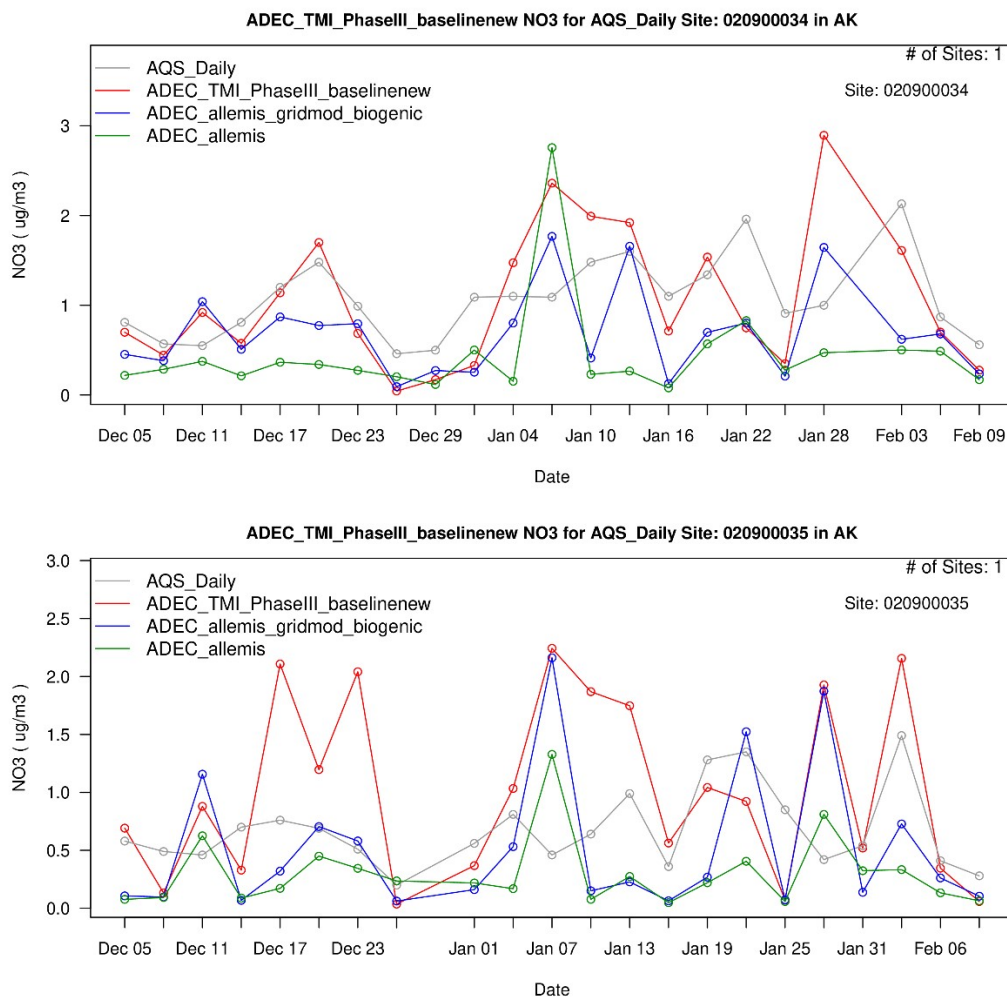
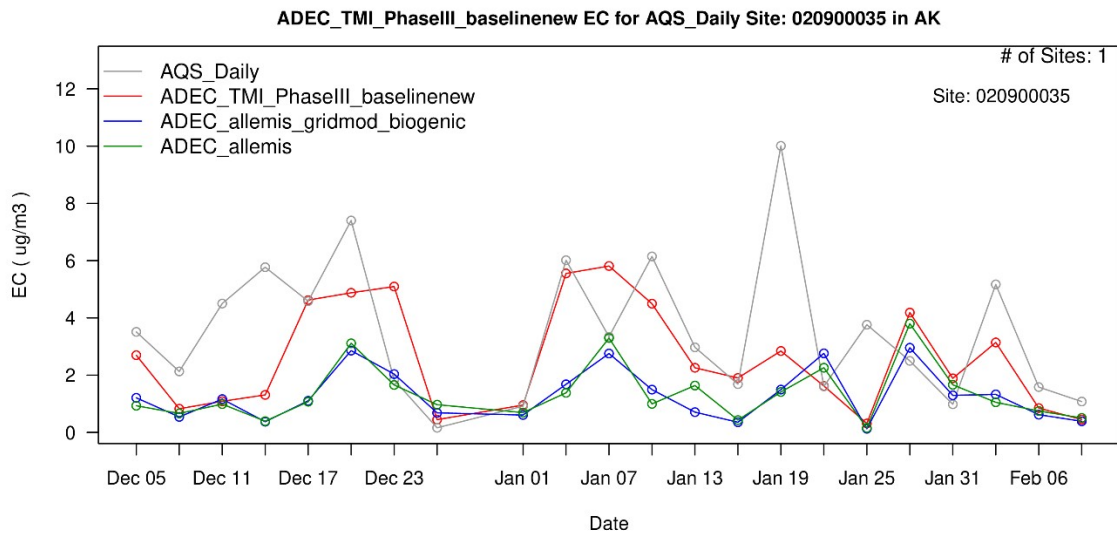
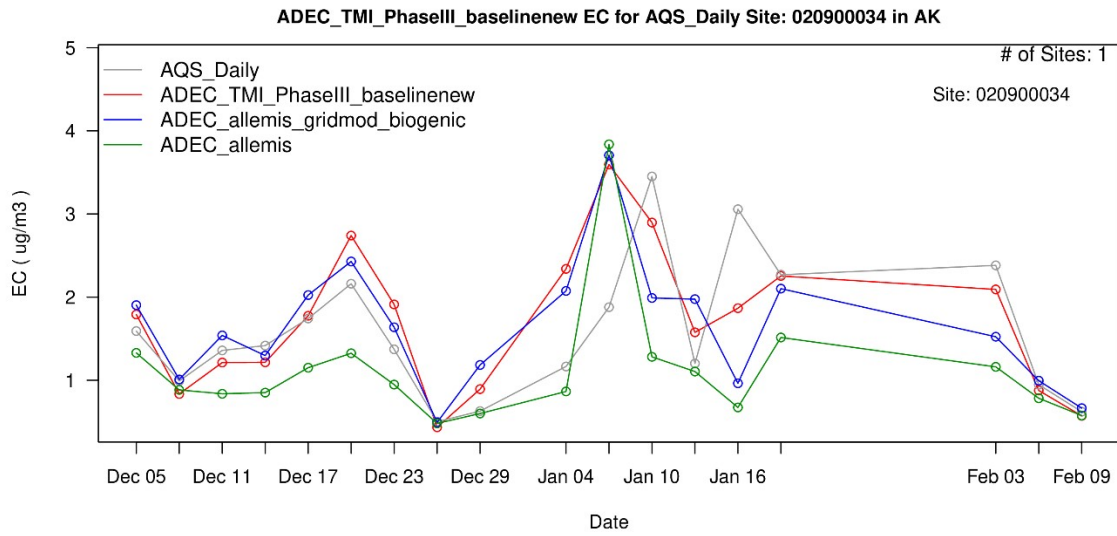


