Supporting information

Deciphering the spectroscopic and thermodynamic aspects of binding of biologically important antioxidants with alkali induced state of human serum albumin

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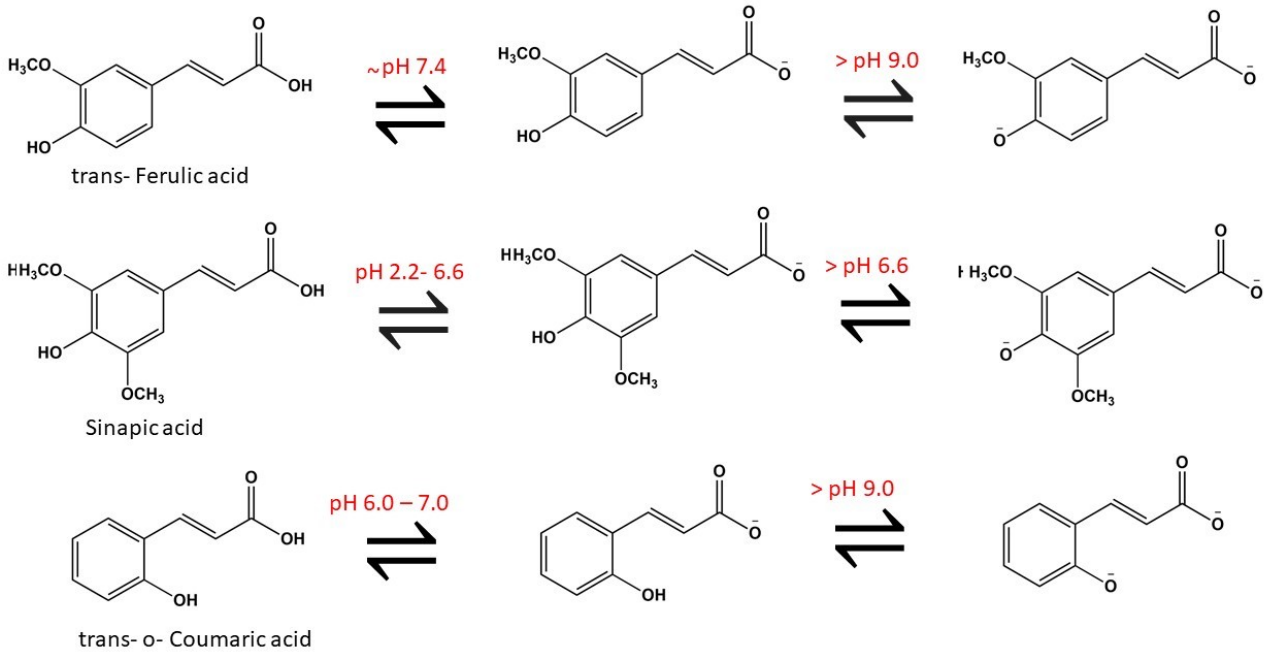
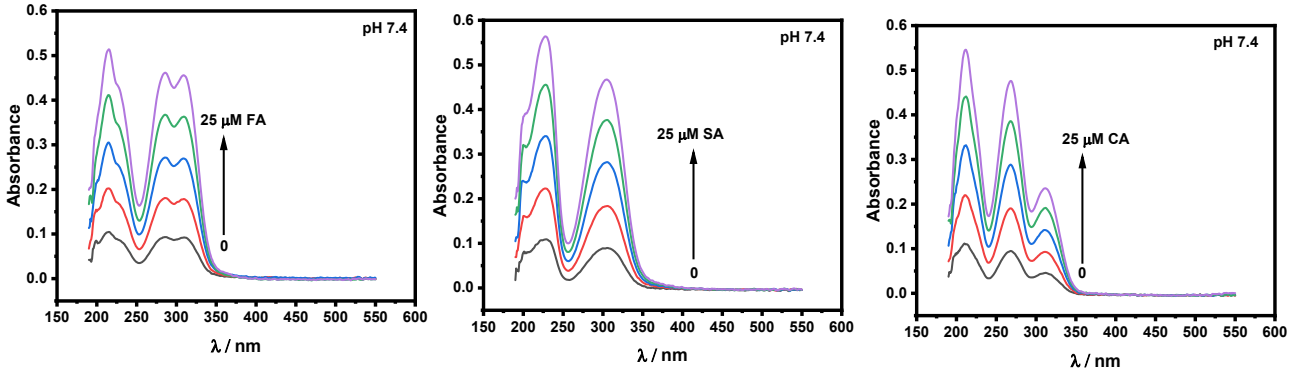


