

Supplementary Information

Feasibility and sustainability of emerging CCU pathways for formic acid production

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VLE validation results for important pairs involved in separation procedure based on the experimental data

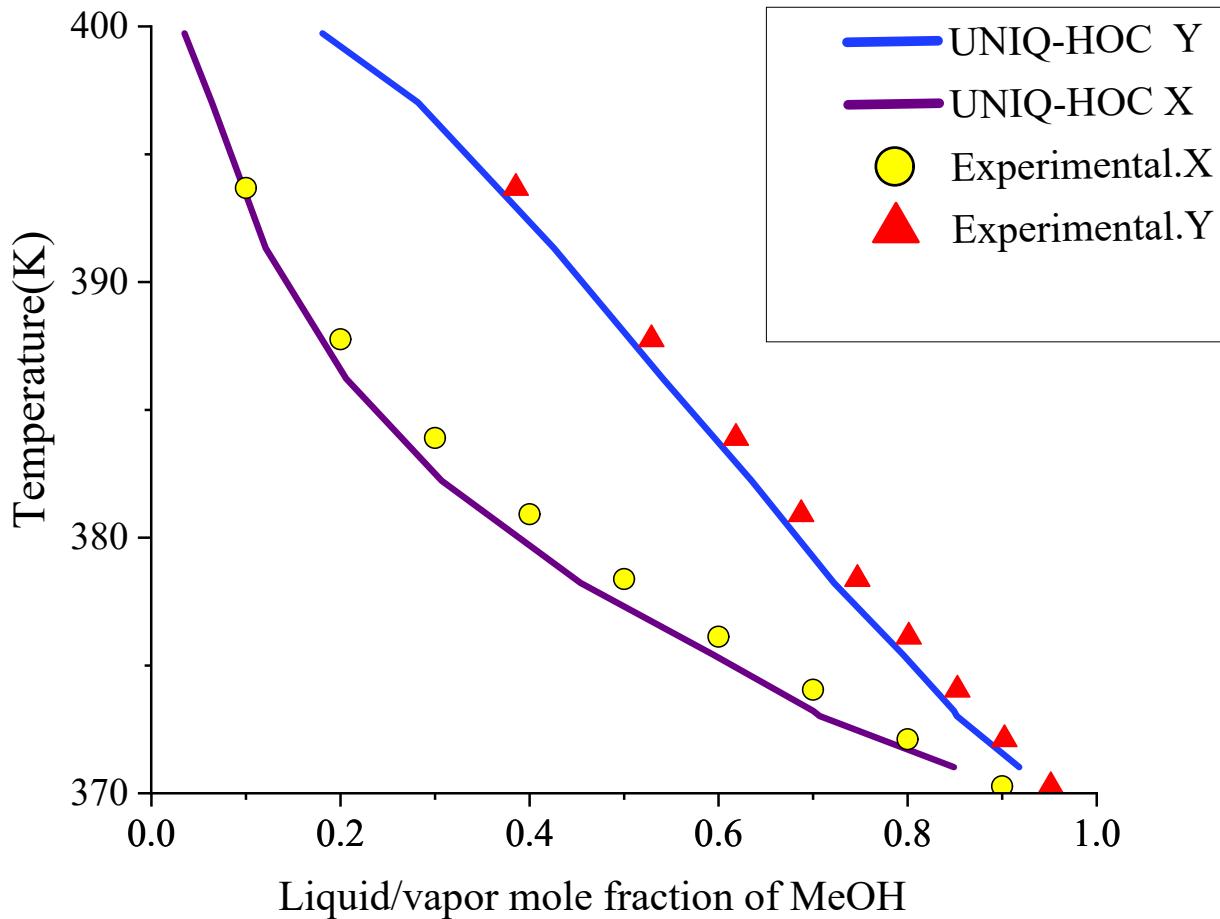


Fig.S1. T_{XY} diagram for H₂O/MeOH (3 bar) for the conventional process. Experimental data is form Hirata et al¹.

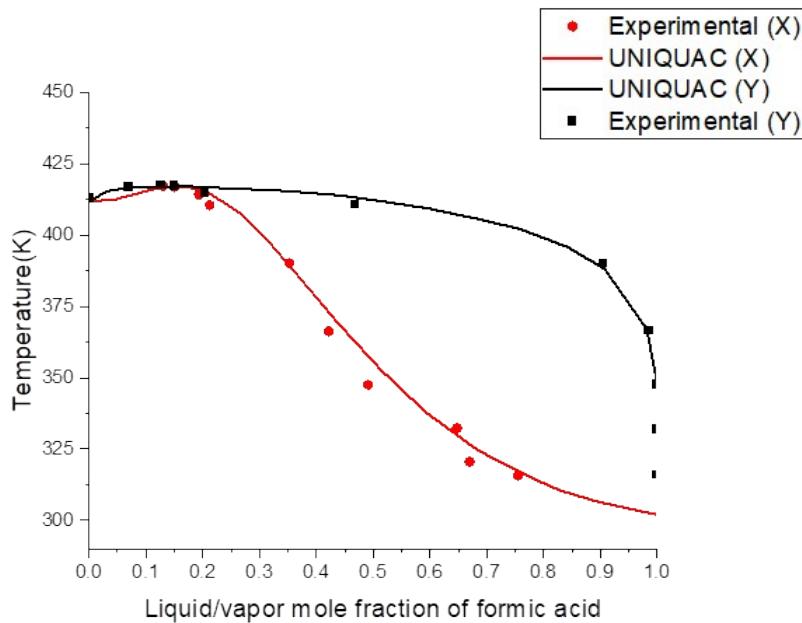


Fig.S2. T_{XY} diagram Formic acid/n-butyleimidazol mixture for the thermo-catalytic process at 50 mmHg².

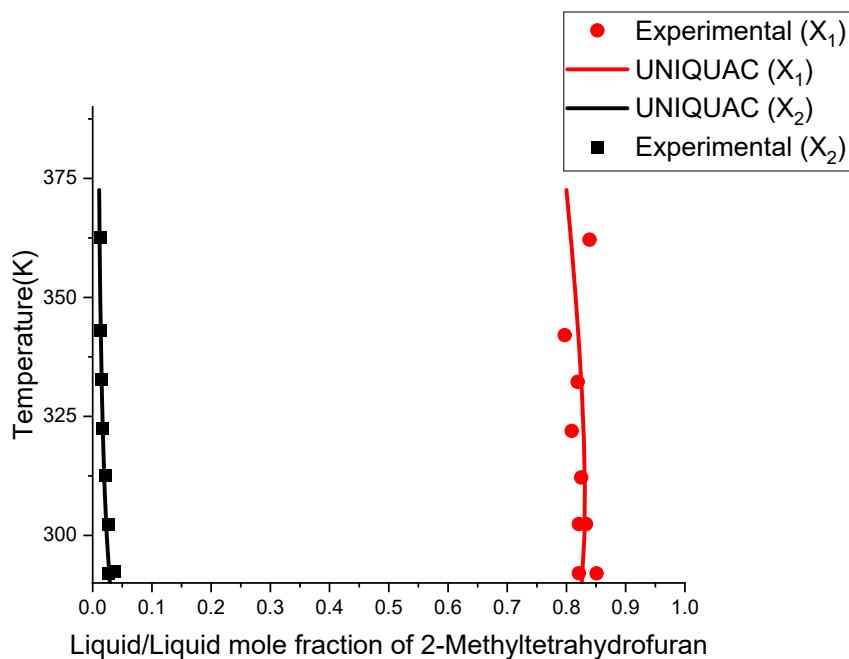


Fig.S3. T_{XX} diagram of water and 2-methyltetrahydrofuran LLE data for the electrochemical reduction process. The Experimental data was obtained from M. Glass et al (2017)³.