Supplementary Figures and Tables

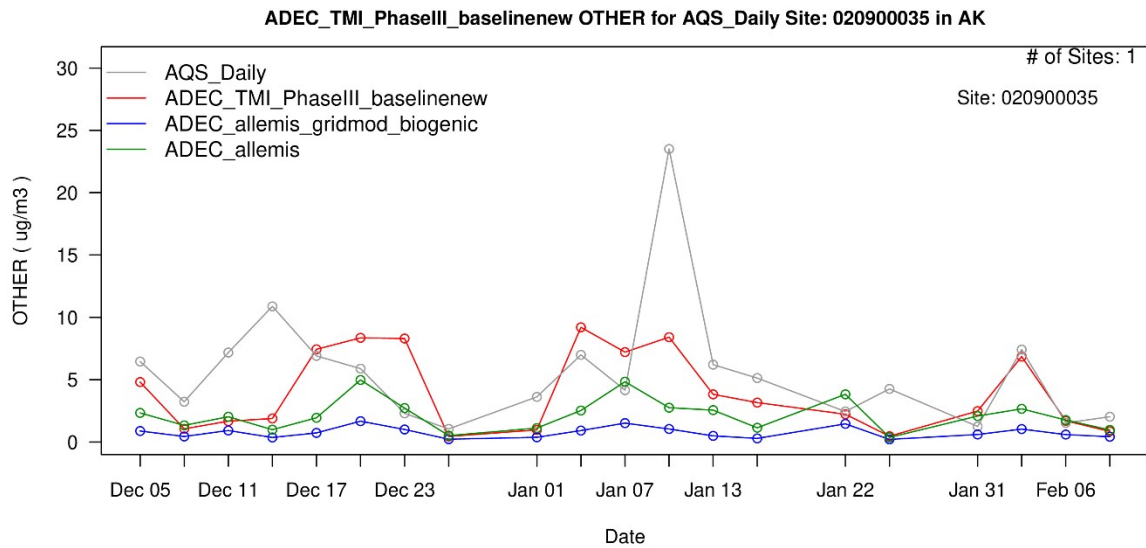
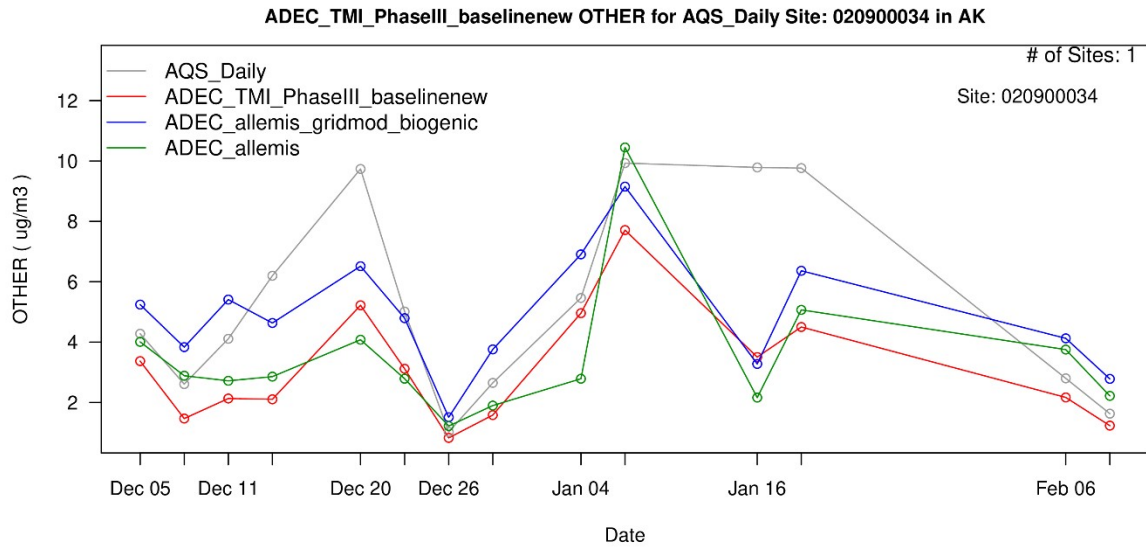
Additional Time Series Plots NO₃, EC and other



