Supporting Information

Integrating Solid Interfaces for Catalysis in All-Solid-State Lithium-Sulfur Batteries

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Keywords: All-solid-state Li-S batteries; Molecular-level Interfaces; Catalysis; Micropores.



Fig. S1 (a) The monolith formed after hydrothermal treatment of glucose with GO and the corresponding

(b) SEM image.



Fig. S2 Glucose/GO weight ratio adjustment of MCS carbon.



Fig. S3 Nitrogen adsorption-desorption isotherms of BP, MCS-1, MCS-2 and MCS-3 and the specific surface areas calculated by Brunauer-Emmett-Teller method.



Fig. S4 Pore size analysis of BP, MCS-1, MCS-2 and MCS-3 by non-local density functional theory method.



Fig. S5 α_s plots of (a) BP, (b) MCS-1, (c) MCS-2 and (d) MCS-3.







Fig. S7 Fourier transform infrared spectrometer spectra of MCS-1, MCS-2 and MCS-3.



Fig. S8 Bulk conductivities of BP, MCS-1, MCS-2 and MCS-3.



Fig. S9 TOF-SIMs spectrum of a) CoS⁻; b) CoSH⁻; c) Co₃S₄⁻; d) CoS₂⁻; e) CoS₂H⁻; f) Co2S⁻.



Fig. S10 HAADF-STEM images of MCS-3-S-Co (colored according to the scattering intensity).



Fig. S11 HAADF-STEM images of BP-S-Co particles (colored according to the scattering intensity).



Fig. S12 (a) TGA curves of MCS-1-S, MCS-2-S and MCS-3-S under N_2 atmosphere. (b-d) TGA curves of MCS-1-S-Co, MCS-2-Co and MCS-3-Co under N_2 and air atmosphere.









Fig. S14 DC polarization measurement of MCS-3-S.



Fig. S15 Linear fittings of DC polarization results for different cathodes: (a) MCS-1-S, (b) MCS-2-S, (c) MCS-

3-S, (d) MCS-1-S-Co, (e) MCS-2-S-Co, (f) MCS-3-S-Co, (g) BP-S and (h) BP-S-Co.



Fig. S16 CV curves at different temperatures for (a) BP-S-Co, (b) MCS-1-S-Co, (c) MCS-2-S-Co, (d) MCS-3-S-Co and (d) MCS-3-S. (f) Corresponding Arrhenius plots calculated by peak currents.



Fig. S17 Comparisons MCS-3-S and MCS-3-S-Co electrodes.



Fig. S18 Specific capacity and overpotential analysis of BP-S, MCS-1-S, MCS-2-S and MCS-3-S cathodes.



Fig. S19 Binding energy and peak area ratio analysis of Co³⁺ during discharge.







Fig. S21 EIS measurement results of BP-S-Co, MCS-1-S-Co, and MCS-3-S-Co cathodes.



Fig. S22 (a) GITT measurement results and (b) corresponding overpotentials for BP-S-Co, MCS-1-S-Co, and MCS-3-S-Co cathodes.



Fig. S23 (a) P 2p, (b) S 2p and (c) Ge 3d XPS results of LGPS, MCS-S-3 and MCS-S-3-Co.



Fig. S24 Ge 3d XPS results of (a) MCS-3-S and (b) MCS-3-S-Co cathodes at different potentials.



Fig. S25 a) Rate performance of MCS-3, MCS-3-LGPS and MCS-3-Co-LGPS cathode; b) Cycling performance of MCS-3-LGPS and MCS-3-Co-LGPS cathode cathode at 1.0 C. (MCS-3 refers to the pristine carbon cathode; MCS-3-LGPS refers to the cathode of ball-milled MCS-3 and LGPS; MCS-3-Co-LGPS refers to the cathode of ball-milled MCS-3, LGPS, CoS₂)



Fig. S26 Galvanostatic charge-discharge profile of MCS-3-S-Co at 0.03 C with a sulfur loading of 8 mg cm⁻².

Supporting Tables

Glucose/GO weight ratio	KOH concentration/g mL ⁻¹	Activation temperature / ℃	Specific surface area / m ² g ⁻¹
15	1	800	2861
18.75	1	800	2951
22.5	1	800	3267
30	1	800	2531
45	1	800	2283
60	1	800	1808
22.5	1	900	2511
22.5	1.5	800	2877
22.5	2	800	2493

Table S1. Specific surface area distribution with different preparation parameters.

Sample name	SSA / m ² g ⁻¹	Micropore volume / cm ³ g ⁻¹	Reference
ACC-10	1094	0.4	[1]
ACC-15	1436	0.5	
APC-1050	1496	0.8	
BP	1528	0	This Work
APC-950	1555	0.8	Ref[1]
APC-1000	1569	0.8	
AEL-1200	1602	0.7	
AEL-1100	1685	0.7	
YP-50F	1694	0.7	
APC-700-2	1700	0.8	
AEL-1000	1701	0.8	
APC-700-1	1708	0.8	
EL-104	1721	0.8	
PW-400	1724	0.8	
APC-1100	1761	0.9	
ACS-PC	1796	0.9	
МС	1800	0.5	Ref[2]
*MCS-1	1808	1.2	This Work
APC-800	1815	0.9	Ref[1]
SC-1800	1821	0.8	
EL-106	1907	0.9	
ACC-20	2004	0.8	
YP-80F	2264	1.1	
*MCS-2	2528	1.8	This Work
aMP	3160	1.6	Ref[3]
*MCS-3	3267	2.0	This Work
BL-1/4	3557	1.4	Ref[4]

 Table S2. SSA and micropore volume of different microporous carbons.

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