Table S1. Process operating conditions and parameters used for designing the electrochemical reduction downstream process.

| Unit operation | Parameter | Value | Reference |
|-------------------------|-----------------------------|--|-----------------------------------|
| PSA | Operating conditions | 400 °C, 30 bar | X. Zhu et al. ⁴ |
| | H ₂ purification | 99.99 mol% | |
| | Catalyst | Fe ₂ O ₃ -Cr ₂ O ₃ | |
| Extractor | Operating conditions | 40 °C, 1 bar | A.T. Laitinen et al. ⁵ |
| Azeotropic Distillation | Operating conditions | 1.1 bar, equilibrium | A.T. Laitinen et al. ⁵ |
| Striper | Operating conditions | 4 bar, equilibrium | A.T. Laitinen et al. ⁵ |

Table S2. CO₂ electrolyzer technoeconomic assumption parameters and process operating conditions.

| Parameter | Value | Reference |
|--|--------------------------|---------------------------------------|
| Electrolyzer faradaic efficiency (HCOOH) | 90% | X. Jiang et al. ⁶ |
| Current density | 500 A/cm ² | X. Jiang et al. ⁶ |
| catalyst | BiO _n cluster | X. Jiang et al. ⁶ |
| HCOOH concentration | ~5.04 g/l | K. Fernández-Caso et al. ⁷ |
| Total CO ₂ conversion | 44% | X. Jiang et al. ⁶ |
| Reference electrolyzer total cost per area | \$48,493/m ² | J.M. Spurgeon et al. ⁸ |
| scaling factor | 0.65 | |

Table S3. PSA gas separation and recycle technoeconomic assumption parameters⁹.

| Parameter | Value |
|---|---|
| Reference total gas flow to separator | 4.40 x 10 ⁶ mol/h |
| Reference separator/recycle cost, 2018 USD | 11,214,299 |
| Scaling exponent | 0.65 |
| CO ₂ recycling efficiency | 97% |
| CO loss to recycle | 0% |
| Case-dependent total gas flow to separator | Y mol/h |
| CO ₂ /CO separator/recycle scaled capital expense | $[Y/(4.40 \times 10^6)]^{0.65} \times (11,214,299)$ |
| Reference power to operate separator/recycle | 10 kW |
| Reference annual energy usage to operate separator/recycle 8.76 x 10 ⁴ kWh/yr | 8.76 x 10 ⁴ kWh/yr |

Table S4. List of parameters with their cost reference

| Materials and utilities | Cost |
|--|-----------------------------|
| Gray H ₂ (USD/kg) | 1.88 ¹⁰ |
| Green H ₂ (USD /kg) | 3.53 ¹¹ |
| Captured CO ₂ (USD /kg) | 0.035 ¹² |
| TREA (USD /kg) | 2.5 (from www.alibaba.com) |
| Catalyst (CH ₃ ONa, in USD /kg) | 3.5 (from www.alibaba.com) |
| Catalyst (Ru/bpyTN-30-CTF, in USD /g) | 175.35/g ¹³ |
| Catalyst (BiO _n , in USD /g) | 1.8/g ¹⁴ |
| KOH (USD /kg) | 0.6 (from www.alibaba.com) |
| 2-Methyltetrahydrofuran (USD /kg) | 2.5 (from www.alibaba.com) |
| O ₂ (Tonne /hr) | 38.07(from www.alibaba.com) |
| CO (USD /Tonne) | 70 ¹⁵ |
| Methanol (USD /Tonne) | 420(from www.alibaba.com) |
| CO ₂ from DAC (USD /kg) | 0.17 ¹⁷ |
| Cooling water (USD 0.028/ Tonne) | 0.354 ¹⁶ |
| Electricity (USD /kWh) | 0.06 ¹⁶ |
| LP Steam (USD /GJ) | 14.05 ¹⁶ |
| MP steam (USD /GJ) | 14.83 ¹⁶ |
| HP Steam (USD/GJ) | 17.7 ¹⁶ |

Table S5. Total capital investment estimation used based on the ratio factor for fluid processing plant.

| | Ratio factor for fluid processing plant (of delivered equipment cost) | Reference |
|---|--|-------------------------------|
| Direct cost | | MS Peters et al ¹⁸ |
| Delivery, % of purchased equipment | 0.1 | |
| Subtotal: Purchased equipment (delivered) | 1 | |
| Purchased equipment installation | 0.47 | |
| Instrumentation and Controls | 0.36 | |
| Piping | 0.68 | |
| Electrical systems | 0.11 | |
| Buildings (including services) | 0.18 | |
| Yard improvements | 0.1 | |
| Service facilities | 0.7 | |
| Indirect cost | | |
| Engineering and supervision | 0.33 | |
| Construction expenses | 0.41 | |
| Legal expences | 0.04 | |
| Contractor's fee | 0.22 | |
| Contingency | 0.44 | |
| Working capital (WC) | 0.89 | |

