

Electronic Supplementary Information for

Highly Lubricative, Self-Healing, Chemically Degradable, Anti-Flaming and Anti-Irradiating Magnetic Gels Simply from Binary Mixtures of Fe₃O₄ Nanoparticles and Water

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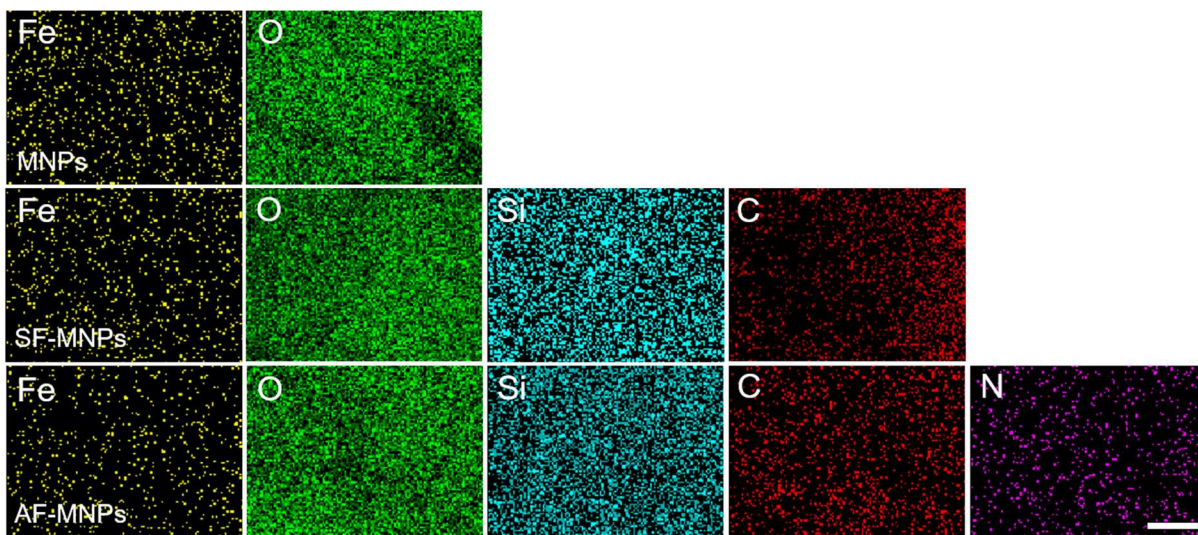


Figure S1. Element distribution on the surface of pristine, SF-, and AF-MNPs. Scale bar = 1 μm .

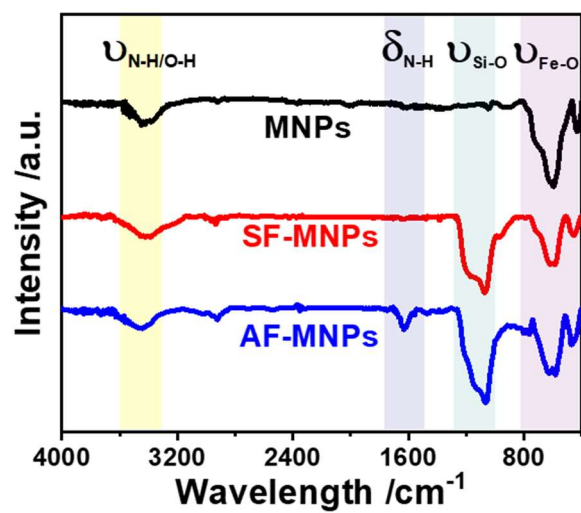


Figure S2. FT-IR spectra of pristine, SF- and AF-MNPs. T = 25 °C.

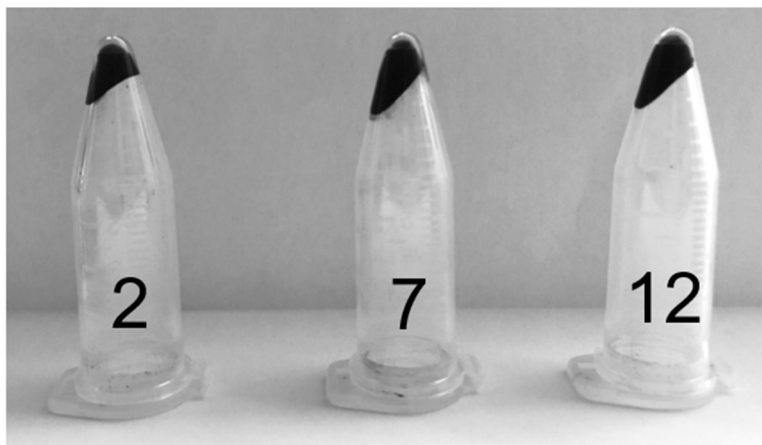


Figure S3. Photographs of the inorganic Fe₃O₄ gels. The numbers refer to the internal pH values of the aqueous gels. T = 25 °C.

