

Electronic Supplementary Information

Solvent effect on the structures of three manganese complexes based on azotetrazolyl-3-hydroxy-2-naphthoic acid: syntheses, structures, magnetic and fluorescent properties

Yun-Jing Zhong,[§] Zhi-Jian Ouyang,[§] Xiao-Man Kuang, You-Hong Li, Wen-Bin Chen,* Meng Yang* and Wen Dong*

Guangzhou Key Laboratory for Environmentally Functional Materials and
Technology, School of Chemistry and Chemical Engineering, Guangzhou University,
Guangzhou 510006, P. R. China

Contents of the Electronic Supplementary Information

Table S1 Selected bond lengths [\AA] and angles [$^\circ$] for complex 1	3
Table S2 Selected bond lengths [\AA] and angles [$^\circ$] for complex 2	3
Table S3 Selected bond lengths [\AA] and angles [$^\circ$] for complex 3	4
Figure S1 Local coordination geometry of Mn(II) ion in complex 1	6
Table S4 SHAPE analysis for complex 1	6
Figure S2 Local coordination geometry of Mn(II) ions in complex 2	6
Table S5 SHAPE analysis for complex 2	7
Figure S3 Local coordination geometry of Mn(II) ions in complex 3	7
Table S6 SHAPE analysis for complex 3	7
Figure S4 the size of CH ₃ OH, DMF and DMSO.....	8
Figure S5 Powder X-ray diffraction (PXRD) patterns for complexes 1-3	8
Figure S6 TG curves of complex 1 under N ₂ atmosphere.	9
Figure S7 <i>M</i> versus <i>H</i> plot at 2 K for complex 2	9
Figure S8 Temperature dependence of the in-phase and out-phase in zero field for complex 2	9
Figure S9 <i>M</i> versus <i>H</i> plot at 2 K for complex 3	10
Figure S10 Temperature dependence of the in-phase and out-phase in zero field for complex 3	10
Figure S11 Magnetic exchange pathways in complexes 1-3	11

Table S1 Selected bond lengths [Å] and angles [°] for complex **1**.

<i>Bond distances</i>			
Mn(1)–N(1)	2.085(2)	Mn(1)–N(2)	2.103(2)
Mn(1)–N(3a)	2.100(2)	Mn(1)–O(1)	2.0554(18)
Mn(1)–O(2)	2.0881(17)	Mn(1)–O(3)	2.1397(19)
<i>Angles</i>			
N(1)–Mn(1)–N(2)	75.45(8)	N(1)–Mn(1)–N(3a)	100.47(8)
N(1)–Mn(1)–O(1)	153.53(8)	N(1)–Mn(1)–O(2)	92.32(8)
N(1)–Mn(1)–O(3)	92.20(8)	N(2)–Mn(1)–N(3a)	175.81(8)
N(2)–Mn(1)–O(1)	78.20(8)	N(2)–Mn(1)–O(2)	88.02(8)
N(2)–Mn(1)–O(3)	88.10(8)	N(3a)–Mn(1)–O(1)	105.91(8)
N(3a)–Mn(1)–O(2)	91.18(8)	N(3i)–Mn(1)–O(3)	93.10(8)
O(1)–Mn(1)–O(2)	89.34(7)	O(1)–Mn(1)–O(3)	84.31(7)
O(2)–Mn(1)–O(3)	173.12(7)		

^a 1-x, 1-y, -z**Table S2** Selected bond lengths [Å] and angles [°] for complex **2**.

<i>Bond distances</i>			
Mn(1)–N(1)	2.311(3)	Mn(1)–N(2)	2.430(3)
Mn(1)–O(1)	2.230(3)	Mn(1)–O(1a)	2.252(2)
Mn(1)–O(2)	2.181(3)	Mn(1)–O(3)	2.161(3)
Mn(1)–O(4a)	2.242(2)	Mn(2)–N(3)	2.153(3)
Mn(2)–N(3c)	2.153(3)	Mn(2)–O(4a)	2.241(2)
Mn(2)–O(4b)	2.241(2)	Mn(2)–O(5a)	2.199(3)
Mn(2)–O(5b)	2.199(3)		
<i>Angles</i>			
N(1)–Mn(1)–N(2)	65.35(10)	N(1)–Mn(1)–O(1a)	147.89(10)
N(1)–Mn(1)–O(1)	132.96(10)	N(1)–Mn(1)–O(2)	88.68(11)