Fig. S1 Ionic structure of hydroxycinnamic acid derivatives with the change of pH.



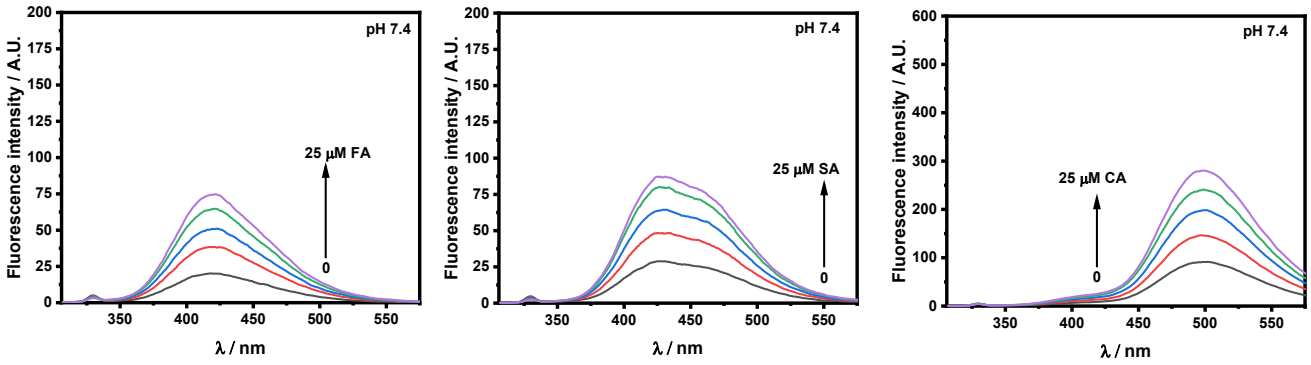


Fig. S2 Absorbance and fluorescence spectra of hydroxycinnamic acid derivatives at pH 7.4 and 11.2 at 298.15 K.

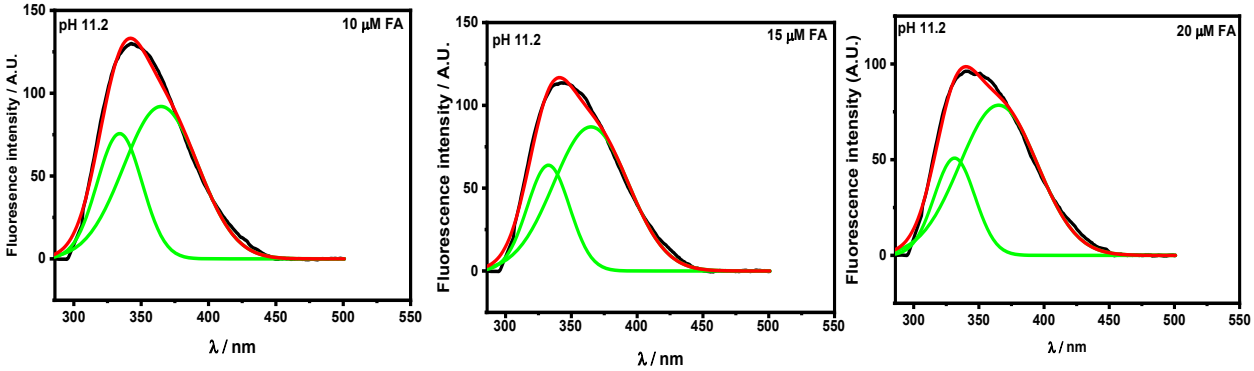


Fig. S3 Deconvoluted fluorescence peaks in the emission spectra of 5×10^{-6} M HSA in the presence of ferulic acid.

