

Small chromium-doped silicon clusters CrSi_n: Structures, IR spectra, charge effect, magnetism and chirality

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Content:

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- Relative energy comparison between B3P86/6-311+G(d) and B3LYP/6-311+G(d) theory levels of anionic CrSi_n⁻ cluster (n = 3-10) (Table S2).

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- The adiabatic ionization energies (IE_a) and the adiabatic electron affinities (AE_a) of neutral isomers having similar shape to the ground-state cationic and anionic structures (Table S4).
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- IR spectra comparison of CrSi_8^+ and CrSi_9^+ clusters between calculation and experiment (Figure S15 and Figure S16).
- Calculated-IR and UV-Vis spectra of ground-state enantiomers $\text{CrSi}_n^{+/0-}$ ($n = 9$ for cation, $n = 7$ for neutral and anion) (Figure S17).
- Calculated-ECD spectra of higher-energy enantiomers of cationic, neutral and anionic $\text{CrSi}_n^{+/0-}$ clusters with $n = 3-10$ (Figure S18).

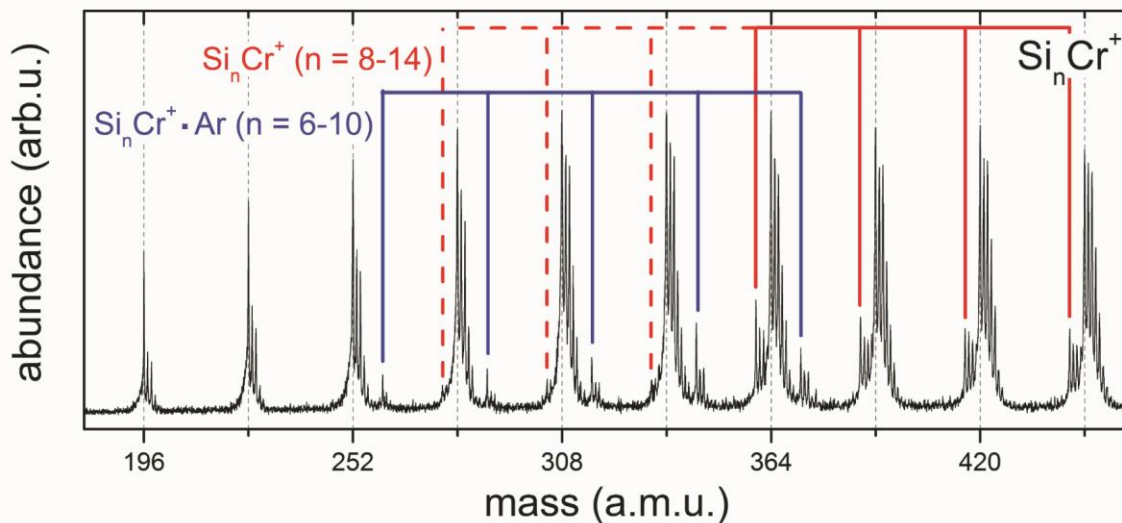


Figure S1. Mass spectrum of CrSi_n^+ clusters with the source at a temperature of 80 K and a 5 % argon mixture in the helium carrier gas. Argon complexes are found for $\text{CrSi}_n^+ \cdot \text{Ar}$ ($n = 6-10$). For these sizes, no CrSi_n^+ clusters are observed in the mass spectrum. The exohedrally doped clusters react completely with the argon atoms. Starting from $\text{CrSi}_{11}\text{Cr}^+$, unreacted doped clusters appear in the mass spectrum. The highest peaks (on the grid lines) in the mass spectrum correspond to pure Si_n^+ clusters (note that Si and Cr have multiple isotopes, of which ^{28}Si and ^{52}Cr have the highest abundances).

Table S1. Relative energy comparison between B3P86/6-311+G(*d*) and B3P86/aug-cc-pVTZ theory levels of cationic, neutral and anionic CrSi_n^{+0/-} clusters (n = 3-10)

Isomer	Electronic state	6-311+G(<i>d</i>)	aug-cc-pVTZ	Isomer	Electronic state	6-311+G(<i>d</i>)	aug-cc-pVTZ
C3-iso1	⁴ B ₂	0.00	0.00	A9-iso1	⁶ A	0.00	0.00
C3-iso2	⁴ A ₁	0.33	0.33	A9-iso2	² A ₁	0.05	0.18
C3-iso3	⁴ A'	0.36	0.38	N3-iso1	⁵ A ₁	0.00	0.00
C4-iso1	⁶ A'	0.00	0.00	N3-iso1	⁷ A ₁	0.39	0.15
C4-iso2	⁶ A ₁	0.22	0.24	N3-iso2	⁵ A'	0.50	0.37
C5-iso1	⁶ A'	0.00	0.00	N4-iso1	⁵ B ₂	0.00	0.00
C5-iso2	⁴ A''	0.36	0.39	N4-iso1	⁷ A	0.31	0.32
C6-iso1	⁶ A'	0.00	0.01	N4-iso2	⁵ A''	0.29	0.25
C6-iso2	⁶ A'	0.03	0.00	N4-iso3	⁷ A	0.31	0.30
C7-iso1	⁶ A ₁	0.00	0.00	N5-iso1	⁵ B ₁	0.00	0.00
C7-iso2	⁶ A	0.16	0.15	N5-iso2	⁵ A'	0.09	0.08
C7-iso3	⁶ A ₁	0.19	0.22	N7-iso1	⁵ A'	0.00	0.05
C8-iso1	⁴ A''	0.00	0.00	N7-iso2	⁵ A	0.02	0.01
C8-iso2	⁶ A'	0.06	0.04	N7-iso3	⁵ A'	0.06	0.00
C10-iso1	⁶ A	0.00	0.00	N7-iso4	⁵ A'	0.07	0.06
C10-iso2	⁶ A'	0.04	0.03	N8-iso1	⁵ A''	0.00	0.00
A6-iso1	⁴ A ₁	0.00	0.00	N8-iso2	⁵ A	0.05	0.00
A6-iso2	⁶ A	0.32	0.33	N8-iso3	³ A'	0.06	0.01
A7-iso1	⁶ A'	0.00	0.04	N8-iso4	⁵ A''	0.16	0.18
A7-iso1	⁴ A'	0.01	0.04	N9-iso1	⁵ A	0.00	0.00
A7-iso2	⁶ A	0.05	0.00	N9-iso2	⁵ A	0.07	0.12
A8-iso1	⁶ A ₁	0.00	0.00	N10-iso1	⁵ A'	0.00	0.00
A8-iso2	⁶ A'	0.18	0.23	N10-iso2	⁷ A ₁	0.01	0.02
A8-iso3	⁶ A'	0.22	0.16	N10-iso3	⁵ A'	0.02	0.01

Table S2. Relative energy comparison between B3P86/6-311+G(*d*) and B3LYP/6-311+G(*d*) theory levels of anionic CrSi_n⁻ cluster (n = 6-10)

Isomer	B3P86/6-311+G(<i>d</i>)	B3LYP/6-311+G(<i>d</i>)	Isomer	B3P86/6-311+G(<i>d</i>)	B3LYP/6-311+G(<i>d</i>)
A6-iso1	0.00	0.00	A8-iso1	0.00	0.00
A6-iso2	0.32	0.46	A8-iso2	0.18	0.02
A7-iso1 (⁶ A')	0.00	0.00	A8-iso3	0.22	0.13
A7-iso1 (⁴ A')	0.01	0.03	A9-iso1	0.00	0.00
A7-iso2-L/R	0.05	0.10	A9-iso2	0.05	0.17

Table S3. Average binding energy, fragmentation energies and natural population analysis for the ground-states of the cationic, neutral, anionic $\text{CrSi}_n^{+/0-}$ clusters.