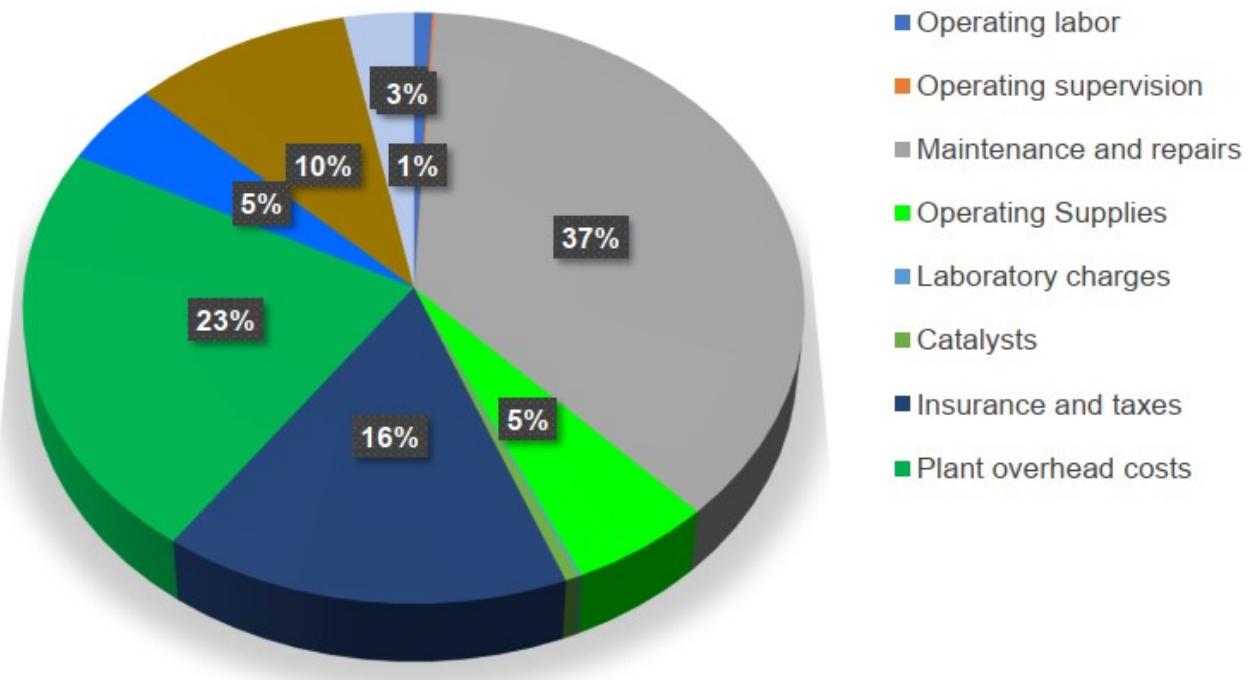


Fig.S4. cost break down of other operational expenses for the electrochemical reduction pathway (CCUER)