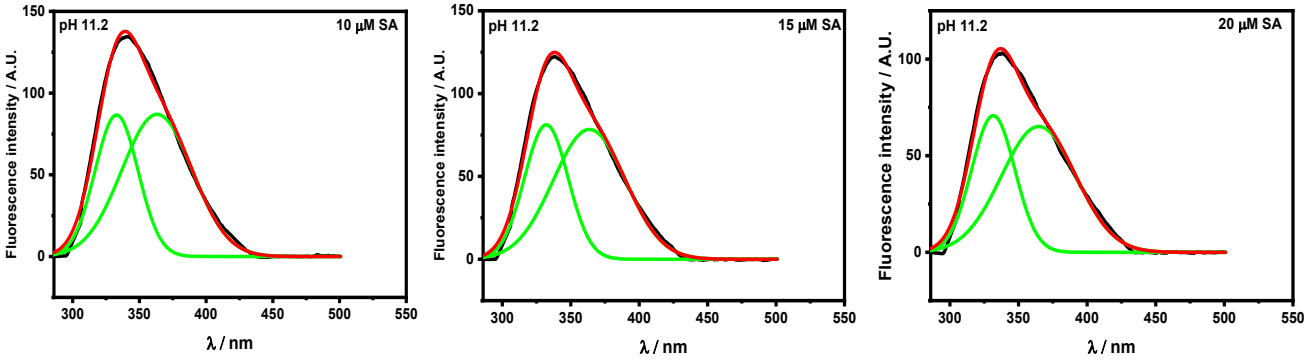
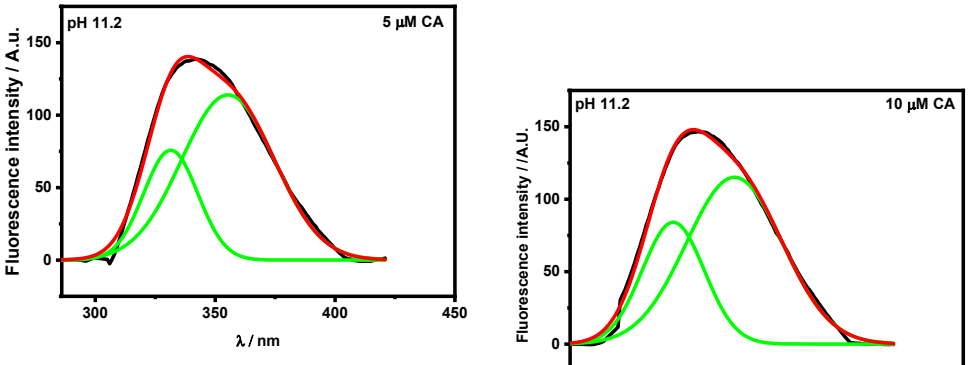


Fig. S4 Deconvoluted fluorescence peaks in the emission spectra of 5×10^{-6} M HSA in the presence of sinapic acid.



