl ether ( $1.29 \mathrm{~mL}, 17.01 \mathrm{mmol}$ ) was added drop wise and the mixture was allowed to stir at $50^{\circ} \mathrm{C}$ for 4 h and then at room temperature for 8 h . The mixture was quenched with saturated solution of $\mathrm{NaHCO}_{3}$ and extracted with ethyl acetate. The organic fractions were washed with brine, dried over sodium sulfate, and concentrated in vacuo to afford $\mathbf{3}$ (2.82 $\mathrm{g}, 76 \%$ ) as pale-yellow oil.
${ }^{1} \mathrm{H}$ NMR ( $600 \mathrm{MHz}, \mathrm{CDCl}_{3}$ ): $\delta 7.62(\mathrm{~s}, 1 \mathrm{H}), 7.56(\mathrm{~s}, 1 \mathrm{H}), 5.36(\mathrm{~s}, 2 \mathrm{H}), 3.31(\mathrm{~s}, 3 \mathrm{H}) ;{ }^{13} \mathrm{C}$ NMR (151 $\left.\mathrm{MHz}, \mathrm{CDCl}_{3}\right): \delta 145.0,134.1,82.1,58.1,56.9 \mathrm{ppm}$.

Table S1: Optomization for synthesis of compounds 5a and 5b


| entry | Catalyst(10 mol\%) | Base (2 equiv.) | solvent | $\operatorname{Temp}\left({ }^{\circ} \mathrm{C}\right)$ | Yield (\%) |
| :---: | :---: | :---: | :---: | :---: | :---: |
| 1 | CuI | $\mathrm{Cs}_{2} \mathrm{CO}_{3}$ | DMF | 120 | NR |
| 2 | CuCl | $\mathrm{Cs}_{2} \mathrm{CO}_{3}$ | DMF | 120 | NR |
| 3 | CuBr | $\mathrm{Cs}_{2} \mathrm{CO}_{3}$ | DMF | 120 | NR |
| 4 | CuOAc | $\mathrm{Cs}_{2} \mathrm{CO}_{3}$ | DMF | 120 | NR |
| 5 | $\begin{gathered} \mathrm{Fe}(\mathrm{acac})_{3}(30 \mathrm{~mol} \%) \\ + \\ \mathrm{Cu}(\mathrm{acac})_{2}(10 \mathrm{~mol} \%) \end{gathered}$ | $\mathrm{Cs}_{2} \mathrm{CO}_{3}$ | DMF | 120 | $\begin{aligned} & 8 \%(\mathbf{5 a}), \\ & 19 \%(\mathbf{5 b}) \end{aligned}$ |
| 6 | $\mathrm{Fe}(\mathrm{acac})_{3}+\mathrm{Cu}(\mathrm{acac})_{2}$ | $\mathrm{K}_{3} \mathrm{PO}_{4}$ | DMF | 120 | $\begin{aligned} & 12 \%(\mathbf{5 a}), \\ & 22 \%(\mathbf{5 b}) \end{aligned}$ |
| 7 | $\mathrm{Cu}_{2} \mathrm{O}$ | $\mathrm{Cs}_{2} \mathrm{CO}_{3}$ | $\mathrm{CH}_{3} \mathrm{CN}$ | 80 | NR |
| 8 | $\mathrm{Cu}_{2} \mathrm{O}$ | $\mathrm{Cs}_{2} \mathrm{CO}_{3}$ | DMF | 80 | NR |
| 9 | $\mathrm{Cu}_{2} \mathrm{O}$ | $\mathrm{Cs}_{2} \mathrm{CO}_{3}$ | DMF | 120 | $\begin{aligned} & 18 \%(\mathbf{5 a}), \\ & 32 \%(\mathbf{5 b}) \end{aligned}$ |
| 10 | $\mathrm{Cu}_{2} \mathrm{O}$ | $\mathrm{Cs}_{2} \mathrm{CO}_{3}$ | DMSO | 120 | $\begin{aligned} & 10 \%(\mathbf{5 a}), \\ & 15 \%(\mathbf{5 b}) \end{aligned}$ |

General procedure for the preparation of compounds 5 a and 5 b (GP-2): ${ }^{3}$


A mixture of $\mathrm{Cu}_{2} \mathrm{O}(0.301 \mathrm{~g}, 2.10 \mathrm{mmol}), \mathrm{K}_{3} \mathrm{PO}_{4}(8.95 \mathrm{~g}, 42.18 \mathrm{mmol}), 1,2,3$-triazole $(4 ; 2.03 \mathrm{~g}, 29.52$ mmol ), and the 4-iodo-1-(methoxymethyl)-pyrazole (3; $5.0 \mathrm{~g}, 21.09 \mathrm{mmol}$ ) in DMF ( 25 mL ) was placed in a 100 mL screw capped Schlenk tube under an argon atmosphere. The resulting mixture was
stirred at $120^{\circ} \mathrm{C}$ for 12 h . The reaction mixture was diluted with EtOAc and filtered through a small plug of silica gel. The filtrate was washed with water and the aqueous layer was extracted withEtOAc. The combined organic extracts were washed with water and brine and dried over $\mathrm{Na}_{2} \mathrm{SO}_{4}$. The solvent was evaporated under reduced pressure. The crude residue was purified by column chromatography on silica gel using hexane-ethyl acetate. Isolation using $10 \%$ ethyl acetate in hexane led to separation of $\mathbf{5 a}(700 \mathrm{mg}, 18 \%)$ and $70 \%$ ethyl acetate in hexane $\mathbf{5 b}(1.24 \mathrm{~g}, 32 \%)$ as pale-yellow liquids.