N(1)-Mn(1)-O(3)	93.24(13)	N(1)-Mn(1)-O(4a)	146.41(10)
N(2)-Mn(1)-O(1)	67.96(9)	N(2)-Mn(1)-O(1a)	86.71(9)
N(2)-Mn(1)-O(2)	91.49(11)	N(2)-Mn(1)-O(3)	86.11(12)
N(2)-Mn(1)-O(4a)	137.40(10)	O(1)-Mn(1)-O(1a)	79.15(9)
O(1)-Mn(1)-O(2)	87.09(10)	O(1)-Mn(1)-O(3)	89.02(12)
O(1)-Mn(1)-O(4a)	154.61(9)	O(1a)-Mn(1)-O(2)	93.55(10)
O(1a)-Mn(1)-O(3)	86.70(12)	O(1a)-Mn(1)-O(4a)	75.80(9)
O(2)-Mn(1)-O(3)	175.98(12)	O(2)-Mn(1)-O(4a)	90.38(10)
O(3)-Mn(1)-O(4a)	93.57(12)	N(3)-Mn(2)-N(3c)	91.50(19)
N(3)-Mn(2)-O(4a)	82.09(11)	N(3)-Mn(2)-O(4b)	123.83(11)
N(3)-Mn(2)-O(5a)	138.87(11)	N(3)-Mn(2)-O(5b)	99.64(12)
N(3c)-Mn(2)-O(4a)	123.83(11)	N(3c)-Mn(2)-O(4b)	82.09(11)
N(3c)-Mn(2)-O(5a)	99.64(12)	N(3c)-Mn(2)-O(5b)	138.87(11)
O(4a)-Mn(2)-O(4b)	145.05(14)	O(4a)-Mn(2)-O(5a)	58.81(9)
O(4a)-Mn(2)-O(5b)	96.99(10)	O(4b)-Mn(2)-O(5a)	96.99(10)
O(4b)-Mn(2)-O(5b)	58.81(9)	O(5a)-Mn(2)-O(5b)	97.44(16)

^a1/2-x, 1/2-y, 1-z; ^b1/2+x, 1/2-y, 1/2+z; ^c1-x, +y, 3/2-z

Table S3 Selected bond lengths [Å] and angles [°] for complex **3**.

<i>Bond distances</i>			
Mn(1)–O(1)	2.292(3)	Mn(1)–O(2)	2.189(4)
Mn(1)–O(3)	2.183(4)	Mn(1)–O(4)	2.169(4)
Mn(1)–O(1b)	2.235(3)	Mn(1)–N(1b)	2.332(4)
Mn(1)–N(6b)	2.408(4)	Mn(2)–O(4)	2.443(4)
Mn(2)–O(5)	2.192(4)	Mn(2)–O(6)	2.293(4)
Mn(2)–O(7)	2.220(5)	Mn(2)–O(8)	2.315(3)
Mn(2)–N(2b)	2.278(5)	Mn(2)–N(8a)	2.283(5)
Mn(3)–O(8)	2.256(4)	Mn(3)–O(9)	2.195(3)
Mn(3)–O(9a)	2.243(4)	Mn(3)–O(10)	2.180(4)

