

**In Situ Built Nanoconfined Nb₂O₅ Particles in 3D Interconnected Nb₂C
MXene@rGO Conductive Framework for High-Performance Potassium-Ion
Batteries**

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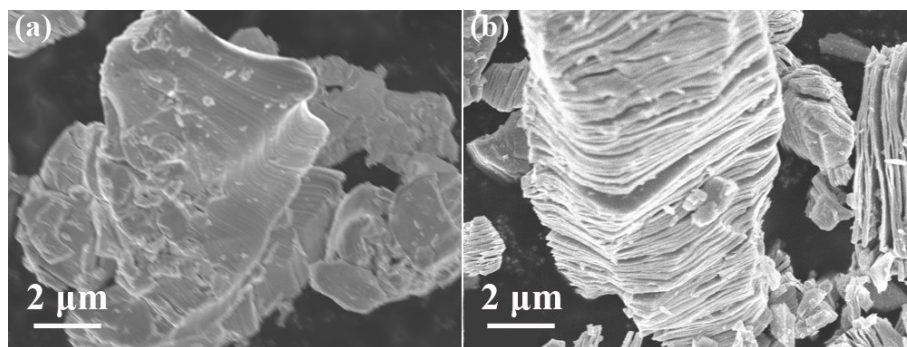


Figure S1 Typical SEM images of (a) Nb₂AlC and (b) multi-layered Nb₂CT_x.

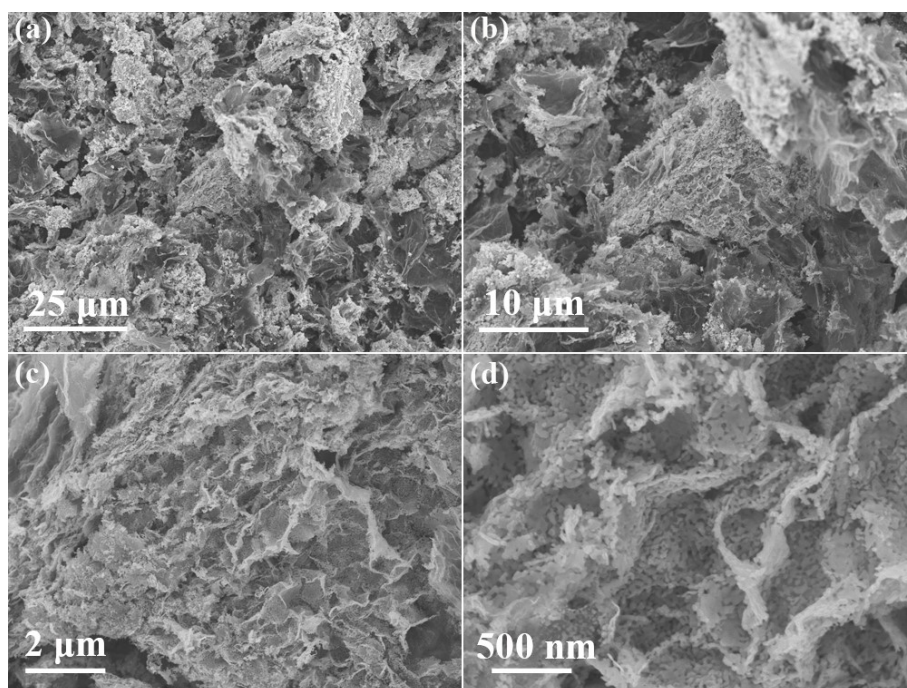


Figure S2 Typical SEM images of Nb₂O₅/rGO.

