

Supporting Information for

Tuning the Selectivity of Benzylamine Photo-oxidation with Different Rhodium Modes Anchoring on BiOIO₃

Jiaping Liu,^{‡a} Yan Wu,^{‡a} Qingqing Chen,^{‡a} Rui Yu,^a Keqing Shi,^b Tao Jing,^d

Zhujie Li,^{b*} Zaizhu Lou,^{c*} and Gang Wang^{a*}

Figs and Tables

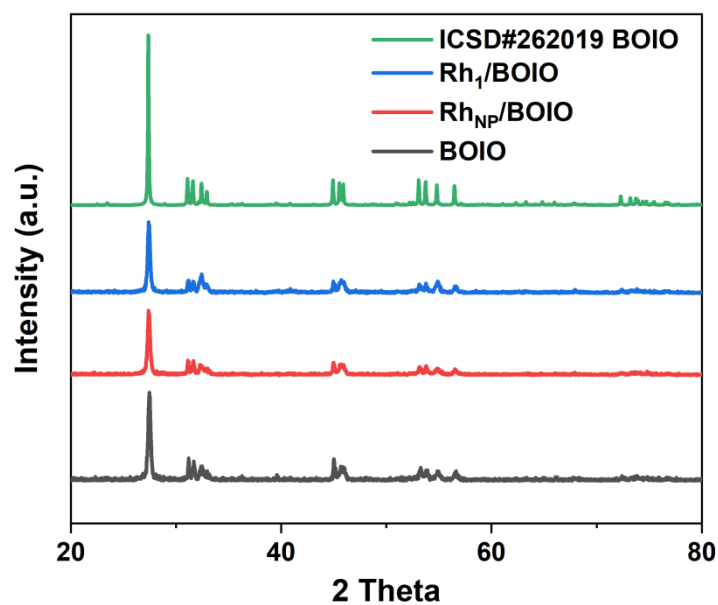


Fig. S1 XRD patterns of the Rh₁/BOIO, Rh_{NP}/BOIO, BOIO and the standard BOIO.

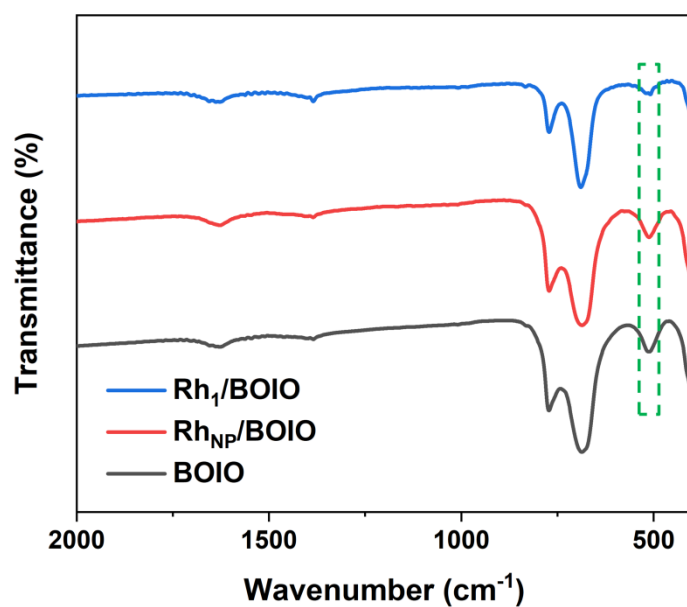


Fig. S2 The Fourier transforms infrared spectroscopy (FT-IR) of samples. The peak at around 514 cm⁻¹ is the stretching vibration of Bi-O of BOIO (green box). In Rh₁/BOIO, the characteristic peak of Bi-O has been reduced, which could attribute to the formation of Rh-O bond.

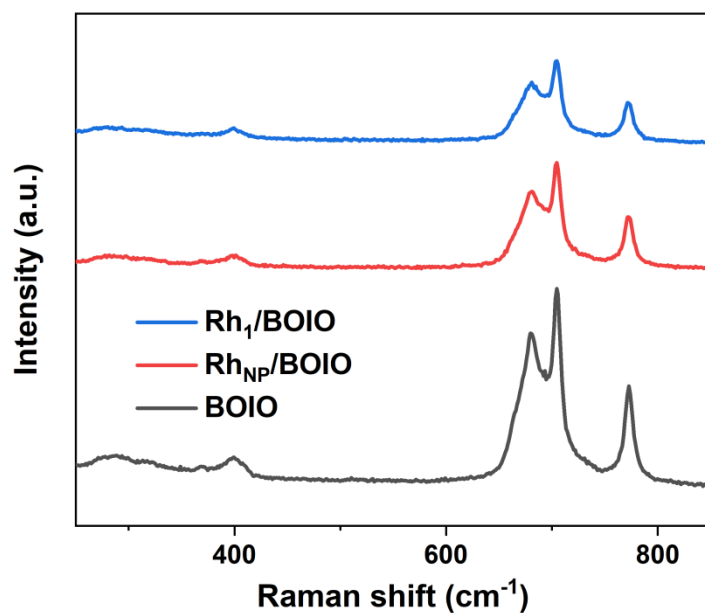


Fig. S3 The Raman spectra of Rh₁/BOIO, Rh_{NP}/BOIO and BOIO.

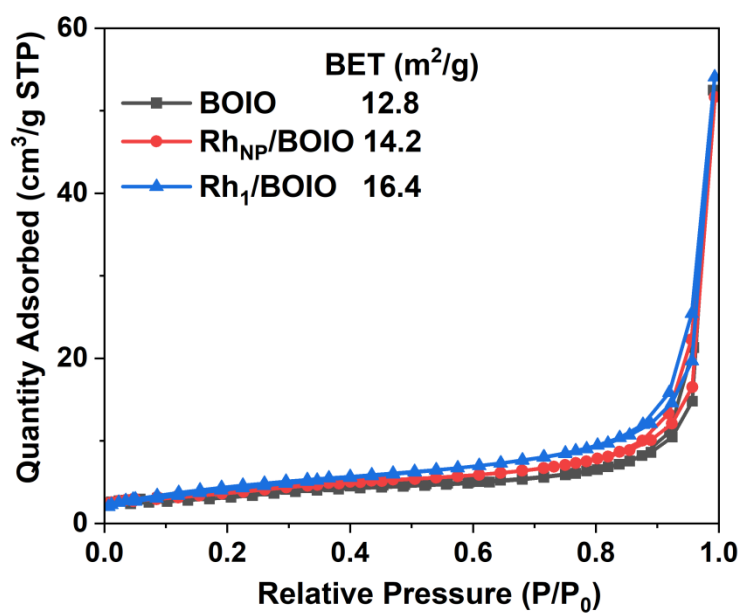


Fig. S4 Nitrogen adsorption and desorption isotherms of Rh₁/BOIO, Rh_{NP}/BOIO and BOIO.

