

Supplementary Information

Catalytic investigation of hyaluronic acid stabilized Ag nanoparticles as non-toxic nanocatalysts in the oxidation of morin

M. Deniz Yilmaz,^{ab*} Nezahat Gokce Ozsamur,^{bc} Sundus Erbas-Cakmak^{bc}

^a *Department of Basic Sciences, Faculty of Engineering, Necmettin Erbakan University, 42140 Konya, Türkiye*

^b *BITAM-Science and Technology Research and Application Center, Necmettin Erbakan University, 42140 Konya, Türkiye.*

^c *Department of Molecular Biology and Genetics, Faculty of Science, Necmettin Erbakan University, 42090 Konya, Türkiye*

*Corresponding author: deniz.yilmaz@erbakan.edu.tr; yilmazdnz@gmail.com

Table of Contents

| | |
|---|---|
| 1. The concentration dependent absorption spectra of HA-AgNPs | 2 |
| 2. FTIR spectra of Hyaluronic acid and HA-AgNPs | 2 |
| 3. UV-vis spectra of 1,3-diphenylisobenzofuran (DPBF) in the presence and absence of HA-AgNPs..... | 3 |
| 4. Comparison of HA-AgNPs with other nanosystems used for the oxidation of morin reported in the literature | 4 |
| 5. Mass spectrum of degradation products of morin..... | 5 |
| 6. Real-time video of the oxidation of morin in the presence of H ₂ O ₂ and HA-AgNPs..... | 5 |

1. The concentration dependent absorption spectra of HA-AgNPs

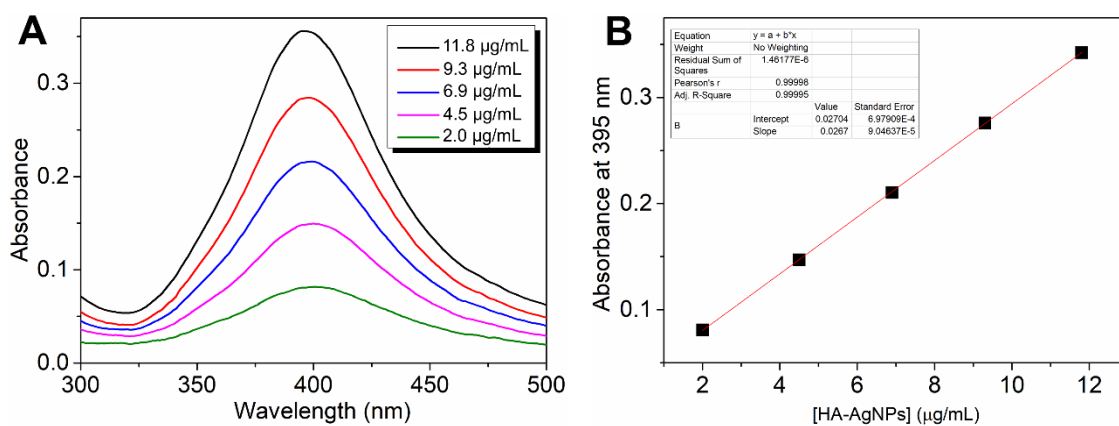


Fig. S1 (A) The concentration dependent UV-vis spectra of HA-AgNPs in H₂O and (B) the plot of the absorbance of HA-AgNPs at 395 nm versus the concentration.

2. FTIR spectra of Hyaluronic acid and HA-AgNPs

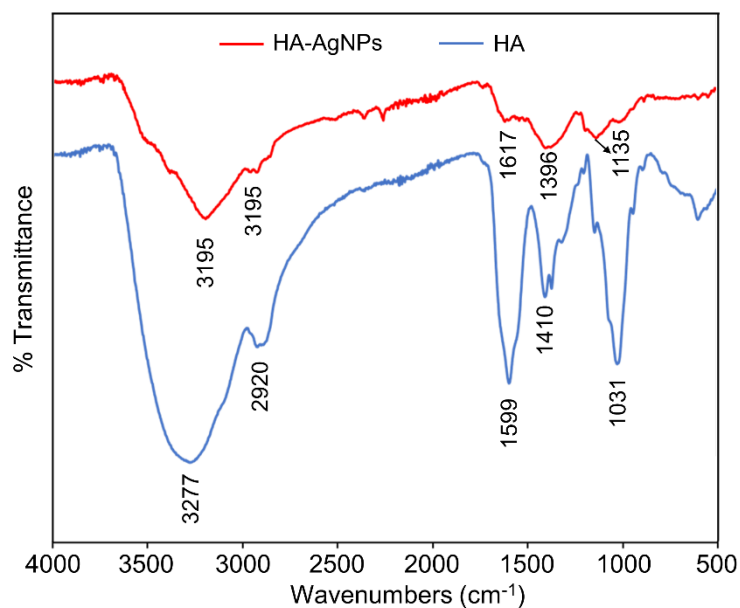


Fig. S2 FTIR spectra of HA and HA-AgNPs.