Mn(3)–O(11)	2.204(4)	Mn(3)–N(7a)	2.280(4)
Mn(3)–N(12a)	2.407(4)		
<i>Angles</i>			
O(1)-Mn(1)-O(2)	93.17(15)	O(1)-Mn(1)-O(3)	95.34(15)
O(1)-Mn(1)-O(4)	77.41(13)	O(1)-Mn(1)-O(1b)	76.12 (14)
O(1)-Mn(1)-N(1b)	149.52(14)	O(1)-Mn(1)-N(6b)	144.55(14)
O(2)-Mn(1)-O(3)	171.46(16)	O(2)-Mn(1)-O(4)	88.07(17)
O(2)-Mn(1)-O(1b)	85.57(16)	O(2)-Mn(1)-N(1b)	85.57(16)
O(2)-Mn(1)-N(6b)	89.37(16)	O(3)-Mn(1)-O(4)	93.05(18)
O(3)-Mn(1)-O(1b)	91.87(15)	O(3)-Mn(1)-N(1b)	86.71(16)
O(3)-Mn(1)-N(6b)	84.13(16)	O(4)-Mn(1)-O(1b)	153.40(14)
O(4)-Mn(1)-N(1b)	72.12(15)	O(4)-Mn(1)-N(6b)	138.03(15)
O(1b)-Mn(1)-N(1b)	134.30(14)	O(1b)-Mn(1)-N(6b)	68.49(14)
N(1b)-Mn(1)-N(6b)	65.92(15)	O(4)-Mn(2)-O(5)	54.98(14)
O(4)-Mn(2)-O(6)	140.79(14)	O(4)-Mn(2)-O(7)	90.3(2)
O(4)-Mn(2)-O(8)	161.52(13)	O(4)-Mn(2)-N(2b)	76.64(14)
O(4)-Mn(2)-N(8a)	85.57(16)	O(5)-Mn(2)-O(6)	86.46(15)
O(5)-Mn(2)-O(7)	90.7(2)	O(5)-Mn(2)-O(8)	139.99(15)
O(5)-Mn(2)-N(2b)	131.40(15)	O(5)-Mn(2)-N(8a)	95.84(18)
O(6)-Mn(2)-O(7)	82.63(19)	O(6)-Mn(2)-O(8)	56.92(13)
O(6)-Mn(2)-N(2b)	140.14(16)	O(6)-Mn(2)-N(8a)	107.42(17)
O(7)-Mn(2)-O(8)	99.06(19)	O(7)-Mn(2)-N(2b)	84.70(19)
O(7)-Mn(2)-N(8a)	168.29(19)	O(8)-Mn(2)-N(2b)	88.31(14)
O(8)-Mn(2)-N(8a)	82.13(14)	N(2b)-Mn(2)-N(8a)	83.69(17)
O(8)-Mn(3)-O(9)	78.48(13)	O(8)-Mn(3)-O(9a)	149.33(13)
O(8)-Mn(3)-O(10)	80.86(14)	O(8)-Mn(3)-O(11)	91.76(15)
O(8)-Mn(3)-N(7a)	75.87(14)	O(8)-Mn(3)-N(12a)	140.04(13)
O(9)-Mn(3)-O(9a)	73.96(14)	O(9)-Mn(3)-O(10)	99.16(15)
O(9)-Mn(3)-O(11)	84.18(16)	O(9)-Mn(3)-N(7a)	150.34(15)

O(9)-Mn(3)-N(12a)	141.19(15)	O(9a)-Mn(3)-O(10)	90.41(15)
O(9a)-Mn(3)-O(11)	98.50(16)	O(9a)-Mn(3)-N(7a)	150.34(15)
O(9a)-Mn(3)-N(12a)	67.56(13)	O(10)-Mn(3)-O(11)	171.06(16)
O(10)-Mn(3)-N(7a)	91.26(15)	O(10)-Mn(3)-N(12a)	86.24(15)
O(11)-Mn(3)-N(7a)	81.98(16)	O(11)-Mn(3)-N(12a)	96.38(16)
N(7a)-Mn(3)-N(12a)	66.73(15)		

^a1-x, 1-y, -z; ^b1-x, -y, 1-z

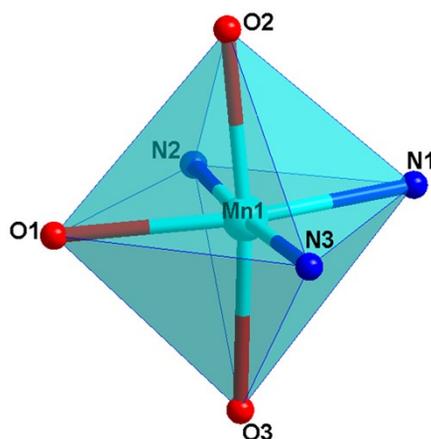


Figure S1 Local coordination geometry of Mn(II) ion in complex **1**.

Table S4 SHAPE analysis for complex **1**.