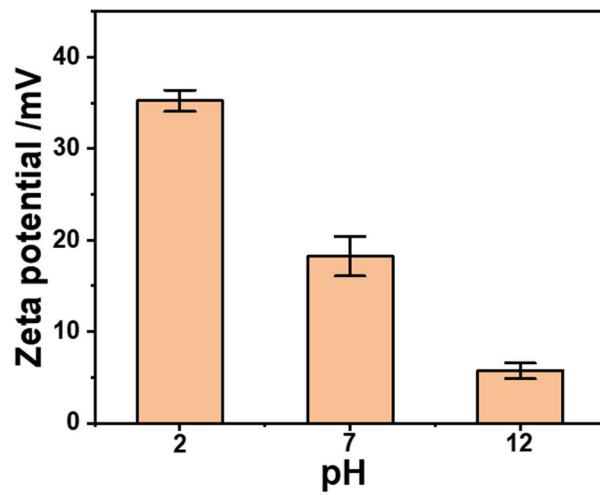


Figure S4. Zeta potential measurements of 1 mg mL^{-1} aqueous suspensions of AF-MNPs at different pH values. $T = 25 \text{ }^\circ\text{C}$. $c_{\text{Fe}_3\text{O}_4} = 48 \text{ wt\%}$.

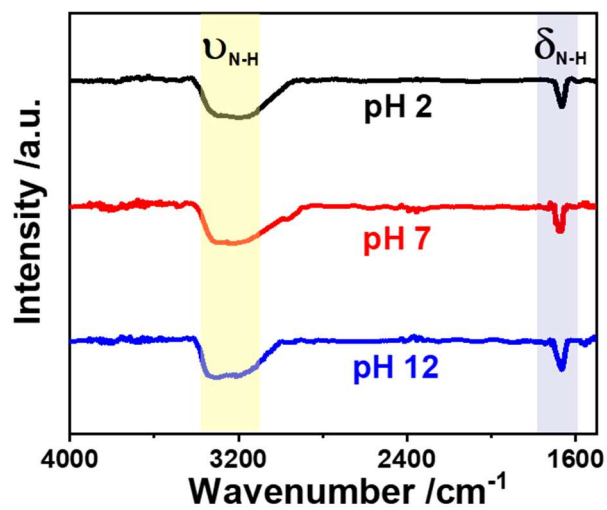


Figure S5. FT-IR spectra of lyophilized Fe₃O₄ gels at pH 2, 7 and 12, respectively. T = 25 °C.

$c_{\text{Fe}_3\text{O}_4} = 48 \text{ wt\%}$.

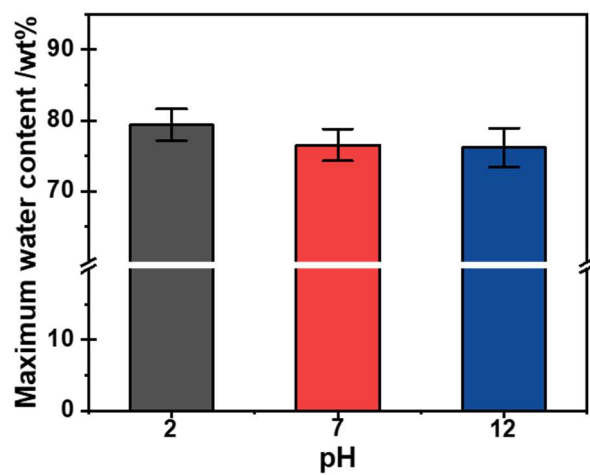


Figure S6. Maximum water content (in wt%) of the magnetic gels with different internal pH values. T = 25 °C.