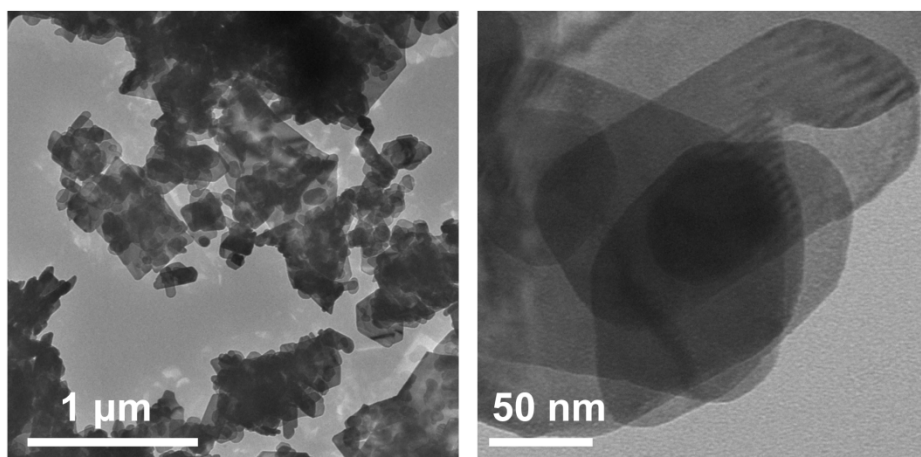


Fig. S5 TEM images of BOIO.

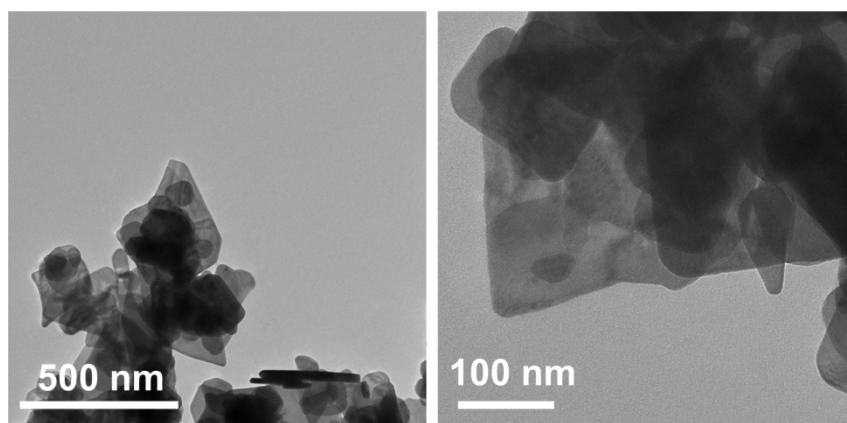


Fig. S6 TEM images of Rh₁/BOIO.

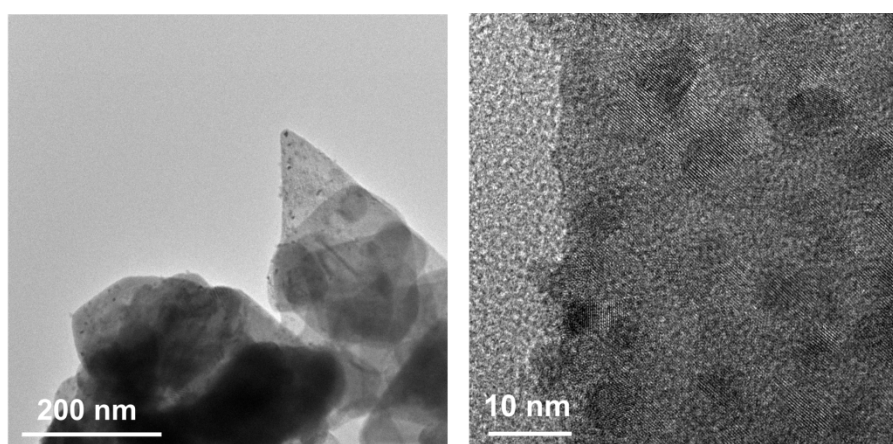


Fig. S7 TEM images of Rh_{NP}/BOIO.

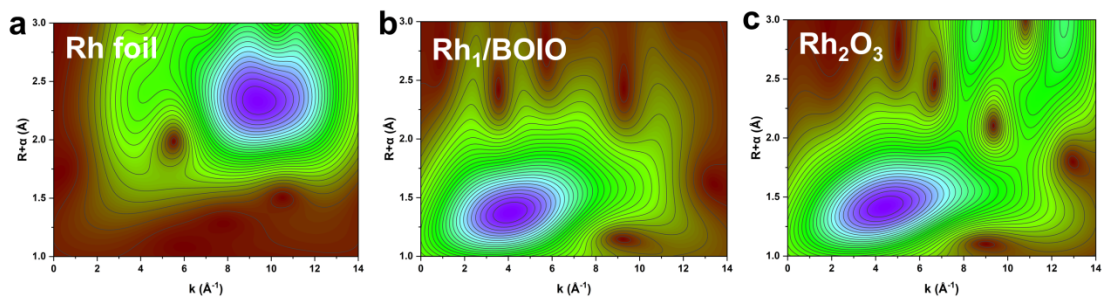


Fig. S8 WT plots of samples. (a) Rh foil, (b) Rh_1/BOIO and (c) Rh_2O_3 .