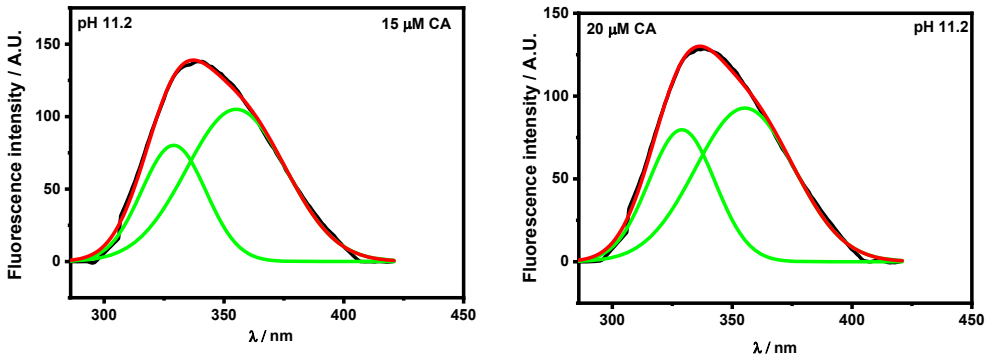


Fig. S5 Deconvoluted fluorescence peaks in the emission spectra of 5 × 10^{-6} M HSA in the presence of trans-o-coumaric acid.

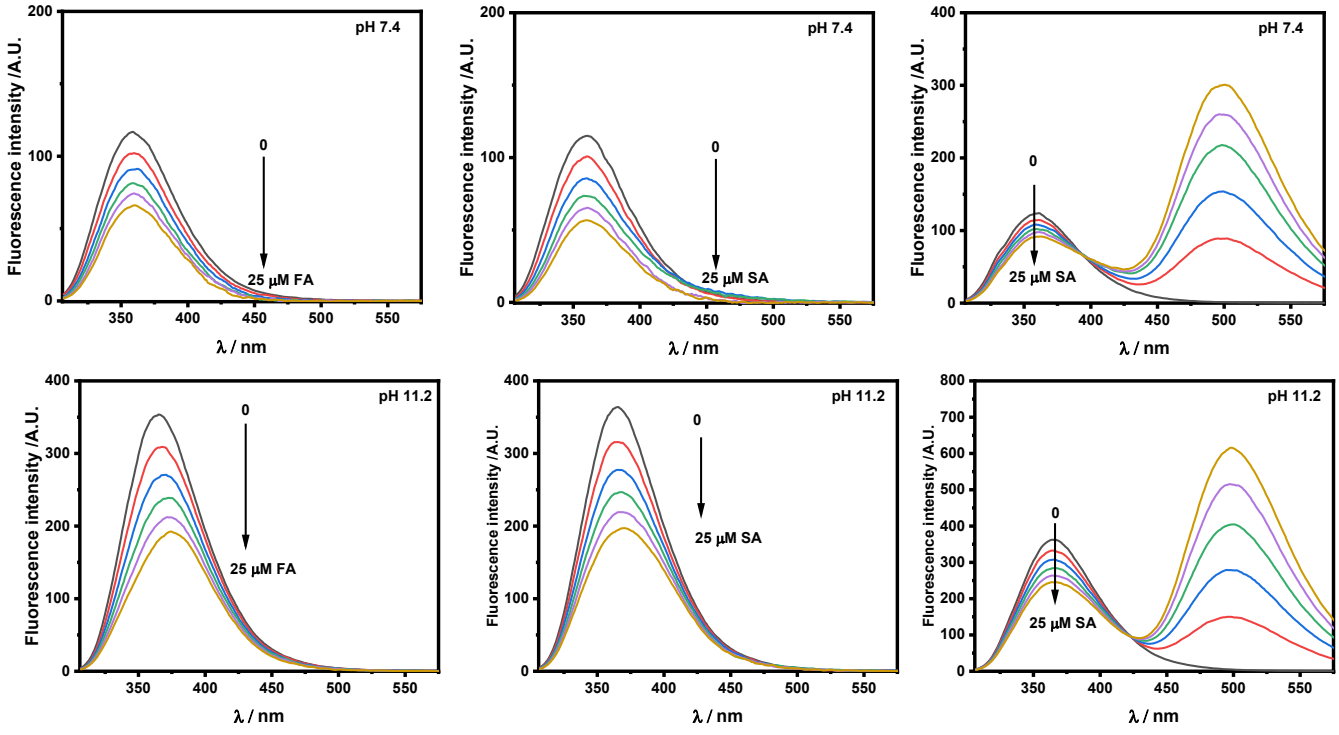


Fig.S6 Emission spectra of 5 × 10^{-6} M tryptophan upon excitation at 295 nm in the presence of ferulic acid (FA), sinapic acid (SA) and trans-o-coumaric acid (CA) at pH 7.4 and 11.2 at 298.15 K.