S 1: Timeseries of NO₃ at NCore (020900034) and Hurst Road (020900035) monitors for the modeling period



S 2: Timeseries of EC at NCore (020900034) and Hurst Road (020900035) monitors for the modeling period



S 3: Timeseries of OTHER at NCore (020900034) and Hurst Road (020900035) monitors for the modeling period

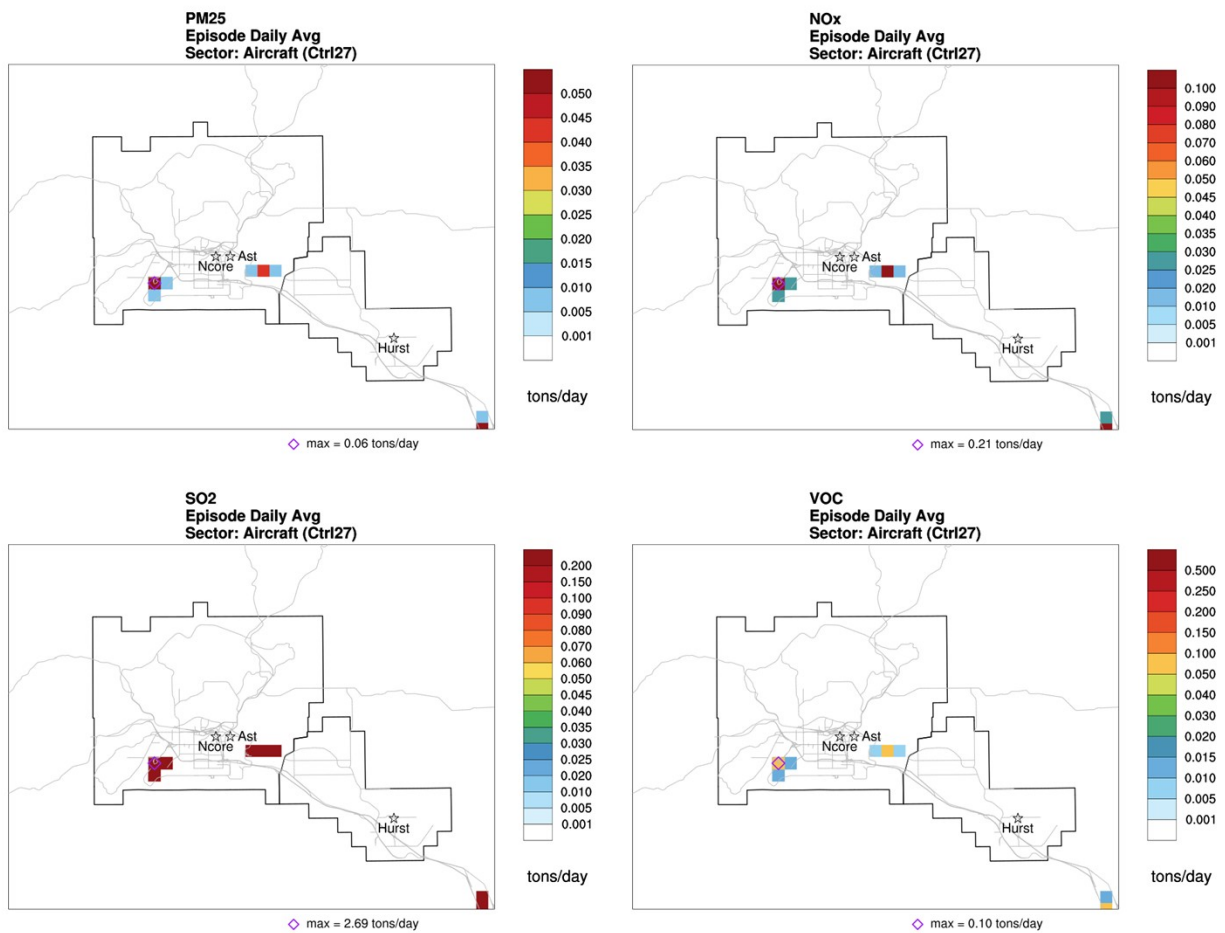
S 4: Statistics of speciated PM and total PM2.5 for the modeling episode at Hurst Road monitor