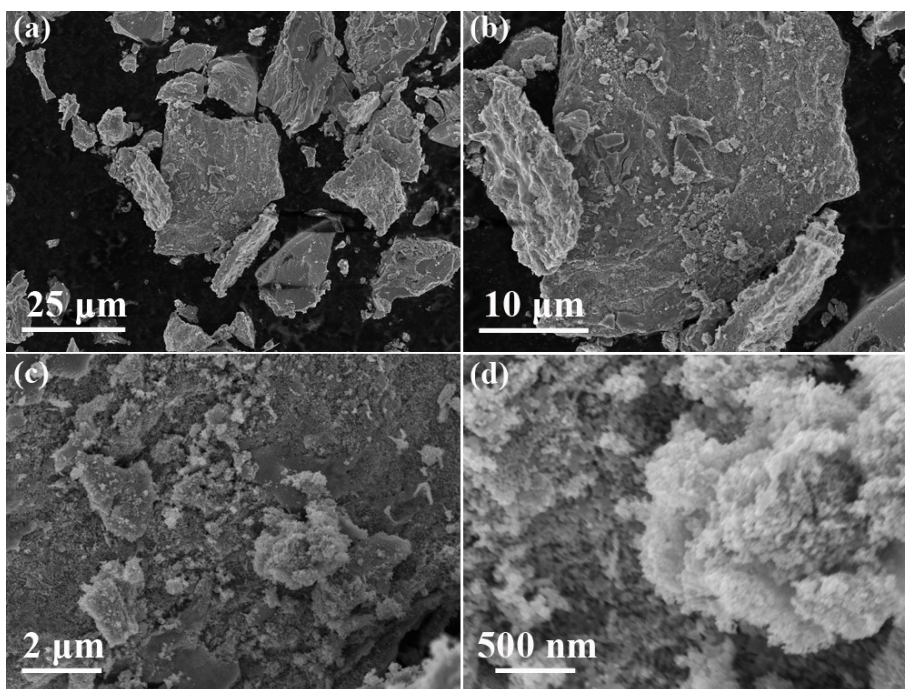


Figure S3 Typical SEM images of Nb_2O_5 .

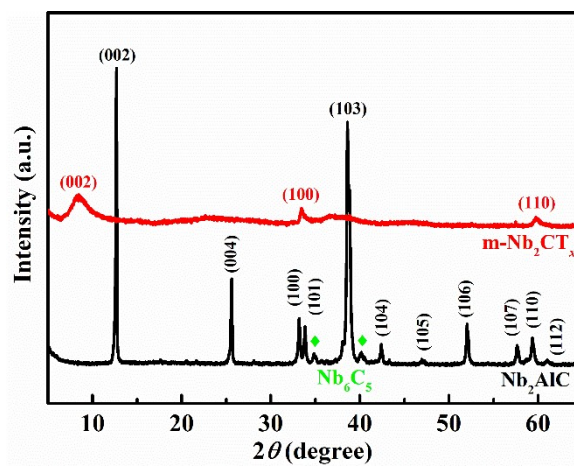


Figure S4 XRD patterns of Nb_2AlC and $m\text{-Nb}_2\text{CT}_x$.

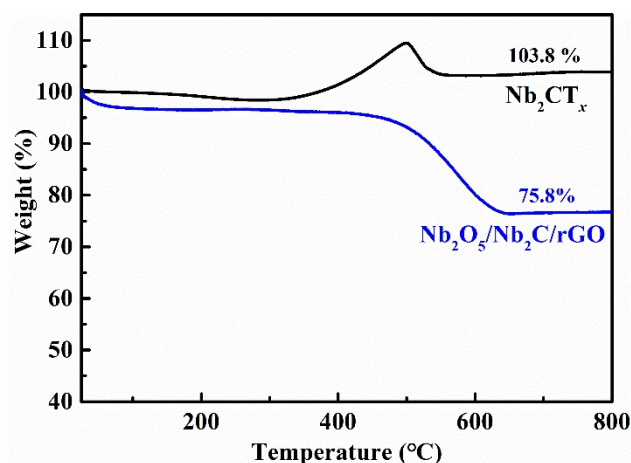


Figure S5 TGA curves of Nb_2CT_x nanosheets and $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}$ aerogel.