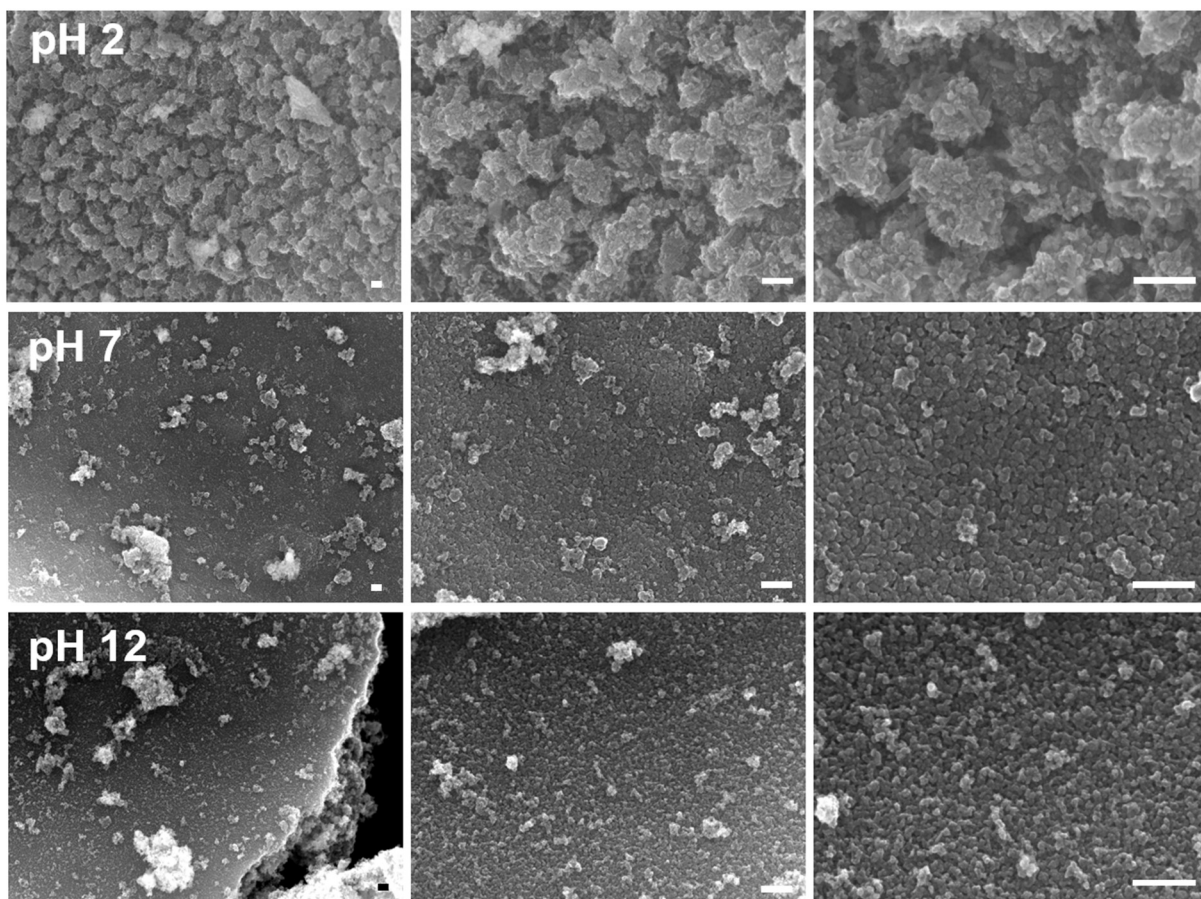


Figure S7. SEM images of freeze-dried Fe₃O₄ gels obtained at different scanning locations and magnifications. Scale bar = 100 nm. $c_{\text{Fe}_3\text{O}_4} = 48 \text{ wt\%}$.

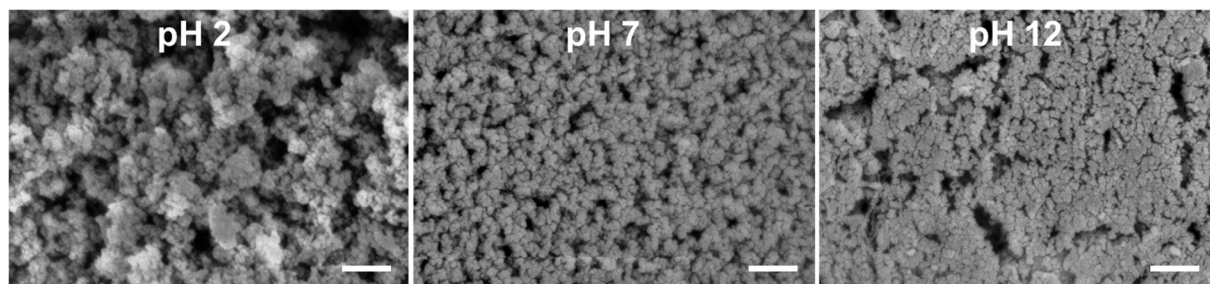


Figure S8. Cryogenic scanning electron microscopy (cryo-SEM) observations of the Fe_3O_4 gels at pH 2, 7 and 12, respectively. Scale bar = 100 nm. $c_{\text{Fe}_3\text{O}_4} = 48$ wt%.

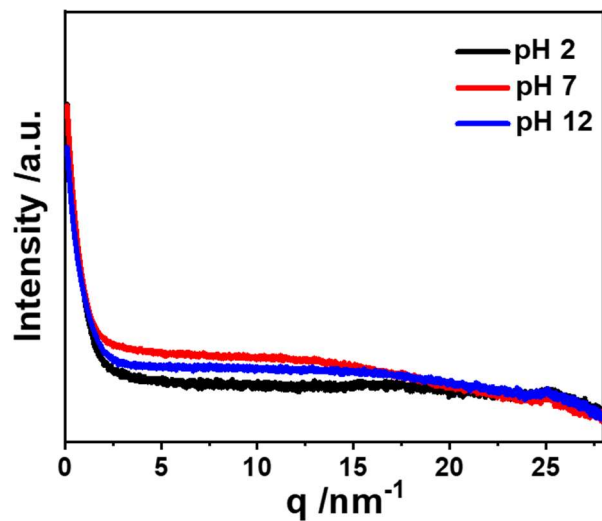


Figure S9. SAXS patterns of lyophilized Fe₃O₄ gels at different pH values. T = 25 °C. $c_{\text{Fe}_3\text{O}_4} = 48$ wt%.