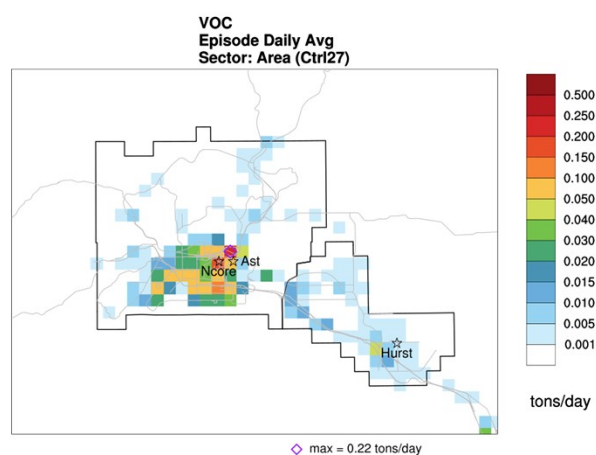
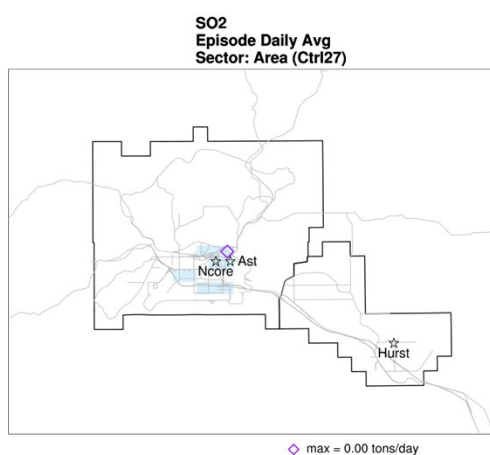
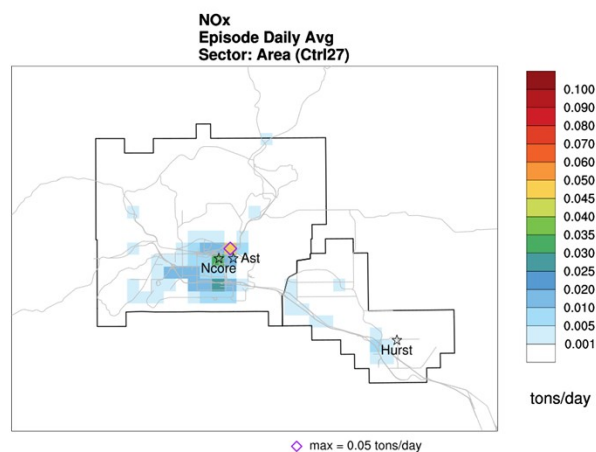
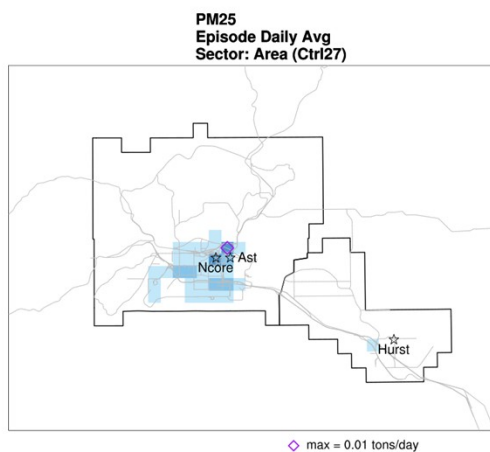
Species	Number of Observations	Mean Obs ($\mu\text{g}/\text{m}^3$)	Mean Model ($\mu\text{g}/\text{m}^3$)	Mean Bias	Mean Error	Normalized Mean Bias (NMB, %)	Normalized Mean Error (NME, %)	Fractional Bias (FB, %)	Fractional Error (FE, %)	Root Mean Square Error (RMSE)	Correlation
NO3	22	0.67	1.01	0.60	0.34	50.2	88.5	5.5	73.3	0.79	0.33
SO4	22	2.26	2.52	1.10	0.26	11.5	48.8	4.9	44.0	1.68	0.50
NH4	22	0.80	1.12	0.61	0.33	41.1	76.5	37.9	66.0	0.86	0.43
EC	22	3.53	2.60	1.74	-0.93	-26.3	49.3	-25.9	59.8	2.45	0.45
OC	22	14.47	13.58	6.42	-0.89	-6.2	44.4	-8.9	57.8	9.24	0.64
PM_TOT	21	25.08	24.96	11.06	-0.12	-0.5	44.1	-9.3	50.1	16.11	0.61