As shown in **Fig. S5**, the final stable weights of Nb_2CT_x nanosheets and $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}$ are 103.8 wt% and 75.8 wt%, respectively. The weight change of pure Nb_2CT_x nanosheets is $(103.8-100)$ wt% = +3.8 wt%. The weight change of $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}$ is $(75.8-100)$ wt% = -24.2 wt%. There is no doubt that the weight change of pure rGO is -100 wt% (*i.e.*, fully oxidized/decomposed into gaseous products). Therefore, the content of rGO in $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}$ can be calculated based on the formula $(1-A) \times 3.8\% - A \times 100\% = B$, where A is the content of rGO in $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}$, and B is the weight change of $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}$ (*i.e.*, -24.2 wt%). Hence, according to TGA results, the content of rGO in $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}$ sample is calculated to be 27.0 wt%.

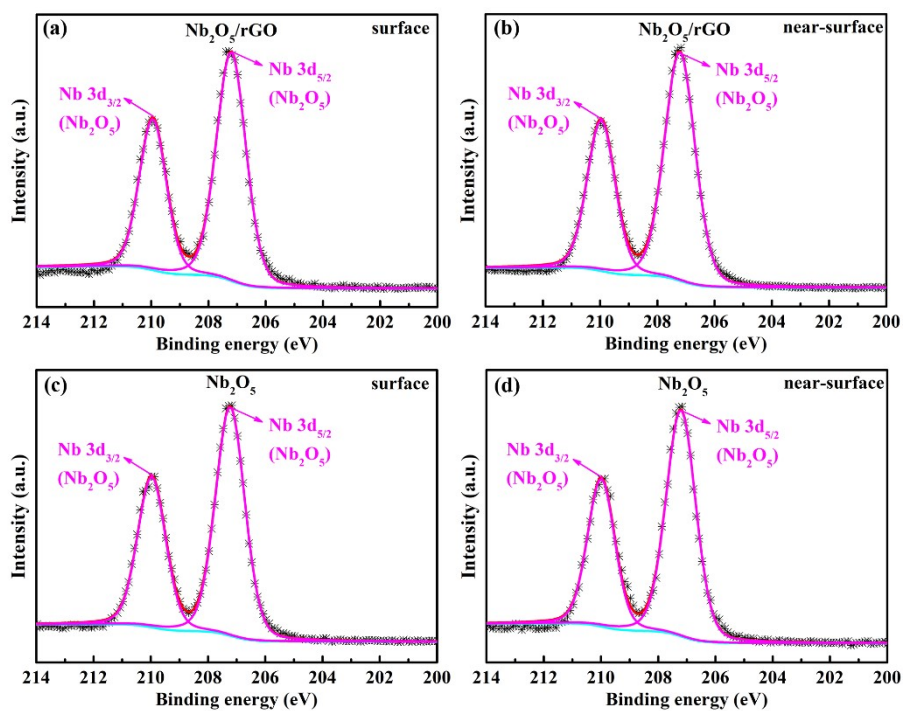


Figure S6 Refined high-resolution Nb 3d XPS spectra of (a, b) $\text{Nb}_2\text{O}_5/\text{rGO}$ and (c, d) Nb_2O_5 before (a, c) and after (b, d) Ar^+ ion sputtering.

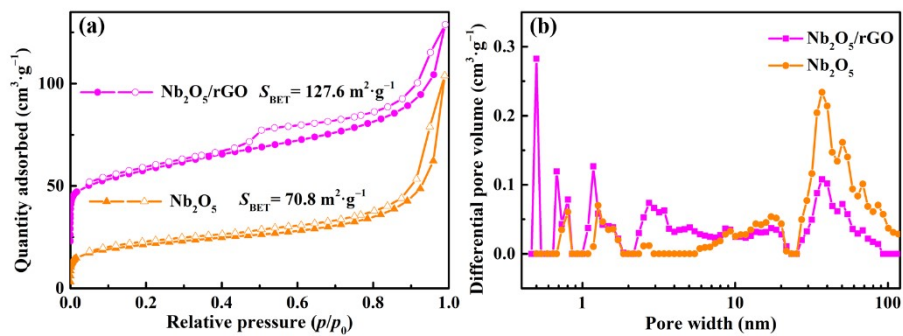


Figure S7 (a) N_2 adsorption and desorption isotherms and (b) the corresponding pore-size distribution curves of $\text{Nb}_2\text{O}_5/\text{rGO}$ and Nb_2O_5 .