## 2-(1-(Methoxymethyl)-1H-pyrazol-4-yl)-2H-1,2,3-triazole (5a):



Following the general procedure (GP-2), compound $\mathbf{5 a}$ ( 700 mg ) was synthesized in $18 \%$ yield as pale-yellow liquid. ${ }^{1} \mathrm{H}$ NMR ( $600 \mathrm{MHz}, \mathrm{CDCl}_{3}$ ): $\delta 8.07(\mathrm{~s}, 1 \mathrm{H}), 8.03(\mathrm{~s}, 1 \mathrm{H}), 7.73$ $(\mathrm{s}, 2 \mathrm{H}), 5.41(\mathrm{~s}, 2 \mathrm{H}), 3.36(\mathrm{~s}, 3 \mathrm{H}) ;{ }^{13} \mathrm{C}$ NMR ( $151 \mathrm{MHz}, \mathrm{CDCl}_{3}$ ): $\delta 134.9,131.6,127.4$, 121.0, 82.7, 56.9 ppm; IR (Neat) $v_{\max } 3124,2998,2938,2832,1604,1518,1431,1342$, 1292, 1094, 935, 819, $750 \mathrm{~cm}^{-1}$; HRMS (ESI) $\mathrm{m} / \mathrm{z}(\mathrm{M}+\mathrm{H})^{+}$calcd for $\mathrm{C}_{7} \mathrm{H}_{10} \mathrm{~N}_{5} \mathrm{O}^{+} 180.0880$, found 180.0885.

## 1-(1-(Methoxymethyl)-1H-pyrazol-4-yl)-1H-1,2,3-triazole (5b):



Following the general procedure (GP-2), compound $\mathbf{5 b}(1.24 \mathrm{~g}$ ) was synthesized in $32 \%$ yield as pale-yellow liquid. ${ }^{1} \mathrm{H}$ NMR ( $600 \mathrm{MHz}, \mathrm{CDCl}_{3}$ ): $\delta 8.06(\mathrm{~s}, 1 \mathrm{H}), 7.87(\mathrm{~s}, 1 \mathrm{H})$, $7.85(\mathrm{~s}, 1 \mathrm{H}), 7.82(\mathrm{~s}, 1 \mathrm{H}), 5.44(\mathrm{~s}, 2 \mathrm{H}), 3.38(\mathrm{~s}, 3 \mathrm{H}) ;{ }^{13} \mathrm{C}$ NMR ( $126 \mathrm{MHz}, \mathrm{CDCl}_{3}$ ): $\delta 133.9,132.3,123.0,122.8,122.6,82.9,57.1 \mathrm{ppm}$; IR (Neat) $v_{\max } 3119,2937,1716$, 1662, 1467, 1388, 1227, 1098, 933, 852, $750 \mathrm{~cm}^{-1}$; HRMS (ESI) $\mathrm{m} / \mathrm{z}(\mathrm{M}-\mathrm{H})^{-}$calcd for $\mathrm{C}_{7} \mathrm{H}_{8} \mathrm{~N}_{5} \mathrm{O}^{-}$ 178.0734, found 178.0701.

## General procedure for the preparation of compound 6 and 7 (GP-3):

A mixture of $98 \%$ sulphuric acid and fuming nitric acid (95\%) was independently added to compound $\mathbf{5 a}$ and $\mathbf{5 b}$ at $0^{\circ} \mathbf{C}$. Later the resulting mixture was stirred at $100^{\circ} \mathrm{C}$ for $8-12 \mathrm{~h}$. The reaction progress was motinored by TLC. Upon reaction completion, the mixture was poured into ice cold water and neutralized with saturated aqueous $\mathrm{NaHCO}_{3}$ solution. After neutralization, the reaction mixture was extracted with EtOAc. The organic layer was separated and dried over $\mathrm{Na}_{2} \mathrm{SO}_{4}$. The solvent was evaporated under reduced pressure. The crude residue was purified by column chromatography on silica gel eluting with hexane: ethyl acetate to afford the desired nitration products 6 and 7. ${ }^{4}$


Sodium 3, 5-dinitro-4-(4-nitro-2H-1,2,3-triazol-2-yl)pyrazol-1-ide (6):
Following the general procedure (GP-3), compound $5 \mathbf{5}(1.00 \mathrm{~g}, 5.58 \mathrm{mmol})$ was dissolved in sulphuric acid ( 10 mL ), and fuming nitric acid ( 7 mL ) was added dropwise to the reaction mixture at $0^{\circ} \mathrm{C}$. The resulting mixture was stirred at $100{ }^{\circ} \mathrm{C}$ for 8 h . The reaction progress was motinored by TLC. Upon reaction completion, the mixture was poured into ice cold water and neutralized with saturated aqueous $\mathrm{NaHCO}_{3}$ solution. After neutralization, the reaction mixture was extracted with EtOAc. The organic layer was separated and dried over $\mathrm{Na}_{2} \mathrm{SO}_{4}$. The solvent was evaporated under reduced pressure. The crude residue was purified by column chromatography on silica gel eluting with hexane: ethyl acetate to afford the desired product $\mathbf{6}(840 \mathrm{mg}, 56 \%)$ as yellow solid. ${ }^{4}$


DSC ( $10{ }^{\circ} \mathrm{C} \mathrm{min}^{-1},{ }^{\circ} \mathrm{C}$ ): $290{ }^{\circ} \mathrm{C}\left(\mathrm{T}_{\mathrm{d}}\right) ;{ }^{1} \mathrm{H}$ NMR ( $600 \mathrm{MHz}, \mathrm{DMSO}-d_{6}$ ): $\delta 9.04$ ( s , $1 \mathrm{H}) ;{ }^{13} \mathrm{C}$ NMR ( $151 \mathrm{MHz}, \mathrm{DMSO}-d_{6}$ ): $\delta 154.2,150.2,132.8,110.0$; IR (Neat) $v_{\text {max }}$ 1621, 1523, 1447, 1359, 1325, 1074, 968, 828, $757 \mathrm{~cm}^{-1}$; HRMS (ESI) m/z (M-Na) ${ }^{-}$calcd for $\mathrm{C}_{5} \mathrm{HN}_{8} \mathrm{O}_{6}-269.0024$, found 269.0019.