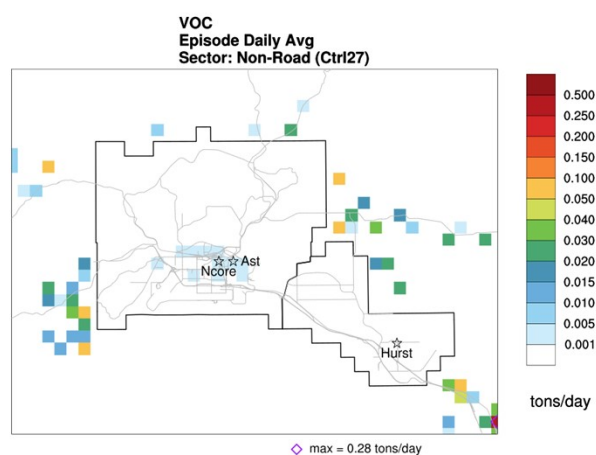
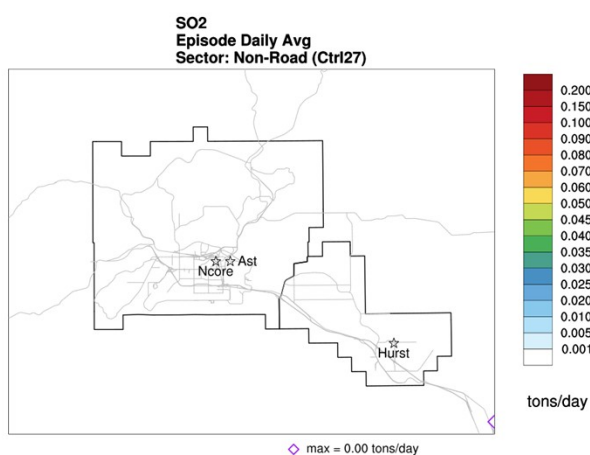
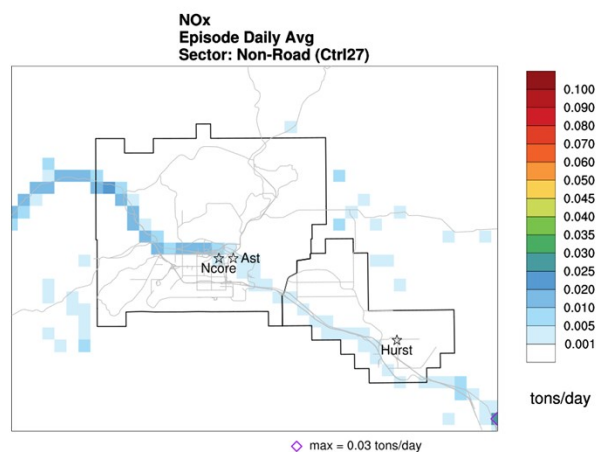
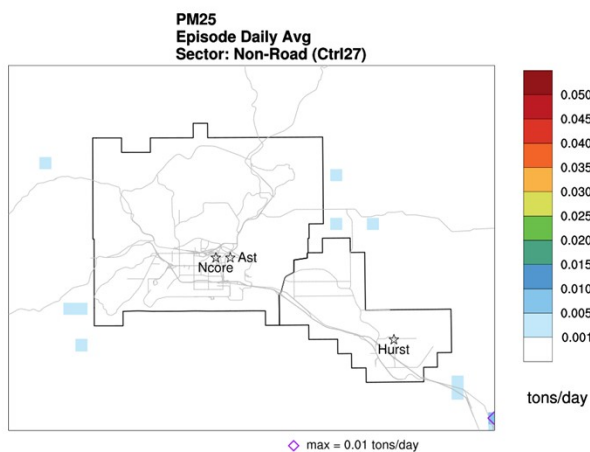
S 5: Statistics of speciated PM and total PM2.5 for the modeling episode at NCore monitor