	Isomer	State	$D_{C/N}(\text{Cr}^+)$	$D_{N/A}(\text{Cr})$	$D_{C/N/A}(\text{Si})$	$E_{b-C/N/A}$	Atomic charge on Cr	Electron population on Cr's 3d shell	Electron population on Cr's 4s shell	Spin density of Cr	Total Wiberg bond of Cr	Total Wiberg bond of cleaved Si
cation	C3-iso1	4B_2	4.77		5.02	2.40	0.83	4.86	0.30	4.54	1.36	2.32
	C4-iso1	$^6A'$	2.10		5.70	2.74	0.84	4.92	0.23	4.82	0.66	2.23
	C5-iso1	$^6A'$	2.58		5.98	2.95	0.77	4.90	0.29	4.62	1.04	2.79
	C6-iso1	$^6A'$	2.21		6.42	3.12	0.86	4.87	0.25	4.77	0.79	2.47
	C6-iso2	$^6A'$	2.19		6.55	3.11	0.82	4.94	0.21	4.73	0.74	2.41
	C7-iso1	6A_1	2.32		6.20	3.27	0.83	4.96	0.20	4.87	0.55	2.83
	C7-iso2	6A	2.20		6.01	3.25	0.82	4.96	0.20	4.74	0.72	2.72
	C7-iso3	6A_1	2.13		6.08	3.25	0.82	4.92	0.26	5.01	0.44	2.76
	C8-iso1	$^4A''$	5.20		5.72	3.23	0.78	4.97	0.22	4.34	1.45	2.48
	C8-iso2	$^6A'$	2.95		5.26	3.22	0.72	4.94	0.28	4.69	0.94	2.82
	C9-iso1-L/R	6A	2.42		6.33	3.33	0.75	5.03	0.18	4.54	0.91	2.76
	C10-iso1	6A	2.52		7.07	3.42	0.82	4.96	0.20	4.71	0.74	2.98
C10-iso2	$^6A'$	2.48		6.94	3.42	0.76	4.98	0.23	4.66	0.87	3.39	
neutral	N3-iso1	5A_1	6.78	3.69	4.96	2.59	0.66	4.94	0.38	4.67	1.22	2.74
	N3-iso1	7A_1	6.63	2.02	5.52	2.54	0.60	5.00	0.38	4.98	0.69	3.02
	N4-iso1	5B_2	6.76	2.87	5.62	2.89	0.61	4.98	0.33	4.56	1.32	2.65
	N5-iso1	5B_1	7.27	4.22	6.07	3.12	0.58	4.99	0.36	4.47	1.50	2.85
	N5-iso2	$^5A'$	6.77	4.19	6.74	3.10	0.63	4.99	0.35	4.73	1.07	2.85
	N6-iso1	5B_2	6.77	3.25	5.92	3.25	0.64	5.05	0.25	4.44	1.34	2.79
	N7-iso1	$^5A'$	7.44	4.14	5.67	3.26	0.44	5.04	0.40	4.32	1.76	2.63
	N7-iso2-L/R	5A	6.81	3.65	5.63	3.25	0.61	4.98	0.35	4.48	1.43	2.82
	N7-iso3	$^5A'$	6.12	2.41	5.41	3.25	0.67	5.07	0.24	4.62	1.01	2.82
	N8-iso1	$^5A''$	6.40	3.85	5.67	3.30	0.58	5.14	0.22	4.31	1.38	2.73
	N8-iso2	$^5A''$	7.00	4.13	6.38	3.29	0.61	5.03	0.29	4.44	1.38	3.52
	N8-iso3	$^3A'$	9.74	7.10	5.89	3.25	0.27	5.23	0.35	3.90	2.45	2.65
	N9-iso1	5A	6.82	4.15	6.40	3.36	0.56	5.10	0.23	4.18	1.60	3.02
	N10-iso1	$^5A'$	6.14	2.80	6.09	3.43	0.67	5.04	0.27	4.59	1.02	2.68
	N10-iso2	7A_1	6.15	1.62	6.65	3.43	0.57	4.96	0.45	5.04	0.77	2.71
	N10-iso3	$^5A'$	6.13	2.89	6.30	3.43	0.60	5.09	0.27	4.53	1.16	3.07
anion	A3-iso1	6A_1		3.06	5.69	2.80	0.30	4.99	0.66	5.08	1.04	2.96
	A4-iso1	4A_2		5.34	6.15	3.00	0.42	5.09	0.40	4.38	1.81	2.78
	A5-iso1	$^6A'$		3.37	7.07	3.29	0.46	5.05	0.44	4.87	1.03	2.84
	A6-iso1	4A_1		7.42	7.29	3.37	0.24	5.04	0.57	4.15	2.41	3.07
	A6-iso2	6A		2.91	5.67	3.33	0.50	5.16	0.27	4.65	1.00	2.85
	A7-iso1	$^6A'$		4.40	6.39	3.38	0.28	5.14	0.41	4.51	1.48	3.06
	A7-iso1	$^4A'$		6.96	6.20	3.38	0.03	5.28	0.47	3.89	2.50	3.06
	A7-iso2-L/R	6A		3.09	5.71	3.37	0.45	5.07	0.43	4.75	1.14	2.95
	A8-iso1	6A_1		3.24	6.33	3.45	0.52	5.21	0.23	4.54	1.03	2.98
	A8-iso2	$^6A'$		2.92	6.68	3.43	0.47	4.97	0.55	5.03	0.94	3.04
	A8-iso3	$^6A'$		3.80	6.66	3.42	0.52	5.09	0.33	4.62	1.12	3.60
	A9-iso1	6A		3.75	6.19	3.50	0.43	5.23	0.23	4.43	1.21	3.06
	A9-iso2	2A_1		10.02	6.60	3.49	-1.73	6.90	0.47	0.96	4.26	3.13
A10-iso1	2A		10.73	6.33	3.55	-2.64	7.77	0.45	0.81	4.22	2.91	

Table S4. The adiabatic ionization energies (IE_a) and the adiabatic electron affinities (AE_a) of neutral isomers having similar shape to the ground-state cationic and anionic structures.

Structure	IE_a (eV)	AE_a (eV)
C5-iso1	7.51	
A5-iso1		3.10
C6-iso1	7.03	
C6-iso2	7.07	
C7-iso1	6.30	
C7-iso2	6.50	
C7-iso3	6.15	
N7-iso2	7.09	2.90
N7-iso3	6.50	2.60
C8-iso2	6.65	
N8-iso2	7.34	3.11
N8-iso3	7.63	3.25
A8-iso2		3.57
C9-iso1	6.67	
N9-iso1	7.87	3.31
A9-iso2		3.52
N10-iso2	6.96	2.95
N10-iso3	6.63	2.98
A10-iso1		3.57

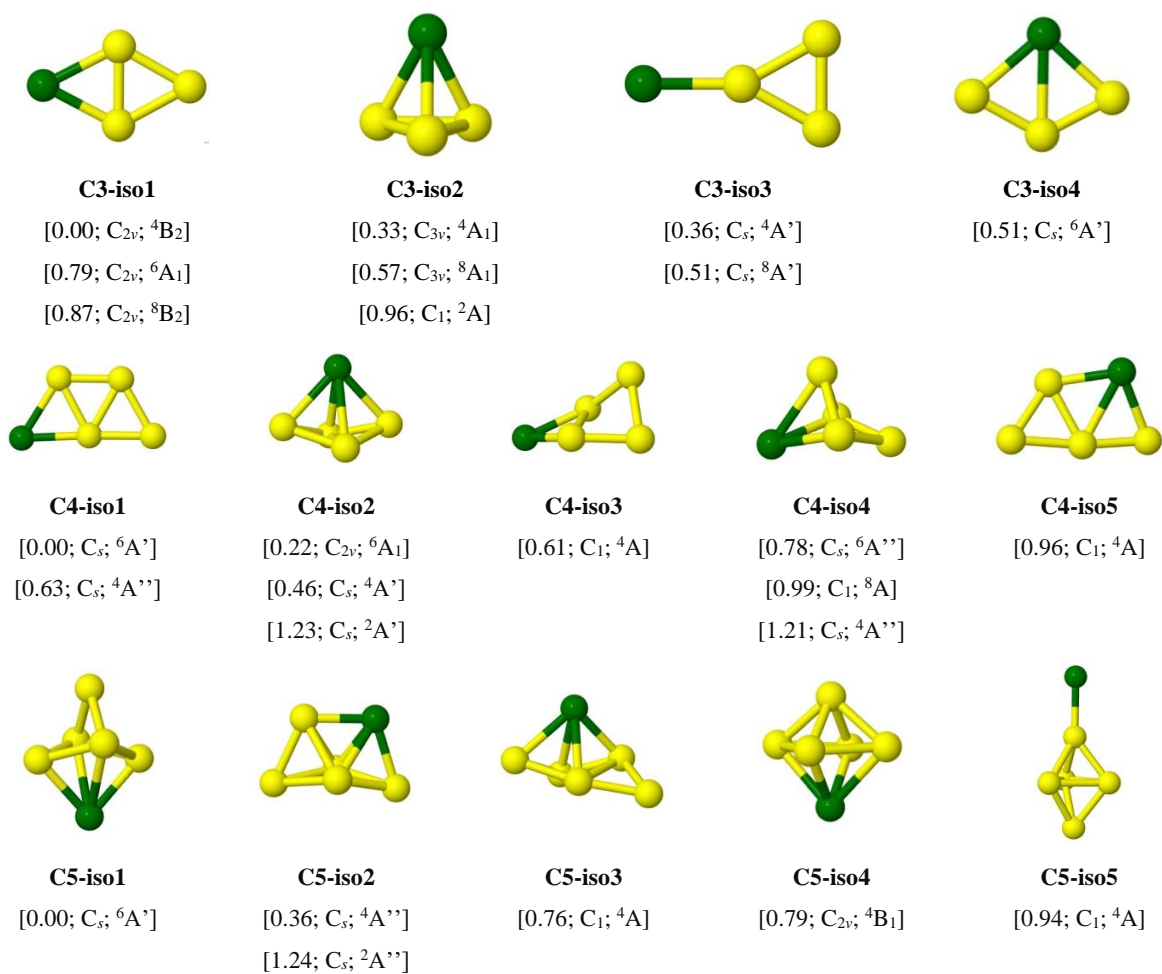


Figure S2. Stable isomers of the cationic CrSi_n⁺ cluster (n = 3-5)