## 1-(3-Nitro-1H-pyrazol-4-yl)-1H-1,2,3-triazole (7):

Following the general procedure (GP-3), compound $\mathbf{5 b}(0.6 \mathrm{~g}, 3.35 \mathrm{mmol})$ was dissolved in sulphuric $\operatorname{acid}(7.5 \mathrm{~mL})$, and fuming nitric acid $(2.5 \mathrm{~mL})$ was added dropwise to the reaction mixture at $0^{\circ} \mathrm{C}$. The resulting mixture was stirred at $100^{\circ} \mathrm{C}$ for 12 h . The reaction progress was motinored by TLC. Upon reaction completion, the mixture was poured into ice cold water and neutralized with saturated aqueous $\mathrm{NaHCO}_{3}$ solution. After neutralization, the reaction mixture was extracted with EtOAc. The organic layer was separated and dried over $\mathrm{Na}_{2} \mathrm{SO}_{4}$. The solvent was evaporated under reduced pressure. The crude residue was purified by column chromatography on silica gel eluting with hexane: ethyl acetate to afford the desired nitration product $7(136 \mathrm{mg}, 23 \%)$ as colorless solid. ${ }^{4,5}$


DSC ( $10{ }^{\circ} \mathrm{C} \mathrm{min}^{-1},{ }^{\circ} \mathrm{C}$ ): $135^{\circ} \mathrm{C}\left(\mathrm{T}_{\mathrm{m}}\right) \& 177{ }^{\circ} \mathrm{C}\left(\mathrm{T}_{\mathrm{d}}\right) ;{ }^{1} \mathrm{H}$ NMR ( $600 \mathrm{MHz}, \mathrm{CD}_{3} \mathrm{CN}$ ): $\delta 12.20(\mathrm{bs}, 1 \mathrm{H}), 8.18(\mathrm{~s}, 1 \mathrm{H}), 8.16(\mathrm{~d}, J=0.6 \mathrm{~Hz}, 1 \mathrm{H}), 7.82(\mathrm{~d}, J=1.2 \mathrm{~Hz}, 1 \mathrm{H}),{ }^{13} \mathrm{C}$

NMR (101 MHz, $\left.\mathrm{CD}_{3} \mathrm{CN}\right): \delta 134.4,130.7,128.4,115.8,79.1 \mathrm{ppm} ; \mathrm{IR}$
(Neat) $v_{\max } 3161,3118,2925,1721,1613,1528,1445,1354,1242,1121,1077,975,827,782$ $\mathrm{cm}^{-1} ;$ HRMS (ESI) $\mathrm{m} / \mathrm{z}(\mathrm{M}+\mathrm{H})^{+}$calcd for $\mathrm{C}_{5} \mathrm{H}_{5} \mathrm{~N}_{6} \mathrm{O}_{2}{ }^{+}$181.0469, found 181.0471.
Synthesis of compounds $8,9,10,11$, and 12 from sodium salt of TNPT derivative (6):


## 2-(1-Methyl-3,5-dinitro-1H-pyrazol-4-yl)-4-nitro-2H-1,2,3-triazole (8):



To a solution of compound $\mathbf{6}(520 \mathrm{mg}, 1.78 \mathrm{mmol})$ in water $(5.0 \mathrm{~mL})$ was added dimethyl sulfate ( $270 \mathrm{mg}, 2.13 \mathrm{mmol}$ ). The reaction mixture was stirred at room temperature for 8 h . The precipitate was collected by filtration and washed with distilled water to yield $\mathbf{8}$ ( $380 \mathrm{mg}, 75 \%$ ) as colorless solid. ${ }^{6}$

DSC-TGA $\left(10^{\circ} \mathrm{C} \mathrm{min}^{-1},{ }^{\circ} \mathrm{C}\right): 134^{\circ} \mathrm{C}\left(\mathrm{T}_{\mathrm{m}}\right) \& 230^{\circ} \mathrm{C}\left(\mathrm{T}_{\mathrm{d}}\right) ;{ }^{1} \mathrm{H}$ NMR $\left(600 \mathrm{MHz}, \mathrm{CD}_{3} \mathrm{CN}\right): \delta 8.66(\mathrm{~s}, 1 \mathrm{H})$, $4.36(\mathrm{~s}, 3 \mathrm{H}) ;{ }^{13} \mathrm{C}$ NMR (151 MHz, $\mathrm{CD}_{3} \mathrm{CN}$ ): $\delta 156.5,147.1,142.4,134.3(\mathrm{~d}, J=10 \mathrm{~Hz}), 113.5,44.3(\mathrm{~d}$, $J=4.4 \mathrm{~Hz}) ;{ }^{15} \mathrm{~N}$ NMR ( 61 MHz , Acetone- $d_{6}$ ): $\delta-26.1,-28.2,-31.7,-32.9,-46.5,-75.0,-151.0$, -180.3 ppm ; IR (Neat) $v_{\max } 3149,1622,1539,1468,1444,1367,1332,1228,1137,1041,961,826,756$ $\mathrm{cm}^{-1} ; \mathrm{HRMS}(\mathrm{ESI}) \mathrm{m} / \mathrm{z}(\mathrm{M}+\mathrm{Na})^{+}$calcd for $\mathrm{C}_{6} \mathrm{H}_{4} \mathrm{~N}_{8} \mathrm{NaO}_{6}{ }^{+}$307.0146, found 307.0148.

## 2-(3,5-Dinitro-1H-pyrazol-4-yl)-4-nitro-2H-1,2,3-triazole (9):



A solution of the corresponding pyrazole $\mathbf{6}(300 \mathrm{mg}, 1.0 \mathrm{mmol})$ in concentrated $\mathrm{HCl}(10 \mathrm{~mL})$ was stirred for 24 h at room temperature. The solvent was evaporated at reduced pressure and the residue was recrystallized from $\mathrm{H}_{2} \mathrm{O}$ to obtain desired product 9 ( $113 \mathrm{mg}, 41 \%$ ) as colorless solid. ${ }^{7}$
DSC ( $\left.10{ }^{\circ} \mathrm{C} \mathrm{min}^{-1},{ }^{\circ} \mathrm{C}\right): 204{ }^{\circ} \mathrm{C}\left(\mathrm{T}_{\mathrm{m}}\right) \& 261{ }^{\circ} \mathrm{C}\left(\mathrm{T}_{\mathrm{d}}\right) ;{ }^{1} \mathrm{H}$ NMR ( $600 \mathrm{MHz}, \mathrm{DMSO}-d_{6}$ ): $\delta 9.03(\mathrm{~s}, 1 \mathrm{H})$; ${ }^{13} \mathrm{C}$ NMR ( 151 MHz, DMSO- $d_{6}$ ): $\delta 154.2,150.2,132.8,110.1 ;{ }^{15} \mathrm{NNMR}\left(61 \mathrm{MHz}, \mathrm{DMSO}-d_{6}\right): \delta-19.5$, $-24.5,-32.4,-47.5,-141.9 \mathrm{ppm}$;IR (Neat) $v_{\max } 3641,3502,3148,1628,1542,1480,1453,1370,1225$, 1067, 960, 829, $759 \mathrm{~cm}^{-1}$; HRMS (ESI) m/z (M-H) ${ }^{-}$calcd for $\mathrm{C}_{5} \mathrm{HN}_{8} \mathrm{O}_{6}{ }^{-}$269.0024, found 269.0016.