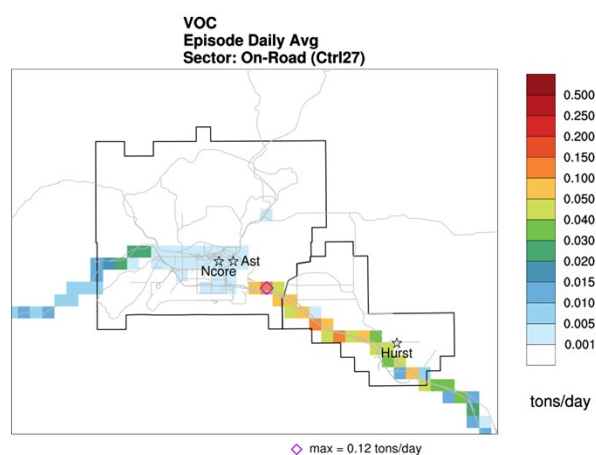
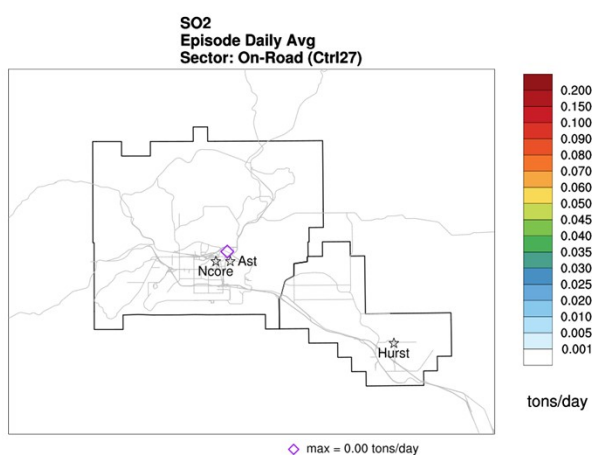
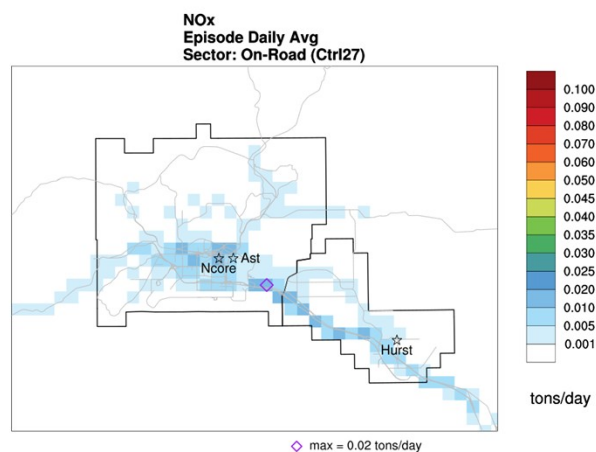
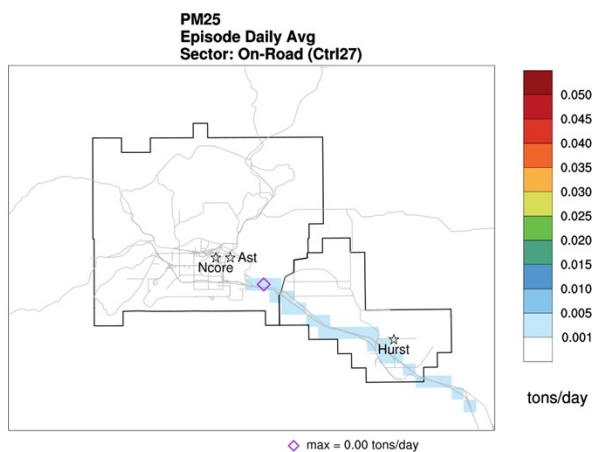
Species	Number of Observations	Mean Obs ($\mu\text{g}/\text{m}^3$)	Mean Model ($\mu\text{g}/\text{m}^3$)	Mean Bias	Mean Error	Normalized Mean Bias (NMB, %)	Normalized Mean Error (NME, %)	Fractional Bias (FB, %)	Fractional Error (FE, %)	Root Mean Square Error (RMSE)	Correlation
NO3	22	1.07	1.06	0.48	-0.01	-1.3	45.0	-22.6	52.2	0.65	0.51
SO4	22	3.41	3.29	1.21	-0.12	-3.4	35.4	-2.1	32.3	1.80	0.69
NH4	22	1.31	1.35	0.56	0.05	3.4	43.0	11.2	40.8	0.85	0.66
EC	18	1.60	1.72	0.42	0.12	7.5	26.5	6.0	23.1	0.63	0.71
OC	18	3.54	7.31	3.84	3.77	106.6	108.5	59.3	61.9	5.06	0.82
PM_TOT	18	15.31	18.16	5.68	2.85	18.6	37.1	9.7	27.8	8.52	0.72

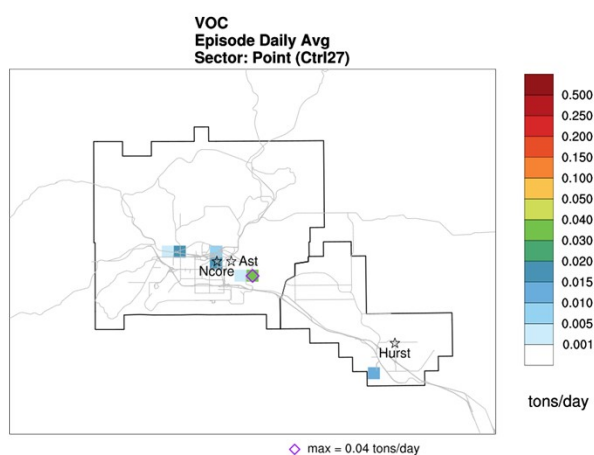
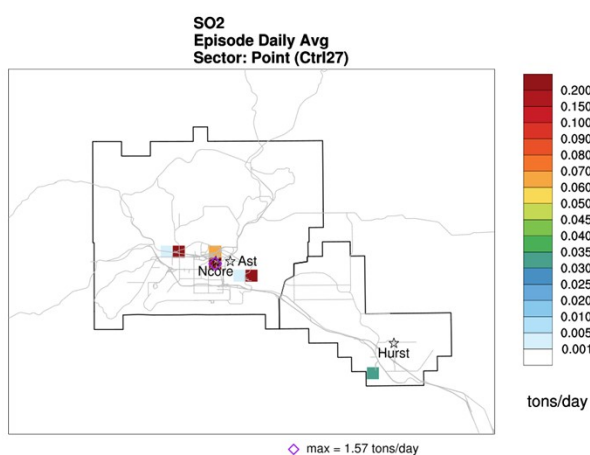
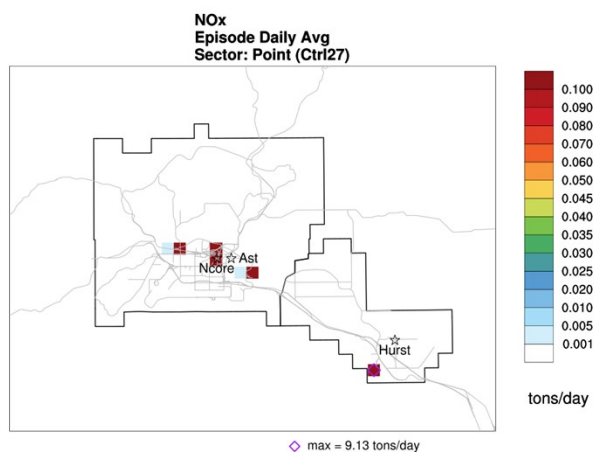
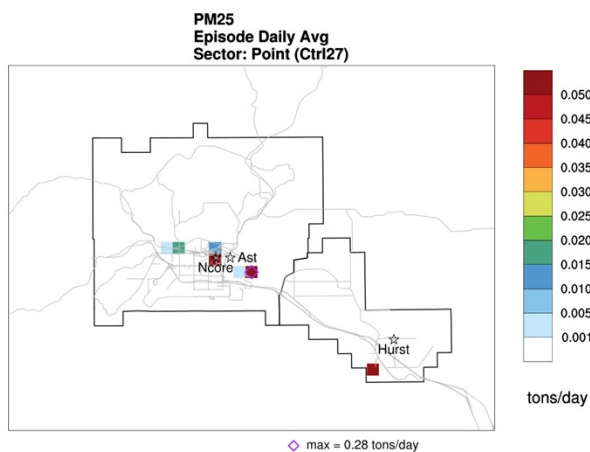
S6 Emission Plots for each sector Aircraft, Area, Non-Road, On-Road, Point, Space-Heat and pollutant (PM2.5, NO_x, SO₂, VOC) for the control year 2027

