Unraveling the effect of carbon morphology evolution in hard

carbons on sodium storage performance

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Fig. S1 SEM and TEM images of CHC-T, (a, e) CHC-700, (b, f) CHC-1100, (c, g) CHC-1300 and (d, h) CHC-1500.



Fig. S2 Interlayer distance of CHC-T, (a) CHC-700, (b) CHC-1100, (c) CHC-1300 and (d) CHC-1500.



Fig. S3 The proportion of highly disordered domains, pseudo-graphitic domains and graphite-like domains for CHC-T.



Fig. S4 (a) N₂ adsorption-desorption isotherms, (b) pore size distribution curves.



Fig. S5 High-resolution C 1s, O 1s and N 1s XPS spectra of (a-c) CHC-700 and (d-f) CHC-1100.



Fig. S6 Electron configurations for pyridinic N, pyrrolic N, and graphitic N.



Fig. S7 Rate performance images of (a) CHC-700, (b) CHC-1100, (c) CHC-1300 and (d) CHC-1500.



Fig. S8 Electrochemical performance images of (a) CHC-900, (b) CHC-1200 and (c) CHC-1400.



Fig. S9 Comparison of rate capability between CHC-1300 and ever reported carbon anode materials in literatures.



Fig. S10 CV curves at different scan rates, (a) CHC-700, (b) CHC-1100, (c) CHC-1300 and (d) CHC-1500.



Fig. S11 *E vs. t* profile for one GITT test. D_{Na^+} was predicted by the following equation:

$$D = \frac{4}{\pi\tau} \left(\frac{n_m V_m}{s}\right)^2 \left(\frac{\Delta E_s}{\Delta E_\tau}\right)^2$$

Where n_m is the amount of active substance of electrode material, V_m is the molar volume, and *S* represents geometric area. $\triangle E_s$ and $\triangle E_{\tau}$ can be obtained from the GITT curves.

Sample	20 (°)	d(002) (nm)	L _a (nm)	L _c (nm)
CHC-700	22.9	0.388	1.86	1.03
CHC-1100	23.6	0.376	2.33	1.09
CHC-1300	24.3	0.365	2.84	1.14
CHC-1500	25.1	0.355	3.02	1.21

Table S1. Physical parameters of CHC-T from XRD

Table S2. Specific surface area and pore diameter of CHC-T from BET

Sample	$S_{BET} (m^2 g^{-1})$	d _{pore} (nm)
CHC-700	0.928	3.08
CHC-1100	9.068	3.17
CHC-1300	4.729	3.00
CHC-1500	4.490	3.44

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