3. UV-vis spectra of 1,3-diphenylisobenzofuran (DPBF) in the presence and absence of HA-AgNPs

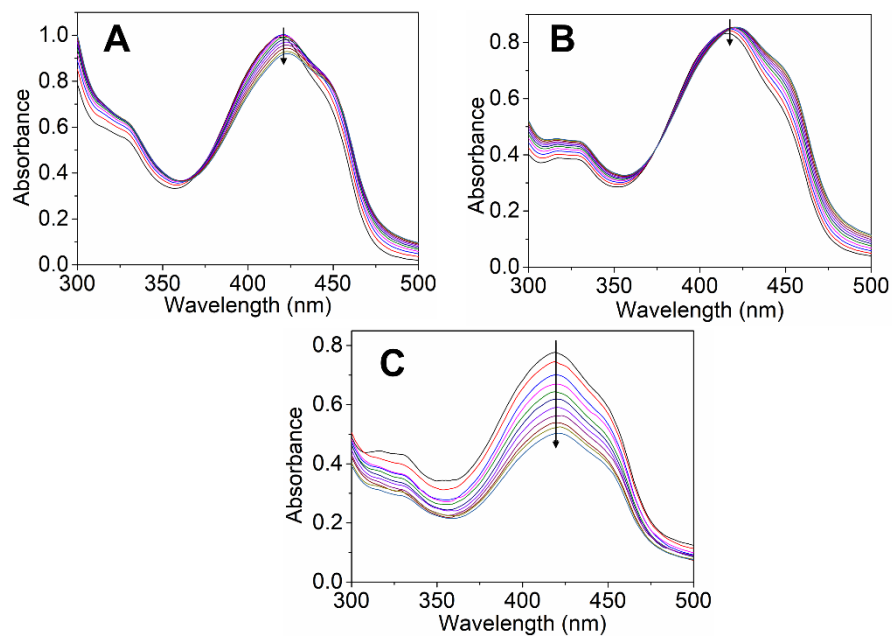


Fig. S3 UV-vis spectra of DPBF in 10 mM carbonate buffer at pH 10 at 298 K (A) in the presence of H_2O_2 (196 mM) and absence of HA-AgNPs; (B) in the presence of HA-AgNPs (4.5 $\mu\text{g}/\text{mL}$) and absence of H_2O_2 ; (C) in the presence of both HA-AgNPs (4.5 $\mu\text{g}/\text{mL}$) and H_2O_2 (196 mM).

4. Comparison of HA-AgNPs with other nanosystems used for the oxidation of morin reported in the literature

Table S1. Comparison of nanocatalysts in the oxidation of morin

| Catalyst | k_{app} ($s^{-1} \times 10^{-2}$) | Catalyst concentration | Morin concentration | H ₂ O ₂ concentration | Cytotoxicity, IC ₅₀ | Reaction time | Ref |
|---------------------------------|--|---------------------------|------------------------|--|--------------------------------|------------------|--------------|
| HA-AgNPs | 1.01 | 4.5 μ g/mL | 150 μ M | 196 mM | > 40 μ g/mL | 250 s | This work |
| Que-AgNPs | 0.14 | 16.6 μ g/mL | 100 μ M | 490 mM | 47.36 μ g/mL | 1200 s | 24 |
| AgDENs | 0.0151 | 490 nM | 100 μ M | 24 mM | na | 7000 s | 21 |
| CH-AuNPs | 0.12 | 2.5 nM | 100 μ M | 0.49 mM | > 6.8 nM | 1800 s | 15 |
| Pt ₂₀ @BSA | 0.80 | 600 nM | 125 μ M | 10 mM | > 2 μ M | 1200 s | 12 |
| MnO _x NP | 0.2 | 770 μ g/mL | 100 μ M | 10 mM | na | 1800 s | 4 |
| Pd ₅₅ -DENs | 0.384 | 300 nM | 900 μ M | 50 mM | na | 1800 s | 10 |
| Co ²⁺ @PEM- MSNPs | na | 133 μ g/mL | 160 μ M | 980 mM | na | 300 s | 13 |

5. Mass spectrum of degradation products of morin

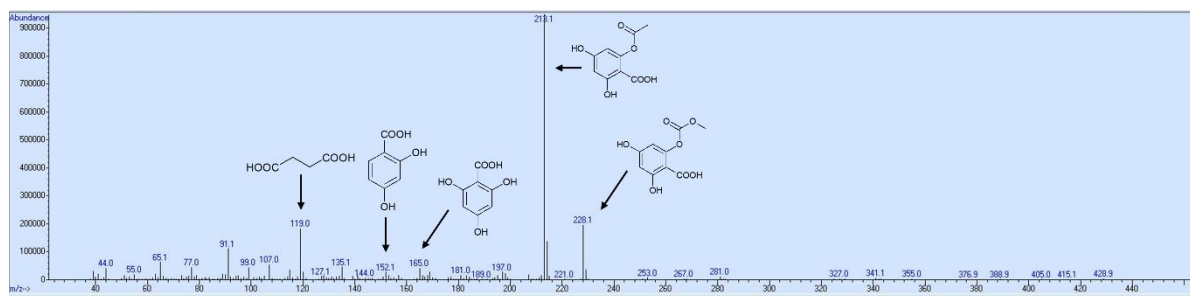


Fig. S4 Mass spectrum of morin after the oxidation by HA-AgNPs in the presence of H₂O₂, collected by GC-MS.

6. Real-time video of the oxidation of morin in the presence of H₂O₂ and HA-AgNPs



Video S1.rar