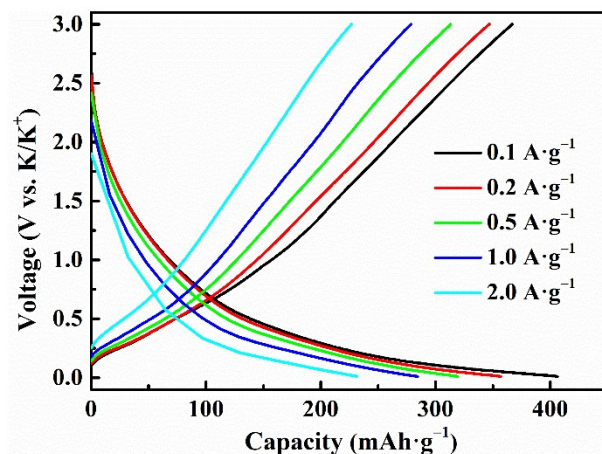


Figure S8 GCD curves of Nb₂O₅/Nb₂C/rGO at different current densities.

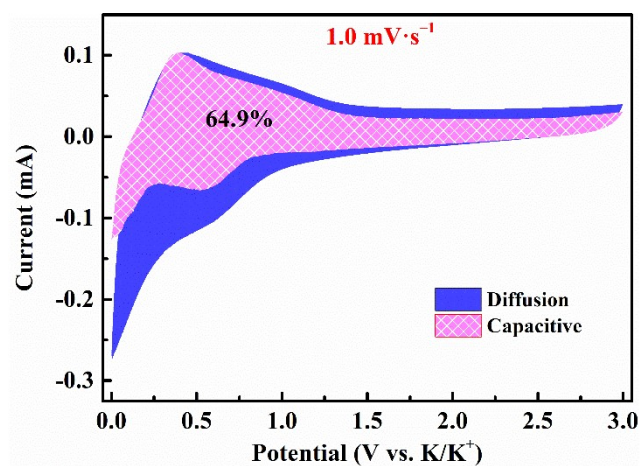


Figure S9 Capacitive- and diffusion-controlled contributions at $1.0 \text{ mV}\cdot\text{s}^{-1}$ for the Nb₂O₅/Nb₂C/rGO.

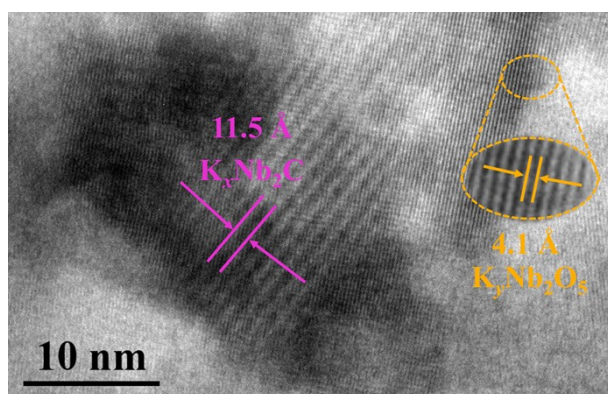


Figure S10 HRTEM of $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}$ electrode after being discharged to 0.01V.

HRTEM test was performed on the electrode after being discharged to 0.01 V to confirm the evolution of structure. As shown in **Figure 2i** and **Figure S10**, it can be observed that the interlayer distance of Nb_2C is increased from 10.4 to 11.5 Å and that of Nb_2O_5 is increased from 3.9 Å to 4.1 Å, indicating that K^+ is successfully intercalated into Nb_2C and Nb_2O_5 layers and thus the formation of $\text{K}_x\text{Nb}_2\text{C}$ and $\text{K}_y\text{Nb}_2\text{O}_5$.