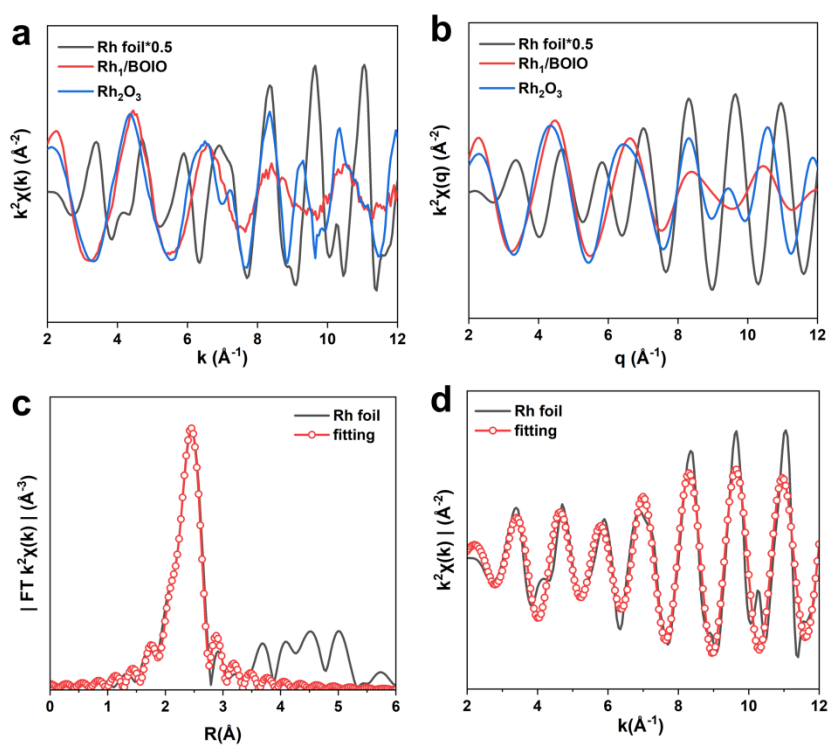


Fig. S9 (a) The k -space spectra of EXAFS of Rh foil, Rh_2O_3 and Rh_1/BOIO . (b) The q -space spectra of EXAFS of Rh foil, Rh_2O_3 and Rh_1/BOIO . (c) R -space fitting results of Rh foil. (d) k -space fitting results of Rh foil.

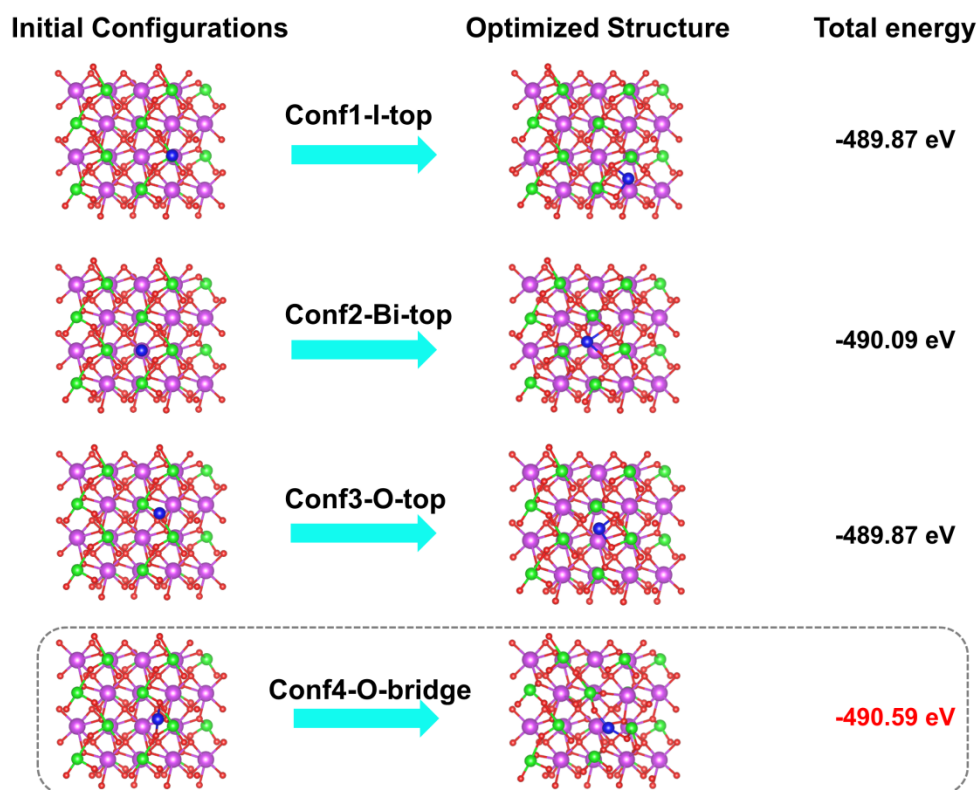


Fig. S10 The optimized structure of Rh_1/BOIO . These are four different initial configurations have been established, Conf1-I-top, Conf2-Bi-top, Conf3-O-top and Conf4-O-bridge. The most stable configuration of Conf4-O-bridge can be confirmed after structural optimization. purple ball: Bi, green ball: I, red ball: O, blue ball: Rh.