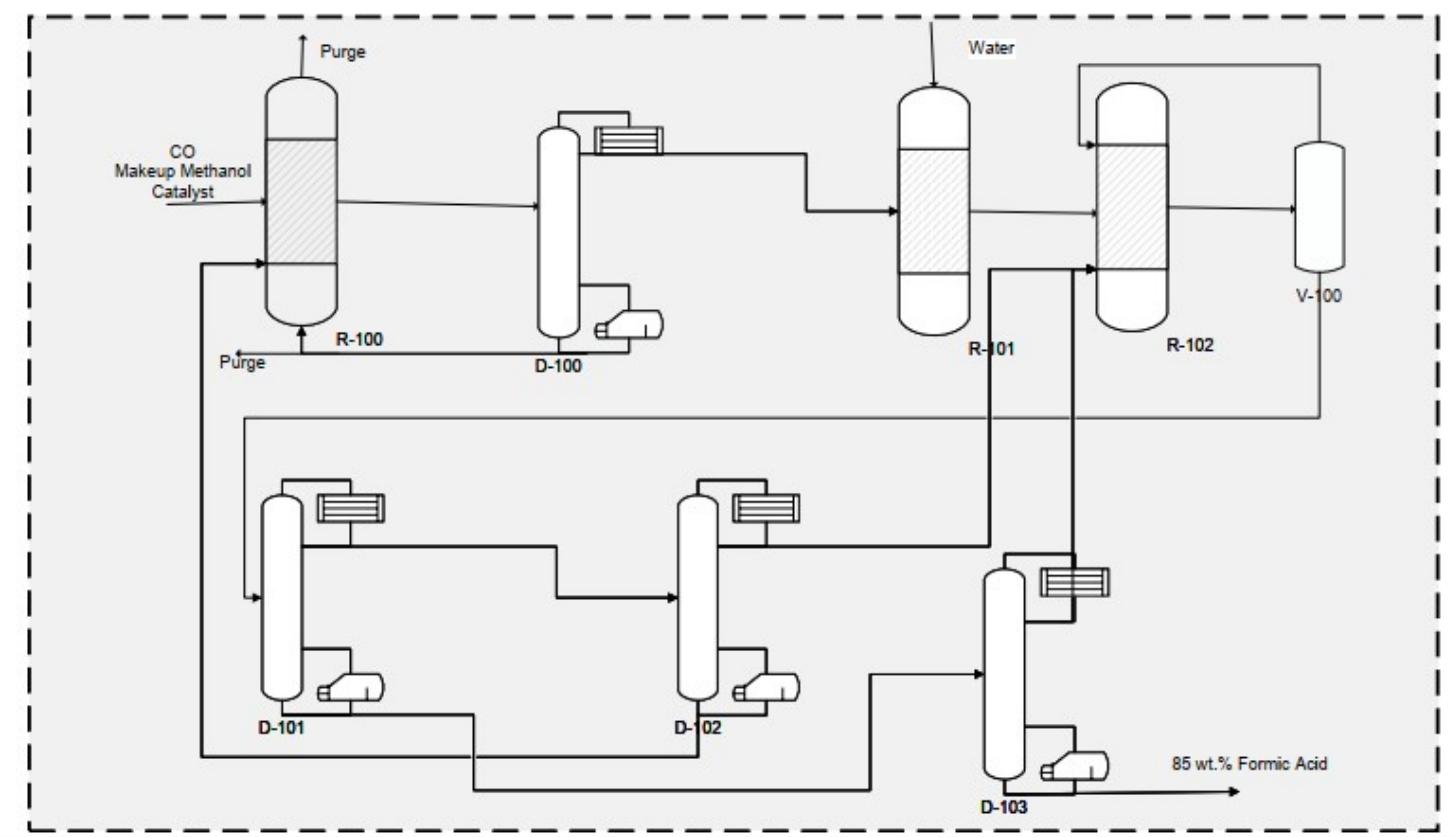


Fig.S5. System boundary of the conventional process pathway

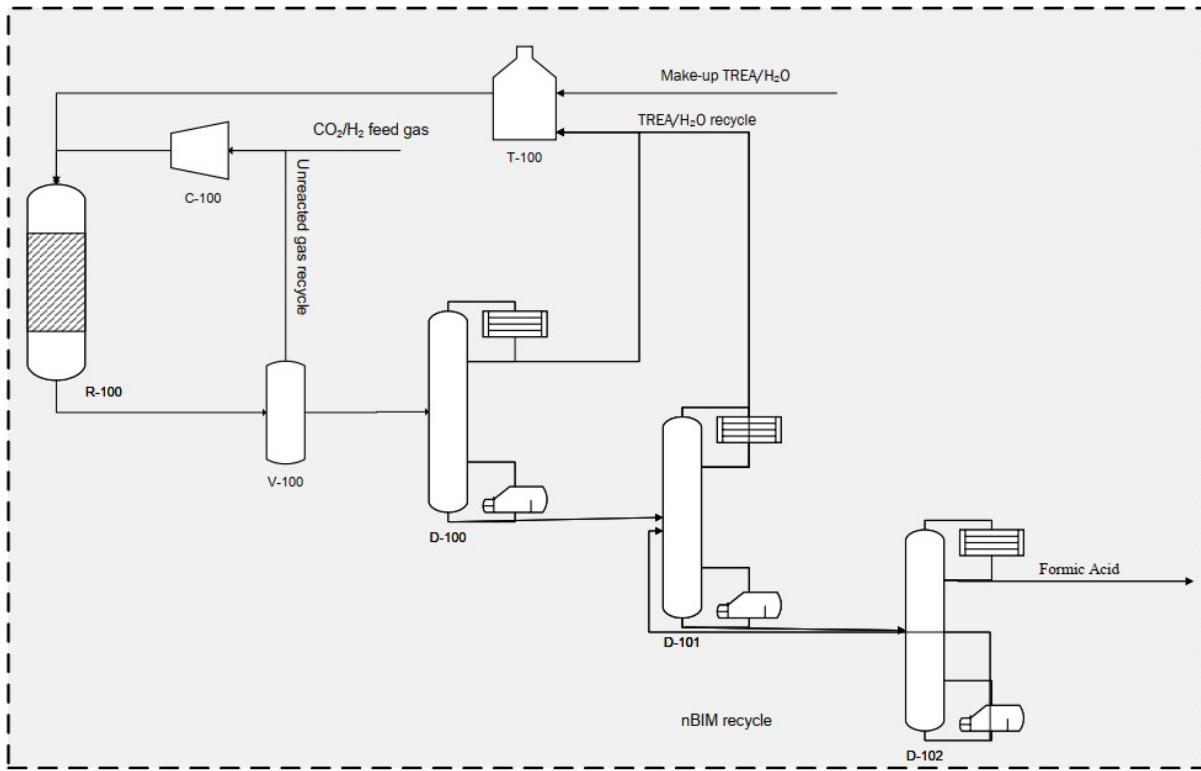


Fig.S6. System boundary of the thermo-catalytic process pathway

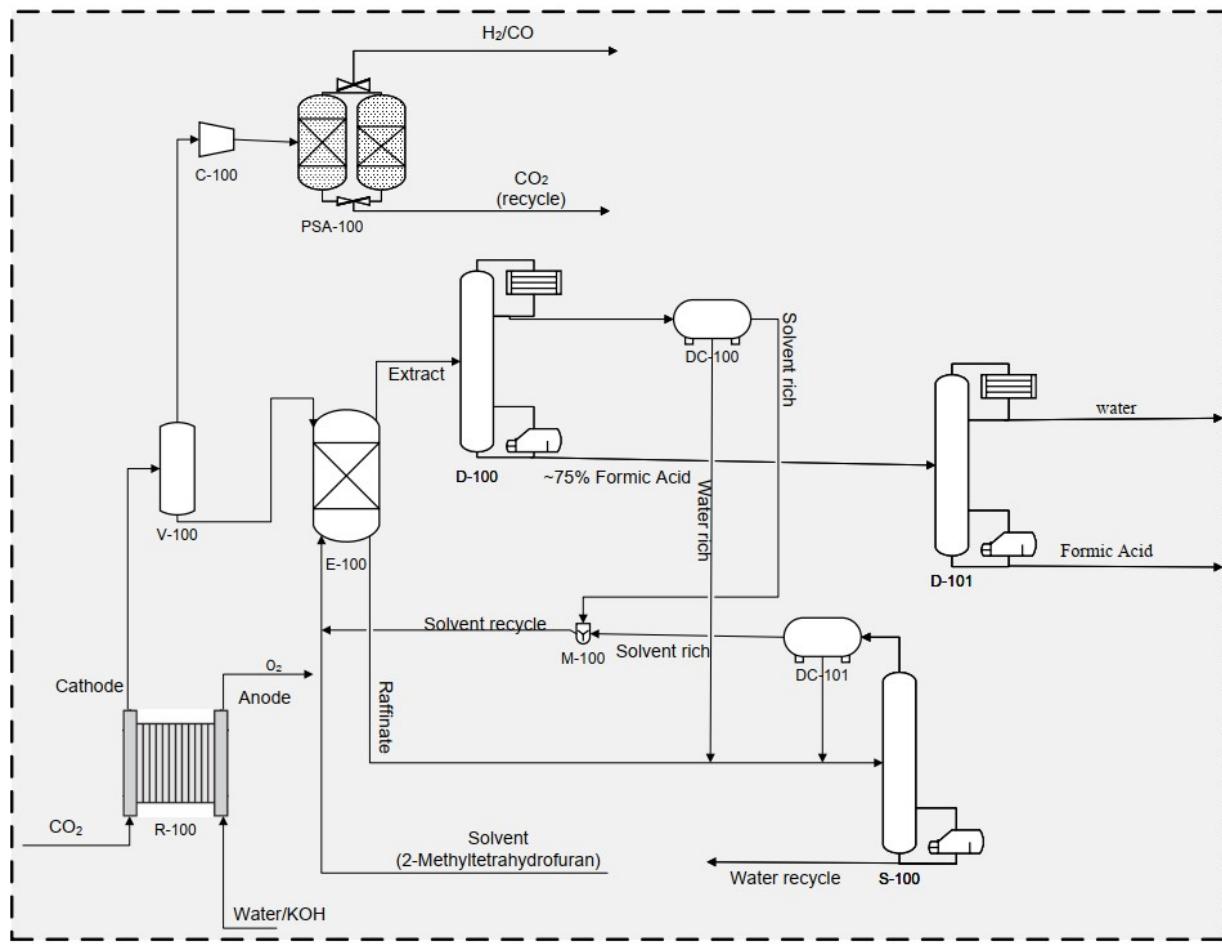


Fig.S7. System boundary of the electrochemical reduction process pathway

Table S6. LCA methodology and GWI values of the conventional pathway compared to previous works.

| Reference | This work | N. Thonemann et al (2019)²⁰ | D. Kang et al (2021)²¹ | Y. Ahn et al (2019)²³ |
|-----------------------------|---------------------------------|---|--|---|
| Scope | Cradle to gate | Cradle to gate | Cradle to gate | Cradle to gate |
| Software | SimaPro LCA | openLCA 1.7.4 | SimaPro LCA | SimaPro LCA |
| Database | Ecoinvent 3.9 | — | Ecoinvent 3 | — |
| LCIA method | ReCiPe 2016 | ILCD 1.0.8 (2016) | ReCiPe 2008 | ReCiPe 1.13 |
| Lifetime horizon of the LCA | 100 years (GWP100) | 100 years (GWP100) | 100 years (GWP100) | 100 years (GWP100) |
| Purity | 85.0 wt% FA | 85.0 wt% FA | 98 wt% | 98% |
| GWI | 4.23 kgCO _{2eq} /kg FA | 4.4 kgCO _{2eq} /kg FA | 2.04 kgCO _{2eq} /kg FA | 2.007 kgCO _{2eq} /kg FA |

Table S7. LCA methodology and GWI values of the thermo-catalytic pathway compared to previous works.