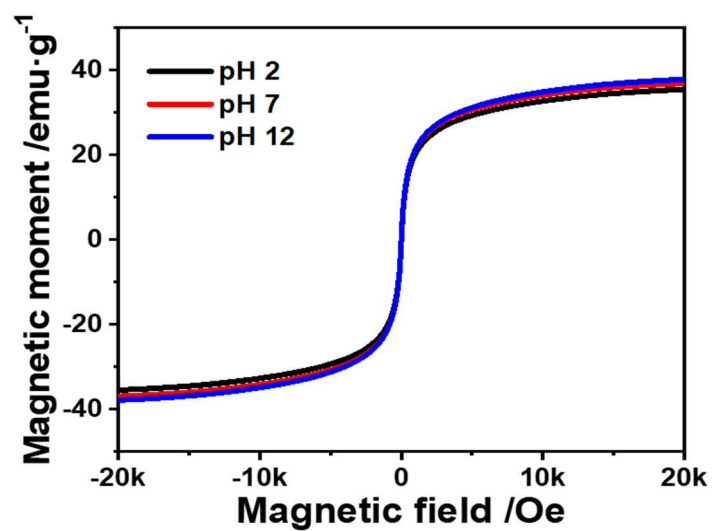


Figure S10. SQUID magnetometry profiles of 48 wt% Fe₃O₄ gels with different internal pH values at 25 °C.

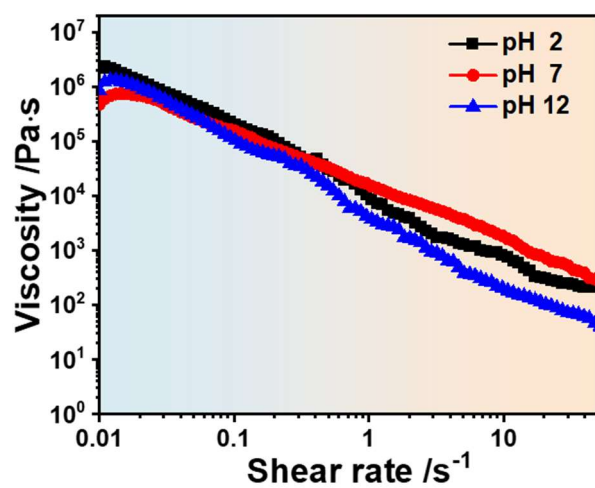


Figure S11. Colloidal viscosity of the Fe₃O₄ gels at shear rates between 0.01 and 50 s⁻¹. T = 25 °C. $c_{\text{Fe}_3\text{O}_4} = 48$ wt%.

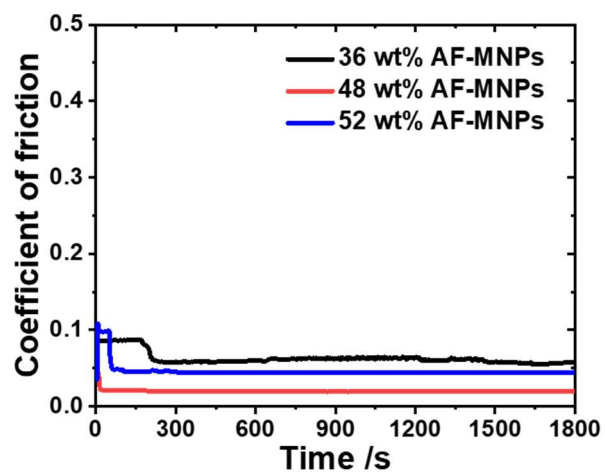


Figure S12. Effect of particle concentration on the CoF of pH 12 Fe_3O_4 gels at 800 N and 20 Hz. $T = 25\text{ }^\circ\text{C}$.

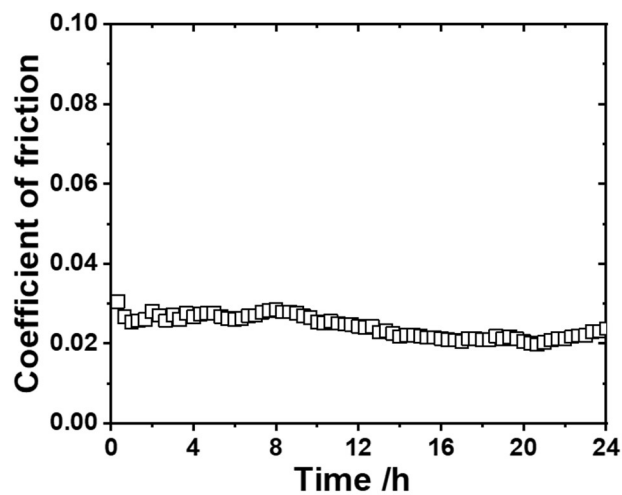


Figure S13. CoF of the pH 12 Fe_3O_4 gel upon continuously used for 24 h under 800 N and 20 Hz. $T = 25\text{ }^\circ\text{C}$. $c_{\text{Fe}_3\text{O}_4} = 48\text{ wt}\%$.