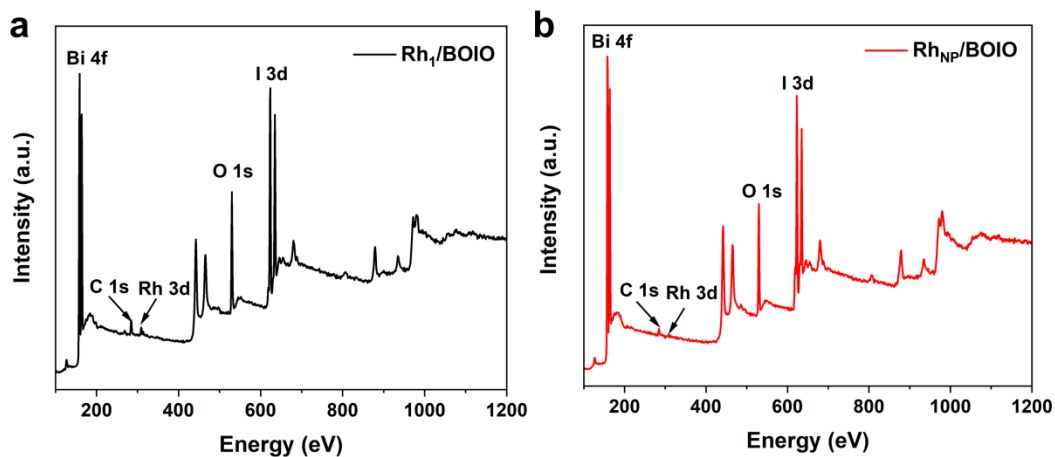


Fig. S11 XPS survey spectrum of the as-prepared (a) Rh₁/BOIO and (b) Rh_{NP}/BOIO.

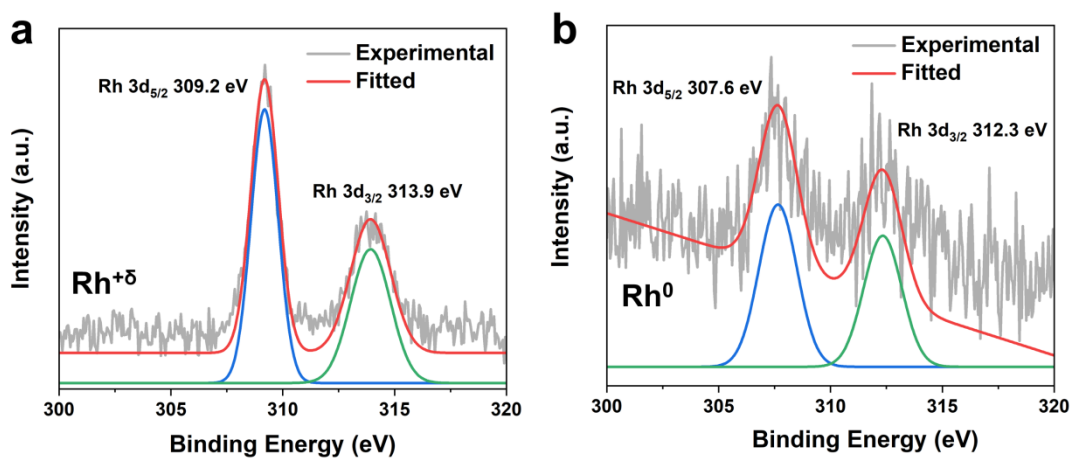


Fig. S12 XPS results of Rh 3d of (a) Rh₁/BOIO (0 < δ < 3) and (b) Rh_{NP}/BOIO.

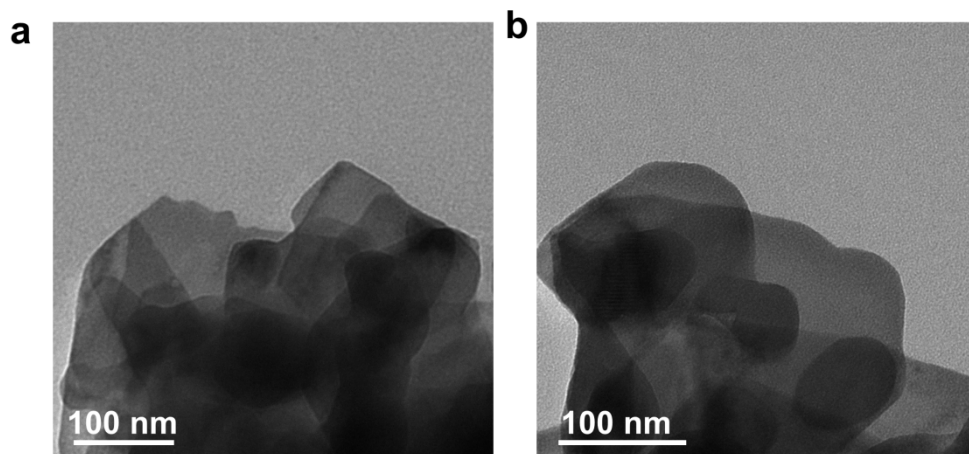


Fig. S13 TEM images of (a) Rh₁-0.34wt% and (b) Rh₁-1.04wt%. Rh₁-0.34wt% sample was obtained after photo-deposition process between 61 μL RhCl₃•xH₂O solution (0.1 g mL⁻¹) and 0.3 g BOIO. 183 μL RhCl₃•xH₂O solution (0.1 g mL⁻¹) was mixed with 0.3 g BOIO to obtain Rh₁-0.34wt% sample after light irradiation.

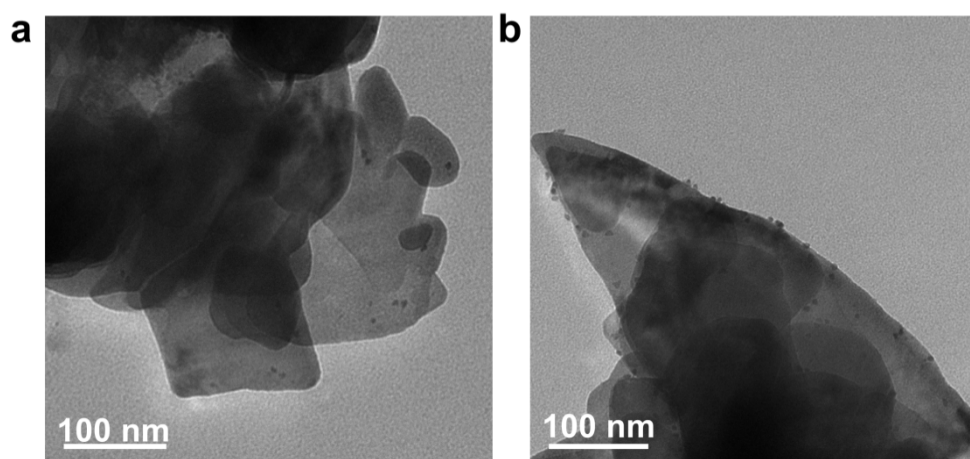


Fig. S14 TEM images of (a) Rh_{NP}-0.35wt% and (b) Rh_{NP}-1.04wt%. The samples were obtained via controlling the amount of as-prepared Rh_{NP}.