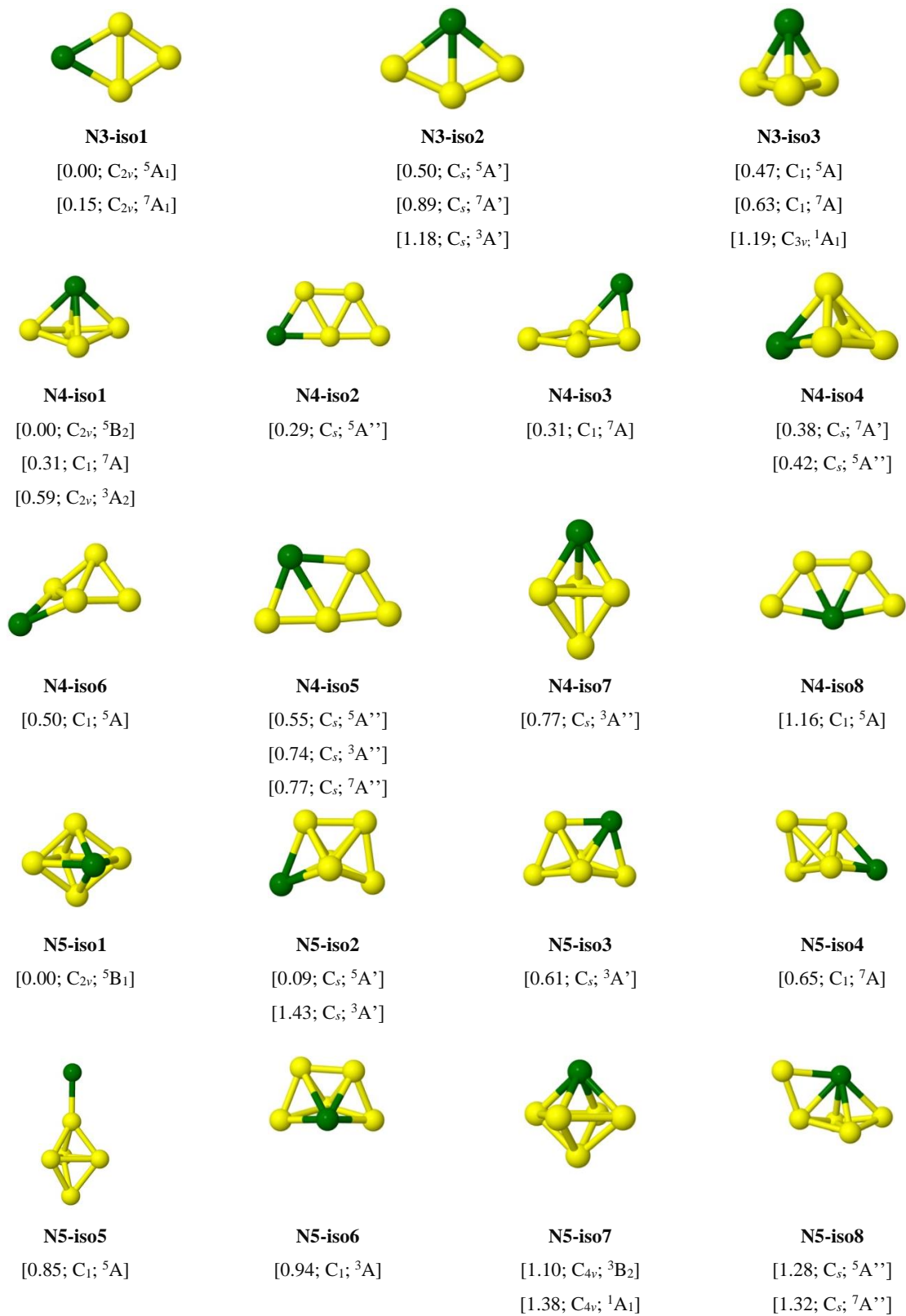


Figure S3. . Stable isomers of the neutral CrSi_n cluster (n = 3-5)

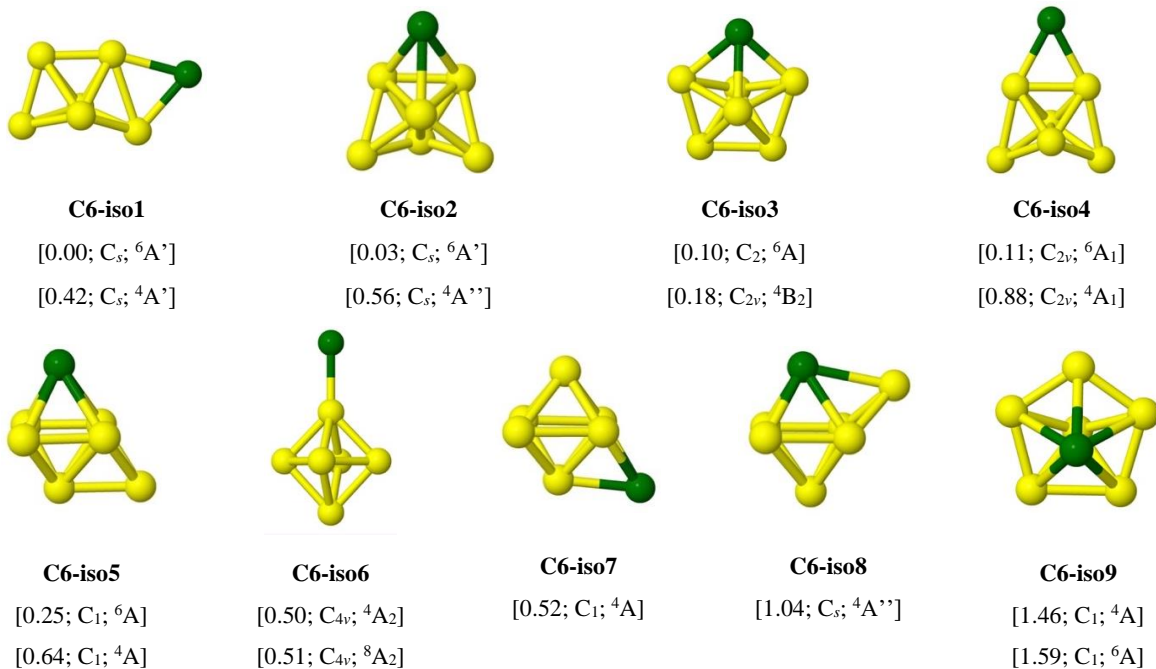


Figure S4. Stable isomers of the cationic CrSi₆⁺ cluster.

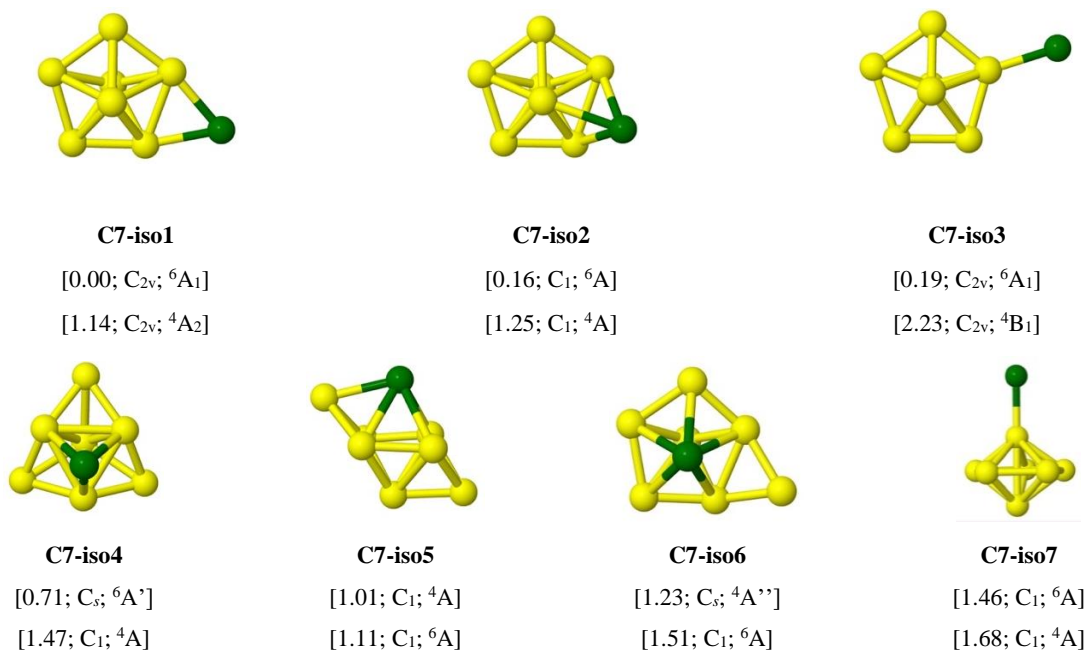
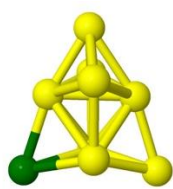
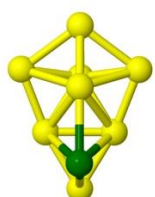


Figure S5. Stable isomers of cationic CrSi₇⁺ cluster.