## Silver (I) 3,5-dinitro-4-(4-nitro-2H-1,2,3-triazol-2-yl)pyrazol-1-ide (10):



A solution of silver nitrate $(410 \mathrm{mg}, 2.44 \mathrm{mmol})$ in distilled water $(4.0 \mathrm{~mL})$ was added dropwise to a suspension of $\mathbf{6}(600 \mathrm{mg}, 2.2 \mathrm{mmol})$ in methanol $(4.0 \mathrm{~mL})$. The reaction mixture was stirred at room temperature for 12 h , then $10(650 \mathrm{mg}$, $78 \%$ ) was collected by filtration, washed with water, and dried in vacuum. ${ }^{6}$

DSC ( $10{ }^{\circ} \mathrm{C} \mathrm{min}^{-1},{ }^{\circ} \mathrm{C}$ ): $\mathrm{T}_{\mathrm{d}}: 302{ }^{\circ} \mathrm{C} ;{ }^{1} \mathrm{H}$ NMR ( $600 \mathrm{MHz}, \mathrm{CD}_{3} \mathrm{CN}$ ): $\delta 8.58(\mathrm{~s}, 1 \mathrm{H})$;
${ }^{13} \mathrm{C}$ NMR ( $151 \mathrm{MHz}, \mathrm{CD}_{3} \mathrm{CN}$ ): $\delta 155.8,150.9,133.3,111.5 \mathrm{ppm}$; IR (Neat) $v_{\max } 1616,1523,1452,1364$, 1326, 1024, 975, 829, $756 \mathrm{~cm}^{-1}$; HRMS (ESI) m/z (M-Ag) ${ }^{-}$calcd for $\mathrm{C}_{5} \mathrm{HN}_{8} \mathrm{O}_{6}{ }^{-}$269.0024, found 269.0038.

## Ammonium 3,5-dinitro-4-(4-nitro-2H-1,2,3-triazol-2-yl)pyrazol-1-ide(11):



A solution of $10(500 \mathrm{mg}, 1.71 \mathrm{mmol})$ in methanol ( 8.0 mL ) was addedammonium chloride ( $100 \mathrm{mg}, 1.88 \mathrm{mmol}$ ). The reaction mixture was stirred at room temperature for 12 h . The suspension was filtered, washed with water, and dried in vacuum to yield $\mathbf{1 1}$ ( $370 \mathrm{mg}, \mathbf{7 5 \%}$ ) as yellow solid. ${ }^{6,8}$

DSC ( $10{ }^{\circ} \mathrm{C} \mathrm{min}^{-1},{ }^{\circ} \mathrm{C}$ ): $\mathrm{T}_{\mathrm{d}}: 268{ }^{\circ} \mathrm{C} ;{ }^{1} \mathrm{H}$ NMR ( $600 \mathrm{MHz}, \mathrm{CD}_{3} \mathrm{CN}$ ): $\delta 8.55(\mathrm{~s}, 1 \mathrm{H}), 6.47(\mathrm{bs}, 4 \mathrm{H}) ;{ }^{13} \mathrm{C}$ NMR (151 MHz, $\mathrm{CD}_{3} \mathrm{CN}$ ): $\delta 155.6,151.4,133.1,111.1 ;{ }^{15} \mathrm{NNMR}\left(61 \mathrm{MHz}, \mathrm{DMSO}-d_{6}\right): \delta-19.1,-24.4$, $-32.5,-47.7,-51.5,-141.6,-352.4 \mathrm{ppm}$; IR (Neat) $v_{\max } 3289,3247,3161,2795,1614,1523,1441$, $1408,1354,1324,1231,1155,1071,966,836,757 \mathrm{~cm}^{-1} ;$ HRMS (ESI) m/z (M+H) ${ }^{+}$calcd for $\mathrm{C}_{5} \mathrm{H}_{6} \mathrm{~N}_{9} \mathrm{O}_{6}{ }^{+}$ 288.0436, found 288.0432.

## 3,5-Dinitro-4-(4-nitro-2H-1,2,3-triazol-2-yl)-1H-pyrazol-1-amine (12):



To a freshly prepared pulverized ethyl $O$-p-tolylsulphonylacetohydroximate (1.34 $\mathrm{g}, 5.22 \mathrm{mmol})$ was added $70 \%$ perchloric acid $(12.5 \mathrm{~mL})$ at room temperature and stirred at ambient conditions for 2 h . The tosylhydroxylamine suspension was poured into 80 mL of ice/water slurry.The mixture was extracted with dichloromethane $(3 \times 20 \mathrm{~mL})$. The combined dichloromethane extracts were dried over sodium sulfate and filtered. Next, a solution of $\mathbf{1 1}(1.0 \mathrm{~g}, 3.48 \mathrm{mmol})$ in acetonitrile ( 50 mL ) was added dropwise to the freshly prepraredtosylhydroxylamine in dichloromethane. The resulting mixture was stirred overnight under ambient conditions, evaporated to dryness, and resuspended in ethyl acetate.