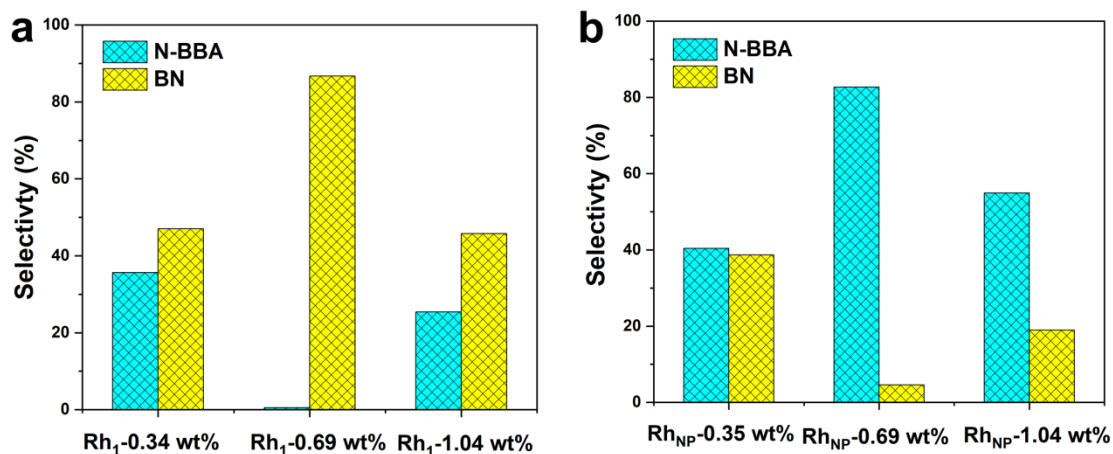


Fig. S15 (a) The selectivity of N-BBA and BN over different Rh₁ amount on BOIO surface, Rh₁-0.69 wt% is the Rh₁/BOIO sample (b) The selectivity of N-BBA and BN over different Rh_{NP} amount on BOIO surface, Rh_{NP}-0.69 wt% is the Rh_{NP}/BOIO sample. The optimized loading amounts of Rh species for Rh₁/BOIO and Rh_{NP}/BOIO have been confirmed.

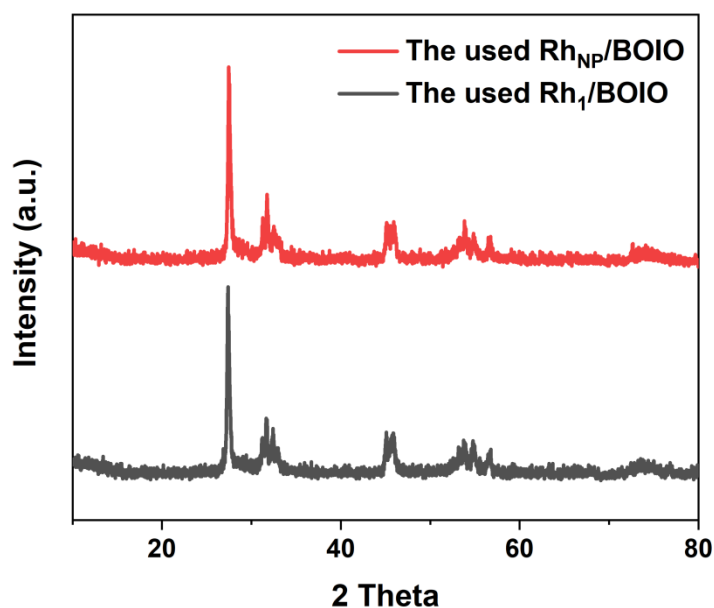


Fig S16. The XRD patterns of the used Rh₁/BOIO and Rh_{NP}/BOIO.

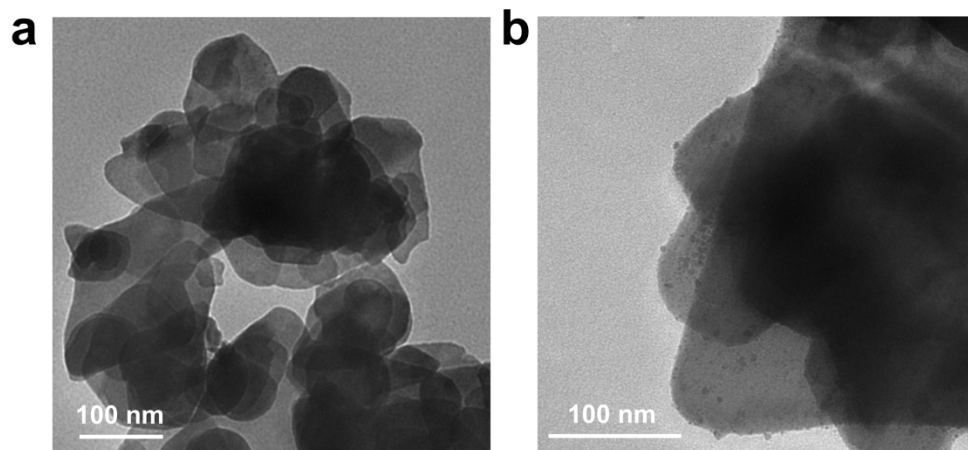


Fig S17. The TEM images of the used (a) Rh₁/BOIO and (b) Rh_{NP}/BOIO.