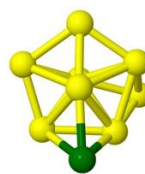
C8-iso1
[0.00; C_s; ⁴A'']
[0.43; C₁; ⁶A/
86i cm⁻¹]



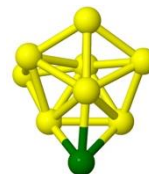
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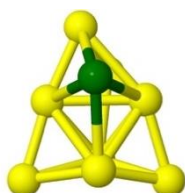
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[0.53; C_s; ⁶A']



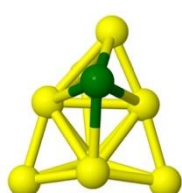
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[0.39; C₁; ⁶A]



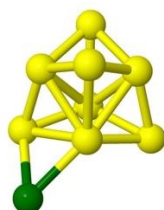
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[0.39; C₁; ⁶A]



C8-iso5-L
[0.23; C₁; ⁶A]
[0.31; C_s; ⁴A'']



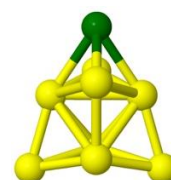
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[0.31; C_s; ⁴A'']



C8-iso6
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[0.49; C₁; ⁴A]



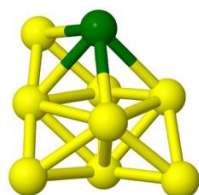
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[0.53; C₁; ⁶A]



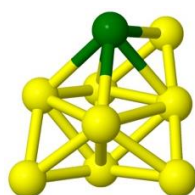
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[0.34; C₁; ⁶A]



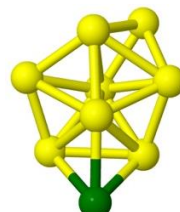
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[0.46; C₁; ⁶A]



C8-iso10-L
[0.29; C₁; ⁴A]
[0.29; C₁; ⁶A]



C8-iso10-R
[0.29; C₁; ⁴A]
[0.29; C₁; ⁶A]



C8-iso11
[0.32; C₁; ⁴A]
[0.55; C₁; ⁶A]



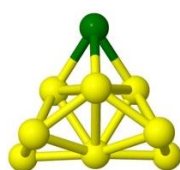
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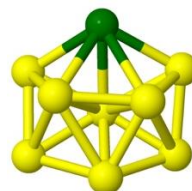
C8-iso13
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[0.60; C₁; ⁶A]



C8-iso14
[0.52; C₁; ⁴A]



C8-iso15
[0.54; C_s; ⁴A'']
[0.73; C_s; ⁶A'']



C8-iso16
[0.56; C₁; ⁴A]
[0.61; C₁; ⁶A]



C8-iso17
[0.67; C_s; ⁴A'']
[0.75; C_s; ⁶A'']

Figure S6. Stable isomers of cationic CrSi₈⁺ cluster.

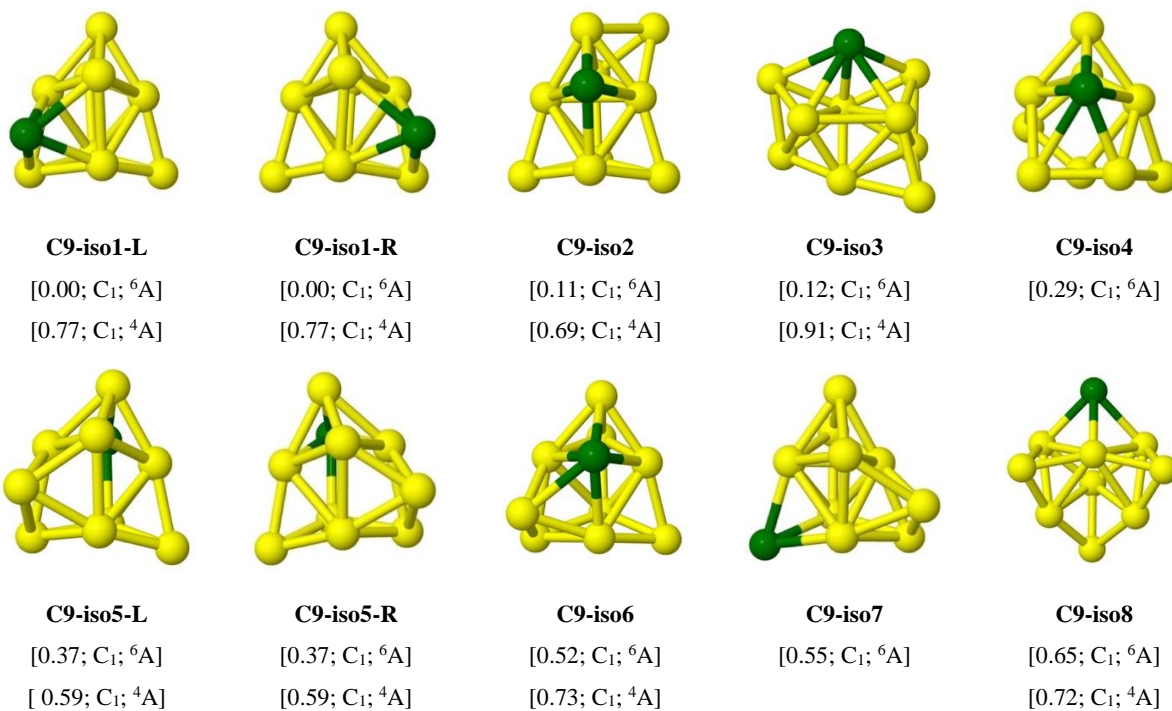


Figure S7. Stable isomers of cationic CrSi₉⁺ cluster.

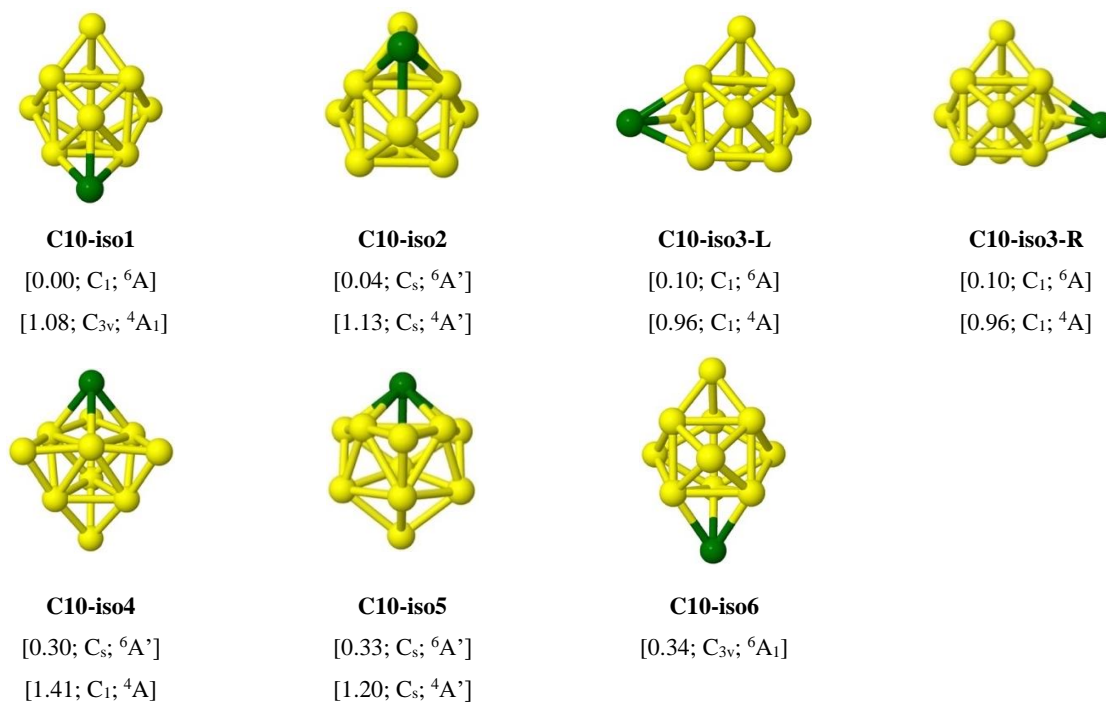


Figure S8. Stable isomers of cationic CrSi₁₀⁺ cluster.

