

## Supplementary Information

### Computational prediction of the critical micelle concentration (CMC) of surfactants using the non-Bornian solvation model

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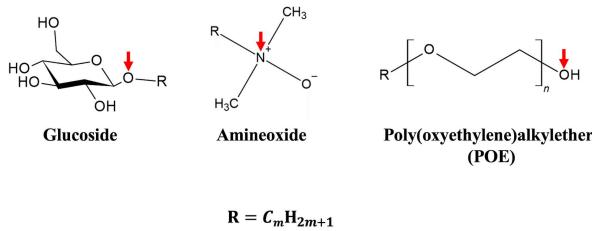


Fig S1 The respective red arrow shows an arbitrarily chosen atom for three different types of non-ionic surfactants, *i.e.*, glucoside, amineoxide, and poly(oxyethylene)alkylether (POE).

Table S1 Theoretical parameters representing the adsorption states of anionic and cationic surfactants at the NB/W interface (25 °C)

Surfactant	d (Å)	θ (°)	ω (°)
<b>Anionic</b>			
C <sub>8</sub> H <sub>17</sub> SO <sub>4</sub> <sup>-</sup> Na <sup>+</sup>	-1.8	164	0
C <sub>9</sub> H <sub>19</sub> SO <sub>4</sub> <sup>-</sup> Na <sup>+</sup>	-1.8	163	118
C <sub>10</sub> H <sub>21</sub> SO <sub>4</sub> <sup>-</sup> Na <sup>+</sup>	-1.8	162	0
C <sub>12</sub> H <sub>25</sub> SO <sub>4</sub> <sup>-</sup> Na <sup>+</sup>	-1.8	160	0
C <sub>6</sub> H <sub>13</sub> SO <sub>3</sub> <sup>-</sup> Na <sup>+</sup>	-1.4	152	0
C <sub>8</sub> H <sub>17</sub> SO <sub>3</sub> <sup>-</sup> Na <sup>+</sup>	-1.4	154	1
C <sub>11</sub> H <sub>23</sub> SO <sub>3</sub> <sup>-</sup> H <sup>+</sup>	-1.4	157	0
C <sub>7</sub> H <sub>15</sub> COO <sup>-</sup> K <sup>+</sup>	-1.0	154	0
C <sub>8</sub> H <sub>17</sub> COO <sup>-</sup> K <sup>+</sup>	-1.0	154	180
C <sub>9</sub> H <sub>19</sub> COO <sup>-</sup> K <sup>+</sup>	-1.0	155	0
C <sub>10</sub> H <sub>21</sub> COO <sup>-</sup> K <sup>+</sup>	-1.0	156	1
C <sub>11</sub> H <sub>23</sub> COO <sup>-</sup> K <sup>+</sup>	-0.8	150	0
C <sub>12</sub> H <sub>25</sub> COO <sup>-</sup> K <sup>+</sup>	-0.8	153	20
C <sub>13</sub> H <sub>27</sub> COO <sup>-</sup> K <sup>+</sup>	-0.8	153	19
<b>Cationic</b>			
C <sub>8</sub> H <sub>17</sub> NH <sub>3</sub> <sup>+</sup> Cl <sup>-</sup>	-0.8	158	180
C <sub>8</sub> H <sub>17</sub> N(CH <sub>3</sub> ) <sub>3</sub> <sup>+</sup> Br <sup>-</sup>	2.0	136	180
C <sub>10</sub> H <sub>21</sub> N(CH <sub>3</sub> ) <sub>3</sub> <sup>+</sup> Br <sup>-</sup>	2.2	136	179
C <sub>12</sub> H <sub>25</sub> NH <sub>3</sub> <sup>+</sup> Cl <sup>-</sup>	-0.8	158	1
C <sub>12</sub> H <sub>25</sub> N(CH <sub>3</sub> ) <sub>3</sub> <sup>+</sup> Cl <sup>-</sup> <sup>a</sup>	2.2	137	180
C <sub>12</sub> H <sub>25</sub> N(CH <sub>3</sub> ) <sub>3</sub> <sup>+</sup> Br <sup>-</sup> <sup>a</sup>	2.2	137	180
C <sub>14</sub> H <sub>29</sub> NH <sub>3</sub> <sup>+</sup> Cl <sup>-</sup>	-0.8	158	0
C <sub>14</sub> H <sub>29</sub> N(CH <sub>3</sub> ) <sub>3</sub> <sup>+</sup> Cl <sup>-</sup>	2.2	133	180
C <sub>18</sub> H <sub>37</sub> N(CH <sub>3</sub> ) <sub>3</sub> <sup>+</sup> Cl <sup>-</sup>	2.0	136	180