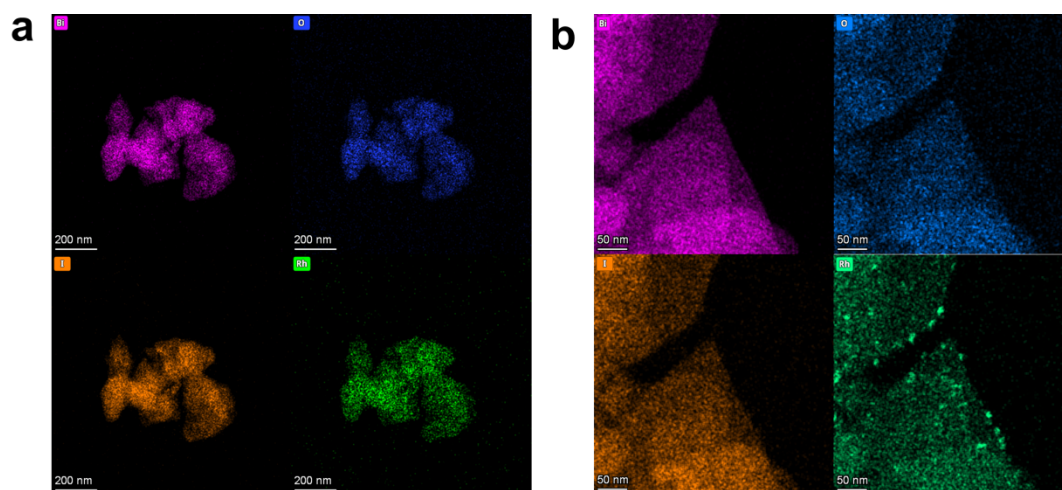


Fig S18. The elements mapping images of the used (a) Rh₁/BOIO and (b) Rh_{NP}/BOIO

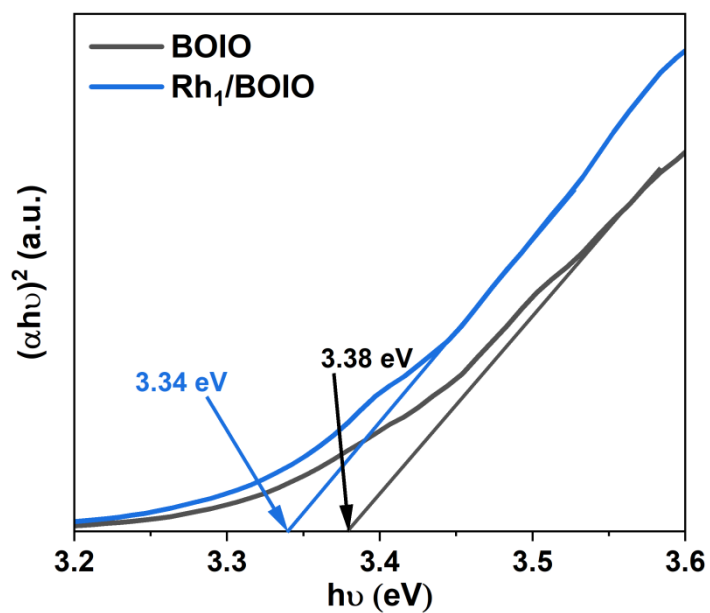


Fig S19. The obtained bandgap of samples via Kubelka–Munk function.

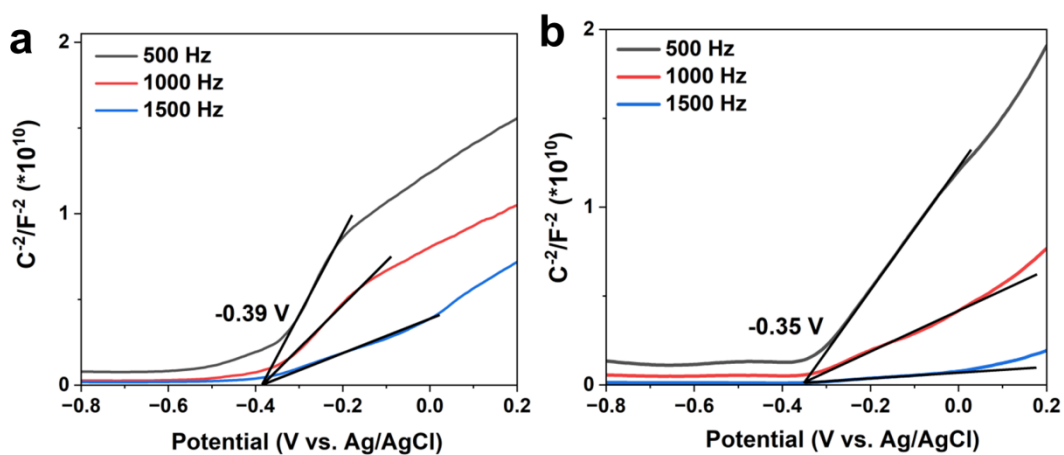


Fig S20. Mott-Schottky plots of (a) BOIO and (b) Rh_1/BOIO under different frequency.

The flat band potential obtained by the M-S plots is approximately 0.1 V below their conductor band positions for n-type semiconductors.

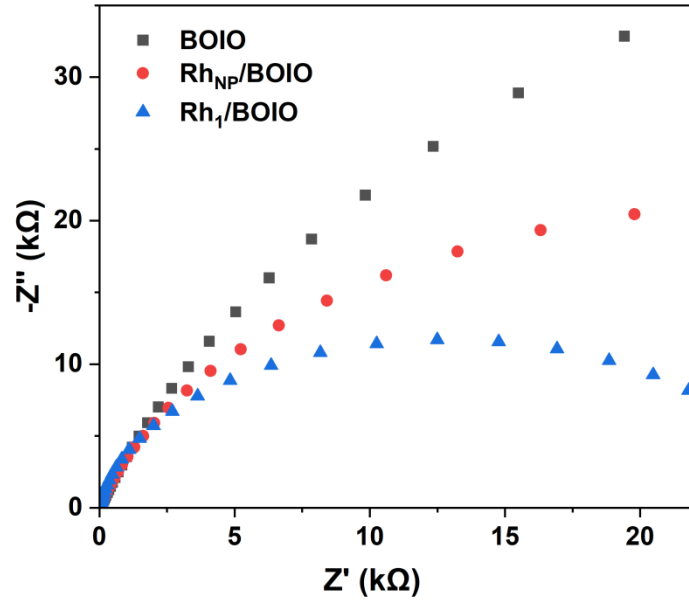


Fig S21. The EIS plots of samples.