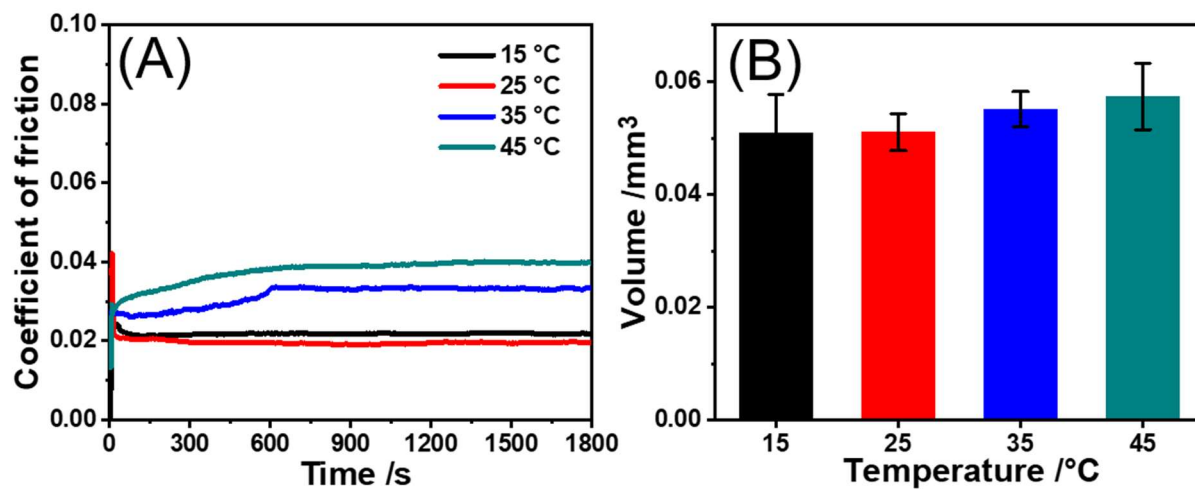


Figure S14. (A) CoF and (B) wear volume of the pH 12 Fe_3O_4 gel at temperatures between 15 and 45 °C under 800 N and 20 Hz. $c_{\text{Fe}_3\text{O}_4} = 48$ wt%.

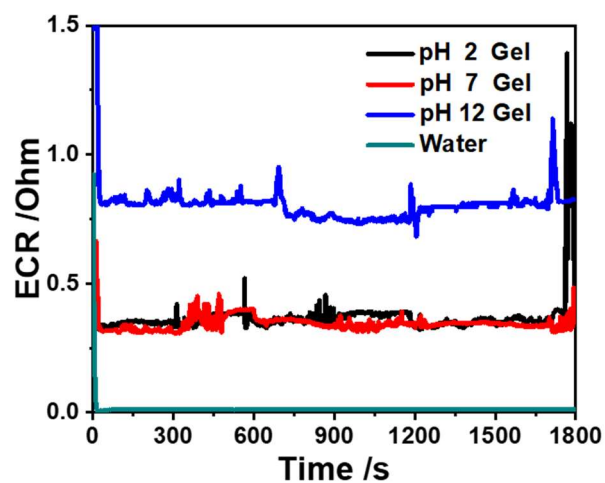


Figure S15. ECR of water and Fe_3O_4 gels with different internal pH values at 25 °C. $c_{\text{Fe}_3\text{O}_4} = 48 \text{ wt\%}$.

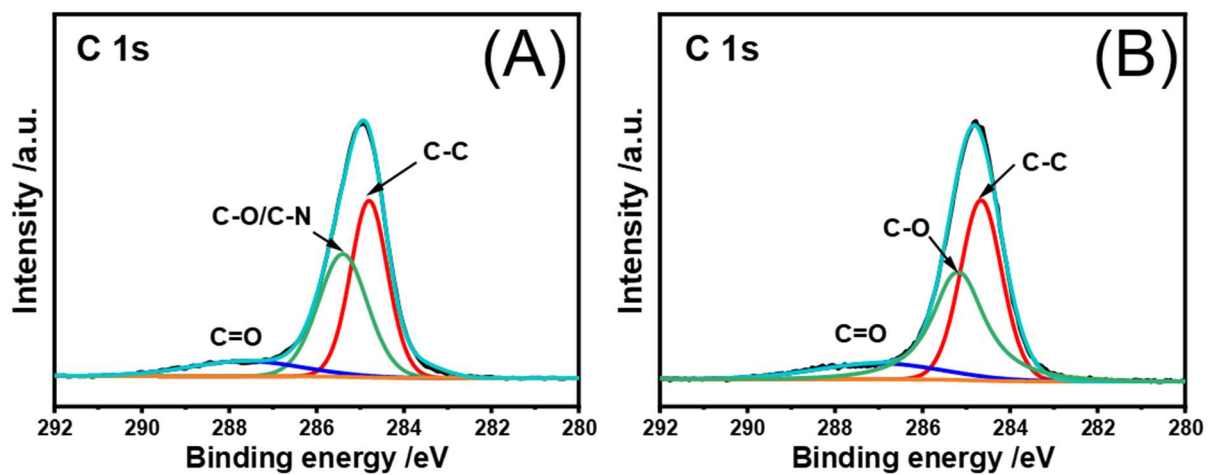


Figure S16. XPS C 1s spectra of steel substrates lubricated with (A) pH 12 Fe_3O_4 gel and (B) ultrapure water, respectively. $T = 25\text{ }^\circ\text{C}$.