| Reference | This work | D. Kang et al (2021)²¹ | Y. Ahn et al (2019)²³ |
|-----------------------------|--|--|---|
| Scope | Cradle to gate | Cradle to gate | Cradle to gate |
| Software | SimaPro LCA | SimaPro LCA | SimaPro LCA |
| Database | Ecoinvent 3.9 | Ecoinvent 3 | — |
| LCIA method | ReCiPe 2016 | ReCiPe 2008 | ReCiPe 1.13 |
| Lifetime horizon of the LCA | 100 years (GWP100) | 100 years (GWP100) | 100 years (GWP100) |
| Purity | 85.0 wt% FA | 99 wt% | 98% |
| GWI | 0.75 to 2.59 kg CO _{2eq} /kg FA | 0.27 kg CO _{2eq} /kg FA | 0.098 kgCO _{2eq} /kg FA |

Table S8. LCA methodology and GWI values of the electrochemical reduction pathway compared to previous works.

| Reference | This work | A. Dominguez-Ramos et al (2015) ¹⁹ | N. Thonemann et al (2019) ²⁰ | D. Kang et al (2021) ²¹ | A. Banu et al (2023) ²² |
|-----------------------------|---|---|---|------------------------------------|------------------------------------|
| Scope | Cradle to gate | Cradle to gate | Cradle to gate | Cradle-to-gate | Cradle-to-gate |
| Software | SimaPro LCA | — | openLCA 1.7.4 | SimaPro LCA | GaBi professional software |
| Database | Ecoinvent 3.9 | Ecoinvent v2.2 | Ecoinvent 3.4 | Ecoinvent 3 | — |
| LCIA method | ReCiPe 2016 | ReCiPe, 2011 | ILCD 1.0.8 (2016) | ReCiPe 2008 | ReCiPe 1.08 |
| Lifetime horizon of the LCA | 100 years (GWP100) | 100 years (GWP100) | 100 years (GWP100) | 100 years (GWP100) | 10 years (GWP10) |
| Purity | 85 wt.% | 84 wt.% formate solution | 85.0 wt.% FA | 85 wt.% | — |
| GWI | 2.91 to 8.23 kgCO _{2eq} /kg FA | 32 to 519 kg CO _{2eq} / kg HCOO ⁻ | 4.2 to 7.2 kg CO2-eq/kg FA | 2 kgCO _{2eq} /kg FA | 3.27 kgCO _{2-eq} /kg FA |

Table S9. Total electricity and heat use per unit of product of the electrochemical reduction pathway

| Reference | Electricity (MWh/Tonne FA) | Heat (MJ/Tonne FA) |
|---|----------------------------|--------------------|
| This work | 2.98 | 13.356 |
| D. Kang et al (2021) ²¹ | 4.79 | 21.78 |
| R. Aldoco et al (2019) ²⁴ | 4.59 | 62 |
| A. Dominguez-Ramos et al (2015) ¹⁹ | 8.259-16 | 35-5169 |

Table S10. Yield of formic acid for the electrochemical reduction pathway

| Reference | CO ₂ Conversion (%) | Faraday's efficiency (%) | Formic acid yield (%) |
|------------------------------------|--------------------------------|--------------------------|-----------------------|
| This work | 44 | 90 | 39.6 |
| Z. Xing et al (2021) ²⁵ | 35 | 83 | 29.05 |
| Qian et al (2020) ²⁶ | 36 | 81 | 29.16 |

Table S11. Heat and mass balance of Conventional pathway

| Stream Name | Units | Feed to R-100(CO) | MeOH with Recycle | R-100 to R- 101 | Water feed to R-101 | D-101 to D-103 | D-103 |
|----------------|---------|----------------------|-------------------|--------------------|------------------------|----------------|--------------|
| Description | | | | | | | |
| From | | | | | | D-101 | D-103 |
| To | | R-100 | R-100 | R-101 | R-101 | D-103 | |
| Stream Class | | CONVEN | CONVEN | CONVEN | CONVEN | CONVEN | CONVEN |
| Temperature | C | 80 | 79.8526401 | 80 | 80 | 158.6559053 | 25 |
| Pressure | bar | 40 | 40 | 3 | 3 | 4.053 | 4 |
| Enthalpy Flow | cal/sec | - | -15053754234 | -29683423.11 | - | -17969715.05 | -9235223.017 |
| | | 156789008.6 | | | 382529113.6 | | |
| Average MW | | 28.0104 | 28.40887903 | 41.87809528 | 18.01528 | 25.91353893 | 36.82065841 |
| Mole Flows | kmol/hr | 21650 | 2022439.133 | 1543.50942 | 20460 | 880.1403768 | 369.6589583 |
| CO | kmol/hr | 21650 | 1996341.702 | 205.7666268 | 0 | 2.91036E-35 | 0 |
| MF | kmol/hr | 0 | 25015.03052 | 827.2613747 | 0 | 1.59179E-21 | 0 |
| METHANOL | kmol/hr | 0 | 1082.4 | 0 | 0 | 4.45641E-13 | 0 |
| WATER | kmol/hr | 0 | 0 | 510.4814185 | 20460 | 631.9619648 | 121.4805463 |
| FA | kmol/hr | 0 | 0 | 3.00921E-09 | 0 | 248.178412 | 248.178412 |
| Mole Fractions | | | | | | | |
| CO | | 1 | 0.987096061 | 0.133310898 | 0 | 3.3067E-38 | 0 |
| MF | | 0 | 0.012368743 | 0.535961338 | 0 | 1.80857E-24 | 0 |
| METHANOL | | 0 | 0.000535195 | 0 | 0 | 5.06329E-16 | 0 |
| WATER | | 0 | 0 | 0.330727764 | 1 | 0.718024058 | 0.328628709 |
| FA | | 0 | 0 | 1.94959E-12 | 0 | 0.281975942 | 0.671371291 |
| Mass Fractions | | | | | | | |
| CO | | 1 | 0.973250493 | 0.089165745 | 0 | 3.57427E-38 | 0 |
| MF | | 0 | 0.026145864 | 0.768560514 | 0 | 4.19121E-24 | 0 |

| | | | | | | |
|----------|---|-------------|-------------|---|-------------|-------------|
| METHANOL | 0 | 0.000603643 | 0 | 0 | 6.26078E-16 | 0 |
| WATER | 0 | 0 | 0.142273741 | 1 | 0.499175527 | 0.150788494 |
| FA | 0 | 0 | 2.14267E-12 | 0 | 0.500824473 | 0.849211506 |