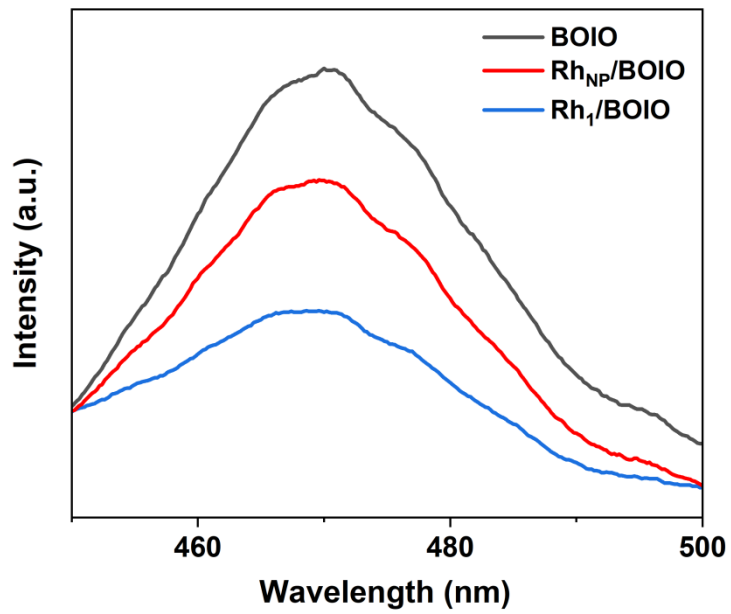


Fig S22. The steady-state PL spectra results of samples.

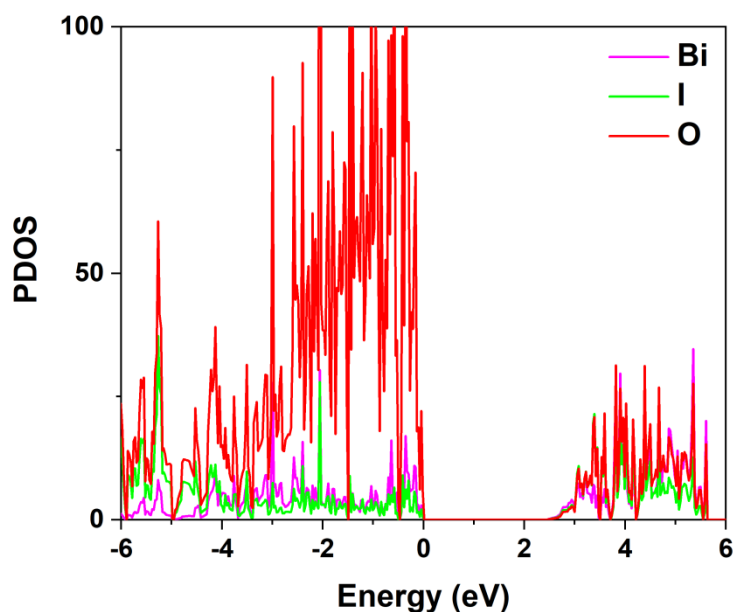


Fig S23. The PDOS results of BOIO.

Table S1. Structural parameters extracted from the Rh K-edge EXAFS fitting.
($S_0^2=0.74$)

Sample	Scattering pair	CN	R(Å)	$\sigma^2(10^{-3}\text{Å}^2)$	$\Delta E_0(\text{eV})$	R factor
Rh ₁ /BOIO	Rh-O	2.32	2.03	4.13	-0.51	0.009
Rh foil	Rh-Rh	12*	2.68	3.82	4.64	0.014

S_0^2 , CN and R is the amplitude reduction factor, coordination number and interatomic distance (the bond length between central atoms and surrounding coordination atoms), respectively; σ^2 is Debye-Waller factor (a measure of thermal and static disorder in absorber-scatterer distances); ΔE_0 is edge-energy shift (the difference between the zero kinetic energy value of the sample and that of the theoretical model). R factor is used to value the goodness of the fitting.

* This value was fixed during EXAFS fitting, based on the known structure.

Error bounds that characterize the structural parameters obtained by EXAFS spectroscopy were estimated as $N \pm 10\%$; $R \pm 1\%$; $\sigma^2 \pm 15\%$; $\Delta E_0 \pm 20\%$.

Rh₁/BOIO (FT range: 2.0-11.1 Å⁻¹; fitting range: 1.05-2.0 Å)

Rh foil (FT range: 3.0-14.0 Å⁻¹; fitting range: 1.0-2.8 Å)

Table S2. BA photo-oxidation performances over Rh₁/BOIO.

Photocatalyst	Product	Conversion (%)	Selectivity (%)	Reference
Rh ₁ /BOIO	BN	99.9	86.7	This work
RuO ₂ /Al ₂ O ₃		75	49	1
Ru-K-ZrO ₂		66	81	2
3%Ru/AC		99.9	81.7	3
[Ru(phen)(PPh ₃) ₂ (CO)(H)]ClO ₄		93	82	4
RuCl ₃		100	53	5
Ru@E-MoS ₂		19	99	6
[RuCl(bhq)(tpy)] [PF ₆]		74	—	7

Table S3. BA photo-oxidation performances over Rh_{NP}/BOIO.

Photocatalyst	Product	Conversion (%)	Selectivity (%)	Reference
Rh _{NP} /BOIO	N-BBA	99.9	82.8	This work
meso-MoOx-300		77	13	8
NH ₂ -MIL-125(Ti)/amorphous TiO ₂		75	73	9
TiO ₂		81	63	10
g-C ₃ N ₄ /BiOBr		63.1	87	11
WS ₂		85	85	12
TiO ₂ nano-flower-4h		99	73.3	13
UiO-66NH ₂ -@Au _{0.5} @COF1		66.9	96	14
WO ₃ /BiOBr		63.8	87	15
BiOI _{0.2} Cl _{0.8}		77.5	96	16
Cu catalysts		82.2	99	17

Table S4. The average lifetimes of different samples.