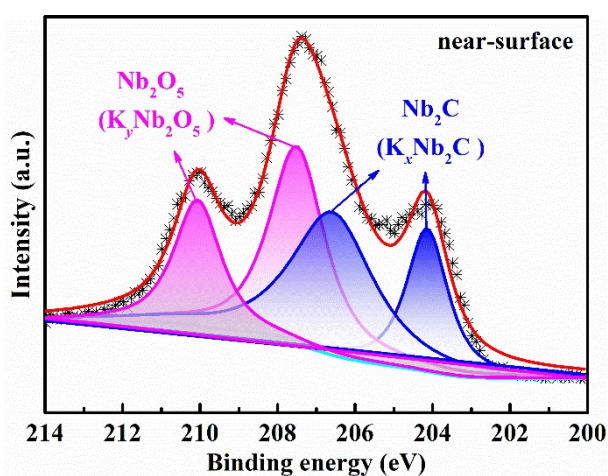


Figure S11 Nb 3d XPS spectrum for the near-surface of $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}$.

As shown in **Figure S11**, the near-surface of $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}$ exhibits four peaks at 210.0, 207.3, 206.4 and 204.1 eV, respectively. Combined with **Figure 5e**, the first two peaks are assigned to Nb_2O_5 ($\text{K}_y\text{Nb}_2\text{O}_5$), while the last two peaks can be indexed to Nb_2C ($\text{K}_x\text{Nb}_2\text{C}$).

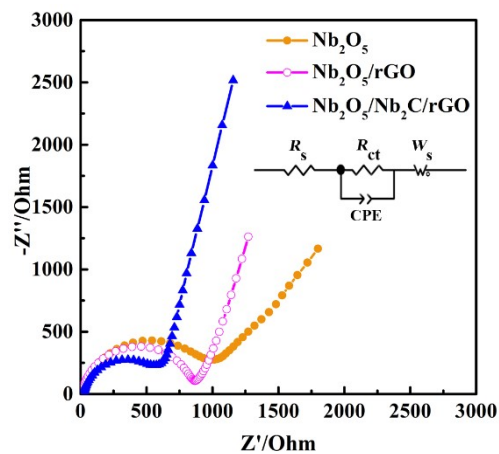


Figure S12 EIS spectra for $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}$, $\text{Nb}_2\text{O}_5/\text{rGO}$ and Nb_2O_5 .