Table S12. Heat and mass balance of Thermo-catalytic pathway

| Stream Name | Units | CO ₂ | H ₂ | TREA | H ₂ O | R-100 feed with recycle | From R-100 to V-100 |
|------------------|---------|-----------------|----------------|----------|------------------|-------------------------|---------------------|
| Description | | | | | | | |
| From | | | | | | | R-100 |
| To | | C100 | C100 | T-100 | T-100 | R-100 | V-100 |
| Stream Class | | | CONVEN | CONVEN | CONVEN | CONVEN | CONVEN |
| Temperature | C | 25 | 25 | 25 | 25 | 120 | 120 |
| Pressure | bar | 1 | 30 | 120 | 120 | 120 | 120 |
| Enthalpy Flow | cal/sec | -15990156.5 | 1025.9174 | -3933304 | -62473484 | -79591740.11 | -81137215.78 |
| Average MW | | 44 | 2 | 101 | 18 | 24.81784387 | 28.26006907 |
| Mole Flows | kmol/hr | 612 | 612 | 324 | 3294 | 4842 | 4254.48 |
| CO ₂ | kmol/hr | 612 | 0 | 0 | 0 | 612 | 318.24 |
| H ₂ | kmol/hr | 0 | 612 | 0 | 0 | 612 | 318.24 |
| H ₂ O | kmol/hr | 0 | 0 | 0 | 3294 | 3294 | 3294 |
| N ₂ | kmol/hr | 0 | 0 | 0 | 0 | 0 | 0 |
| O ₂ | kmol/hr | 0 | 0 | 0 | 0 | 0 | 0 |

| | | | | | | | |
|------------------|---------|---|---|-----|---|-------------|-------------|
| NTRIET | kmol/hr | 0 | 0 | 0 | 0 | 0 | 0 |
| TREA | kmol/hr | 0 | 0 | 324 | 0 | 324 | 30.24 |
| AMMON-01 | kmol/hr | 0 | 0 | 0 | 0 | 0 | 0 |
| HCOOH | kmol/hr | 0 | 0 | 0 | 0 | 0 | 0 |
| TREA-FO | kmol/hr | 0 | 0 | 0 | 0 | 0 | 293.76 |
| N-ISO-01 | kmol/hr | 0 | 0 | 0 | 0 | 0 | 0 |
| Mole Fractions | | | | | | | |
| CO ₂ | | 1 | 0 | 0 | 0 | 0.126394052 | 0.074801151 |
| H ₂ | | 0 | 1 | 0 | 0 | 0.126394052 | 0.074801151 |
| H ₂ O | | 0 | 0 | 0 | 1 | 0.680297398 | 0.774242681 |
| N ₂ | | 0 | 0 | 0 | 0 | 0 | 0 |
| O ₂ | | 0 | 0 | 0 | 0 | 0 | 0 |
| NTRIET | | 0 | 0 | 0 | 0 | 0 | 0 |
| TREA | | 0 | 0 | 1 | 0 | 0.066914498 | 0.007107802 |
| AMMON-01 | | 0 | 0 | 0 | 0 | 0 | 0 |
| HCOOH | | 0 | 0 | 0 | 0 | 0 | 0 |
| TREA-FO | | 0 | 0 | 0 | 0 | 0 | 0.069047216 |
| N-ISO-01 | | 0 | 0 | 0 | 0 | 0 | 0 |
| Mass Fractions | | | | | | | |
| CO ₂ | | 1 | 0 | 0 | 0 | 0.224086279 | 0.116462937 |

| | | | | | | |
|------------------|---|---|---|---|-------------|-------------|
| H ₂ | 0 | 1 | 0 | 0 | 0.01018574 | 0.00529377 |
| H ₂ O | 0 | 0 | 0 | 1 | 0.493409227 | 0.493146999 |
| N ₂ | 0 | 0 | 0 | 0 | 0 | 0 |
| O ₂ | 0 | 0 | 0 | 0 | 0 | 0 |
| NTRIET | 0 | 0 | 0 | 0 | 0 | 0 |
| TREA | 0 | 0 | 1 | 0 | 0.272318754 | 0.025402909 |
| AMMON-01 | 0 | 0 | 0 | 0 | 0 | 0 |
| HCOOH | 0 | 0 | 0 | 0 | 0 | 0 |
| TREA-FO | 0 | 0 | 0 | 0 | 0 | 0.359693385 |
| N-ISO-01 | 0 | 0 | 0 | 0 | 0 | 0 |

| Stream Name | Units | V-100 to D-100 | D-100 to D-101 | nBIM | D-101 to D-102 | D-102 Distillate |
|---------------|---------|----------------|----------------|--------------|----------------|------------------|
| Description | | | | | | |
| From | | V-100 | D-100 | D-102 | D-101 | D-102 |
| To | | D-100 | D-101 | D-101 | D-102 | |
| Stream Class | | CONVEN | CONVEN | CONVEN | CONVEN | CONVEN |
| Temperature | C | 100 | 34.61082716 | 25 | 100 | 29.11059499 |
| Pressure | bar | 120 | 0.199983553 | 0.199983553 | 0.133322368 | 0.066661184 |
| Enthalpy Flow | cal/sec | -74789011.65 | -11598123.81 | -1048297.032 | -9558658.557 | -8686064.137 |
| Average MW | | 27.98776683 | 66.62370624 | 124.18576 | 76.24208293 | 37.50160635 |

| | | | | | | |
|------------------|---------|-------------|-------------|-----|-------------|-------------|
| Mole Flows | kmol/hr | 3912.356832 | 661.8297544 | 276 | 628.3459354 | 356.4989791 |
| CO ₂ | kmol/hr | 265.203606 | 0 | 0 | 0 | 0.37497292 |
| H ₂ | kmol/hr | 0.012593845 | 0 | 0 | 0 | 2.16122E-11 |
| H ₂ O | kmol/hr | 3077.64008 | 92.32920241 | 0 | 92.26638277 | 108.402864 |
| N ₂ | kmol/hr | 0 | 0 | 0 | 0 | 0 |
| O ₂ | kmol/hr | 0 | 0 | 0 | 0 | 0 |
| NTRIET | kmol/hr | 0 | 0 | 0 | 0 | 0 |
| TREA | kmol/hr | 295.0464022 | 295.0464022 | 0 | 9.17039E-21 | 0 |
| AMMON-01 | kmol/hr | 0 | 0 | 0 | 0 | 0 |
| HCOOH | kmol/hr | 274.4541498 | 274.4541498 | 0 | 260.0795527 | 247.7211422 |
| TREA-FO | kmol/hr | 0 | 0 | 0 | 0 | 0 |
| N-ISO-01 | kmol/hr | 0 | 0 | 276 | 276 | 6.2393E-224 |
| Mole Fractions | | | | | | |
| CO ₂ | | 0.06778615 | 0 | 0 | 0 | 0.00105182 |
| H ₂ | | 3.21899E-06 | 0 | 0 | 0 | 6.06236E-14 |
| H ₂ O | | 0.786646058 | 0.139505971 | 0 | 0.146840104 | 0.304076226 |
| N ₂ | | 0 | 0 | 0 | 0 | 0 |
| O ₂ | | 0 | 0 | 0 | 0 | 0 |
| NTRIET | | 0 | 0 | 0 | 0 | 0 |
| TREA | | 0.075413981 | 0.445804076 | 0 | 1.45945E-23 | 0 |