Samples	τ_1 (ns)	A_1	τ_2 (ns)	A_2	τ_{avg} (ns)
BOIO	0.8538	228.6940	6.8506	1.5270	1.1587
Rh _{NP} /BOIO	0.8642	218.7880	4.7132	4.0910	1.2204
Rh ₁ /BOIO	0.8817	205.5880	4.0689	7.6420	1.3484

References

1. Nordvang, E. C.; Schill, L.; Riisager, A.; Fehrmann, R. Ruthenium Dioxide Catalysts for the Selective Oxidation of Benzylamine to Benzonitrile: Investigating the Effect of Ruthenium Loading on Physical and Catalytic Properties. *Top. Catal.* **2017**, *60*, 1449–1461.
2. Zhu, G.; Shi, S.; Feng, X.; Zhao, L.; Wang, Y.; Cao, J.; Gao, J.; Xu, J. Switching Amine Oxidation from Imines to Nitriles by Carbon-Hydrogen Bond Activation via Strong Base Modified Strategy. *ACS Appl. Mater. Interfaces* **2022**, *14*, 52758–52765.
3. Niu, B.; Lu, F.; Zhang, H.-Y.; Zhang, Y.; Zhao, J. Synthesis of Nitriles from Aerobic Oxidation of Amines Catalyzed by Ruthenium Supported on Activated Carbon. *Chem. Lett.* **2017**, *46*, 330–333.
4. Ray, R.; Chandra, S.; Yadav, V.; Mondal, P.; Maiti, D.; Lahiri, G. K. Ligand controlled switchable selectivity in ruthenium catalyzed aerobic oxidation of primary amines. *Chem. Commun.* **2017**, *53*, 4006–4009.
5. Tang, R.; Diamond, S. E.; Neary, N.; Mares, F., Homogeneous catalytic oxidation of amines and secondary alcohols by molecular oxygen. *J. Chem. Soc., Chem. Commun.* **1978**, 562.
6. Asif Hussain, M.; Yang, M.; Jang, H.-S.; Hwang, S.-Y.; Um, B.-H.; Choi, B. G.; Kim, J. W. Two-Dimensional Heterogeneous Ruthenium–Molybdenum Disulfide Nanocatalyst for the Selective Aerobic Oxidation of Amines. *Ind. Eng. Chem. Res.* **2016**, *55*, 7043–7047.
7. Aiki, S.; Taketoshi, A.; Kuwabara, J.; Koizumi, T.-A.; Kanbara, T., The catalytic activity of a cyclometalated ruthenium(III) complex for aerobic oxidative dehydrogenation of benzylamines. *J. Organomet. Chem.* **2011**, *696*, 1301–1304.

8. Shubhashish, S.; Khanna, H. S.; Achola, L. A.; Amin, A. S.; Willis, W. S.; Suib, S. L. Selective Oxidative Coupling of Amines Using Mesoporous MoO_x Catalysts. *ACS Appl. Nano Mater.* **2021**, *4*, 2086–2097.
9. Sheng, W.; Huang, F.; Lang, X. NH₂-MIL-125(Ti)/amorphous TiO₂ microspheres for enhanced visible light photocatalytic selective oxidation of amines. *Mater. Today Chem.* **2023**, *30*, 101505.
10. Li, N.; Lang, X.; Ma, W.; Ji, H.; Chen, C.; Zhao, J., Selective aerobic oxidation of amines to imines by TiO₂ photocatalysis in water. *Chem. Commun.* **2013**, *49*, 5034.
11. Juntrapirom, S.; Anuchai, S.; Thongsook, O.; Pornsuwan, S.; Meepowpan, P.; Thavornytikarn, P.; Phanichphant, S.; Tantraviwat, D.; Inceesungvorn, B., Photocatalytic activity enhancement of g-C₃N₄/BiOBr in selective transformation of primary amines to imines and its reaction mechanism. *Chem. Eng. J.* **2020**, *394*, 124934.
12. Raza, F.; Park, J. H.; Lee, H.-R.; Kim, H.-I.; Jeon, S.-J.; Kim, J.-H. Visible-Light-Driven Oxidative Coupling Reactions of Amines by Photoactive WS₂ Nanosheets. *ACS Catal.* **2016**, *6*, 2754–2759.
13. Bu, J.; Fang, J.; Leow, W. R.; Zheng, K.; Chen, X. Single-crystalline rutile TiO₂ nano-flower hierarchical structures for enhanced photocatalytic selective oxidation from amine to imine. *RSC Adv.* **2015**, *5*, 103895–103900.
14. Zhang, K.; Xi, Z.; Wu, Z.; Lu, G.; Huang, X., Visible-Light-Induced Selective Oxidation of Amines into Imines over UiO-66-NH₂@Au@COF Core-Shell Photocatalysts. *ACS Sustainable Chemistry & Engineering* **2021**, *9*, 12623–12633.
15. Khampuanbut, A.; Santalelat, S.; Pankiew, A.; Channei, D.; Pornsuwan, S.; Faungnawakij, K.; Phanichphant, S.; Inceesungvorn, B. Visible-light-driven WO₃/BiOBr heterojunction photocatalysts for oxidative coupling of amines to imines: Energy band alignment and mechanistic insight. *J. Colloid Interface Sci.* **2020**, *560*, 213–224.
16. Anuchai, S.; Tantraviwat, D.; Nattestad, A.; Chen, J.; Inceesungvorn, B., Tuning product selectivity and visible-light-driven activity in oxidative coupling of amines to imines: A case study of BiOI_xCl_{1-x} photocatalyst. *Colloids and Surfaces A: Physicochemical and Engineering Aspects* **2021**, *629*, 127481.
17. Chai, Y.; Zhang, L.; Liu, Q.; Yang, F.; Dai, W.-L. Insights into the Relationship of the Heterojunction Structure and Excellent Activity: Photo-Oxidative Coupling of Benzylamine on CeO₂-rod/g-C₃N₄ Hybrid under Mild Reaction Conditions. *ACS. Sustain. Chem. Eng.* **2018**, *6*, 10526–10535.