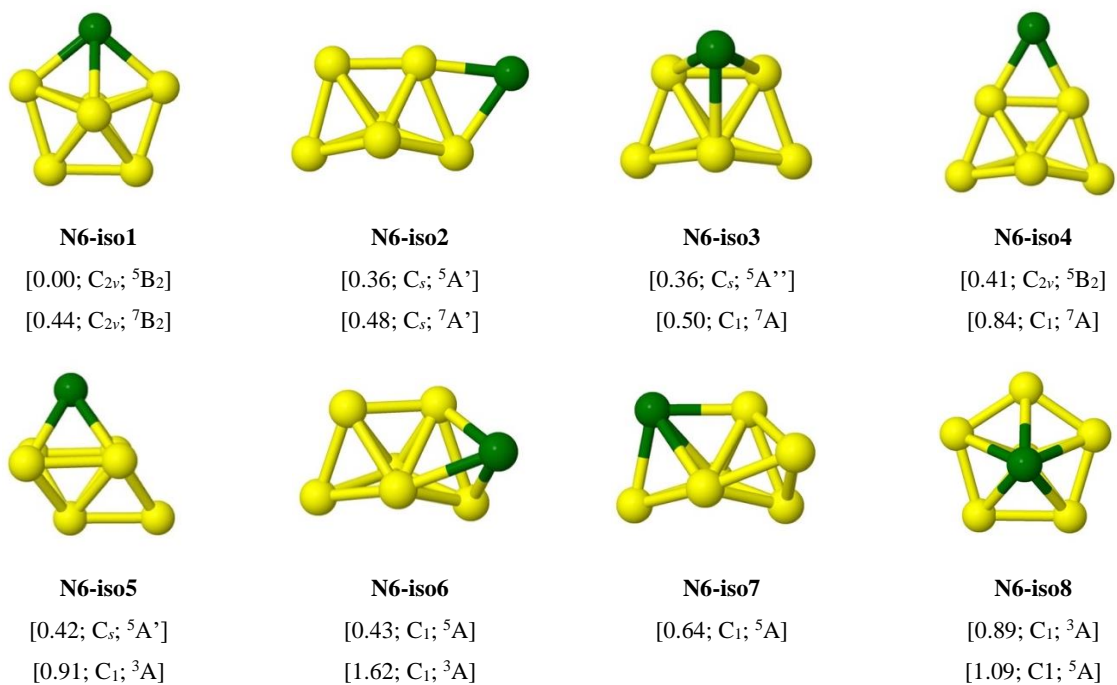


Figure S9. Stable isomers of neutral CrSi₆ cluster

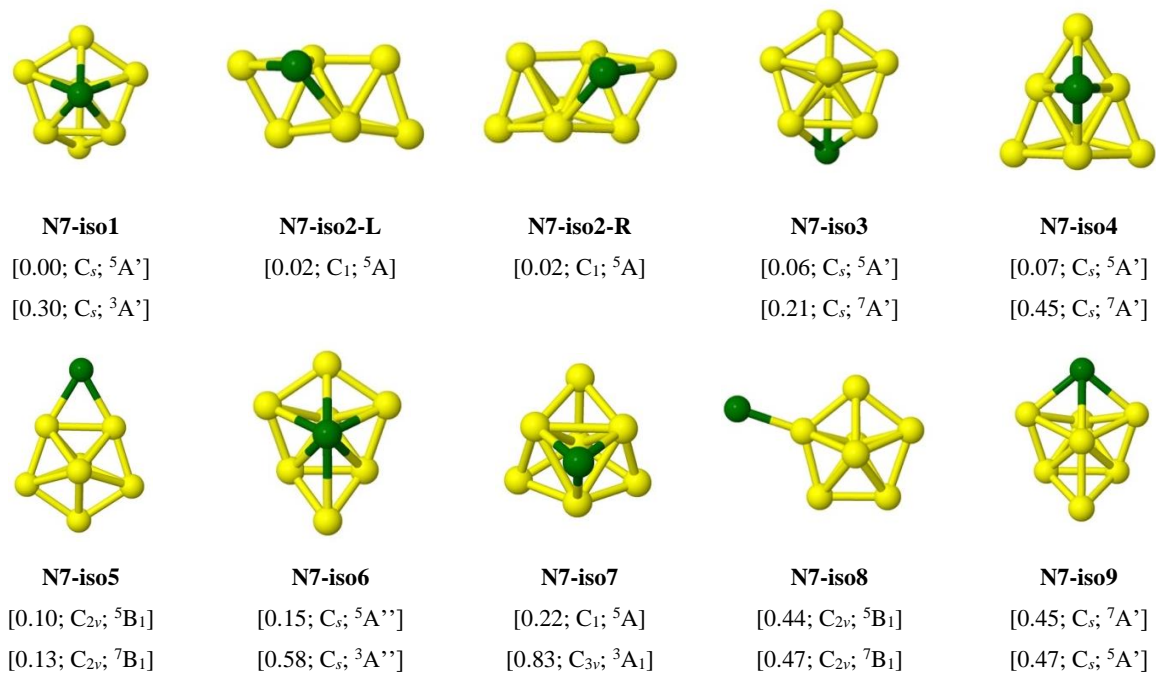


Figure S10. Stable isomers of neutral CrSi₇ cluster

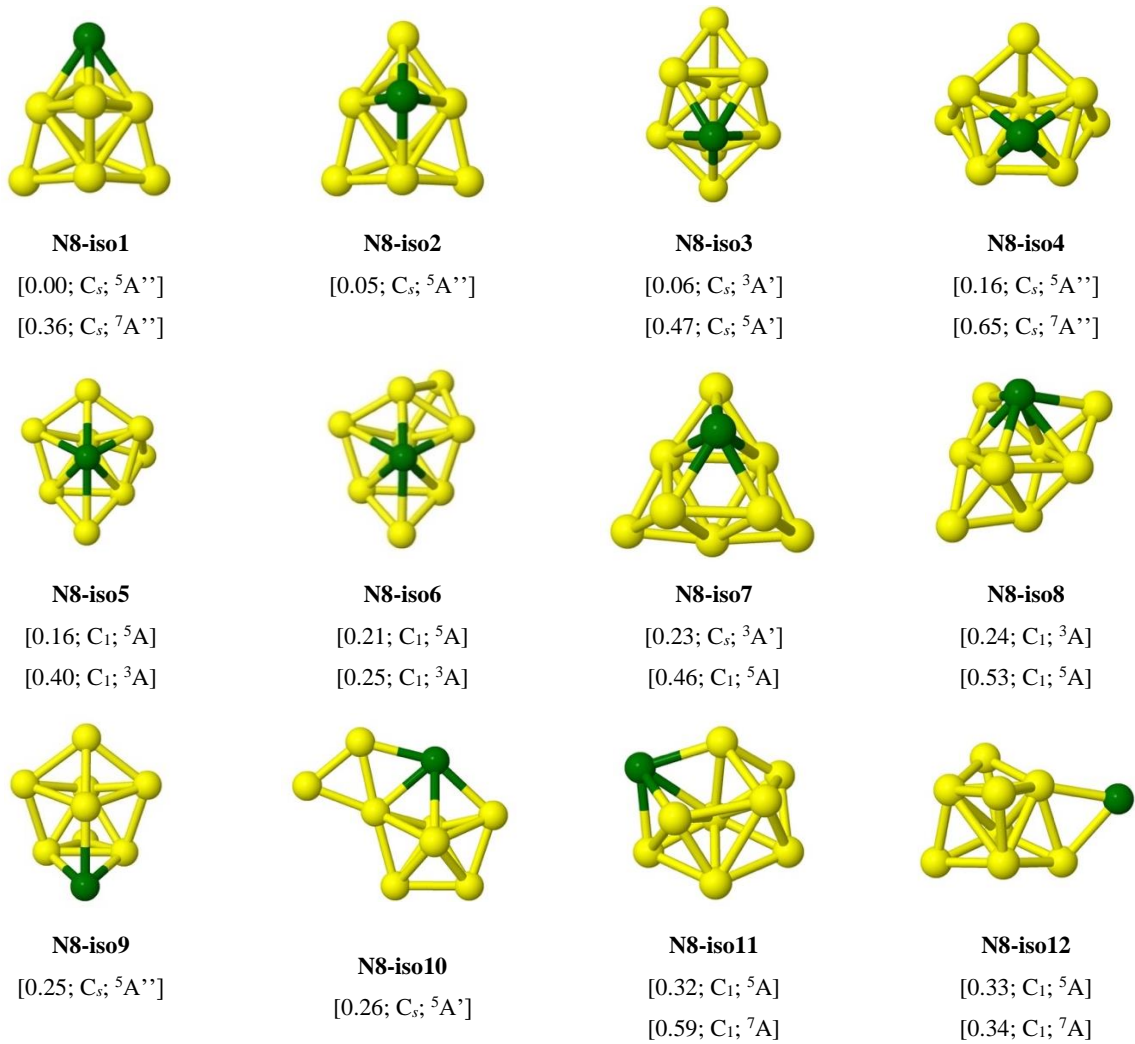