Complex	M	PPY-6	OC-6	TPR-6	JPPY-6
1	Mn1	20.007	1.282	12.881	28.716

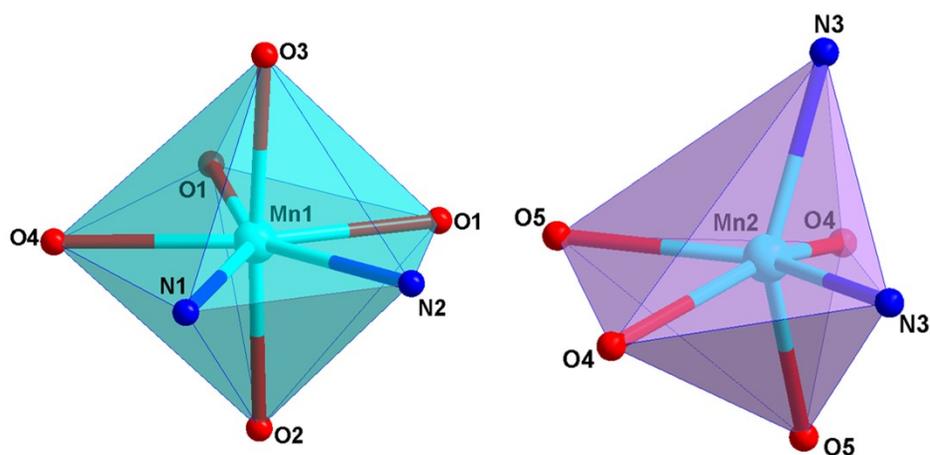


Figure S2 Local coordination geometry of Mn(II) ions in complex **2**.

Table S5 SHAPE analysis for complex 2.

Complex	M	PBPY-7	COC-7	CTRR-7	JPBPY
2	Mn 1	0.415	7.055	5.235	3.147
	Mn 2	20.007	1.282	12.881	28.716

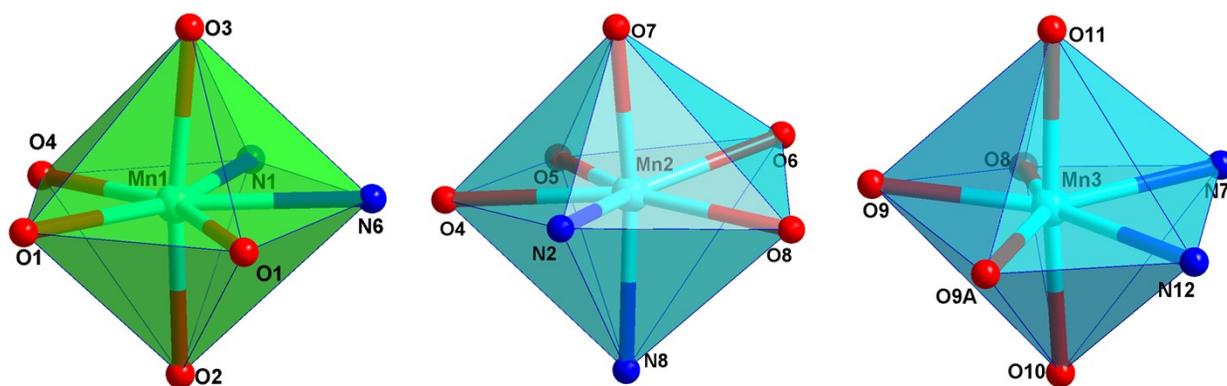


Figure S3 Local coordination geometry of Mn(II) ions in complex 3.

Table S6 SHAPE analysis for complex 3.

Complex	M	PBPY-7	COC-7	CTRR-7	JPBPY-7
3	Mn 1	0.261	7.880	5.982	3.094
	Mn 2	2.652	5.882	4.656	5.916
	Mn 3	0.989	6.902	5.046	3.983

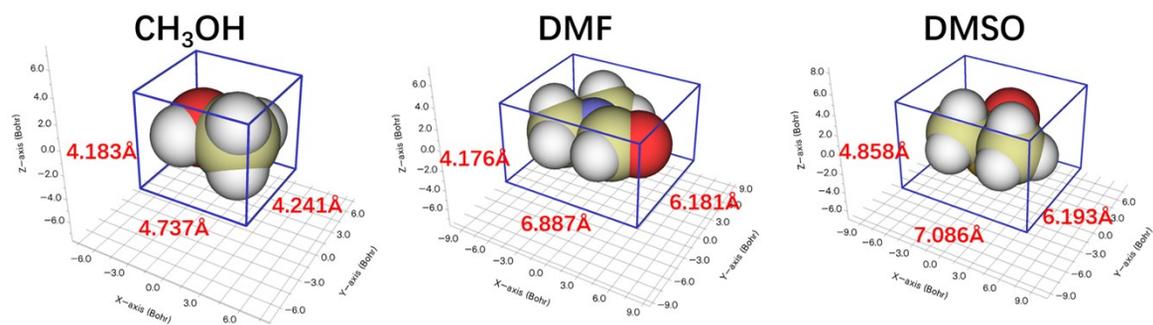


Figure S4 The size of CH₃OH, DMF and DMSO.