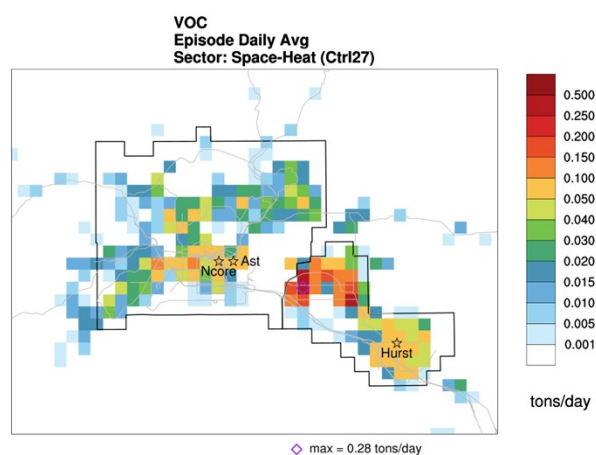
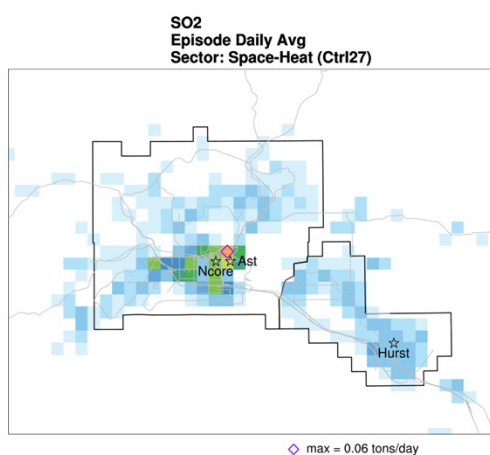
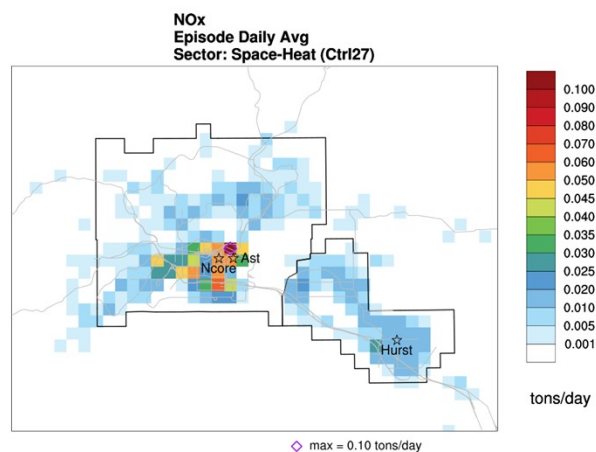
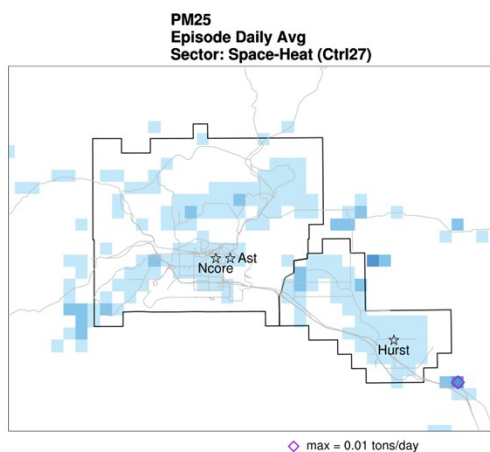












S 7 SO₂ Precursor Summary

		Episode Average			Max Daily Value		
CMAQ Sensitivity 100%		A Street	NCore	Hurst	A Street	NCore	Hurst
CMAQ - Absolute							
	SOx	-0.08703	-0.14015	-0.06109	-0.23067	-0.37986	-0.60623
CMAQ - Design Value							
	SOx	-0.21665	-0.14372	-0.21052			

The Emission Factors for space heating and for point sources are attached in excel spreadsheets:

EFs_SpaceHeating_RevSIP.xlsx

PointSource_EFs_By_EmissionUnit.xlsx

The SCC codes that were replaced with OMNI locally tested wood profiled are attached: in Speciations_SpaceHeating_OMNI.xlsx

From SMOKE, gspro_ptnonipm_cmaq_cb05_omni_100518.txt

The species specific RRFs and Future DV calculations are in the manual SMAT spreadsheet:

SMAT_091523.xls

Details on the updated to CMAQv5.3.3+

Updates were made to CMAQv5.3.3+ to add a representation of heterogeneous sulfur chemistry in aerosol water, including the formation and loss of hydroxymethanesulfonate (HMS). Aqueous aerosol sulfur reactions were parameterized with a reactive probability formulation (Cheng et al., 2016; Shao et al., 2019; Jacob, 2000) as follows:

$$\frac{dSO_4^{2-}}{dt} = k \times [SO_2 \text{ or reactant}] , \quad (1)$$

$$k = \left(\frac{r}{D_g} + \frac{4}{v\gamma} \right)^{-1} A , \quad (2)$$

$$\gamma = \left(\frac{1}{\alpha} + \frac{v}{4H^*RT\sqrt{D_a k_{chem}}} \frac{1}{f_q} \right)^{-1} , \quad (3)$$

$$f_q = \left(\coth(q) - \frac{l}{q} \right) , \quad (4)$$

$$q = r \left(\frac{k_{chem}}{D_a} \right)^{0.5} , \quad (5)$$

where r = particle radius, D_g = gas-phase diffusivity of SO_2 or reactant, v = mean molecular speed of SO_2 or reactant, A = aerosol surface area, γ = reactive uptake coefficient, α = accommodation coefficient of SO_2 or reactant, H^* = effective Henry's law coefficient of SO_2 or reactant, R = ideal gas constant, T = temperature (K), D_a = aqueous diffusion coefficient, k_{chem} =

(pseudo) first-order aqueous-phase reaction rate coefficient, and q = the diffuso-reactive parameter (Schwartz and Friberg, 1981).

The uptake gas was determined by the limiting mass transfer rate, $J_{aq,lim}$:

$$J_{aq,lim} = \min\{J_{SO_2}J_{RXT}\}, \quad (6)$$

where

$$J_x = k_{mt,x} H_x^* p_{x,\infty}, \quad (7)$$

$$k_{mt,x} = \left[\frac{r^2}{3D_g} + \frac{4r}{3\alpha v} \right]^{-1}, \quad (8)$$

where “x” = SO₂ or reactant (RXT) for a given reaction, $p_{x,\infty}$ = partial pressure of species x in the bulk gas phase, and v = mean molecular speed of x. The aqueous concentration of the other (non-limiting) reactant is rolled into the pseudo-first order reaction rate coefficient, k_{chem} [s⁻¹]. For species originating in the gas phase, the aqueous concentration [M] is estimated with the effective Henry’s law coefficient and partial pressure, $[X] = H_x^* p_{x,\infty}$. Concentrations [M] of species that primarily reside in the aerosol phase are estimated using the mass concentration of the species in air [$\mu\text{g m}^{-3}$ air], molecular weight [g mol⁻¹], and the accumulation mode volume concentration [$\text{m}^3 \text{m}^{-3}$ air]. [H⁺], inorganic aerosol liquid water content (ALW), and the phase distribution of semi-volatile inorganic species is provided by ISORROPIA II (Fountoukis and Nenes, 2007). Water associated with organics is estimated as described in Pye et al. (2017) with total aerosol liquid water content equal to the sum of inorganic and organic water.

Significant uncertainty lies in the S(IV) oxidation rates in aerosol water, where concentrations are orders of magnitude higher than the dilute (cloud-like) conditions under which many of the aqueous-phase oxidation rates were determined. The rates may be enhanced or inhibited with increasing ionic strength, and this is an active area of research (Liu et al., 2021). Here we applied ionic strength (enhancement and/or inhibition) factors (ISF) to adjust rate coefficients when shifting from a cloud- to aerosol-water environment. Ionic strength used to calculate aerosol water ISFs is estimated from the concentrations of the transported dissociating aerosol species but limited to the maximum ionic strength in the experiments investigating a particular ISF. As in Farrell et al. (2024), ISFs were used to adjust the rate coefficients for the oxidation of SO₂ by H₂O₂, O₃, and transition metal-catalyzed O₂ (TMI) as well as the effective Henry’s law coefficients for SO₂ and H₂O₂. Also included is the ionic-strength dependent reaction rate for the reaction of S(IV) with NO₂ as described in Chen et al. (2019). In the cases where multiple oxidation rate expressions have been suggested in the literature (i.e., TMI and NO₂), the rates have been averaged with equal weighting. For more information on rate expressions, coefficients, and other parameters, please refer to Fahey et al. (in preparation) and Farrell et al. (2024).

Cloud/fog chemistry is modeled using CMAQ's KMT2 cloud option (Fahey et al., 2017; Fahey et al., submitted). This cloud module includes the formation and loss of HMS as well as inorganic S(IV) oxidation by H₂O₂, O₃, transition metal-catalyzed O₂, methyl hydroperoxide, peroxyacetic acid, peroxyntic acid, NO₂, OH and NO₃ radicals, and N(III). ISFs are not applied to chemical rates in cloud/fog water.

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