The suspension was filtered and evaporated, and purified by silica gel chromatography eluting with hexane-ethyl acetate ( $70: 30$ )to afford $\mathbf{1 2}$ ( $655 \mathrm{mg}, 67 \%$ ) as pink solid. ${ }^{9}$


DSC $\left(10^{\circ} \mathrm{C} \mathrm{min}^{-1},{ }^{\circ} \mathrm{C}\right): 185{ }^{\circ} \mathrm{C}\left(\mathrm{T}_{\mathrm{m}}\right) \& 280^{\circ} \mathrm{C}\left(\mathrm{T}_{\mathrm{d}}\right),{ }^{1} \mathrm{H}$ NMR ( 600 MHz , Acetone- $\left.\mathrm{d}_{6}\right): \delta 8.88(\mathrm{~s}, 1 \mathrm{H})$, $7.91(\mathrm{~s}, 2 \mathrm{H}),{ }^{13} \mathrm{C}$ NMR ( 151 MHz , Acetone- $\mathrm{d}_{6}$ ): $\delta 156.2,143.0,137.9,134.1,111.8 ;{ }^{15} \mathrm{~N}$ NMR ( 61 MHz , $\mathrm{CD}_{3} \mathrm{CN}$ ): $\delta-30.4,-32.9,-36.4,-37.9,-51.1,-88.8,-155.9,-168.6,-286.9 \mathrm{ppm} ; \mathrm{IR}$ (Neat) $v_{\max } 3363$, 3277 , 1620, 1484, 1370, 1334, 1311, 1233, 1128, 1059, 968, 827, $758 \mathrm{~cm}^{-1} ;$ HRMS (ESI) m/z (M+Na) ${ }^{+}$ calcd for $\mathrm{C}_{5} \mathrm{H}_{3} \mathrm{~N}_{9} \mathrm{NaO}_{6}{ }^{+}$308.0099, found 308.0078.

## X-ray crystallography: ${ }^{10}$

Single crystal X-ray data for the compounds $\mathbf{6 , 7 , 9}, \mathbf{1 1}$, were collected using 'Bruker D8 VENTURE Photon III detector' system [6, 7, 11 using $\lambda(\mathrm{Mo}-\mathrm{K} \alpha)=0.71073 \AA ; 9$ using $\lambda(\mathrm{Cu}-\mathrm{K} \alpha)=1.54 \AA$ ] at 108 K, $104 \mathrm{~K}, 296 \mathrm{~K}, 107 \mathrm{~K}$, and 8, 12 were collected using Rigaku Oxford Hypix-3000 system [6, 7 using $\lambda(\mathrm{Mo}-\mathrm{K} \alpha)=0.71073 \AA \AA$; at $298 \mathrm{~K}, 298 \mathrm{~K}$, graphite monochromator with a $\omega$ scan. Data reduction was performed using Bruker SAINT ${ }^{2}$ software. Intensities for absorption were corrected using SADABS 2014/5. Structure solution and refinement were carried out using Bruker SHELXTL/ Olex2 1.3 version. All non-hydrogen atoms were refined anisotropically. Thermal ellipsoid plots of all the compounds $\mathbf{6}, \mathbf{7}, \mathbf{8}, \mathbf{9}, \mathbf{1 1}$ and $\mathbf{1 2}$ with $50 \%$ probability and hydrogen atoms are unlabelled for clarity shown (Manuscript Figure 3).

CCDC-2241660 (6), CCDC-2241851 (7), CCDC-2289671 (8), CCDC-2241662 (9), CCDC-2241663 (11), and CCDC-2289806 (12) contain the supplementary crystallographic data for this paper. These data can be obtained free of charge from the Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/data_request/cif.

Table S2 Crystallographic data for compounds 6, 7, and $\mathbf{8}$

| Compound | 6 | 7 | 8 |
| :---: | :---: | :---: | :---: |
| CCDC | 2241660 | 2241851 | 2289671 |
| Formula | C5 H3 N8 Na O7 | $\mathrm{C}_{5} \mathrm{H}_{4} \mathrm{~N}_{6} \mathrm{O}_{2}$ | $\mathrm{C}_{6} \mathrm{H}_{4} \mathrm{~N}_{8} \mathrm{O}_{6}$ |
| $\mathbf{M w}_{\mathbf{w}}$ | 310.14 | 180.14 | 284.17 |
| Crystal system | Monoclinic | Monoclinic | Monoclinic |
| Space group | C2/c | P21 | P21/n |
| T [K] | 108 K | 104 K | 298 K |
| a [ $\AA$ ] | 11.4210 (11) | 6.6812 (10) | 9.1552(5) |
| b [ $\AA$ ] | 9.3158 (10) | 7.0171 (11) | 8.1515(4) |
| c [ $\AA$ ] | 21.599 (2) | 7.8472 (12) | 15.7139(7) |
| $\alpha\left[{ }^{\circ}\right]$ | 90 | 90 | 90 |
| $\left.\boldsymbol{\beta}{ }^{\circ}{ }^{\circ}\right]$ | 96.498 (4) | 104.571 (5) | 106.884(5) |
| $\gamma\left[{ }^{\circ}\right]$ | 90 | 90 | 90 |
| Z | 8 | 2 | 4 |
| $\mathrm{V}\left[\AA^{\mathbf{3}}\right]$ | 2283.3 (4) | 356.07 (9) | 1122.16(10) |
| D calc $^{\text {[g/cc] }}$ | 1.804 | 1.68 | 1.682 |
| $\mu\left[\mathrm{mm}^{-1}\right]$ | 0.196 | 0.136 | 0.151 |
| Total reflns | 2838 | 2082 | - |
| Unique reflns | 2829 | 1941 | 2347 |
| Observed reflns | 2553 | 1924 | 1974 |
| $R_{1}[I>2 \sigma(I)]$ | 0.0290 | 0.0303 | 0.0427 |
| $w R_{2}$ [all] | 0.0789 | 0.0804 | 0.1329 |
| GOF | 1.060 | 1.064 | 1.065 |
| Diffractometer | Bruker D8 <br> VENTURE <br> Photon III <br> detector | Bruker D8 <br> VENTURE <br> Photon III <br> detector | Rigaku Oxford Hypix-3000 detector |

