## Supporting information for: Study on the reaction pathway and mechanism of urea alcoholysis by the disassociation and conjugation of groups

Jie Wang<sup>a,b</sup>, Jielin Huang<sup>a,b</sup>, Songsong Chen<sup>a,b</sup>, Junping Zhang<sup>b,c</sup>, Li Dong<sup>b,c,\*</sup>,

Xiangping Zhang<sup>b,c</sup>

<sup>a</sup> School of Chemical Engineering, University of Chinese Academy of Sciences, Beijing 100049, China

<sup>b</sup> Beijing Key Laboratory of Ionic Liquids Clean Process, CAS Key Laboratory of Green Process and Engineering, State Key Laboratory of Multiphase Complex Systems, Institute of Process Engineering, Chinese Academy of Sciences, Beijing 100190, China

<sup>c</sup> Longzihu New Energy Laboratory, Zhengzhou Institute of Emerging Industrial Technology, Henan University, Zhengzhou 450000, China



Figure S1 In-situ FT-IR for urea alcoholysis catalyzed by catalyst [P<sub>4444</sub>][Br].



Figure S2 In-situ FT-IR for urea alcoholysis catalyzed by catalyst ZnBr<sub>2</sub>.



Figure S3 In-situ FT-IR for urea alcoholysis catalyzed by the binary catalyst [P<sub>4444</sub>][Br]/ZnBr<sub>2</sub>.



Figure S4 The trend of temperature changes over time during the reaction process.



Figure S5 Reaction mechanism for the formation of 2-OX using  $[P_{4444}][Br]/ZnBr_2$  as catalyst.