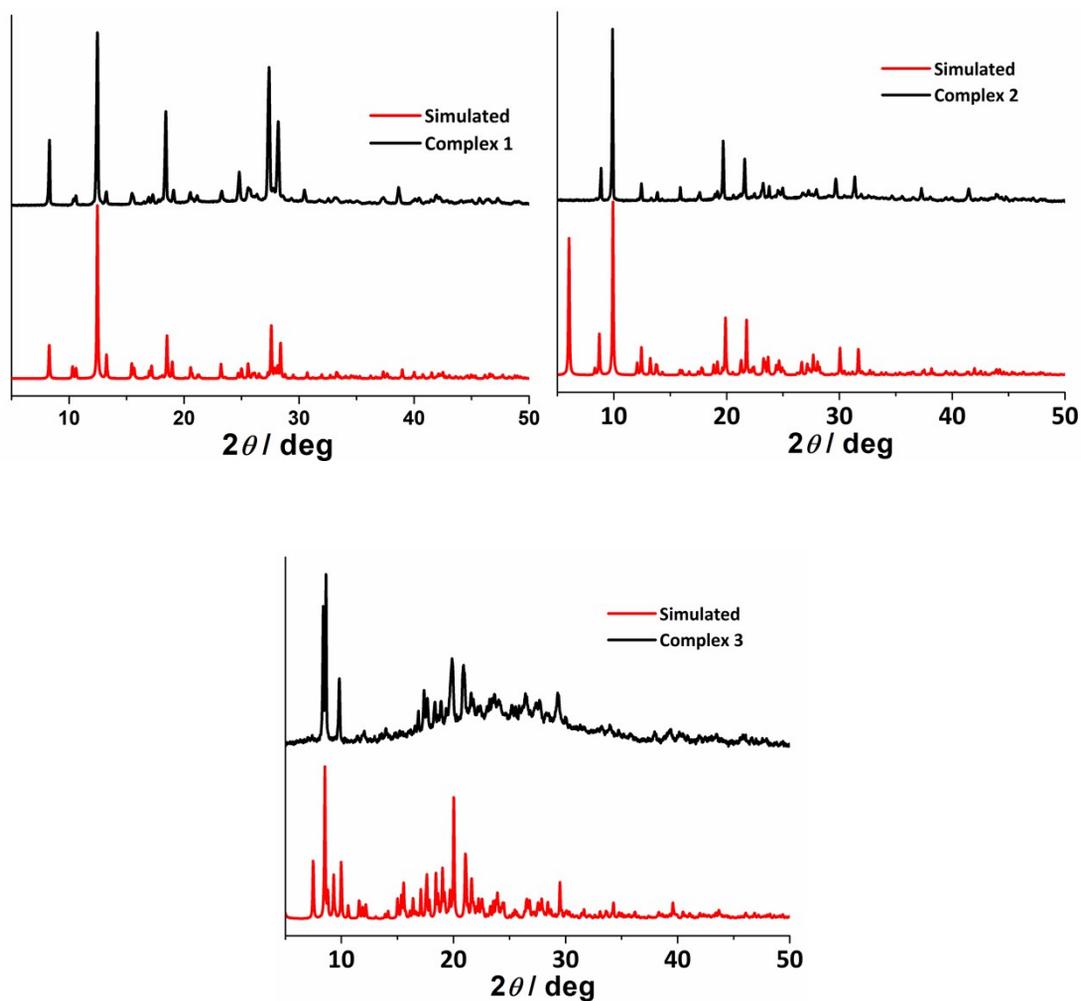


Figure S5 Powder X-ray diffraction (PXRD) patterns for complexes 1-3.