Table S3 Crystallographic data for compounds 9,11 , and $\mathbf{1 2}$

| Compound | 9 | 11 | 12 |
| :---: | :---: | :---: | :---: |
| CCDC | 2241662 | 2241663 | 2289806 |
| Formula | $\mathrm{C}_{5} \mathrm{H}_{2} \mathrm{~N}_{8} \mathrm{O}_{6}$ | $\mathrm{C}_{5} \mathrm{H}_{5} \mathrm{~N}_{9} \mathrm{O}_{6}$ | $\mathrm{C}_{5} \mathrm{H}_{3} \mathrm{~N}_{9} \mathrm{O}_{6}$ |
| $\mathbf{M w}_{\mathbf{w}}$ | 270.15 | 287.18 | 285.16 |
| Crystal system | Monoclinic | Triclinic | Monoclinic |
| Space group | P21/c | $P-1$ | $P 21 / n$ |
| T [K] | 296 K | 107 K | 298 K |
| a [ $\AA$ ] | 11.9516 (2) | 5.1838 (8) | 8.8502(4) |
| b [ $\AA$ ] | 8.6036 (2) | 9.1275 (12) | 7.3093(4) |
| c [ $\AA$ ] | 10.7241 (2) | 12.3907 (18) | 16.7588(7) |
| $\alpha\left[{ }^{\circ}\right]$ | 90 | 107.311 (4) | 90 |
| $\boldsymbol{\beta}{ }^{\circ}{ }^{\circ}$ | 104.501 (1) | 99.537 (5) | 98.440(4) |
| $\gamma\left[{ }^{\circ}\right]$ | 90 | 102.736 (4) | 90 |
| Z | 4 | 2 | 4 |
| $\mathrm{V}\left[\AA^{\mathbf{3}}\right]$ | 1067.6 (4) | 528.83 (13) | 1072.37(9) |
| $D_{\text {calc }}[\mathrm{g} / \mathrm{cc}]$ | 1.68 | 1.804 | 1.766 |
| $\mu\left[\mathrm{mm}^{-1}\right]$ | 1.371 | 0.163 | 0.161 |
| Total reflns | - | 4021 | - |
| Unique reflns | 1926 | 3999 | 2204 |
| Observed reflns | 1779 | 3603 | 1129 |
| $R_{1}[I>2 \sigma(I)]$ | 0.0352 | 0.0328 | 0.0686 |
| $w R_{2}$ [all] | 0.0993 | 0.0904 | 0.2247 |
| GOF | 1.057 | 1.051 | 1.011 |
| Diffractometer | Bruker D8 <br> VENTURE <br> Photon III detector | Bruker D8 <br> VENTURE <br> Photon III <br> detector | Rigaku Oxford Нуріх-3000 detector |



Figure S1: Ball and stick packing diagram of compound 6 viewed down the b axis. Dashed lines indicate hydrogen bonding.


Figure S2: Ball and stick packing diagram of compound 7 viewed down the C axis. Dashed lines indicate hydrogen bonding.


Figure S3: Ball and stick packing diagram of compound $\mathbf{8}$ viewed down the b axis. Dashed lines indicate hydrogen bonding.


Figure S4: Ball and stick packing diagram of compound 9 viewed down the a axis. Dashed lines indicate hydrogen bonding.


Figure S5: Ball and stick packing diagram of compound $\mathbf{1 1}$ viewed down the a axis. Dashed lines indicate hydrogen bonding.


Figure S6: Ball and stick packing diagram of compound $\mathbf{1 2}$ viewed down the $b$ axis. Dashed lines indicate hydrogen bonding.

In most cases, the X-Ray crystallography data of the energetic compounds are collected at low temperature. The strong intermolecular interactions in the crystal unit cell affect the X-ray diffraction [see Effect of Strong Intermolecular Interaction in 2D Inorganic Molecular Crystals, Xin Feng, Xingliang Peng, Baixin Peng, Zexin Li, Wentao Huang, Sijie Yang, Ke Pei, Zongdong Sun, Fuqiang Huang, Huiqiao Li, Zhigang Shuai, and Tianyou Zhai, Journal of the American Chemical Society 2021, 143 (48), 2019220201]. Thus, most of the X-ray crystal data are collected in different temperature; we report the best diffraction results. Whereas X-ray data for compound $\mathbf{8}$ and $\mathbf{1 2}$ are collected at 298 K .
The Hirschfield surfaces analysis of compound $\mathbf{1 2}$ was independently examined at low temperature and room temperature as shown in Table 1. We did not see significant variation of atomic interactions (listed in the table) in the crystal packing. We therefore believe that the thermal expansion is not greatly affected with respect to temperature.
Table 1:

| S. No | 2D fingerprint plots of compound 12 at low <br> temperature |  |  | 2D fingerprint plots of compound 12 <br> at room temperature |
| :--- | :--- | :---: | :--- | :---: |
|  | Atomic interactions | Percentage (\%) | Atomic <br> interactions | Percentage (\%) |
|  | $\mathrm{O} \ldots \mathrm{H} / \mathrm{H} \ldots \mathrm{O}$ | 27.2 | $\mathrm{O} \ldots \mathrm{H} / \mathrm{H} \ldots \mathrm{O}$ | 25.9 |
| 2 | $\mathrm{~N} \ldots \mathrm{H} / \mathrm{H} \ldots \mathrm{N}$ | 9.7 | $\mathrm{~N} \ldots \mathrm{H} / \mathrm{H} \ldots \mathrm{N}$ | 6.7 |
| 3 | $\mathrm{~N} \ldots \mathrm{O} / \mathrm{O} \ldots \mathrm{N}$ | 22.7 | $\mathrm{~N} \ldots \mathrm{O} / \mathrm{O} \ldots \mathrm{N}$ | 23.3 |
| 4 | $\mathrm{C} \ldots \mathrm{O} / \mathrm{O} \ldots \mathrm{C}$ | 11.7 | $\mathrm{C} \ldots \mathrm{O} / \mathrm{O} \ldots \mathrm{C}$ | 11.6 |
| 5 | $\mathrm{C} \ldots \mathrm{N} / \mathrm{N} \ldots \mathrm{C}$ | 4.6 | $\mathrm{C} \ldots \mathrm{N} / \mathrm{N} \ldots \mathrm{C}$ | 4.6 |
| 6 | $\mathrm{O} \ldots \mathrm{O}$ | 17.7 | $\mathrm{O} \ldots \mathrm{O}$ | 19.4 |

## Hirshfeld Surface Analysis ${ }^{11}$

The Hirshfeld surface image shown in Figure4 (in the manuscript body), the red spots signify the high contact populations, while blue and white spots are for low contact populations. This suggests that the negative (red) or positive value (blue and white) of $\mathrm{d}_{\text {norm }}$ depends on the intermolecular contacts being shorter (red) or longer (blue and white) than the Vander Waals separations. For each point on the Hirshfeld surface, the normalized contact distance ( $\mathrm{d}_{\text {norm }}$ ) was determined by the equation shown below.