| | | | | | |
|------------------|-------------|-------------|---|-------------|-------------|
| AMMON-01 | 0 | 0 | 0 | 0 | 0 |
| HCOOH | 0.070150592 | 0.414689953 | 0 | 0.413911411 | 0.694871954 |
| TREA-FO | 0 | 0 | 0 | 0 | 0 |
| N-ISO-01 | 0 | 0 | 1 | 0.439248485 | 1.7502E-226 |
| Mass Fractions | | | | | |
| CO ₂ | 0.106567652 | 0 | 0 | 0 | 0.001234083 |
| H ₂ | 2.30028E-07 | 0 | 0 | 0 | 3.23312E-15 |
| H ₂ O | 0.505922074 | 0.0376909 | 0 | 0.034667493 | 0.145950336 |
| N ₂ | 0 | 0 | 0 | 0 | 0 |
| O ₂ | 0 | 0 | 0 | 0 | 0 |
| NTRIET | 0 | 0 | 0 | 0 | 0 |
| TREA | 0.272147902 | 0.675828684 | 0 | 1.93337E-23 | 0 |
| AMMON-01 | 0 | 0 | 0 | 0 | 0 |
| HCOOH | 0.115362142 | 0.286480416 | 0 | 0.249869277 | 0.852815581 |
| TREA-FO | 0 | 0 | 0 | 0 | 0 |
| N-ISO-01 | 0 | 0 | 1 | 0.71546323 | 5.7956E-226 |

Table S13. Heat and mass balance of Electrochemical reduction pathway

| Stream Name | Units | CO ₂ | water/KOH | R-100 to V-100 | O ₂ | Solvent |
|-------------|-------|-----------------|-----------|----------------|----------------|---------|
| Description | | | | | | |
| From | | | | R-100 | R-100 | |

| To | | R-100 | R-100 | V-100 | E-100 |
|------------------|---------|--------------|--------------|--------------|-------------|
| Stream Class | | CONVEN | CONVEN | CONVEN | CONVEN |
| Temperature | C | 25 | 25 | 25 | 25 |
| Pressure | bar | 1 | 1 | 1 | 81.00000278 |
| Enthalpy Flow | cal/sec | -7637021.468 | -155085860.9 | -165358205.7 | 2.31099E-11 |
| Average MW | | 44.0098 | 18.08527022 | 19.92982037 | 31.9988 |
| Mole Flows | kmol/hr | 292.3261806 | 8163.360481 | 8561.018596 | 187.004873 |
| H ₂ O | kmol/hr | 0 | 8148.360481 | 7780.472928 | 0 |
| CO ₂ | kmol/hr | 292.3261806 | 0 | 391.4228634 | 0 |
| HCOOH | kmol/hr | 0 | 0 | 260.6421236 | 0 |
| O ₂ | kmol/hr | 0 | 0 | 0 | 187.004873 |
| CO | kmol/hr | 0 | 0 | 6.122193882 | 0 |
| H ₂ | kmol/hr | 0 | 0 | 107.3584858 | 0 |
| KOH | kmol/hr | 0 | 15.00000049 | 15.00000049 | 0 |
| 2-MET-01 | kmol/hr | 0 | 0 | 0 | 1620 |
| Mole Fractions | | | | | |
| H ₂ O | | 0 | 0.998162521 | 0.908825608 | 0 |
| CO ₂ | | 1 | 0 | 0.045721529 | 0 |
| HCOOH | | 0 | 0 | 0.030445223 | 0 |
| O ₂ | | 0 | 0 | 0 | 1 |
| CO | | 0 | 0 | 0.000715124 | 0 |
| H ₂ | | 0 | 0 | 0.012540387 | 0 |
| KOH | | 0 | 0.001837479 | 0.001752128 | 0 |
| 2-MET-01 | | 0 | 0 | 0 | 1 |
| Mass Fractions | | | | | |
| H ₂ O | | 0 | 0.99429962 | 0.821520089 | 0 |
| CO ₂ | | 1 | 0 | 0.100964049 | 0 |
| HCOOH | | 0 | 0 | 0.070309821 | 0 |
| O ₂ | | 0 | 0 | 0 | 1 |

| | | | | | |
|----------------|---|------------|-------------|---|---|
| CO | 0 | 0 | 0.001005073 | 0 | 0 |
| H ₂ | 0 | 0 | 0.001268447 | 0 | 0 |
| KOH | 0 | 0.00570038 | 0.004932521 | 0 | 0 |
| 2-MET-01 | 0 | 0 | 0 | 0 | 1 |

| Stream Name | Units | V-100 to E-100 | E-100 to D-100 | D-100 to D-101 | D-101 |
|------------------|---------|----------------|----------------|----------------|------------------|
| Description | | | | | |
| From | | V-100 | E-100 | D-100 | D-101 |
| To | | E-100 | D-100 | D-101 | |
| Stream Class | | CONVEN | CONVEN | CONVEN | CONVEN |
| Temperature | C | 40 | 45.88817494 | 114.8619566 | 25 |
| Pressure | bar | 1 | 1.1 | 5.1 | 4 |
| Enthalpy Flow | cal/sec | -152321371.8 | -50598505.48 | -11326836.34 | - 8431871.724 |
| Average MW | | 19.03372859 | 58.65799418 | 32.8046018 | 41.49039323 |
| Mole Flows | kmol/hr | 7946.903461 | 2707.793959 | 514.4808522 | 324.1229369 |
| H ₂ O | kmol/hr | 7662.107075 | 926.5722468 | 248.322803 | 57.96489617 |
| CO ₂ | kmol/hr | 14.23911279 | 14.01372085 | 5.80619E-50 | 0 |
| HCOOH | kmol/hr | 255.3933502 | 255.1616312 | 250.916971 | 250.9169625 |
| O ₂ | kmol/hr | 0 | 0 | 0 | 0 |
| CO | kmol/hr | 0.037354884 | 0.023749338 | 4.43649E-64 | 0 |
| H ₂ | kmol/hr | 0.126567157 | 0.070387709 | 3.38629E-74 | 0 |
| KOH | kmol/hr | 15.00000049 | 14.99999804 | 15.24107824 | 15.24107824 |
| 2-MET-01 | kmol/hr | 0 | 1496.952225 | 3.3195E-28 | 0 |
| Mole Fractions | | | | | |
| H ₂ O | | 0.964162597 | 0.342187131 | 0.482666754 | 0.178836144 |
| CO ₂ | | 0.001791781 | 0.005175328 | 1.12855E-52 | 0 |
| HCOOH | | 0.032137467 | 0.094232292 | 0.487709056 | 0.774141333 |

| | | | | |
|------------------|-------------|-------------|-------------|-------------|
| O ₂ | 0 | 0 | 0 | 0 |
| CO | 4.70056E-06 | 8.77073E-06 | 8.62323E-67 | 0 |
| H ₂ | 1.59266E-05 | 2.59945E-05 | 6.58196E-77 | 0 |
| KOH | 0.001887528 | 0.005539564 | 0.029624189 | 0.047022523 |
| 2-MET-01 | 0 | 0.55283092 | 6.45214E-31 | 0 |
| Mass Fractions | | | | |
| H ₂ O | 0.912572598 | 0.105093893 | 0.26506576 | 0.077651306 |
| CO ₂ | 0.004142958 | 0.003882934 | 1.51404E-52 | 0 |
| HCOOH | 0.077711983 | 0.073938862 | 0.684268052 | 0.858762198 |
| O ₂ | 0 | 0 | 0 | 0 |
| CO | 6.91743E-06 | 4.18821E-06 | 7.36299E-67 | 0 |
| H ₂ | 1.6868E-06 | 8.93344E-07 | 4.04469E-78 | 0 |
| KOH | 0.005563857 | 0.005298524 | 0.050666188 | 0.063586497 |
| 2-MET-01 | 0 | 0.811780705 | 1.69411E-30 | 0 |