Figure S11. Stable isomers of neutral CrSi₈ cluster

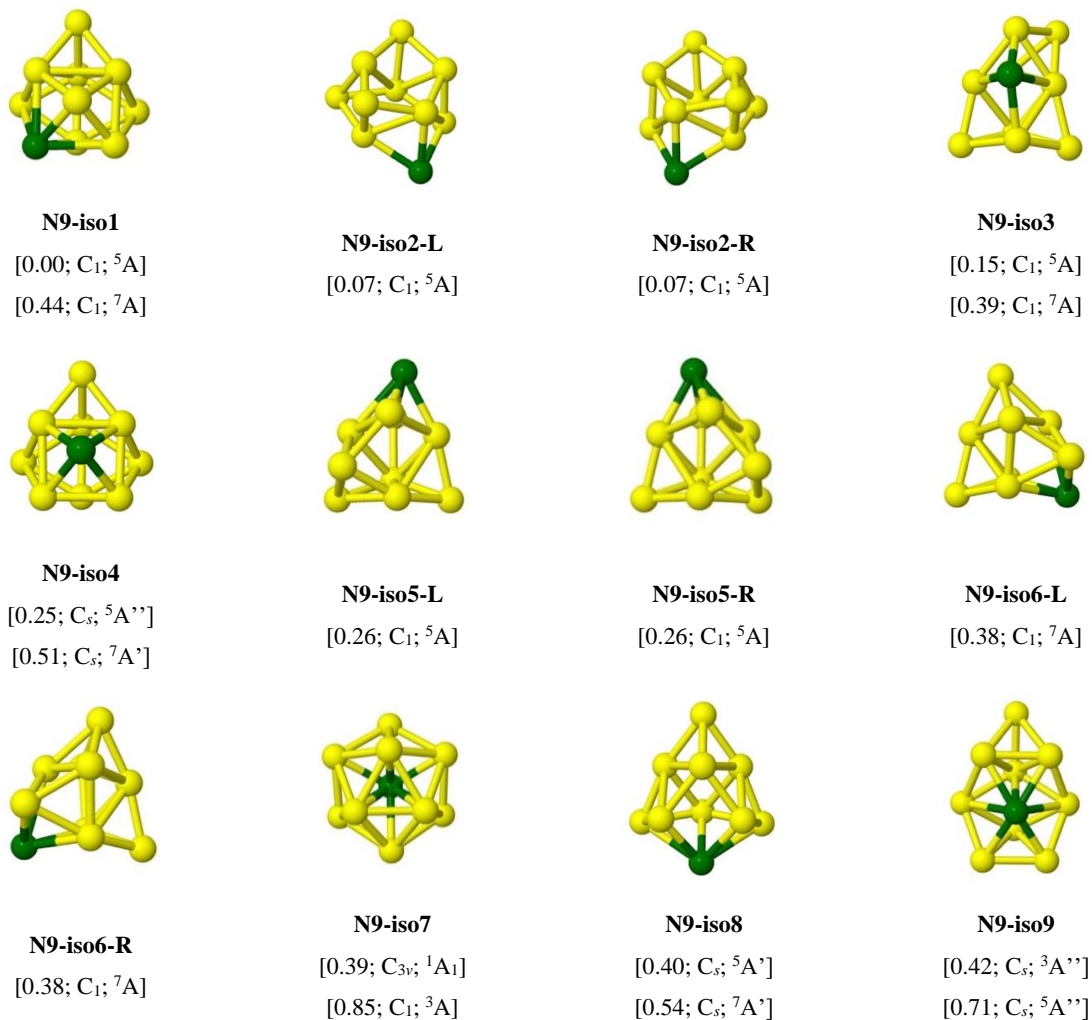


Figure S12. Stable isomers of neutral CrSi₉ cluster

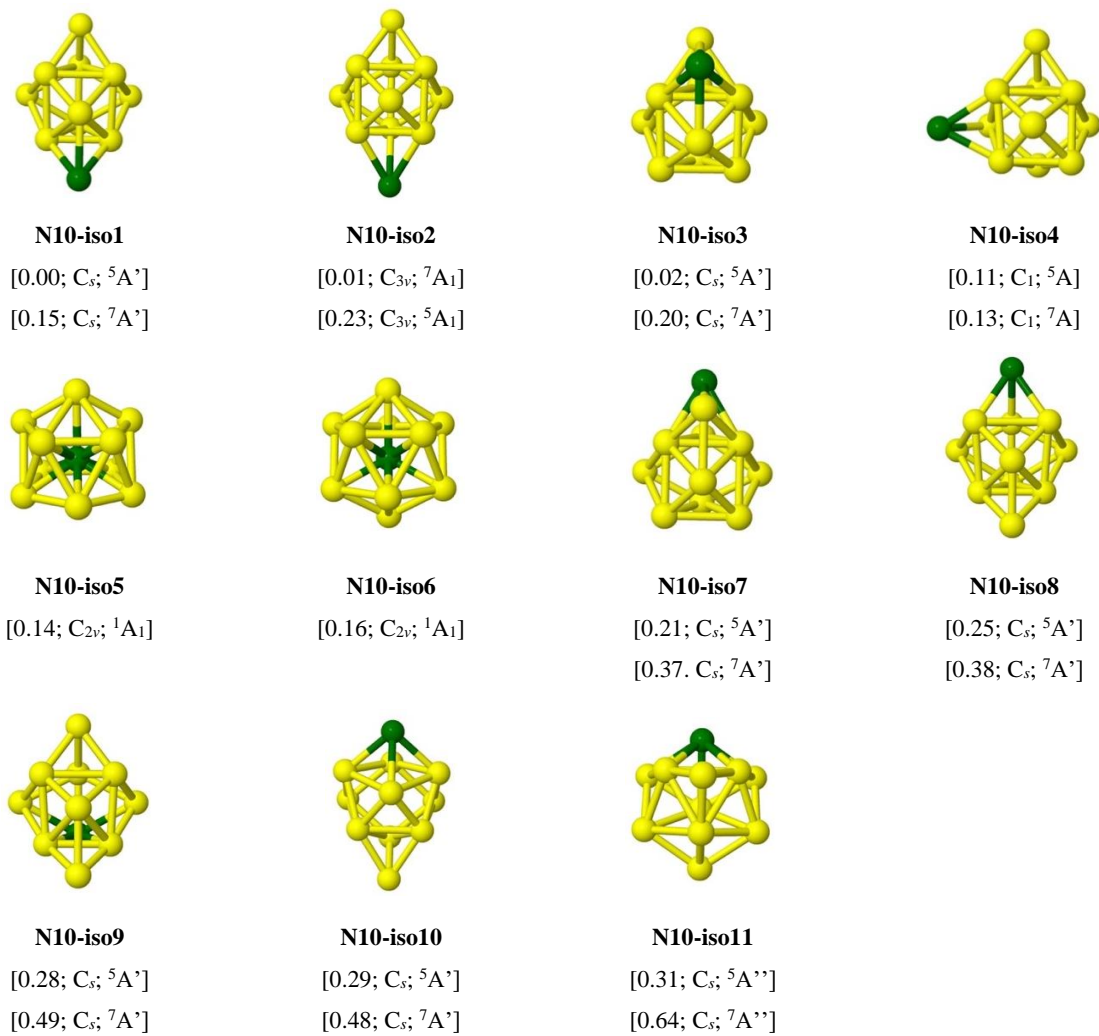


Figure S13. Stable isomers of neutral CrSi₁₀ cluster