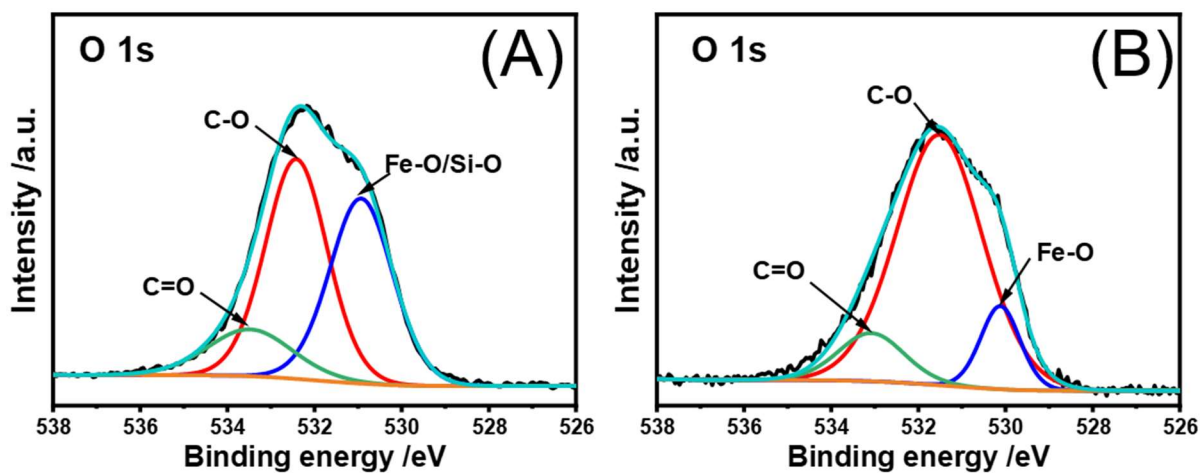


Figure S17. XPS O 1s spectra of steel substrates lubricated with (A) pH 12 Fe_3O_4 gel and (B) deionized water, respectively. $T = 25^\circ\text{C}$.

