

Supporting Information

The effect of side chain spacer length on the thermoresponsive behavior of poly(methylamide acrylate)s

Alexander Rajakanthan^a, Paul Wilson ^{b*} and Kristian Kempe ^{a,c*}

^a Monash Institute of Pharmaceutical Sciences, Monash University, 381 Royal Parade, Parkville VIC 3052, Australia

^b Department of Chemistry, The University of Warwick, CV4 7AL, United Kingdom

^c Department of Material Science and Engineering, Monash University, Clayton, VIC 3800, Australia

Equation S1. Mathematic equation to calculate monomer self-association calculating, K_a .

$$\Delta\delta_{obs} = \frac{\Delta\delta_{max} \left\{ (2[M] + K_d) - [(2[M] + K_d)^2 - 4[M]^2]^{\frac{1}{2}} \right\}}{2[M]}$$

Where $K_d = [M]^2/[D]$ and $K_a = 1/K_d$. K_d , $[M]$ and $[D]$ are dissociation constants, monomer and dimer concentrations (mol L^{-1} , M) respectively.

Table S1. RAFT conditions used to prepare poly(methylamide acrylate) homopolymers (time: 24 h).

Sample	Target DP	Conversion (%)	[M] (mol L^{-1})	[BCTP] (mol L^{-1})	[tBuOOH] ($\mu\text{mol L}^{-1}$)	[AsAc] ($\mu\text{mol L}^{-1}$)
P(MAmEA) ₁₀₀	100	100	3.09	0.031	7.7	3.9
P(MAmPA) ₉₇	100	97	3.09	0.031	7.7	3.9
P(MAmBA) ₂₅	25	100	3.09	0.124	30.9	15.5
P(MAmBA) ₄₈	50	96	3.09	0.062	15.5	7.7
P(MAmBA) ₉₉	100	99	3.09	0.031	7.7	3.9
P(MAmBA) ₁₉₀	200	95	3.09	0.015	3.9	1.9

