

Supporting Information

Varying the degree of oxidation of graphite: Effect of the oxidation time and the oxidant mass

Ioannis Karnis,^{1,2} Fanourios Krasanakis,¹ Labrini Sygellou,³

Anastassia N. Rissanou,^{1,4} Konstantinos Karatasos,⁵

and Kiriaki Chrissopoulou^{1,*}

¹Institute of Electronic Structure and Laser, Foundation for Research and Technology-Hellas, P.O. Box 1527, 711 10, Heraklion Crete, Greece

²Department of Chemistry, University of Crete, Heraklion Crete, Greece

³Institute of Chemical Engineering Studies, Foundation for Research and Technology-Hellas, Stadiou Str, 26504 Patras, Greece

⁴Institute of Theoretical and Physical Chemistry, National Hellenic Research Foundation, 48 Vassileos Konstantinou Ave, 11635 Athens, Greece

⁵Department of Chemical Engineering, Aristotle University of Thessaloniki, Thessaloniki, Greece

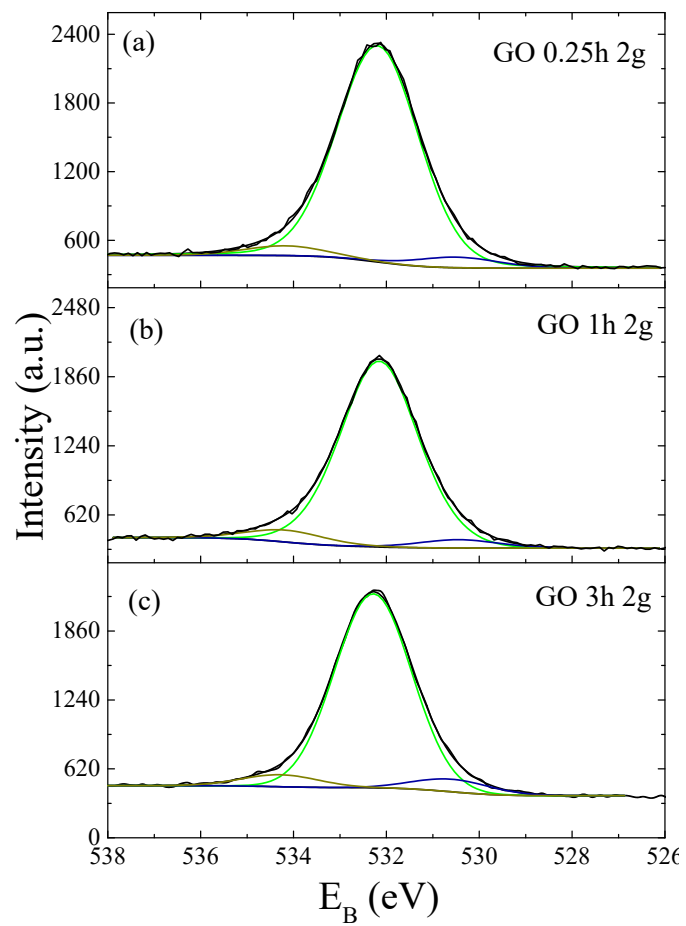
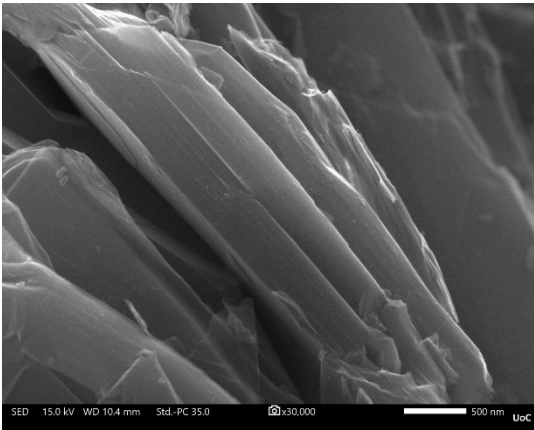
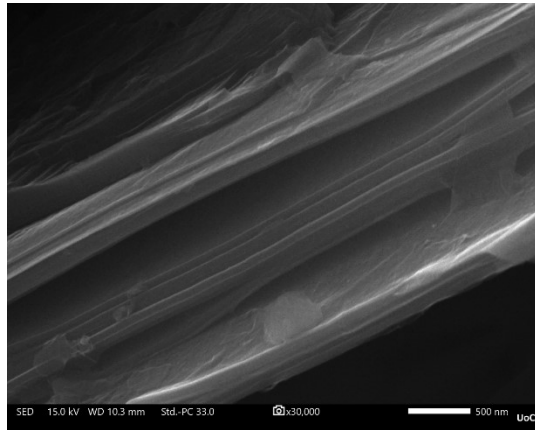