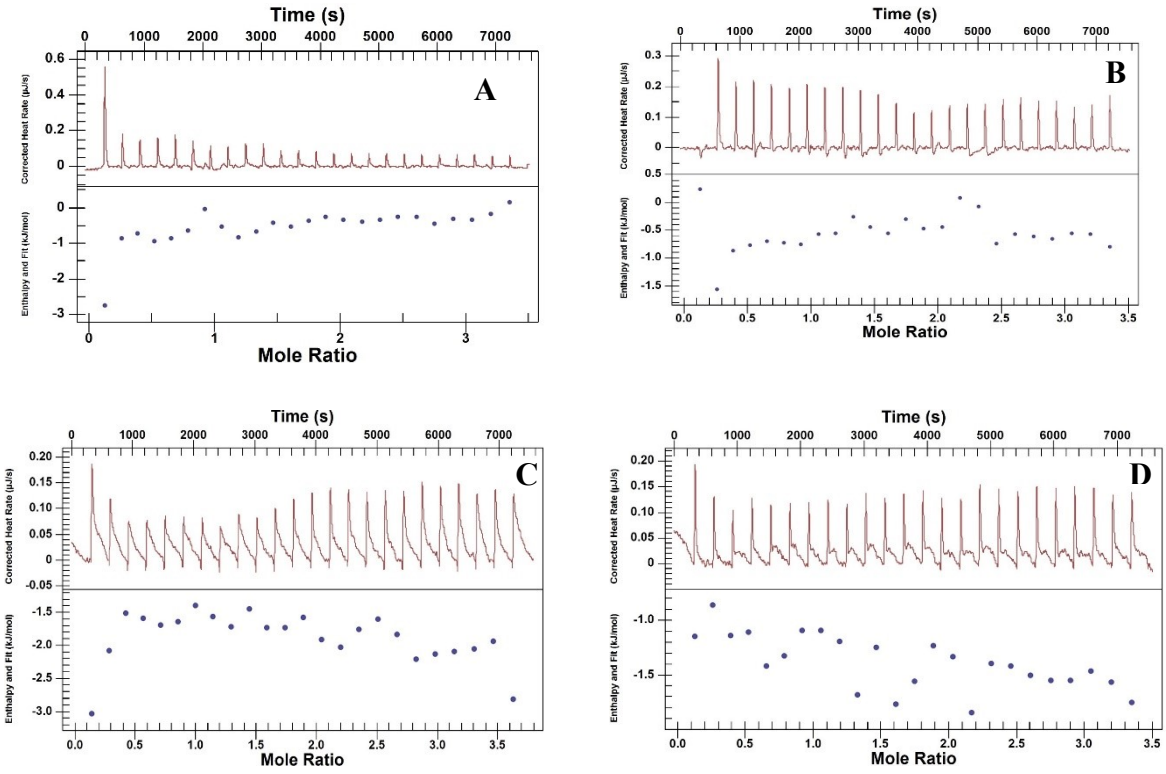


Fig.S7 Dilution ITC profiles of (A) 0.13 mM HSA and 2.5 mM of (B) FA (C) SA and (D) CA at pH 7.4 at 298.15 K.