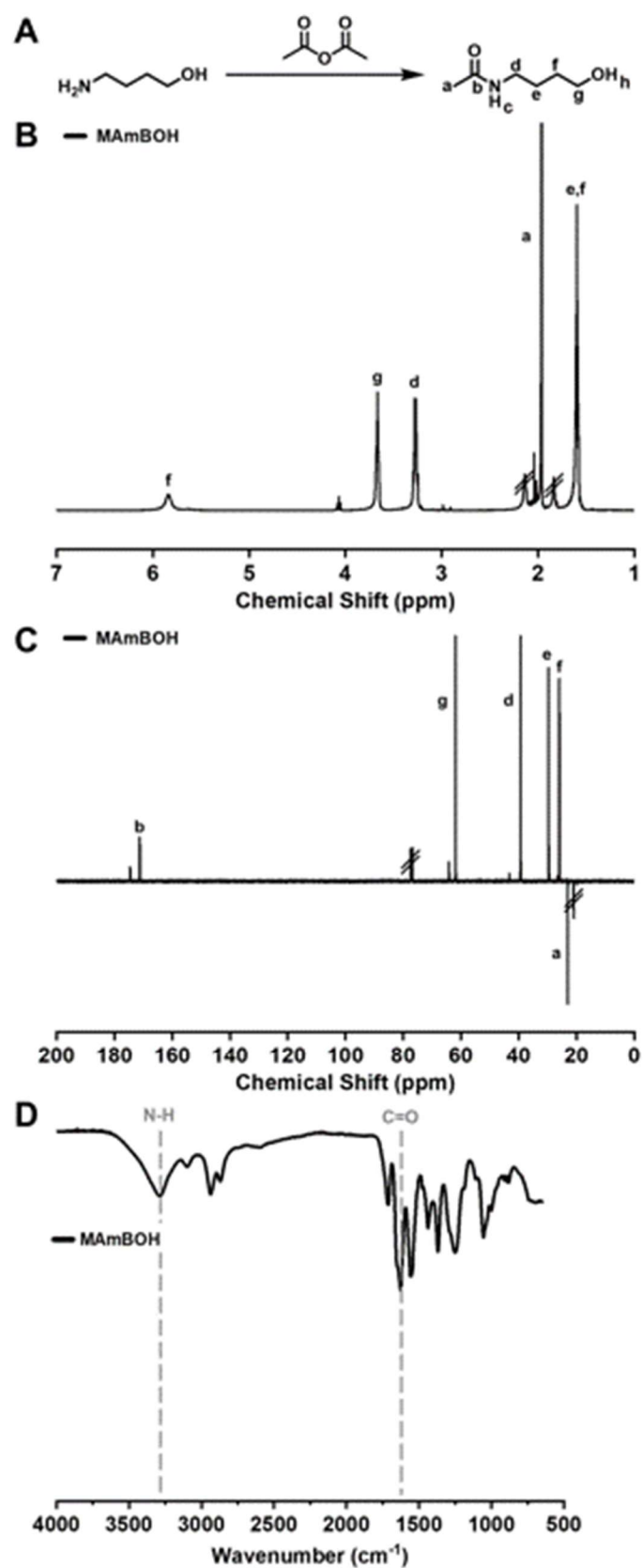


Figure S1. (A) Synthesis pathway of MAmBOH; (B) ^1H NMR spectrum (400 MHz, CDCl_3) of MAmBOH; (C) ^{13}C NMR spectrum (400 MHz, CDCl_3) of MAmBOH; (D) FT-IR spectrum of MAmBOH.

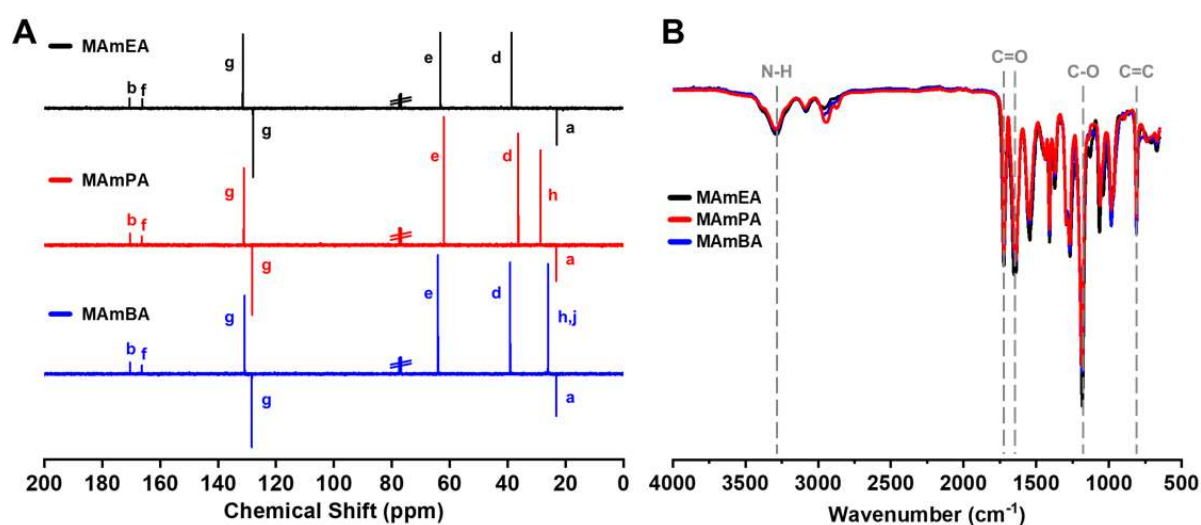


Figure S2. (A) ^{13}C NMR spectra (400 MHz, CDCl_3) of MAmEA, MAmPA and MAmBA; (B) FT-IR spectra of MAmEA, MAmPA and MAmBA.