References

1. Hirata M, Suda S; Y. Kagaku Kogaku. Vapor Pressure of Methanol in High Pressure Regions. 1967;31:339–41.
2. Park K, Gunasekar GH, Kim SH, Park H, Kim S, Park K, et al. CO₂ hydrogenation to formic acid over heterogenized ruthenium catalysts using a fixed bed reactor with separation units. *Green Chemistry*. 2020 Mar 7;22(5):1639–49.
3. Glass M, Aigner M, Viell J, Jupke A, Mitsos A. Liquid-liquid equilibrium of 2-methyltetrahydrofuran/water over wide temperature range: Measurements and rigorous regression. *Fluid Phase Equilib.* 2017 Feb 15;433:212–25.
4. Zhu X, Shi Y, Li S, Cai N. Elevated temperature pressure swing adsorption process for reactive separation of CO/CO₂ in H₂-rich gas. *Int J Hydrogen Energy*. 2018 Jul 19;43(29):13305–17.
5. Laitinen AT, Parsana VM, Jauhainen O, Huotari M, Van Den Broeke LJP, De Jong W, et al. Liquid-Liquid Extraction of Formic Acid with 2-Methyltetrahydrofuran: Experiments, Process Modeling, and Economics. *Ind Eng Chem Res*. 2021 Apr 21;60(15):5588–99.
6. Jiang X, Lin L, Rong Y, Li R, Jiang Q, Yang Y, et al. Boosting CO₂ electroreduction to formate via bismuth oxide clusters. *Nano Res.* 2023 Oct 1;16(10):12050–7.
7. Fernández-Caso K, Díaz-Sainz G, Alvarez-Guerra M, Irabien A. Electroreduction of CO₂: Advances in the Continuous Production of Formic Acid and Formate. Vol. 8, ACS Energy Letters. American Chemical Society; 2023. p. 1992–2024.
8. Spurgeon JM, Kumar B. A comparative technoeconomic analysis of pathways for commercial electrochemical CO₂ reduction to liquid products. *Energy Environ Sci.* 2018 Jun 1;11(6):1536–51.
9. Li X, Anderson P, Jhong HRM, Paster M, Stubbins JF, Kenis PJA. Greenhouse Gas Emissions, Energy Efficiency, and Cost of Synthetic Fuel Production Using Electrochemical CO₂ Conversion and the Fischer-Tropsch Process. *Energy and Fuels*. 2016 Jul 21;30(7):5980–9.
10. Al-Qahtani A, Parkinson B, Hellgardt K, Shah N, Guillen-Gosalbez G. Uncovering the true cost of hydrogen production routes using life cycle monetisation. *Appl Energy*. 2021 Jan 1;281.
11. Yang H, Zhang C, Gao P, Wang H, Li X, Zhong L, et al. A review of the catalytic hydrogenation of carbon dioxide into value-added hydrocarbons. Vol. 7, Catalysis Science and Technology. Royal Society of Chemistry; 2017. p. 4580–98.
12. Atsbha TA, Yoon T, Yoo BH, Lee CJ. Techno-economic and environmental analysis for direct catalytic conversion of co2 to methanol and liquid/high-calorie-sng fuels. *Catalysts*. 2021 Jun 1;11(6).

13. Sigma-Aldrich. ruthenium(ii) catalyst. [accessed 2023 May 4]. Available from: [https://www.sigmaaldrich.com/KR/en/search/ruthenium\(ii\)](https://www.sigmaaldrich.com/KR/en/search/ruthenium(ii))
14. Sigma-Aldrich. Bismuth(III) oxide catalyst. [accessed 2023 May 4]. Available from: <https://www.sigmaaldrich.com/KR/en/product/aldrich/202827>
15. Sengupta D, Pike RW. Chemicals From Biomass – Integrating Bioprocesses into Chemical Production Complexes for Sustainable Development. 1st ed. Boca Raton, FL: CRC Press; 2013.
16. Turton Richard. Analysis, synthesis, and design of chemical processes. Prentice Hall; 2012. 1007 p.
17. Energy Agency I. Direct Air Capture: A key technology for net zero [Internet]. [accessed 2024 Jan 1]. Available from: www.iea.org/t&c/
18. Peters MS, Timmerhaus KD, West RE. Plant Design and Economics for Chemical Engineers. Vol. 35, Journal of Chemical Education. McGraw-Hill Companies, Inc., 1221 Avenue of the Americas, New York, NY 10020; 2003. A506 p.
19. Dominguez-Ramos A, Singh B, Zhang X, Hertwich EG, Irabien A. Global warming footprint of the electrochemical reduction of carbon dioxide to formate. J Clean Prod. 2015 Oct 1;104:148–55.
20. Thonemann N, Schulte A. Supporting Information: From laboratory to industrial scale: A prospective LCA for electrochemical reduction of CO₂ to formic acid.
21. Kang D, Byun J, Han J. Evaluating the environmental impacts of formic acid production from acid production from CO₂: Catalytic hydrogenation: Vs. electrocatalytic reduction. Green Chemistry. 2021 Dec 7;23(23):9470–8.
22. Banu A, Mir N, Ewis D, El-Naas MH, Amhamed AI, Bicer Y. Formic acid production through electrochemical reduction of CO₂: A life cycle assessment. Energy Conversion and Management: X. 2023 Oct 1;20.
23. Ahn Y, Byun J, Kim D, Kim BS, Lee CS, Han J. System-level analysis and life cycle assessment of CO₂ and fossil-based formic acid strategies. Green Chemistry. 2019;21(12):3442–55.
24. Aldaco R, Butnar I, Margallo M, Laso J, Rumayor M, Dominguez-Ramos A, et al. Bringing value to the chemical industry from capture, storage and use of CO₂: A dynamic LCA of formic acid production. Science of the Total Environment. 2019 May 1;663:738–53.
25. Xing Z, Hu X, Feng X. Tuning the Microenvironment in Gas-Diffusion Electrodes Enables High-Rate CO₂Electrolysis to Formate. ACS Energy Lett. 2021 May 14;6(5):1694–702.
26. Qian Y, Liu Y, Tang H, Lin BL. Highly efficient electroreduction of CO₂to formate by nanorod@2D nanosheets SnO. Journal of CO₂ Utilization. 2020 Dec 1;42.