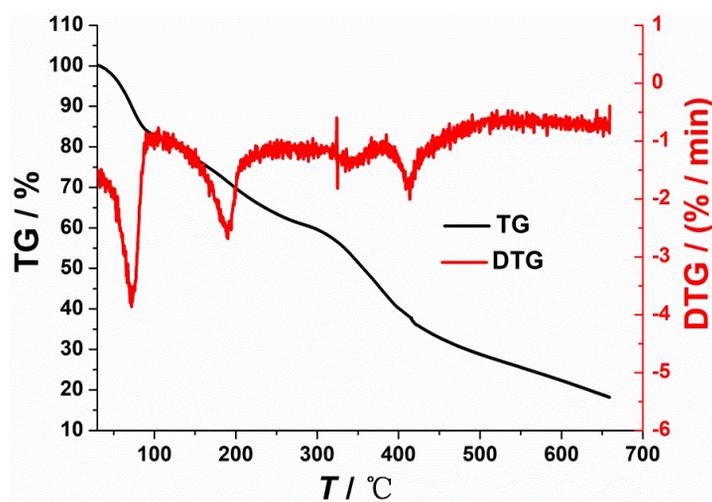


Figure S6 TG curves of complex 1 under N₂ atmosphere.

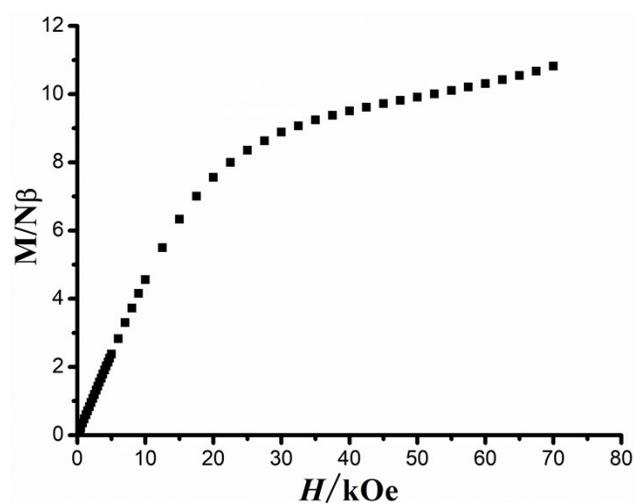


Figure S7 M versus H plot at 2 K for complex 2.

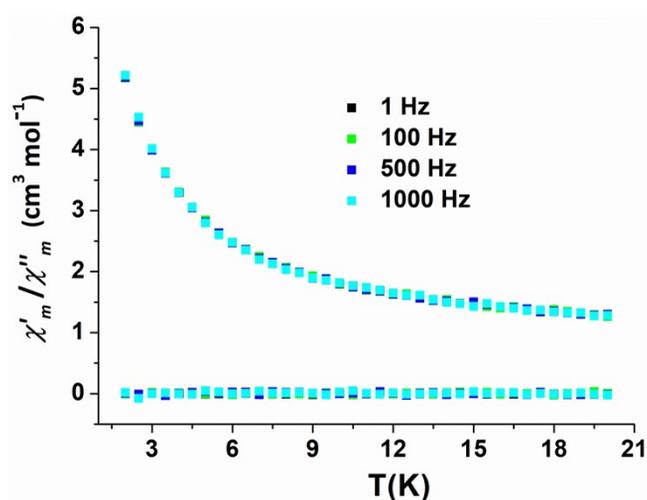


Figure S8 Temperature dependence of the in-phase and out-phase in zero field for complex 2.

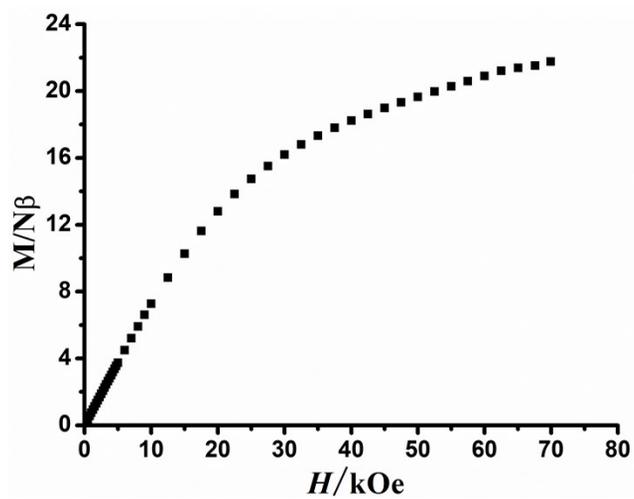


Figure S9 M versus H plot at 2 K for complex 3.