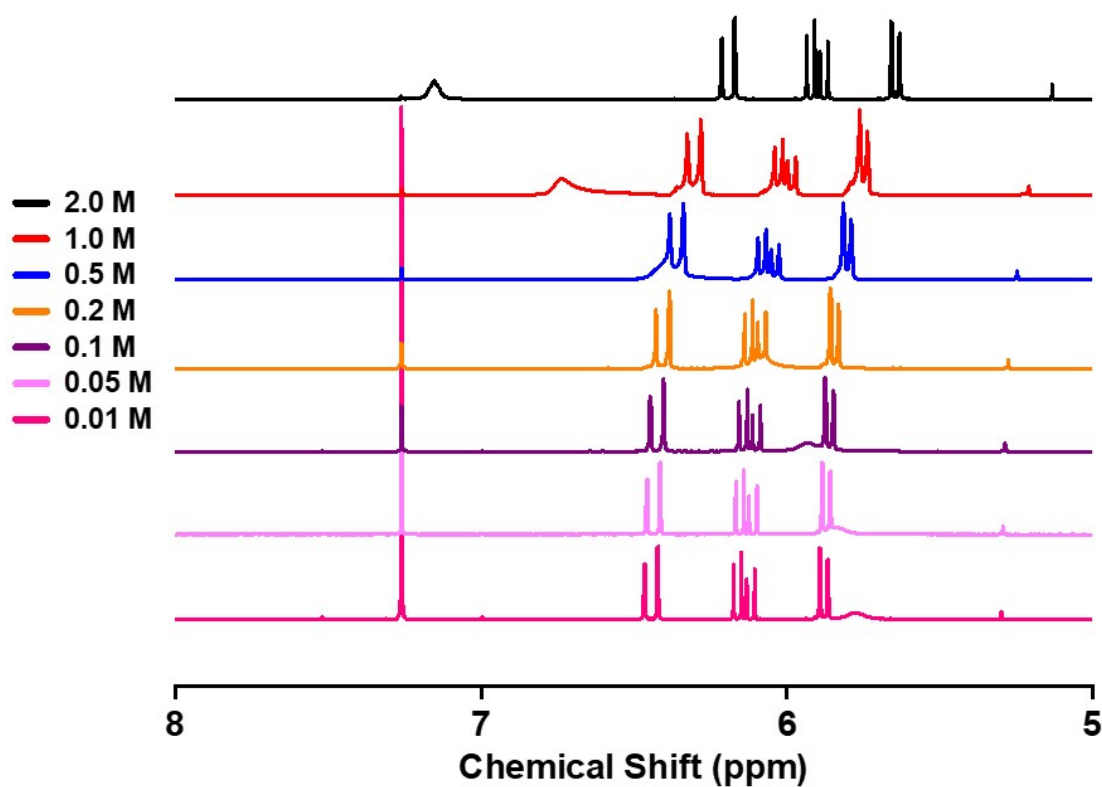


Figure S3. ^1H NMR spectra (400 MHz, CDCl_3) of MAmEA at varying $[\text{M}]$, highlighting the concentration-dependent change in chemical shift of H_c .

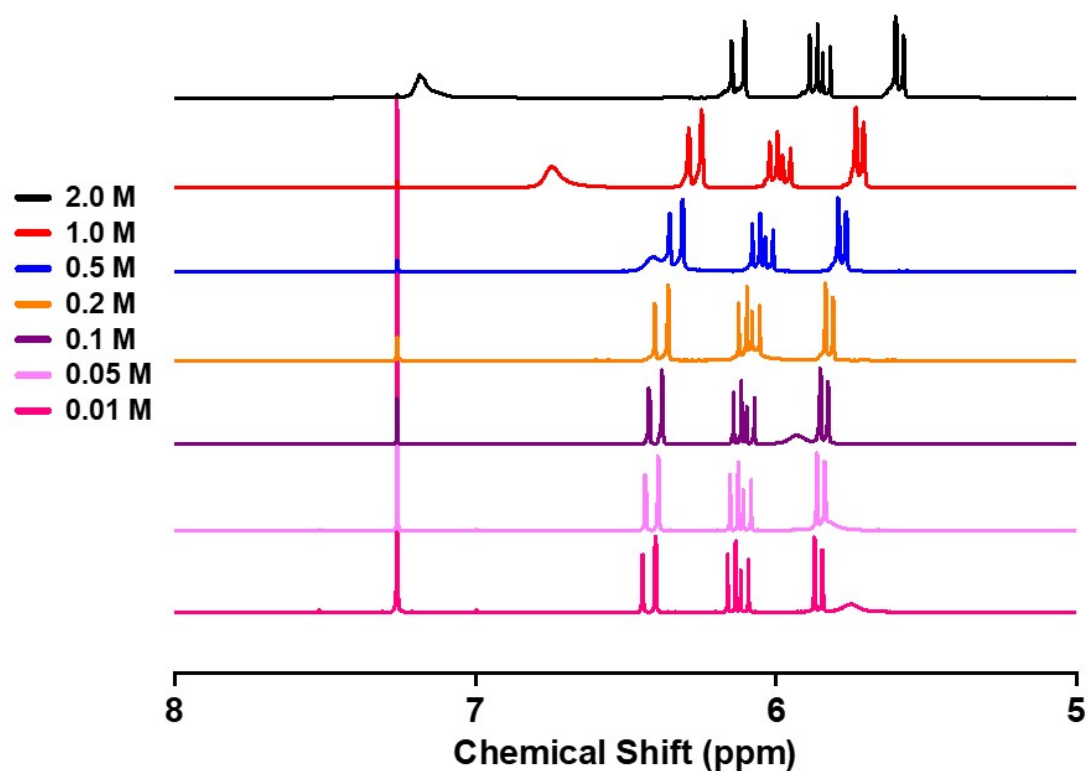


Figure S4. ^1H NMR spectra (400 MHz, CDCl_3) of MAmPA at varying $[\text{M}]$, highlighting the concentration-dependent change in chemical shift of H_c .