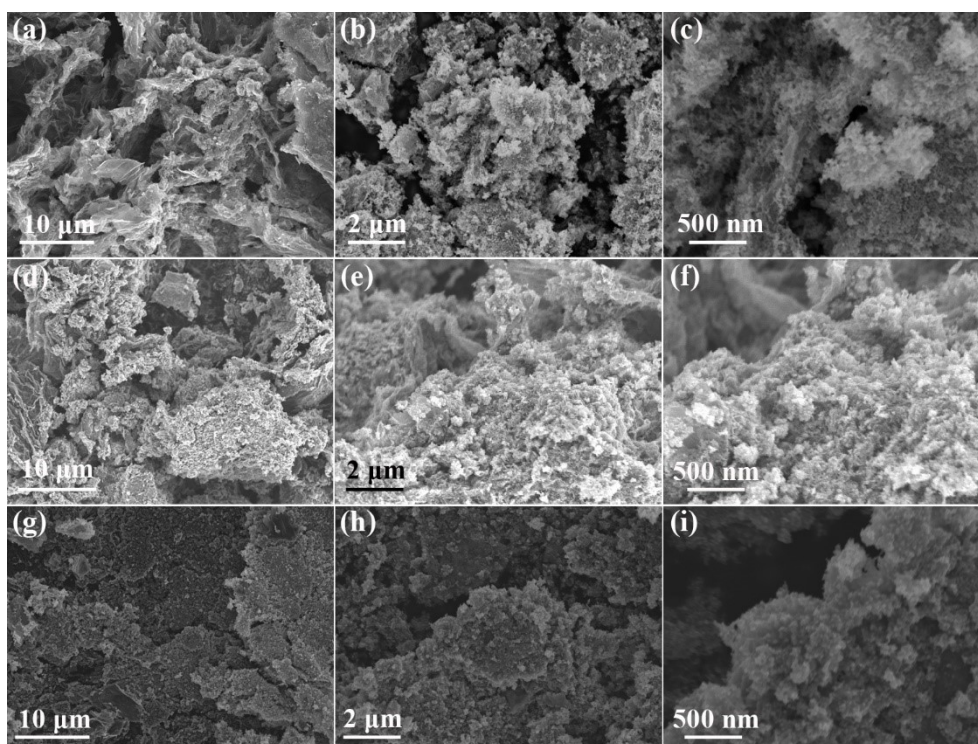


Figure S13 SEM images of (a-c) $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}$, (d-f) $\text{Nb}_2\text{O}_5/\text{rGO}$ and (g-i) Nb_2O_5 electrodes after 100 cycles at a current density of $0.1 \text{ A} \cdot \text{g}^{-1}$.

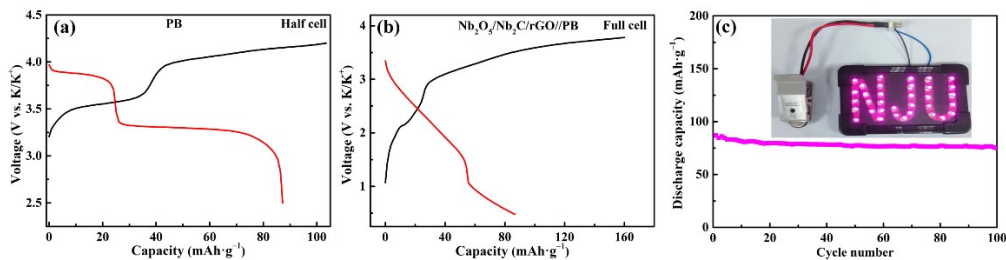


Figure S14 GCD curves of (a) PB in half-cell and (b) $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}/\text{PB}$ full-cell at a current density $0.1 \text{ A}\cdot\text{g}^{-1}$; (c) cycle performance of $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}/\text{PB}$ full-cell at a current density of $0.1 \text{ A}\cdot\text{g}^{-1}$ (inset shows LEDs powered by the full cell).

PB was prepared by a simple precipitation method. **Figure S14a** shows the GCD curves of PB in half-cell with charge and discharge capacities of 103.5 and $87.2 \text{ mAh}\cdot\text{g}^{-1}$, respectively. Before assembling the full-cells, the charge balance between anode and cathode needs to be optimized by controlling the mass ratio of anode and cathode based on the specific discharge capacities of $\text{Nb}_2\text{O}_5/\text{Nb}_2\text{C}/\text{rGO}$ and PB in half-cells.

Table S1 A detailed comparison of the electrochemical performance of different anode materials for KIBs.

Samples	Capacity ($\text{mA}\cdot\text{g}^{-1}$) @rate ($\text{A}\cdot\text{g}^{-1}$)	Capacity retention @rate ($\text{A}\cdot\text{g}^{-1}$)@cycles	Ref.
Sn/rGO	222.4 @ 0.1 67.1 @ 2	69% @ 0.5 @ 500	[1]
MoS ₂ /graphene	511 @ 0.1 234 @ 2	75% @ 1 @ 800	[2]
N-Doped graphene	320 @ 0.05 170 @ 0.5	88% @ 0.5 @ 500	[3]
TiO _x N _y /C	150 @ 0.2 75 @ 1.6	23% @ 0.2 @ 120	[4]
Onion-like carbon	245 @ 0.05 78 @10	71% @ 2 @ 980	[5]
TiS ₂	124 @ 0.05 92.1 @ 0.25	64% @ 0.05 @ 450	[6]
Co ₃ O ₄ -F ₂ O ₃ /C	420 @ 0.05 278 @ 1	60% @ 0.05 @ 100	[7]
KTiOPO ₄	102 @ 0.005 50 @ 0.2	77% @ 0.005 @ 200	[8]
BiOCl	367 @ 50 175 @ 1	58% @ 0.05 @ 50	[9]
Sn ₄ P ₃ /MGS	378 @ 0.1 113 @ 5	77% @ 0.5 @ 1000	[10]
Nb ₂ O ₅ /Nb ₂ C/rGO-2	410 @ 0.1 230 @ 2	89% @ 2 @ 1000	This work

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