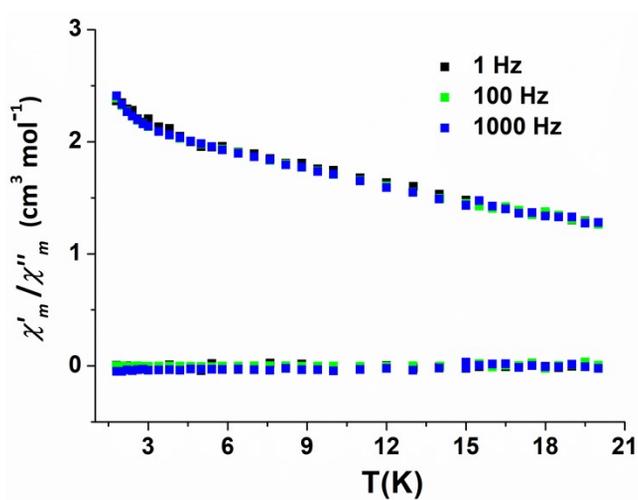


Figure S10 Temperature dependence of the in-phase and out-phase in zero field for complex 3.

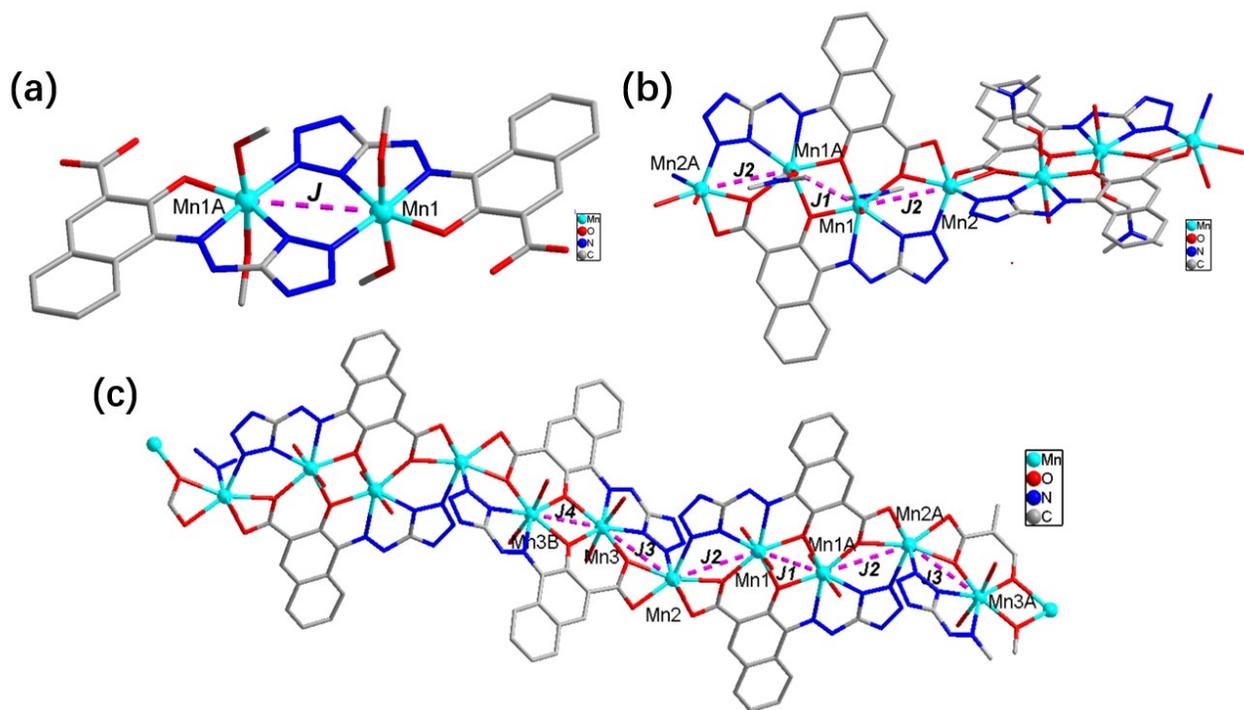


Figure S11 Magnetic exchange pathways between metal ions on the figure of X-ray structure in complexes 1-3.