<sup>a</sup> Possible effects of hydrophilic counterions were not taken into account in the present theoretical simulation.

Table S2 Theoretical parameters representing the adsorption states of non-ionic surfactants at the NB/W interface (25 °C)

Surfactant	<i>d</i> (Å)	<i>θ</i> (°)	<i>ω</i> (°)
<b>Non-ionic</b>			
C <sub>8</sub> H <sub>17</sub> (OC <sub>2</sub> H <sub>4</sub> )OH	2.8	29	180
C <sub>8</sub> H <sub>17</sub> (OC <sub>2</sub> H <sub>4</sub> ) <sub>3</sub> OH	9.0	26	1
C <sub>8</sub> H <sub>17</sub> (OC <sub>2</sub> H <sub>4</sub> ) <sub>6</sub> OH	17.6	36	181
C <sub>8</sub> H <sub>17</sub> (OC <sub>2</sub> H <sub>4</sub> ) <sub>9</sub> OH	27.2	27	0
C <sub>10</sub> H <sub>21</sub> (OC <sub>2</sub> H <sub>4</sub> ) <sub>3</sub> OH	9.0	25	187
C <sub>10</sub> H <sub>21</sub> (OC <sub>2</sub> H <sub>4</sub> ) <sub>6</sub> OH	17.6	36	45
C <sub>10</sub> H <sub>21</sub> (OC <sub>2</sub> H <sub>4</sub> ) <sub>9</sub> OH	14.8	179	152
C <sub>12</sub> H <sub>25</sub> (OC <sub>2</sub> H <sub>4</sub> ) <sub>6</sub> OH	17.4	37	0
C <sub>13</sub> H <sub>27</sub> (OC <sub>2</sub> H <sub>4</sub> ) <sub>5</sub> OH	14.8	29	4
C <sub>14</sub> H <sub>29</sub> (OC <sub>2</sub> H <sub>4</sub> ) <sub>6</sub> OH	18.0	34	5
<i>n</i> -octyl-β-D-glucoside	1.8	102	124
<i>n</i> -decyl-β-D-glucoside	1.8	63	99
<i>n</i> -dodecyl-β-D-glucoside	1.6	69	89
C <sub>9</sub> H <sub>19</sub> N(CH <sub>3</sub> ) <sub>2</sub> -O	0.6	67	106
C <sub>10</sub> H <sub>21</sub> (OC <sub>2</sub> H <sub>4</sub> )N(CH <sub>3</sub> ) <sub>2</sub> -O	0.6	68	138
C <sub>12</sub> H <sub>25</sub> (OC <sub>2</sub> H <sub>4</sub> )N(CH <sub>3</sub> ) <sub>2</sub> -O	0.6	68	302
C <sub>14</sub> H <sub>29</sub> (OC <sub>2</sub> H <sub>4</sub> )N(CH <sub>3</sub> ) <sub>2</sub> -O	0.6	68	299
C <sub>16</sub> H <sub>33</sub> (OC <sub>2</sub> H <sub>4</sub> )N(CH <sub>3</sub> ) <sub>2</sub> -O	0.4	65	304
C <sub>18</sub> H <sub>37</sub> (OC <sub>2</sub> H <sub>4</sub> )N(CH <sub>3</sub> ) <sub>2</sub> -O	0.4	65	304