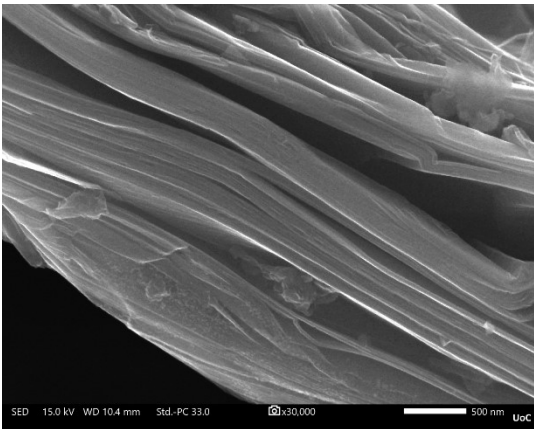
Figure S1: O1s peak deconvolution for O functional groups determination in GO oxidized with 2 g of oxidant during (a) 15 min, (b) 1h and (c) 3h of oxidation reaction.



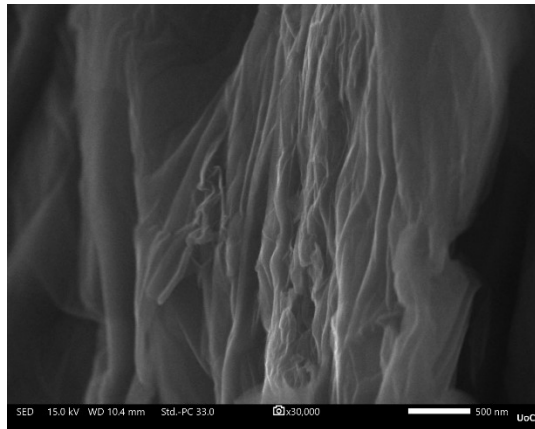
(a)



(b)



(c)



(d)

Figure S2: SEM images of graphite (a) and of GOs oxidized during 1h with 0.5g (b), 1.5g (c) and 6g (d) of oxidizing agent.

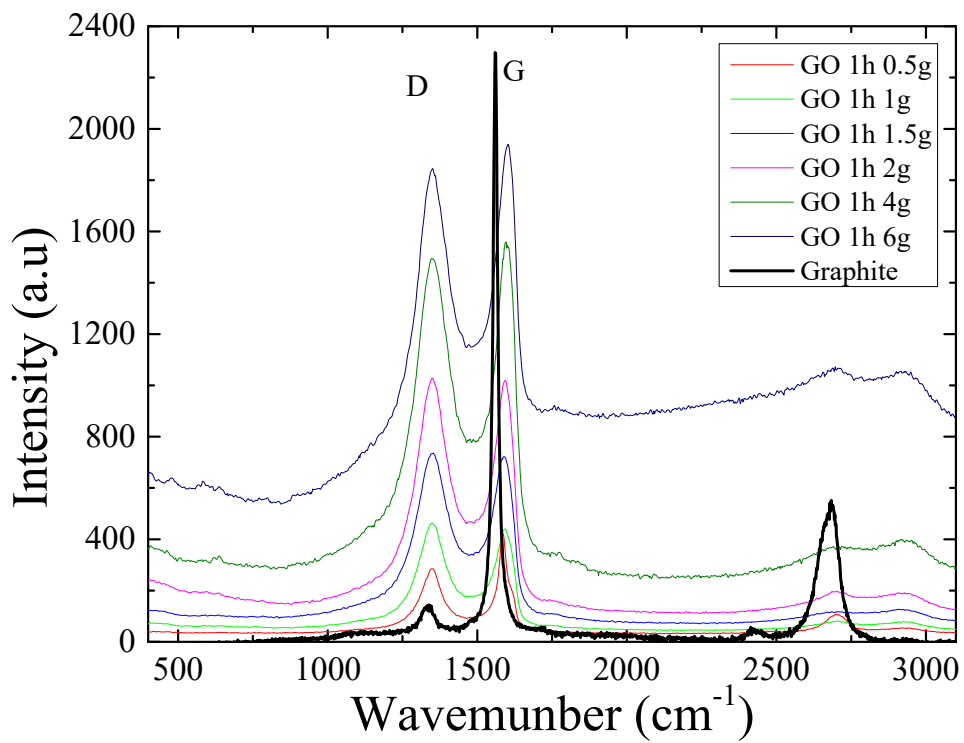


Figure S3: Raman spectra for GO samples synthesized during 1h oxidation time and different amounts of oxidant agent. The spectrum of graphite is shown for comparison as well.