| 1  | Supplementary Information  |
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| 9  | A validated analytical method to measure metals dissolved in deep        |
| 10 | eutectic solvents  |
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Figure S1: Calibration curve data at the 0.01, 0.04, 0.1, 0.4, 1  $\mu$ g/mL concentration levels, which were contracted in Figure 2, for Li, Mg, Fe, Co, Ni, Cu, Zn, Pd, Al, Sn, and Pb in ChCl:EG:I<sub>2</sub>. The markers represent the data points corresponding to the intensity-concentration values of each sample. The black straight line represents the linear part of each calibration curve, as obtained from the linear regression equations, R<sup>2</sup> values, and the F-test results.

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35 Figure S2: Inter-run precision of the analytical method developed for the determination of Cu in

36 ChCl:EG:I<sub>2</sub>. Measurements of 40 samples of 2  $\mu$ g/mL were obtained at different days in a 2-month

37 period and the accuracy (as % recovery) of the analytical method for each element was estimated.

38 The acceptability criterion for the recovery was within 90–110%.

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Figure S3: Calibration curve data at the 0.01, 0.04, 0.1, 0.4, 1  $\mu$ g/mL concentration levels, which were contracted in Figure 5, for Li, Co, Ni, Cu, and Al in ChCl:EG. The markers represent the data points corresponding to the intensity-concentration values of each sample. The black straight line represents the linear part of each calibration curve, as obtained from the linear regression equations, R<sup>2</sup> values, and the F-test results.





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49 **Figure S4:** Calibration curve data at the 0.01, 0.04, 0.1, 0.4, 1  $\mu$ g/mL concentration levels, which 50 were contracted in Figure 6, for Li, Co, Ni, Cu, and Al in ChCl:LA. The markers represent the data 51 points corresponding to the intensity-concentration values of each sample. The black straight line 52 represents the linear part of each calibration curve, as obtained from the linear regression

53 equations,  $R^2$  values, and the F-test results.