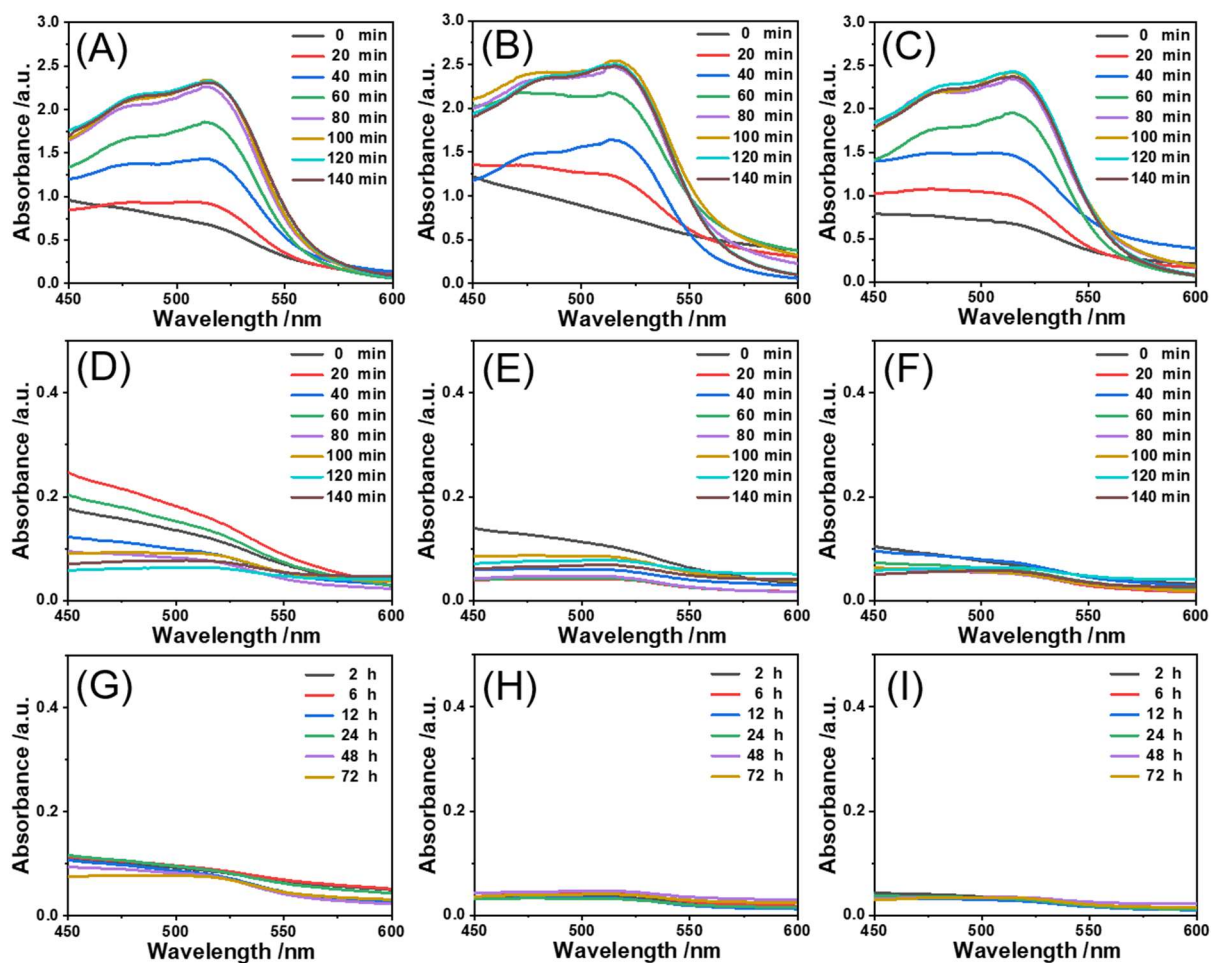


Figure S18. UV-vis spectra of the pH 2 (A), 7 (B) and 12 (C) Fe_3O_4 gels after immersed in hydrochloric acid for 140 min. UV-vis spectra of the pH 2 (D, G), 7 (E, H) and 12 (F, I) Fe_3O_4 gels after placed in deionized water for different durations. $T = 25\text{ }^\circ\text{C}$.

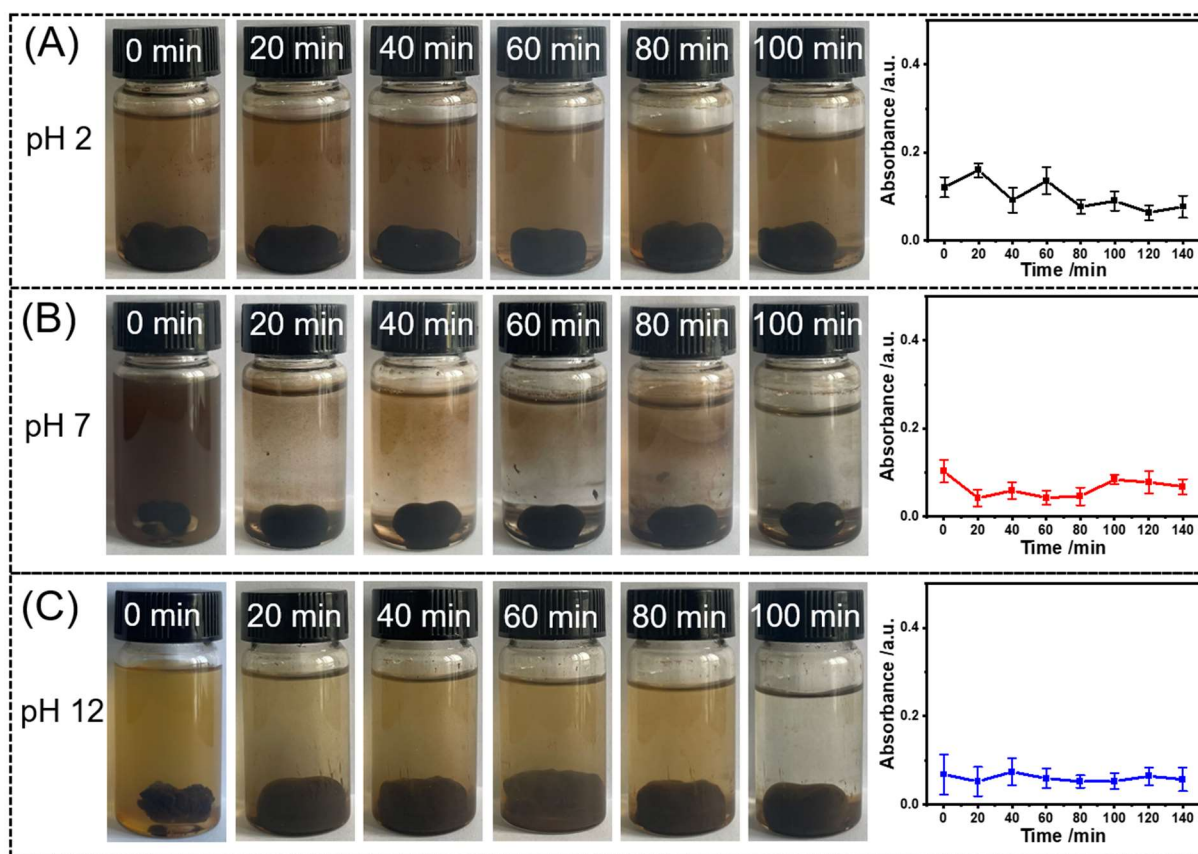


Figure S19. Variation of the pH 2 (A), 7 (B) and 12 (C) magnetic gels after placed in ultrapure water for 100 min under magnetic stirring. The UV-vis spectra on the right were used for monitoring the decomposition process by detecting the absorbance of released Fe-based ions in solutions.