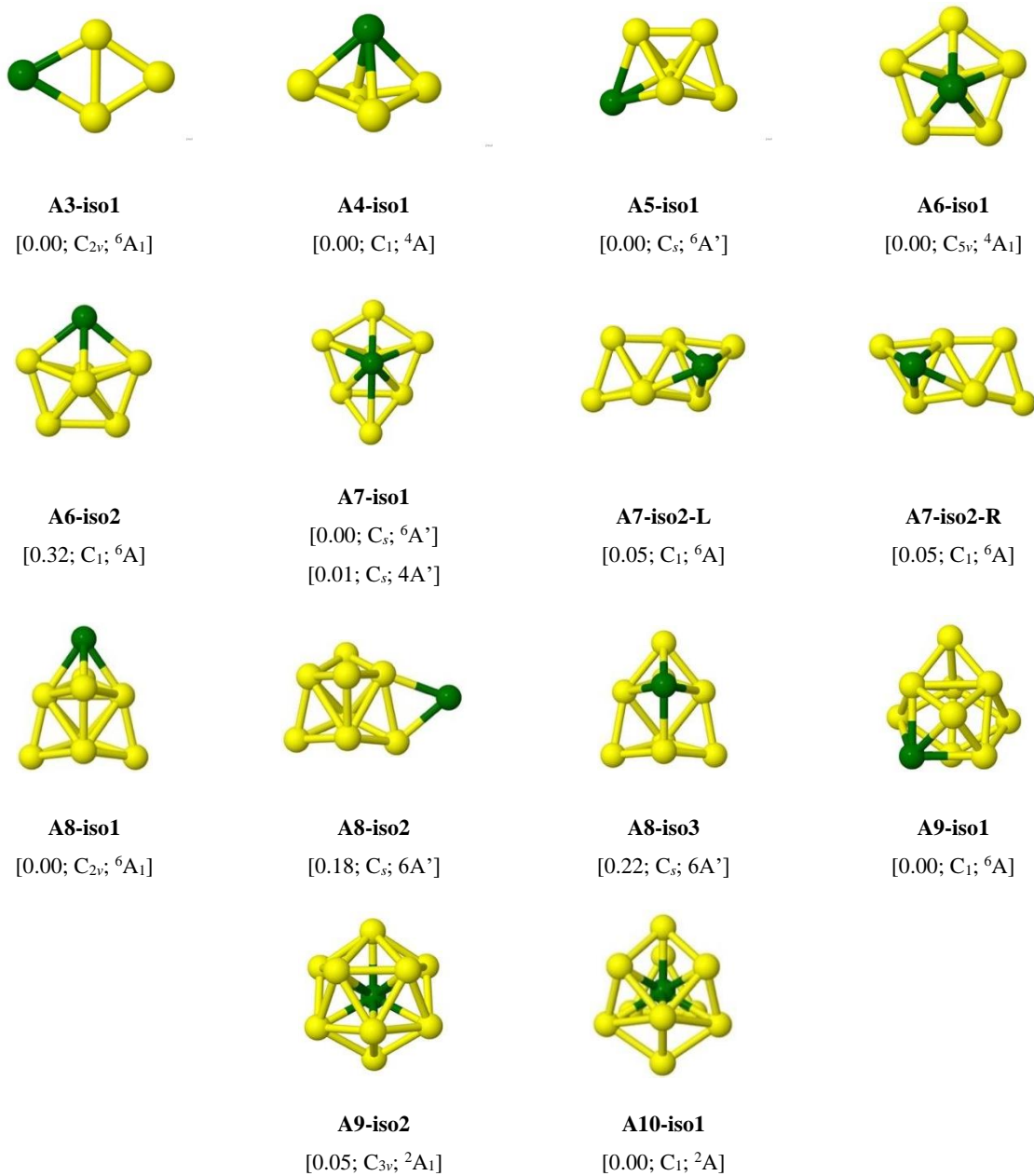


Figure S14. Stable structures of anionic CrSi_n⁻ cluster (n = 6-10).

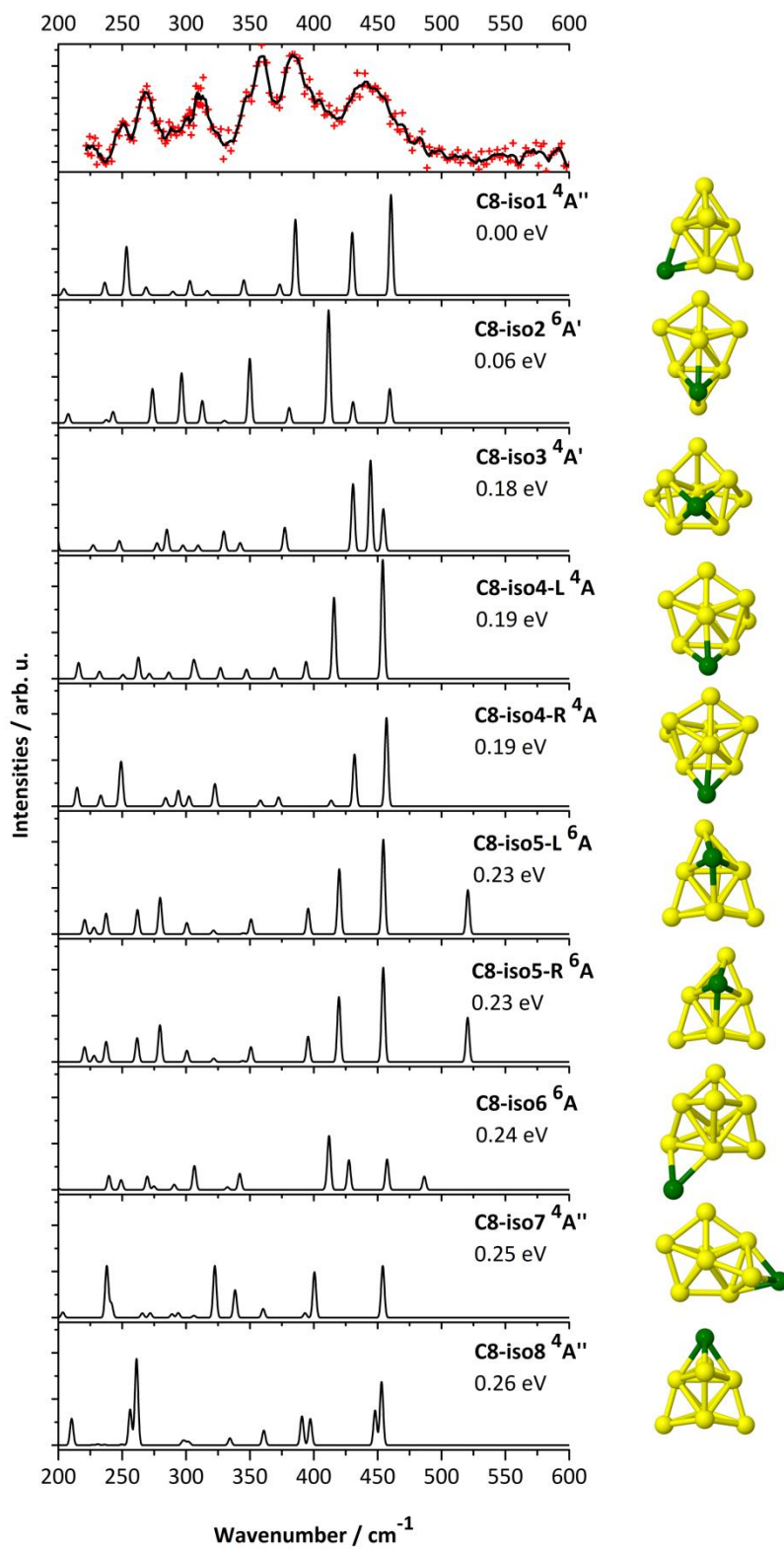


Figure S15. IR spectra comparison of CrSi_8^+ clusters between calculation and experiment.