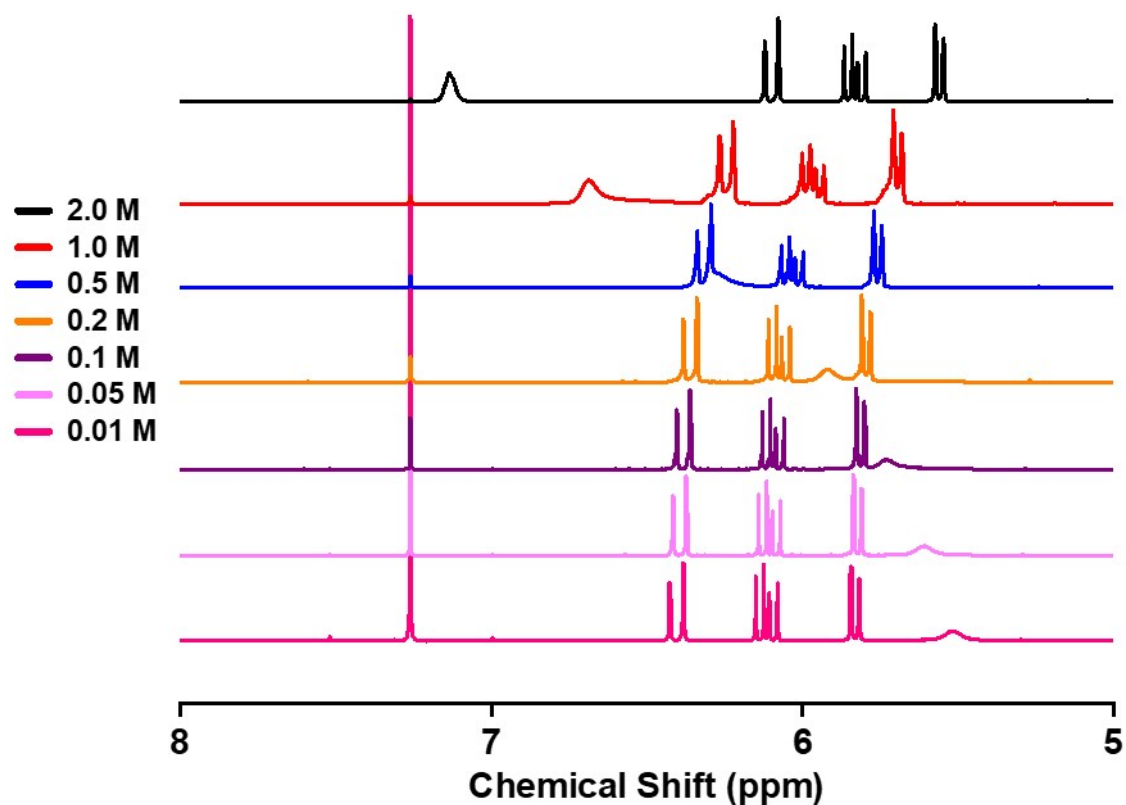


Figure S5. ^1H NMR spectra (400 MHz, CDCl_3) of MAmBA at varying $[\text{M}]$, highlighting the concentration-dependent change in chemical shift of H_c .

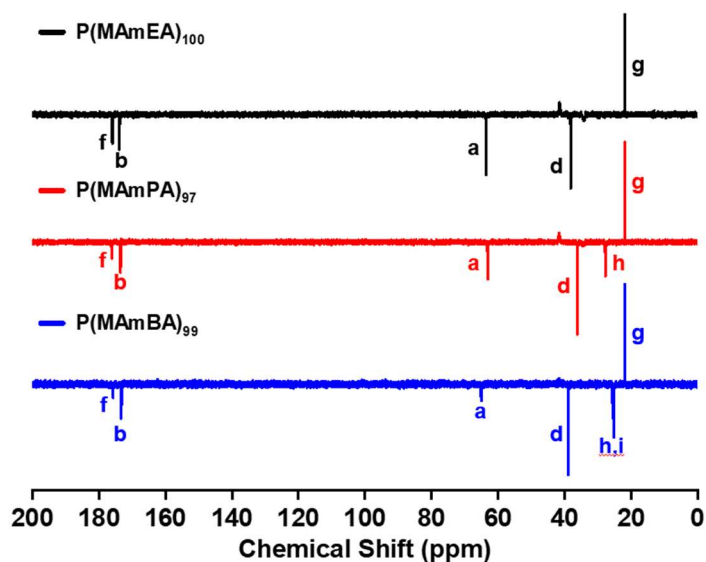


Figure S6. ^{13}C NMR spectra (400 MHz, D_2O) of P(MAmEA)_{100} , P(MAmPA)_{97} and P(MAmBA)_{99} .