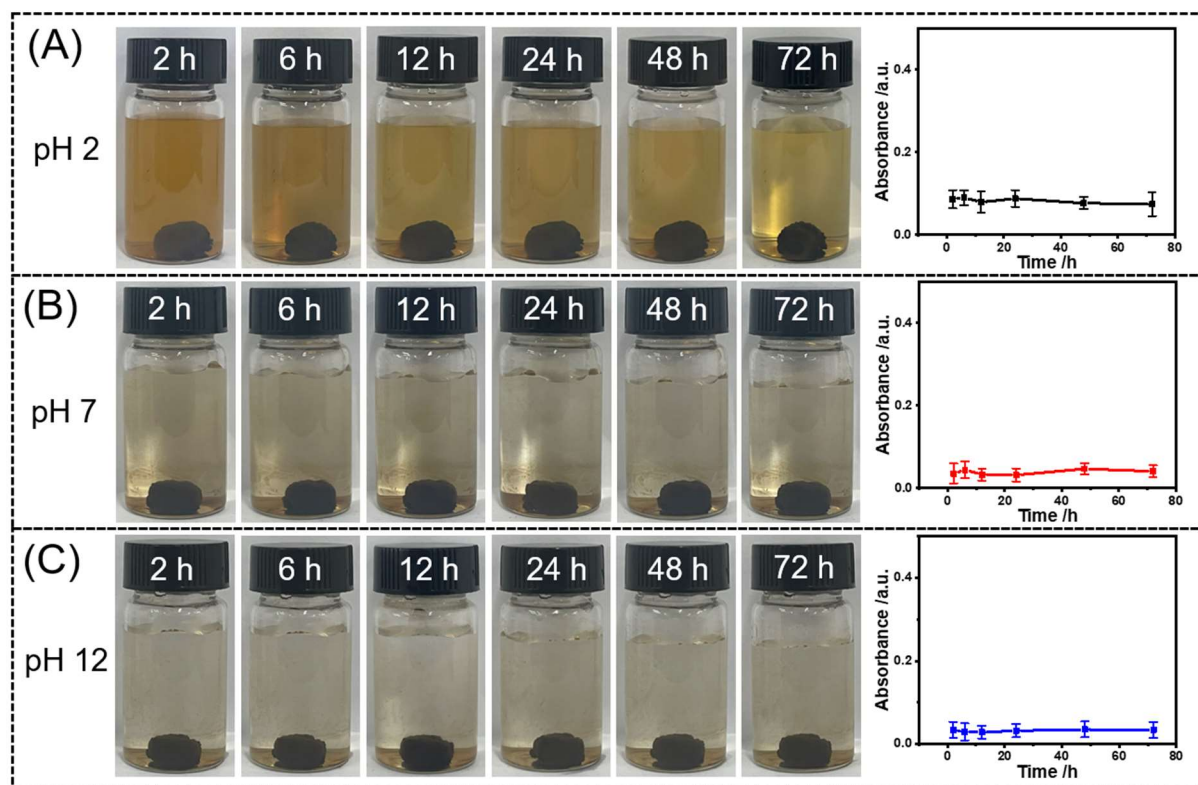


Figure S20. Variation of the pH 2 (A), 7 (B) and 12 (C) magnetic gels after placed in ultrapure water for continuous 72 h without magnetic stirring. The UV-vis spectra on the right were used for monitoring the decomposition process by detecting the absorbance of released Fe-based ions in solutions.



Figure S21. Variation of an equal amount of commercially available magnetic greases after placed in 1 mol L⁻¹ HCl solutions for 100 min under magnetic stirring. The UV-vis spectra on the right were utilized to monitor the degradation process by detecting the absorbance of released Fe-based ions in solutions.

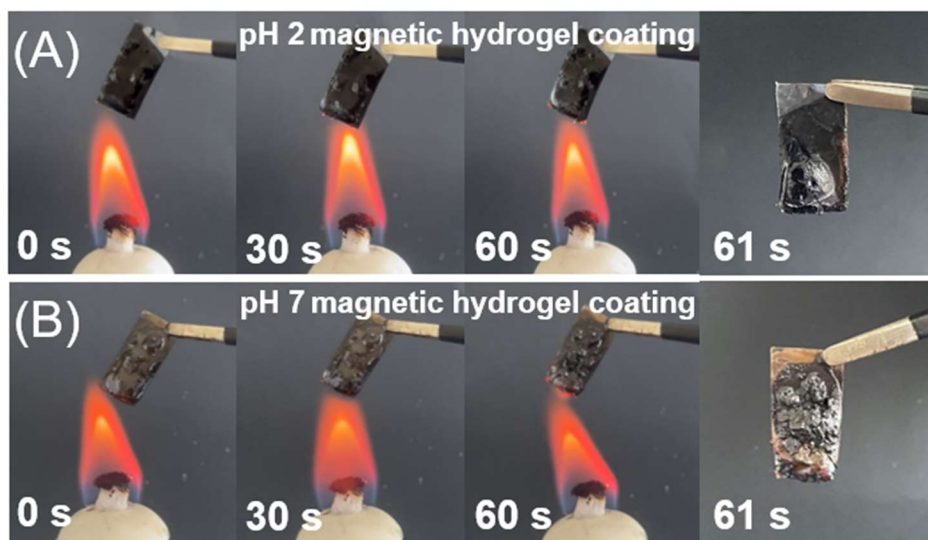


Figure S22. Photographs of plastic substrates coated with pH 2 (A) and 7 (B) Fe_3O_4 gels upon exposure to an external alcohol lamp.

Movie S1. Magnetically controlled migration of a pH 12, 48 wt% Fe₃O₄ gel by an external 0.25 T magnet at 25 °C.