$$
\left[\mathrm{d}_{\text {norm }}=\left(\mathrm{d}_{\mathrm{i}}-\mathrm{d}_{\mathrm{i}}^{\mathrm{vdW}}\right) / \mathrm{r}_{\mathrm{i}}^{\mathrm{vdW}}+\left(\mathrm{d}_{\mathrm{e}}-\mathrm{d}_{\mathrm{e}}^{\mathrm{vdW}} / \mathrm{r}_{\mathrm{e}}^{\mathrm{vdW}}\right]\right.
$$

In which, $d_{i}$ is measured from the surface to the nearest atom interior to the surface interior, while $d_{e}$ is measured from the surface to the nearest atom exterior to the surface interior, where $\mathrm{r}_{\mathrm{i}}{ }^{\mathrm{vdW}}$ andre $_{\mathrm{e}}{ }^{\mathrm{vdW}}$ are the Van der Waals radii of the atoms. Hirshfeld surface graphs and two-dimensional fingerprint plots of $\mathbf{8}, \mathbf{9}, \mathbf{1 1}$, and $\mathbf{1 2}$ were analyzed using Crystalexplorer17.5 softwareand Hirshfeld surface calculation of 11as shown in Figure S12 (Hirshfeldplots of 8, 9, and 12are shown inFigure 4, manuscript body).


Figure S7: Hirshfeld surface calculation of 11 as well as two-dimensional finger print plot in the crystal structure. Hirshfeld surface graph (centre of pie graph) with proximity of close contacts around 11 molecules (white, $\mathrm{d}=$ Van der Waals ( vdW ) distance; blue, $\mathrm{d}>\mathrm{vdW}$ distance; red, $\mathrm{d}<\mathrm{vdW}$ distance). The pie graph (b) for $\mathbf{1 1}$ and show percentage contributions of the individual atomic contacts to the Hirshfeld surface.

Hirschfield surface and the related 2D fingerprint plots mainly reveals the intermolecular interactions [such as hydrogen bonding (HB), and $\pi$-stacking] of the energetic crystals; these interactions generally determine the packing structure and further the properties and performances of an energetic compound. In general, the large population of $\mathrm{O} . . \mathrm{H} / \mathrm{H} . . . \mathrm{O}$ and $\mathrm{N} . . \mathrm{H} / \mathrm{H} . . . \mathrm{N}$ interactions makes the molecule more stable and therefore is less sensitive. Accordingly, compound $\mathbf{8}$ ( N -methyl trinitro-pyrazolo-triazole; IS: 40 J ) and $\mathbf{1 2}$ ( N -amino trinitro-pyrazolo-triazole; IS: 35 J ) are relatively less sensitive than trinitro-pyrazolo-triazole (TNPT) (IS: 35 J ) due to more O...H/H...O and N...H/H...N interactions in the crystal packing \{see: Hirshfeld Surface Method and Its Application in Energetic Crystals, Shijie Li, Rupeng Bu, Rui-jun Gou, and Chaoyang Zhang, Crystal Growth \& Design, 2021, 21 (12), 6619-6634].

Isodesmic reactions for the prediction of heat of formation: ${ }^{12}$



## Impact Sensitivity Test Results



Table S4

| Weight | Height | IS | $\mathbf{1}$ | $\mathbf{2}$ | $\mathbf{3}$ | $\mathbf{4}$ | $\mathbf{5}$ | $\mathbf{6}$ | Result |
| :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- |
| 5 kg | 100 cm | 50 J | Negative | Positive | Positive |  |  |  | 2 Out of 3 |
|  | 90 cm | 45 J | Negative | Negative | Negative | Positive | Negative | Positive | 1 Out of 6 |
|  | 80 cm | 40 J | Negative | Negative | Negative | Negative | Negative | Negative | 0 Out of 6 |

Final Result: Impact Sensitivity of 2-(1-Methyl-3,5-dinitro-1H-pyrazol-4-yl)-4-nitro-2H-1,2,3-

$$
\text { triazole }=40 \mathrm{~J}
$$



Table S5

| Weight | Height | IS | $\mathbf{1}$ | $\mathbf{2}$ | $\mathbf{3}$ | $\mathbf{4}$ | $\mathbf{5}$ | $\mathbf{6}$ | Result |
| :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- |
| 5 kg | 100 cm | 50 J | Negative | Positive | Negative | Positive |  |  | 2 Out of 4 |
|  | 80 cm | 40 J | Positive | Negative | Negative | Negative | Positive |  | 1 Out of 4 |
|  | 70 cm | 35 J | Negative | Negative | Negative | Negative | Negative | Negative | 0 Out of 6 |

Final Result: Impact Sensitivity of 2-(3,5-Dinitro-1H-pyrazol-4-yl)-4-nitro-2H-1,2,3-triazole= 35 J


Table S6

| Weight | Height | IS | $\mathbf{1}$ | $\mathbf{2}$ | $\mathbf{3}$ | $\mathbf{4}$ | $\mathbf{5}$ | $\mathbf{6}$ | Result |
| :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- |
| 5 kg | 100 cm | 50 J | Positive | Positive |  |  |  |  | 2 Out of 2 |
|  | 80 cm | 40 J | Negative | Positive | Negative | Negative | Positive |  | 2 Out of 5 |
|  | 70 cm | 35 J | Negative | Negative | Negative | Negative | Negative | Positive | 1 Out of 6 |
|  | 60 cm | 30 J | Negative | Negative | Negative | Negative | Negative | Negative | 0 Out of 6 |

Final Result: Impact Sensitivity of Ammonium 3,5-dinitro-4-(4-nitro-2H-1,2,3-triazol-2-yl)pyrazol-1-ide= 30 J


12
Table S7

| Weight | Height | IS | $\mathbf{1}$ | $\mathbf{2}$ | $\mathbf{3}$ | $\mathbf{4}$ | $\mathbf{5}$ | $\mathbf{6}$ | Result |
| :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- |
| 5 kg | 100 cm | 50 J | Negative | Positive | Positive |  |  |  | 2 Out of 3 |
|  | 80 cm | 40 J | Positive | Negative | Negative | Negative | Positive |  | 2 Out of 5 |
|  | 70 cm | 35 J | Negative | Negative | Negative | Negative | Negative | Negative | 0 Out of 6 |