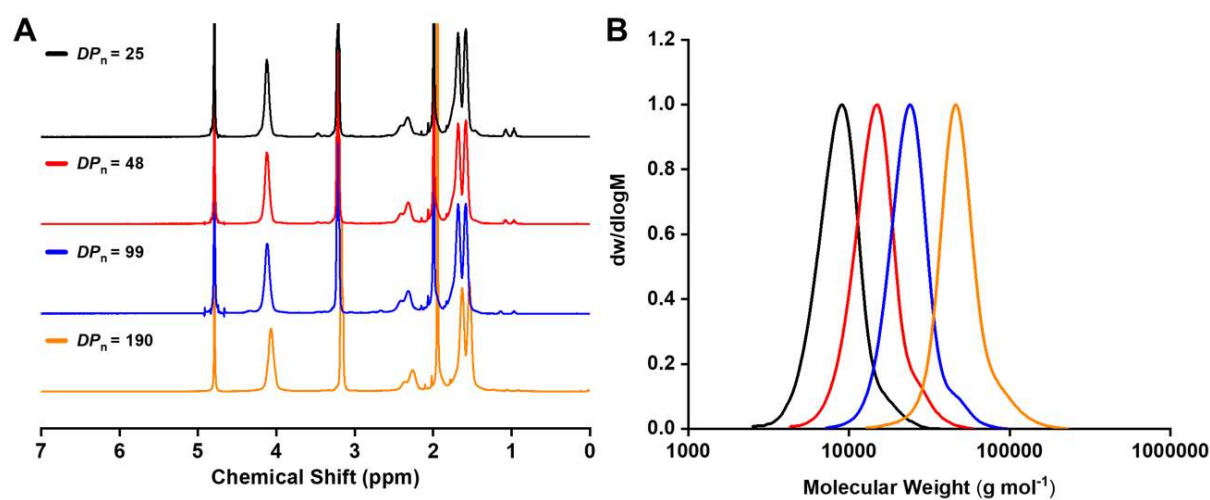


Figure S7. (A) ^1H NMR spectra (400 MHz, D_2O) of P(MAmBA)_{25} , P(MAmBA)_{48} , P(MAmBA)_{99} and P(MAmBA)_{190} ; (B) SEC traces of P(MAmBA)_{25} ($M_n = 8,200$, $\bar{D}_M = 1.12$), P(MAmBA)_{48} ($M_n = 12,800$, $\bar{D}_M = 1.12$), P(MAmBA)_{99} ($M_n = 22,700$, $\bar{D}_M = 1.11$) and P(MAmBA)_{190} ($M_n = 48,300$, $\bar{D}_M = 1.12$).

Table S2. Conditions used for turbidity measurements of P(MAmBA) and the associated cloud point temperatures (T_{cp}). Temperature range measured: 30 - 80 °C.

DP_n	Concentration (mg ml ⁻¹)	Media	T_{cp} (° C)
190	5	PBS	39.6
99	5	PBS	52.8
48	5	PBS	66.0
25	5	PBS	N/A
99	1	PBS	57.2
99	0.1	PBS	N/A
99	5	0.137 M NaCl	59.6
99	5	Milli-Q Water	65.8

**Uncoated
(Control)
 $59.0 \pm 2.7^\circ$**

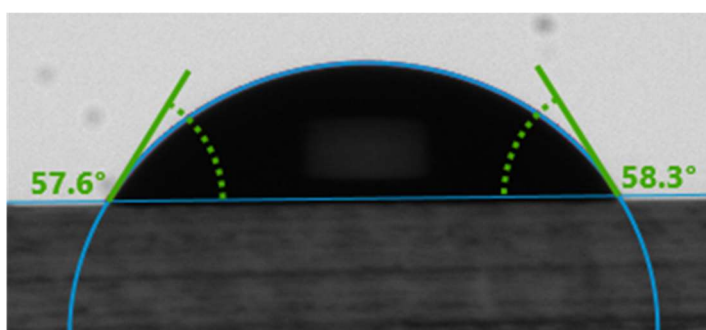


Figure S8. Water contact angle image of an uncoated glass slide.