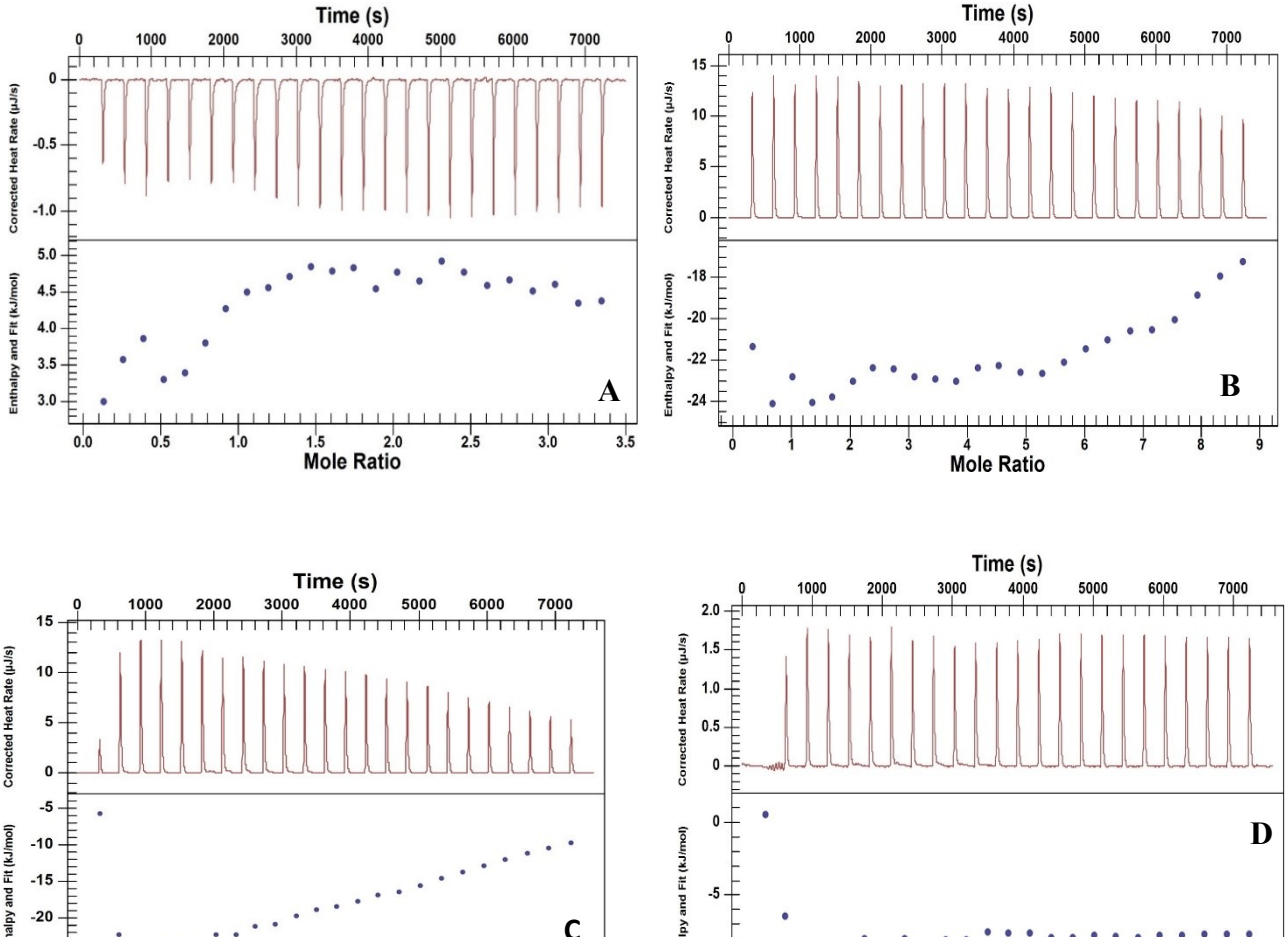


Fig.S8 Dilution ITC profiles of (A) 0.13 mM HSA and (B) 6.5 mM FA (C) 7.8 mM SA And (D) 2.5 mM CA at pH 711.2 at 298.15 K.