Final Result: Impact Sensitivity of 3,5-Dinitro-4-(4-nitro-2H-1,2,3-triazol-2-yl)-1H-pyrazol-1-amine = 35 J

## Friction Sensitivity Results



Table S8

| Load | Notch <br> No. | FS | 1 | 2 | 3 | 4 | 5 | 6 | Result |
| :--- | :---: | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- |
| B9 | 6 | 360 N | Negative | Negative | Negative | Negative | Negative | Negative | 0 Out of 6 |

Final Result: Friction Sensitivity of 2-(1-Methyl-3,5-dinitro-1H-pyrazol-4-yl)-4-nitro-2H-1,2,3-triazole = 360 N


Table S9

| Load | Notch <br> No. | FS | $\mathbf{1}$ | $\mathbf{2}$ | $\mathbf{3}$ | $\mathbf{4}$ | $\mathbf{5}$ | $\mathbf{6}$ | Result |
| :--- | :---: | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- |
| B9 | 5 | 324 N | Negative | Positive | Positive |  |  |  | 2 Out of 3 |
| B9 | 4 | 288 N | Negative | Negative | Positive | Negative | Positive |  | 2 Out of 5 |
| B9 | 3 | 252 N | Negative | Negative | Negative | Negative | Positive |  | 1 Out of 5 |
| B9 | 2 | 216 N | Negative | Negative | Negative | Negative | Negative | Negative | 0 out of 6 |

Final Result: Friction Sensitivity of 2-(3,5-Dinitro-1H-pyrazol-4-yl)-4-nitro-2H-1,2,3-triazole= 216 N


Table S10

| Load | Notch <br> No. | FS | 1 | 2 | 3 | 4 | 5 | 6 | Result |
| :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :---: |
| B9 | 5 | 324 N | Negative | Positive | Negative | Negative | Negative | Negative | 1 Out of 6 |
| B9 | 4 | 288 N | Negative | Negative | Negative | Negative | Negative | Negative | 0 Out of 6 |

Final Result: Friction Sensitivity of Ammonium 3,5-dinitro-4-(4-nitro-2H-1,2,3-triazol-2-yl) pyrazol-1-ide= 288 N


Table S11

| Load | Notch <br> No. | FS | $\mathbf{1}$ | $\mathbf{2}$ | $\mathbf{3}$ | $\mathbf{4}$ | $\mathbf{5}$ | $\mathbf{6}$ | Result |
| :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- |
| B9 | 5 | 324 N | Positive | Positive |  |  |  |  | 2 Out of 2 |
| B9 | 4 | 288 N | Negative | Positive | Negative | Positive |  |  | 2 Out of 4 |
| B9 | 3 | 252 N | Negative | Negative | Positive | Negative | Negative | Negative | 1 Out of6 |
| B9 | 2 | 216 N | Negative | Negative | Negative | Negative | Negative | Positive | 1 out of 6 |
| B9 | 1 | 180 N | Negative | Negative | Negative | Negative | Negative | Negative | 0 out of 6 |

Final Result: Friction Sensitivity of 3,5-Dinitro-4-(4-nitro-2H-1,2,3-triazol-2-yl)-1H-pyrazol-1-amine= 180 N

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DSC






























| 気 | MBR-5-110-15N-R_TVK-PC-34-DNP-3S-15N_TVK-NA-TRNPY-1S 1_single_pulse_dec-1-2.jdf <br>  |  | ```= MBR-5-110-15N-R_TV = delta = single pulse_dec.j = MBR-5-110-15 = DMSO-D6 = 9-JAN-2021 17:46: = 6-JAN-2023 10:09: = single pulse decou = 1D REAL = 52429 = Nitrogen15 = Nitrogen15 = [ppm] = X = ACRHEM_UOH = JNM-ECZ = 14.09636928[T] = 0.85983232[s] = Nitrogen15 = 60.81520929[MHz] = 214.12266[ppm] = 32768 = 4 = 1.45377182[Hz] = 38.1097561[kHz] = 30.48780488[kHz] = Proton = 600.1723046[MHz] = 5[ppm] = 5[us] = TRUE = 10000 = 10000 = 15[s] = 56 = 19.4[dC] = 24.21[us] = 0.68786586[s] = 30[deg] = 9.9[dB] = 8.07[us] = 33.452[dB] = 33.452[dB] = 33.452[dB] = 7.23684211[kHz] = 12.05794078[ppm] = 600.1723046[MHz] = 2.2 = TRUE = FALSE = WALTZ = 5 [ppm] = 76[us] = 76[us] = 76[us] = FAISE = 0``` |
| :---: | :---: | :---: | :---: |





## XEVO-G2XSQTOF\#YEA1155

17012019_22 -ve 16 (0.194) AM2 (Ar,22000.0,554.26,0.00,LS 10); Cm (16:23-114:161)
(100
270.0032


## Display Report

Analysis Info
Analysis Name
Method
Sample Name MBR-1-141 +VE
Comment


MBR-1-141 +VE.d


UNIVERSITY OF HYDERABAD

| $12072022 \_05-V E ~$ | $39(0.450)$ |
| :---: | :---: |
| 100 | AM2 (Ar,22000.0,554.26,0.00,LS 10); Cm (39:48-73:89) |
| 269.0016 | 1:TOF MS ES- |
| $6.33 e 7$ |  |





## Display Report

Analysis Info
Analysis Name
Method TL-P.m
Sample Name MPG-112
Comment

## Acquisition Parameter

| Acquisition Parameter |  |  |  |  |
| :--- | :--- | :--- | :--- | :--- |
| Source Type | ESI | Ion Polarity | Positive | Set Nebulizer |
| Focus | Not active | Set Capillary | 4000 V | 0.3 Bar |
| Scan Begin | $100 \mathrm{~m} / \mathrm{z}$ | Set End Plate Offset | -500 V | Set Dry Heater |
| Scan End | $2000 \mathrm{~m} / \mathrm{z}$ | Set Charging Voltage | 0 V | Set Dry Gas |
|  |  | Set Corona | 0 nA | Set Divert Valve |
|  |  |  | Set APCI Heater | $0^{\circ} \mathrm{C}$ |



MPG-112.d