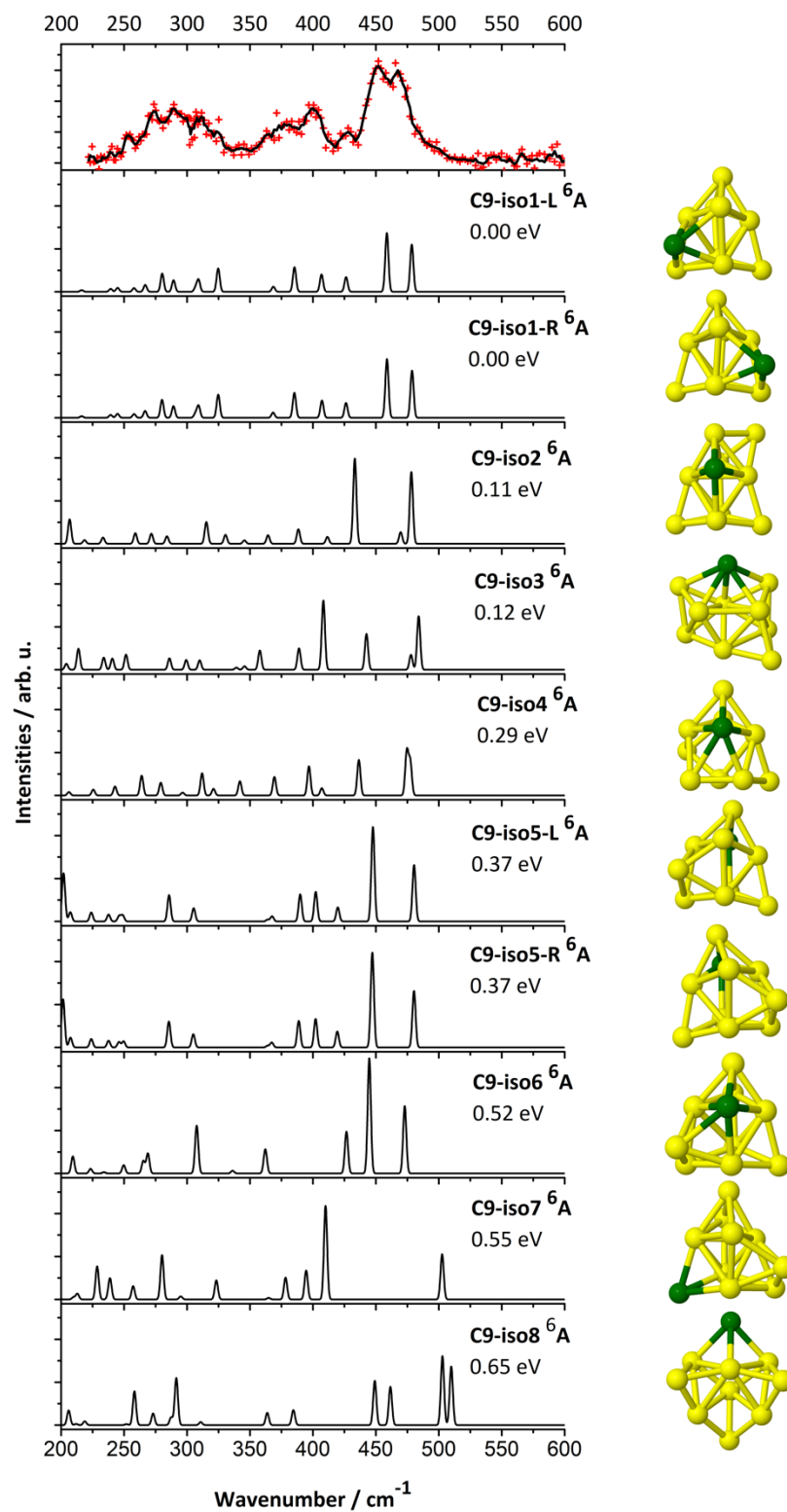


Figure S16. IR spectra comparison of CrSi_9^+ clusters between calculation and experiment.

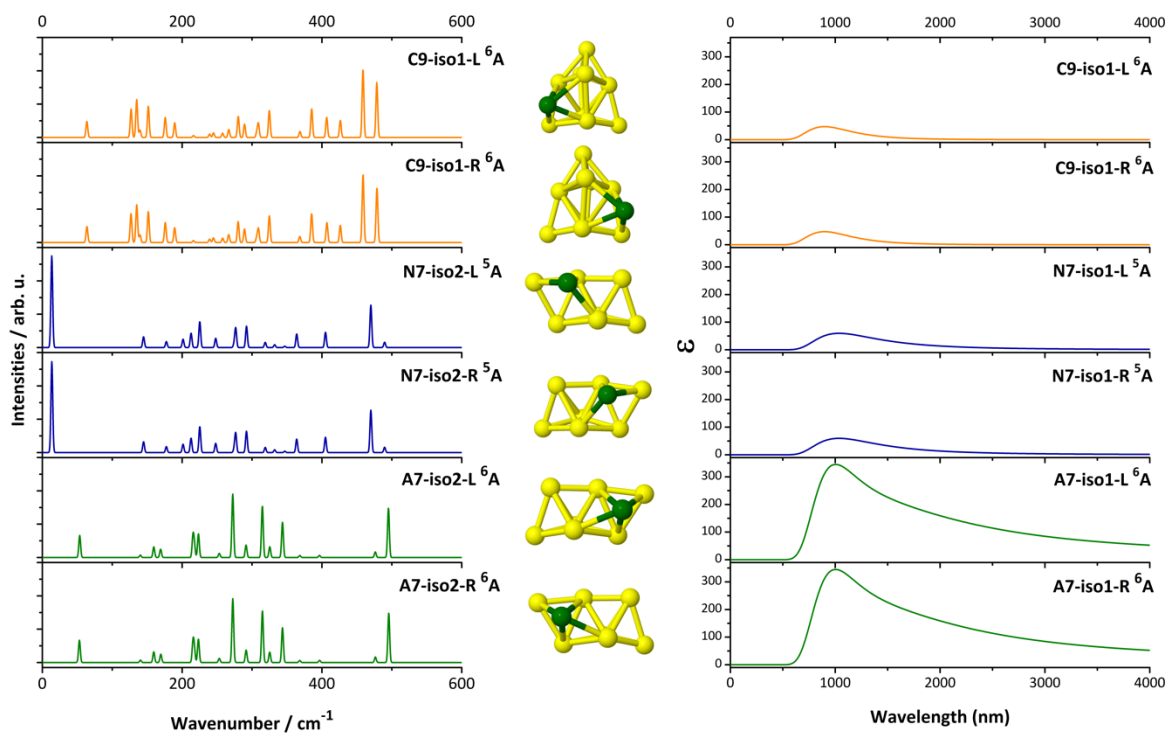


Figure S17. Calculated-IR and UV-Vis spectra of ground-state enantiomeric clusters $\text{CrSi}_n^{+/0/-}$ ($n = 9$ for cation, $n = 7$ for neutral and anion).

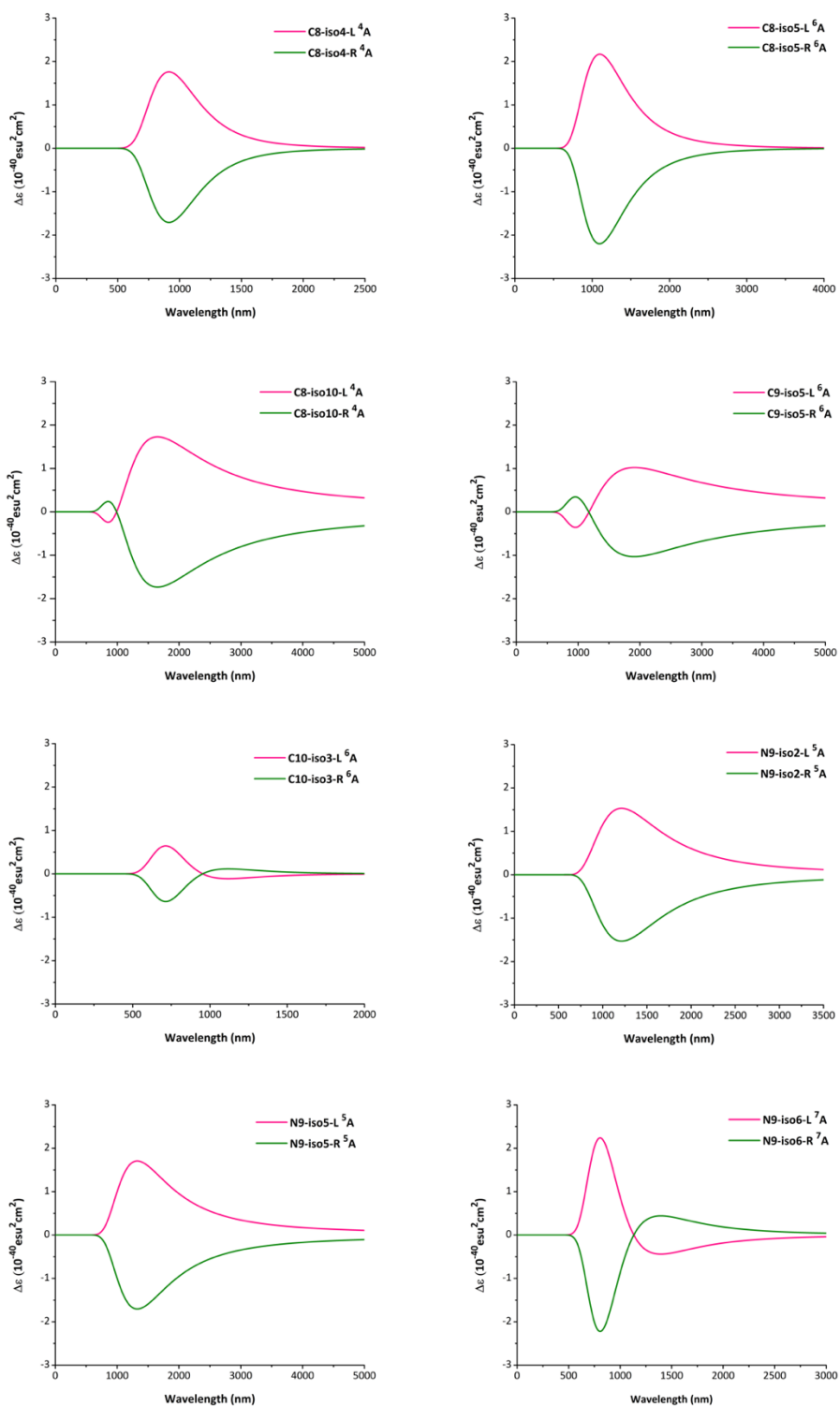


Figure S18. Calculated Electronic Circular Dichroism spectra of higher-energy enantiomeric clusters $\text{CrSi}_n^{+/0/-}$ ($n = 3-10$).